Adsorption and reduction of hexavalent chromium on magnetic greigite (Fe₃S₄)-CTAB: leading role of Fe(II) and S(−II)

Yanxia Zhou,a Yiting Zhao,a Xiaoge Wua, Weiqin Yin,a Jianhua Hou,a,b Shengsen Wanga,ab Ke Fengab and Xiaozhi Wangc*abc

In this study, a facile one-step route was used to synthesize a novel magnetic mesoporous greigite (Fe₃S₄)-CTAB composite, which was utilized to remove hexavalent chromium (Cr(VI)). The optimized Fe₃S₄-CTAB₀.₇₅ composite with a CTAB dosage of 0.75 g possessed the maximum specific surface, showing the highest Cr(VI) adsorption capacity of 330.03 mg g⁻¹. The mechanism analysis revealed that Fe(II) and S(−II) were critical for the reduction of Cr(VI). CTAB can promote the removal of Cr(VI) by Fe₃S₄-CTAB composites, possibly due to increased S(−II) concentration, better dispersion of nanoparticles, and greater zeta potential. Besides, there is mild effect of Fe(0) on Cr(VI) removal, which is confirmed by the disappearance of the Fe(0) peak from the XPS analysis. The pseudo-second-order kinetic model could explain the Cr(VI) removal processes well. The adsorption of Cr(VI) at different initial concentrations was more consistent with a Langmuir isotherm. The existence of H⁺ was beneficial for Cr(VI) removal by Fe₃S₄-CTAB₀.₇₅. Our work confirmed that the obtained Fe₃S₄-CTAB₀.₇₅ composites exhibit considerable potential for Cr(VI) removal from aqueous solution.

1. Introduction

As a metal contaminant, chromium frequently occurs in wastewater. Since it could seriously harm human health, the U.S. Environmental Protection Agency (EPA) has listed hexavalent chromium as a group 'A' human carcinogen.¹ Cr(III) and Cr(VI) are the most common and stable forms of chromium,²,³ and the poisonousness of the latter is 100 times larger than the former. It has been reported that long exposure to Cr(VI) could induce many diseases, for example liver failure, reproductive failure, chronic headaches, respiratory disease, nose bleeds and any type of cancer.⁴

A lot of technologies have been utilized to remove Cr(Ⅵ) from the environment, such as photoreduction,⁶,⁷ electro-chemical precipitation,⁸ ion exchange,⁹ and adsorption.¹⁰ Combining with adsorption, chemical reduction from Cr(Ⅵ) to Cr(Ⅲ) to reduce toxicity is one of the most simple and economical methods.¹¹ The combined strategy includes many advantages, such as high efficiency, easy use, and low cost.¹²

Greigite (Fe₃S₄) is isostuctural with Fe₃O₄.¹³⁻¹⁵ Due to its low cost and good chemical reduction ability, Fe₃S₄ has been applied on heavy metal removal.¹⁶ For example, Kong and his coworkers investigated lead (Pb²⁺) transformations on β-cyclo-dextrin stabilized magnetic Fe₃S₄ nanoparticles, and found that the surface adsorption and chemical precipitation (PbS) were the dominant mechanisms in the process of removing lead.¹⁷ Mahamudur Islam et al. utilized greigite-conducting poly-pyrrole nanocomposite to remove arsenate and arsenite from aqueous solution.¹⁸ However, the application of Fe₃S₄ in the adsorption and reduction of hexavalent chromium is rarely investigated. According to the previous reports,¹⁹ Fe₃S₄ can release Fe(Ⅱ) and S(−Ⅱ) in aqueous solution, which can promote Cr(VI) reduction effectively.

Various cationic surfactants have been applied to removal anionic metal.¹⁹ Appropriate surfactants or stabilizers such as β-CD, PEG, and CTAB, could affect the purity and stability of Fe₃S₄.²⁰ Cetyltrimethylammonium bromide (CTAB) is one of the common cationic surfactants and it is known that the dissociated cetyltrimethylammonium (CTA⁺) maintains the properties of surfactant. CTAB could help to obtain high purity greigite microcrystals,²¹ improve the content of S(−Ⅱ) in greigite, and CTA⁺ could combine with Cr(Ⅵ) anions and form ion-pairs.²² This is the first work about the application of Fe₃S₄-CTAB to removal Cr(Ⅵ).

The task of this study was to prepare greigite (Fe₃S₄) doped with CTAB for the removal of Cr(VI). The detailed works were listed as following: (1) synthesize and characterize greigite...
doped with CTAB, (2) determine the effects of solution pH, reaction time, CTAB doping amount, and initial concentrations of contaminant on the Cr(VI) removal effectiveness, (3) investigate the reductive and adsorptive capacity of Cr(VI). This work could provide evidence to explain Cr(VI) removal mechanisms by Fe3S4-CTAB.

2. Materials and methods

2.1 Chemical reagents

K2Cr2O7 was dissolved to prepare Cr(VI) stock solutions. Ferric chloride hexahydrate (FeCl3·6H2O), thiourea, ethylene glycol (EG), ethanol, potassium dichromate (K2Cr2O7), cetyltrimethylammonium bromide (CTAB) were purchased from Yangzhou Chemicals Corporation (Yangzhou, China). Solution pH was regulated by 1 M HCl and/or NaOH. Deionized water was used in the whole experiment. All reagents used were of analytical grade and had not been further purified before use.

2.2 Synthesis of greigite and greigite-CTAB composite

Fe3S4 composite was synthesized by the modified hydrothermal method. Briefly, FeCl3·6H2O [3.0 mmol] and thiourea [6.0 mmol] were added to EG [60 mL] and stirred for 20 min to form a carmine solution. Typically, CTAB (0.75 g) was then added to the carmine solution, followed by 15 min magnetic stirring and 15 min sonication. After the mixing, the above solution was transferred into an autoclave (the capacity is 100 mL). After 12 h heating at 180 °C, a magnet was used to collect solid products. The collected samples were repeatedly washed with distilled water followed by ethanol. The obtained Fe7/3Fe5/3FeS4-CTAB0.75 (referred to as Fe3S4-CTAB0.75 in the article) were dried in a vacuum oven at 60 °C for 12 h. Following the similar fabricated method, other Fe3S4-CTAB samples were prepared with varying amounts of CTAB and named as Fe3S4-CTAB0.25, Fe3S4-CTAB0.50, and Fe3S4-CTAB1.00.

2.3 Batch experiments of Cr(VI) removal

In order to determine the best pH in the reaction process, the pH value of the Cr(VI) solutions was set to 2.0, 3.0, 4.0, 5.0 and 6.0, respectively, with 1 M HCl and 1 M NaOH. Typically, 0.75 g of CTAB was added to the carmine solution, followed by 15 min magnetic stirring and 15 min sonication. The suspension was filtered through a 0.22 mm membrane to achieve clarified samples. Adsorption isotherms were performed by mixing the composite with Cr(VI) solution (initial concentration ranging from 100 to 300 mg L−1) and shaking for 5 h to obtain adsorptive equilibrium.

2.4 Analysis

Total Cr(VI) in solution was confirmed through the 1,5-diphenylcarbazide colorimetric method (with potassium permanganate). Moreover, the 1,5-diphenylcarbazide colorimetric method was applied to measure the Cr(VI) concentrations at 540 nm through a UV/Vis spectrophotometer. The Cr(VI) concentration corresponded to the difference between total Cr and Cr(VI). The Cr speciation of the reaction product was confirmed by XPS. The o-phenanthroline method was utilized to determine the concentration of ferrous ions at 510 nm. Through adding hydroxylamine hydrochloride to reduce Fe from Fe(III) to Fe(II), then the total dissolved iron was measured.

2.5 Characterization

XRD analysis (D8 Advance Bruker AXS, Germany) was applied to determine the phase structure of the obtained composites. HRTEM was obtained using a Tecnai G2 F30 S-TWIN (FEI, USA) to ensure the crystalline lattice of the reaction product. Scanning electron microscope (S-4800II, Japan) was used to observe the surface morphology of the obtained composite. Fourier transform infrared (FTIR) spectra were obtained by using a Thermo Nicolet iS5 spectrometer, using KBr as a reference. The magnetization of the obtained composites was obtained on a vibrating sample magnetometry (VSM-EV7, ADE). Micromeritics ASAP 2460 was applied to confirmed nitrogen adsorption–desorption isotherms and specific surface areas. The valence state of related elements of the adsorbent before and after reaction was identified by XPS (ESCALAB 250 Xi, USA). Zeta potentials of the obtained composites were measured through zeta potentiometer (DTS1060).

3. Results and discussion

3.1 Characterization of greigite and greigite-CTAB composite

The crystal structure and phase purity of the obtained composites and Cr(VI)-treated Fe3S4-CTAB0.75 were determined by XRD pattern. From Fig. 1, FeS shows three major diffraction peaks at 17.61°, 38.99° and 49.59°, which correspond to [001], [111] and [200] crystalline planes of FeS. Other diffraction peaks can correspond well to cubic Fe3S4 (JCPDS file no. 89-1998). No other impurity peaks were observed between bare Fe3S4 and CTAB modified Fe3S4 composite, indicating that the existence of CTAB did not change the structure of Fe3S4. Meanwhile, the product from the reaction of Fe3S4-CTAB0.75 with Cr(VI) was shown in Fig. 1. The XRD pattern showed that the diffraction peaks of FeS disappeared and a new diffraction peak attributed to Cr2O3 emerged.
HRTEM (Fig. S1†) was applied to determine the composition of Cr(VI)-treated Fe3S4-CTAB0.75. The d-spacing in HRTEM was about 0.25 nm and 0.334 nm in accordance with the (110) faces of Cr2O3 and Cr(OH)3, respectively, which demonstrated the existence of Cr2O3 and Cr(OH)3.

SEM images under high magnification (Fig. S2†) clearly displayed that a large amount of nanosheets formed 3D flower-shaped microspheres, which agreed with previous reports.† The FTIR spectra of the obtained composites and pure CTAB were presented in Fig. S3.† Interestingly, the peaks at 2400 cm⁻¹ and 2300 cm⁻¹ became more obvious with the increase of CTAB doping amount. These coincided with the spectra of pure CTAB between 2400 cm⁻¹ and 2300 cm⁻¹ indicating Fe3S4 crystals doped with CTAB has been successfully fabricated.

The magnetization of obtained composites were gained under the condition of an applied magnetic field. The magnetic hysteresis loop of the obtained Fe3S4, Fe3S4-CTAB0.25, Fe3S4-CTAB0.50, Fe3S4-CTAB0.75 and Fe3S4-CTAB1.00 were shown in Fig. 2, which indicates that all the five composites exhibit ferromagnetic nature. The saturation magnetization of greigite.

Table 1  Surface area, average pore size and pore volume of Fe3S4, Fe3S4-CTAB0.25, Fe3S4-CTAB0.50, Fe3S4-CTAB0.75 and Fe3S4-CTAB1.00 composites

| Sample                  | $S_{BET}$ (m² g⁻¹) | Average pore size (nm) | Pore volume (cm³ g⁻¹) |
|-------------------------|---------------------|------------------------|-----------------------|
| Fe3S4                   | 18.82               | 15.035                 | 0.071                 |
| Fe3S4-CTAB0.25          | 15.13               | 6.934                  | 0.026                 |
| Fe3S4-CTAB0.50          | 16.69               | 15.799                 | 0.066                 |
| Fe3S4-CTAB0.75          | 21.27               | 14.895                 | 0.079                 |
| Fe3S4-CTAB1.00          | 20.95               | 6.116                  | 0.032                 |

The nitrogen adsorption–desorption method was applied to determine the specific surface area of the obtained composites. A type IV isotherm shown in Fig. S4† displayed the presence of mesopores structure. Table 1 summarizes the BET surface area of the obtained composites with varied doping amount of CTAB.

But when we increase CTAB dosage, it acts as a mesopore directing agent, leading to the reduction of the pore size and specific surface areas.²⁷ But when we increase CTAB dosage, it acts as a mesopore directing agent, which will enter into the material structure, thus increasing the specific surface area of the material.²⁹

XPS spectroscopy was used to examine the valence change of related elements before and after treatment of Cr(VI) (Fig. 3). For Fe3S4-CTAB0.75, the Fe 2p spectra in Fig. 3[a] was decomposed to five peaks. The binding energies of Fe(II) were at 707.6 (ref. 30) and 710.3 eV.²⁴,³¹ The Fe peak at 2p3/2 = 720.2 eV which can be assigned to Fe0.²¹,³³ The binding energies of Fe 2p3/2 and Fe 2p1/2 are at 713.3 eV and 724.4 eV, respectively, proving of the existence of Fe(III).³²,³³ The S 2p spectra in Fig. 3(b) exhibited different peaks at binding energies of 530 eV and 531 eV, consisting with O²⁻ (ref. 30) and C=O,²⁴ respectively. Furthermore, the proportion of the reduced sulfur in Fe3S4-CTAB0.75 was increased and the diffraction peak –OH was converted into C-O after CTAB doping (Fig. S5†). These changes further proved that CTAB had been successfully doped into greigite.

Upon Cr(VI) sorption, the reductive state of iron(Fe(III)) and Fe0 decreased from 59.84% for Fe3S4-CTAB0.75 to 49.85% for Cr(VI)-spent Fe3S4-CTAB0.75, whereas Fe(II) increased from 40.16% to 50.15% (Table S1†). Furthermore, the peak assigned to Fe0 disappeared, while a new peak ascribed to Fe(III) appeared at 718.9 eV.²² It displays the oxidation of Fe(II) and Fe0 by Cr(VI) under the condition of experiment. The percentage of S(−II) decreased from 47.36% to 24.25%, and S(II) increased from 11.13% to 27.57% after contacting with Cr(VI) (Table S1†), revealing the oxidation of S(−II) by Cr(VI). The peak assigned to

![Fig. 2 Hysteresis (M–H) analysis of Fe3S4, Fe3S4-CTAB0.25, Fe3S4-CTAB0.50, Fe3S4-CTAB0.75 and Fe3S4-CTAB1.00.](image)

![Fig. 3 XPS spectra of (a) the Fe 2p region, (b) the S 2p region, (c) the O 1s region and (d) the Cr 2p region of the fresh and Cr(VI)-treated Fe3S4-CTAB0.75 sample.](image)
S(–I) vanished, indicating it could react with Fe(III) to form Fe(II) and release H⁺ (FeS₂ + 14Fe(III) + 8H₂O → 15Fe(II) + 2SO₄²⁻ + 16H⁺). Simultaneously, a new peak ascribed to S(VI) (n > 2) resulting from the oxidation of FeS and/or FeS₂ (ref. 39) appeared at 163.5 eV. For O 1s, C–O disappeared, the binding energy at 531.6 eV suggests the existence of SO₄²⁻/SO₃²⁻, which is consistent with the increased proportion of S(VI). Additionally, a new peak assigned to Cr₂O₃ emerged at 530.8 eV (ref. 31) suggesting the existence of SO₃²⁻ and reduction of Cr(VI).

3.2 Effect of pH on Cr(Ⅵ) removal

Solution pH is a critical parameter in the process of adsorption and reduction of Cr(Ⅵ). As shown in Fig. 4(a), the removal efficiency of Cr(Ⅵ) in the solution was 93.73% and 22.67%, respectively, corresponding to the pH value of 2.0 and 6.0. Under the pH range from 2.0 to 4.0, the removal efficiency of Cr(Ⅵ) showed a significant drop, which is consistent with previous experimental results. This trend coincided with the marked decrease of concentration of dissolved Fe(Ⅲ) from Fe₃S₄-CTAB₀.₇₅ (Fig. 4(b)). The result may suggest facilitated dissolution of FeS⁷ at lower pHs played important roles in Cr(Ⅵ) removal. Besides, the isoelectric point of the material is between pH 2.0 and 3.0, and thus material becomes positively charged and protonated at pH 2.0 which is favourable for electrostatic attraction between HCrO₄⁻. This was also further confirmed with positive zeta potential at pH 2.0 (Fig. S6f). However, higher pH values will change the surface charge of the material, impairing the electrostatic attraction between the material and Cr(Ⅵ) anions. Furthermore, hydroxyls from the higher pH solution could compete the adsorption sites on Fe₃S₄-CTAB₀.₇₅ surface with Cr(Ⅵ) anions, leading to a reduction of the removal efficiency of Cr(Ⅵ). However, under the pH range from 4.0 to 6.0, the removal efficiency slightly increased, which may be due to the formation of a small amount of Cr(OH)₃. In addition, coexistence of H⁺ at lower pH and reduced sulphur was able to reduce Cr(Ⅵ) to Cr(Ⅲ) more easily. Based on the pH-dependent sorption of Fe₃S₄-CTAB, we choose the pH value at 2.0 for further removal studies.

3.3 Effects of CTAB dosage on Cr(Ⅵ) removal kinetics

Fig. 5(a) shows the adsorption data of Cr(Ⅵ) at different time intervals. The initial concentration of Cr(Ⅵ) was set to 100 mg L⁻¹ and the amount of Fe₃S₄-CTAB₂ which added was 0.3 g L⁻¹. It can be seen that adsorption reached equilibrium after 30 min reaction. The removal effectiveness of Fe₃S₄-CTAB₂ is affected by the amount of CTAB doping. Cr(Ⅵ) removal efficiency was about 70% and 94% for Fe₃S₄ and Fe₃S₄-CTAB₀.₇₅ after 60 min, respectively. The enhanced removal of Cr(Ⅵ) by CTAB composites might be attributed to several reasons. First, the higher reduced sulfur in the composites as revealed by the XPS analysis (Fig. S5(a)) facilitated its reduction reaction with Cr(Ⅵ). This could be a very important mechanism associated with our composites. Second, the CTAB improved the dispersion and reduced the aggregation of Fe₃S₄ to Cr(Ⅵ) (Fig. S2(a) & (d)). Third, the cationic surfactant CTAB could increase electrostatic attraction with negatively charged HCrO₄⁻ on one hand. On the other hand, CTAB could increase zeta potential of Fe₃S₄ compared to pristine Fe₃S₄ which provide more positive surface for attraction with HCrO₄⁻ at
Table 2  Adsorption kinetics parameters for the adsorption of Cr(VI)

| Sample                 | Pseudo-first-order model | Pseudo-second-order model |
|------------------------|--------------------------|---------------------------|
|                        | $K_1$ (min$^{-1}$)       | $Q_e$ (mg g$^{-1}$)       | $R^2$     | $K_2 \times 10^{-4}$ (g (mg min$^{-1}$)) | $Q_e$ (mg g$^{-1}$) | $R^2$     |
| Fe$_3$S$_4$             | 0.117                    | 204.925                   | 0.883     | 7.023                                   | 235.849            | 0.998     |
| Fe$_3$S$_4$-CTAB$_{0.25}$ | 0.092                    | 253.764                   | 0.941     | 5.302                                   | 281.690            | 0.999     |
| Fe$_3$S$_4$-CTAB$_{0.50}$ | 0.105                    | 280.186                   | 0.908     | 4.659                                   | 319.489            | 0.998     |
| Fe$_3$S$_4$-CTAB$_{0.75}$ | 0.094                    | 304.922                   | 0.977     | 3.881                                   | 346.021            | 0.999     |
| Fe$_3$S$_4$-CTAB$_{1.00}$ | 0.113                    | 232.121                   | 0.924     | 4.491                                   | 280.899            | 0.997     |

As shown in Table 2, the pseudo-second-order adsorption model was more suitable than the pseudo-first-order adsorption model to explain Cr(VI) removal from correlation coefficients. Thus, the difference in the removal rate by obtained composites is expressed by the pseudo-first-order reaction model (Fig. 5(b)). The kinetic parameters calculated from the pseudo-first reaction model are listed in Table S2,† where reaction rate of Fe$_3$S$_4$-CTAB$_{0.75}$ is 3–4 times greater than pristine Fe$_3$S$_4$.

To gain a better explanation of transformation of chromium species, the total chromium concentrations were examined at different sampling time points (shown in Fig. 6). The total chromium concentration dropped sharply within 5 min, and then gradually reached equilibrium within 30 min. The residual total chromium concentration dropped to 40 mg L$^{-1}$ at 60 min, about 85% of the chromium was Cr(III), which is less toxic than Cr(VI). Therefore, it is suggested that Cr(VI) could be reduced to Cr(III) effectively by Fe$_3$S$_4$-CTAB$_{0.75}$.

3.4 Adsorption isotherm

The adsorption isotherm of Cr(VI) removal was discussed by changing the initial concentrations of contaminant, with the Fe$_3$S$_4$-CTAB$_{0.75}$ dosage of 0.3 g L$^{-1}$ (initial pH = 2.0). The related model parameters were shown in Table S1.† The adsorption capacity of Fe$_3$S$_4$-CTAB$_{0.75}$ reached to 303.53 mg g$^{-1}$ and 329.95 mg g$^{-1}$ with initial Cr(VI) concentrations of 100 and 300 mg L$^{-1}$, respectively. The Freundlich model could not fit the experimental data well (Fig. S7†), compared with the Langmuir model ($R^2 = 0.99$) (Fig. 7).

According to the fitting results, it could be speculated that the adsorption of Cr(VI) by Fe$_3$S$_4$-CTAB$_{0.75}$ was mononuclear layer adsorption. The calculated $Q_{\text{max}}$ value (the adsorption capacity) was 330.03 mg g$^{-1}$. Fe$_3$S$_4$-CTAB$_{0.75}$ exhibited a excellent removal good capacity for Cr(VI), compared with other Fe-based adsorbents (Table 3).

3.5 Proposed Cr(VI) removal mechanism

Based on the above discussion, a possible mechanism for Cr(VI) removal on Fe$_3$S$_4$-CTAB$_{0.75}$ was preliminarily proposed as follows. The mesoporous Fe$_3$S$_4$-CTAB$_{0.75}$ composite could adsorb Cr(VI) rapidly as CTA$^+$ cations can combine the HCrO$_4^-$ (main species of hexavalent chromium at pH 2.0 (ref. 41)) to form ion-pairs (eqn (1)). Moreover Fe$^0$ could be easily oxidized by Cr(VI). The absorbed Cr(VI) is in situ reduced by Fe$^0$ to generate Cr(III), accompanied by the oxidation of Fe$^0$ to form Fe(II), which is consistent with the disappearance of Fe$^0$ peak (Fig. 3(b)) (eqn (2)). Furthermore, dissolved FeS provides Fe(II) and S(−II) (eqn (3)) donating the electrons to Cr(VI), which was evidenced by the increase of SO$_4^{2-}$ content in Fig. 3(b) (eqn (4)). In addition, we measured the concentration of Fe(II) and Fe(III) in deionized water (pH = 2.0) and in the presence of Cr(VI) (pH = 2.0) at different time intervals, respectively. Without

![Fig. 6](image1.png) Concentrations of the total chromium, Cr(III) and Cr(VI) in the solution during the Cr(VI) transformation by Fe$_3$S$_4$-CTAB$_{0.75}$ remaining in the solution ($C_0 = 100$ mg L$^{-1}$, pH = 2.0, Fe$_3$S$_4$-CTAB$_{0.75}$ dosage = 0.3 g L$^{-1}$), the corresponding Cr(III) and Cr(VI).

![Fig. 7](image2.png) Langmuir isotherm of Cr(VI) adsorption. The reactions were developed at initial Cr(VI) concentrations of 100–300 mg L$^{-1}$ with 0.3 g L$^{-1}$ Fe$_3$S$_4$-CTAB$_{0.75}$, with pH$_0$ = 2.0.
When the doping content of CTAB was 0.75 g, the highest Cr(VI) adsorption capacity varies with the doping content of CTAB. The predicted maximum Cr(Ⅵ) removal capacity was achieved. Langmuir isotherm model matches better than Freundlich isotherm model in removal capacity. Fe3S4-CTAB0.75 was the optimum condition for Cr(VI) removal. FeS-coated iron (Fe/FeS) has an adsorption capacity of 62.1 mg g⁻¹ and a pH of 5.0, with an equilibrium time of 72 h. MnFe2O4@SiO2-CTAB and FeS coated with CTAB have adsorption capacities of 76.65 mg g⁻¹ and 330.03 mg g⁻¹, respectively, with pH values of 2.0 and 3.0, and equilibrium times of 12 h.

4. Conclusions

Novel greigite-CTAB composites applied to remove Cr(Ⅵ) from aqueous solutions were successfully synthesized by a simple one-step modification method. The effectiveness of greigite-CTAB composites varies with the doping content of CTAB. When the doping content of CTAB was 0.75 g, the highest Cr(Ⅵ) removal capacity was achieved. Langmuir isotherm model matches better than Freundlich isotherm model in fitting Cr(Ⅵ) adsorption data. The predicted maximum Cr(Ⅵ) removal capacity is 330.03 mg g⁻¹. The removal process is in accordance with pseudo-second-order kinetic model. Furthermore, low pH value of the system is crucial to the reduction of Cr(Ⅵ) to Cr(Ⅲ). It could be inferred that the release of Fe(Ⅱ) under acidic conditions is favorable. In the case of Fe3S4-CTAB0.75, Cr(Ⅵ) was adsorbed on the adsorbent surface through electrostatic attraction followed by reduced to less toxic Cr(Ⅲ). Therefore, the novel greigite-CTAB composites have great potential for the efficient removal of Cr(Ⅵ) in wastewater.

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Table 3  Comparison of the maximum Cr(Ⅵ) adsorption capacity of various adsorbents

| Adsorbents                | S_BET (m² g⁻¹) | q_m (mg g⁻¹) | pH  | Equilibrium time (h) | Ref. |
|---------------------------|----------------|--------------|-----|----------------------|------|
| FeS@Fe₄                | 53.11          | 66.7         | 5.0 | 2                    | 24   |
| FeS-coated iron (Fe/FeS) | 62.1           | 69.7         | 5.0 | 72                   | 38   |
| CMC-FeS@biochar         | 51.5           | 130.5        | 5.5 | 72                   | 31   |
| FeS                     | —              | 240          | 5   | 4.5                  | 50   |
| MnFe3O4@SiO2-CTAB       | 53.4           | 25.044       | 3.0 | 12                   | 22   |
| Fe₃O₄ capped with CTAB  | —              | 10.05        | 4.0 | 12                   | 19   |
| FeS₃S₄-CTAB0.75         | 21.273         | 330.03       | 2.0 | 1                    | This study |

Cr(Ⅵ), dominant iron species in acidic solution (pH = 2.0) is in the Fe(Ⅱ) state, while after 60 min reaction in the presence of Cr(Ⅵ), iron exists mainly in trivalent iron (Fig. S8). Thus, it is suggested that Cr(Ⅵ) is reduced by ferrous ions dissolved from Fe3S4-CTAB0.75 (eqn (5)).

C16H33(CH3)N+ + HCrO4⁻ → C16H33(CH3)N⁺···HCrO4⁻ (1)

3Fe⁰ + 2HCrO4⁻ + 6H₂O → 3Fe²⁺ + 3Cr(OH)₃ + 8OH⁻ (2)

FeS + H⁺ → Fe²⁺ + HS⁻ (3)

8HS⁻ + 3HCrO4⁻ + 29H⁺ → 3SO₄²⁻ + 8Cr³⁺ + 20H₂O (4)

3Fe²⁺ + HCrO₄⁻ + 7H⁺ → 3Fe³⁺ + Cr³⁺ + 4H₂O (5)

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