Effective Platinum-Copper Catalysts for Methanol Oxidation and Oxygen Reduction in PEM FC

Vladislav Menshchikov¹, Anastasia Alekseenko¹, Vladimir Guterman¹*, Andrey Nechitailov², Nadezhda Glebova², Aleksandr Tomasov³, Olga Spiridonova¹, Sergey Belenkov¹, Natalia Zelenina² and Olga Safronenko¹

¹Southern Federal University, Chemistry Faculty, 7 Zorge st., Rostov-on-Don, 344090, Russia; men.vlad@mail.ru (V.M.); an-an-alekseenko@yandex.ru (A.A.); gut57@mail.ru (V.G.); spiridonova_olga_alex@mail.ru (O.S.); serg1986chem@mail.ru (S.B.); ya.safronenko2014@yandex.ru (O.S.)
²Ioffe Institute, Polytechnicheskaya 26, St. Petersburg, 194021, Russia; a.a.shuv@mail.ioffe.ru (A.N.); glebova@mail.ioffe.ru (N.G.); Alex.Tomasov@mail.ioffe.ru (A.T.); natochka56@mail.ru (N.Z.).
³PROMETHEUS R&D Ltd., Rostov-on-Don, Russia, 344091, Zhmailova st., 4g, 36. prometheus.rd.ltd@gmail.com

* Correspondence: gut57@mail.ru, guter@sfedu.ru

Abstract: Behavior of supported alloyed and de-alloyed platinum-copper catalysts, which contained 14% - 27% wt. of Pt, was studied in the reactions of methanol electrooxidation (MOR) and oxygen electroreduction (ORR) in 0.1 M HClO₄ solutions. Alloyed PtCu/C catalysts were prepared by a multistage sequential deposition of copper and platinum onto a Vulcan XC72 dispersed carbon support. De-alloyed PtCu/C catalysts were prepared by PtCu/C materials pretreatment in acid solutions. The effects of the catalysts initial composition and the acid treatment condition on their composition, structure, and catalytic activity in MOR and ORR were studied. Functional characteristics of platinum-copper catalysts were compared with those of commercial Pt/C catalysts when tested, both in an electrochemical cell and in H₂/Air membrane-electrode assembly (MEA). It was shown that the acid pretreatment of platinum-copper catalysts practically does not have negative effect on their catalytic activity, but it reduces the amount of copper passing into the solution during the subsequent electrochemical study. The activity of platinum-copper catalysts in the MOR and the current-voltage characteristics of the H₂/Air PEMFC MEAs measured in the process of their life tests were much higher than those of the Pt/C catalysts.

Key words: platinum electrocatalyst, PtCu/C, oxygen electroreduction, methanol electrooxidation, catalyst activity, durability, fuel cell life tests, de-alloyed catalysts, PEM FC

1. Introduction

Low-temperature proton-exchange membrane fuel cells are the promising sources of electrical energy that can be used in various types of devices, including automobiles, unmanned aerial vehicles, portable chargers, stationary units with a regular power supply, etc [1-4]. The study and application of hydrogen-air (PEM FC) and methanol (DMFC) fuel cells are of particular interest [3-7]. During the operation of a hydrogen-air fuel cell with a proton exchange membrane, hydrogen is oxidized at the anode: H₂ - 2e⁻ = 2H⁺, while two protons and two electrons are formed from one molecule. In a direct methanol fuel cell, an electrooxidation reaction of the simplest monohydric alcohol proceeds at the anode: CH₃OH + H₂O → CO₂ + 6H⁺ + 6e⁻. Electrons pass through the external circuit, while protons migrate through a polymer electrolyte membrane (Nafion) [5-7]. Both electrons and protons reach the cathode, and next they react with the continuously supplied oxygen: O₂+4H⁺ + 4e⁻ = 2 H₂O. High speed current-forming reactions require the presence of highly efficient electrocatalysts in the cathode and anode layers [8-11]. In an acidic environment, the best electrocatalyst for the above reactions is platinum, which is usually used in the form of nanoparticles deposited on the particles of dispersed carbon carriers [9-12].

It is well known that platinum alloying with the base metals such as Ni, Co, Cu, etc. contributes to an increase in the functional characteristics of catalysts in MOR and ORR [6, 10-14]. The increase in the
activity of bimetallic catalysts compared to Pt/C in the methanol electrooxidation reaction is usually explained by the bifunctional catalysis mechanism, in which OH groups adsorbed on the atoms of the doping component facilitate electrochemical desorption of the intermediate products of methanol oxidation from the Pt surface [6, 15-17]. Optimization of catalysts for ORR is associated with an increase in the bond strength of Pt-O, which facilitates the adsorption of oxygen and intermediate oxygen-containing products (OH, HOz, etc.) formed during its electroreduction. At the same time, an excessively strong adsorption interaction can hinder the desorption of reaction products [18, 19]. Platinum doping with copper results in a decrease in the interatomic distance of Pt-Pt in the nanoparticles, as well as the change in the energy of free d-orbitals, which facilitates the adsorption of O₂ on the surface of metal nanoparticles [11, 18–20].

To reduce the content of the precious metal in the bimetallic catalyst, while increasing its activity in current-forming reactions, attempts are made to optimize the composition and structure of platinum-containing nanoparticles [10, 13, 18-24]. Note that the features of the hierarchical structural organization of PtCu/C catalysts are determined not only by the shape and size of the metal nanoparticles, but also by their spatial distribution on the surface of the support, and have a crucial effect on the activity of catalysts in ORR and MOR [20-23]. Evaluation of some PtCu/C catalysts activity and stability in ORR and MOR obtained in electrochemical cells clearly indicates their high functional characteristics [20-22, 24]. Unfortunately, the composition of bimetallic nanoparticles changes in the process of their functioning: there is a selective dissolution of atoms of the alloying component. When conducting studies in an electrochemical cell in the presence of a large amount of liquid electrolyte, the concentration of M⁺ cations in the solution is low and it doesn’t have a noticeable effect on the electrode characteristics. A membrane-electrode assembly with a proton-conducting polymer electrolyte is another matter. Under the conditions of the MEA operation, the selective dissolution of the base metal from the nanoparticles is dangerous due to the possible poisoning of both the proton conducting polymer, which is a part of the catalytic layer, and the polymer membrane [25-27].

In our opinion, catalysts based on de-alloyed PtCu nanoparticles containing a relatively small amount of a "strongly bounded" copper could be very promising for use in membrane-electrode assemblies of the fuel cells. Direct synthesis of such catalysts is rather problematic, but one could try to obtain them from the copper rich PtCuₓ/C (ₓ> 1) materials by treating them in acids. It is important to make it sure that, firstly, the predominant dissolution of Cu atoms allows the formation of a secondary Pt shell that protects the internal copper atoms from dissolution [22, 28]. Secondly, it is necessary to determine, whether such de-alloyed PtCuₓ/C catalysts retain high stability and activity in ORR and MOR, which are characteristic of the previously studied alloyed PtCuₓ/C catalysts [28-32].

This study is aimed at a comparative study of the functional characteristics of de-alloyed PtCuₓ/C catalysts obtained by the acid pretreatment of PtCuₓ/C materials, as well as their comparison with the characteristics of commercial Pt/C catalysts. It has been carried out not only in electrochemical cells with liquid electrolyte, but also in the MEAs of a hydrogen-air fuel cell with a proton exchange membrane.

2. Materials and Methods

Vulcan XC-72 carbon-supported platinum-copper catalysts were prepared by a four-stage synthesis in the liquid phase, sodium borohydride being used as a reducing agent as described in [22]. Initially, copper nanoparticles were deposited on a carbon support from a CuSO₄ solution. At the second and third stages, the calculated amounts of copper (CuSO₄) and platinum (H₂PtCl₆) precursors were added to the suspension, and next they were reduced with the excess sodium borohydride solution. At the fourth stage formation of a platinum layer on the surface of nanoparticles was carried out by reducing Pt (IV) from the solution, which did not contain copper ions. The PtCuₓ/C suspension was filtered, and the catalyst was dried over P₂O₅. Heterogeneity of the components distribution (an increase in the platinum concentration from the center to the surface) in the platinum-copper nanoparticles, formed in accordance with the described synthesis technique, was proved in [22]. The composition of the obtained PtCuₓ/C materials samples, labeled S1 - S5 below, corresponded to 1.7 ≤ x ≤ 2.9 values (see the "Results and discussions" section). A portion of each of the synthesized catalysts was kept in solutions of different acids at 25 °C for 2 – 6 hours with constant stirring. After being obtained in the liquid phase or treated in acid solutions, all materials were filtered, washed with the distilled water, dried at room temperature in a desiccator over P₂O₅. De-alloyed PtCuₓ/C materials, obtained from samples S1 - S5 as a result of acid
treatment, were labeled S1A - S5A, respectively. The more detailed information on the composition of specific materials and conditions of their acid treatment is given in the Results and Discussion section.

The ratio of platinum and copper in the platinum-copper catalysts was determined by the method of X-ray fluorescence analysis (XRF) on a spectrometer with the total external reflection of X-ray radiation RFS-001 (Research Institute of Physics, SFedU). The mass fraction of metals was determined by the mass of the residue obtained after 40 minutes at 800 °C in air. For some samples the changes in the mass and the kinetics of high-temperature oxidation were studied by thermogravimetry. For this, the combined TGA/DSC/DTA NETZSCH STA 449 C analyzer was used. Oxidation was carried out in the atmosphere consisting of N2 (80%) and air (20%), in the temperature range from 20 to 750 °C at the heating rate of 10 °C/min and the gas flow rate of 20 ml/min, using corundum crucibles. When calculating the mass fractions of metals in PtCu/C samples, it was assumed that the non-combustible residue consists of Pt and CuO. The accuracy of mass fractions determination was ± 0.4%.

The catalysts phase composition was determined by X-ray phase analysis (the voltage on the X-ray tube was 50 kV, the current was 150 μA, the time for the spectrum recording was 300 s, molybdenum anode was used) Diffraactograms were taken on an ARL X`TRA powder diffractometer with the Bragg-Brentano geometry (θ-θ), CuKα radiation (λ = 0.15405618 nm). Measurements were carried out at room temperature. Samples were thoroughly mixed and placed in a cuvette1.5 mm. in depth. The shooting was carried out in the angles range of 15-55 degrees with the step of 0.02 degrees and the speed of 8 to 0.5 degrees per minute, depending on the task set. The average crystallite size of the metal phase was determined using the Scherrer equation for a more intense reflection (111), as described in [22, 33]. Electrochemical behavior of the catalysts in a standard three-electrode cell was studied by cyclic voltammetry (CV) at the solution temperature of 23 °C on an AFCBP bipotentiostat (Pain Research Instrumentation). A 0.1 M HClO4 solution (ChP, manufactured by Vector, Russia) was used as an electrolyte. A saturated silver chloride electrode was used as a reference electrode, a platinum wire – as a counter electrode. All the potentials were given relative to the potential of the reversible hydrogen electrode (RHE). The studied electrode was a catalytic layer formed at the end of the glassy carbon disk electrode a drop of suspension being applied first. Before applying the suspension, the electrode surface was polished and then washed in isopropyl alcohol.

To obtain a catalyst suspension (catalytic “ink”), 900 μl of isopropyl alcohol and 100 μl of a 0.5% aqueous emulsion of Nafton® polymer were added to 0.0060 g of each sample. Next, the suspension was dispersed by the ultrasound under cooling for 15 minutes. With continuous stirring, a 6 μl aliquot of “ink” was taken using a microdoser and applied to the end face of a polished and degreased glassy carbon electrode with an area of 0.196 cm², the exact weight of the drop being recorded. After being dried, 7 μl of a 0.05% Nafton® emulsion were applied to fix the catalytic layer, next the electrode was dried in air for another 15 minutes.

The studied electrode was standardized by means of 100 cycles of potential scanning in the range of values from 0.04 V to 1.2 V at the rate of 200 mV/s. Then, 2 cyclic voltammograms (CV) were recorded on the fixed electrode at the potential sweep rate of 20 mV/s. The calculation of the electrochemically active surface area (ECSA) was performed for the second CV. For this purpose, the amount of electricity spent on electrochemical adsorption Qads and desorption Qdes of hydrogen was estimated as described in [22]. In some cases, ECSA was additionally measured by the oxidation of a chemisorbed CO monolayer, as described in [22]. In the latter case, the electrode was kept at a potential of 0.1 V in an electrolyte saturated with CO for 20 minutes. Then, the solution was purged with argon for 20 min, after that, two cyclic voltammograms were recorded, the purge not being stopped, and according to these voltammograms the calculation was performed. ECSA values are given with an accuracy of ± 10%.

While studying the catalysts activity in MOR, methanol and perchloric acid were added to the electrochemical cell, thereby obtaining a solution of 0.1 M HClO4 + 0.5 M CH3OH. Cyclic voltammograms were recorded in the potential range of 0.04 – 1.3 V with a potential sweep rate of 20 mV/s. Chronoamperograms were measured at the potential of 0.70 V. The measurements were carried out in the argon atmosphere.

To evaluate the catalysts activity in ORR, a 0.1 M HClO4 solution was saturated with oxygen for 1 hour, after that a series of voltammograms was measured in the range from 0.12 to 1.19 V with a linear potential sweep at the speed of 20 mV/s at the electrode rotation speeds of 400, 900, 1600, and 2500 rpm. To take into account the contribution of the ohmic potential drop and non-ORR related processes, the
voltammograms, obtained at the potential scanned towards more positive values were normalized according to the generally accepted methods [34–36]. For this, the potential of the electrode under study was refined by the formula: \( E = E_{so} - I_0 + R \), where: \( E_{so} \) is the set value of the potential, \( I_0 \) is the ohmic potential drop equal to the product of the current strength by the resistance \( (R) \) of the solution layer between the reference electrode and the studied electrode, which in this case was 23 ohms. This resistance value is in good agreement with the published data [35]. The contribution of the processes, occurring on the electrode in the oxygen-free solution (Ar atmosphere), was taken into account by subtracting from the voltammogram a similar curve recorded on the same electrode during measurements in the Ar atmosphere: \( I(V) - 1(Ar) \), as described in [36, 37]. The catalytic activity of the catalysts in the ORR (kinetic current) was determined by the normalized voltammograms, taking into account the contribution of mass transfer under the conditions of RDE [36, 37]. The kinetic current was calculated according to the Koutecky-Levich equation: \( 1/j = 1/j_s + 1/j_k \), where \( j \) is the experimentally measured current, \( j_s \) is the diffusion current, and \( j_k \) is the kinetic current. Kinetic currents were calculated for the potential of 0.90 V (RHE).

It is well known that the ECSA values and values of the catalyst activity depend on the composition, structure, and method of the catalytic layer deposition [34, 37–39]. Therefore, the electrodes, on which similar catalytic layers were formed by the same method but with commercial Pt/C HiSPEC3000 and HiSPEC4000 (Johnson Matthey) electrocatalysts containing 20% and 40% wt. platinum, respectively, were used as the reference samples.

The tests of electrocatalysts in the membrane-electrode assembly (MEA) were carried out as follows. To form the cathode and anode layers, the studied catalyst was applied in the necessary amount to achieve platinum loading of 0.4 mg/cm². A commercial Nafion solution (DE-2020, 20% by weight) was used.

Technological operations to prepare the dispersion of the electrode material, required for the subsequent fabrication of the MEA, included two stages: mechanical and ultrasonic dispersion of a mixture of precisely weighed components in an isopropanol-water mixture. The volume ratio of the liquid components of isopropanol to water was ~1:1. The ratio of the masses in the solid and liquid phases in the final dispersion was in the range 1:40 - 1:80.

Mechanical dispersion was performed on a magnetic mixer of the Milaform MM-5M type for ~0.5 hours with an arm rotation speed of ~ 400 rpm until a visually homogeneous (without visible lumps) mass was obtained. Subsequent ultrasonic dispersion was carried out in a Branson 3510 ultrasonic bath for 40-100 hours to obtain a homogeneous dispersion that did not delaminate for one minute.

Membrane-electrode assemblies were made by applying a uniform dispersion of components directly to the proton-conducting membrane through a stainless-steel mask. Before applying the electrode material, the membrane was kept in 0.5 M sulfuric acid for 15 min at temperature 70-80 °C, followed by five times washing with water. The electrodes were made by smearing the dispersion of components in an isopropanol–water mixture at room temperature onto a 50 μm (46–50 μm) thick Nafion proton-conducting membrane pre-treated at 85 °C. The amount of supported catalyst was controlled gravimetrically. Before conducting electrochemical measurements, the MEA was kept in 0.5 M sulfuric acid for 15 min at temperature 70-80°C, followed by five times washing with water. After that, the MEA was placed in a standard measuring cell (FC-05-02, ElectroChem, Inc.) with graphite collector electrodes. Toray060 standard carbon paper was used as the gas diffusion layer.

The change in the MEA characteristics during aging was monitored by the current-voltage characteristics. We used the two-electrode method, potentiostat P-150 (LLC Elins, Russia). Potential sweep rate - 10 mV/s. Before starting the main measurements, the MEA was activated as described in [40]. MEA aging was carried out at room temperature and atmospheric pressure for a given number of cycles (0, 100, 300, 1000, etc.) of a voltage scan in the range 0.6 - 1.0 V with a potential scan rate of 50 mV/s. Wet (~100%) N2 and H2 were applied to the electrodes. An aging electrode was supplied with N2. The MEA aging regime corresponded to that of the DOE protocols (Table A-1) [41].

3. Results and Discussion

3.1. Methanol electro-oxidation at the alloyed and de-alloyed platinum-copper catalysts
The initial alloys and the platinum-copper catalysts obtained from them after two-hour treatment in 1 M HNO₃ at 25 °C (de-alloyed), which contained from 18.5 to 26.6% wt. of platinum, were studied as catalysts for the methanol electro-oxidation (Table 1).

**Table 1.** Composition and structural characteristics of the obtained PtCu/C and commercial Pt/C materials.

| Sample  | Metal component composition of the catalysts (according to XRF) | Platinum loading, ω (Pt), % wt. | 2 θ (111), grad | Average crystallite size, D_{av}, nm |
|---------|---------------------------------------------------------------|---------------------------------|----------------|----------------------------------|
|         | In “as-prepared” state | After electrochemical standardization |                               |                  |                                  |
| S1      | Pt₇₅Cu₂₅          | Pt₆₅Cu₃₅         | 19.9 ± 0.5 | 41.2 | 3.0 ± 0.2 |
| S2      | Pt₇₅Cu₂₅          | Pt₆₅Cu₃₅         | 23.4 ± 0.5 | 41.3 | 3.0 ± 0.2 |
| S3      | Pt₇₅Cu₂₅          | Pt₆₅Cu₃₅         | 26.6 ± 0.5 | 41.2 | 2.8 ± 0.2 |
| S1A     | Pt₆₇Cu₃₃          | Pt₆₅Cu₃₅         | 18.5 ± 0.5 | 40.9 | 2.7 ± 0.2 |
| S2A     | Pt₆₇Cu₃₃          | Pt₆₅Cu₃₅         | 22.4 ± 0.5 | 41.0 | 3.0 ± 0.2 |
| S3A     | Pt₆₇Cu₃₃          | Pt₆₅Cu₃₅         | 26.4 ± 0.5 | 40.8 | 2.7 ± 0.2 |
| JM20 (Pt/C) | Pt | Pt | 20.0 ± 0.5 | 39.9 | 2.0 ± 0.2 |

* D_{av}, nm – average crystallite size, calculated by the Scherrer formula.

The X-ray diffraction patterns of platinum-copper materials contain reflections corresponding to the phases of carbon and platinum (Figure 1). In this case, platinum characteristic reflections are shifted toward large values of the 2 theta angles, this being a consequence of a Pt-Cu solid solution formation. A significant difference in the composition of the initial S1-S3 samples had practically no effect on the position of the characteristic reflections of the X-ray diffraction pattern, in particular, on 2 theta angles of the characteristic maximum 111 (Table 1, Figure 1 b). Apparently, a larger or a smaller fraction of copper in these samples is contained in X-ray amorphous oxides [42], while the composition of the metal component of various catalysts is indeed similar. At the same time, the 2 theta angle of the maximum 111 in the diffraction patterns decreases with the transition from the starting materials (S1, S2, S3) to the de-alloyed catalysts obtained from them (Table 1, Figure 1). This indicates a decrease of the copper content in the catalysts metal component and correlates with the obtained results, when their composition was determined (Table 1).

The average crystallite diameters in the studied bimetallic materials, calculated by the Scherrer formula, are from 2.7 to 3.0 nm (Table 1). Given the accuracy of the measurements, this indicates the absence of a significant effect of the initial composition and acid treatment on their size. At the same time, it should be noted that the application of the Scherrer equation to calculate the average size of bimetallic nanoparticles may provide results which are not quite accurate, since such calculation does not take into account the contribution of the inhomogeneous structure of nanoparticles to the broadening of reflection 111 [42].

Acid treatment of all the initial PtCu/C samples leads to the dissolution of a significant amount of copper (Table 1). A decrease in the concentration of copper in the samples cannot be associated only with the dissolution of its amorphized oxides, since judging by the data of the X-ray diffractometry, the content of metallic copper in the solid solution based on platinum, decreases. This leads to the observed decrease in the 2-theta angle for the maximum (111) during the transition from the initial to the acid-treated samples (Table 1, Figure 1). Note that despite the significant difference in the initial compositions of the platinum-copper catalysts, the compositions of the three de-alloyed samples are similar to each other: Pt₆₇Cu₃₃ - Pt₆₇Cu₃₃. Thus, the amount of “firmly bound” copper, whose selective dissolution from platinum-copper nanoparticles does not occur at the stage of acid treatment, slightly depends on the initial composition of the materials, obtained by the synthesis method described above.
The behavior of the initial and de-alloyed catalysts in the process of electrochemical standardization has significant differences. CVs of S1 – S3 samples during the first few cycles show anodic maxima in the potential range 0.25-0.35 V, associated with the dissolution of copper from the intrinsic phase (Figure 2) [17, 24]. For the samples of de-alloyed catalysts S1A - S3A, similar maxima on CVs were not observed. It should be noted that the materials surface development in the “as prepared” state continues during a greater number of cycles than for catalysts treated in acid (Figure 2). This indirectly indicates a long-term reorganization of the nanoparticles surface, due to a more intensive anodic dissolution of copper atoms. According to the published data [17, 21], the anodic dissolution of copper from the Pt-Cu solid solution mainly occurs in the potential range 0.60–0.85 V and is reflected by a gradually decreasing local maximum on the CVs at the above potentials (Figure 2, samples S1 – S3).

It was found that during the standardization stage the copper content in all studied catalysts (alloyed and de-alloyed) decreased due to its selective dissolution, the composition of the materials, as a result, being approximately the same Pt63Cu37 - Pt67Cu33 (Table 1). Moreover, 70% to 83% and 33 to 45% of the copper content are dissolved from the alloyed and de-alloyed catalysts, respectively. Given the lower content of the alloying component in de-alloyed catalysts (Table 1), we can assume a significant decrease in the concentration of copper (II) cations in the electrolyte and, therefore, a decrease in their possible negative effect on Nafion. Nevertheless, in order to exclude the influence of copper (II) cations, which were present in the solution, on the behavior of the catalysts, prior to further electrochemical studies, the electrolyte was replaced with a freshly prepared 0.1 M HClO4 solution, which was then saturated with argon for 20 minutes.

Figure 3 a, c shows cyclic voltammograms of the standardized catalysts. They have the form characteristic of Pt/C and PtM/C electrocatalysts [17, 21, 24]. The ECSA values of PtCu/C materials, determined by the areas of electrochemical adsorption/desorption of atomic hydrogen and the oxidation peaks of the monolayer of chemisorbed CO, are in the range from 30 to 40 m²/g (Pt) (Table 2), which is much lower than that of the commercial Pt/C sample. The lower ECSA values of the platinum-copper catalysts, compared to the commercial Pt/C sample, are largely due to the larger average crystallite size (Table 1), as well as the agglomeration (aggregation) of the nanoparticles.
Figure 2. Cyclic voltammograms of catalysts in the standardization process (100 cycles). The electrolyte solution is 0.1 M HClO₄ saturated with argon at atmospheric pressure. The sweep rate of potential is 200 mV/s.

The acid treatment of S1 - S3 catalysts does not lead to a considerable change in the platinum ECSA (Table 2). Apparently, the stage of electrochemical standardization, leading to the formation of catalysts with approximately the same composition, eliminates the differences in ECSA values that could occur in the initial and de-alloyed samples.

The oxidation of chemisorbed CO on all platinum-copper catalysts begins at the potentials lower than those on Pt/C, and ends in a narrower range of potentials, approximately 0.63–0.88 V (Figure 3b, d). Bimodal maxima of CO oxidation (Figure 3b, d) can be related to heterogeneity of the composition and structure of bimetallic nanoparticles, as well as their size dispersion [31].

Electrocatalysts activity in the MOR was determined by the methods of cyclic voltammetry and chronopotentiometry. Using the CVs measured in 0.5 M solutions of CH₃OH, we calculated the specific amount of electricity $Q_{CH3OH}$ (Cl/g(Pt)) spent on methanol oxidation in the direct course of potential sweep (Figure 4 a, c, Table 2). Despite the smaller ECSA values for all platinum-copper catalysts, $Q_{CH3OH}$ is greater than that for Pt/C (Table 2).

Table 2. ECSA values and parameters characterizing the catalysts behavior in the methanol electrooxidation reaction.

| Sample | ECSA ($H_{ads/desor}$, (CO)) | ECSA (CO) | $Q_{CH3OH} \times 10^5$ (Cl/g(Pt)) | $I_{max}$ oxidation CH₃OH | Chronoamperometry results at E=0.70 V |
|--------|-----------------------------|-----------|-----------------------------------|---------------------------|--------------------------------------|
| S1     |                             |           |                                   |                           |                                      |
| S1K    |                             |           |                                   |                           |                                      |
| S2     |                             |           |                                   |                           |                                      |
| S2K    |                             |           |                                   |                           |                                      |
| S3     |                             |           |                                   |                           |                                      |
| S3K    |                             |           |                                   |                           |                                      |
|       | $m^2$/g(Pt) | $m^2$/g(Pt) | $A$/g(Pt) | $I_{initial}$/A/g(Pt) | $I_{final}$/A/g(Pt) |
|-------|-------------|-------------|------------|----------------------|-------------------|
| S1    | 39 ± 4      | 31 ± 3      | 127.8      | 970                  | 492.3             |
| S2    | 31 ± 3      | 33 ± 3      | 116.6      | 816                  | 464.4             |
| S3    | 37 ± 4      | 29 ± 3      | 70.1       | 600                  | 432.3             |
| S1A   | 38 ± 4      | 31 ± 3      | 96.0       | 868                  | 529.5             |
| S2A   | 43 ± 4      | 37 ± 4      | 107.8      | 997                  | 578.7             |
| S3A   | 33 ± 3      | 32 ± 3      | 72.3       | 647                  | 411.7             |
| JM20  | 81 ± 8      | 78 ± 8      | 32.9       | 350                  | 291.6             |

*When calculating the specific values of the electric current and the amount of electricity, ECSA values were used, calculated by the adsorption/desorption of atomic hydrogen.

![Cyclic voltammograms of electrocatalysts in the argon atmosphere](image)

**Figure 3.** Cyclic voltammograms of electrocatalysts in the argon atmosphere (a, c) after completion of standardization and (c, d) after CO purging (b, d). Commercial Pt/C (JM20) and initial PtCu/C materials - a, b; commercial Pt/C (JM20) and de-alloyed PtCu/x/C materials.

The same is true for the maximum currents of methanol oxidation on CVs: for bimetallic catalysts they are essentially higher than for Pt/C (Figure 4 and Table 2). The potentials for the onset of methanol oxidation, determined for platinum-copper catalysts by the corresponding CVs of the forward scan, are approximately 150-200 mV less than for commercial Pt/C. Despite a considerably higher methanol oxidation rate and, possibly, a higher concentration of intermediate reaction products on the electrode surface, platinum-copper catalysts, as a whole, showed higher current values when chronocoulometry was performed at the potential of 0.7 V (Table 2, Figure 4 b, d) compared to Pt/C. Together with a lower oxidation potential of a monolayer of chemisorbed CO (Figure 3 b, d), this indicates a higher tolerance of bimetallic catalysts to intermediate products of methanol oxidation. The least activity in MOR was demonstrated by S3 and S3A catalysts.
Figure 4. (a, c) Cyclic voltammograms and (b, d) chronoamperograms of commercial Pt/C (JM20) and the studied (a, b) alloyed and (c, d) de-alloyed PtCu/C catalysts. The currents are normalized to ECSA. The potential sweep rate during CVs registration is 20 mV/s. The electrolyte is a solution of 0.1 M HClO$_4$ + 0.5 M CH$_3$OH saturated with Ar at atmospheric pressure.

It should be noted that de-alloyed PtCu$_{x-y}$/C catalysts are not inferior in activity to MOR electrocatalysts, which were deposited on a glass-graphite electrode in the “as-prepared” state and initially contained a significantly larger amount of copper. Obviously, the reason for their similar behavior is the proximity of the composition / structure of the platinum-copper catalysts, formed after a portion of copper being selectively dissolved due to a chemical and/or electrochemical treatment of the initial materials (Table 1) [42]. A comparative analysis of the electrochemical behavior of the platinum-copper catalysts in MOR (Figure 4, Table 2) suggests a noticeably higher activity of samples A1, A1K, A2 and A2K compared to A3 and A3K: initial A1 and A2, and de-alloyed A1K, A2K catalysts exhibit higher values of Q$_{CH3OH}$ (table 2), electric current of “methanol maximum” on CVs (Figure 4 a, c) and currents on chronoamperograms of methanol oxidation (Figure 4 b, d). This difference in the catalysts behavior can be partially due to the difference in the composition of the metal component after electrochemical standardization: for example, A3K sample contains slightly less copper than A1K and A2K (see Table 1).

In this case, as was noted above, reliable differences in the ECSA values of different platinum-copper catalysts were not established (Table 2). In our opinion, one should take into account possible differences in the character of atoms localization of a “firmly bound” copper in the standardized nanoparticles, their surface being enriched in platinum. The smaller the thickness of the secondary platinum shell, the stronger the promoting effect of the alloying component atoms on the activity of platinum in electrochemical reactions can be expressed.

This study has shown that platinum-copper catalysts exhibit much higher mass activity (A/g(Pt)) and specific activity (A/m$^2$) in MOR compared to the commercial Pt/C material, despite a significantly lower ECSA. The pretreatment of catalysts in the nitric acid promotes leaching of a larger copper fraction from the nanoparticles, but does not lead to a decrease in the activity of catalysts in MOR. Therefore, such a procedure can be used for the pretreatment of catalysts to reduce the subsequent pollution of MEAs with copper cations.
3.2. Electroreduction of oxygen on alloyed and de-alloyed platinum-copper catalysts

The next step in our research was to test the hypothesis that changes in the composition and structure of the platinum-copper nanoparticles, and, as a result, their activity in the electroreduction of oxygen may depend on the processing conditions in acids, in particular, on the nature and concentration of the acid. At this stage, the Pt₃₁Cu₆₉/C (S₄) catalyst with a low platinum content (about 14 wt%) was used as the initial sample. Equal portions of this material were subjected to a 6-hour treatment in solutions of different acids at room temperature ~ 25 °C (Table 3). Then the composition, structural characteristics (Table 3) and the electrochemical behavior of the obtained de-alloyed catalysts were studied.

The initial S₄ sample treatment in acids leads to a selective dissolution of a greater amount of copper and, apparently, to a small amount of platinum. As a result, the composition of the metal component varies from Pt₃₁Cu₆₉ (S₄) to Pt₅₇Cu₄₃-Pt₅₈Cu₄₂ for most of the de-alloyed samples (Table 3). The total mass fraction of metals also changes accordingly: from 23.4% mass (S₄) up to 16.6 - 17.8 % mass for pre-treated samples (Table 3). Note that the sample treated in a 5 M solution of the nitric acid is characterized by the lowest residual copper content (Pt₆₃Cu₃₇) in comparison with other de-alloyed materials (Table 3). This is due to the high concentration and aggressiveness of the acid, used for pretreatment. It is also impossible to exclude completely the possibility of corrosion of the carbon carrier, which occurs under the conditions of acid treatment, especially when S₄-5N is obtained.

**Table 3.** Composition and structural characteristics of alloyed PtCu/C catalyst (S₄) and samples obtained after its treatment in various acids.

| Sample  | Composition of the solution for treatment | Composition of metallic component | Metals loading, % wt. | Average size of crystallites (XRD), nm |
|---------|------------------------------------------|----------------------------------|----------------------|-------------------------------------|
|         |                                          |                                  | Metals               | Platinum                            |                                    |
| S₄      | -                                        | Pt₃₁Cu₆₉                         | 23.4 ± 0.2           | 13.6 ± 0.2                          | 2.8 ± 0.2                           |
| S₄-1N   | 1M HNO₃                                  | Pt₅₇Cu₄₃                         | 17.4 ± 0.2           | 13.9 ± 0.2                          | 2.4 ± 0.2                           |
| S₄-1Cl  | 1M HClO₄                                 | Pt₅₇Cu₄₃                         | 17.2 ± 0.2           | 13.8 ± 0.2                          | 2.6 ± 0.2                           |
| S₄-1S   | 1M H₂SO₄                                 | Pt₅₈Cu₄₂                         | 17.8 ± 0.2           | 14.4 ± 0.2                          | 2.5 ± 0.2                           |
| S₄-5N   | 5M HNO₃                                  | Pt₆₃Cu₃₇                         | 16.6 ± 0.2           | 13.9 ± 0.2                          | 1.8 ± 0.2                           |

The thermograms of high-temperature oxidation of PtCu/C materials (Figure 5) have the form characteristic of platinum-carbon catalysts [43, 44]. In this case, the most rapid oxidation of the carbon carrier is observed for the initial Pt₃₁Cu₆₉/C (A₄) sample and the Pt₅₇Cu₄₃/C (A₄-1N) material subjected to treatment in 1 M nitric acid. Taking into account the data on the composition and size of the nanoparticles given in Table 3, as well as the results of Ref. [43], one can make an assumption that a longer or shorter duration (temperature range) of intense combustion of the studied materials may be due to a more (rapid oxidation) or less (slow oxidation) ordered distribution of the nanoparticles over the surface of the carbon carrier microparticles.
A change in the composition of the metal component of the catalysts obtained after acid treatment is also confirmed by the results of X-ray diffractometry (Figure 6 a, b): the reflection maximum 111 for all post-treated samples (Figure 6b) shifts to lower 2 theta angles compared to the maximum for the starting material S4, which is due to a decrease of copper concentration in the nanoparticles.

All the catalysts obtained after the acid treatment are characterized by a decrease in the average crystallite sizes from the initial 2.8 to 2.6 - 1.8 nm (Table 3). The most significant decrease in this parameter compared to the initial sample is also observed for sample S4-5N obtained after processing the starting material in the most aggressive medium. Probably under conditions of rapid structural restructuring caused by intense corrosion of copper, the crystallites of this material acquire the highest defectiveness, which makes an additional contribution to the peak broadening in the X-ray diffraction pattern (Figure 6b) and underestimates their average size, calculated by the Scherrer equation.
CVs of the initial (S4) and acid-treated electrocatalysts recorded during standardization have the form characteristic of platinum nanoparticles or its alloys deposited on a carbon support (Figure 7) [17, 21, 24]. In the hydrogen region of CV (0.03–0.3 V), a gradual increase in currents is observed due to the cleaning and development of the surface of the PtCu/C materials, as was noted above (Figure 7). Moreover, there are no peaks in the potential range 0.25–0.35 V associated with the dissolution of copper from the intrinsic phase [17, 24] on the anode CV branches of the initial (S4) and all de-alloyed PtCu/C materials (Figure 7). This indirectly indicates the absence of direct contact of the copper phase with the electrolyte. Note that the development of the surface of sample S4 (“as prepared” state) is much more pronounced (continues for a larger number of cycles) than for PtCu/C catalysts treated in acid (Figure 7).

Moreover, there are no peaks in the potential range 0.25–0.35 V associated with the dissolution of copper from the intrinsic phase [17, 24] on the anode CV branches of the initial (S4) and all de-alloyed PtCu/C materials (Figure 7). This indirectly indicates the absence of direct contact of the copper phase with the electrolyte. Note that the development of the surface of sample S4 (“as prepared” state) is much more pronounced (continues for a larger number of cycles) than for PtCu/C catalysts treated in acid (Figure 7). Obviously, this is largely due to a more intense anodic dissolution of copper atoms. The differences in the character of changes in the anode part of CV observed for S4, as compared to de-alloyed catalysts at potentials of about 1 V (Figure 7), are also apparently due to a significant difference in the surface composition of nanoparticles.

![Cyclic voltammograms of PtCu/C electrocatalysts corresponding to 1, 25, 50, 75 and 100d standardization cycles. The arrows indicate the direction of CV displacement in the standardization process.](image)

**Table 4.** Composition and parameters characterizing electrocatalysts electrochemical behavior

| Sample      | Pt and Cu ratio in the sample after completion of electrochemical measurements | ECSA, m²/g(Pt) | $E_{1/2}$, V (1600 min⁻¹) | Kinetic current, $j_k$ (at $E = 0.90$ V) | $n$ |
|-------------|---------------------------------------------------------------------------------|----------------|--------------------------|----------------------------------------|-----|
| S4          | Pt₆₇Cu₃₃                                                                         | 39.4           | 0.90                     | 161                                    | 4.1 |
| S4-1N       | Pt₆₅Cu₃₅                                                                         | 41.2           | 0.90                     | 200                                    | 4.9 |
| S4-1Cl      | Pt₆₅Cu₃₁                                                                         | 35.7           | 0.89                     | 130                                    | 3.6 |
| S4-1S       | Pt₆₅Cu₃₂                                                                         | 35.2           | 0.90                     | 117                                    | 3.3 |
| S4-5N       | Pt₇₂Cu₂₈                                                                         | 39.2           | 0.90                     | 140                                    | 3.6 |
From the comparison of the metal component composition of the platinum-copper catalysts before (Table 3) and after measurement completion (Table 4), it follows that in the process of conducting electrochemical measurements (standardization of electrodes, ECSA determination, evaluation of the catalysts activity in ORR at different speeds of the disk electrode rotation) all the studied samples lose some amount of copper. The loss is from 77% (catalyst S4) to 28–39% (acid pretreated materials) of the initial copper concentration in the catalysts deposited on the electrode. In this case, the lowest degree of copper recovery is characteristic of catalysts A4-1N and A4-5N, treated in the nitric acid solutions (Tables 3, 4).

Electrocatalysts activity in the reaction of oxygen electroreduction (ORR) was studied by recording voltammograms with a linear scan of potential for RDE in an oxygen atmosphere. Normalized, as described in the Experimental section, LSVs are shown in Figure 8a. The relationship between the current strength and rotation speed was analyzed at potential 0.90 V in the Koutetsky-Levich coordinates (Figure 8b).

The voltammograms obtained for all studied PtCu/C catalysts (Figure 8a) are characterized by close values of a half-wave potential of the oxygen reduction reaction (Table 4). The analysis of the straight-line dependences in the coordinates 1/I - ω⁻¹/² (Figure 8, b) made it possible to calculate the kinetic currents of ORP and confirmed the four-electron reaction mechanism characteristic of platinum (the number of electrons n ~ 4) (Table 4). The order in the increase of catalysts mass-activity can be presented as follows: S4-1S<S4-1Cl<S4-5N<JM20<S4-1N (Table 4). The combination of the minimum specific and relatively high mass activity observed for JM20 is due to a higher ECSA of this catalyst compared to platinum-copper materials. Note that the highest residual concentration of the copper atoms, mainly localized in the inner layers of PtCu nanoparticles, may be one of the reasons for the higher activity of S4-1N and S4 samples in ORR.

It is important to note that the selective dissolution of copper during standardization of the electrodes and electrochemical measurements that followed, takes place on all PtCu/C catalysts. However, for the samples previously treated in acid, as was previously established for samples of the S1 - S3 series, the amount of copper passing into the solution, and the change in the composition for all de-alloyed catalysts obtained from S4, is much less pronounced than for the initialPt1:Cu9/C (S4) catalyst (compare the composition of the catalysts before (Table 3) and after (Table 4) electrochemical measurements). This confirms our assumption that an “acid treatment” can be used to wash out “loosely bound” copper from the platinum catalysts, as its presence makes their use in the MEA problematic.

It is known that the catalysts exhibiting high ORR activity in the electrochemical cell do not always demonstrate similar results when tested in the MEA [1-3]. In the case of platinum-copper catalysts, the negative effect can also be associated with the poisoning of Nafion by copper cations, formed as a result.
of its selective dissolution from the bimetallic nanoparticles [25-27]. Based on the results of electrochemical behavior of alloyed/de-alloy PtCu/C materials, we synthesized the platinum-copper catalyst S5, an analogue of S4, which contained a greater loading of metals. Through a three-hour treatment of portions of this material in 1 M sulfuric and nitric acids at 25 °C, de-alloyed samples of similar composition— S5N and S5S, respectively, were obtained. The mass fraction of platinum in these catalysts was 18.3 ± 0.5 wt%, and the ratio of platinum and copper corresponded to the formula Pt₅₃Cu₄₇. These catalysts were used to form the cathodic and anodic catalytic layers during the assembly of H₂/air PEMFC MEAs. The MEAs were assembled using a similar technology based on the commercial Pt/C PM40 catalyst (40% wt. Pt, PROMETHEUS R&D, Russia), which previously demonstrated high functional characteristics [45]. The conditions for conducting MEAs life tests, which included 5000 cycles of potential changes, are described in the Experimental section, and some test results are presented in Figures 9, 10.

![Image](image1.png)

**Figure 9.** Volt-ampere and watt-ampere characteristics of MEAs after 100, 300 and 2000 stress testing cycles. Catalyst layers were prepared using PM40, S5N, and S5S catalysts. The loading of platinum in each catalytic layer was of 0.4 mg/cm².

![Image](image2.png)

**Figure 10.** The maximum power of MEAs based on the studied PtCu/C and Pt/C catalysts after 100, 300, 2000 and 5000 cycles of stress testing in the membrane-electrode assembly.

Initially, all three MEAs showed similar current-voltage and watt-ampere characteristics (Figure 9a). However, as the cycling progressed, a decrease in the characteristics of the Pt / C catalyst based MEAs was observed, while both MEAs with Pt/C catalysts during 2000 cycles even slightly increased their characteristics (Figure 9b, c). After 2800 sweep cycles, the Pt/C-based MEA tests had to be stopped due to electrode degradation. At the same time, for the MEA based on A55 after 5000 cycles of potential scanning, only a small decrease in power was recorded, while for the MEA containing A5N, the current-voltage characteristics even slightly increased (Figure 10). Thus, the test results of electrocatalysts in the MEAs (Figure 9, 10) correlate well with the results obtained by studying their activity in ORR (Tables 2, 4) and stability, when tested in an electrochemical cell [46]. The negative influence of Cu²⁺ cations formed during the operation of the de-alloyed PtCu/C electrocatalysts on proton conductivity (membrane resistance) of Nafion was not recorded.

4. Conclusions
PtCu/C catalysts of various compositions (1.7 ≤ x ≤ 2.9), containing bimetallic nanoparticles in which the platinum concentration increases from the center to the surface of the particles were synthesized by sequential chemical reduction of copper, and next, of copper and platinum from the solutions of their precursors. By treating portions of PtCu/C materials in the solutions of acids of different nature and concentration, the de-alloyed PtCu_{x-y}/C catalysts (13.8 - 26.4% wt. Pt) with a reduced copper content (x-y = 0.58 - 0.90) were obtained.

Both alloyed and de-alloyed catalysts are characterized by a decrease in the copper content in their composition as a result of electrochemical standardization of their surface. Nevertheless, for de-alloyed materials, the amount of copper transferred to the electrolyte is considerably less. The composition of the post-treated PtCu_{x-y}/C catalysts weakly depends on their initial composition and, to a greater extent, on the acid treatment conditions: the higher the acid concentration and the processing time, the lower the residual copper content is. We have not established a significant effect of the initial composition and pretreatment conditions of the catalysts on their ECSA, measured after standardization. The ECSA values of the alloyed and de-alloyed electrocatalysts measured after standardization of the electrodes ranged from 31 to 43 m²/g (Pt), while the ECSA values of the commercial Pt/C catalyst (40% wt. Pt) used for comparison, were about 80 m²/g(Pt).

Despite the lower values of ECSA, the supported platinum-copper catalysts exhibit significantly higher mass activity and specific activity in MOR compared to the commercial Pt/C electrocatalyst. The difference in mass activity of the platinum-copper and platinum catalysts in ORR is not that crucial, however, some samples of PtCu/C catalysts were also more active than Pt/C.

Pretreatment of PtCu/C catalysts in acids leads to the leaching of a significant fraction of copper from the nanoparticles, but it does not cause a decrease in their activity in MOR and ORR. The fact is that the electrochemical standardization of the catalysts causes an even higher degree of extraction of a “weakly bound” copper from the nanoparticles, therefore their initial composition does not play the leading role. At the same time, the conditions of acid pretreatment of the initial PtCu/C catalysts can have some effect on the change in the composition/structure of the nanoparticles and, as a consequence, on their electrochemical behavior. For example, the highest activity in ORR was demonstrated by a de-alloyed catalyst pretreated in 1 M HNO₃. The influence of acid treatment conditions on the structural and electrochemical characteristics of platinum-copper catalysts was also noted by Hodnik et al. [29, 47]. Therefore, acid treatment conditions can be optimized.

The combination of high activity in ORR and satisfactory resistance to selective dissolution of copper made it possible to conduct life tests of de-alloyed PtCu/C electrocatalysts in the H₂/air MEA. At the initial stage of stress testing, the current-voltage characteristics of MEAs containing bimetallic catalysts and MEA, based on commercial Pt/C, were close but during stress testing the platinum-copper catalysts showed much higher stability. As a result, the life tests of the MEA with a Pt/C catalyst were completed after 2800 cycles of potential change; when de-alloyed PtCu/C catalysts were used, the MEAs successfully withstood 5000 cycles, and one of them even slightly increased its power. These results correlate well with the data on the increased stability of PtCu/C catalysts, which were obtained in [46] from stress testing of PtCu/C and Pt/C catalysts in an electrochemical cell.

Note that we did not detect the negative effect of copper cations formed during the stress testing of the de-alloyed catalysts on the current-voltage characteristics of the MEA. This fact confirms the assumption that the acid pretreatment of PtCu/C materials allows one to obtain the de-alloyed PtCu_{x-y}/C catalysts that possess not only high durability, but also stability which is sufficient from the point of the copper selective dissolution.

The results of the study are in good agreement with the data, that we have previously obtained [22, 46] and with those published in literature [11, 28, 29]. In our opinion, they also confirm the feasibility of further de-alloyed PtCu_{x-y}/C catalysts testing in the fuel cell MEAs.

Author Contributions: V.M. - studied electrochemical behavior of the catalysts in the methanol electrooxidation reaction, took part in the results analysis and the article design; V.G. and A.A. - worked out and developed the general idea and methodology of the project, analyzed the results obtained (with A.N. participation); Olga Spiridonova - prepared materials, studied their electrochemical behavior in ORR; Olga Safronenko, A.A., V.G. – wrote the script, reviewed and edited the article; S.B. - analyzed the composition and studied of the structural characteristics of the materials obtained, participated in the results discussion; A.N. – developed methodology of the
MEAs performance study; A.T., N.G. and N.Z. studied the catalysts behavior in MEAs, analyzed and described the results.

**Funding:** This research was funded by Southern Federal University, Russia.

**Conflicts of Interest:** The authors declare no conflict of interest.

**References**

1. O’Hayre, R.P. Fuel cell fundamentals, 2nd ed.; Publisher: Wiley & Sons Ltd., New York, USA, 2009; pp. 57–67.
2. Qin, C.; Wang, J.; Yang, D.; Li, B.; Zhang, C. Proton Exchange Membrane Fuel Cell Reversal: A Review. *Catalysts* 2016, 6, 197–218.
3. Kongkanand, A.; Mathias, M.F. The Priority and Challenge of High-Power Performance of Low-Platinum Proton-Exchange Membrane Fuel Cells. *J. Phys. Chem. Lett.* 2016, 7, 1127–1137.
4. Crawley, G. Proton exchange membrane (PEM) fuel cells: opening doors to fuel cell commercialization. *Fuel Cell Today* 2006, 1–12.
5. Gong, L.; Yang, Z.; Li, K.; Xing, W.; Liu, C.; Ge, J. Recent development of methanol electrooxidation catalysts for direct methanol fuel cell. *J. Energy Chem.* 2018, 27, 1618–1628.
6. Basri, S.; Kamarudin, S.K.; Daud, W.R.W.; Yaakub, Z. Nanocatalyst for direct methanol fuel cell (DMFC). *Int. J. Hydrogen Energy* 2010, 35, 7957–7970.
7. Xia, Z.; Zhang, X.; Sun, H.; Wang, S.; Sun, G. Recent advances in multi-scale design and construction of materials for direct methanol fuel cells. *Nano Energy* 2019, 65, 104048.
8. Holton, O.T.; Stevenson, J.W.; The Role of Platinum in Proton Exchange Membrane Fuel Cells. *Platin. Met. Rev.* 2013, 57, 259–271.
9. Kyu, K.; Lee, J.H.; Yoon, Ch.; Kwon, Y.; The Effect of Platinum-Based Bimetallic Electrocatalysts on Oxygen Reduction Reaction of Proton Exchange Membrane Fuel Cells. *Int. J. Electrochem. Sci.* 2013, 8, 11752–11767.
10. Long, N.V.; Yang, Y.; Minh, Thi.C.; Minh, N.V.; Cao, Y.; Nogami, M. The development of mixture, alloy, and core-shell nanocatalysts with nanomaterial supports for energy conversion in low-temperature fuel cells. *Nano Energy* 2013, 2, 636–676.
11. Bele, M.; Gatalo, M.; Jovanović, P.; Ruiz-Zepeda, F.; Šala, M.; Šest, E.; Gaberšček, M. Insight on Single Cell Proton Exchange Membrane Fuel Cell Performance of Pt-Cu/Cathode. *Catalysts* 2019, 9, 544–557.
12. Lv, H.; Li, D.; Strmcnik, D.; Paulikas, A.P.; Markovic, N.M.; Stamenkovic, V.R. Recent advances in the design of tailored nanomaterials for efficient oxygen reduction reaction. *Nano Energy* 2016, 29, 149–165.
13. Caballero-Manrique, G.; Brillas, E.; Centellas, F.; Garrido, J.; Rodríguez, R.; Cabot, P.-L. Electrochemical Oxidation of the Carbon Support to Synthesize Pt(Cu) and Pt-Ru(Cu) Core-Shell Electrocatalysts for Low-Temperature Fuel Cells. *Catalysts* 2015, 5, 815–837.
14. Kaewsai, D.; Hunsom, M. Comparative Study of the ORR Activity and Stability of Pt and PtM (M = Ni, Co, Cr, Pd) Supported on Polyaniline/Carbon Nanotubes in a PEM Fuel Cell. *Nanomaterials* 2018, 8, 299–319.
15. Singh, R.N.; Awasthi, R.; Sharma, C.S. Review: An overview of recent development of platinum based cathode materials for direct methanol fuel cells. *Int. J. Electrochem. Sci.* 2014, 9, 5607–5639.
16. Ramli, Z.A.C.; Kamarudin, S.K. Platinum-Based Catalysts on Various Carbon Supports and Conducting Polymers for Direct Methanol Fuel Cell Applications: a Review. *Nanoscale Res. Let.* 2018, 13, 410–435.
17. Wang, Y.; Zhou, H.; Sun, P.; Chen, T. Exceptional methanol electro-oxidation activity by bimetallic concave and dendritic Pt-Cu nanocrystals catalysts. *J. Power Sources* 2014, 245, 663–670.
18. Gómez–Marín, A.M.; Ticianelli, E.A. A reviewed vision of the oxygen reduction reaction mechanism on Pt-based catalysts. *Curr. Opin. Electrochem.* 2018, 9, 129–136.
19. Neergat, M.; Rahul, R. Unsupported Cu-Pt Core-Shell Nanoparticles: Oxygen Reduction Reaction (ORR) Catalyst with Better Activity and Reduced Precious Metal Content. *J. Electrochem. Soc.* 2012, 159, F234–F241.
20. Gasteiger, H.A.; Kocha, Sh.S.; Sompalli, B.; Wagner, F.T. Activity benchmarks and requirements for Pt, Pt-alloy, and non-Pt oxygen reduction catalysts for PEMFCs. *Appl. Catal. B: Environ.* 2005, 56, 9–35.
21. Oezaslan, M.; Hasche, F.; Strasser, P.; PtCu3, PtCu and Pt3Cu alloy nanoparticle electrocatalysts for oxygen reduction reaction in alkaline and acidic media. *J. Electrochem. Soc.* 2012, 159, B444–B454.
22. Alekseenko, A.A.; Guterman, V.E.; Belenov, S.V.; Menshikov, V.S.; Tabachkova, N.Yu.; Safronenko, O.I.; Moguchikh, E.A. Pt/C electrocatalysts based on the nanoparticles with the gradient structure. *Int. J. Hydrogen Energy* 2018, 43, 3676–3687.
23. Stamenkovic, V.R.; Mun, S.B.; Mayrhofer, K.J.J.; Ross, Ph.N.; Markovic N.M. Effect of surface composition on electronic structure, stability, and electrocatalytic properties of Pt-transition metal alloys: pt-skin versus pt-skeleton surfaces. *J. Am. Chem. Soc.* 2006, 128, 8813–8819.
24. Zhu, H.; Li, X.; Wang, F. Synthesis and characterization of Cu@Pt/C core-shell structured catalysts for proton exchange membrane fuel cell. *Int. J. Hydrogen Energy* 2011, 36, 9151–9154.
25. Shabani, B.; Hafttananian, M.; Khamani, S.; Ramiar, A.; Ranjbar, A.A. Poisoning of proton exchange membrane fuel cells by contaminants and impurities: Review of mechanisms, effects, and mitigation strategies. J. Power Sources 2019, 427, 21–48.

26. Zamel, N.; Li, X. Effect of contaminants on polymer electrolyte membrane fuel cells. Prog. Energy Combust. Science 2011, 37, 292–329.

27. Cheng, X.; Shi, Z.; Glass, N.; Zhang, L. A review of PEM hydrogen fuel cell contamination: Impacts, mechanisms, and mitigation. J. Power Sources 2007, 165, 739–756.

28. Kühl, S.; Strasser, P. Oxygen Electrocatalysis on Dealloyed Pt Nanocatalysts. Top. Catal. 2016, 59, 1628–1637.

29. Gatalo, M.; Bel, M.; Ruiz-Zepeda, F.; Šest, E.; Šala, M.; Kamšek, A.R.; Maselj, N.; Galun, T.; Hodnik, N.; Gabersček, M. Double Passivation Water Based Galvanic Displacement Method for Reproducible Gram Scale Production of High Performance Pt-allyl Electrocatalysts. Angew. Chem. Int. Ed. 2019, 58, 18096–18101.

30. Guterman, V.E.; Belenov, S.V.; Alekseeenko, A.A.; Tabakhova, N.Yu.; Volochaev, V.A.; The relationship between activity and stability of deposited platinum-carbon electrocatalysts. Russ. J. Electrochem. 2017, 53, 531–539.

31. Ghavidel, Z.M.R.; Monteverde Videla, A.H.A.; Specchia, S.; Easton, E.B. The relationship between the structure and ethanol oxidation activity of Pt-Cu/C alloy catalysts. Electrochim. Acta 2017, 230, 58–72.

32. Marcu, A.; Toth, G.; Srivistava, R.; Strasser, P. Preparation, characterization and degradation mechanisms of PtCu alloy nanoparticles for automotive fuel cells. J. Power Sources 2012, 208, 288–295.

33. Volochaev, V.A.; Belenov, S.V.; Alekseeenko, A.A.; Guterman, V.E. On the possibilities of recognizing the architecture of binary Pt-M nanoparticles. Nanotechnologies in Russia 2017, 5-6, 2278–235.

34. Shinozaki, K.; Zack, J.W.; Pylypenko, S.; Pivovar, B.S.; Kocha, S.S. Oxygen reduction reaction measurements on platinum electrocatalysts utilizing rotating disk electrode technique: II. Influence of ink formulation, catalyst layer uniformity and thickness. J. Electrochem. Soc. 2015, 162, F1384–F1396.

35. Van der Vliet, D.; Strmcnik, D.S.; Wang, C.; Stamenkovic, V.R. On the importance of correcting for the uncompensated Ohmic resistance in model experiments of the oxygen reduction reaction. J. Electroanal. Chem. 2010, 647, 29–34.

36. Khudhayer, W.J.; Kariuki, N.N.; Wang, X.; Myers, D.J. Oxygen Reduction Reaction Electrocatalytic Activity of Glancing Angle Deposited Platinum Nanorod. J. Electrochem. Soc. 2011, 158, 1029–1041.

37. Shinozaki, K.; Zack, J.W.; Pylypenko, S.; Richards, R.M.; Pivovar, B.S.; Kocha, S.S. Benchmarking the oxygen reduction reaction activity of Pt-based catalysts using standardized rotating disk electrode methods. Int. J. Hydrogen Energ. 2015, 40, 16820–16830.

38. Trindell, J.A.; Duan, Zh.; Henkelman, G.; Crooks, R.M. Well-Defined Nanoparticle Electrocataltysts for the Refinement of Theory. Chem. Rev. 2020, 120, 814–850.

39. Ishikawaa, H.; Sugawaraa, Y.; Inoueab, G.; Kawasec, M. Effects of Pt and ionomer ratios on the structure of catalyst layer: A theoretical model for polymer electrolyte fuel cells. J. Power Sources, 2018, 374, 196–204.

40. Nechitaiov, A.A.; Glebova, N.V.; Tomasov, A.A.; Krasnova, A.O.; Zelenina N.K. Study of the Heterogeneity of a Mixed-Conducting Electrochemical Electrode. Tech. Phys. 2019, 64, 899–907.

41. Available online: https://www.energy.gov/sites/prod/files/2014/02/18/fctt_roadmap_june2013.pdf (accessed on 10 Desemver 2019).

42. Pryadchenko, V.V.; Srbionyan, V.V.; Belenov, S.V.; Volochaev, V.A.; Kurzin, A.A.; Avakyan, L.A.; Zizak, I.; Guterman, V.E.; Bugaev, L.A. Bimetallic PtCu Nanoparticles in PtCu/C Electrocatalysts: Structural and Electrochemical Characterization. Appl. Catal. A-Gen, 2016, 525, 226–236.

43. Guterman, V.E., Belenov, S.V., Krikov, V.V., Vysochina, LL., Yohannes, W., Tabakhova, N.Yu., Balakshina, E.N. Reasons for the differences in the kinetics of thermal oxidation of Pt/C nanomaterials. J. of Phys. Chem. C. 2014, 118, 23835–23844.

44. Leontyev, I.N.; Leontyeva, D.V.; Kuriganova, A.B.; Popov, Y.V.; Maslova, O.A.; Glebova, N.V.; Smirnova, N.V. Characterization of the electrocatalytic activity of carbon-supported platinum-based catalysts by thermal gravimetric analysis. Mendeleev Commun. 2015, 25, 468–469.

45. Guterman, V.E.; Belenov, S.V.; Alekseenko, A.A.; Moguchikh, E.A.; Kirakosyan, S.A.; Danilenko, M.V. New platinum-containing electrocatalysts for PEM Fuel Cells. XXI Mendeleev Congress on General and Applied Chemistry, Saint Petersburg, Russia, September 9-13, Publisher: Saint Petersburg, Russia, 2019.

46. Alekseenko, A.A.; Moguchikh, E.A.; Safronenko, O.I.; Guterman, V.E. Durability of de-allyloyed Pt/C electrocatalysts. Int. J. Hydrogen Energ. 2018, 43, 22885–22895.

47. Hodnik, N.; Cherevko, S. Spot the Difference at the Nanoscale: Identical Location Electron Microscopy in Electrocataltysis, Curr. Op. Electrochem. 2019, 15, 73–82.