Title: WATER/MAGMA INTERACTION: SOME THEORY AND EXPERIMENTS ON PEPERITE FORMATION

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Water/magma interaction: some theory and experiments on peperite formation

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Abstract

Experiments, using molten thermite as a magma analog, produce peperite when the melt interacts with wet sand. These experiments also show explosive behavior, developing Strombolian- and Surtseyan-like bursts. The results demonstrate that the application of fuel–coolant interaction (FCI) theory is appropriate for interpretation of peperites. The theory described includes discussion of the importance of mass interaction ratios of wet sediment and magma ($R_s$), which determine thermal equilibrium temperature limits and contact interface dynamics. The dynamics of the interface between magma and wet sediments involves heat transfer over a wide range of rates from passive quenching to explosive fragmentation. A vapor film layer develops at the interface and acts both as an insulating barrier, promoting passive quenching, as well as a potential energy reservoir that can cause magma fragmentation, mingling of the magma with wet sediments, and explosive quenching when the vapor film becomes unstable. An important parameter in determining the behavior of the vapor film is the value of $R_s$, which controls whether heat can be convectively removed from the layer as more is being added from its contact with magma. If $R_s > 1$ for fully saturated sediments, there is enough water in the sediments to make convective heat flow effective in quenching the magma, but below that value, there is the potential that the vapor film will be unstable, producing highly dynamic phenomena, including explosive fragmentation. At values of $R_s < 0.1$ there is insufficient water to allow the escalation of explosive fragmentation.

Keywords: peperite; hydrovolcanism; fragmentation; explosion phenomena; experiments; fuel–coolant interaction

1. Introduction

Experimental volcanology at Los Alamos National Laboratory first began in 1975 under the direction of Thomas McGetchin. I followed a twin approach of exploring both novel computer simulations and scaled model volcano experiments. The scaled models provided an experimental basis for analyzing and adapting the numerical simulations. These experiments simulated both effusive and explosive eruption physics. The early models (McGetchin et al., 1976; Widdicombe et al., 1976) employed Carbowax® as a lava analog for effusive eruptions. This polymer forms a two-phase (liquid and solid) system at temperatures near 40°C and can develop flows comparable to those of lava while satisfying most criteria of similitude. In later experiments Wohletz and McQueen (1984) and Wohletz et al. (1995) uti-
lized molten thermite as a magma analog for produc-
ing gas-rich explosions by interaction with ex-
ternal water. We chose thermite because large
batches of melt can be generated easily by its
exothermic reaction. Its physical and chemical
properties are also well suited for a basalt analog.
The thermite experiments continued periodically
over 10 yr and produced dramatic results, which
paved the way for a much deeper appreciation of
the importance and evolution of hydrovolcanism
— volcanism resulting from the interaction of mag-
ma or lava with external water. Some variations
of these experiments demonstrated the potential
for dynamic interaction of magma with wet sedi-
ments and produced the experimental equivalent
of peperite. It is these experiments that are the
subject of this paper.

While the experiments described below give
some basis for understanding of peperite forma-
tion (Kokelaar, 1982; Busby-Spera and White,
1987; Brooks, 1995; White, 1996), a theoretical
understanding is quite important to analyzing
field observations. The theory is important be-
cause the quenching of magma during contact
with water or wet sediments is really quite com-
plex, and interpretations should fall within the
realm of theoretical possibility. For this reason I
address some phenomenological considerations
before describing the experiments.

2. Phenomenology of wet sediment/magma
interaction

Because the thermodynamic behavior of water
is so well known from the steam-locomotion era,
many volcanological studies have addressed the
problem from a rather simplistic approach by di-
rect applications of the First Law of Thermody-
namics. While this approach does provide some
limiting conditions, it completely ignores the com-
plex issues concerning multi-phase fluid mechan-
ics. Fluid mechanics place further constraints on
the application of thermodynamics to this prob-
lem, as described by Delaney (1982) and Wohletz
(1986). Such physical considerations are embodied
in the concept of fuel—coolant interaction (FCI;
Buchanan and Dullforce, 1973; Buchanan, 1974),
a field of study concerning the interaction of two
fluids for which the temperature of one (fuel) is
above the vaporization temperature of the other
(coolant), more recently termed molten fuel—cool-
ant interaction or MFCI by Zimanowski et al.
(1997). For more reading on these considerations,
refer to Wohletz et al. (1995), Zimanowski et al.
(1991), and Zimanowski (1998).

There are important thermodynamic con-
straints for the phenomenology of wet sediment/
magma interaction. These include the temperature
and heat content of the wet sediments and magma
and the rate of heat transfer between them. The
amount of heat energy that is gained by the wet
sediments is limited by the temperature of thermal
equilibrium between the two, that is, the wet sedi-
ments can never gain more heat energy than the
magma loses during their interaction. This ther-
mal equilibrium is expressed as a function of the
water/magma mass ratio, \( R = m_w / m_m \). For satu-
rated sediments or impure coolants (White, 1996),
sediment particles and impurities act as heat sinks,
and the mass ratio, denoted by \( R_s \), is given as:

\[
R_s = \frac{m_s}{m_m} = \frac{x_w}{R_s} \tag{1}
\]

where \( m_s \) is the sediment mass including pore
water and \( m_m \) is the magma mass. In these situa-
tions the water/magma mass ratio is:

\[
R = x_w R_s \tag{2}
\]

where \( x_w \) is the sediment water mass fraction.
Using cgs units for density, the following equa-
tions can be derived, starting with an expression
for \( x_w \):

\[
x_w = \frac{S \phi}{\rho_s'} \tag{3}
\]

In Eq. 3 \( S \) is the sediment saturation (the volume
fraction of pores filled by water), \( \phi \) is the sediment
porosity, and \( \rho_s' \) is the sediment bulk density.
From Eqs. 2 and 3 \( R \) may be approximated as:

\[
R \approx \left( \frac{S \phi}{\rho_m'} \right) \frac{V_s'}{V_m'} \tag{4}
\]
where $V_s$ is the sediment bulk volume (including pore space) and $\rho_m$ accounts for magma composition by its density (cgs units). For peperites the mass of sediments involved in their formation can be calculated. Using values for $S$ and $\phi$ appropriate for the sediments involved by defining the term $S\psi = S\phi(1-\phi)$, one may determine a value for $R_s$ by estimating the volumes of sediment and magma fragments in the peperite, $V_s$ and $V_m$ respectively, and measuring their respective particle densities, $\rho_s$ and $\rho_m$ (cgs units):

$$R_s = \left(\frac{\rho_s + S\psi}{\rho_m}\right)\frac{V_s}{V_m}. \quad (5)$$

For a simple estimate of thermal equilibrium for which there is no phase change, one may apply energy conservation for heat exchange:

$$m_sC_s(T_e - T_s) + m_mC_m(T_e - T_m) = 0, \quad (6)$$

where $C$ is the constant volume heat capacity, $T$ is temperature, and the subscripts $s$, $m$, and $e$ indicate equilibrium, sediment, and magma, respectively. The heat capacity of the wet sediment, $C_s$, can be approximated (e.g. Buntebarth and Schopper, 1998) as:

$$C_s = x_w C_w + (1-x_w)C'_s, \quad (7)$$

where $C_w$ and $C'_s$ denote the constant volume heat capacities of water and solid constituents, respectively. By using the mass ratio, $R_s$, given in Eq. 1 or 5, the thermal equilibrium limit becomes:

$$T_e = \frac{R_s C_s T_s + C_m T_m}{R_s C_s + C_m}. \quad (8)$$

For a typical magma contacting 298 K water-saturated sediments, $C_m$ and $C'_s$ are $\sim 25\%$ of $C_w$ and $T_e$ decreases with $R_s$, as shown in Fig. 1. It is evident that for magma interacting with wet sediments, $T_e$ can exceed critical temperature [647 K (pure) to 720 K (5 wt% dissolved solids)] where $R_s < 1.0$ (basalt) and $R_s < 0.5$ (rhyolite). Where critical temperature is exceeded during interaction prior to explosive expansion, supercritical pressures will be created.

Thermal equilibrium is probably never reached during the time span of interaction because of the insulating property of a vapor film that rapidly forms at the magma–sediment interface. Because of its relatively low thermal conductivity, a vapor film can greatly decrease the rate of heat transfer from the magma to the wet sediments, allowing gradual quenching. With gradual quenching, the film slowly heats water in the sediments near the magma at a rate balanced by the heat transfer away from it by the convective movement of

Fig. 1. Thermal equilibrium ($T_e$) is the maximum possible temperature water can reach during interaction with magma and is shown as a function of $R_s$ (sediment/magma mass ratio). This example is for cases where the initial temperature of water and sediments is 298 K and the magma has typical values of heat capacity about 25% that of water. The critical temperature is shown as a range reflecting the effect of dissolved solids. (a) $T_e$ is shown for several different initial magma temperatures interacting with water (100% volume fraction in sediments). (b) For basalt (solid curve; 1473 K) and rhyolite (dashed curve; 1173 K) interacting with wet sediments, $T_e$ is shown as a function of sediment water volume fraction for which 10 and 70% bound a range centered on 40%, representing a porous, fully saturated sandstone, $x_w \approx 0.2$. 

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pore water. However, such passive quenching is not always the case. Consider the hypothetical instantaneous interface temperature, $T_i$, attained by the initial contact of water with magma; it can be estimated by:

$$T_i = \frac{T_m(\alpha_m/\sqrt{K_m}) + T_w(\alpha_w/\sqrt{K_w})}{(\alpha_m/\sqrt{K_m}) + (\alpha_w/\sqrt{K_w})}.$$  \hspace{1cm} (9)

where $\alpha$ and $\kappa$ are thermal conductivity and diffusivity, respectively, and subscripts $m$ and $w$ refer to magma and water, respectively. For the contact of a typical basalt magma with pure water, $T_i$ approaches 1000 K. Because of the rapidity of heat exchange, water may exist in the metastable state of superheat in which it is a liquid well above its vaporization temperature. A consequence of this superheated state is that it continues to absorb heat at a high rate, reaching temperatures well in excess of its spontaneous nucleation temperature, experimentally measured at about 570 K (Reid, 1976). With temperatures approaching the critical temperature (647 K) the likelihood of vaporization by homogeneous nucleation increases to a point where a vapor film is instantly created. A vapor film formed this way rapidly expands and is highly unstable – that is, it can expand well beyond the thickness required for thermodynamic equilibrium. In doing so the vapor becomes supercooled, leading to its spontaneous condensation. The condensation then leads to a rapid collapse such that liquid water impacts the magma surface with a finite amount of kinetic energy, leading to a second spontaneous vaporization event. This cyclic vapor film growth and collapse is repeated continuously, typically with a frequency of up to 1 kHz (analogous to the Leidenfrost phenomena of a drop of water vibrating on a hot metal surface). Vapor film instability can generate enough kinetic energy to distort the interface between the magma and wet sediments as well cause failure of the host sediments. In some cases, film collapse can lead to jets of water-saturated sediment that actually penetrate the magma surface (White, 1996). In addition, the rapid heat loss from the magma by this continued vapor film instability leads to magma quenching and possible granulation. With these interface phenomena, the magma gradually gets increasingly fragmented, leading to higher surface areas for heat transfer and larger volumes of superheated water and vapor.

The above system of hot magma, wet sediments, and an interface consisting of fragmented magma, superheated water, and steam can evolve in several fashions, depending upon the balance of rates of heat transfer from the magma and convective cooling by the wet sediments. If convective cooling by the sediments matches the heat transfer from the magma, the system can stabilize, forming a mixture of quenched magma and sediments. Alternatively, if the wet sediments cannot convect heat away from the growing interface fast enough, the heat transfer can grow exponentially and reach explosive rates for which the vapor film volume and pressure become too large for the enclosing sediments to contain. This situation leads to a thermohydraulic explosion (Büttner and Zimanowski, 1998) whose thermodynamic work is expended mainly in further magma fragmentation, seismic wave propagation, and vapor expansion, which accelerates sediment, and magma particles in directions of least resistance. For subsurface situations, fragment acceleration may occur along bedding planes, fractures, or even into the magma body itself, forming breccia dikes in the sediments as well as in the magma by a mechanism analogous to hydrofracture (Heiken et al., 1988). If the dikes penetrate the surface, a hydrovolcanic (phreatomagmatic) eruption will occur.

How the interaction of magma with water or water-saturated sediments leads to explosive phenomena is a problem that has been pondered by volcanologists for years. Wohletz et al. (1995) considered both theory and experimental evidence (Wohletz and McQueen, 1984; Wohletz, 1986) to conclude that in a general sense $R$ is perhaps the most important controlling factor. For low values of wet sediment/magma interaction ($R_s < 0.1$) there is not enough water to produce the vapor expansion work required for an explosion. For high values ($R_s > 1.0$) there is enough water to form a convective system that balances heat transfer, so that it never grows to explosive rates. But in-between these two values, thermohydraulic ex-
A plosion is possible. In Fig. 2 conversion ratios (the fraction of magma heat converted to thermodynamic work), calculated by the method described by Wohletz (1986), are shown as a measure of how much mechanical energy is released by the interaction of a basaltic magma with water-saturated sediments (40% porosity, 100% saturated, 40% volume saturation, $x_w \approx 0.2$) compared to those for pure water/magma interaction (100% saturation by volume). Both isentropic (steam separates from fragmented magma) and pseudo-isothermal (steam and magma fragments remain at the same temperature, denoted in figures as isothermal) expansions are calculated. Note that compared to pure water/magma interactions, sediment/magma interactions have lower conversion ratios (less energetic) with optimum conversions near $R_s$ values of 1.0. Fig. 2 also shows the fraction of vaporized water that condenses back to liquid during expansion. For interactions at low $R_s$ values, most of the water remains in the vapor state after expansion, leading to the likelihood of explosive behavior. In contrast for $R_s$ values...
> 3.0 (wet sediments), a dominant portion of the steam formed condenses during expansion leaving the system as a fluid and particle system. This latter behavior allows convective heat transfer that promotes passive cooling, not likely to be explosive. Wet sediment/magma interaction may also occur at depths below the Earth’s surface where hydrothermal systems of elevated temperature and pressure occur. Fig. 3 shows that for situations where sediment pore water is at elevated temperatures (e.g. 358 K) a slightly greater fraction of the magma’s thermal energy is converted to thermodynamic work and optimum peaks occur at slightly higher values of $R_s$ when compared to values shown in Fig. 2; however, elevated hydrostatic pressure (e.g. 30 MPa) does not have much effect. Still, conversion ratios at high hydrostatic pressure in excess of critical pressure are great enough that such interactions could be explosive.

In order to better illustrate the above calculations, Fig. 4 depicts water phase diagrams illustrating pressure–temperature–volume–entropy ($P$–$T$–$V$–$S$) relationships for theoretical initial equilibrium and final states. Both isentropic (Fig. 4a) and pseudo-isothermal (Fig. 4b and c) paths are shown. By comparison to phase diagrams presented by Kieffer and Delany (1979), one can see that isentropic expansion follows isentropes (Fig. 4a). In contrast, pseudo-isothermal expansion takes the vapor and magma fragment mixture along a path intermediate to that for isentropic and pure isothermal expansions. Peperite formation likely involves isentropic expansions for pore water heated at a distance from the magma, but pseudo-isothermal expansion may take place close to the magma contact where hot fragments are mingled with wet sediment.

The phase diagrams of Fig. 4 also illustrate an important fact that experimental studies confirm: the critical point of water is not necessarily a limiting factor in vapor explosions. It has been com-

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**Fig. 4.** Phase diagrams illustrating calculations of wet sediment/magma interaction, using the method of Wohletz (1986). Labeled points (diamonds) are the theoretical initial equilibrium conditions for interactions of various $R_s$ (from 0.01 to 50.0; $x_w = 0.2$), calculated by the method of Wohletz (1986). The critical point (CP), liquid (L), vapor (V), and liquid plus vapor two-phase (L+V) regions are shown. (a) A $P$–$V$ diagram shows expansion volumes and release isentropes (kJ kg$^{-1}$ K$^{-1}$) followed during isentropic expansion of vapor. For all interactions, water expands into the two-phase region. (b) A $T$–$S$ diagram for pseudo-isothermal expansion shows the increase in entropy as water stays in thermal equilibrium with magma fragments. For $R_s < 1.3$ water expands into the vapor field. (c) A $P$–$T$ diagram shows the variation in water temperature during pseudo-isothermal expansion. For more thermodynamic details for these diagrams see Kieffer and Delany (1979).
monly assumed by volcanologists that interactions occurring at confining pressures above the critical point of water (22 MPa) cannot result in explosions, because water exists as a supercritical fluid for which there is no liquid–vapor phase boundary. Considering the thermodynamic paths illustrated in Fig. 4 where interactions are nearly an order of magnitude in excess of critical pressure, release of the interaction pressure involves a specific volume increase, especially for pseudo-isothermal expansion. This volume change occurs at a rate determined by local sound speeds discussed by Kieffer and Delany (1979). The presence of other pressure perturbations, such as those caused by seismicity, host media failure, and the vapor film dynamics, discussed above, adds to the likelihood that expansion will lead to thermohydraulic explosion. The factors governing thermohydraulic explosions are complex (Zimanowski et al., 1997), and I will not pursue them.

Fig. 5. Photographs depict the setup (a) and progress (b) of an experiment designed to visually assess the nature of the contact between molten thermite and water. The setup consisted of a 15-cm-diameter by 60-cm-long Plexiglas cylinder. Thermite is supported by an aluminum disk about half-way up the cylinder, and the bottom of the cylinder is filled with water. High current wires extend into the thermite to initiate its melting. When the thermite melted from the top all the way down and through the aluminum partition, interaction began. The right photo shows a bright, incandescent lobe of molten thermite penetrating down into the water. Higher in the cylinder, dark areas show clots of quenched (dark) thermite. The interaction lasted nearly 1 min during which the molten thermite was fragmented and ejected in a cloud of steam. Ejecta followed ballistic trajectories in a Strombolian-type eruption.
here, but I do wish to stress the point that mag-
ma–wet sediment interactions beneath the Earth’s
surface or in submarine environments where hy-
drostatic pressure is greater than critical pressure
may in certain circumstances be explosive.

Overall, when one considers geological environ-
ments where there is clear evidence of the inter-
action of magma with water and/or wet sedi-
ments, one may perceive shortcomings in the
above physical analysis. For example, White
(1996) concludes that the vent environment influ-
ences the fragmentation processes, explosive ef-
ciency, and resultant particle populations more
than water/magma mass ratio. He finds that cool-
ant impurities, including sediment particles, en-
hance the ability of magma to mingle with cool-
ant but dampen the explosivity, the latter of
which is supported by results shown in Figs. 2
and 3. Furthermore, it is diﬃcult to calculate
from geological information appropriate values
for \( R_s \) and the actual magma fragmentation
mechanism. While there are a number of ﬂuid
dynamic mechanisms proposed to explain how
interaction promotes magma fragmentation (e.g.
Wohletz, 1983), rapid magma quenching evident
in peperite exposures indicates that magma sur-
fACES quench prior to much interaction, thus beg-
ging a question about ﬂuid dynamics. Zimanow-
ski et al. (1997) have recently shown the ﬁrst
experimental evidence that dynamic interaction
can involve a brittle mechanism. To be sure, the
physical analysis presented here is overly simple
to explain much about geological occurrences
where there can be a multitude of external in-
fluences on how interaction proceeds. The under-
standing of how external controls affect the basic
physical behavior described above will require
careful ﬁeld observation and laboratory sample
examination to determine the controls of pressure,
temperature, heat exchange, and fragmentation.

3. Experimental approach

Thermite produces a melt, simulating basaltic
magma, by an exothermic reaction of the ﬁne-
grained aluminum (\( \sim 24 \) wt\%) with magnetite
(\( \sim 76 \) wt\%) as described by Wohletz et al. (1995):

\[
3\text{Fe}_3\text{O}_4 + 8\text{Al} = 9\text{Fe} + 4\text{Al}_2\text{O}_3 + \text{heat.}
\]

In general this thermite composition can be
shown by consideration of heats of formation to
yield about 1130 kJ excess heat per mole of iron
oxide at 1800 K. This excess heat allows addition
of quartzo-feldspathic sand to the thermite at a
ratio of \( \sim 1:3 \) to produce a silicate melt (Table 1).
The physical and thermal properties of Fe–Al
thermite melt are compared with those of a typi-
cal tholeiitic basaltic melt in Table 2. Using the
methods of Bottinga and Weil (1972) and Shaw
(1972), which estimate viscosity of polymeric
melts containing SiO \(_2\) and Al \(_2\)O \(_3\), viscosities in
the range of about 1 to 200 Pa s are predicted
for the thermite melt over a temperature range
from 1500 to 1800 K; the wide range reﬂects sus-
pension eﬀects of crystallites. Indeed, petro-
graphic inspection of quenched samples of this

| Oxide     | Blocky particle | Silicate coating | Spindle particle |
|-----------|-----------------|------------------|------------------|
| SiO\(_2\)  | 14.3            | 36.6             | 18.6             |
| TiO\(_2\)  | 2.3             | 1.4              | 1.9              |
| Al\(_2\)O\(_3\) | 11.4           | 34.4             | 42.5             |
| FeO       | 57.0            | 17.3             | 23.3             |
| MnO       | 1.6             | 1.0              | 1.3              |
| MgO       | 6.4             | 3.7              | 5.8              |
| CaO       | 2.1             | 1.7              | 1.9              |
| Na\(_2\)O  | 3.1             | 2.0              | 3.2              |
| K\(_2\)O   | 2.0             | 2.0              | 1.9              |

\( ^{a} \) Standardless energy dispersive spectral analyses (EDS)
with total Fe expressed as FeO.

| Properties                  | Fe–Al melt | Basaltic melt |
|-----------------------------|------------|---------------|
| Liquidus temperature        | 1000–2000 K | 1370–1520 K    |
| Enthalpy                    | 3700 kJ kg\(^{-1}\) | 1150 kJ kg\(^{-1}\) |
| Viscosity \( \sim 10^2 \) Pa s | \( 10^2–10^4 \) Pa s |
| Density                     | 3.0–4.0 Mg m\(^{-3}\) | 2.5–2.7 Mg m\(^{-3}\) |
| Surface tension             | 0.5 N m\(^{-1}\) | 0.35 N m\(^{-1}\) |
| Thermal conductivity        | 2.4 J m\(^{-1}\) s\(^{-1}\) | 2.1 J m\(^{-1}\) s\(^{-1}\) |

\( ^{a} \) Properties for thermite from Buxton and Benedict
(1979) and for basalt from McBinney and Murase (1984).
melt reveals abundant crystallites, which indicates subliquidus melt temperatures. Accordingly, 100 Pa s is considered typical of the viscosity of the thermitic melt used in the experiments. The reader is referred to Buxton and Benedict (1979) for further discussion on the thermochemical properties of the Fe–Al thermitic reaction.

By designing experiments to make molten thermitite contact water (Wohletz and McQueen, 1984) we found that the two commonly interacted in a dynamic manner with explosive production of steam and fragmented melt. Not all experimental interactions produced explosions. In some experiments, the melt contacted the water, built up pressure in the confinement vessel, and then passively quenched. However, utilizing a number of different experimental designs, Wohletz et al. (1995) found that in general the explosiveness depends largely on the mass ratio of water and melt interacting. It was in experiments where this ratio was high ($R > 5$) that passive behavior was most likely; such experiments produced the equivalent of pillow breccia.

The experiments clearly showed that during explosive interaction the melt and water intimately mingled, breaking the melt into tiny particles that produced an extremely high surface area for rapid heat transfer. This process was observed most readily in a series of experiments for which the interaction was confined in a transparent Plexiglas cylinder that allowed video documentation. One might wonder how Plexiglas with a relatively low melting temperature could confine molten thermitite. The answer is that the heat from the thermitite vaporizes the Plexiglas surface layer, producing a thin insulating film between the thermitite and the Plexiglas. Fig. 5 shows this experimental setup and some results that demonstrate that the melt can move some distance into water before quenching and explosive interaction.

Inspection of video documentation from the experiments revealed that a vapor film served to insulate the water from the melt. But pressure records show that this film is not stable – it grows and collapses repeatedly, each cycle of growth lasting a period of milliseconds or less. With each film collapse, the melt interface is distorted and gradually fragmented, such that heat transfer to water increases with time until it is so large that a vaporization wave propagates and causes nearly instantaneous flashing of relatively large volumes of water with steam pressures large enough to breach confinement. In some cases though, the vapor film never fragments enough of the melt surface to reach heat transfer rates high enough to produce an explosion.

Realizing that in nature magma does not always come in contact with a body of water, we wondered how melt would interact with sediments containing water. Would the presence of sand grains prevent explosive interaction? The easiest experimental design called for using wet sand instead of water. Four experimental designs involved burial of a thermitite-containing vessel in wet sand with and without a separate water compartment (Table 3). In general, Table 3 shows

| Experiment type | $R_s$ | Confinement | Description |
|-----------------|------|-------------|-------------|
| 1.              | 0.9-1.9 | 0.5 m water and wet sand | 1–2 mm diameter melt fragments ejected in a fountain 2–3 m high for 4–5 s; Hawaiian to weak Strombolian |
| 2.              | 0.6–1.3 | 0.5 m water and wet sand | 1 cm diameter melt fragments ejected in a fountain 7–10 m high for 2.5 s using a restricted vent nozzle; strong Strombolian |
| 3.              | 1.3–2.5 | 0.2 m water and wet sand | Strombolian burst of cm-sized melt fragments ballistically ejected 5 m high for ~1 s; Surtseyan burst of <1 mm ejecta to 10 m high for <1 s; ending with lava flow |
| 4.              | 0.48–0.55 | 0.2 m wet sand | Strombolian activity for several seconds ballistically ejected cm-sized fragments of melt 2–5 m into the air |

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a The wet sand/thermite mass ratio $R_s = 5–10R$. Experiments 3 and 4 were repeated twice.
Fig. 6. A series of three photographs depicting the eruption produced by a Type 3 experimental design. The volcano edifice was constructed of sand and is about 0.5 m in diameter and 0.2 m high. Note the cm-size incandescent melt fragments following ballistic trajectories, the black finer grained melt quenched and carried up from the vent in convecting plumes while some moves downslope in a density current, and a small lava flow rivulet running downslope on the right.
results indicating that explosiveness increases with wet sand/melt mass ratio, $R_s$.

Fig. 6 shows the progress of a Type 3 experiment in which thermite is buried in a cone-like mound of wet sand. The thermite first interacted with a small container of water, separated from the thermite by an aluminum partition that was subsequently melted, and then with wet sand contained in the mound. This particular experiment resembled a Strombolian eruption that near its end, when the water source was depleted, formed a small lava flow. Other experiments like this one produced brief Surtseyan blasts, followed by Strombolian activity, but overall we could not tell whether the activity was related to the water body or the wet sand.

In Type 4 experiments, approximately equal volumes of wet sand and thermite ($V_s/V_m = 1.0$) were placed in the vessel shown in Fig. 7. Having $\sim 40\%$ porosity ($V_s/V_m = 0.6$) that was $\sim 50$–$100\%$ saturated, the wet sand contained a mass fraction of water, $x_w = 0.1$–$0.2$, and from Eqs. 4 and 5 $R_s = 0.48$–$0.55$ ($R = 0.05$–$0.11$). The thermite was held above the wet sand by an aluminum plate. A smaller diameter steel ring, placed above the thermite, served as a nozzle to provide some pressure confinement. Above the nozzle another open cylinder directed ejecta upward. This whole assembly was situated on a platform of wet sand and stabilized in a hill of dry sand. The results of these experiments are of particular relevance to peperite formation, as discussed below.

![Fig. 7. Schematic illustration of the Type 4 experimental setup for the interaction of molten thermite with wet sediments (sand). This assembly was emplaced in a hill of sand with the top just at the surface of the hill. High current passed through the igniter loop initiated the thermite reaction, completely melting the thermite in less than half a minute. The melting proceeded from the top downward. When the molten thermite melted and penetrated the aluminum plate at its base, it contacted the wet sand and began the interaction.](image-url)
4. Experimental results

The thermite melt with wet sand experiments produced dramatic evidence of the potential for explosive interaction. Fig. 8 depicts Strombolian interaction typical of Type 4 experiments. The activity lasted several seconds and was characterized by several Strombolian bursts, reaching 2–5 m
above the vent nozzle, and fragmentation of the melt into cm- and mm-sized clasts. In one Type 4 experiment, explosive interaction culminated with Surtseyan explosive activity bursting from the bottom of the cylinder, forming a crater in the sand surrounding the steel assembly (Fig. 9a).

Fig. 9a shows the crater and ejecta blanket on
the sand hill surrounding the experimental assembly from the Type 4 experiment that produced culminating Surtseyan activity. The ejecta are mm-sized fragments of quenched thermite in a matrix of sand grains and fine (<0.1 mm) thermite dust particles. The ejecta blanket extends only about 1 m outward from the vent. Fine ray-like striations visible in the ejecta are the result of alignment of coarser grains of agglomerated thermite fragments. The alignment of these striations does not point to the vent nozzle, but towards the crater that formed along side of the cylinder assembly. This observation suggests that at least a portion of the ejecta came from the venting from underneath the experimental assembly; hence, its origin is related to the peperite dike we found there.

Inspection of the Type 4 experimental assembly afterwards (Fig. 9a) revealed that most of the thermite and a similar volume of wet sand had been completely removed. Excavation into wet sand below the experimental assembly revealed that the thermite melt had penetrated completely through the wet sand compartment into the substrate, which was also wet sand. Here we found a slab-like and tuberous dike of quenched melt, extending about 0.5 m downward and outward into the packed wet sand below the base of the experimental assembly (Fig. 9b). Apparently it was during formation of this dike that ejecta vented around the outsides of the cylinders.

Fig. 10 shows two samples of the dike. The surfaces of the dike are uneven, showing small, bulbous protrusions and wrinkles. Such surface textures are similar to the texture of shark skin or elephant hide quench textures on lava surfaces (described by Lavine and Aalto, 2002). Close inspection of the dike reveals that it is in fact composed of a quenched melt matrix enclosing sand grains. The volume fraction of quenched melt averages from 30 to 35% for these samples. The sand and quenched melt completely and intimately mingled (Fig. 11), but this mixture shows variations in the relative proportions of the constituents with clot-like quenched melt-rich and sediment-rich zones. Fig. 11 also shows that the quenched melt matrix itself appears granular, indicating that it was fragmented during intrusion and quenching. No thermal alteration or oxidation stains are visible, which indicates the rapidity of the quenching process.
5. Discussion and conclusions

From the discussion of phenomenology above, I conclude that peperite can be formed in a wide variety of geological environments where magma interacts with wet sediments. From Fig. 2 it is evident that such interaction can convert at least several percent of the magma's heat energy into thermodynamic work, which in turn is available for fragmentation of the host rock and magma, leading to mingling of the two and formation of a variety of breccia textures, including peperite. This interaction occurs over a range of wet sediment/magma mass ratios ($R_s$) at various confining pressures (lithostatic/hydrostatic) and initial water temperatures. Because interactions that involve $R_s$ between 0.1 and about 1.0 have the greatest potential of being explosive (Figs. 2 and 3), characteristic peperite occurrences not easily related to explosive phenomena are likely to be formed in saturated environments where $R_s > 1$. The reader however should note that other external controls, discussed above, may be more important than $R_s$ in controlling peperite formation.

There is some ambiguity concerning the value of $R_s$ because one does not know a priori how much of the wet sediment and magma will interact. This value can only be estimated, as described above, from field occurrences or samples where the volume ratio of sediment fragments to magma fragments in the peperite can be measured. With estimates of sediment porosity and saturation and magma and sediment clast densities, Eq. 5 can then be applied to make an estimate of $R_s$, but this value will likely be only a minimum estimate because not all sediments contributing to the interaction mingle with magma fragments to form peperite. Consider a very wet, saturated sediment with 50% porosity and bulk density about two-thirds that of a basaltic magma. Applying Eqs. 2–4 one can find that its interacting bulk volume ($V_s')$ must exceed that of magma ($V_m$) by nearly five times in order to achieve values of $R_s > 3$, enough to predict non-explosive peperite (Fig. 2). The peperite formed by this interaction would then have a volume ratio of sediment clasts to magma fragments greater than 3. Such situations are indeed possible, and for environments where $R_s$ is much larger than 3 the sediments are likely to be very fluid, like quick sand or liquefied mud.

Fig. 11. Close-up photographs of experimental peperite textures, showing the intimate mixing of sand grains (from 65 to 70% by volume) within the thermite magma analog. Little or no thermal alteration is visible on any of the sand grains, evidence that quenching was rapid.
In these conditions interaction may result in the formation of globular peperite (e.g. Busby-Spera and White, 1987) or intrusive pillow breccia (e.g. Snyder and Fraser, 1963). The abundant water in these systems allows easy penetration of magma into them but they also rapidly quench the surface of the magma so that blobs of magma detach to form globular shapes like pillows (e.g. Walker, 1992) as the intrusion continues.

For the experimental results described above, only a small range of interaction ratios $R_s \approx 0.48 - 0.54$ was explored. During interaction of the thermit melt with wet sand in the Type 4 experimental vessel, this range of $R_s$ resulted in mostly Strombolian-type bursts, which corresponds to an effective $R$ of $0.5 - 0.1$ in accordance to phenomena predicted by Wohletz and McQueen (1984). Near the end of the Type 4 experiment described above, a portion of the thermit melt intruded wet sand below the vessel. As the intrusion progressed the volume ratio of wet sand to melt increased as more melt surface area contacted wet sand. From inspections of the experimental peperite, the volume fraction of sand is about two times that of the thermite ($V_s/V_m = 2$), which from Eq. 5 corresponds to a value for $R_s$ of about 1.8. When the intrusion stopped growing $R_s \approx 1.8$, which is high enough, at least theoretically, to explain why the thermite quenched to form the peperite dike.

These peperite experiments have not systematically explored the effects of varying the degree of saturation, the sediment strength, the rate at which melt is introduced to the sediments, and the confining pressure. In conclusion, however, these peperite experiments do demonstrate that the interaction of wet sediment with melt can produce vapor explosions while peperite is being formed. The experiments also show that interaction results in the intimate mingling of the melt and sediments. These results suggest that in analysis of peperite formation, the application of thermodynamic and fluid mechanic principles of FCI's as described above is appropriate.

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