A highly selective fluorescent probe for detection of Cd\(^{2+}\) and HSO\(_3^-\) based on photochromic diarylethene with a triazole-bridged coumarin-quinoline group

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A novel photochromic diarylethene containing a quinoline-linked 3-aminocoumarin Schiff base unit (10) was synthesized and used for the selective detection of Cd\(^{2+}\) and HSO\(_3^-\). The synthesized probe exhibited a straightforward response for the selective detection of Cd\(^{2+}\). Its fluorescence emission redshifted ~126 nm and was enhanced 24.9 fold in the presence of Cd\(^{2+}\). Meanwhile, the fluorescence color of 10 changed from dark cyan to golden yellow. The binding stoichiometry between 10 and Cd\(^{2+}\) was determined to be 1 : 1. A molecular logic circuit with three inputs and one output was successfully constructed with its light and metal-responsive behaviors. In addition, 10 was able to selectively recognize HSO\(_3^-\) with a 135-fold enhanced fluorescence emission and a notable fluorescence color change from dark cyan to bright cyan. The \(^1\)H NMR and mass spectrometry analyses suggest that the HSO\(_3^-\) sensing of 10 is based on the hydrolysis of the Schiff base group of 10.

1. Introduction

Among various heavy metal ions, cadmium\([iv]\) is one of the most dangerous ions due to its high toxicity and carcinogenicity.\(^{1,2}\) As we know, Cd\(^{2+}\) has been widely used in many fields such as industry, agriculture, metallurgy, etc.\(^{3,4}\) However, high levels of Cd\(^{2+}\) can impose huge threats to the environment and human health due to its bioaccumulation through the food chain.\(^{5,6}\) Excessive intake of Cd\(^{2+}\) can not only increase the risk of cancer, heart disease, cardiovascular diseases and diabetes,\(^{7-10}\) but also damage the liver and kidneys. As a highly toxic heavy metal, cadmium has been listed on the “CERCLA Priority List of Hazardous Substances” of Toxic Substances and Disease Registry (ATSDR).\(^{11}\) Thus, efficient methodologies for selective detection of Cd\(^{2+}\) are desperately needed. One of the greatest challenges in Cd\(^{2+}\) detection is the highly similar binding properties of Cd\(^{2+}\) and Zn\(^{2+}\) that are located in the same group of the periodic table.\(^{12}\) Therefore, the development of highly selective sensors of Cd\(^{2+}\) without interferences from Zn\(^{2+}\) under physiological conditions has become a hot research topic.

Anions play important roles in biological, medical, environmental and chemical sciences.\(^{13-15}\) Among various anions, bisulfite (HSO\(_3^-\)) is one of the most concerned anions, which has been widely used as an enzyme inhibitor, antimicrobial agent, beverages, and an antioxidant for foods.\(^{16}\) In fact, addition of bisulfite to beer and wine has been customary for centuries in most countries. However, it has been found that some individuals are sensitive to high concentrations of HSO\(_3^-\) with the syndromes of asthmatic attack, gastrointestinal distress, allergic reaction and skin allergy.\(^{17-22}\) Thus, the bisulfite content in foodstuffs has been strictly limited in many countries. For this reason, the selective and sensitive chemosensors of bisulfite are highly desired. To date, many of the reported HSO\(_3^-\) sensors have based on different reaction mechanisms, including coordination to metal ions,\(^{23,24}\) Michael additions,\(^{25,26}\) complexation with amines,\(^{27-32}\) and inductively coupled plasma-atomic emission spectrometry (ICP-AES).\(^{28}\) For instance, Zeng et al. reported a mitochondria-targeted probe derived from the conjugation of carbazole and benzox[e]indolium, to which 1,4-nucleophilic addition reaction with bisulfite occurred.\(^{33}\) However, there are no reports for HSO\(_3^-\) sensors, which based on Schiff base with hydrolysis mechanisms.

A variety of advanced technologies, such as inductively coupled plasma-atomic emission spectrometry (ICP-AES),\(^{37}\) atomic absorption spectroscopy (AAS)\(^{38}\) and inductively coupled plasma-mass spectrometry (ICP-MS),\(^{39}\) have been widely used for the analysis of Cd\(^{2+}\) and HSO\(_3^-\). By comparison, fluorescent probes are superior and have attracted increasing attentions in ion detections in both chemistry and biology due to their high sensitivity, good selectivity, easy operation, and low cost.\(^{40-46}\)

Among the reported photo-responsive materials, diarylethene is one of the most promising photo-switchable molecules, which have exhibited great potentials in optical information storage media and photonic switch devices with.
their excellent photo reactivity, thermal stability, and fatigue resistance. Moreover, diarylethenes have also been extensively introduced to application in electronic logic devices in recent years, because they exhibit excellent optical performances and multi-response to various ions owing to their special structures. As known to all, electronic equipment has gradually become a necessity in people’s lives and the logic circuits are essential for the fabrication of optoelectronic devices. According to the above advantages, an increasing number of diarylethenes were introduced to fluorescent sensor for detection of different ions in recent years. For instance, Li et al. reported a highly selective fluorescent probe for Cd$^{2+}$ and Zn$^{2+}$ based on a new diarylthene with quinoline–benzimidazole conjugated system, Zhang et al. reported a fluorescent sensor for Cd$^{2+}$ based on a new diarylthene with a 1,8-naphthyridine unit, Duan et al. studied a fluorescent probe for Cd$^{2+}$ based on a diarylthene with pyridinepiperazine-linked hydroxyquinoline group. However, there are no diarylthene-based chemosensors that exhibit fluorescence response to both Cd$^{2+}$ and HSO$_3^−$. Thus, development of fluorescent probes toward Cd$^{2+}$ and HSO$_3^−$ are still challenging.

In this study, we designed and synthesized a novel photochromic diarylthene bearing a quinoline-linked 3-amino-coumarin Schiff base unit (1O), and its photochromic and fluorescent responses to Cd$^{2+}$ and HSO$_3^−$ were systematically discussed. Scheme 1 shows the photochromic process of 1O.

2. Experiments

2.1 General methods

All solvents were of analytical grade and distilled before use. Other reagents were used as-received. $^1$H NMR and $^{13}$C NMR spectra were recorded on a Bruker AV400 spectrometer (400 MHz) using DMSO-d$_6$, CDCl$_3$ and CD$_3$CN as the solvents and tetramethylsilane (TMS) as the internal standard. Mass spectra were obtained using an Agilent 1100 Ion Trap LC/MS MSD system. Infrared spectra (IR) were collected on a BrukerVertex-70 spectrometer. Melting points were measured using a WRS-1B melting point apparatus. All metal ions, except for Hg$^{2+}$ and K$^+$ that were prepared with their chloride salts, were obtained by dissolving the corresponding metal nitrates (0.1 mmol) in distilled water (10 mL). All anions were obtained by dissolving the corresponding potassium or sodium salts (0.1 mmol) in distilled water (10 mL). The EDTA solution was prepared with ethylenediaminetetraacetic acid disodium salt (Na$_2$EDTA) (1.0 mmol) in distilled water (10 mL). Fluorescence spectra were measured with a Hitachi F-4600 fluorescence spectrophotometer. UV-vis absorption spectra were recorded on an Agilent 8453
UV/vis spectrophotometer equipped with an MUA-165 UV lamp and MVL-210 visible lamp for photoirradiation. Fluorescence quantum yield was measured using a QY C11347-11 absolute PL quantum yield spectrometer. Photoirradiation was conducted on a setup consisting of an SHG-200 UV lamp, Cx-21 ultraviolet fluorescence analysis cabinet, and BMH-250 visible lamp.

**Fig. 1**  (A) Changes in the absorption spectrum of 1O in acetonitrile (2.0 × 10^{-5} mol L^{-1}) upon alternating irradiation with UV and visible light; (B) fatigue resistance of 1O in acetonitrile under the alternating irradiation at room temperature.

**Fig. 2**  Changes in the absorption spectrum and color of 1O in acetonitrile (2.0 × 10^{-5} mol L^{-1}) in response to Cd^{2+}/EDTA and light stimuli. (A) Spectral changes of 1O induced by Cd^{2+}/EDTA; (B) spectral changes of 1O + Cd^{2+} induced by UV/vis light; (C) spectral changes of 1C induced by Cd^{2+}/EDTA.
2.2 Synthesis of 1O

1O was synthesized following the route shown in Scheme 2. Compounds 2 and 3 were prepared by the methods reported in literature.58,59

2.2.1 Synthesis of compound 4. The NaVc (0.04 g, 0.20 mmol) and CuSO4 (0.025 g, 0.10 mmol) solution in water (10 mL) was added to a solution of 2 (0.55 g, 1.00 mmol) and 3 (0.21 g, 1.00 mmol) in THF (40 mL) under stirring at room temperature. The reaction mixture was stirred at room temperature for 12 h to allow the Cu(I)-catalyzed 1,3-dipolar cycloaddition reaction of 2 and 3, and extracted with CH2Cl2. The organic phase was dried over Na2SO4, filtered and evaporated. The crude product was purified by silica gel column chromatography using petroleum dichloro/ethyl acetate (v/v = 1/1) as the eluent to afford 0.51 g compound 4 as a purple solid in 67% yield.1H NMR (400 MHz, CDCl3, ppm): δ 1.94 (s, 3H, –CH3), 2.29 (s, 3H, –CH3), 5.30 (s, 1H, –CH2−N−), 5.52 (s, 2H, –CH2−O−), 5.59 (s, 2H, Ar−H), 7.16 (s, 1H, Ar−H), 7.23 (s, 1H, Ar−H), 7.25 (s, 1H, Ar−H), 7.33 (q, 2H, Ar−H), 7.41 (q, 3H, Ar−H), 7.50 (d, 1H, Ar−H), 7.59 (m, 2H, Ar−H), 7.73 (t, 1H, Ar−H), 8.06 (d, 1H, Ar−H), 8.28 (d, 1H, Ar−H), 10.26 (s, 1H, −CHO).

2.2.2 Synthesis of (1O). Compound 4 (0.23 g, 0.3 mmol) was refluxed with 3-aminocoumarin (0.048 g, 0.3 mmol) in anhydrous methanol (5.0 mL) at 80 °C under stirring for 1 h, cooled to room temperature and filtered. The residue was washed with cold methanol to afford 1O (0.22 g) in 81% yield. Mp: 439–440 K; 1H NMR (400 MHz, DMSO, ppm): δ 2.35 (s, 3H, −CH3), 3.00 (s, 3H, −CH3), 5.45 (s, 2H, −CH2−N−), 5.72 (s, 2H, −CH2−O−), 6.18 (d, 1H, −CH=−N−), 7.15 (s, 1H, Ar−H), 7.29 (d, 1H, Ar−H), 7.35 (t, 3H, Ar−H), 7.40 (s, 1H, Ar−H), 7.42 (s, 2H, Ar−H), 7.47 (q, 2H, Ar−H), 7.52 (s, 1H, Ar−H), 7.54 (s, 1H, Ar−H), 7.60 (q, 2H, Ar−H), 7.62 (s, 1H, Ar−H), 7.79 (d, 1H, Ar−H), 7.94 (d, 1H, Ar−H), 8.51 (d, 2H, Ar−H); 13C NMR (100 MHz, DMSO, TMS): δ = 14.56, 15.07, 50.17, 53.01, 63.50, 83.50, 109.65, 111.73, 116.18, 119.37, 120.60, 120.75, 121.62, 121.92, 123.07, 123.36, 124.59, 124.86, 125.29, 125.71, 125.99, 126.36, 127.07, 127.98, 129.35, 130.98, 132.67, 136.35, 137.81, 138.11, 138.42, 141.20, 142.07, 143.62, 144.23, 148.55, 154.11, 154.52, 159.14; IR (KBr, ν, cm−1): 1715 (–C=O), 1628 (–C=N), 1338 (–C–N), 1275 (–C=O); MS (m/z): Calculated for C48H31F6N5O3S2 [M]+: 903.1773, found: 904.1836 for [1O + H]+ and 926.1658 for [1O + Na]+.

3. Results and discussion

3.1 Photochromic of diarylethene 1O

The photochromic behavior of diarylethene 1O induced by UV/vis light was investigated in acetonitrile (C = 2.0 × 10−5 mol L−1) at room temperature. As shown in Fig. 1A, the 1O...
solution in acetonitrile is colorless, exhibiting a sharp absorption peak at 242 nm due to the \( \pi-\pi^* \) transition. A new absorption band centered at 547 nm appeared upon the irradiation with 297 nm light, accompanied with a visible color change from colorless to purple, due to the photocyclization of \( 1O \) to \( 1C \).\(^{60-62} \) The photocyclization reaction reached the photostationary state (PSS) in 4.5 min under the irradiation of UV light, where a clear isosbestic point was observed at 300 nm. These observations suggest that the irritation caused a two-component photochromic reaction.\(^{63} \) The absorption spectrum was successfully restored to the original state, accompanied by the color change from purple to colorless, by irradiating the \( 1C \) solution with visible light. The cyclization and cycloreversion quantum yields were calculated to be 0.139 and 0.324, respectively. The fatigue resistance of \( 1O \) was tested by alternating irradiation with UV and visible light at room temperature. Ten coloration–decoloration cycles of \( 1O \) caused only 6.5% degradation (Fig. 1B).

### 3.2 Changes in absorption spectrum induced by \( Cd^{2+} \)

Fig. 2 shows the absorption spectra and colors of \( 1O \) induced by \( Cd^{2+}/EDTA \) and UV/vis stimuli in acetonitrile. A new absorption band appeared at 392 nm as \( Cd^{2+} \) added and the band intensity increased with the increase of \( Cd^{2+} \) amount and reached the maximum at 1.8 equiv. of \( Cd^{2+} \), accompanied by a notable color change from colorless to yellow, due to the formation of complex \( 1O + Cd^{2+} \). The addition of an excess amount of EDTA did not restore the original color of \( 1O \), suggesting that the \( Cd^{2+} \) sensing process of \( 1O \) was irreversible (Fig. 2A). As depicted in Fig. 2B, the \( 1O + Cd^{2+} \) complex was also photoisomerized under the alternating irradiation with UV and visible light. Upon the irradiation with 297 nm UV light, the yellow solution of \( 1O + Cd^{2+} \) turned plum and a new absorption band emerged at 547 nm (\( \varepsilon_{max} = 8.5 \times 10^3 \text{ mol}^{-1} \text{L cm}^{-1} \)) due to the formation of the closed-ring isomer \( 1C + Cd^{2+} \).\(^{66,67} \) Reversely, the plum solution turned yellow, and the absorption spectrum was restored to that of the open-ring isomer \( 1O + Cd^{2+} \) upon irradiation with visible light. Directly adding \( Cd^{2+} \) to the solution of \( 1C \) weakened the absorption band centered at 322 nm and enhanced the absorption at 393 nm. Meanwhile, the solution color changed from purple to plum, indicating that \( 1C \) and \( Cd^{2+} \) formed \( 1C + Cd^{2+} \) complex. It is worth noting that the complexation between \( 1C \) and \( Cd^{2+} \) was not be reversed by EDTA either (Fig. 2C). These results indicate that both isomers of \( 1O \) have stronger \( Cd^{2+} \) binding ability than EDTA.

### 3.3 Fluorescence response of \( 1O \) to metal ions

The fluorescence and color responses of \( 1O \) (2.0 \( \times \) 10^{-5} mol L^{-1} in acetonitrile) to various metal cations (2.4 equiv.) including \( Cd^{2+} \), \( Cr^{3+} \), \( Al^{3+} \), \( Fe^{3+} \), \( Ni^{2+} \), \( Sr^{2+} \), \( Ag^{+} \), \( Ba^{2+} \), \( Ca^{2+} \), \( Mg^{2+} \), \( Pb^{2+} \), \( K^{+} \), \( Mn^{2+} \), \( Co^{2+} \), \( Zn^{2+} \), \( Hg^{2+} \), \( Pb^{2+} \), \( K^{+} \),
Ca\textsuperscript{2+}, Mg\textsuperscript{2+}, Ba\textsuperscript{2+}, Sr\textsuperscript{2+}, Ag\textsuperscript{+} and Ni\textsuperscript{2+} were investigated with fluorescence spectroscopy. All metal ions except for Cr\textsuperscript{3+}, Al\textsuperscript{3+} and Fe\textsuperscript{3+} caused no obvious changes in the fluorescence emission at 507 nm of 1O (Fig. 3A). The fluorescence emission of 1O at 507 nm was dramatically enhanced as Cr\textsuperscript{3+}, Al\textsuperscript{3+} or Fe\textsuperscript{3+} added, companied by the color change from dark cyan to bright

![Fig. 5](image_url)  
**Fig. 5** Fluorescence responses of 1O (2.0 × 10\textsuperscript{-5} mol L\textsuperscript{-1}) to Cd\textsuperscript{2+}/EDTA in acetonitrile, excited at 392 nm. (A) Fluorescent emission spectra of 1O titrated with different amounts of Cd\textsuperscript{2+}; (B) fluorescence intensities of 1O at 633 nm as titrated with different equiv. of Cd\textsuperscript{2+}; (C) fluorescent emission spectra of 1O + Cd\textsuperscript{2+} under alternating photoirradiation with UV and visible light; (D) Job’s plot showing the 1 : 1 complex of 1O with Cd\textsuperscript{2+}.

![Fig. 6](image_url)  
**Fig. 6** ESI-MS spectra of 1O in the absence of presence of Cd\textsuperscript{2+} in acetonitrile.
cyan (Fig. 3C). However, the maximum emission peak had an obvious red-shift from 507 nm to 633 nm upon the addition of Cd$^{2+}$ to 10 in acetonitrile, we can see this phenomenon more clearly through Fig. 3B. The structural rigidity of the Cd$^{2+}$ complexes and metal binding close to the diarylethene might be the factors that cooperate to induce the large emission peak shift. Meanwhile, it’s emission intensity enhanced by 24.9 folds with the stimulation of 2.4 equiv. of Cd$^{2+}$, accompanied by a color change from dark cyan to golden yellow. These results indicate that 10 can selectively recognize Cd$^{2+}$ over other metal ions in acetonitrile.

To further demonstrate the selectivity of 10 towards Cd$^{2+}$ in acetonitrile, competitive experiments were conducted on Cd$^{2+}$ in the presence of equal equivalent of other competitive metal cations listed above. As shown in Fig. 4, the fluorescence response of 10 to Cd$^{2+}$ was not significantly affected by other competing metal ions, except for Cu$^{2+}$ that caused a reduction in emission intensity. However, compared with the variations of emission intensity and fluorescence color induced by Cd$^{2+}$ alone, the interference of Cu$^{2+}$ is negligible. As mentioned above, most Cd$^{2+}$ fluorescent probes encounter the interference from Zn$^{2+}$ that is located in the same group in the periodic table and has similar chemical properties to Cd$^{2+}$. Yet such interference was not observed in the Cd$^{2+}$ chemosensing of 10 too. These results suggested that 10 is highly selective to Cd$^{2+}$ over other competing cations, even Cr$^{3+}$, Al$^{3+}$, Fe$^{3+}$, and Zn$^{2+}$ in acetonitrile.

**Fig. 7**  (A) Changes in the photochromism, color, and fluorescent of 10 induced by Cd$^{2+}$/EDTA and UV/vis lights; (B) the combinational logic circuit equivalents to the truth table in Table 1, where In1 (297 nm UV light), In2 (500 nm visible light), In3 (Cd$^{2+}$) are inputs, and O1 is output.
The fluorescence responses of 10 to Cd\(^{2+}\)/EDTA and light stimuli were investigated in acetonitrile at room temperature. The emission peak of 10 at 507 nm disappeared and a new peak centered at 633 nm appeared as Cd\(^{2+}\) added due to the formation of 10 + Cd\(^{2+}\) complex (Fig. 5A). The emission intensity at 633 nm linearly increased with the increase of Cd\(^{2+}\) concentration and reached the maximum at 2.4 equiv. Cd\(^{2+}\), followed by a plateau with further titration (Fig. 5B). Meanwhile, the fluorescence color changed from dark cyan to golden yellow. The fluorescence quantum yield of 10 + Cd\(^{2+}\) was determined to be 0.031, 14.5-fold greater than that of 10 (\(\phi_f, 10 = 0.002\)). The 126 nm re-shift of emission peak and up to 24.9-fold enhanced emission intensity suggest that 10 is an ideal fluorescent chemosensor of Cd\(^{2+}\). The subsequent addition of excess EDTA did not restore the fluorescence to the original state of 10, suggesting that the Cd\(^{2+}\) sensing process of 10 was irreversible. The 10 + Cd\(^{2+}\) complex also exhibited photoswitchable fluorescence behaviors upon alternating irradiation with UV and visible light. The fluorescence of 10 + Cd\(^{2+}\) was quenched dramatically with a clear color change from golden yellow to fawn brown upon the irradiation with 297 nm light, due to the formation of weak fluorescence closed-ring isomer of 1C + Cd\(^{2+}\) (Fig. 5C). However, the fluorescence intensity of 10 + Cd\(^{2+}\) was only reduced to ca. 34% at the photostationary state, possibly owing to the incomplete cyclization and the formation of isomers with parallel conformations. The back irradiation with appropriate visible light regenerated the open-ring isomer 10 + Cd\(^{2+}\) and the original emission intensity was recovered.

Additionally, in order to determine the binding stoichiometry between 10 and Cd\(^{2+}\), the Job’s plot with emission intensity at 633 nm as a function of molar fraction of 10 was drawn. The

| Input | Output λ\(_{ab}\) = 393 nm |
|-------|----------------------------|
| In1 (UV) | In2 (Vis) | In3 (Cd\(^{2+}\)) |
| 0 | 0 | 0 | 0 |
| 0 | 0 | 1 | 1 |
| 1 | 0 | 0 | 0 |
| 0 | 1 | 0 | 0 |
| 1 | 0 | 1 | 0 |
| 0 | 1 | 1 | 1 |
| 1 | 1 | 0 | 0 |
| 1 | 1 | 1 | 1 |

The output is defined as 1 if the absorbance at 393 nm is greater than 0.18, otherwise it is defined as 0.

The fluorescence responses of 10 to Cd\(^{2+}\)/EDTA and light stimuli were investigated in acetonitrile at room temperature. The emission peak of 10 at 507 nm disappeared and a new peak centered at 633 nm appeared as Cd\(^{2+}\) added due to the formation of 10 + Cd\(^{2+}\) complex (Fig. 5A). The emission intensity at 633 nm linearly increased with the increase of Cd\(^{2+}\) concentration and reached the maximum at 2.4 equiv. Cd\(^{2+}\), followed by a plateau with further titration (Fig. 5B). Meanwhile, the fluorescence color changed from dark cyan to golden yellow. The fluorescence quantum yield of 10 + Cd\(^{2+}\) was determined to be 0.031, 14.5-fold greater than that of 10 (\(\phi_f, 10 = 0.002\)). The 126 nm re-shift of emission peak and up to 24.9-fold enhanced emission intensity suggest that 10 is an ideal fluorescent chemosensor of Cd\(^{2+}\). The subsequent addition of excess EDTA did not restore the fluorescence to the original state of 10, suggesting that the Cd\(^{2+}\) sensing process of 10 was irreversible. The 10 + Cd\(^{2+}\) complex also exhibited photoswitchable fluorescence behaviors upon alternating irradiation with UV and visible light. The fluorescence of 10 + Cd\(^{2+}\) was quenched dramatically with a clear color change from golden yellow to fawn brown upon the irradiation with 297 nm light, due to the formation of weak fluorescence closed-ring isomer of 1C + Cd\(^{2+}\) (Fig. 5C). However, the fluorescence intensity of 10 + Cd\(^{2+}\) was only reduced to ca. 34% at the photostationary state, possibly owing to the incomplete cyclization and the formation of isomers with parallel conformations. The back irradiation with appropriate visible light regenerated the open-ring isomer 10 + Cd\(^{2+}\) and the original emission intensity was recovered.

Additionally, in order to determine the binding stoichiometry between 10 and Cd\(^{2+}\), the Job’s plot with emission intensity at 633 nm as a function of molar fraction of 10 was drawn. The
concentration of \(1O + Cd^{2+}\) reached the maximum at the molar fraction of \([Cd^{2+}] / ([Cd^{2+}] + [1O]) = \sim 0.5\) (Fig. 5D), indicating that the binding stoichiometry was 1 : 1. The 1 : 1 coordination stoichiometry of \(1O\) with \(Cd^{2+}\) was further confirmed by the ESI mass spectrometry (ESI-MS) analysis. \(1O\) exhibited a characteristic peak at 926.1658 \(m/z\) for \([1O + Na]^+\) (calcd 926.1659) and 904.1836 \(m/z\) for \([1O + H]^+\) (calcd 904.1852), which disappeared as 2.4 equiv. of \(Cd^{2+}\) were added, accompanied by the appearance of a new peak at 1016.0887 \(m/z\) for \([1O + Cd^{2+} + H]^+\) (calcd 1016.0884) (Fig. 6).

### 3.4 Application in logic circuits

Based on the photochromic behaviors of the diarylethene \(1O\) in response to \(Cd^{2+}\) and UV/vis lights, a logic circuit was constructed with light irradiations and \(Cd^{2+}\) as the inputs and the absorbance of \(1O\) at 393 nm as the output (Fig. 7). As shown in Fig. 7A, the absorption of \(1O\) (output) can be effectively modulated by a logic circuit constructed with the three inputs including In1: 297 nm UV light, In2: \(\lambda > 500\ \text{nm}\) visible light, and In3: \(Cd^{2+}\) (Fig. 7B). The output can be either ‘on’ or ‘off’ with the Boolean value of ‘1’ or ‘0’. For example, In1 is ‘on’ with the Boolean value of ‘1’ when 297 nm light is used. Similarly, In2 is ‘1’ as irradiation with appropriate visible light (\(\lambda > 500\ \text{nm}\)) is used and In3 is ‘1’ as \(Cd^{2+}\) is added. The output is considered at ‘on’ state with the Boolean value of ‘1’ if the absorbance at 393 nm is greater 0.18. Otherwise, it is regarded at ‘off’ state with the Boolean value of ‘0’. Table 1 lists all possible strings of the binary inputs and the corresponding output digits. Upon the stimulation of different inputs, the diarylethene can exhibit an on–off–on photochromic switching behavior.

### 3.5 Fluorescence response of \(1O\) to anions

The fluorescence responses of \(1O\) to a variety of anions including \(HSO_3^-\), \(HSO_4^-\), \(F^-\), \(Cl^-\), \(Br^-\), \(I^-\), \(H_2PO_4^-\), \(CO_3^{2-}\), \(HPO_4^{2-}\), \(SO_3^{2-}\), \(SCN^-\), \(SO_4^{2-}\), \(S_2O_3^{2-}\), \(PO_4^{3-}\), \(ClO_4^-\), \(HCO_3^-\), \(NO_3^-\), \(Ac^-\) and \(CN^-\) were measured in acetonitrile to explore its potential application as a selective anion sensor. As shown in Fig. 8A and C for the fluorescence spectra of \(1O\) in the presence of 4.3 equiv. anions, none of them, except for \(HSO_3^-\) caused notably fluorescent emission at 510 nm and a fluorescent color.

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**Fig. 9** Changes in the fluorescence of \(1O\) (2.0 \(\times 10^{-5}\) mol L\(^{-1}\) in acetonitrile) induced by \(HSO_3^-\)/EDTA and light stimuli at room temperature. Excitation: 370 nm. (A) Fluorescence emission spectra of \(1O\) titrated with different amounts of \(HSO_3^-\); (B) the fluorescence intensities of \(1O\) at 510 nm in the presence of different equiv. \(HSO_3^-\); (C) fluorescence emission spectra of \(1O + HSO_3^-\) under the irradiation with UV and visible lights.
change from dark cyan to bright cyan. Compared with that of 1O, the fluorescence intensity was enhanced by 135.5 folds, suggesting that 1O can recognize HSO$_3^-$ . The selectivity of the reorganization was further confirmed by competitive experiments. As depicted in Fig. 8B, all competing anions, except for CN$^-$ showed no obvious interference with the HSO$_3^-$ sensing.

To further illuminate the responsive fluorescent emission of 1O induced by HSO$_3^-$ and UV/vis irradiation, fluorescence titration was conducted at room temperature. As shown in Fig. 9A and B, the fluorescence emission of 1O at 510 nm linearly increased with the increase of HSO$_3^-$ amount, accompanied by a notable fluorescent color change from dark cyan to bright cyan, and reached the maximum at 4.3 equiv. HSO$_3^-$, followed by a plateau as HSO$_3^-$ further added. The addition of excess EDTA did not recover its original fluorescence intensity, indicating that the HSO$_3^-$ sensing process of 1O was

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Fig. 10  (A) Partial $^1$H NMR spectra of 1O in the absence (a) and presence of 4.3 eq. (b) of HSO$_3^-$ in deuterated acetonitrile; (B) ESI-MS spectrum of 1O + HSO$_3^-$.
irreversible (Fig. 9A). The absolute fluorescence quantum yield of IO + HSO₃⁻ was determined to be 0.048, 23 folds higher than that of IO (Φf, IO = 0.002). These results suggest that the diarylethene can also be potentially used as a fluorescent probe for the quantitative detection of HSO₃⁻. The IO + HSO₃⁻ complex also exhibited photoswitchable fluorescence behavior upon alternating irradiation with UV and visible light. The fluorescence of IO + HSO₃⁻ was quenched dramatically with a clear color change from bright cyan to light cyan under the irradiation of 297 nm light, due to the formation of closed-ring isomer 1C + HSO₃⁻ (Fig. 9C). The fluorescence intensity of IO + HSO₃⁻ was weakened to ca. 32% at PSS due to the incomplete cyclization and existence of isomers with parallel conformations. Upon the irradiation of visible light, the fluorescent spectrum was restored to the initial state of IO + HSO₃⁻.

To further explore the interaction between IO and HSO₃⁻ in solution and determine the sensing mechanism, the ¹H NMR titration of IO with HSO₃⁻ was performed in deuterated acetonitrile. IO exhibited a singlet at δ 6.18 ppm of the imine protons (N=CH) (Fig. 10A). As 4.3 equiv. HSO₃⁻ was added to the IO solution, the singlet at δ 6.18 ppm disappeared and a new singlet appeared at δ 10.26 ppm that was attributed to aldehyde protons (−CHO). These results suggest that IO was hydrolyzed to afford compound 4 in the presence of HSO₃⁻ in acetonitrile, similar phenomena were also found in other reported probes.⁶⁶⁻⁷⁰

The mass spectra of IO and IO + HSO₃⁻ more directly evidenced the reaction. As depicted in Fig. 10B, the peaks of IO at 926.1658 m/z for [IO + Na⁺]⁺ (calcd 926.1659) and 904.1836 m/z for [IO + H⁺]⁺ (calcd 904.1852) disappeared after 4.3 equiv. HSO₃⁻ added, and two new peaks appeared at m/z = 759.1340 for compound 4, (calcd C₁₉H₁₉F₂N₃O₃S₂ [M – H⁺]: 759.13216) and 160.0401 for compound 5 (calcd for C₃H₆N₂O₂ [M – H⁺]: 160.03976). Therefore, the enhanced fluorescence and color change of IO in recognizing HSO₃⁻ was due to the cleavage of imine bond in IO that generated aldehyde and amine, e.g. compound 4 and 3-aminocoumarin 3, as depicted in Scheme 3.

### 4. Conclusion

In summary, a novel photochromic chemosensor, a diarylethene bearing a quinoline-linked 3-aminocoumarin group (IO), was designed and synthesized. It exhibited typical photochromism and fluorescent switching properties and high selectivity to Cd²⁺ and HSO₃⁻. IO showed a straightforward response for the selective detection of Cd²⁺, ~126 nm red shifted emission peak and up to 24.9-fold enhanced fluorescence intensity, accompanied by a notable fluorescent color change from dark cyan to golden yellow. The competitive experiments suggest that such selective fluorescent response of IO to Cd²⁺ is not interfered by other competing metal cations including Zn²⁺. The binding stoichiometry between IO and Cd²⁺ was determined to be 1 : 1. Similarly, it was found that HSO₃⁻ was able to enhance the fluorescence emission of IO in acetonitrile up to 135.5-fold and caused a notable fluorescence color change from dark cyan to bright cyan due to the hydrolysis of the C= N bond in IO that produced an aldehyde and amine. These results suggest that the diarylethene derivative has a great application potential as a fluorescent sensor of both Cd²⁺ and HSO₃⁻.

### Conflicts of interest

There are no conflicts of interest to declare.

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