Lead Leaching of Perovskite Solar Cells in Aqueous Environments: A Quantitative Investigation

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1. Introduction

Lead halide perovskite solar cells (PSCs) have emerged as a highly promising next-generation photovoltaic (PV) technology that combines high device performance with ease of processing and low cost. However, the potential leaching of lead is recognized as a major environmental concern for their large-scale commercialization, especially for application areas with significant overlap with human life. Herein, a quantitative kinetic analysis of the Pb leaching behavior of five types of benchmark PSCs, namely, MAPbI$_3$, FA$_{0.95}$MA$_{0.05}$Pb(I$_{0.95}$Br$_{0.05}$)$_3$, Cs$_{0.05}$(FA$_{0.85}$MA$_{0.15}$)$_{0.95}$Pb(I$_{0.85}$Br$_{0.15}$)$_3$, CsPbI$_3$, and CsPbI$_2$Br, under laboratory rainfall conditions is reported. Strikingly, over 60% of the Pb contained in the unencapsulated perovskite devices is leached within the first 120 s under rainfall exposure, suggesting that very rapid leaching of Pb can occur when indoor and outdoor PV devices are subject to physical damage or failed encapsulation. The initial Pb leaching rate is found to be strongly dependent on the types of PSCs, pointing to a potential route toward Pb leaching reduction through further optimization of their materials design. The findings offer kinetic insights into the Pb leaching behavior of PSCs upon aqueous exposure, highlighting the urgency to develop robust mitigation methods to avoid a potentially catastrophic impact on the environment for their large-scale deployment.
Lead (Pb) plays a crucial role in the efficient operation of PSCs. Their ideal electronic configuration (e.g., Pb 6s lone-pair states and inactive Pb 6p orbital) and ionic radii are known as ideal valent cations for perovskite light-absorbing materials. However, Pb is widely recognized as a harmful substance, which poses potential risks both to the natural and to the built environment. To minimize Pb content in PSCs, there has been increasing research effort in the development of Pb-free PSCs and Pb-replacement methods have been explored to reduce Pb leakage of PSCs through further optimization of their device architecture. For example, using Pb-absorbing or self-healing encapsulation materials, Pb leakage from damaged PSCs was significantly reduced without affecting the efficiencies of the cells. Huang et al. used self-healing epoxy resin encapsulation and Pb-absorbing ionogel-based encapsulation respectively, achieving simultaneous Pb leakage reduction and stability improvement of perovskite modules. Zhu et al. reported on-device sequestration of more than 96% of Pb leakage, achieved through Pb-absorbing coatings on both sides of PSCs.

Concerning the potential environmental impact of Pb-based PSCs, the leaching of Pb from PSCs upon aqueous exposure can be triggered by the degradation and decomposition of the perovskite materials, further accelerated by the accidental damage and disposal of PSCs. The leaked Pb can migrate in the environment and subsequently accumulate in living organisms via bioaccumulation and human via the food chain, known to be tenfold the mobility of other Pb compounds in daily human life. There has been increasing research effort in Pb leakage analyses and life cycle assessments of PSCs to understand their leaching process and evaluate their potential environmental impacts. Cahn et al. studied MAPbI3 perovskite films exposed under aqueous environments with different pH values from 4.2 to 8.1 and found that the difference in Pb leaching amount is not significant. Nazeeruddin et al. demonstrated a two-stage dissolution process of Pb2+ ions from three different types of PSCs (MAPbI3, FA0.85MA0.15Pb(I0.85Br0.15)3, and Cs0.15FA0.85Pb(I0.85Br0.15)3) and found that the amount of Pb leaching is within the same order of magnitude under simulated normal and acidic rainwater (pH = 5.6 and 4.5). Zhao et al. investigated the leaching concentrations of several heavy metals (e.g., Pb, Zn, Ni) in PSCs and confirmed that the Pb leaching amount surpasses the hazardous waste standard of 5 mg L−1. Cheng et al. evaluated the potential Pb contamination of cracked perovskite modules under simulated working conditions and demonstrated that a widely used “sandwich” encapsulation strategy (utilizing polyethylene and polyisobutylene as encapsulants) could effectively prevent Pb leaching from perovskite modules. Among the abovementioned studies, only a few have conducted quantitative analysis of Pb leaching of PSCs. Furthermore, a comprehensive comparison of the kinetic Pb leaching behavior over a broad range of state-of-the-art Pb-free PSCs is still lacking within the community.
the Pb leaching rate) could be a potential strategy toward reduction of Pb leaching, in addition to the existing effort in the development of Pb-free materials and device packaging and encapsulation approaches.

2. Results and Discussion

Water molecules can solvate (surface hydration with MAI or PbI) and infiltrate to interact with the Pb atoms, resulting in the decomposition of the PSC crystal structure.[42,43] To better understand the kinetic Pb leaching behavior of PSCs, both static immersion and dynamic spraying experiments were performed. The experimental design of Pb leaching evaluation is illustrated in Figure 1, covering the experimental steps of: 1) PSC fabrication; 2) static soaking in acid and dynamic rainfall exposure as driving forces for Pb leaching; 3) pretreatment for dissolution of precipitates; and 4) quantitative Pb leaching analysis with inductively coupled plasma mass spectrometry (ICP–MS). The size of all PSC samples used for our study was 2.5 × 2.5 cm (length × width). To allow for a straightforward analysis of the intrinsic Pb leaching behavior, we focus on analysis of PSCs without encapsulation, which is relevant to more extreme environmental scenarios where the device encapsulation fails or the devices are subjected to catastrophic physical damage.

We selected five types of high-performance PSCs with similar device architectures, namely, MAPbI3, FAMA, FAMACs, CsPbI3, and CsPbI2Br, which are representatives of state-of-the-art organic–inorganic hybrid and inorganic PSCs.[44–48] These benchmark high-performance PSCs were fabricated following recipes reported in the literatures,[44–48] with their device parameters summarized in Table 1. J–V characteristics of the five types of PSCs under AM 1.5G (100 mW cm−2) are shown in Figure S1, Supporting Information. The corresponding surface scanning electron microscope (SEM) images and X-ray diffraction patterns of the perovskite films are also shown in Figure S2 and S4, Supporting Information, respectively.

Since the processing method and antisolvents used are similar among all types of PSCs, the different amount of Pb measured is most likely attributed to the different thickness of the perovskite films, which has been optimized to obtain the best device efficiency for each kind of PSC. As shown in the cross-sectional SEM images in Figure S3, Supporting Information, the thickness of MAPbI3, CsPbI3, and CsPbI2Br-based PSCs is in the range of 420 ± 30 nm, while the thickness of FA0.95MA0.05Pb(I0.95Br0.05) and Cs0.05(FA0.85MA0.15)0.95Pb(I0.85Br0.15)3-based PSCs is in the range of 900 ± 30 nm. The total Pb amount measured is therefore closely correlated to the thickness of the perovskite layers. Beyond film thickness, other factors such as the density of the perovskite films and residual PbI2 during perovskite

![Figure 1. Schematic experimental procedures of Pb leaching evaluation.](image-url)

| PSCs       | Device structure                          | Jsc [mA cm−2] | Voc [V] | FF  | PCE [%] | PCE [%] in refs |
|------------|------------------------------------------|---------------|--------|-----|---------|-----------------|
| MAPbI5     | FTO/TiO2/MAPbI5/Spiro-OMeTAD/Ag          | 22.69         | 1.10   | 0.77| 19.22   | 17.80[44]       |
| FAMA       | FTO/TiO2/FA0.95MA0.05Pb(I0.95Br0.05)/Spiro-OMeTAD/Ag | 22.90 | 1.09 | 0.76| 18.97   | 19.18[45]       |
| FAMACs     | FTO/SnO2/(Cs0.05(FA0.85MA0.15)0.95Pb(I0.85Br0.15)3)/Spiro-OMeTAD/Ag | 22.69 | 1.09 | 0.78| 19.30   | 19.00[46]       |
| CsPbI3     | FTO/TiO2/CsPbI3/Spiro-OMeTAD/Ag          | 22.87         | 1.09   | 0.76| 18.94   | 17.06[47]       |
| CsPbI2Br   | FTO/TiO2/CsPbI2Br/Spiro-OMeTAD/Ag        | 16.33         | 1.14   | 0.80| 14.89   | 13.23[48]       |

Figure 1. Schematic experimental procedures of Pb leaching evaluation.
formation might also play a role in the different amount of Pb measured.

It is reported that PbI$_2$ or PbBr$_2$ is among the major decomposition product of perovskite materials, which only have modest solubility in pure water.$^{[42,49]}$ For accurate determination of the total amount of Pb leached, different concentrations of sodium thiosulfate (0.05, 0.1, 0.2, 0.5, 50 mL) and nitric acid (0.3%, 1%, 3%, 50 mL) were selected as eluents to transform any precipitates into soluble ions, as shown in Figure S5, Supporting Information. As shown in Figure S6a,b, Supporting Information, at least 0.1 mol L$^{-1}$ sodium thiosulfate is required to obtain a dissolution rate over 90%, while the high sodium thiosulfate concentration may lead to pipe blocking and instrument corrosion. In comparison, 0.3% nitric acid is sufficient to dissolve all precipitates (Figure S6b, Supporting Information), which was therefore selected throughout the study for the pretreatment of all extracted solutions.

Static soaking tests were first conducted to determine the total amount of Pb contained in each type of PSC. Unencapsulated PSC samples were submerged in 30 mL of 0.3% nitric acid (by volume) for 1, 5, or 10 days. The Pb leaching amount per unit area $L_{\text{total}}$ is given by

$$L_{\text{total}} = \frac{c}{V \times S_{\text{PSC}}}$$  \hspace{1cm} (1)

where $c$ denotes the Pb density measured by ICP–MS, $V$ denotes the volume of the pretreated solution, and $S_{\text{PSC}}$ denotes the effective area of PSCs. The total amount of Pb contained in the PSCs is determined by the accumulated Pb leaching amount after 10 days of soaking. PSCs were found to lose all the color, indicating that all the precipitates were fully dissolved. After 10 days of static soaking, the total Pb leaching amount of MAPbI$_3$, FAMA-, FAMACs-, CsPbI$_3$, and CsPbI$_3$Br-based PSCs was calculated to be 438.26, 917.50, 943.45, 439.77, and 493.54 mg m$^{-2}$, respectively (Figure 2). The Pb leaching amount after different days of static soaking is divided by the total Pb amount to obtain the quantitative leaching percentage. CsPbI$_3$-based PSCs leached 99.1% of total amount of Pb after 1 day of soaking, further increasing to 99.7% after 5 days of soaking, the fastest in the percentage of Pb leaching. In comparison, FAMA-based PSCs leached 85.5% of total amount of Pb after 1 day of soaking, increasing to 98.5% after 5 days of soaking, the slowest in the percentage of Pb leaching. MAPbI$_3$, FAMACs-, and CsPbI$_3$Br-based PSCs leached 96.3%, 97.4%, and 93.4% of total Pb amount after 1 day of soaking, increasing to 98.9%, 99.7%, and 99.5% after 5 days of soaking, respectively. It is obvious that all types of PSCs suffer from almost complete leaching of Pb within the first day of static soaking (see further details in Table S1, Supporting Information).

We further extend our studies to Pb leaching analysis under dynamic laboratory rainfall conditions, which is relevant to real-life conditions. The amount of Pb leached from unencapsulated PSCs under rainfall (freshwater) exposure is shown in Figure 3a. The initial Pb leaching dynamics can be revealed by varying the duration of rainfall spray at 10, 30, 60, 120, 240, 480, and 960 s. The absolute amount of Pb leaching after 960 s of spray time is strongly correlated to the total amount of Pb in the PSCs, with CsPbI$_3$-based PSCs exhibiting the least amount of Pb leaching and FAMA-based PSCs exhibiting the most. This result suggests that liquid water can easily infiltrate throughout the whole perovskite film, causing strong tendency for Pb leaching upon aqueous exposure across all types of PSCs investigated, thereby resulting in a strong dependence of the Pb leaching rate upon the total amount of Pb contained in the PSCs. The leakage ratio/percentage can be calculated with $L_{\text{total}}/L_{\text{leach}}$, where $L_{\text{total}}$ is the total amount of Pb in the PSCs determined in the static soaking test. As shown in Figure 3b, CsPbI$_3$Br-based PSCs leached 97.3% of total amount of Pb after 960 s of rainfall exposure, compared with 91.2%, 91.8%, 60.4%, and 84.9% for MAPbI$_3$-, FAMA-, FAMACs-, and CsPbI$_3$-based PSCs, respectively. Surprisingly, over 60% of total Pb amount was leached within the first 120 s of rainfall exposure for all types of PSCs investigated, demonstrating that the leaching of Pb in aqueous environments can be extremely fast, with most of the Pb in the PSCs leached from unencapsulated devices within a timescale from seconds to minutes.

However, it should be noted that the Pb leaching rate (especially within the first 120s of rainfall exposure) appears to be strongly dependent on the types of PSCs studied, as shown in Figure 3b’s inset. It is thus plausible that the different material compositions, crystal structures, film quality, and/or crystal stability of PSCs could have a strong impact upon Pb leaching. First, the moisture stability of the perovskite crystal structure appears to play a major role in the initial Pb leaching rate of PSCs. It has been reported that while all-inorganic CsPbI$_3$ and CsPbI$_3$Br-based PSCs are considered to have greater thermal stability than organic–inorganic PSCs, they are known to be more sensitive to moisture, which accelerates the phase degradation of α-phase to the nonperovskite δ-phase.$^{[50,51]}$ As shown in Figure 3b, it’s apparent that CsPbI$_3$ and CsPbI$_3$Br-based PSCs exhibit higher initial Pb leaching rates within the first 60 s than other types of PSCs. This result suggests that moisture-induced crystal decomposition may serve as the first step for Pb leaching upon aqueous exposure. It has been reported that the original perovskite film could rapidly decompose into solid PbI$_2$/PbBr$_2$ precipitates as well as water-soluble byproducts within seconds of exposure to water.$^{[49,52]}$ After 60 s, we observe a steady increase in the Pb leaching amount for FAMA- and FAMACs- based PSCs, which contain a higher total amount of Pb than the other types of PSCs. This results suggest that the total Pb amount contained in the perovskite film also plays a significant role, especially in the longer
It is therefore likely that the pH-dependent Pb leaching rate is attributed to the accelerated decomposition reaction of the perovskite materials in the presence of hydrogen ions and hydroxide ions.\textsuperscript{[53]}

We further investigate the effect of temperature upon the Pb leaching behavior. The temperature of the rainwater was varied at 10, 20, 30, and 40 °C, covering a temperature range commensurate to typical real-life conditions. As illustrated in Figure 4c,d, the amount of Pb leaching upon 60 s of rainfall exposure does not differ significantly over the temperature range, indicating general insensitivity of Pb leaching to temperature. The results above demonstrate that Pb, a major hazardous substance of PSCs, can easily leak out of PSCs under aqueous exposure. The leaked Pb can subsequently migrate within the environment and uptake by other living organisms, resulting in potentially significant environmental and ecotoxicological impacts. As a considerable amount of Pb was found to leach out of PSCs, their potential environmental risks were evaluated. If damaged PSCs are discarded, the increment amount of Pb leaching into the soil $L_i$ can be calculated with the equation\textsuperscript{[40]}

$$L_i = \frac{c \times S_{PSC}}{V_{soil} \times \rho_{soil}}$$

where $c$ denotes the amount of Pb leaching per unit area and $V_{soil}$ denotes the soil volume. The dry bulk density of soil is denoted as $\rho_{soil}$, and the average value is 1.3 g cm\textsuperscript{-3}.\textsuperscript{[57] If 1000 m\textsuperscript{2} PSC modules are constructed during a large-scale ground installation and 80% of soil area is covered, the effective size of PSCs, $S_{PSC}$, is assumed to be 800 m\textsuperscript{2}. It was asserted that the Pb species formed during perovskite decomposition by rain would strongly adsorb on soil particles, remaining concentrated in the top few centimeters, resulting in a greater Pb concentration near the soil surface.\textsuperscript{[16]} In this case, we assumed that Pb penetrated the soil 1 cm below the surface, since Pb tends to be highly absorbed by soil particles. The maximum Pb concentration in the polluted soil is calculated to rise by 43.5 mg kg\textsuperscript{-1}. The standard Pb concentration of uncontaminated soil ranges from 10 to 50 mg kg\textsuperscript{-1}, while urban soils often have much greater Pb concentrations, typically above 150 mg kg\textsuperscript{-1}.\textsuperscript{[58] While this level of leaching appears below the limit, we need to note that very little is known about the exact chemical and

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**Figure 3.** Pb leaching dynamics of the five different types of PSCs under pure water rainfall exposure. a) Pb leaching per unit area from 0 to 960 s. b) Pb leaching percentage at 0–960 s intervals and inset is 0–120 s intervals.
structural properties of the leachates as well as their solubility and environmental mobility, which could have a significant impact on their ecotoxicity. Moreover, Pb does not biodegrade or vanish over time, but rather persists for thousands of years in soils\(^{[54,58]}\). For indoor photovoltaic applications or wearable electronics, human exposure to Pb through skin contact or inhalation is highly likely. Therefore, it is necessary to recognize that other possible sources of Pb exposure exist to cumulate, such as home interiors or school or daycare playgrounds. Finally, it’s worth mentioning that Pb leaching of PSCs is on a similar level to solar cells based on PbS QDs, another class of emerging indoor PV technology\(^{[59,60]}\) (Figure S7, Supporting Information), suggesting that mitigating the environmental risks caused by Pb leaching can be a general critical challenge toward the commercialization of multiple technologies based on emerging Pb-based semiconductors.

3. Conclusion

We performed a quantitative analysis of the Pb leaching dynamics of five types of state-of-the-art PSCs. We found that significant Pb leaching could occur rapidly, with more than 60% of total amount of Pb found to leach from the PSCs within the first 120 s upon aqueous exposure. Our results suggest that the Pb leaching process of PSCs in aqueous environments could be a complicated process mediated by multiple competing factors, including in particular 1) the moisture stability of the perovskite crystal structure, with crystal decomposition serving as the first step for Pb leaching; 2) film quality, with grain boundaries and other defects facilitating water infiltration; and 3) the total Pb amount contained in the perovskite film contributing to Pb leaching in the longer term. Our results provide quantitative and kinetic insights for the environmental impact assessment, materials screening, and encapsulation and packaging of PSCs, highlighting the importance of incorporating effective mitigation methods to minimize Pb leaching toward their large-scale commercialization.

4. Experimental Section

Materials: Methylammonium iodide (MAI, 99.5%), Lead iodide (PbI\(_2\), 99.99%), Lead(II) Bromide (PbBr\(_2\), 99.99%), Cesium iodide (CsI, 99.9%), Spiro-OMeTAD (99.8%), 4-tert-butylypyridine (tBP, 99%), and lithium bis(trifluoromethanesulfonyl)imide (LiTFSI, 96%) were purchased from Xi’an Polymer Light Technology Corp. Methylammonium bromide (MABr, 99%) and formamidinium iodide (FAI, 99%) were purchased from Greatcell Solar Materials Pty Ltd. Titanium tetrachloride (TiCl\(_4\)), N,N-dimethylformamide (DMF, 99.9%), dimethyl sulfoxide (DMSO, 99.9%), isopropanol (IPA, 99.5%), chlorobenzene (CB, 99.8%), and acetonitrile (99.8%) were purchased from Sigma-Aldrich. 15% SnO\(_2\) colloidal solution was purchased from Alfa Aesar. The lead standard solution (1000 mg L\(^{-1}\)± 0.3%) was purchased from o2si smart solutions (Charleston, USA). Rhenium Standard Solution (50 μgL\(^{-1}\)) was obtained from Guobiao (Beijing) Testing & Certification Co., Ltd. (Beijing, China). Sodium thiosulfate was purchased from Tianjin Kernel Chemical Reagent Co., Ltd. (Tianjin, China). Nitric acid (68%) was produced from Guangzhou Chemical Reagent Factory (Guangzhou, China). Nylon filter
(0.45 μm) was purchased from Tianjin jinteng Experimental Equipment Co., Ltd. (Tianjin, China). All chemicals were used as received without further purification.

**Electron Transport Layer (ETL) Solution Preparation:** 3% SnO2 electron transport layer (ETL) colloidal solution was prepared by diluting 15% SnO2 colloidal solution in deionized water and then ultrasound for 10 min. Then, the SnO2 ETL solution was deposited on cleaned FTO-coated glass at 3000 rpm for 40 s and annealed at 150 °C for 30 min. Then, the SnO2 ETL solution was deposited on the cleaned FTO-coated glass at 3000 rpm for 40 s and annealed at 150 °C for 30 min. The TiO2 ETL was prepared by immersing the substrates in 200 mL of TiCl4 solution at 70 °C for 1 h.

**Perovskite Solution Preparation:** The MAPbI3 precursor solution was prepared by dissolving 461.01 mg PbI2, 223.00 mg FAI, 73.40 mg PbBr2, 77.00 mg MABr, and 19.48 mg CsI in a mixture of solvent of DMF and DMSO (4:1 v/v). The CsPbI2Br precursor solution was prepared by dissolving 138.00 mg PbBr2, 23.00 mg FAI, 73.40 mg PbBr2, 77.00 mg MABr, and 19.48 mg CsI in a mixture of solvent of DMF and DMSO (4:1 v/v). The CsPbBr3 precursor solution was prepared by dissolving 259.00 mg DMI, 461.00 mg PbI2, and 156.00 mg CsI in DMSO solvent. The CsPbBr3 precursor solution was prepared by dissolving 138.00 mg CsI and 98.00 mg PbBr2, and 57.00 mg CsBr in DMSO.

**Spiro-OMeTAD Solution Preparation:** After dissolving 90 mg Spiro-OMeTAD powder in 1 mL CB, 36 μL TBP and 22.5 μL Li-TFSI (520 mg mL\(^{-1}\) in acetonitrile) were added.

**Devices Fabrication:** The MAPbI3 precursor solution was spin coated on the TiO2 ETL for 40 s (4000 rpm, 2000 s\(^{-1}\)), treated with 200 μL CB by drop casting at 20 s, and then annealed at 100 °C for 10 min. Subsequently, the 100 μL spiro-OMeTAD solution was spin coated on the perovskite film for 30 s (5000 rpm, 2000 s\(^{-1}\)). After that, the 10 nm MoO3 (0.03–0.05 nm s\(^{-1}\)) and 100 nm Ag (0.2 nm s\(^{-1}\)) were deposited. The similar fabrication steps were employed by other PSCs except the following perovskite film preparation process.

The FA\(_{0.95}\)MA\(_{0.05}\)PbI\(_{3}\) precursor solution was spin coated on TiO2 ETL at 1000 rpm for 10 s and then 5000 rpm for 20 s. It was treated with 200 μL CB by drop casting at 20 s and then 100 °C annealed for 10 min.

The Cs\(_{0.05}\)(FA\(_{0.85}\)MA\(_{0.15}\))PbI\(_{3}\) precursor solution was spin coated on SnO2 ETL at 1000 rpm for 10 s (1000 rpm s\(^{-1}\)) at 6000 rpm for 30 s (3000 rpm s\(^{-1}\)) and treated with 200 μL CB at 25 s. It was then 100 °C annealed for 10 min.

The CsPbBr3 precursor solution was spin coated on 90 °C TiO2 ETL at 4000 rpm for 40 s, treated with 200 μL CB drop cast at 20 s, and then 210 °C annealed for 5 min.

The CsPbBr3 precursor solution was spin coated on 90 °C TiO2 ETL at 1000 rpm for 10 s (1000 rpm s\(^{-1}\)) and then 3000 rpm for 30 s (2000 rpm s\(^{-1}\)) before it was treated with 200 μL CB drop cast for 30 s and then 260 °C annealed for 10 min.

**Solution Pretreatment:** Nitric acid was added to the solutions after static soaking or dynamic spraying to achieve a 0.3% concentration. The mixed solution was left still for 24 h, and then sonicated for 2 h, and at last left still for 12 h. Furthermore, solutions were filtered with a 0.45 μm nylon filter and then diluted for the ICP–MS test.

**Spraying Test Setup:** Dynamic rain spraying was simulated with ultra-pure water delivered via a peristaltic pump (Longer WT600-2J). The device was attached to a bracket at a 30° angle to simulate the installation situation for solar cells.

**ICP–MS Test:** Pb concentration was measured by ICP–MS (NexION 300, Perkin Elmer), with rhodium (50 μg L\(^{-1}\)) as the internal standard. 10 mg L\(^{-1}\) Pb standard solution was prepared by diluting 1 mL 1000 mg L\(^{-1}\) Pb standard solution in a 100 mL volumetric flask with ultrapure water. Then, it was further diluted stepwise to 0, 1, 5, 10, 50, and 100, and 300 μg L\(^{-1}\) to get the standard curve before testing. During the ICP–MS test, KED mode was used, and the gas flow was 3.5 L min\(^{-1}\). All results reported were repeated with at least three different samples.

**Supporting Information**

Supporting Information is available from the Wiley Online Library or from the author.

**Acknowledgements**

D.Y. and X.L. contributed equally to this work. This research was supported by the National Natural Science Foundation of China (22106021 and 61704027), the Guangdong Basic and Applied Basic Research Foundation (2021A15151012372), Research Fund of Guangdong-Hong Kong-Macao Joint Laboratory for Intelligent Micro-Nano Optoelectronic Technology (no. 2020B1212030010), and Engineering and Physical Sciences Research Council (EPSRC, EP/V039771/1) and Royal Society of Chemistry (E21-9668828170).

**Conflict of Interest**

The authors declare no conflict of interest.

**Data Availability Statement**

The data that support the findings of this study are available from the corresponding author upon reasonable request.

**Keywords**

indoor photovoltaics, kinetic studies, lead leaching, perovskite solar cells, quantum dots

Received: April 14, 2022
Revised: June 3, 2022
Published online:

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