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Why do models overestimate surface ozone in the Southeast United States?

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Abstract. Ozone pollution in the Southeast US involves complex chemistry driven by emissions of anthropogenic nitrogen oxide radicals (NOx = NO + NO2) and biogenic isoprene. Model estimates of surface ozone concentrations tend to be biased high in the region and this is of concern for designing effective emission control strategies to meet air quality standards. We use detailed chemical observations from the SEAC4RS aircraft campaign in August and September 2013, interpreted with the GEOS-Chem chemical transport model at 0.25° × 0.3125° horizontal resolution, to better understand the factors controlling surface ozone in the Southeast US. We find that the National Emission Inventory (NEI) for NOx from the US Environmental Protection Agency (EPA) is too high. This finding is based on SEAC4RS observations of NOx and its oxidation products, surface network observations of nitrate wet deposition fluxes, and OMI satellite observations of tropospheric NO2 columns. Our results indicate that NEI NOx emissions from mobile and industrial sources must be reduced by 30–60%, dependent on the assumption of the contribution by soil NOx emissions. Upper-tropospheric NO2 from lightning makes a large contribution to satellite observations of tropospheric NO2 that must be accounted for when using these data to estimate surface NOx emissions. We find that only half of isoprene oxidation proceeds by the high-NOx pathway to produce ozone; this fraction is only moderately sensitive to changes in NOx.
emissions because isoprene and NOx emissions are spatially segregated. GEOS-Chem with reduced NOx emissions provides an unbiased simulation of ozone observations from the aircraft and reproduces the observed ozone production efficiency in the boundary layer as derived from a regression of ozone and NOx oxidation products. However, the model is still biased high by 6 ± 14 ppb relative to observed surface ozone in the Southeast US. Ozone sondes launched during midday hours show a 7 ppb ozone decrease from 1.5 km to the surface that GEOS-Chem does not capture. This bias may reflect a combination of excessive vertical mixing and net ozone production in the model boundary layer.

1 Introduction

Ozone in surface air is harmful to human health and vegetation. Ozone is produced when volatile organic compounds (VOCs) and carbon monoxide (CO) are photochemically oxidized in the presence of nitrogen oxide radicals (NOx ≡ NO + NO2). The mechanism for producing ozone is complicated, involving hundreds of chemical species interacting with transport on all scales. In October 2015, the US Environmental Protection Agency (EPA) set a new National Ambient Air Quality Standard (NAAQS) for surface ozone as a maximum daily 8 h average (MDA8) of 0.070 ppm not to be exceeded more than three times per year. This is the latest in a succession of gradual tightening of the NAAQS from 0.12 ppm (1 h average) to 0.08 ppm in 1997 and to 0.075 ppm in 2008, responding to accumulating evidence that ozone is detrimental to public health even at low concentrations (EPA, 2013). Chemical transport models (CTMs) tend to significantly overestimate surface ozone in the Southeast US (Lin et al., 2008; Fiore et al., 2009; Reidmiller et al., 2009; Brown-Steiner et al., 2015; Canty et al., 2015), and this is an issue for the design of pollution control strategies (McDonald-Buller et al., 2011). Here we examine the causes of this overestimate by using the GEOS-Chem CTM to simulate NASA SEAC4RS aircraft observations of ozone and its precursors over the region in August–September 2013 (Toon et al., 2016), together with additional observations from surface networks and satellite.

A number of explanations have been proposed for the ozone model overestimates in the Southeast US. Fiore et al. (2003) suggested excessive modeled ozone inflow from the Gulf of Mexico. Lin et al. (2008) proposed that the ozone dry deposition velocity could be underestimated. McDonald-Buller et al. (2011) pointed out the potential role of halogen chemistry as a sink of ozone. Isoprene emitted from vegetation is the principal VOC precursor of ozone in the Southeast US in summer, and Fiore et al. (2005) found that uncertainties in isoprene emissions and in the loss of NOx from formation of isoprene nitrates could also affect the ozone simulation. Horowitz et al. (2007) found a large sensitivity of ozone to the fate of isoprene nitrates and the extent to which they release NOx when oxidized. Squire et al. (2015) found that the choice of isoprene oxidation mechanism can alter both the sign and magnitude of the response of ozone to isoprene and NOx emissions.

The SEAC4RS aircraft campaign in August–September 2013 provides an outstanding opportunity to improve our understanding of ozone chemistry over the Southeast US. The SEAC4RS DC-8 aircraft hosted an unprecedented chemical payload including isoprene and its oxidation products, NOx and its oxidation products, and ozone. The flights featured extensive boundary layer mapping of the southeast as well as vertical profiling to the free troposphere (Toon et al., 2016). We use the GEOS-Chem global CTM with high horizontal resolution over North America (0.25° × 0.3125°) to simulate and interpret the SEAC4RS observations. We integrate into our analysis additional Southeast US observations during the summer of 2013, including from the NOMADSS aircraft campaign, the SOAS surface site in Alabama, the SEACIONS ozonesonde network, the EPA Clean Air Status and Trends Network (CASTNET) ozone network, the National Acid Deposition Program (NADP) nitrate wet deposition network, and NO2 satellite data from the OMI instrument. Several companion papers apply GEOS-Chem to simulate other aspects of SEAC4RS and concurrent data for the Southeast US, including aerosol sources and optical depth (Kim et al., 2015), isoprene organic aerosol (Marais et al., 2016), organic nitrates (Fisher et al., 2016), formaldehyde and its relation to satellite observations (Zhu et al., 2016), and sensitivity to model resolution (Yu et al., 2016).

2 GEOS-Chem model description

We use the GEOS-Chem global 3-D CTM (Bey et al., 2001) in version 9.02 (http://www.geos-chem.org) with modifications described below. GEOS-Chem is driven with assimilated meteorological data from the Goddard Earth Observing System (GEOS-5.11.0) of the NASA Global Modeling and Assimilation Office (GMAO). The GEOS-5.11.0 data have a native horizontal resolution of 0.25° × 0.3125° latitude and a temporal resolution of 3 h (1 h for surface variables and mixing depths). We use a nested version of GEOS-Chem (Chen et al., 2009) with native 0.25° × 0.3125° horizontal resolution over North America and adjacent oceans (130–60° W, 9.75–60° N) and dynamic boundary conditions from a global simulation with 4° × 5° horizontal resolution. Turbulent boundary layer mixing follows a non-local parameterization based on K-theory (Holtslag and Boville, 1993) implemented in GEOS-Chem by Lin and McElroy (2010). Daytime mixing depths are reduced by 40 % from the GEOS-5.11.0 data as described by Kim et al. (2015) and Zhu et al. (2016) to match aircraft lidar observations. The GEOS-Chem nested model simulation is conducted for August–
September 2013, following 6 months of initialization at 4° × 5° resolution.

2.1 Chemistry

The chemical mechanism in GEOS-Chem version 9.02 is described by Mao et al. (2010, 2013). We modified aerosol reactive uptake of HO2 to produce H2O2 instead of H2O in order to better match H2O2 observations in SEAC4RS. We also include a number of updates to isoprene chemistry, listed comprehensively in the Supplement (Tables S1 and S2) and described here more specifically for the low-NOx pathways. Companion papers describe the isoprene chemistry updates relevant to isoprene nitrates (Fisher et al., 2016) and organic aerosol formation (Marais et al., 2016). Oxidation of biogenic monoterpenes is also added to the GEOS-Chem mechanism (Fisher et al., 2016) but does not significantly affect ozone.

A critical issue in isoprene chemistry is the fate of the isoprene peroxy radicals (ISOPO2) produced from the oxidation of isoprene by OH (the dominant isoprene sink). When NOx is sufficiently high, ISOPO2 reacts mainly with NO to produce ozone (high-NOx pathway). At lower NOx levels, ISOPO2 may instead react with HO2 or other organic peroxy radicals, or isomerize, in which case ozone is not produced (low-NOx pathways). Here we increase the molar yield of isoprene hydroperoxide (ISOPOOH) from the ISOPO2 + HO2 reaction to 94% based on observations of the minor channels of this reaction (Liu et al., 2013). Oxidation of ISOPOOH by OH produces isoprene epoxides (IEPOX) that subsequently react with OH or are taken up by aerosol (Paulot et al., 2009b; Marais et al., 2016). We use updated rates and products from Bates et al. (2014) for the reaction of IEPOX with OH.

ISOPO2 isomerization produces hydroperoxy-aldehydes (HPALDs) (Peeters et al., 2009; Crounse et al., 2011; Wolfe et al., 2012), and we explicitly include this in the GEOS-Chem mechanism. HPALDs go on to react with OH or photolyze at roughly equal rates over the Southeast US. We use the HPALD + OH reaction rate constant from Wolfe et al. (2012) and the products of the reaction from Squire et al. (2015). The HPALD photolysis rate is calculated using the absorption cross section of MACR, with a quantum yield of 1, as recommended by Peeters and Müller (2010). The photolysis products are taken from Stavrakou et al. (2010). Self-reaction of ISOPO2 is updated following Xie et al. (2013).

A number of studies have suggested that conversion of NO2 to nitrous acid (HONO) by gas-phase or aerosol-phase pathways could provide a source of HO2 radicals following HONO photolysis (Li et al., 2014; Zhou et al., 2014). This mechanism would also provide a catalytic sink for ozone when NO2 is produced by the NO + ozone reaction, viz.,

\[
\text{NO} + \text{O}_3 \rightarrow \text{NO}_2 + \text{O}_2, \quad (\text{R}1)
\]

\[
\text{NO}_2 \rightarrow \text{HONO} \quad \text{(by various pathways)}, \quad (\text{R}2)
\]

\[
\text{HONO} + h\nu \rightarrow \text{NO} + \text{OH}. \quad (\text{R}3)
\]

Observations of HONO from the NOMADSS campaign (https://www2.acom.ucar.edu/campaigns/nomadss) indicate a mean daytime HONO concentration of 10 ppt in the Southeast US boundary layer (Zhou et al., 2014), whereas the standard gas-phase mechanism in GEOS-Chem version 9.02 yields less than 1 ppt. We add the pathway proposed by Li et al. (2014), in which HONO is produced by the reaction of the HO2 + H2O complex with NO2, but with a slower rate constant \(k_{\text{HO}_2, \text{H}_2\text{O}, \text{NO}_2} = 2 \times 10^{-12} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}\) to match the observed ~10 ppt daytime HONO in the Southeast US boundary layer. The resulting impact on boundary layer ozone concentrations is negligible.

2.2 Dry deposition

The GEOS-Chem dry deposition scheme uses a resistance-in-series model based on Wesely (1989) as implemented by Wang et al. (1998). Underestimation of dry deposition has been invoked as a cause for model overestimates of ozone in the eastern US (Lin et al., 2008; Walker, 2014). Daytime ozone deposition is determined principally by stomatal uptake. Here, we decrease the stomatal resistance from 3.9 and 0.6 cm s\(^{-1}\) for both coniferous and deciduous forests (Wesely, 1989) by 20% to match summertime measurements of the ozone dry deposition velocity for a pine forest in North Carolina (Finkelstein et al., 2000) and for the Ozarks oak forest in southeastern Missouri (Wolfe et al., 2015), both averaging 0.8 cm s\(^{-1}\) in the daytime. The mean ozone deposition velocity in GEOS-Chem along the SEAC4RS boundary layer flight tracks in the Southeast US averages 0.7 ± 0.3 cm s\(^{-1}\) for the daytime (09:00–16:00 local) surface layer. Deposition is suppressed in the model at night due to both stomatal closure and near-surface stratification, consistent with the Finkelstein et al. (2000) observations.

Deposition flux measurements for isoprene oxidation products at the Alabama SOAS site (http://soas2013.rutgers.edu) indicate higher deposition velocities than simulated by the standard GEOS-Chem model (Nguyen et al., 2015). The diurnal cycle of dry deposition in GEOS-Chem compares well with the observations from SOAS (Nguyen et al., 2015). As an expedient, Nguyen et al. (2015) scaled the Henry’s law coefficients for these species in GEOS-Chem to match their observed deposition velocities and we follow their approach here. Other important depositing species include HNO3 and peroxyacetyl nitrate (PAN), with mean deposition velocities along the SEAC4RS Southeast US flight tracks in daytime of 3.9 and 0.6 cm s\(^{-1}\), respectively.
2.3 Emissions

We use hourly US anthropogenic emissions from the 2011 EPA National Emissions Inventory (NEI11v1) at a horizontal resolution of 0.1° × 0.1° and adjusted to 2013 using national annual scaling factors (EPA NEI, 2015). The scaling factor for NO\textsubscript{x} emissions is 0.89, for a 2013 US NEI total of 3.5 Tg N a\textsuperscript{-1}. Further information on the use of the NEI11v1 in GEOS-Chem can be found at http://wiki.seas.harvard.edu/geos-chem/index.php/EPA/NEI11_North_American_emissions. Soil NO\textsubscript{x} emissions, including emissions from fertilizer application, are computed according to Hudman et al. (2012), with a 50 % reduction in the Midwestern US based on a previous comparison with OMI NO\textsubscript{2} observations (Vinken et al., 2014). Open fire emissions are from the daily Quick Fire Emissions Database (QFED) (Darmenov and da Silva, 2014) with diurnal variability from the Western Regional Air Partnership (Air Sciences, 2005). We emit 40 % of open fire NO\textsubscript{x} emissions as PAN and 20 % as HNO\textsubscript{3} to account for fast oxidation taking place in the fresh plume (Alvarado et al., 2010). Following Fischer et al. (2014), we inject 35 % of fire emissions above the boundary layer, evenly between 3.5 and 5.5 km altitude. Lightning is an additional source of NO\textsubscript{x}, but is mainly released in the upper troposphere, as described below.

Initial implementation of the above inventory in GEOS-Chem resulted in an 60–70 % overestimation of NO\textsubscript{x} and HNO\textsubscript{3} measured by the SEAC\textsuperscript{4}RS DC-8 aircraft and a 70 % overestimation of nitrate (NO\textsubscript{3}\textsuperscript{-}) wet deposition fluxes measured by the NADP across the Southeast US. Correcting this bias required a ~40 % decrease in surface NO\textsubscript{x} emissions. Assuming strongly reduced soil and fertilizer NO\textsubscript{x} emissions (18 % of total NO\textsubscript{x} emissions in the southeast) and open fires (2 %), also considering the large uncertainty in these emissions, would be insufficient to correct this bias. Emissions from power plant stacks are directly measured but account for only 12 % of NEI NO\textsubscript{x} emissions on an annual basis (EPA NEI, 2015). Several local studies in recent years have found that NEI NO\textsubscript{x} emissions for mobile sources may be too high by a factor of 2 or more (Castellanos et al., 2011; Fujita et al., 2012; Brioude et al., 2013; Anderson et al., 2014). We can achieve the required 40 % decrease in total NO\textsubscript{x} emissions by reducing NEI emissions from mobile and industrial sources (all sources except power plants) by 60 % or alternatively by reducing these sources by 30 % and zeroing out soil and fertilizer NO\textsubscript{x} emissions. Since it is apparent that there is some minimum contribution by soil NO\textsubscript{x} emissions, we assessed the impact of the approach of reducing the non-power-plant NEI emissions by 60 %. The spatial overlap between anthropogenic and soil NO\textsubscript{x} emissions is such that we cannot readily arbitrate between these two scenarios. Comparisons with observations will be presented in the next section.

We constrain the lightning NO\textsubscript{x} source with satellite data as described by Murray et al. (2012). Lightning NO\textsubscript{x} is mainly released at the top of convective updrafts following Ott et al. (2010). The standard GEOS-Chem model uses higher NO\textsubscript{x} yields for midlatitudes lightning (500 mol flash\textsuperscript{-1}) than for tropical (260 mol flash\textsuperscript{-1}) (Huntrieser et al., 2007, 2008; Hudman et al., 2007; Ott et al., 2010) with a fairly arbitrary boundary between the two at 23° N in North America and 35° N in Eurasia. Zhang et al. (2014) previously found that this leads GEOS-Chem to overestimate background ozone in the southwestern US and we find the same here for the eastern US and the Gulf of Mexico. We treat here all lightning in the 35° S–35° N band as tropical and thus remove the distinction between North America and Eurasia.

Figure 1 gives the resulting surface NO\textsubscript{x} emissions for the Southeast US for August and September 2013. With the original NEI inventory, fuel combustion accounted for 81 % of total surface NO\textsubscript{x} emissions in the Southeast US (not including lightning). If the required reduction of non-power-plant
NEI emissions is 60%, the contribution from fuel combustion would be 68%.

Biogenic VOC emissions are from MEGAN v2.1, including isoprene, acetone, acetaldehyde, monoterpenes, and \( \geq \text{C}_2 \) alkenes. We reduce MEGAN v2.1 isoprene emissions by 15% to better match SEAC4RS observations of isoprene fluxes from the Ozarks (Wolfe et al., 2015) and observed formaldehyde (Zhu et al., 2016). Yu et al. (2016) show the resulting isoprene emissions for the SEAC4RS period.

3 Overestimate of NO\(_x\) emissions in the EPA NEI inventory

Figure 2 shows simulated and observed median vertical distributions of NO\(_x\), total inorganic nitrate (gas-phase HNO\(_3\) + aerosol NO\(_3^-\)), ozone, isoprene nitrate (ISOPN), isoprene hydroperoxide (ISOPOOH), and hydroperoxyaldehydes (HPALDs) for the SEAC4RS flights over the Southeast US (domain of Fig. 1). Observations from the DC-8 aircraft are compared to GEOS-Chem model results. The dashed red line shows model results before adjustment of NO\(_x\) emissions from fuel combustion and lightning (see text). The 25th and 75th percentiles of the DC-8 observations are shown as grey bars. The SEAC4RS observations have been filtered to remove open fire plumes, stratospheric air, and urban plumes as described in the text. Model results are sampled along the flight tracks at the time of flights and gridded to the model resolution. Profiles are binned to the nearest 0.5 km. The NOAA NO\(_x\)O\(_3\) four-channel chemiluminescence (CL) instrument made measurements of ozone and NO\(_3^-\) (Ryerson et al., 1998), NO (Ryerson et al., 2000), and NO\(_2\) (Polliack et al., 2010). Total inorganic nitrate was measured by the University of New Hampshire Soluble Acidic Gases and Aerosol (UNH SAGA) instrument (Dibb et al., 2003) and was mainly gas-phase HNO\(_3\) for the SEAC4RS conditions. ISOPOOH, ISOPN, and HPALDs were measured by the Caltech single mass analyzer CIMS (Crounse et al., 2006; Paulot et al., 2009a; Crounse et al., 2011).

Further support for decreasing NO\(_x\) emissions is offered by observed nitrate wet deposition fluxes from the NADP network (NADP, 2007). Figure 3 compares simulated and observed fluxes for the model with decreased NO\(_x\) emissions. Model values have been corrected for precipitation bias following the method of Paulot et al. (2014), in which the monthly deposition flux is assumed to scale to the 0.6th power of the precipitation bias. We diagnose precipitation bias in the GEOS-5.11.0 data relative to high-resolution PRISM observations (http://prism.oregonstate.edu). For the Southeast US, the precipitation bias is \(-34\%\) in August and \(-21\%\) in September 2013. We see from Fig. 3 that the model with decreased NO\(_x\) emissions reproduces the spatial variability in the observations with only \(+8\%\) bias over the Southeast US and \(+7\%\) over the contiguous US. In comparison, the model with original emissions had a \(63\%\) overestimation of the nitrate wet deposition flux nationally and a \(71\%\) overestimation in the southeast. The high deposition fluxes along the Gulf of Mexico in Fig. 3, both in the model and in the observations, reflect particularly large precipitation.

The model with decreased NO\(_x\) emissions also reproduces the spatial distribution of NO\(_x\) in the Southeast US boundary layer as observed in SEAC4RS. This is shown in Fig. 4 with simulated and observed concentrations of NO\(_x\) along the flight tracks below 1.5 km altitude. The spatial correlation coefficient is 0.71. There are no obvious spatial patterns of model bias that would point to specific source sectors as responsible for the NO\(_x\) emission overestimate, beyond the
blanket 30–60 % decrease of non-power-plant NEI emissions needed to correct the regional emission total.

4 Using satellite NO$_2$ data to verify NO$_x$ emissions: sensitivity to upper troposphere

Observations of tropospheric NO$_2$ columns by solar backscatter from the OMI satellite instrument offer an additional constraint on NO$_x$ emissions (Duncan et al., 2014; Lu et al., 2015). We compare the tropospheric columns simulated by GEOS-Chem with the NASA operational retrieval (Level 2, v2.1) (NASA, 2012; Bucsela et al., 2013) and the Berkeley High-Resolution (BEHR) retrieval (Russell et al., 2011). The NASA retrieval has been validated to agree with surface measurements to within ±20 % (Lamsal et al., 2014). Both retrievals fit the observed backscattered solar spectra to obtain a slant tropospheric NO$_2$ column, $\Omega_s$, along the optical path of the backscattered radiation detected by the satellite. The slant column is converted to the vertical column, $\Omega_v$, by using an air mass factor (AMF) that depends on the vertical profile of NO$_2$ and on the scattering properties of the surface and the atmosphere (Palmer et al., 2001):

$$\Omega_v = \Omega_s \frac{AMF}{AMF_G \int_0^{z_T} w(z) S(z) dz}.$$  \hspace{1cm} (1)

In Eq. (4), AMF$_G$ is the geometric air mass factor that depends on the viewing geometry of the satellite, $w(z)$ is a scattering weight calculated by a radiative transfer model that describes the sensitivity of the backscattered radiation to NO$_2$ as a function of altitude, $S(z)$ is a shape factor describing the normalized vertical profile of NO$_2$ number density, and $z_T$ is the tropopause. Scattering weights for NO$_2$ retrievals typically increase by a factor of 3 from the surface to the upper troposphere (Martin et al., 2002). Here we use our GEOS-Chem shape factors to recalculate the AMFs in the NASA and BEHR retrievals as recommended by Lamsal et al. (2014) for comparing model and observations. We filter out cloudy scenes (cloud radiance fraction > 0.5) and bright surfaces (surface reflectivity > 0.3).

Figure 5 shows the mean NO$_2$ tropospheric columns from BEHR, NASA, and GEOS-Chem (with NO$_x$ emission reductions applied) over the Southeast US for August–September 2013. The BEHR retrieval is on average 6 % higher than the NASA retrieval. GEOS-Chem is on average 11 ± 19 % lower than the NASA retrieval and 16 ± 18 % lower than the BEHR retrieval. With the original NEI NO$_x$ emissions, GEOS-Chem would be biased high against both retrievals by 26–31 %. The low bias in the model with reduced NO$_x$ emissions does not appear to be caused by an overcorrection of surface emissions but rather by the upper troposphere. Figure 6 (top left panel) shows the mean vertical profile of NO$_2$ number density as measured from the aircraft by two independent instruments (NOAA and UC Berkeley) and simulated by GEOS-Chem. At the surface, the median difference is $1.8 \times 10^9$ molecules cm$^{-3}$, which is
within the NOAA and UC Berkeley measurement uncertainties of ±0.030 ppbv + 7% and ±5%, respectively. The observations show a secondary maximum in the upper troposphere above 10 km, absent in GEOS-Chem. It has been suggested that aircraft measurements of NO$_2$ in the upper troposphere could be biased high due to decomposition in the instrument inlet of thermally unstable NO$_3$ reservoirs such as HNO$_4$ and methyl peroxy nitrate (Browne et al., 2011; Reed et al., 2016). This would not affect the UC Berkeley measurement (Nault et al., 2015) and could possibly account for the difference with the NOAA measurement in Fig. 6.

The top right panel of Fig. 6 shows the cumulative contributions from different altitudes to the slant NO$_2$ column measured by the satellite, using the median vertical profiles from the left panel and applying mean altitude-dependent scattering weights from the NASA and BEHR retrievals. The boundary layer below 1.5 km contributes only 19–28% of the column. The upper troposphere above 8 km contributes 32–49% in the aircraft observations and 23% in GEOS-Chem. Much of the observed upper-tropospheric NO$_2$ likely originates from lightning and is broadly distributed across the southeast because of the long lifetime of NO$_3$. At that altitude (Li et al., 2005; Bertram et al., 2007; Hudman et al., 2007). The NO$_2$ vertical profile (shape factor) assumed in the BEHR retrieval does not include any lightning influence, and the Global Modeling Initiative (GMI) model vertical profile assumed in the NASA retrieval has little contribution from the upper troposphere (Lamsal et al., 2014). These underestimates of upper-tropospheric NO$_2$ in the retrieval shape factors will cause a negative bias in the AMF and therefore a positive bias in the retrieved vertical columns.

The GEOS-Chem underestimate of observed upper-tropospheric NO$_2$ in Fig. 6 is partly driven by NO/NO$_2$ partitioning. The bottom left panel of Fig. 6 shows the [NO]/[NO$_2$] concentration ratio in GEOS-Chem and in the observations (NOAA for NO, UC Berkeley for NO$_2$). One would expect the [NO]/[NO$_2$] concentration ratio in the

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Figure 5. NO$_2$ tropospheric columns over the Southeast US in August–September 2013. GEOS-Chem (sampled at the 13:30 local time overpass of OMI) is compared to OMI satellite observations using the BEHR and NASA retrievals. Values are plotted on the 0.25° × 0.3125° GEOS-Chem grid. The GEOS-Chem mean bias over the figure domain and associated spatial standard deviation are inset in the bottom panel.

Figure 6. Vertical distribution of NO$_2$ over the Southeast US during SEAC4RS (August–September 2013) and contributions to tropospheric NO$_2$ columns measured from space by OMI. The top left panel shows median vertical profiles of NO$_2$ number density measured from the SEAC4RS aircraft by the NOAA and UC Berkeley instruments and simulated by GEOS-Chem. The top right panel shows the fractional contribution of NO$_2$ below a given altitude to the total tropospheric NO$_2$ slant column measured by OMI, accounting for increasing sensitivity with altitude as determined from the retrieval scattering weights. The bottom left panel shows the median vertical profiles of the daytime [NO]/[NO$_2$] molar concentration ratio in the aircraft observations (NOAA for NO and UC Berkeley for NO$_2$) and in GEOS-Chem. Also shown is the ratio computed from NO–NO$_2$–O$_3$ photochemical steady state (PSS) as given by Reactions (4) and (6) (blue) and including Reaction (5) with doubled HO$_2$ and RO$_2$ concentrations above 8 km (purple). The bottom right panel shows the median H$_2$O$_2$ profile from the model and from the SEAC4RS flights over the Southeast US. H$_2$O$_2$ was measured by the Caltech CIMS (see Fig. 2).
daytime upper troposphere to be controlled by photochemical steady state:

\[ \text{NO} + \text{O}_3 \rightarrow \text{NO}_2 + \text{O}_2. \quad (R4) \]

\[ \text{NO} + \text{HO}_2/\text{RO}_2 \rightarrow \text{NO}_2 + \text{OH}/\text{RO}. \quad (R5) \]

\[ \text{NO}_2 + h\nu \rightarrow \text{O}_2 \rightarrow \text{NO} + \text{O}_3. \quad (R6) \]

If Reaction (R5) plays only a minor role then \([\text{NO}] / [\text{NO}_2] \approx k_6 / (k_4[\text{O}_3]),\) defining the NO–NO\(_2–\text{O}_3\) photochemical steady state (PSS). The PSS plotted in Fig. 6 agrees closely with GEOS-Chem. Such agreement has previously been found when comparing photochemical models with observed \([\text{NO}] / [\text{NO}_2]\) ratios from aircraft in the marine upper troposphere (Schultz et al., 1999) and lower stratosphere (Del Negro et al., 1999). The SEAC\(^4\)RS observations show large departure. The NO\(_2\) photolysis frequencies \(k_6\) computed locally by GEOS-Chem are on average within 10% of the values determined in SEAC\(^4\)RS from measured actinic fluxes (Shetter and Muller, 1999), so this is not the problem.

A possible explanation is that the model underestimates peroxy radical concentrations and hence the contribution of Reaction (5) in the upper troposphere. Zhu et al. (2016) found that GEOS-Chem underestimates the observed HCHO concentrations in the upper troposphere during SEAC\(^4\)RS by a factor of 3, implying that the model underestimates the HO\(_3\) source from convective injection of HCHO and peroxides (Jaegele et al., 1997; Prather and Jacob, 1997; Müller and Brasseur, 1999). HO\(_3\) observations over the central US in summer during the SUCCESS aircraft campaign suggest that this convective injection increases HO\(_3\) concentrations in the upper troposphere by a factor of 2 (Jaegele et al., 1998). The bottom right panel of Fig. 6 shows median modeled and observed vertical profiles of the HO\(_3\) reservoir hydrogen peroxide (H\(_2\)O\(_2\)) during SEAC\(^4\)RS over the Southeast US. GEOS-Chem underestimates observed H\(_2\)O\(_2\) by a mean factor of 1.7 above 8 km. The bottom left panel of Fig. 6 shows the \([\text{NO}] / [\text{NO}_2]\) ratio in GEOS-Chem with H\(_2\)O\(_2\) and RO\(_2\) doubled above 8 km. Such a change corrects significantly the bias relative to observations.

The PSS and GEOS-Chem simulation of the NO / NO\(_2\) concentration ratio in Fig. 6 use \(k_4 = 3.0 \times 10^{-12} \exp[-1500/T] \text{cm}^3\text{molecule}^{-1} \text{s}^{-1}\) and spectroscopic information for \(k_6\) from Sander et al. (2011). It is possible that the strong thermal dependence of \(k_4\) has some error, considering that only one direct measurement has been published for the cold temperatures of the upper troposphere (Borders and Birks, 1982). Cohen et al. (2000) found that reducing the activation energy of \(k_4\) by 15% improved model agreement in the lower stratosphere. Correcting the discrepancy between simulated and observed \([\text{NO}] / [\text{NO}_2]\) ratios in the upper troposphere in Fig. 6 would require a similar reduction to the activation energy of \(k_4\), but this reduction would negatively impact the surface comparison. This inconsistency of the observed \([\text{NO}] / [\text{NO}_2]\) ratio with basic theory needs to be resolved, as it affects the inference of NO\(_x\) emissions from satellite NO\(_2\) column measurements. Notwithstanding this inconsistency, we find that NO\(_2\) in the upper troposphere makes a significant contribution to the tropospheric NO\(_2\) column observed from space.

### 5 Isoprene oxidation pathways

Measurements aboard the SEAC\(^4\)RS aircraft included first-generation isoprene nitrates (ISOPN), isoprene hydroperoxide (ISOPOOH), and hydroperoxy-aldehydes (HPALDs) (Crounse et al., 2006; Paulot et al., 2009a; St. Clair et al., 2010; Crounse et al., 2011; Beaver et al., 2012; Nguyen et al., 2015). Although measurement uncertainties are large (30, 40, and 50%, respectively; Nguyen et al., 2015), these are unique products of the ISOPO\(_2\) + NO, ISOPO\(_2\) + HO\(_2\), and ISOPO\(_2\) isomerization pathways and thus track whether oxidation of isoprene proceeds by the high-NO\(_x\) path (producing ozone) or the low-NO\(_x\) path. Figure 2 (bottom row) compares simulated and observed concentrations. All three gases are restricted to the boundary layer because of their short lifetimes. Mean model concentrations in the lowest altitude bin (Fig. 2, approximately 400 m above ground) differ from observations by +19% for ISOPN, +70% for ISOPOOH, and −40% for HPALDs. The GEOS-Chem simulation of organic nitrates including ISOPN is further discussed in Fisher et al. (2016). Our HPALD source is based on the ISOPO\(_2\) isomerization rate constant from Crounse et al. (2011). A theoretical calculation by Peeters et al. (2014) suggests a rate constant that is 1.8× higher, which would reduce the model bias for HPALDs and ISOPOOH and increase boundary layer OH by 8%. St. Clair et al. (2015) found that the reaction rate of ISOPOOH + OH to form IEPOX is approximately 10% faster than the rate given by Paulot et al. (2009b), which would further reduce the model overestimate. For both ISOPOOH and HPALDs, GEOS-Chem captures much of the spatial variability (\(r = 0.80\) and 0.79, respectively).

Figure 7 shows the model branching ratios for the fate of the ISOPO\(_2\) radical by tracking the mass of ISOPO\(_2\) reacting via the high-NO\(_x\) path (ISOPO\(_2\) + NO) and the low-NO\(_x\) paths over the Southeast US domain. The mean branching ratios for the Southeast US are ISOPO\(_2\) + NO 54%, ISOPO\(_2\) + HO\(_2\) 26%, ISOPO\(_2\) isomerization 15%, and ISOPO\(_2\) + RO\(_2\) 5%. The lack of dominance of the high-NO\(_x\) path is due in part to the spatial segregation of isoprene and NO\(_x\) emissions (Yu et al., 2016). This segregation also buffers the effect of changing NO\(_x\) emissions on the fate of isoprene. Our original simulation with higher total NO\(_x\) emissions (unadjusted NEI11v1) had a branching ratio for the ISOPO\(_2\) + NO reaction of only 62%.

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6 Implications for ozone: aircraft and ozonesonde observations

Figure 2 compares simulated and observed median vertical profiles of ozone concentrations over the Southeast US during SEAC4RS. There is no significant bias through the depth of the tropospheric column. The median ozone concentration below 1.5 km is 49 ppb in the observations and 51 ppb in the model. We also find excellent model agreement across the US with the SEACIONS ozonesonde network (Fig. 8). The successful simulation of ozone is contingent on the decrease in NOx emissions. As shown in Fig. 2, a simulation with the original NEI emissions overestimates boundary layer ozone by 12 ppb.

The model also has success in reproducing the spatial variability of boundary layer ozone seen from the aircraft, as shown in Fig. 4. The correlation coefficient is $r = 0.71$ on the $0.25^\circ \times 0.3125^\circ$ model grid, and patterns of high and low ozone concentration are consistent. The highest observed ozone (> 75 ppb) was found in air influenced by agricultural burning along the Mississippi River and by outflow from Houston over Louisiana. GEOS-Chem does not capture the extreme values and this probably reflects a dilution effect (Yu et al., 2016).

A critical parameter for understanding ozone production is the ozone production efficiency (OPE) (Liu et al., 1987), defined as the number of ozone molecules produced per molecule of NOx emitted. This can be estimated from atmospheric observations by the relationship between odd oxygen ($O_3 \equiv O_3 + NO_2$) and the sum of products of NOx oxidation, collectively called NOx and including inorganic and organic nitrates (Trainor et al., 1993; Zaveri, 2003). The $O_3$ vs. NOx linear relationship (as derived from a linear regression) provides an upper estimate of the OPE because of rapid deposition of NOx, mainly HNO3 (Trainor et al., 2000; Rickard et al., 2002).

Figure 8 shows the observed and simulated daytime (09:00–16:00 local) $O_3$ vs. NOx relationship in the SEAC4RS data below 1.5 km, where NOx is derived from the observations as NOx – NOy = HNO3 + aerosol nitrate + PAN + alkyl nitrates. The resulting OPE from the observations (17.4 ± 0.4 mol mol$^{-1}$) agrees well with GEOS-Chem (16.7 ± 0.3 mol mol$^{-1}$). Previous work during the INTEX-NA aircraft campaign in summer 2004 found an OPE of 8 below 4 km (Men-Carrasco et al., 2007). By selecting INTEX-NA data only for the southeast and below 1.5 km we find an OPE of 14.1 ± 1.1 (Fig. 9, right panel). The median NOx was 1.1 ppb during SEAC4RS and 1.5 ppb during INTEX-NA, a decrease of approximately 40 %. With
INTEX-NA: $[Ox] = 29.0 + 14.1[NO_x]$, $r = 0.58$

SEAC4RS: $[Ox] = 23.9 + 17.4[NO_z]$, $r = 0.72$

GEOS-Chem: $[Ox] = 27.0 + 16.7[NO_z]$, $r = 0.83$

Figure 9. Ozone production efficiency (OPE) over the Southeast US in summer estimated from the relationship between odd oxygen ($O_x$) and the sum of NO$_x$ oxidation products (NO$_z$) below 1.5 km altitude. The left panel compares SEAC4RS observations to GEOS-Chem values for August–September 2013 (data from Fig. 2). The right panel compares SEAC4RS observations to INTEX-NA aircraft observations collected over the same Southeast US domain in summer 2004 (Singh et al., 2006). NO$_z$ is defined here as HNO$_3$ + aerosol nitrate + PAN + alkyl nitrates, all of which were measured from the SEAC4RS and INTEX-NA aircraft. The slope and intercept of the reduced-major-axis (RMA) regression are provided inset with the correlation coefficient ($r$). Observations for INTEX-NA were obtained from ftp://ftp-air.larc.nasa.gov/pub/INTEXA/.

Figure 10. Maximum daily 8 h average (MDA8) ozone concentrations at the 30 CASTNET sites in the Southeast US in June–August 2013. The left panels show seasonal mean values in the observations and GEOS-Chem. The right panel shows the probability density functions (pdfs) of daily values at the 30 sites.

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systematic bias. The median ozone concentration in the marine boundary layer is 26 ppb in the observations and 29 ppb in the model. This successful simulation is due to our adjustment of lightning NO\textsubscript{x} emission (Sect. 2.3); a sensitivity test with the original (twice higher) GEOS-Chem lightning emissions in the southern US increases surface ozone over the Gulf of Mexico by up to 6 ppb. The aircraft observations in Fig. 4 further show no indication of a coastal depletion that might be associated with halogen chemistry. Remarkably, the median ozone over the Gulf of Mexico is higher than approximately 8% of MDA\textsubscript{8} values at sites in the southeast.

It appears instead that there is a model bias in boundary layer vertical mixing and chemistry. Figure 12 shows the median ozonesonde profile at a higher vertical resolution over the Southeast US (Huntsville, Alabama, and St. Louis, Missouri, sites) during SEAC\textsuperscript{4}RS as compared to GEOS-Chem below 1.5 km. The ozonesondes indicate a decrease of 7 ppb from 1.5 km to the surface, whereas GEOS-Chem features a reverse gradient of increasing ozone from 1.5 to 1 km with flat concentrations below. This implies a combination of two model errors in the boundary layer: (1) excessive vertical mixing and (2) net ozone production whereas observations indicate net ozone loss.

8 Conclusions

We used aircraft (SEAC\textsuperscript{4}RS), surface, satellite, and ozonesonde observations from August and September 2013, interpreted with the GEOS-Chem chemical transport model, to better understand the factors controlling surface ozone in the Southeast US. Models tend to overestimate ozone in that region. Determining the reasons behind this overestimate is critical to the design of efficient emission control strategies to meet the ozone NAAQS.

A major finding from this work is that NEI1v1 for NO\textsubscript{x} (the limiting precursor for ozone formation) is biased high across the US by as much as a factor of 2. Evidence for this comes from (1) SEAC\textsuperscript{4}RS observations of NO\textsubscript{x} and its oxidation products, (2) NADP network observations of nitrate wet deposition fluxes, and (3) OMI satellite observations of NO\textsubscript{2}. Presuming no error in emissions from large power plants with continuous emission monitors (14% of unadjusted NEI inventory), we find that emissions from other industrial sources and mobile sources must be 30–60% lower than NEI values, depending on the assumption of the contribution from soil NO\textsubscript{x} emissions. We thus estimate that anthropogenic fuel NO\textsubscript{x} emissions in the US in 2013 were 1.7–2.6 Tg N\textsubscript{a}\textsuperscript{-1}, as compared to 3.5 Tg N\textsubscript{a}\textsuperscript{-1} given in the NEI.

OMI NO\textsubscript{2} satellite data over the Southeast US are consistent with this downward correction of NO\textsubscript{x} emissions but interpretation is complicated by the large contribution of the free troposphere to the NO\textsubscript{2} tropospheric column retrieved from the satellite. Observed (aircraft) and simulated vertical profiles indicate that NO\textsubscript{2} below 2 km contributes only 20–35% of the tropospheric column detected from space while NO\textsubscript{2} above 8 km (mainly from lightning) contributes 25–50%. Current retrievals of satellite NO\textsubscript{2} data do not properly account for this elevated pool of upper-tropospheric NO\textsubscript{2}, so that the reported tropospheric NO\textsubscript{2} columns are biased high. More work is needed on the chemistry maintaining high levels of NO\textsubscript{2} in the upper troposphere.

Isoprene emitted by vegetation is the main VOC precursor of ozone in the southeast in summer, but we find that only 50% reacts by the high-NO\textsubscript{x} pathway to produce ozone. This is consistent with detailed aircraft observations of isoprene oxidation products from the aircraft. The high-NO\textsubscript{x}}
fraction is only weakly sensitive to the magnitude of NO\(_x\) emissions because isoprene and NO\(_x\) emissions are spatially segregated. The ability to properly describe high- and low-NO\(_x\) pathways for isoprene oxidation is critical for simulating ozone and it appears that the GEOS-Chem mechanism is successful for this purpose.

Our updated GEOS-Chem simulation with decreased NO\(_x\) emissions provides an unbiased simulation of boundary layer and free-tropospheric ozone measured from aircraft and ozonesondes during SEAC\(^4\)RS. Decreasing NO\(_x\) emissions is critical to this success as the original model with NEI emissions overestimated boundary layer ozone by 12 ppb. The ozone production efficiency (OPE) inferred from O\(_3\) vs. NO\(_x\) aircraft correlations in the mixed layer is also well reproduced. Comparison to the INTEX-NA aircraft observations over the southeast in summer 2004 indicates a 14 % increase in OPE associated with a 40 % reduction in NO\(_x\) emissions.

Despite the successful simulation of boundary layer ozone (Figs. 2 and 9), GEOS-Chem overestimates MDA8 surface ozone observations in the Southeast US in summer by 6 ± 14 ppb. Daytime ozonesonde data indicate a 7 ppb decrease from 1.5 km to the surface that GEOS-Chem does not capture. This may be due to excessive boundary layer mixing and net ozone production in the model. Excessive mixing in GEOS-Chem may be indicative of an overestimate of sensible heat flux (Holtslag and Boville, 1993), and thus an investigation of boundary layer meteorological variables is warranted. Such a bias may not be detected in the comparison of GEOS-Chem with aircraft data, generally collected under fair-weather conditions with minimal sampling in the lower part of the boundary layer. An investigation of relevant meteorological variables and boundary layer source and sink terms in the ozone budget to determine the source of bias and its prevalence across models will be the topic of a follow-up paper.

9 Data availability

The SEAC\(^4\)RS airborne trace gas and particle measurements and SEACIONS ozonesonde measurements are available from the NASA LaRC Airborne Science Data for Atmospheric Composition (http://www-air.larc.nasa.gov/cgi-bin/ArcView/seac4rs) with doi:10.5067/Aircraft/SEAC4RS/Aerosol-TraceGas-Cloud.

Observations for INTEX-NA were also obtained from NASA LaRC (http://www-air.larc.nasa.gov/cgi-bin/ArcView/intexna) with doi:10.5067/Aircraft/INTEXA/Aerosol-TraceGas.

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References

Air Sciences, Inc.: 2002 Fire Emission Inventory for the WRAP Region – Phase II, Western Governors Association/Western Regional Air Partnership, Denver and Portland, 2005.

Alvarado, M. J., Logan, J. A., Mao, J., Apel, E., Riemer, D., Blake, D., Cohen, R. C., Min, K.-E., Perring, A. E., Browne, E. C., Wooldridge, P. J., Diskin, G. S., Sachse, G. W., Fuelberg, H., Sessions, W. R., Harrigan, D. L., Huey, G., Liao, J., Case-Hanks, A., Jimenez, J. L., Cubison, M. J., Vay, S. A., Weinghein, A. J., Knapp, D. J., Montzka, D. D., Flocke, F. M., Pollack, I. B., Wennberg, P. O., Kurten, A., Crounse, J., Clair, J. M. St., Wisthaler, A., Mikoviny, T., Yantosca, R. M., Carouge, C. C., and Le Sager, P.: Nitrogen oxides and PAN in plumes from boreal fires during ARCTAS-B and their impact on ozone: an integrated analysis of aircraft and satellite observations, Atmos. Chem. Phys., 10, 9739–9760, doi:10.5194/acp-10-9739-2010, 2010.

Anderson, D. C., Loughner, C. P., Diskin, G., Weinghein, A., Canty, T., P., Salawitch, R. J., Worden, H. M., Fried, A., Mikoviny, T., Wisthaler, A., and Dickerson, R. R.: Measured and modeled CO and NO\(_x\) in DISCOVER-AQ: An evaluation of emissions and chemistry over the eastern US, Atmos. Environ., 96, 78–87, doi:10.1016/j.atmosenv.2014.07.004, 2014.

Bates, K. H., Crounse, J. D., St Clair, J. M., Bennett, N. B., Nguyen, T. B., Seinfeld, J. H., Stoltz, B. M., and Wennberg, P. O.: Gas Phase Production and Loss of Isoprene Epoxydiols, J. Phys. Chem. A, 118, 1237–1246, doi:10.1021/jp4107958, 2014.
Bucsela, E. J., Krotkov, N. A., Celarier, E. A., Lamsal, L. N., Brown-Steiner, B., Hess, P. G., and Lin, M. Y.: On the capabilities of GCCM simulations in the US using CESM CAM-Chem, Atmos. Environ., 105, 134–148, doi:10.1016/j.atmosenv.2014.11.001, 2015.

Bey, I., Jacob, D. J., Yantosca, R. M., Logan, J. A., Field, B. D., Fiore, A. M., Li, Q. B., Liu, H. G. Y., Mickley, J. L., and Schultz, M. G.: Global modeling of tropospheric chemistry with assimilated meteorology: Model description and evaluation, J. Geophys. Res.-Atmos., 106, 23073–23095, doi:10.1029/2001jd000807, 2001.

Borders, R. A., and Birks, J. W.: High-precision measurements of activation energies over small temperature intervals: Curvature in the Arrhenius plot for the reaction NO + O₃ → NO₂ + O₂, J. Phys. Chem. A, 86, 3295–3302, 1982.

Brioude, J., Angevine, W. M., Ahmadov, R., Kim, S.-W., Evan, S., Bey, I., Jacob, D. J., Yantosca, R. M., Logan, J. A., Field, B. D., Fiore, A. M., Li, Q. B., Liu, H. G. Y., Mickley, J. L., and Schultz, M. G.: Global modeling of tropospheric chemistry with assimilated meteorology: Model description and evaluation, J. Geophys. Res.-Atmos., 106, 23073–23095, doi:10.1029/2001jd000807, 2001.

Brown-Steiner, B., Hess, P. G., and Lin, M. Y.: Ozone, oxides of nitrogen, and carbon monoxide during pollution events over the eastern United States: An evaluation of emissions and vertical mixing, J. Geophys. Res., 116, D16307, doi:10.1029/2010JD014540, 2011.

Chen, D., Wang, Y., McElroy, M. B., He, K., Yantosca, R. M., and Le Sager, P.: Regional CO pollution and export in China simulated by the high-resolution nested-grid GEOS-Chem model, Atmos. Chem. Phys., 9, 3825–3839, doi:10.5194/acp-9-3825-2009, 2009.

Cohen, R. C., Perkins, K. K., Koch, L. C., Stimpfle, R. M., Weng, B. P., Hanisco, T. F., Lanzendorf, E. J., Bonne, G. P., Voss, P. B., Salawitch, R. J., Del Negro, L. A., Wilson, J. C., McElroy, C. T., and Bui, T. P.: Quantitative constraints on the atmospheric chemistry of nitrogen oxides: An analysis along chemical coordinates, J. Geophys. Res., 105, 24283–24304, 2000.

Crounse, J. D., McKinney, K. A., Kwan, A. J., and Wennberg, P. O.: Measurement of gas-phase hydroperoxides by chemical ionization mass spectrometry (CIMS), Anal. Chem., 78, 6726–6732, 2006.

Crounse, J. D., Paulot, F., Kjaergaard, H. G., and Wennberg, P. O.: Peroxy radical isomerization in the oxidation of isoprene, Phys. Chem. Chem. Phys., 13, 13607–13613, doi:10.1039/c1cp21330j, 2011.

Darmenov, A. and da Silva, A.: The Quick Fire Emissions Dataset (QFED) – Documentation of versions 2.1, 2.2 and 2.4, NASA Technical Report Series on Global Modeling and Data Assimilation, NASA TM-2013-104606, 32, 183 pp., Draft Document (12939 kb), 2013.

Dibb, J. E., Talbot, R. W., Scheuer, E. M., Seid, G., Avery, M. A., and Singh, H. B.: Aerosol chemical composition in Asian continental outflow during the TRACE-P campaign: Comparison with PEM-West B, J. Geophys. Res., 108, 8815, doi:10.1029/2002jd002748, 2003.

Dibb, J. E., Talbot, R. W., Scheuer, E. M., Seid, G., Avery, M. A., and Singh, H. B.: Aerosol chemical composition in Asian continental outflow during the TRACE-P campaign: Comparison with PEM-West B, J. Geophys. Res., 108, 8815, doi:10.1029/2002jd002748, 2003.

Del Negro, L. A., Fahey, D. W., Gao, R. S., Donnelly, S. G., Keim, E. R., Neuman, J. A., Cohen, R. C., Perkins, K. K., Koch, L. C., Salawitch, R. J., Lloyd, S. A., Proffitt, M. H., Margitan, J. J., Stimpfle, R. M., Bonne, G. P., Voss, P. B., Wennberg, P. O., McElroy, C. T., Swartz, W. H., Kusterer, T. L., Anderson, D. E., Lait, L. R., and Bui, T. P.: Comparison of modeled and observed values of NO₂ and JNO₂ during the Photochemistry of Ozone Loss in the Arctic Region in Summer (POLARIS) mission, J. Geophys. Res., 104, 26687, doi:10.1029/1999jd900246, 1999.

Duncan, B. N., Prados, A. I., Lamsal, L. N., Liu, Y., Streets, D. G., Gupta, P., Hilsenrath, E., Kahn, R. A., Nielsen, J. E., Beyersdorf, A. J., Burton, S. P., Fiore, A. M., Fishman, J., Henze, D. K., Hostetler, C. A., Krotkov, N. A., Lee, P., Lin, M., Pawson, S., Pfister, G., Pickering, K. E., Pierce, R. B., Yoshida, Y., and Ziemba, L. D.: Satellite data of atmospheric pollution for U.S. air quality applications: Examples of applications, summary of data end-user resources, answers to FAQs, and common mistakes to avoid, Atmos. Environ., 94, 647–662, doi:10.1016/j.atmosenv.2014.05.061, 2014.

EPA: Integrated Science Assessment for Ozone and Related Photochemical Oxidants, US Environmental Protection Agency, Research Triangle Park, NC, 2013.

EPA: US Environmental Protection Agency Clean Air Markets Division Clean Air Status and Trends Network (CASTNET)
Fujita, E. M., Campbell, D. E., Zielinska, B., Chow, J. C., Lind- 
Fisher, J. A., Jacob, D. J., Travis, K. R., Kim, P. S., Marais, E. 
Fiore, A. M., Horowitz, L. W., Purves, D. W., Levy, H., Evans, M. 
Fiore, A. M., Jacob, D. J., Liu, H., Yantosca, R. M., Fairlie, T. D., 
Finkelstein, P. L., Ellestad, T. G., Clarke, J. F., Meyers, T. P., 
Schwe, D. B., Hebert, E. O., and Neal, J. A.: Ozone and sulfur dioxide dry deposition to forests: Observations and model evaluation, J. Geophys. Res.-Atmos., 105, 15365–15377, doi:10.1029/2000jd901885, 2000. 

Fiore, A. M., Jacob, D. J., Liu, H., Yantosca, R. M., Fairlie, T. D., and Li, Q.: Variability in surface ozone background over the United States: Implications for air quality policy, J. Geophys. Res.-Atmos., 108, 4787, doi:10.1029/2003jd003855, 2003. 

Fiore, A. M., Horowitz, L. W., Purves, D. W., Levy, H., Evans, M. J., Wang, Y., Li, Q., and Yantosca, R.: Evaluating the contribution of changes in isoprene emissions to surface ozone trends over the eastern United States, J. Geophys. Res., 110, D12303, doi:10/2009jd045485, 2005. 

Fiore, A. M., Dentener, F. J., Wild, O., Cuvelier, C., Schultz, M. G., Hess, P., Teutkr, C., Schulz, M., Doherty, R. M., Horowitz, L. W., MacKenzie, I. A., Sanderson, M. G., Shindell, D. T., Stevenson, D. S., Szopa, S., Van Dingenen, R., Zeng, G., Ather- ton, C., Bergmann, D., Bey, I., Carmichael, G., Collins, W. J., Dunnaren, B. N., Faluveci, G., Golberth, G., Gauss, M., Gong, S., Hauglustaine, D., Holloway, T., Isaksen, I. S. A., Jacob, D. Jacob, Jonson, J. E., Kaminiski, J. W., Keating, T. J., Lupu, A., Marmer, E., Montanaro, V., Park, R. J., Pitari, G., Pringle, K. J., Pyle, J. A., Schroeder, S., Vivanco, M. G., Wind, P., Wojcik, G., Wu, S., and Zuber, A.: Multimodel estimates of intercontinental source-receptor relationships for ozone pollution, J. Geophys. Res., 114, D04301, doi:10.1029/2008jd010816, 2009. 

Fischer, E. V., Jacob, D. J., Yantosca, R. M., Sulprizio, M. P., Milet, D. B., Mao, J., Paulot, F., Singh, H. B., Roiger, A., Ries, L., Talbot, R. W., Dzepina, K., and Pandey Deolal, S.: Atmospheric peroxyacetyl nitrate (PAN): a global budget and source attribution, Atmos. Chem. Phys., 14, 2679–2698, doi:10.5194/acp-14-2679-2014, 2014. 

Fisher, A. J., Jacob, D. J., Travis, K. R., Kim, P. S., Marais, E. A., Chan Miller, C., Yu, K., Zhu, L., Yantosca, R. M., Sulprizio, M. P., Mao, J., Wennberg, P. O., Crouse, J. D., Teng, A. P., Nguyen, T. B., St. Clair, J. M., Cohen, R. C., Romer, P., Nault, B. A., Wooldridge, P. J., Jimenez, J. L., Campuzano-Jost, P., Day, D. A., Hu, W., Shepson, P. B., Xiong, F., Blake, D. R., Goldstein, A. H., Misztal, P. K., Hanisco, T. F., Wolfe, G. M., Ryerson, T. B., Wisthaler, A., and Mikoviny, T.: Organic nitrogen chemistry and its implications for nitrogen budgets in an isoprene- and monoterpen-rich atmosphere: constraints from aircraft (SEAC4RS) and ground-based (SOAS) observations in the Southeast US, Atmos. Chem. Phys., 16, 5969–5991, doi:10.5194/acp-16-5969-2016, 2016. 

Fujita, E. M., Campbell, D. E., Zielinska, B., Chow, J. C., Lind- hjem, C. E., DenBleyker, A., Bishop, G. A., Schuchmann, B. G., Stedman, D. H., and Lawson, D. R.: Comparison of the MOVES2010a, MOBILE6.2, and EMFAC2007 mobile source emission models with on-road traffic tunnel and remote sensing measurements, J. Air Waste Manage., 62, 1134–1149, doi:10.1080/10962247.2012.699016, 2012. 

Holstlag, A. and Bovile, B.: Local versus nonlocal boundary-layer diffusion in a global climate model, J. Climate, 6, 1825–1842, 1993. 

Horowitz, L. W., Fiore, A. M., Milly, G. P., Cohen, R. C., Pering, A., Wooldridge, P. J., Hess, P. G., Emmons, L. K., and Lamarque, J. F.: Observational constraints on the chemistry of isoprene nitrates over the eastern United States, J. Geophys. Res.-Atmos., 112, D12S08, doi:10.1029/2006jd007747, 2007. 

Hudman, R. C., Jacob, D. J., Turquety, S., Leibensperger, E. M., Murray, T. L., Wu, S., Gilliland, A. B., Avery, M., Bertram, T. H., Brune, W., Cohen, R. C., Dibb, J. E., Flocke, F. M., Fried, A., Holloway, J., Neuman, J. A., Orville, R., Perring, A., Ren, X., Sachse, G. W., Swanson, A., and Wooldridge, P. J.: Surface and lightning sources of nitrogen oxides over the United States: Magnitudes, chemical evolution, and outflow, J. Geophys. Res., 112, D12S05, doi:10.1029/2006jd007912, 2007. 

Hudman, R. C., Moore, N. E., Mebust, A. K. Martin, R. V., Russell, A. R., Valin, L. C., and Cohen, R. C.: Steps towards a mechanistic model of global soil nitric oxide emissions: implementation and space-based constraints, Atmos. Chem. Phys., 12, 7779–7795, doi:10.5194/acp-12-7779-2012, 2012. 

Huntrieser, H., Schlager, H., Roiger, A., Lichtenstern, M., Schu- mann, U., Kurz, C., Brunner, D., Schwierz, C., Richter, A., and Stohl, A.: Lightning-produced NOx over Brazil during TROC- CINOX: airborne measurements in tropical and subtropical thunderstorms and the importance of mesoscale convective systems, Atmos. Chem. Phys., 7, 2987–3013, doi:10.5194/acp-7-2987-2007, 2007. 

Huntrieser, H., Schumann, U., Schlager, H., Höller, H., Giez, A., Betz, H.-D., Brunner, D., Forster, C., Pinto Jr., O., and Calheiros, R.: Lightning activity in Brazilian thunderstorms during TROC- CINOX: implications for NOx production, Atmos. Chem. Phys., 8, 921–953, doi:10.5194/acp-8-921-2008, 2008. 

Jaeglé, L., Jacob, D. J., Wennberg, P. O., Spivakovsky, C. M., Hanisco, T. F., Lanzendorf, E. L., Hintsa, E. J., Fahey, D. W., Keim, E. R., Proffitt, M. H., Atlas, E., Flocke, F., Schaufler, S., McElroy, C. T., Midwinter, C., Pfister, L., and Wilson, J. C.: Observed OH and HO2 in the upper troposphere suggest a ma- jor source from convective injection of peroxides, Geophys. Res. Lett., 24, 3181–3184, 1997. 

Jaeglé, L., Jacob, D. J., Wang, Y., Weinheimer, A. J. Ridley, B. A., Campos, T. L., Sachse, G. W., and Hagen, D. E.: Sources and chemistry of NOx in the upper troposphere over the United States, Geophys. Res. Lett., 25, 1705–1708, 1998. 

Kim, P. S., Jacob, D. J., Fisher, A. J., Travis, K. Yu, K., Zhu, L., Yanosca, R. M., Sulprizio, M. P., Jimenez, J. L., Campuzano-Jost, P., Froyd, K. D., Liao, J., Hair, J. W., Fenn, M. A., But- ler, C. F., Wagner, N. L., Gordon, T. D., Wetti, A., Wennberg, P. O., Crouse, J. D., St. Clair, J. M., Teng, A. P., Millet, D. B., Schwarz, J. P., Markovic, M. Z., and Perring, A. E.: Sources, seasonality, and trends of southeast US aerosol: an integrated analysis of surface, aircraft, and satellite observations with the GEOS-Chem chemical transport model, Atmos. Chem. Phys., 15, 10411–10433, doi:10.5194/acp-15-10411-2015, 2015. 

Lamsal, L. N., Krotkov, N. A., Celarier, E. A., Swartz, W. H., Pick- ering, K. E., Bucsela, E. J., Gleason, J. F., Martin, R. V., Philip, S., Irie, H., Cede, A., Herman, J., Weinheimer, A., Szykman, J. J., and Knepp, T. N.: Evaluation of OMI operational standard NO2 column retrievals using in situ and surface-based NO2 observa-
Mao, J., Jacob, D. J., Park, R., Wang, Y., Heald, C. L., Hudson, R., and Yantosca, R. M.: North American pollution outflow and the trapping of convectively lifted pollution by upper-level anticyclone. J. Geophys. Res., 110, D10301, doi:10.1029/2004JD005039, 2005.

Li, X., Rohrer, F., Hofzumahaus, A., Brauers, T., Haseler, R., Bohn, B., Broch, S., Fuchs, H., Gomm, S., Holland, F., Jager, K., Kaiser, J., Keutsch, F. N., Lohse, I., Lu, K., Tillmann, R., Wegener, R., Wolfe, G. M., Mentel, T. F., Kiendler-Scharr, A., and Wahner, A.: Missing gas-phase source of HONO inferred from Zepelin measurements in the troposphere, Science, 344, 292–296, doi:10.1126/science.1248999, 2014.

Lin, J., Youn, D., Liang, X., and Wuebbles, D.: Global model simulation of summertime U.S. ozone diurnal cycle and its sensitivity to PBL mixing, spatial resolution, and emissions. Atmos. Environ., 42, 8470–8483, doi:10.1016/j.atmosenv.2008.08.012, 2008.

Lin, J.-T. and McElroy, M. B.: Impacts of boundary layer mixing on pollutant vertical profiles in the lower troposphere: Implications to satellite remote sensing. Atmos. Environ., 44, 1726–1739, doi:10.1016/j.atmosenv.2010.02.009, 2010.

Liu, S. C., Trainer, M., Fehsenfeld, F. C., Parrish, D. D., Williams, E. J., Fahey, D. W., Hubler, G., and Murphy, P. C.: Ozone Production in the Rural Troposphere and the Implications for Regional and Global Ozone Distributions, J. Geophys. Res., 92, 4191–4207, 1987.

Liu, Y. J., Herdlinger-Blatt, I., McKinney, K. A., and Martin, S. T.: Production of methyl vinyl ketone and methacrolein via the hydroperoxy pathway of isoprene oxidation. Atmos. Chem. Phys., 13, 5715–5730, doi:10.5194/acp-13-5715-2013, 2013.

Lu, Z., Streets, D. G., de Foy, B., Lamsls, L. N., Duncan, B. N., and Xing, J.: Emissions of nitrogen oxides from US urban areas: estimation from Ozone Monitoring Instrument retrievals for 2005–2014. Atmos. Chem. Phys., 15, 10367–10383, doi:10.5194/acp-15-10367-2015, 2015.

Mao, J., Jacob, D. J., Evans, M. J., Olson, J. R., Ren, X., Brune, W. H., Clair, J. M. St., Crounse, J. D., Spencer, K. M., Beaver, M. R., Wennberg, P. O., Cubison, M. J., Jimenez, J. L., Fried, A., Weinbrin, P., Walega, J. G., Hall, S. R., Weinheimer, A. J., Cohen, R. C., Chen, G., Crawford, J. H., McNaughton, C., Clarke, A. D., Jaegle, L., Fisher, J. A., Yantosca, R. M., Le Sager, P., and Carouge, C.: Chemistry of hydrogen oxide radicals (HOx) in the Arctic troposphere in spring. Atmos. Chem. Phys., 10, 5823–5838, doi:10.5194/acp-10-5823-2010, 2010.

Mao, J., Paulot, F., Jacob, D. J., Cohen, R. C., Crounse, J. D., Wennberg, P. O., Keller, C. A., Hudman, R. C., Barkley, M. P., and Horowitz, L. W.: Ozone and organic nitrates over the eastern United States: Sensitivity to isoprene chemistry. J. Geophys. Res.-Atmos., 118, 11256–11268, doi:10.1002/jgrd.50817, 2013.

Marais, E. A., Jacob, D. J., Jimenez, J. L., Campuzano-Jost, P., Day, D. A., Hu, W., Kreczmer, J., Zhu, L., Kim, P. S., Miller, C. C., Fisher, J. A., Travis, K., Yu, K., Hanisco, T. F., Wolfe, G. M., Arkinson, H. L., Pye, H. O. T., Froyd, K. D., Liao, J., and McNeill, V. F.: Aqueous-phase mechanism for secondary organic aerosol formation from isoprene: application to the south-east United States and co-benefit of SO2 emission controls, Atmos. Chem. Phys., 16, 1603–1618, doi:10.5194/acp-16-1603-2016, 2016.
Paulot, F., Crounse, J. D., Kjaergaard, H. G., Kurten, A., St. Clair, J. M., Seinfeld, J. H., and Wennberg, P. O.: Unexpected Epoxide Formation in the Gas-Phase Photooxidation of Isoprene, Science, 325, 730–733, doi:10.1126/science.1172910, 2009b.

Paulot, F., Jacob, D. J., Pinder, R. W., Bash, J. O., Travis, K., and Henze, D. K.: Ammonia emissions in the United States, European Union, and China derived by high-resolution inversion of ammonium wet deposition data: Interpretation with a new agricultural emissions inventory (MASAGE_NH3), J. Geophys. Res.-Atmos., 119, 4343–4364, doi:10.1002/2013jd021130, 2014.

Peeters, J. and Müller, J. F.: HO radical regeneration in isoprene oxidation via peroxy radical isomerisations. II: experimental evidence and global impact, Phys. Chem. Chem. Phys., 12, 14227–14235, doi:10.1039/c0cp00811g, 2010.

Peeters, J., Nguyen, T. L., and Vereecken, L.: HO radical regeneration in the oxidation of isoprene, Phys. Chem. Chem. Phys., 11, 5935–5939, doi:10.1039/b908511d, 2009.

Peeters, J., Müller, J. F., Stavrakou, T., and Nguyen, V. S.: Hydroxyl radical recycling in isoprene oxidation driven by hydrogen bonding and hydrogen tunneling: the upgraded LIM1 mechanism, J. Phys. Chem. A, 118, 8625–8643, doi:10.1021/jp5033146, 2014.

Pollack, J. B., Lerner, B. M., and Ryerson, T. B.: Evaluation of ultraviolet light-emitting diodes for detection of atmospheric NO2 by photolysis – chemiluminescence, J. Atmos. Chem., 65, 111–125, doi:10.1007/s10874-011-9184-3, 2010.

Prather, M. J. and Jacob, D. J.: A persistent imbalance in HOx and NOx photochemistry of the upper troposphere driven by deep tropical convection, Geophys. Res. Lett., 24, 3189–3192, 1997.

Reed, C., Evans, M. J., Di Carlo, P., Lee, J. D., and Carpenter, L. J.: Interferences in photolytic NO2 measurements: explanation for an apparent missing oxidant?, Atmos. Chem. Phys., 16, 4707–4724, doi:10.5194/acp-16-4707-2016, 2016.

Reidmiller, D. R., Fiore, A. M., Jaffe, D. A., Bergmann, D., Cuvier, C., Dentener, F. J., Duncan, B. N., Folberth, G., Gauss, M., Gong, S., Hess, P., Jonson, J. E., Keating, T., Lupu, A., Marmer, E., Park, R., Schultz, M. G., Shindell, D. T., Szopa, S., Vivanco, M. G., Wild, O., and Zuber, A.: The influence of foreign vs. North American emissions on surface ozone in the US, Atmos. Chem. Phys., 9, 5027–5042, doi:10.5194/acp-9-5027-2009, 2009.

Rickard, A. R., Salisbury, G., Monks, P. S., Lewis, A. C., Baugitte, S., Bandy, B. J., Clemitshaw, K. C., and Penkett, S. A.: Comparison of Measured Ozone Production Efficiencies in the Marine Boundary Layer at Two European Coastal Sites under Different Pollution Regimes, J. Atmos. Chem., 43, 107–134, 2002.

Russell, A. R., Perring, A. E., Valin, L. C., Bucsela, E. J., Browne, E. C., Wooldridge, P. J., and Cohen, R. C.: A high spatial resolution retrieval of NO2 column densities from OMI: method and evaluation, Atmos. Chem. Phys., 11, 8543–8554, doi:10.5194/acp-11-8543-2011, 2011.

Ryerson, T. B., Buhr, M. P., Frost, G. J., Goldan, P. D., Holloway, J. S., Hübner, G., Jobson, B. T., Kuster, W. C., McKeeen, S. A., Parrish, D. D., Roberts, J. M., Suarez, D. T., Trainer, M., Williams, J., and Fehsenfeld, F. C.: Emissions lifetimes and ozone formation in power plant plumes, J. Geophys. Res., 103, 22569–22583, 1998.

Ryerson, T. B., Williams, E. J., and Fehsenfeld, F. C.: An efficient photolysis system for fast-response NO2 measurements, J. Geophys. Res., 105, 26447, doi:10.1029/2000jd900389, 2000.

Sander, S. P., Abbatt, J., Barker, J. R., Burkholler, J. B., Friedl, R. R., Golden, D. M., Huie, R. E., Kolb, C. E., Kurylo, M. J., Moortgat, G. K., Orkin, V. L., and Wine, P. H.: Chemical Kinetics and Photochemical Data for Use in Atmospheric Studies, Evaluation No. 17, JPL Publication 10-6, Jet Propulsion Laboratory, Pasadena, 2011.

Shultz, M. G., Jacob, D. J., Wang, Y., Logan, J. A., Atlas, E. L., Blake, D. R., Blake, N. J., Bradshaw, J. D., Browell, E. V., Fenn, M. A., Flocke, F., Gregory, G. L., Heikes, B. G., Sachse, G. W., Sandholm, S. T., Shetter, R. E., Singh, H. B., and Talbot, R. W.: On the origin of tropospheric ozone and NOx over the tropical South Pacific, J. Geophys. Res., 104, 5829, doi:10.1029/98jd02309, 1999.

Shetter, R. E. and Muller, M.: Photolysis frequency measurements using actinic flux spectroradiometry during the PEM-Tropics mission: Instrumentation description and some results, J. Geophys. Res., 104, 5647–5661, doi:10.1029/98JD01381, 1999.

Singh, H. B., Brune, W. H., Crawford, J. H., Jacob, D. J., and Russell, P. B.: Overview of the summer 2004 Intercontinental Chemical Transport Experiment–North America (INTEX-A), J. Geophys. Res., 111, doi:10.1029/2006jd007905, 2006.

Squire, O. J., Archibald, A. T., Griffiths, P. T., Jenkin, M. E., Smith, D., and Pyle, J. A.: Influence of isoprene chemical mechanism on modelled changes in tropospheric ozone due to climate and land use over the 21st century, Atmos. Chem. Phys., 15, 5123–5143, doi:10.5194/acp-15-5123-2015, 2015.

St. Clair, J. M., McCabe, D. C., Crounse, J. D., Steiner, U., and Wennberg, P. O.: Chemical ionization tandem mass spectrometer for the in situ measurement of methyl hydrogen peroxide, Rev. Sci. Instrum., 81, 094102, doi:10.1063/1.3480552, 2010.

St. Clair, J. M., Rivera-Rios, J. C., Crounse, J. D., Knap, H. C., Bates, K. H., Teng, A. P., Jorgensen, S., Kjaergaard, H. G., Keutsch, F. N., and Wennberg, P. O.: Kinetics and Products of the Reaction of the First-Generation Isoprene Hydroxy Hydroperoxide (ISOPOOH) with OH, J. Phys. Chem. A, 120, 1441–1451, doi:10.1021/acs.jpca.5b06532, 2015.

Stavrakou, T., Peeters, J., and Müller, J.-F.: Improved global modelling of HOx recycling in isoprene oxidation: evaluation against the GABRIEL and INTEX-A aircraft campaign measurements, Atmos. Chem. Phys., 10, 9863–9878, doi:10.5194/acp-10-9863-2010, 2010.

Toon, O. B., Maring, H., Dibb, J., Ferrare, R., Jacob, D. J., Jensen, E. J., Luo, Z. J., Mace, G. G., Pan, L. L., Pfister, L., and Rosenlof, K. H.: Planning, implementation, and scientific goals of the Studies of Emissions and Atmospheric Composition, Clouds and Climate Coupling by Regional Surveys (SEAC4RS) field mission, J. Geophys. Res.-Atmos., 121, 4967–5009, doi:10.1002/2015JD024297, 2016.

Trainer, M., Parrish, D. D., Buhr, M. P., Norton, R. B., Fehsenfeld, F. C., Anlauf, K. G., Bottenheim, J. W., Tang, Y. Z., Wiebe, H. A., Roberts, J. M., Tanner, R. L., Newman, L., Bowersox, C., Meagher, J. F., Olszyna, K. J., Rodgers, M. O., Wang, T., Berresheim, H., Demerjian, K. L., and Roychowdhury, U. K.: Correlation of ozone with NOx in photochemically aged air, J. Geophys. Res., 98, 2917–2925, 1993.

Atmos. Chem. Phys., 16, 13561–13577, 2016

www.atmos-chem-phys.net/16/13561/2016/
Trainer, M., Parrish, D. D., Goldan, P. D., Roberts, J., and Fehsenfeld, F. C.: Review of observation-based analysis of the regional factors influencing ozone concentrations, Atmos. Environ., 34, 2045–2061, 2000.

Vinken, G. C. M., Boersma, K. F., Maasakkers, J. D., Adon, M., and Martin, R. V.: Worldwide biogenic soil NO x emissions inferred from OMI NO 2 observations, Atmos. Chem. Phys., 14, 10363–10381, doi:10.5194/acp-14-10363-2014, 2014.

Walker, T. W.: Applications of Adjoint Modeling in Chemical Composition: Studies of Tropospheric Ozone at Middle and High Northern Latitudes, Graduate Department of Physics, University of Toronto, 2014.

Wang, Y., Jacob, D. J., and Logan, J. A.: Global simulation of tropospheric O 3 -NO x -hydrocarbon chemistry. 1. Model formulation, J. Geophys. Res., 103, 10727–10755, doi:10.1029/97JD03144, 2003.

Zaveri, R. A.: Ozone production efficiency and NO x depletion in an urban plume: Interpretation of field observations and implications for evaluating O 3 -NO x -VOC sensitivity, J. Geophys. Res., 108, 4436, doi:10.1029/2002jd003144, 2003.

Zhou, X., Ye, C., Pu, D., Stutz, J., Festa, J., Spolaor, M., Weinheimer, A. J., Campos, T. L., Haggerty, J. A., Cantrell, C. A., Mauldin, L., Guenther, A. B., Hornbrook, R. S., Apel, E. C., and Jensen, J. B.: Tropospheric HONO Distribution and Chemistry in the Southeastern US, American Geophysical Union, Fall Meeting 2014, abstract #A31J-08, 2014.

Yu, K., Jacob, D. J., Fisher, J. A., Kim, P. S., Marais, E. A., Miller, C. C., Travis, K. R., Zhu, L., Yantosca, R. M., Sulprizio, M. P., Cohen, R. C., Dibb, J. E., Fried, A., Mikoviny, T., Ryerson, T. B., Wennberg, P. O., and Wisthaler, A.: Sensitivity to grid resolution in the ability of a chemical transport model to simulate observed oxidant chemistry under high-isoprene conditions, Atmos. Chem. Phys., 16, 4369–4378, doi:10.5194/acp-16-4369-2016, 2016.

Zaveri, R. A.: Parameterization of Surface Resistances to Gaseous Dry Deposition in Regional-Scale Numerical-Models, Atmos. Environ., 23, 1293–1304, doi:10.1016/0004-6981(89)90153-4, 1989.

Wolfe, G. M., Crounse, J. D., Parrish, J. D., St Clair, J. M., Beaver, M. R., Paulot, F., Yoon, T. P., Wiberg, P. O., and Keutsch, F. N.: Photolysis, OH reactivity and ozone reactivity of a proxy for isoprene-derived peroxyalkyl nitrate (HPANs), Phys. Chem. Chem. Phys., 14, 7276–7286, doi:10.1039/c2cp40388a, 2012.

Zhang, L., Jacob, D. J., Yue, X., Downey, N. V., Wood, D. A., and Blewitt, D.: Sources contributing to background surface ozone in the US Intermountain West, Atmos. Chem. Phys., 14, 5295–5309, doi:10.5194/acp-14-5295-2014, 2014.

Zhu, L., Jacob, D. J., Kim, P. S., Fisher, J. A., Yu, K., Travis, K. R., Mickley, L. J., Yantosca, R. M., Sulprizio, M. P., De Smedt, I., Gonzalez Abad, G., Chance, K., Li, C., Ferrare, R., Fried, A., Hair, J. W., Hanisco, T. F., Richter, D., Scarno, A. J., Walega, J., Weibring, P., and Wolfe, G. M.: Observing atmospheric formaldehyde (HCHO) from space: validation and intercomparison of six retrievals from four satellites (OMI, GOME2A, GOME2B, OMPS) with SEAC4RS aircraft observations over the Southeast US, Atmos. Chem. Phys. Discuss., doi:10.5194/acp-2016-162, in review, 2016.