Growth and characterization of thin epitaxial Co$_3$O$_4$(111) films

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Abstract

The growth and characterization of epitaxial Co$_3$O$_4$(111) films grown by oxygen plasma-assisted molecular beam epitaxy on single crystalline α-Al$_2$O$_3$(0001) is reported. The Co$_3$O$_4$(111) grows single crystalline with the epitaxial relation Co$_3$O$_4$(111)[121]||α-Al$_2$O$_3$(0001)[1010], as determined from in situ electron diffraction. Film stoichiometry is confirmed by x-ray photoelectron spectroscopy, while ex situ x-ray diffraction measurements show that the Co$_3$O$_4$ films are fully relaxed. Post-growth annealing induces significant modifications in the film morphology, including a sharper Co$_3$O$_4$/α-Al$_2$O$_3$ interface and improved surface crystallinity, as shown by x-ray reflectometry, atomic force microscopy and electron diffraction measurements. Despite being polar, the surface of both as-grown and annealed Co$_3$O$_4$(111) films are (1 × 1), which can be explained in terms of inversion in the surface spinel structure.

Key words: Co$_3$O$_4$, spinel, interface structure, polar surfaces, oxide film growth, molecular beam epitaxy

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1. Introduction

Recent developments in the growth of high quality epitaxial thin metal oxide films have led to a renewed interest in the properties of metal oxides as one or more physical dimensions is reduced to the nanoscale. As the contribution of the interface becomes a significant part of the whole system, new physical phenomena are expected to emerge due to symmetry-breaking and the ensuing changes in electronic structure. Surface states and perturbed orbital states of the interface atoms often give rise to strongly anisotropic behavior and novel surface effects. Critical to this effort is the ability to grow and characterize high quality epitaxial thin films. It is in this context that we report here a detailed study of the growth and structural and electronic characterization of epitaxial [111]-oriented Co$_3$O$_4$ thin films.

Of the three known cobalt oxides, the mixed valence compound, Co$^{2+}$Co$^{3+}$O$_4$, is stable at ambient pressure and temperature and crystallizes in the cubic spinel structure (with lattice constant $a = 8.086$ Å [1]), while the high temperature CoO phase crystallizes in the rock salt structure. Also reported in the literature is a sesquioxide of cobalt, Co$_2$O$_3$, crystallizing in the corundum structure [2][3] (ab initio calculations suggest this phase to be a stable energy minimum [4]). One aspect of particular interest in compounds with strong ionic character is the effect of surface charge on the stability of polar surfaces and interfaces. This occurs along crystal directions where an electric dipole moment is present (arising from alternating charged crystal planes), where a divergent electrostatic energy would develop in clean, bulk-terminated crystals. One general mechanism for quenching such an increase in electrostatic self energy is via charge compensation, whereby a modification in the surface charge distribution cancels the overall electric dipole [5][6][7]. Charge compensation is bound to result in important modifications of the surface atomic and electronic structure, including changes in the valence state of surface ions, surface reconstructions, surface roughening and faceting, among others [5][6][7][8][9][10]. In the case of the spinel structure, all low index surfaces are polar, and we expect charge compensation processes to modify the atomic and electronic surface structure of spinel crystals. Recently, we have shown that Co$_3$O$_4$(110) thin films grown by molecular beam epitaxy (MBE) exhibit (1 × 1) surfaces, and we attributed the stability of this surface structure to modified cationic valence states at the surface, a process equivalent to an inversion in the spinel structure at the film surface [11]. Motivated by these findings, we study here the (111) surface of Co$_3$O$_4$ grown on α-Al$_2$O$_3$(0001) single crystals.

The crystal structure of Co$_3$O$_4$ along the [111] direction is particularly intricate: while the hexagonal primitive (oblique) cell in the (111) plane is relatively small, with a lattice constant of 5.72 Å (see Fig. [1]), the repeat unit along the [111] direction contains 18 atomic planes, in the form $[O_4-Co^{3+}_4-O_4-Co^{2+}_4-Co^{3+}_4-Co^{2+}_4]_3$, with four basic types of planes: one hexagonal oxygen plane, two octahedral Co$^{3+}$ planes and one tetrahedral Co$^{2+}$ plane. Along the [111] direction, the O sublattice in Co$_3$O$_4$ presents a...
face-centered cubic close-packing (fcc) stacking sequence, or A-B-C-A. All planes have non-zero charge per unit cell, and therefore all (111) planes are polar; since the (111) planes are composed of only O anions or Co cations, the total charge per unit area is large.

Very few studies have addressed the stability of the CoO₃(111) surface. Meyer et al. [12] have reported the growth of twinned (1 × 1) [111]-oriented Co₃O₄ films on Ir(001)-(1 × 1); from scanning tunneling microscopy and quantitative low energy electron diffraction (LEED) analysis, they conclude that the film surface is terminated at a Co²⁺-O plane, with a strong inward relaxation of the Co atoms to almost level with the O plane, and they suggest that a modified ionicity (inversion) of the surface cations leads to charge compensation and stabilization of the (1 × 1) Co₃O₄(111) surface. Other recent studies of the (111) surface of Co₃O₄ include that of Petitto et al. [13], where a detailed study of the interaction of the low index surfaces of Co₃O₄ to oxygen and water is reported, and the study by Tang et al. [14] on the reactivity of Co₃O₄ nanoparticles, which is found to be strongly reduced for particles terminated by (111) facets, as compared to irregularly shaped nanoparticles. Studies of the surface energies of Co₃O₄ could not be found in the literature. As a proxy system, we may consider the case of MgAl₂O₄, another prototypical spinel with a nearly identical lattice constant, which has been studied more often. Theoretically, atomistic calculations of the surface energy of MgAl₂O₄ for the low index planes indicate that charge compensation and surface stability can be achieved by surface vacancies; for the (110) and (111) surfaces, significant surface relaxations are predicted. For the (111) surface, the lowest energy termination is that which truncates the crystal at the O layer between the Al³⁺ and the Al³⁺-Mg²⁺-Al³⁺ layers, with nine O per unit cell on top of the Al layer, and seven O on the Mg layer at the opposite face of the crystal [15]. An earlier study predicted that the lowest energy (111) surface is that between the Mg²⁺ and Al³⁺ planes, with two vacant Al³⁺ cations per unit cell; the possibility of surface inversion was also considered and calculated to lower the surface energy [16]. While such inversion may be chemically difficult to achieve in MgAl₂O₄ during cleavage, it could be produced during crystal growth [7]. In the case of Co₃O₄, as noted, the mixed Co valency may make this process more easy to accommodate, even in bulk crystals, since it does not involve atomic diffusion. Cleavage of MgAl₂O₄ has been shown to occur preferentially along (001) planes, which also exhibit the lowest fracture surface energy [17, 15]. These results suggest that the (100) surface has the lowest energy, which is in disagreement with the results of the most recent theoretical calculations, which indicate the compensated (111) surface to be the most stable; this discrepancy between the fracture experimental results and the atomistic calculations has been attributed to the effect of water adsorption, which was found to reduce significantly the free energy of the (111) surface [7, 19, 20].

Figure 1: (a) Schematic diagram of the crystal structure of corundum (α-Al₂O₃), showing the primitive hexagonal (oblique) cell of the (0001) surface (shadowed area, top) and the stacking sequence along the [0001] direction in the hexagonal motive (bottom). Large circles stand for O atoms and small circles for Al: white small circles correspond to Al atoms lying in the center of the hexagonal motive, grey and dark circles to Al atoms lying at the vertices. White dots on the Al atoms designate positions where two atoms overlap along the direction perpendicular to the plane of the page. (b) Schematic diagram of the crystal structure of the spinel Co₃O₄ for the primitive oblique cell of the (111) surface in the hexagonal representation (shadowed area, top) and the respective atomic stacking along the [111] direction (bottom). Large circles represent O atoms, small circles correspond to Co (light circles are Co²⁺, tetrahedrally coordinated; darker circles are Co³⁺, octahedrally coordinated). To the left, in the bottom diagram, is shown the cationic arrangement of the (111) planes along the [111] direction. The oblique surface unit cell of the O sublattices is drawn in white (top).

Sapphire, α-Al₂O₃, crystallizes in the corundum structure and is the most stable aluminum oxide. α-Al₂O₃ is rhombohedral (with two formula units in the primitive cell), but it is more conveniently described in the hexagonal representation, with lattice parameters \( a = 4.7577 \) Å, \( c = 12.9877 \) Å [21, 22, 23] for the primitive (oblique) hexagonal cell, shown in Fig. 1, following Wyckoff’s convention for the crystal axes [24]. The hexagonal primitive cell contains six formula units, with a stacking sequence of eighteen layers, in the form \([Al-O₂-Al]₆\) along the c-axis. This results in three possible bulk terminations of the unit cell along the [0001] direction, consisting of planes exposing Al atoms for a truncation between two Al atomic planes, planes exposing two Al planes (Al-Al), and planes exposing O atoms, see Fig. 1. In the α-Al₂O₃ terminations yields non-polar surfaces. Experimentally, it is observed that (0001) surfaces prepared at tem-


peratures below ∼1500 K show a stable (1 × 1) surface, corresponding to the non-polar Al-termination; this surface is also characterized by large inward relaxation of the Al atoms [27, 28, 29, 30, 31, 32]. Annealing to higher temperatures produces a rotated (√31 × √31) surface reconstruction, associated with the Al-rich (Al-Al) surface [31, 32, 33, 34, 35, 36, 37, 38, 39, 40, 41]. Figure 1 shows the bulk terminated α-Al2O3(0001) surface exposing an Al-Al plane. The O sublattice is close to a hexagonal close packed (hcp) lattice (A-B-A stacking, with the Al atoms occupying 2/3 of the octahedrally coordinated interstitial sites), with an O inter-plane distance of 2.166 Å along the [0001] direction.

Here, we consider the surface and interface properties of [111]-oriented epitaxial films of the prototypical Co3O4 spinel grown on α-Al2O3(0001) substrates. Since the oxygen sublattices of both α-Al2O3 and Co3O4 are close packed (hexagonal and face-centered cubic, respectively) with close basal lattice constants, 2.746 Å for α-Al2O3 and 2.858 Å for Co3O4 (lattice mismatch of −3.9%), we may expect epitaxial growth of Co3O4 on α-Al2O3(0001) to proceed as a continuation of the O-sublattice. Indeed, we show that Co3O4(111) thin films can be grown epitaxially on α-Al2O3(0001) substrates by oxygen assisted molecular beam epitaxy. The as-grown Co3O4(111) surface exhibits a (1 × 1) surface diffraction pattern with no evidence of periodic reconstructions. The as-grown film surface shows a significant amount of disorder, which is slightly reduced upon annealing in air.

2. Sample growth and characterization techniques

The samples for this study were grown by O-assisted molecular beam epitaxy (MBE) in a ultrahigh vacuum (base pressure of 1 × 10−9 mbar). The substrates consist of polished α-Al2O3(0001) single crystal wafers, which were annealed at 870 K in ultrahigh vacuum for 60 min, and then cleaned under an O-plasma flux at 470 K for 30 min prior to film growth. No impurities other than trace amounts of Ca were detected by Auger electron spectroscopy (AES) taken after the substrate cleaning procedure. Low energy electron diffraction (LEED) and reflection high energy electron diffraction (RHEED) of the α-Al2O3(0001) substrate after cleaning display pattern characteristic of highly ordered surfaces (see Figs. 2 and 3), indicating a good crystalline order of the surface. The LEED and Laue diffraction patterns show 3-fold symmetric patterns (inner ring of the six diffraction spots in the LEED pattern show alternating intensity), suggesting that the surface is composed, predominantly, of double-layer atomic steps in the singlecrystalline α-Al2O3(0001) surface. The cobalt oxide films were grown by oxygen-assisted molecular beam epitaxy by simultaneous exposure of the substrate to a thermally evaporated Co atomic beam and an atomic O flux. The oxygen partial pressure during growth was 3 × 10−5 mbar, and the electron cyclotron resonance oxygen plasma source magnetron power was set to 175 W, yielding an atomic O flux of the order of 1 × 1014 cm−2 s−1 at the sample [32, 33]. The Co evaporation rate was ∼2 Å/min, estimated by means of a calibrated quartz crystal microbalance. Sample growth was monitored with RHEED, and film crystallinity and electronic structure were determined immediately after growth by RHEED, LEED and XPS, entirely in ultrahigh vacuum. After film growth, the samples were characterized ex situ using x-ray diffraction (XRD), x-ray reflectometry (XRR), and atomic force microscopy (AFM). Post-growth annealing was performed at 820 K for 14 h in air; this temperature and oxygen partial pressure favor the formation of Co3O4 over CoO [44, 45, 46]. For this study, two Co3O4 samples (22 and 38 nm thick) were grown independently, which were found to have similar spectroscopic, structural and morphological characteristics (due to sample charging, no LEED patterns could be obtained for the annealed 22 nm film and for the as-grown 38 nm film).

3. Results and discussion

The surface crystallinity of the films was probed during growth by RHEED. The RHEED pattern evolution showed a gradual fading of the α-Al2O3 sharp diffraction spots with increasing Co oxide thickness, became streaky at about 15 Å and finally broadened at about 30 Å. This indicates that film growth occurs via three-dimensional island growth. Typical RHEED patterns of the Co3O4 films after growth are shown in Fig. 2 where the in-plane crystal directions refer to those of the α-Al2O3 substrate, determined independently from Laue diffraction measurements. The RHEED patterns exhibited relatively broad diffraction spots, suggesting a relatively rough surface. The diffraction patterns are characteristic of a transmission pattern of the spinel {112} planes along the (1010) azimuths of the α-Al2O3 substrate (similar to the diffraction pattern of as-grown Co3O4(110) films along the [112] direction observed in [41], but rotated by 90°); and of the spinel [110] planes along the (1100) azimuths. In particular, we infer the epitaxial growth relation as Co3O4(111)[112]∥α-Al2O3(0001)[1010] and Co3O4(111)[011]∥α-Al2O3(0001)[1100]. The same epitaxial relationship is observed for MgAl2O4/α-Al2O3(0001) grown by solid state reactions [47, 48]. LEED patterns for the α-Al2O3(0001) substrate (Fig. 3) and for the as-grown Co3O4 film (not shown) exhibit a (1 × 1) diffraction pattern. Compared with the α-Al2O3 LEED patterns, the diffraction spots of the as-grown Co3O4 film are much broader, and the background is also more intense, indicating that the as-grown Co3O4 films have a significant amount of surface disorder (charging of the surface also contributes to the poor patterns, especially at lower electron beam energies). Motivated by our recent results that demonstrated a significant improvement in the bulk and surface structure of Co3O4/MgAl2O4(110) thin films upon annealing [41], we have also studied the effect
of post-growth annealing on the properties of Co$_3$O$_4$/α-Al$_2$O$_3$ (0001). RHEED patterns obtained after annealing (shown in Fig. 2 for the 38 nm film) show that annealing induces significant transformations in the film structure, as indicated by sharper and streakier RHEED diffraction patterns. However, the RHEED characteristics indicate that these surfaces are not atomically flat. LEED patterns for the annealed sample show well defined (1 × 1) patterns with sharp diffraction spots; although the patterns are not as sharp as for the α-Al$_2$O$_3$ substrate and while the background is more intense, it can be inferred that the Co$_3$O$_4$ film surface is well ordered. Unlike the 3-fold symmetric α-Al$_2$O$_3$ LEED patterns, the LEED diffraction patterns of the Co$_3$O$_4$ are 6-fold symmetric; this is usually associated with the presence of rotational twinning, i.e., the presence of both ABCA and ACBA stakings, which is known to occur in the growth of fcc (111) films on the (0001) planes of hexagonal crystals [49, 50]. Based on the geometrical configuration of our LEED system we are able to deduce the reciprocal space unit cell of the Co$_3$O$_4$(111) and α-Al$_2$O$_3$(0001) surfaces (Fig. 3 left) and that of the respective O sublattices (Fig. 3 right). It is readily seen that the unit cell of the α-Al$_2$O$_3$(0001) surface is larger (in reciprocal space) than that of Co$_3$O$_4$(111) and rotated by 30°, while the surface unit cell of the O sublattice are juxtaposed one on the other as expected from growth of Co$_3$O$_4$ as a continuation of the O sublattice. The relative spacing of the diffraction spots in the α-Al$_2$O$_3$ and Co$_3$O$_4$ RHEED patterns along the direction perpendicular to the electron beam also agree with this structural model: the spacing ratio is ≈ 0.7 along the ⟨1010⟩ azimuth and 2.0 along ⟨1100⟩, which correspond to $\sqrt{3}a_{\text{Co}_3\text{O}_4}/a_{\alpha-\text{Al}_2\text{O}_3} = 0.69$ and $a_{\text{Co}_3\text{O}_4}/\sqrt{3}a_{\alpha-\text{Al}_2\text{O}_3} = 2.08$, respectively.

Core level XPS measurements of the Co$_3$O$_4$ films after growth and after annealing were carried out to assess the film stoichiometry. The XPS spectra were obtained using the Mg Kα line ($h\nu = 1253.6$ eV) of a double anode x-ray source and a double pass cylinder mirror analyzer (PHI 15-255G) set at a pass energy of 25 eV (energy resolution of about 0.8 eV). XPS spectra of the O 1s and Co 2p lines of the as-grown and annealed films are shown in Fig. 4. Corrections to the data include a five-point adjacent smoothing, x-ray satellite correction and correction of energy shifts due to charging (aligned with respect to the Co 2p peaks, using the energy assignments given in [51, 52]). One observation is that the Co 2p spectra for both samples are very similar, showing that no significant changes in stoichiometry or in the ionic state of the Co cations occur as a consequence of annealing. A second observation is that the Co 2p spectra are characteristic of a Co$_3$O$_4$ ionic environment [51, 52], with strongly suppressed shake-up peaks compared to those of CoO [13, 16, 51, 52, 53, 54, 55, 56]. The O 1s photoemission line is also similar before and after annealing. The additional shoulder observed at higher binding energies has been attributed to adsorbed oxygen [13, 51, 57, 58, 59, 60, 61, 62] or to surface hydroxylation [13, 52]; for the annealed sample the shoulder is much more pronounced compared to the as-grown film, consistent with surface hydroxylation through water adsorption.

![Figure 2: Reflection high energy electron diffraction (RHEED) patterns of (a) the α-Al$_2$O$_3$(0001) substrate and Co$_3$O$_4$ films (b) before and (c) after annealing, along the ⟨1010⟩ and ⟨1100⟩ azimuths of the α-Al$_2$O$_3$ crystal (parallel to the electron beam, set at a grazing angle of incidence). Patterns (a) and (b) are from the 22 nm sample, (c) is from the 38 nm Co$_3$O$_4$ sample. The incident beam energy was set to 15 keV.](image1)

![Figure 3: Low energy electron diffraction (LEED) patterns of the α-Al$_2$O$_3$(0001) (top panel) and for the annealed 38 nm Co$_3$O$_4$ film (bottom panel) at three values of the incident beam energy. The in-plane α-Al$_2$O$_3$(0001) crystal orientation, inferred from Laue diffraction, is shown in the inset. The dashed hexagon indicates the symmetry of the LEED pattern; based on the geometrical configuration of our LEED system, the hexagon on the left images correspond to the respective unit cells of α-Al$_2$O$_3$ and Co$_3$O$_4$, while the larger hexagons on the right correspond to the hexagonal unit cell of the O sublattice, identical for both α-Al$_2$O$_3$ and Co$_3$O$_4$.](image2)
upon exposure to air.

The surface morphology was also probed ex situ using atomic force microscopy (AFM) in contact mode at room temperature. The AFM data were corrected and analyzed using the Gwyddion freeware package [63]. corrections to the data included planarisation, background-correction and removal of faulty scan lines. Typical AFM surface morphology profiles for the 38 nm sample are shown in Fig. 5 for $10 \times 10 \mu$m$^2$ and $2 \times 2 \mu$m$^2$ scanning areas. The surface of the α-Al$_2$O$_3$ substrate is atomically flat (average roughness of 0.2 nm), in agreement with the RHEED and LEED results. The surface profile of the as-grown Co$_3$O$_4$ film reveals a pronounced surface texture (rms roughness parameter of 2.4 nm), with a characteristic in-plane correlation length of about 80 nm determined from the $10 \times 10 \mu$m scans. Although regular in-plane and square in shape, these islands are not flat, in agreement with the spotty LEED and transmission-like RHEED patterns. Annealing leads to a significant change in the surface morphology, including a meandering island shape and a significant flattening of the islands, resulting in very uniform contrast in the higher magnification images. Overall, the surface is more uniform for the annealed film (rms roughness amplitude of 1.5 nm), although the island size is reduced in the process (giving an in-plane correlation length of 60 ± 20 nm for the annealed surface). The observation in AFM of smoother surfaces for the annealed film agree with the streakier RHEED patterns observed for the annealed films, and contrast with the more 3D transmission-like patterns of the as-grown films.

Structural characterization of the Co$_3$O$_4$ films was carried out ex situ by x-ray scattering measurements on a Shimadzu diffractometer using the Cu K$_\alpha$ line ($\lambda_{K\alpha1} = 1.540606$ Å) with a Ni filter to remove the Cu K$_\beta$ lines.

Comparison with α-Al$_2$O$_3$ XRD spectra allows one to determine readily the diffraction peaks originating from the Co$_3$O$_4$ films (not shown). The XRD results show that the spectra for the as-grown and annealed films are identical in the range from $2\theta = 20 - 100^\circ$, with diffraction peaks that coincide with the (hhl) planes ($h = 1, 2, 3, 4$) of Co$_3$O$_4$, at $2\theta$ values corresponding to the bulk lattice parameter. This confirms the spinel crystalline phase and epitaxy of the Co$_3$O$_4$ film, and also that the films are fully relaxed both after growth and after annealing. In Fig. 6 we show a detail of the XRD spectra of the annealed film around the (0006) α-Al$_2$O$_3$ peak, showing the (222) Co$_3$O$_4$ diffraction peak.

The XRR spectra of the substrate, as-grown, and annealed 22 nm Co$_3$O$_4$ films are shown in Fig. 7. The spectrum of the substrate (as-received) allows one to determine its surface roughness, which is estimated as 1 nm from the fit to the data (fits are shown as solid lines in Fig. 7). The reflectivity spectrum for the as-grown films show a rapid drop in the Kiessig fringes’ amplitude with momentum transfer; considering a single Co$_3$O$_4$ layer with bulk scattering density does not yield a good fit to the data, suggesting the presence of a non-uniform or graded interface layer. In contrast, the reflectivity spectrum of the annealed films show oscillations over the entire momentum trans-
fer range probed, indicating a significant improvement of the interface sharpness. Fits to the data indicate that the changes occur most significantly at the α-Al₂O₃ interface, where the roughness amplitude is found to decrease from 2.0 to 0.6 nm for the 22 nm sample, and from 4.0 to 1.5 nm for the 38 nm Co₃O₄ film. Hence, we interpret the strongly damped XRR spectra of the as-grown films as resulting from chemical or structural disorder at the Co₃O₄/α-Al₂O₃ interface, as also suggested from the RHEED pattern evolution during growth, which shows the α-Al₂O₃ sharp diffraction spots becoming streaky at about 15 Å and broadening at about 30 Å. In XRR, interface roughness may originate from graded interfaces, interdiffusion, and morphological roughness proper [64]. The improvement in interface sharpness upon annealing also suggests that the interface roughness in the as-grown film does not result from a reacted interface region, which should increase in thickness with annealing. Fits to the data give film thicknesses of 22.5 nm and 38.5 nm for the two samples, in good agreement with the nominal thicknesses estimated from the Co evaporation rate and the ratio between the Co₃O₄ and Co mass densities. The rms surface roughness of the Co₃O₄ films is found to increase slightly upon annealing from 1 to 2 nm for the 22 nm film, and from 2 to 2.5 nm for the 38 nm sample. The accuracy of the fits are within ±0.5 nm for the interface roughness and ±0.2 nm for the thickness.

These results are in marked contrast with those obtained for Co₃O₄(110)/MgAl₂O₄(110) films, where annealing under similar conditions results in atomically flat surfaces, associated with one particular termination of Co₃O₄ along the [110] direction [11]. Two factors can be expected to give rise to this difference in behaviour: the relative surface stability of the (110) and (111) crystal planes, and the role of strain in determining the surface morphology and surface roughness in particular. Surface roughening starts at the very early stages of the Co₃O₄(111) growth, and while this could be driven by misfit strain relaxation, we find that a similar roughening process also occurs in Co₃O₄/MgAl₂O₄(110), where the lattice mismatch is virtually zero [11]. Transmission electron microscopy results for the Co₃O₄/MgAl₂O₄(110) interface show sharp and relatively well defined interfaces [11], suggesting that the disorder observed in the XRR data could be due to modified film stoichiometry at the interface, leading to a graded interface structure; for the Co₃O₄/MgAl₂O₄(110) system the starting growth temperature was set to 770 K for the first 1 nm interface layer to optimize crystalline growth, which could lead to variations in interface composition and to a degree of interface grading. In the present study, the growth temperature was constant throughout the film growth; however, differences in the substrate and film crystal structure may favor an initial cationic arrangement closer to the corundum structure than to the spinel; for example, if the Co in the first few layers occupy an excess of octahedral sites, the film stoichiometry could be closer to Co₂O₅, resulting in a graded interface. This also suggests that film roughening may arise from a kinetic process common to both Co₃O₄/α-Al₂O₃(0001) and Co₃O₄/MgAl₂O₄(110) systems, for example, induced by limitations in atomic surface diffusion during growth. At annealing temperatures of about 800 K, atomic diffusion is sufficient to promote surface ordering in the Co₃O₄ films, but while the (110) films become atomically smooth, the (111) films retain their pronounced island structure and are not atomically flat, as indicated from the AFM, XRR.
References

[1] J. F. Picard, G. Baud, J. P. Besse, R. Chevalier, Journal of the Less-Common Metals 75 (1980) 99.
[2] J. Chenavas, J. C. Joubert, M. Marezio, Solid State Commun. 9 (1971) 1057.
[3] G. V. Samsonov (Ed.), The oxide handbook, 2nd Edition, IFI/Plenum, New York, 1982.
[4] M. Catti, G. Sandrone, Faraday Discuss. 106 (1997) 189.
[5] P. W. Tasker, J. Phys.: Condens. Matter 12 (1997) 4977.
[6] C. Noguera, J. Phys.: Condens. Matter 12 (2000) R367.
[7] J. Goniakowski, F. Finocchi, C. Noguera, Rep. Prog. Phys. 71 (2008) 016501.
[8] W. A. Harrison, E. A. Kraut, J. R. Waldrop, R. W. Grant, Phys. Rev. B 18 (1978) 4402.
[9] M. Gajdardziska-Josifovska, R. Plass, M. A. Schofield, D. R. Giese, R. Sharma, J. Electron Microsc. 51 (2002) S13.
[10] V. K. Lazarov, R. Plass, H.-C. Poon, D. K. Saldin, M. Weinert, S. A. Chambers, M. Gajdardziska-Josifovska, Phys. Rev. B 71 (2005) 115434.
[11] C. A. F. Vaz, H. Wang, C. H. Ahn, V. E. Henrich, M. Z. Baykara, T. Schwendemann, N. Filet, B. J. Albers, U. Schwarz, L. H. Zhang, Y. Zhu, J. Wang, E. I. Altman, Surf. Sci. 603 (2009) 291.
[12] W. Meyer, K. Bierdemann, M. Gubo, L. Hammer, K. Heinz, J. Phys.: Condens. Matter 20 (2008) 265011.
[13] S. C. Petitto, E. M. Marsh, G. A. Carson, M. A. Langell, J. Molecular Catalysis A: Chemical 281 (2008) 49.
[14] X. Tang, J. Li, J. Hao, Materials Research Bulletin 43 (2008) 2912.
[15] C. M. Fang, S. C. Parker, G. de With, J. Am. Ceram. Soc. 83 (2000) 2082.
[16] M. J. Davies, S. P. Parker, G. W. Watson, J. Mater. Chem. 4 (1994) 813.
[17] R. L. Stewart, R. C. Bradt, J. Mater. Sci. 15 (1980) 67.
[18] R. C. Bradt, Cleavage of ceramic and mineral single crystals, in: K. S. Chan (Ed.), George R. Irving Symposium on Cleavage Fracture, Warrendale, PA, 1997, p. 355.
[19] C. M. Fang, G. de With, S. C. Parker, J. Am. Ceram. Soc. 84 (2001) 1553.
[20] N. J. van der Laag, A. J. M. van Dijk, N. Louisberg, G. de With, L. J. M. G. Dortmans, J. Am. Ceram. Soc. 88 (2005) 660.
[21] R. E. Newnham, Y. M. D. Haan, Zeitschrift für Kristallographie 117 (1962) 235.
[22] C. S. G. Cousins, J. Phys. C: Solid State Phys. 14 (1981) 1585.
[23] A. Kirfel, K. Eichhorn, Acta Cryst. A 46 (1990) 271.
[24] R. W. G. Wyckoff, Crystal structures, 2nd Edition, Vol. 2, Interscience Publishers, New York, 1964.
[25] J. Guo, D. E. Ellis, J. D. Lam, Phys. Rev. B 45 (1992) 13647.
[26] T. J. Godin, P. Labèfemina, Phys. Rev. B 49 (1994) 7691.
[27] I. Manassidis, A. D. Vita, M. Gillan, Surf. Sci. Lett. 285 (1993) L517.
[28] J. Ahn, J. Rabalais, Surf. Sci. 388 (1997) 121.
[29] P. Guénard, G. Renaud, A. Barbier, M. Gautier-Soyer, Surf. Rev. Lett. 5 (1997) 321.
[30] G. Renaud, Surface Science Reports 32 (1998) 1.
[31] E. A. Jarvis, E. A. Carter, J. Phys. Chem. B 105 (2001) 4045.
[32] E. Wallin, J. M. Andersson, E. P. Münger, V. Chirita, U. Helmersson, Phys. Rev. B 74 (2006) 125409.
[33] C. C. Chang, J. Appl. Phys. 39 (1968) 5570.
[34] T. M. French, G. A. Somorjai, J. Appl. Phys. 74 (1970) 2489.
[35] V. E. Henrich, P. A. Cox, The surface science of metal oxides, Cambridge University Press, Cambridge, 1994.
[36] T. Suzuki, S. Hishita, K. Oyoshi, R. Soulé, Surf. Sci. 437 (1999) 289.
[37] C. F. Walters, K. F. McCarty, E. A. Soares, M. A. Van Hove, Surf. Sci. 464 (2000) L732.
[38] C. Barth, M. Reichling, Nature 414 (2001) 54.
[39] E. A. Soares, M. A. Van Hove, C. F. Walters, K. F. McCarty, Phys. Rev. B 70 (2004) 115409.
[40] A. Marmier, S. C. Parker, Rev. Lett. 65 (1990) 2082.
[41] J. A. Kelber, Surf. Sci. Rep. 62 (2007) 271.
[42] W. Gao, R. Klei, E. Altman, Thin Solid Films 485 (2005) 115.
[43] R. Anton, T. Wiegner, W. Naumann, M. Liebmann, C. Klein, C. Bradley, Rev. Sci. Instrum. 71 (2000) 1177.
[44] H. P. Tripp, B. W. King, J. American Ceramic Soc. 38 (1955) 432.
[45] K. Koumoto, H. Yanagida, Jpn. J. Appl. Phys. 20 (1981) 445.
[46] M. Oku, Y. Sato, Appl. Surf. Sci. 55 (1992) 37.
[47] R. C. Rossi, R. M. Fulrath, J. Am. Ceram. Soc. 46 (1963) 145.
[48] C.-M. Liu, J.-C. Chen, C.-J. Chen, J. Crystal Growth 285 (2005) 7.
275.

[49] Y. Gao, Y. J. Kim, S. A. Chambers, G. Bai, J. Vac. Sci. Technol. A 15 (1997) 332.

[50] M. J. Stowell, Defects in epitaxial deposits, in: J. W. Matthews (Ed.), Epitaxial growth, Part B, Academic Press, Inc., 1975, p. 437.

[51] T. J. Chuang, C. R. Brundle, D. W. Rice, Surf. Sci. 59 (1976) 413.

[52] J. Haber, L. Ungier, J. Electron Spectrosc. Relat. Phenom. 12 (1977) 305.

[53] J. P. Bonnelle, J. Grimblot, A. D'Huysser, J. Electron Spectrosc. Relat. Phenom. 7 (1975) 151.

[54] M. A. Langell, M. D. Anderson, G. A. Carson, L. Peng, S. Smith, Phys. Rev. B 59 (1999) 4791.

[55] H. A. Hagelin-Weaver, G. B. Hofflund, D. M. Minahan, G. N. Salaita, Appl. Surf. Sci. 235 (2004) 420.

[56] H.-Q. Wang, E. I. Altman, V. E. Henrich, Phys. Rev. B 77 (2008) 085313.

[57] Y. Jugnet, T. M. Duc, J. Phys. Chem. Solids 40 (1979) 29.

[58] Y. M. Kolotyrkin, I. D. Belova, Y. E. Roginskaya, V. B. Kozhevnikov, D. S. Zakhar' in, Y. N. Venevtsev, Materials Chemistry and Physics 11 (1984) 29.

[59] B. Klingenberg, F. Grellner, D. Borgmann, G. Wedler, Surf. Sci. 296 (1993) 374.

[60] B. Klingenberg, F. Grellner, D. Borgmann, G. Wedler, J. Electron Spectrosc. Relat. Phenom. 71 (1995) 107.

[61] V. M. Jiménez, A. Fernández, J. P. Espinós, A. R. González-Elipe, J. Electron Spectrosc. Relat. Phenom. 71 (1995) 61.

[62] G. A. Carson, M. H. Nassir, M. A. Langell, J. Vac. Sci. Technol. A 14 (1996) 1637.

[63] http://gwyddion.net

[64] M. Wormington, I. Pape, T. P. A. Hase, B. K. Tanner, D. K. Bowen, Phil. Mag. Lett. 74 (1996) 211.

[65] M. Ritter, W. Weiss, Surf. Sci. 432 (1999) 81.

[66] I. V. Shvets, N. Berdunov, G. Mariotto, S. Murphy, Europhys. Lett. 63 (2003) 567.