Monitoring the Reductive and Oxidative Half-Reactions of a Flavin-Dependent Monoxygenase using Stopped-Flow Spectrophotometry

Elvira Romero¹, Reeder Robinson¹, Pablo Sobrado¹

¹Department of Biochemistry, Virginia Polytechnic Institute and State University

Correspondence to: Pablo Sobrado at psobrado@vt.edu

URL: http://www.jove.com/video/3803
DOI: doi:10.3791/3803

Keywords: Bioengineering, Issue 61, Stopped-flow, kinetic mechanism, SidA, C4a-hydroperoxyflavin, monoxygenase, Aspergillus fumigatus

Date Published: 3/18/2012

Citation: Romero, E., Robinson, R., Sobrado, P. Monitoring the Reductive and Oxidative Half-Reactions of a Flavin-Dependent Monoxygenase using Stopped-Flow Spectrophotometry. J. Vis. Exp. (61), e3803, doi:10.3791/3803 (2012).

Abstract

Aspergillus fumigatus siderophore A (SidA) is an FAD-containing monoxygenase that catalyzes the hydroxylation of ornithine in the biosynthesis of hydroxamate siderophores that are essential for virulence (e.g. ferricrocin or N,N',N''-triacetylfusarinine C). The reaction catalyzed by SidA can be divided into reductive and oxidative half-reactions (Scheme 1). In the reductive half-reaction, the oxidized FAD bound to Af SidA, is reduced by NADPH²,³. In the oxidative half-reaction, the reduced cofactor reacts with molecular oxygen to form a C4a-hydroperoxyflavin intermediate, which transfers an oxygen atom to ornithine. Here, we describe a procedure to measure the rates and detect the different spectral forms of SidA using a stopped-flow instrument installed in an anaerobic glove box. In the stopped-flow instrument, small volumes of reactants are rapidly mixed, and after the flow is stopped by the stop syringe (Figure 1), the spectral changes of the solution placed in the observation cell are recorded over time. In the first part of the experiment, we show how we can use the stopped-flow instrument in single mode, where the anaerobic reduction of the flavin in Af SidA by NADPH is directly measured. We then use double mixing settings where Af SidA is first anaerobically reduced by NADPH for a designated period of time in an aging loop, and then reacted with molecular oxygen in the observation cell (Figure 1). In order to perform this experiment, anaerobic buffers are necessary because when only the reductive half-reaction is monitored, any oxygen in the solutions will react with the reduced flavin cofactor and form a C4a-hydroperoxyflavin intermediate that will ultimately decay back into the oxidized flavin. This would not allow the user to accurately measure rates of reduction since there would be complete turnover of the enzyme. When the oxidative half-reaction is being studied the enzyme must be reduced in the absence of oxygen so that just the steps between reduction and oxidation are observed. One of the buffers used in this experiment is oxygen saturated so that we can study the oxidative half-reaction at higher concentrations of oxygen. These are often the procedures carried out when studying either the reductive or oxidative half-reactions with flavin-containing monoxygenases. The time scale of the pre-steady-state experiments performed with the stopped-flow is milliseconds to seconds, which allow the determination of intrinsic rate constants and the detection and identification of intermediates in the reaction.⁵ The procedures described here can be applied to other flavin-dependent monoxygenases.⁶

Protocol

1. Preparation of Anaerobic Buffer

1. Prepare 1 L of 100 mM potassium phosphate buffer, pH 7.5. Pour 250 mL of the buffer into a 500-mL Büchner flask with a stir bar.
2. Seal the flask with a rubber stopper and place it on a stir plate. Connect the short tube of the flask to a Schlenk line.
3. Degas the buffer under vacuum for 5 hours at room temperature with agitation. During this time period, perform 5 consecutive rounds of vacuum degassing and flushing with argon every hour.
4. Flush the flask with argon for 10 seconds and disconnect it from the vacuum manifold. Place the flask inside the glove box.
5. Leave the flask open in the glove box overnight with vigorous agitation.

2. Removing Oxygen from the Stopped-flow System

1. Prepare 1 L of 0.1 M sodium acetate, pH 5.0. Pour 125 mL of this buffer into a 250-mL Büchner flask and perform steps 1.2-1.4.
2. Weigh out 5 mg glucose oxidase from Aspergillus niger (181300 U/g) and 0.9 g glucose using a 1.5 mL Eppendorf tube and 50 mL Falcon conical tube, respectively. Place both containers into the glove box.
3. Dissolve glucose oxidase and glucose in 50 mL of anaerobic 0.1 M sodium acetate, pH 5.0 (final concentrations of 18.13 U/mL and 100 mM, respectively). Fill the reservoir syringes with this solution (Figure 1).
4. Ensure that the drive valves are in the "load" position and fill the drive syringes (Figure 1). Turn the F valve to the "drive" position. Empty the stop syringe by clicking on Empty in the Pro-Data control software. Push the buffer from drive syringe F through the flow circuit by manually
3. Preparation of Oxygen Saturated Buffer

1. Prepare 100 mM potassium phosphate buffer (pH 7.5) and pour 50 mL into a 50-mL vial with a stir bar. Cap the vial with a Wheaton stopper and a Wheaton aluminum seal.
2. Place the vial on ice. Immerse a long needle connected to a tank containing 100% oxygen into the solution. Insert another short needle through the stopper of the vial as a vent.
3. Bubble the solution with 100% oxygen for 1 hour with agitation.
4. First remove the short needle from the vial and wait 10 seconds before removing the long needle. Place the closed vial in the glove box.

5. Removal of Oxygen from the Enzyme Solution

1. Takes 400 μL of enzyme stock solution from the freezer and thaw in the glove box. The enzyme stock solution (360 μM) was previously prepared in 100 mM potassium phosphate buffer (pH 7.5) containing 100 mM NaCl, and frozen in liquid nitrogen.
2. Dissolve the NADPH in 300 μL of anaerobic 100 mM potassium phosphate buffer, pH 7.5. Remove 30 μL of this solution from the glove box to determine the NADPH concentration with a spectrophotometer (ε = 13,700 M⁻¹ cm⁻¹) and prepare 4 mL of a NADPH solution 1.5-fold more concentrated.
3. Place the vial on ice and connect it to a Schlenk line by inserting a needle through the stopper of the vial.
4. Degas the enzyme solution by performing 5 consecutive rounds of vacuum degassing and flushing with argon every 20 minutes for 1 hour.
5. Flush the vial with argon for 10 seconds and disconnect it from the vacuum manifold. Place the closed vial in the glove box.

6. Reductive Half-reaction: Monitoring Flavin Reduction

1. Prepare the stopped-flow instrument and the Pro-Data control software for single mixing mode following the manufacturer's protocol.
2. Turn on a circulating water bath (15 °C). Remove oxygen from the water by bubbling nitrogen through it for 20 minutes. This prevents the oxygen contamination through the flow circuit.
3. Maintain the temperature of the drive syringes and the observation cell at 15 °C by connecting the water bath housing of the stopped-flow to the circulating bath.
4. In the Control Panel window of the Pro-Data control software select Photodiode Array, External trigger (to activate stopped-flow operation), and Logarithmic scale.
5. Launch the Pro-Data viewer software and define the directory where the data files will be stored.
6. Place the vial on ice. Immerse a long needle connected to a tank containing 100% oxygen into the solution. Insert another short needle through the stopper of the vial as a vent. Perform this three times in total. In the Pro-Data control software, set the acquisition time to 60 s.
7. Turn both valves to the "drive" position. Empty the stop syringe by clicking Empty in the Pro-Data control software. Push the buffer from both drive syringes through the flow circuit by manually raising the corresponding drive ram five times in total (click Empty before each movement of the ram).
8. Degas the enzyme solution by performing 5 consecutive rounds of vacuum degassing and flushing with argon every 20 minutes for 1 hour.
9. Mix 100 μL of enzyme stock solution with 1100 μL of anaerobic 100 mM potassium phosphate buffer (final concentration after mixing in the stopped flow of 15 μM).
10. Replace the solution of reservoir syringe C with the enzyme solution. Turn the F valve to the "drive" position. Empty the stop syringe by clicking Empty in the Pro-Data control software. Push the enzyme solution from drive syringe F through the flow circuit by manually raising the corresponding drive ram. Perform this three times in total. In the Pro-Data control software, set the acquisition time to 60 s.
11. Ensure that both drive syringes are filled with their corresponding solutions and the drive rams are in contact with the drive syringe pistons. Turn both valves to the "drive" position and click Acquire in the Pro-Data control software to perform the drive. Repeat this step two times more to obtain the data in triplicate. This yields spectra of the oxidized form of the enzyme.
12. Using the A₄₅₂ nm value obtained from the drives, determine the enzyme concentration before mixing (ε = 13,700 M⁻¹ cm⁻¹) and prepare 4 mL of a NADPH solution 1.5-fold more concentrated.
13. Replace the solution of reservoir syringe C with the NADPH solution and empty and refill the stop syringe three times as described in 6.10.
14. Repeat step 6.11. This yields spectra during the reduction of the enzyme.

7. Oxidative Half-reaction: Monitoring Flavin Oxidation

1. Prepare the stopped-flow instrument and the Pro-Data control software for double mixing mode following the manufacturer's protocol.
2. Repeat steps 6.2-6.5.
3. Replace the solution in the four reservoir syringes with anaerobic 100 mM potassium phosphate buffer (pH 7.5). Turn the F valve to the "drive" position. Empty the stop syringe by clicking Empty in the Pro-Data control software. Push the buffer from the drive syringe F through the flow circuit by manually raising the corresponding drive ram. Perform this step five times in total with each drive syringe.
4. Perform step 7.3 four times in total to completely remove the contents from the flow circuit.
5. Click Baseline in the Pro-Data control software.

6. Mix 200 μL of enzyme stock solution with 1000 μL of anaerobic 100 mM potassium phosphate buffer (final concentration after mixing in the stopped flow of 15 μM).

7. Replace the solution of reservoir syringe A with the enzyme solution. Turn the A valve to the “drive” position. Empty the stop syringe by clicking Empty in the Pro-Data control software. Push the enzyme solution from drive syringe A through the flow circuit by manually raising the corresponding drive ram. Perform this step three times in total. In the Pro-Data control software, set the delay time to 0.5 s and acquisition time to 60 s.

8. Ensure that all drive syringes are filled with their corresponding solutions and the drive rams are in contact with the drive syringe pistons. Turn all valves to the “drive” position and click Acquire in the Pro-Data control software to perform the drive. Repeat this step two more times to obtain the data in triplicate. This yields spectra of the oxidized form of the enzyme.

9. Replace the solution of reservoir syringe B with an NADPH solution 1.5-fold more concentrated than the enzyme solution in syringe A. Turn the B valve to the “drive” position and empty and refill the stop syringe three times as described in 7.7.

10. Replace the solution of reservoir syringe C with oxygenated buffer. Turn the C valve to the “drive” position and empty and refill the stop syringe three times as described in 7.7. In the Pro-Data control software, set the delay and acquisition time at 15 and 900 s, respectively.

11. Repeat step 7.8. This yields spectra during the reoxidation of the fully reduced enzyme.

8. Data Analysis

1. Launch the Pro-Data converter software. Click on the Options icon and select Save in ProDataCSV.

2. Open the directory where the data files were stored. Drag the files to the APL Pro-Data converter window. The data will be saved automatically as CSV format files in the same directory.

3. Open the CSV format files using Microsoft Excel (Microsoft, Redmond, WA, USA). Normalize the baseline of the stopped-flow traces by subtracting the corresponding absorption at 700 nm.

4. Analyze the corrected traces using KaleidaGraph (Synergy Software, Reading, PA) by fitting the plot of the absorbance at the corresponding maximum versus time to the appropriate exponential function. As shown in Figure 2B, the absorbance at 452 nm was plotted as a function of time in order to determine the rate of flavin reduction after mixing the enzyme and NADPH in the stopped-flow instrument. In this case, the best fit of the data was obtained using a double exponential equation and rates of 0.65 and 0.23 s$^{-1}$ were calculated for the first and second phase of the reduction, respectively. To determine the rate of formation of the C4a-hydroperoxyflavin intermediate and the reoxidation after reacting the reduced enzyme with oxygen, we analyzed the stopped-flow traces at 380 and 452 nm, respectively (Figure 3B). Using a single exponential equation, we calculated rates of 1.4 and 0.006 s$^{-1}$ for the first and second step of the oxidative half-reaction, respectively.

9. Representative Results

The results from the experiments described in the previous sections show how the reductive half-reaction of Af SidA can be monitored by measuring the changes in the absorbance at 452 nm, which correspond to the changes in the redox state of the flavin. The rate of this step can be determined by fitting the data to the appropriate equation (Figure 2; Step 8.4). The reduction rate obtained (0.65 s$^{-1}$) is similar to the $k_{cat}$ value calculated with steady-state experiments, suggesting that the reduction is the rate-determining step of the reaction. Taking advantage of the double mixing mode of the stopped-flow, the rate of oxidation and the intermediates in this half-reaction can be determined (Figure 3; Step 8.4).

In the reaction catalyzed by Af SidA, the C4a-hydroperoxyflavin is clearly detected (λ$_{max}$ of 380 nm), and the rate of formation and decay can be calculated. The slow rate of the reoxidation obtained (0.006 s$^{-1}$) indicates that the C4a-hydroperoxyflavin is very stable in absence the ornithine.
Figure 1. The stopped-flow instrument. A) Picture of the components of the Applied Photophysics SX20 stopped-flow spectrophotometer. B) Picture of the sample handling unit. C) Scheme of the flow circuit in double mixing mode.

Figure 2. Anaerobic reduction of SidA with NADPH in the stopped-flow instrument. A) Spectral changes recorded after mixing equal volumes of 30 μM SidA and 45 μM NADPH. The first spectrum (oxidized SidA) and last spectrum (fully reduced SidA) is highlighted in blue and red, respectively. B) Absorbance trace at 452 nm recorded as a function of time.
Figure 3. Oxidation of SidA in the stopped-flow instrument. A) Spectral changes recorded after mixing equal volumes of the fully reduced SidA and oxygenated buffer. The final concentrations were 15 μM SidA, 22.5 μM NADPH, and 0.95 mM oxygen. The spectrum recorded at 0.034, 4.268, and 727.494 s corresponds to the fully reduced enzyme, the C4a-hydroperoxyflavin intermediate ($\lambda_{\text{max}}$ of 380 nm), and the oxidized enzyme ($\lambda_{\text{max}}$ of 450 nm), respectively. B) Absorbance trace at 382 and 452 nm recorded as a function of time.

Discussion

Enzymes that catalyze redox reactions usually contain cofactors such as hemes and flavins that undergo significant absorbance changes during the catalytic cycle. The oxidized form of the flavin presents absorbance maxima at ~360 and 450 nm, and its reduction is typically monitored by following the absorbance decrease at 450 nm$^7$. In general, some transient intermediates are present but form and decay too fast to be measured in regular spectrophotometers. Using the Applied Photophysics SX20 stopped-flow spectrophotometer (or similar instruments), it is possible to measure absorbance changes in the millisecond time scale (dead-time, 2 ms). Here we studied the reductive and oxidative half-reactions of the flavin-dependent monoxygenase Af SidA, serving as a model. The rate of hydride transfer was determined by measuring the change of absorbance at 452 nm after mixing the enzyme with NADPH under anaerobic conditions. Subsequently, taking advantage of the double mixing mode of the stopped-flow instrument, the enzyme was first reacted with NADPH, until full reduction was achieved, then the reduced enzyme-NADP$^+$ complex was mixed with oxygen. Following this procedure, it is possible to detect transient oxygenated flavin intermediates and to measure rates of formation and decay. The identification of these intermediates provides experimental data about the nature of the reacting species in catalysis. In the case of Af SidA, the formation of the C4a-hydroperoxyflavin (typically monitored at 370-380 nm), which is the hydroxylating species. In addition, measuring the rate constant of each step allows one to obtain information about the rate-determining step of the reaction and help to elucidate the kinetic and chemical mechanisms of the enzyme.

In general, similar approaches can be used for other flavoenzymes, or proteins for which absorbance changes occurred, such as proteins that contain heme, pyridoxal-phosphate, or non-heme iron$^8$$^{10}$. A limitation to this method is that large amounts of purified enzyme are required, but this can be overcome by using an expression system with high yields. One determines the optimal protein concentration for recording spectra by using enough protein so that a strong enough signal can be observed, but not too much so that enzyme is not wasted. Typically, the lowest enzyme concentration for flavin-containing enzymes used in stopped-flow experiments is 6-10 μM (after mixing) and is determined using the corresponding molar extinction coefficient of the enzyme. In the case of Af SidA, the percentage of the enzyme bound FAD is 50-65%$^2$. Apo-protein is regarded as inactive in these experiments because a bound FAD cofactor is necessary for catalysis. Another possible limitation to this method is if processes in an enzyme occur faster than 2 ms (dead time of the stopped-flow) they will not be observed, but there are reported strategies where rates can be decreased to overcome this issue. One example for this includes using a high NaCl concentration in the reaction of a ferredoxin-NADP$^+$ reductase$^1$. The scrubbing of oxygen from the flow circuit of the stopped-flow is often a tricky step in this experiment and requires special attention. The glucose oxidase-glucose system described here is used successfully in most laboratories as it is an effective and inexpensive method. However, there are some drawbacks which include the production of H$_2$O$_2$ and for some applications other alternatives as the protocatechuate dioxygenase-protocatechuate system should be considered$^{12}$. The utilization of an anaerobic glove box makes it easier to ensure anaerobic conditions, but is not essential. Oxygen must be removed from the flow circuit of the stopped-flow as we want the enzyme to be reduced in the absence of oxygen or react with oxygen at concentrations that we specify. Although the stopped-flow is in the glove box, there is oxygen in the flow circuit if we used aerobic buffers in previous experiments. In addition to absorbance measurements, fluorescence and circular dichroism assays can be performed in the Applied Photophysics SX20 stopped-flow spectrophotometer with the corresponding accessories.
Disclosures

We have nothing to disclose.

Acknowledgements

Research supported by NSF award MCB-1021384.

References

1. Hissen, A.H., et al. The Aspergillus fumigatus siderophore biosynthetic gene sidA, encoding L-ornithine N5-oxygenase, is required for virulence. *Infect. Immun.* **73**(9), 5493-5503 (2005).

2. Chocklett, S.W. & Sobrado, P. Aspergillus fumigatus SidA is a highly specific ornithine hydroxylase with bound flavin cofactor. *Biochemistry.* **49**(31), 6777-6783 (2010).

3. Mayfield, J.A., et al. Comprehensive spectroscopic, steady state, and transient kinetic studies of a representative siderophore-associated flavin monooxygenase. *J. Biol. Chem.* **285**(40), 30375-30388 (2010).

4. Fierke, C.A. & Hammes, G.G. Transient kinetic approaches to enzyme mechanisms. In: Contemporary enzyme kinetics and mechanism. Elsevier, USA, (2009).

5. Palffy, B.A. & McDonald, C.A. Control of catalysis in flavin-dependent monooxygenases. *Arch. Biochem. Biophys.* **493**(1), 26 (2010).

6. van Berkel, W.J.H., Kamerbeek, N.M., & Fraaije, M.W. Flavoprotein monooxygenases, a diverse class of oxidative biocatalysts. *J. Biotechnol.* **124**(4), 670-689 (2006).

7. Chapman, S.K. & Reid, G.A. In: Flavoprotein Protocols. Humana Press, Totowa, USA, (1999).

8. Sobrado, P. & Fitzpatrick, P.F. Solvent and primary deuterium isotope effects show that lactate CH and OH bond cleavage are concerted in Y254F flavocytochrome b2, consistent with a hydride transfer mechanism. *Biochemistry.* **42**, 15208-15214 (2003).

9. Kumar, S., Gawandi, V.B., Capito, N., & Phillips, R.S. Substituent effects on the reaction of &beta;-benzoylalanines with *Pseudomonas fluorescens* kynureninase. *Biochemistry.* **49**(36), 7909-7913 (2010).

10. Yun, D., Garcia-Serres, R., Chicales, B.M., An, Y.H., Huynh, B.H., & Bollinger, J.M.J. (μ-1,2-Peroxo)diiron(III/III) complex as a precursor to the diiron(III/IV) intermediate X in the assembly of the iron-radical cofactor of ribonucleotide reductase from mouse. *Biochemistry.* **46**(7), 1925-32 (2007).

11. Pennati, A., Zanetti, G., Aliverti, A., & Gadda, G. Effect of salt and pH on the reductive half-reaction of *Mycobacterium tuberculosis* FprA with NADPH Biochemistry. *Biochemistry.* **47**(11), 3418-3425 (2008).

12. Patil, P.V. & Allou, D.P. The use of protocatechuate dioxygenase for maintaining anaerobic conditions in biochemical experiments. *Analytical Biochemistry.* **286**(2), 187-192 (2000).