High Efficiency Dye-sensitized Solar Cells Constructed with Composites of TiO$_2$ and the Hot-bubbling Synthesized Ultra-Small SnO$_2$ Nanocrystals

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An efficient photo-anode for the dye-sensitized solar cells (DSSCs) should have features of high loading of dye molecules, favorable band alignments and good efficiency in electron transport. Herein, the 3.4 nm-sized SnO$_2$ nanocrystals (NCs) of high crystallinity, synthesized via the hot-bubbling method, were incorporated with the commercial TiO$_2$ (P25) particles to fabricate the photo-anodes. The optimal percentage of the doped SnO$_2$ NCs was found at ~7.5% (SnO$_2$/TiO$_2$ \( w/w \)), and the fabricated DSSC delivers a power conversion efficiency up to 6.7%, which is 1.52 times of the P25 based DSSCs. The ultra-small SnO$_2$ NCs offer three benefits, (1) the incorporation of SnO$_2$ NCs enlarges surface areas of the photo-anode films, and higher dye-loading amounts were achieved; (2) the high charge mobility provided by SnO$_2$ was confirmed to accelerate the electron transport, and the photo-electron recombination was suppressed by the highly-crystallized NCs; (3) the conduction band minimum (CBM) of the SnO$_2$ NCs was uplifted due to the quantum size effects, and this was found to alleviate the decrement in the open-circuit voltage. This work highlights great contributions of the SnO$_2$ NCs to the improvement of the photovoltaic performances in the DSSCs.

Dye-sensitized solar cells (DSSCs) based on semiconductor electrodes are of great interest as alternatives to the conventional silicon based solar cells in view of the ease of fabrication, cost-effectiveness and environmental benignity$^{1-3}$. An ideal photo-anode for DSSCs should combine features of high specific surface areas, fast electron transport and less interfacial electron recombination$^{3,4}$. Intensive work has been devoted to the fabrication of TiO$_2$ photo-anodes$^{5,6}$. However, the TiO$_2$ based anodes suffer from sluggish electron mobility and high density of surface states which gave rise to the charge recombination. To explore materials for more efficient photo-anodes, efforts have been made to utilize metal oxides such as ZnO, SnO$_2$, Nb$_2$O$_5^{7-12}$, and bi-functional materials including ZnO/TiO$_2$, SnO$_2$/TiO$_2$, ZnO/SnO$_2$, SrTiO$_3$/TiO$_2$, et al.$^{13-18}$. Particularly, SnO$_2$ has attracted great attention due to the two remarkable advantages: (1) SnO$_2$ possesses a high electron mobility (100–500 cm$^2$ V$^{-1}$ S$^{-1}$), two orders of magnitude higher than that of TiO$_2$ (0.1–10 cm$^2$ V$^{-1}$ S$^{-1}$), which would give rise to improved charge transfer$^{19}$; (2) Compared to TiO$_2$, SnO$_2$ has a larger band gap of 3.6 eV and a more negative conduction band minimum (CBM = −4.56 V vs. vacuum), which would facilitate the electron injection from the sensitizer to the semiconductor electrodes$^{20}$. However, the efficiency of SnO$_2$ based DSSCs is still low up to date$^{21}$. The low open-circuit voltage (\( V_{OC} \)) was thought as reasoned by the more negative position of CBM$^{22}$. The sluggish photo-to-electron efficiency (\( \eta \)) was mainly caused by the charge recombination which is usually trapped by the surface states$^{23}$. After a thorough survey of literatures, we found that almost all of the reported photo-anodes based on SnO$_2$ have barely considered the quantum-size effects. Nevertheless, this could possibly alleviate the decrement in \( V_{OC} \)$.$^{24,25}$ Herein, we propose our strategy to satisfy the following two requirements: (1) to improve the \( V_{OC} \), the level of CBM or the flat band potential (\( V_{fb} \)) of the

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Figure 1. (a) TEM and HRTEM (inset) images of the colloidal SnO2 NCs synthesized via the hot-bubbling method; (b) XRD pattern of the colloidal SnO2 NCs; (c,d) The TEM (c) and HRTEM (d) images of the air-annealed SnO2 NCs (T = 450 °C, t = 2.0 hrs).

Results and Discussion

Structural characterization of the SnO2 NCs. For a brief introduction of the hot-bubbling synthesis, the Sn-oleate was prepared from precursors of oleic acid and the newly prepared SnOx·xH2O in 1-octadecene (ODE) solutions at temperatures of 280–320 °C. When a flow of room-temperature air was bubbled into these hot solutions, an explosive nucleation in form of SnO2 clusters occurs in the O2-induced hydrolysis reactions31. The high diffusion rate of the gas benefits the fast nucleation. The air flow cools down the micro-environment of the SnO2 clusters, which would avert the Oswald ripening growth 34. Therefore, it is necessary to find strategies to synthesize highly crystalline SnO2 plus uniform size-distributed nanocrystals (NCs). As suggested in our previous report31, the hot-bubbling synthesis, which was conducted by bubbling air into surfactant solutions dissolved with Sn-precursors, yields SnO2 NCs of high crystallinity and desired sizes31,32. Because the exciton Bohr radius (~3.0 nm) of SnO2 is much smaller than other semiconductors, the size effects of SnO2 are not easily observed33. In this research, we utilize this hot-bubbling method to get ultra-small SnO2 NCs (average size 3.4 nm), and incorporate these NCs in the TiO2 photo-anode to construct the DSSCs. Due to the high temperature performed in the synthesis and the fast diffusion rate provided by the gas reactants, homogeneous nucleation was achieved. The colloidal SnO2 NCs were found not to be aggregated during the annealing processes because of the well-protection by the ligand molecules. The SnO2 NCs were found not only to enlarge the surface area of the semiconductor anodes, but also to facilitate the charge transfer across the DSSCs. The SnO2/TiO2 composite films are demonstrated to be efficient anodes to boost the photovoltaic performances owning to the increased dye loading, the facilitated electron injection and the efficient charge collection.
the diameter ranges from 3.0 to 4.0 nm. The average size (3.4 nm) of the annealed particles was evaluated by
the Scherrer equation. We thus named the hot-bubbling synthesized sample as the 3.4 nm-SnO2 NCs. The HRTEM
image is shown in Fig. 1d. Well-resolved lattice fringes were found, and this suggests the high-crystallinity of the
SnO2 particles. The distances of the fringes are estimated to be 0.142 and 0.330 nm, which can be well indexed to
the d-spacing correlated with the (301) and (101) of the tetragonal SnO2. In contrast, the commonly used hydro-
thermal synthesis yields SnO2 particles of a spherical shape, and the average diameter is about 20 nm. We name
this sample as S20 for a control study.

Morphological studies on the fabricated SnO2/TiO2 films. The above annealed SnO2 were mixed
with the commercial TiO2 particles (P25, D = ~25 nm) to fabricate photo-anodes for the DSSCs. Figure 2a–c
show schematic diagrams for the SnO2/TiO2 photo-anodes prepared by using different ratios of SnO2/TiO2. The
possible pathways followed by the electron transfer are also demonstrated, which is obtained from the following
analyses. Percentages of SnO2/TiO2 = 0%, 7.5% and ∞% (wt.) are denoted, respectively in Fig. 2a–c. We named
these films as S0, S7.5 and S∞. The data in the sample name means the percentage of SnO2 doped with respect to
TiO2. For example, the sample of S0 denotes the anode films were made with pure TiO2 particles while the film
S∞ comprised totally with the SnO2 NCs. The SEM images for films S0, S7.5 and S∞ were shown in Fig. 2d–f. The
porous structures were clearly found in these anodes. Specifically the composite film S7.5 showed more porosity
with tiny holes, and these holes (shown as circles) are uniformly distributed in the film (Fig. 2e). Although the
SEM image can only give views in a micrometre sized area, the smaller but uniform distribution can be further
confirmed by the N2 adsorption-desorption characterizations. Herein, it is easy to understand that the mixed
particles of different size levels (D_{SnO2} = 3.4 nm, D_{TiO2} = 25 nm) provide more chances for particles approaching,
and usually assemble into networks in the photo-anodes. Thus we might deduce that paths for electrons trans-
port change from point-to-point in the pure TiO2 films to point-to-surface or surface-to-surface in the composite
SnO2/TiO2 films. Closer contact of the semiconductor oxides is expected to accelerate the charge transfer in the
photo-anodes. The TEM images (Fig. 2g) illustrates that the SnO2 NCs are well dispersed in the composite films.

**Figure 2.** (a–f) Schematic presentations and SEM images of the SnO2/TiO2 photo-anodes prepared with
the percentage of 0% (a,d), 7.5% (b,e) and ∞% (c,f) SnO2; (g,h) TEM images for S7.5 at different magnifications.
Characterizations on the SnO$_2$/TiO$_2$ films. The incorporation of SnO$_2$ NCs in the composite photo-anodes were further studied by means of XRD, FT-IR, N$_2$ adsorption-desorption and the electrical resistivity measurements. Figure 3a shows the XRD profiles of the fabricated films in which different ratios of SnO$_2$ were incorporated (S$_0$ $\rightarrow$ S$_{12.5}$). The strong diffraction peaks can be assigned to the anatase TiO$_2$ (JCPDS No. 84–1285, labelled as ‘A’). We also found weaker peaks (110), (101) and (211) attributing to the cassiterite type SnO$_2$ in films of S$_{2.5}$–S$_{12.5}$ (labelled as ‘C’)$^{37}$. Figure 3b displays the FT-IR spectra of the composite films with three typical ratios of SnO$_2$/TiO$_2$ (S$_0$, S$_{7.5}$ and S$_{12.5}$). The broad bands in range of 400–800 cm$^{-1}$ are the stretching and bending modes of Ti-O-Ti$^{38}$. The band at 1063 cm$^{-1}$ observed in sample S$_{7.5}$ and S$_{12.5}$ can be attributed to the stretching vibrations of Sn-O-Sn demonstrating the presence of SnO$_2$.$^{38}$ The strongest intensity of this band in sample S$_{12.5}$ suggests uniform incorporation of SnO$_2$. Figure 3c shows the N$_2$ adsorption-desorption isotherms (type II hysteresis loop), and the derived data such as surface area, pore volume and size distribution are summarized in Table 1$^{16,39,40}$. The Barrett-Joyner-Halenda (BJH) pore size distribution is plotted as an inset in Fig. 3c. We found a peak representing smaller sizes in the either pure SnO$_2$ or composites of TiO$_2$ and SnO$_2$, yet the pure TiO$_2$ has no such pores. This means the ultra-small SnO$_2$ NCs produce smaller pores as well as enlarge the surface area of the composite. Or we can say the incorporation of SnO$_2$ NCs would cause rougher surface of the TiO$_2$ films. This would surely benefit a higher loading of dye molecules (N$_{719}$). To study the saturated amount of N$_{719}$ that are adsorbed by the composite, we recorded the UV-vis absorption spectra of the N$_{719}$ solutions. The solutions were obtained by desorbing from a pre-fabricated photo-anode with the aid of KOH. The concentration of N$_{719}$ was evaluated by recording the absorbance at a particular wavelength ($\lambda$ = 500 nm)$^{27}$. The data were also summarized in Table 1 (the 3rd column). The sample S$_{7.5}$ has the highest amount of N$_{719}$. It was indicated that the adsorption capacity of dye molecules are related to the surface areas and pore volumes. As expected the incorporation of SnO$_2$ NCs was found to increase the loading amount of N$_{719}$ molecules. However, excessive incorporation of SnO$_2$ would block the pores, and less amount of N$_{719}$ molecules are immobilized (see samples S$_{12.5}$ or higher ratios of SnO$_2$/TiO$_2$). Therefore, a certain amount of SnO$_2$ should be incorporated in order to get a higher loading of N$_{719}$. We also studied the conductivity of the fabricated films or the sheet resistance, and these measurements were conducted by means of the 4-point probe method. The Hall-effects were also evaluated. The collected data (Table 2) reveals that the resistivity of the SnO$_2$/TiO$_2$ composite films decreases with incorporation of SnO$_2$ NCs, while the Hall mobility of charge carriers increases significantly. For example, the resistivity of the pure P25 film (S$_0$) is 4.57 $\Omega$ cm, and such a value decreases to 3.74 $\times$ 10$^{-2}$ $\Omega$ cm in sample S$_{7.5}$. The charge carrier concentration was found to increase enormously (3 ~ 5 orders of magnitude) after the incorporation of SnO$_2$. It is interesting to note that the carrier concentration is the highest (2.57 $\times$ 10$^{19}$ cm$^{-3}$) when the incorporation ratio of SnO$_2$ is 7.5%, which is even higher than the pure SnO$_2$. This suggests a synergistic effect might occur between TiO$_2$ and SnO$_2$, and this would provide advantages in enhancing the diffusion of charge carriers. On condition that measurements performed under the same lights and temperatures the mobility or concentration of charge carriers.
The Nyquist plots are listed in Table S1 (see the supporting information). The circuit (Fig. S1) has been used to fit the Nyquist plots. The circuit and the impedance parameters derived from S7.5 photo-anode could be related to the better Ohmic contact with SnO2 NCs at the percentage of 7.5%. For compared with the P25 film (129.4 SnO2/TiO2 composites, the characteristic of S0&S7.5). Detailed reasons for the synergistic effects are still to be exploited, but the special structure and the effects were not observable, and was over weighted by the brilliant conductivity afforded by the SnO2 (see Table 2, S0&S7.5). Detailed reasons for the synergistic effects are still to be exploited, but the special structure and the ultra-small size of the SnO2 NCs could have a great impact on these effects. This feature would surely benefit the photon-electron conversion when the film was used as photo-anodes of the DSSCs.

### Photovoltaic study on the DSSCs.

To study the photovoltaic properties of the DSSCs fabricated with the SnO2/TiO2 composites, the characteristic of J–V curves and the IPCE plots were recorded (Fig. 4a,b). Parameters such as V_{OC} short-circuit current density (J_{SC}), fill factor (FF), η and the maximum values of IPCE are summarized in Table 3. Although the incorporation of SnO2 caused a slight decrease in the V_{OC} (0.82 → 0.79 V for films S0 → S7.5), the J_{SC} was increased greatly from 8.2 to 15.4 mA cm^{-2}. The η value was found to be improved from 4.4% to 6.7% after the incorporation of SnO2 NCs. The highest η (6.7%) was found in case of film S7.5, and this value is even higher than the photo-anode constructed by the TiO2-coated SnO2 hollow microspheres (MHSs) as reported (η ≈ 5.56%) by Cao et al.\(^\text{22}\). This comparison was carefully done because one may doubt on the experimental errors that would be caused by factors such as thickness of films, electrolytes, amounts of N719, counter electrodes, different batches of samples, and so on. Herein, we found the η value (2.1%) in our sample S\(\infty\) (the pure SnO2 NCs) is equivalent to that (η = 1.4%) of MHSs by Cao et al.\(^\text{22}\). Therefore, we think the comparison is trustable because all our results were obtained in the same batch of experiments. For other comparison, we have fabricated DSSCs by using the hydrothermally synthesized SnO2 (S_20) as dopants for the photo-anodes. The ratio of SnO2/TiO2 was also set at 7.5% in this composite film. As shown in Fig. 4c, the values of V_{OC}, J_{SC} and η for the film S_20 were lower than those of fabricated with the hot-bubbling synthesized SnO2 NCs. Therefore, these superior performances illustrate that better photovoltaic properties can be achieved by the incorporation of SnO2 NCs.

To understand the interfacial reactions of photo-excited electrons and resistance across the DSSCs, we conducted the electrochemical impedance spectroscopic (EIS) measurements on the fabricated DSSCs in darkness. The bias voltage was set at the V_{OC}, and the frequency ranges from 0.1 Hz to 1 MHz. The Nyquist plots are presented in Fig. 4d. All plots exhibit double semicircles of which a small-radius semicircle was found in the high-frequency region (> 1 kHz, see the inset of Fig. 4d) and a larger one within 100–1.0 Hz. An equivalent circuit (Fig. S1) has been used to fit the Nyquist plots. The circuit and the impedance parameters derived from the Nyquist plots are listed in Table S1 (see the supporting information). The R_s representing the Ohmic serial resistance, can be read directly from the onset of the first semicircle\(^\text{41}\). The derived resistance R_s corresponds to the charge transfer resistance across the counter-electrode/electrolyte interface (smaller semicircle), and another part R_i is ascribed to the resistance between the oxide/electrolyte interface and the photo-anode film (larger semicircle)\(^\text{34,35}\). As can be found, all the R_i are of the same magnitude order, i.e., values range between 10.8 and 22.3 Ω cm\(^2\). In accordance with the photovoltaic performances of the S12.5 which is inferior to the S7.5 fabricated DSSCs, the R_i of the former was found larger than the later. A possible explanation for the low R_i in the S7.5 photo-anode could be related to the better Ohmic contact with SnO2 NCs at the percentage of 7.5%. For the charge transfer resistance, we got smaller R_i from the SnO2/TiO2 composite (e.g., 25.3 Ω cm\(^2\) for S7.5) as compared with the P25 film (129.4 Ω cm\(^2\)), but the R_i is equivalent. The smaller R_i implies that the incorporation of SnO2 accelerates the charge transfer from the electrolyte to the photo-anode. The accelerated transfer can be resorted to the improvement in the full access of the electrolyte and the better charge mobility provide by the

### Table 1. The surface density of adsorbed dye molecules N719 and results derived from the N\(\infty\) adsorption-desorption isotherms of the SnO2/TiO2 films. \(\text{a}^\text{The total pore volume was evaluated for a P/P_0 ratio of 0.99.}\)

| Film | Weight ratio of SnO2/TiO2 (%) | Adsorbed N719 (10\(^2\) mol cm\(^{-2}\)) | Surface Area (m\(^2\) g\(^{-1}\)) | Pore Volume\(\text{b}\) (cm\(^3\) g\(^{-1}\)) | Average Pore Size (nm) |
|------|-----------------------------|-----------------------------|-----------------------------|-----------------------------|-----------------------------|
| S0   | 0.0                         | 1.70                        | 45.4                        | 0.22                        | 28.4                        |
| S7.5 | 7.5                         | 1.90                        | 53.0                        | 0.27                        | 20.5                        |
| S12.5| 12.5                        | 1.57                        | 53.6                        | 0.25                        | 18.7                        |
| S\(\infty\) | \(-\)                | \(-\)                      | 0.71                        | 64.8                        | 0.06                        |

### Table 2. Resistivity, carrier concentrations and Hall mobility derived from the 4-point probe resistivity/ Hall-effect measurements.

| Film | Resistivity (Ω cm) | Hall coefficient (cm\(^2\) C\(^{-1}\)) | Carrier concentration (cm\(^{-3}\)) | Hall mobility (cm\(^2\) V\(^{-1}\) s\(^{-1}\)) | P/N type |
|------|-------------------|----------------------------------------|-------------------------------------|--------------------------------------|----------|
| S0   | 4.57              | 9.1 × 10\(^{-1}\)                     | 6.86 × 10\(^{14}\)                 | 6.48                                 | N        |
| S7.5 | 3.74 × 10\(^{-2}\) | 2.43 × 10\(^{-1}\)                   | 2.57 × 10\(^{13}\)                 | 1.98 × 10\(^{7}\)                   | N        |
| S12.5| 5.87 × 10\(^{-2}\) | 6.99 × 10\(^{-1}\)                   | 8.93 × 10\(^{13}\)                 | 1.19 × 10\(^{7}\)                   | N        |
| S\(\infty\) | 7.14 × 10\(^{-2}\) | 3.57 × 10\(^{-1}\)                   | 1.75 × 10\(^{13}\)                 | 4.99 × 10\(^{7}\)                   | N        |
| S_20 | 4.71 × 10\(^{-3}\) | 1.81 × 10\(^{-1}\)                   | 3.45 × 10\(^{14}\)                 | 3.83 × 10\(^{7}\)                   | N        |
SnO2 NCs. Probably the special structures and the larger surface area can also be additional reasons. However, more incorporation of SnO2 NCs was found not to benefit the photovoltaic performance due to the increment of interfacial recombinations20. The equivalent $R_1$ (see Table S1, S0–S∞) means that the resistance of charge transfer at the counter electrodes (Pt) are the same. Results from the EIS measurements were found to be consistent with the 4-points resistivity characterizations, which further validates the incorporation of SnO2 being favourable to improve the photovoltaic performances. The optimum percentage of SnO2 was found to be 7.5%, and the corresponding cell yields the highest $\eta$ up to 6.7%.

The electron transport and recombination at the interface of photo-anode were further studied by techniques of the intensity-modulated photovoltage and photocurrent spectroscopies (IMVS and IMPS)44–46. Figure 5a shows the IMPS plots recorded under illumination of 470 nm. The electron transport time $\tau_d$, which demonstrates the average time intervals from the generation to collection of electrons, can be derived from the equation $\tau_d = 1/(2\pi f_{\text{min}})$, where $f_{\text{min}}$ is the frequency of the minimum point in the IMPS semicircle. The $\tau_d$ for films S0, S5, S7.5, S12.5 and $S\infty$ were evaluated to be 3.99, 1.26, 1.12, 1.00 and 317.47 ms. It was found that the $\tau_d$ was greatly reduced when the SnO2 NCs were incorporated, which means the shorter time period for the photo-electrons to reach the FTO substrate. The recombination time constant ($\tau_r$) for photo-electrons and ions of I3− or other redox couples in the electrolyte can be derived from the IMVS plots. The $\tau_r$ was obtained according to the equation $\tau_r = 1/(2\pi f_{\text{max}})$36,37, where $f_{\text{max}}$ is the frequency of the maximum point in the IMVS semicircle. As listed in Table 4, the $\tau_r$ for the SnO2/TiO2 composite film becomes longer due to the SnO2 NCs, which indicates that less recombination occurs at the interface of oxide/electrolyte on the basis that the same thickness for the photo-anodes are used. The charge collection efficiency ($\eta_{\text{cc}}$) was obtained from the relation of $\eta_{\text{cc}} = 1 - \tau_d/\tau_r^{37}$. The maximum $\eta_{\text{cc}}$ (99.7%) was obtained.

**Figure 4.** The $J$-$V$ curves (a) and IPCE plots (b) of the DSSCs constructed with the SnO2/TiO2 photo-anodes (film S0, S2.5, S5, S7.5, S10, S12.5 and S$\infty$); (c) The $J$-$V$ curves of DSSCs constructed with P25 + S_20 and film S7.5; (d) The Nyquist plots of DSCs constructed with photo-anode films of S0, S7.5, S12.5, and S$\infty$.

| Film | $V_{oc}$ (V) | $J_{sc}$ (mA cm$^{-2}$) | FF | $\eta$ (%) | IPCE (%) |
|------|--------------|-------------------|----|------------|---------|
| S0   | 0.82         | 8.34              | 0.64 | 4.4        | 42.5    |
| S2.5 | 0.80         | 10.60             | 0.49 | 4.5        | 61.9    |
| S5   | 0.80         | 12.79             | 0.57 | 6.0        | 59.4    |
| S7.5 | 0.79         | 14.53             | 0.58 | 6.7        | 61.2    |
| S10  | 0.78         | 13.01             | 0.56 | 6.6        | 57.5    |
| S12.5| 0.77         | 12.85             | 0.48 | 4.8        | 57.9    |
| S$\infty$ | 0.72     | 6.80              | 0.53 | 2.1        | 14.1    |
| S_20 | 0.72         | 12.45             | 0.51 | 4.8        | 56.4    |

**Table 3.** Photovoltaic parameters of DSSCs constructed with different photo-anodes (S0–S$\infty$, and S_20).
in the film of S12.5, and data is only a slightly higher than S7.5 ($\eta_{cc} = 98.0$). Thus the maximum $\eta$ was found in film of S7.5. Base on the above analyses, we deduced that more percentages of SnO$_2$ would accelerate the electron transport, and alleviate the interface charge recombination. However, the $V_{OC}$ decreases notably if more SnO$_2$ NCs were use, which would result in reduction of the photo-to-electron conversion. Therefore, the optimal doping amount of SnO$_2$ NCs should be set at a certain amount.

Discussions
It is generally accepted that the decrement in the $V_{OC}$ was mainly determined by the two factors$^{1-3}$: (1) the conduction band shifts to more positive values because of the SnO$_2$ incorporation; (2) the photo-electrons recombine with the redox couples in the electrolytes. Herein, we conducted experiments to verify these two factors and have tried to avoid the disadvantages. The first disadvantage can be avoided by decreasing the size of SnO$_2$. As is known, the band gaps can be enlarged by decreasing the size of a crystal. Due to the effective mass of electrons in SnO$_2$ is much lighter than those of holes ($m_e^* = 0.275 m_e$, $m_h^* = 10 m_e^*$, $m_e = 9.11 \times 10^{-31}$ kg), the CBM is raised prominently$^{28,38}$, which uplifts the Fermi level of the semiconductor electrode$^{41}$. The position of CBM in bulk SnO$_2$ is located at $-4.56$ V vs. vacuum$^{42}$. According to the semi-empirical pseudo-potential method (SEPM) calculations$^{43}$, the CBM value is uplifted to $-4.34$ V vs. vacuum when the sized of SnO$_2$ NCs is 3.4 nm. The position lies between the CBM of TiO$_2$ and the FTO (InSnO$_x$ particles) electrode, which would enhance the $V_{OC}$ and facilitate the electron transfer as well as the photo-electron injection. For a better understanding, we present a schematic diagram for the electron band alignments at the interface (Fig. 6). Based on this indication, the $V_{OC}$ of the DSSCs will be enlarged as compared to the bulk SnO$_2$ or micro-sized SnO$_2$$^{27}$. The second disadvantage is undertaken by reducing the trapping states. Herein, we managed to remove the trapping states by improving the crystallinity of the SnO$_2$, and this was realized by utilizing a high-temperature synthesis. Due to the ultra-small SnO$_2$ inserted among the P25 particles, full access of the semiconductor particles was realized$^{16,27}$. This special structures of networks would surely provide efficient charge transfer pathways because of the high mobility in SnO$_2$ NCs. The better conductivity of the composite films were confirmed by the 4-points conductivity and the EIS measurements. The diffusion lengths of electron ($L_n = d(\tau_r/2.35\tau_d)^{1/2}$ is the film thickness) were also calculated, which represents the average travel distance of electrons$^{44}$. The estimated values of $L_n$ as listed in Table 4 showed that the electron diffusion lengths were greatly lengthened due to the presence of SnO$_2$. Moreover, the finding of more dye molecules adsorbed on the composite film would provide more chances for photo-electron generation, which has been testified by enhancements in the IPCE values (Fig. 4b). Therefore, the incorporation of ultra-small SnO$_2$ NCs has advantages in fabricating DSSCs. The hot-bubbling synthesis can fulfil this task in preparing such ultra-small particles.

In summary, we have demonstrated the construction of highly efficient photoconversion DSSCs by using the hot-bubbling synthesized SnO$_2$ NCs. The SnO$_2$ NCs are of ultra-small sizes ($\sim$3.4 nm). Due to the quantum size effects the CBM of SnO$_2$ is uplifted, which alleviates the decrement in $V_{OC}$. The SnO$_2$ NCs are of high crystallinity, and this would help to suppress the interfacial recombination of the photo-electrons. The composite film of SnO$_2$/TiO$_2$ exhibits a higher photo-current density and photo-conversion efficiency as compared to the hydrothermally

### Table 4. Photovoltaic parameters of the electron transport time ($\tau_d$), the recombination time ($\tau_r$) and the collection efficiency derived from measurements of IMPS and IMVS.

| Film | $\tau_r$ (ms) | $\tau_d$ (ms) | $\eta_{cc}$ (%) | $L_n$ (μm) |
|------|---------------|---------------|-----------------|-------------|
| S0   | 31.7          | 3.99          | 87.4            | 14.7        |
| S5   | 50.4          | 1.26          | 97.5            | 33.0        |
| S7.5 | 50.4          | 1.12          | 98.0            | 35.0        |
| S12.5| 318.3         | 1.00          | 99.7            | 93.1        |
| S∞   | 1004.2        | 317.47        | 68.4            | 9.28        |

Figure 5. (a) The IMPS complex plane plots and (b) The IMVS complex plane plots of the composite photo-electrode for the films of S0, S5, S7.5, S12.5 and S∞.
synthesized SnO2/TiO2 photo-anodes. The higher mobility and concentration of the charge carriers in SnO2 were confirmed to improve the current density of the DSSCs. Compared to the pure TiO2 (P25) photo-anode, the transport time (τd) for electrons injection to the FTO substrates was greatly reduced. The optimal percentage of SnO2 NCs was found at 7.5%, and the fabricated DSSC delivers a higher η of 6.7%, which is 1.52 times as that as that of pure TiO2 based photo-anode.

Methods

Materials and chemicals. The TiO2 (P25) (Degussa product with a mean size ~25 nm and a BET surface area of 45.4 m2/g) particles were used in this research. The tin (IV) chloride pentahydrate (SnCl4·5H2O, A.R.), oleylamine (OLA, 80–90% C18 content), oleic acid (OA), 1-octadecene (ODE, 90%), stannous sulfate (SnSO4), sodium citrate (Na3C6H5O7·2H2O), were obtained from Sigma Aldrich. The dye sensitizer—cis-bis(isothiocyanato) bis(2,20-bipyridyl-4,40-dicarboxylato) ruthenium (bis-tertrabutylammonium) (N719) was purchased from Solaronix SA, Switzerland.

Synthesis of ~3.4 nm SnO2 NCs and the ~20 nm SnO2. For a typical hot-bubbling reaction, the synthesis of colloidal SnO2 NCs were performed in a three-neck flask linked with the Schlenk line31. Reagents including new-prepared SnO·xH2O (1.0 mmol, x = 1.0), OA (5.2 mL, 4.0 mmol), OLA (2.0 mL) and ODE (20 mL) were loaded in a three-neck flask. The synthesis of ~3.4 nm SnO2 NCs was undertaken as reported31. Before the synthesis all the volatile substances were removed by vacuum distillation (~0.01 bar) at 100 °C in order to purify the solvent. Under atmosphere of N2 the mixture solution was then heated up to 220 °C until a clear and colourless solution was obtained. A flow of air was bubbled at temperatures of 280~320 °C in order to get uniform sized NCs. The air was bubbled through a glass delivery tube. The tube has multiple pinholes (D = 0.5 mm) at one end which was exposed to the hot solutions. Samples were purified by precipitation employing toluene as solvent and iso-propanol/methanol (1.0, v/v) as non-solvent. The obtained ~3.4 nm SnO2 samples were dried in vacuum at 60 °C.

For a control experiment we also synthesized the SnO2 particles (~20 nm) via a hydrothermal pathway27. Briefly, the ligand Na3C6H5O7·2H2O (4.412 g, 15.0 mmol) dissolved in a solution of 10.0 ml ethanol and 90.0 ml deionized water was mixed with SnSO4 (1.076 g, 5.0 mmol). After homogenization the dispersion was transferred to a Teflon-lined autoclave. The dispersion was maintained at 180 °C for 12 hrs. The product in form of light yellow precipitates was collected by centrifugation, washed with distilled water/ethanol. The SnO2 microspheres were obtained after drying in vacuum at 70 °C for 24 hrs. The powders were annealed in air at 450 °C for 2 hrs.

DSSCs fabrication. To fabricate the photo-anodes of the DSSCs, pastes of TiO2 (P25) and SnO2 were prepared. Into an agate mortar 1.0 g P25 powders and a certain amount of SnO2 NCs were loaded, and the mixture was grinded for ~0.5 hrs before the addition of acetic acid (99.7%, 100 μl), deionized water (50 μl) and ethanol (50 ml). The suspension was further grinded for another 0.5 hrs, and was then transferred to a flask for sonication (~1.0 hrs). A mixture of terpineol (3.5 g) and ethyl cellulose (0.5 g) was added, and the final homogeneous paste was prepared by the repeated procedures of magnetic stirring and sonication. The excessive ethanol was removed by rotary-evaporator in a round-bottomed flask until a percentage of ethanol (~15-20% wt.) was left. In this experiment different ratios of SnO2 were added to prepared the composites of SnO2/TiO2, i.e., percentages of SnO2/TiO2 = 0%, 2.5%, 5%, 7.5%, 10%, 12.5% and ∞% were achieved in the paste. The sample of ∞% denotes the pure SnO2 NCs. These samples were labelled as S0, S2.5, S5, S7.5, S10, S12.5 and S∞. For comparison, the S_20 nm was also employed to prepare a paste in a similar procedure.

Films of photo-anodes were prepared by applying pastes of S0–S∞ and S_20 on an electric conducting glass plate. The fluorine-doped tin oxide (FTO, Geao Co., Wuhan, China) glass plates were used as the electric conductive substrates. A layer of mixture film (0.5 × 0.5 cm2) was fabricated on a FTO glass via the method of screen-printing. The film was then dried at an oven at 125 °C for 6.0 min. Subsequently the film was annealed at

Figure 6. Schematic diagrams of the band alignments proposed at interfaces of oxide/dye/electrolyte for the pure TiO2 (left panel) and TiO2/SnO2 photo-anodes (right panel).
500 °C for ~30 min to generate the porous structure. The porous films in thickness of ~8.0 μm were further coated with another layer of TiO₂ by dipping in a 0.05 M TiCl₄ aqueous solution for 30 min. After drying, the plate was sintered in air for 30 min at 500 °C. The electrodes were sensitized with dye molecules by immersing in a 50 mM N719 solution of acetonitrile-butyl alcohol (v:v = 1:1). The adsorption time is 24 hrs, and the temperature is room temperature (~25 °C). Finally, the electrodes were washed with ethanol and dried in vacuum (~0.01 mbar). The DSSCs were fabricated with a sensitized photo-anode, a platinized (Pt) counter electrode and the electrolyte. The DSSCs were sealed with a hot-molten gasket. The electrolyte consists of lithium iodide (LiI, 0.045 M), iodide (I₂, 0.032 M), 4-ter-butylpyridine (TBP, 0.5 M), guanidinium thiocyanate (0.1 M), 1-butyl-3-methylimidazolium iodide (BMII, 0.6 M) and acetonitrile/valeronitrile (85/15 vol %), which is similar to our previous reports.

Characterizations. Phases of materials were determined by the X-ray diffraction (XRD, Rigaku Co., Japan) using Cu Kα (λ = 0.15418 nm) as the irradiation. The morphology and microstructures were studied by the field emission scanning electron microscopy (FESEM, SU8020) and transmission electron microscopy (TEM, JEM-2100F). For the Fourier transform infrared (FT-IR) analysis, the spectrum was collected on the BRUKER TENSOR 27 instrument. The surface area and pore size distribution were obtained from a MICROMERITICS-ASAP 2010 unit, and the sample was activated in the N₂ atmosphere (T = 100 °C). The surface area and pore size distribution were further studied by the field emission scanning electron microscopy (FESEM, SU8020) and transmission electron microscopy (TEM, JEM-2100F). The Fourier transform infrared (FT-IR) analysis, the spectrum was collected on the BRUKER TENSOR 27 instrument. The surface area and pore size distribution were obtained from a MICROMERITICS-ASAP 2010 unit, and the sample was activated in the N₂ atmosphere (T = 100 °C). The surface area and pore size distribution were further studied by the field emission scanning electron microscopy (FESEM, SU8020) and transmission electron microscopy (TEM, JEM-2100F). The incident photon-to-current conversion efficiency (IPCE) plotted as a function of illumination wavelength, were recorded on a QTest Station 1000 ADI system (Crowntech, Inc.) equipped with a 300 W Xe lamp. The measurements were carried out by using a monochromator, assisted by an automatic filter wheel. Electrochemical impedance spectroscopic (EIS) measurements were recorded on Autolab320N electrochemical workstation. The frequency range explored was from 100 kHz to 1 Hz at a set potential of 0.78 V. The dynamic measurements of IMVS and IMPS were also recorded on the same potentiostat but linked with an intensity modulated blue LED light (470 nm).

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Author Contributions
S.M., J.X. and R.Z. conceived and coordinated the research. X.M., S.Z., L.D. and S.Q contributed to the synthesis, structural and electrochemical characterization of the materials and DSSCs. L.W. contributed to the theoretical calculation. The manuscript was primarily written by S.M. and R.Z. and revised by J.X. and Z.C. All authors contributed to discussions and manuscript review.

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