Highly selective conversion of CO\textsubscript{2} to methanol on the CuZnO–ZrO\textsubscript{2} solid solution with the assistance of plasma

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CuZnO–ZrO\textsubscript{2}–C was prepared by a co-precipitation method. For comparison, CuZnO–ZrO\textsubscript{2}–PC and CuZnO–ZrO\textsubscript{2}–CP were prepared by glow discharge plasma. The catalysts were characterized via the XRD, N\textsubscript{2} adsorption–desorption, TEM, SEM, EDX, XPS, CO\textsubscript{2}–TPD and H\textsubscript{2}–TPR techniques. The catalysts were comparatively investigated for CO\textsubscript{2} conversion and methanol selectivity in a fixed-bed reactor under the condition of 2 MPa, 250 °C, H\textsubscript{2}/CO\textsubscript{2} = 3/1 and GHSV = 12 000 mL g\textsuperscript{-1} h\textsuperscript{-1}. The results showed that the activities of the catalysts increased in the order of CuZnO–ZrO\textsubscript{2}–PC > CuZnO–ZrO\textsubscript{2}–CP > CuZnO–ZrO\textsubscript{2}–C. Moreover, the CO\textsubscript{2} conversion of CuZnO–ZrO\textsubscript{2}–C increased by 38.9% via treatment with glow discharge plasma. The results are well explained based on the CO\textsubscript{2}–TPD and H\textsubscript{2}–TPR characterizations of the catalysts.

1. Introduction

The excessive consumption of fossil fuels leads to the increase in atmospheric CO\textsubscript{2} concentration. CO\textsubscript{2} is one of the main greenhouse gases, which causes the global temperature to rise more seriously. Therefore, conversion of CO\textsubscript{2} into value-added chemicals has attracted extensive attention of the researchers.\textsuperscript{1–4} As an excellent C\textsubscript{1} resource, CO\textsubscript{2} is often used to prepare high value-added chemicals, such as methane, methanol, formic acid, and so on.

Methanol from CO\textsubscript{2} hydrogenation is one of the most important means of utilization of CO\textsubscript{2} resource. Methanol is an important raw material for the synthesis of organic chemicals and can be used to synthesize formaldehyde and other chemical products.\textsuperscript{5–7} Methanol is also a type of clean fuel, which can be blended into or substituted for gasoline as automobile fuels. Actually, the use of methanol has already been established for different applications such as in cars and fuel cells. In the 1960s, commercial Cu/ZnO/Al\textsubscript{2}O\textsubscript{3} catalysts were used in the production of industrial methanol at operating temperatures of 240–260 °C and reaction pressures of 5–10 MPa. However, at present, the selectivity of copper-based catalysts towards the methanol production from CO\textsubscript{2} hydrogenation is low. Therefore, it is urgent to either improve copper-based catalysts or develop new catalysts. Consequently, the research on the process route of methanol production from CO\textsubscript{2} hydrogenation is focused on the catalysts. Some studies have shown that conversion of CO\textsubscript{2} to methanol reaction is a structure-sensitive reaction. Besides, different preparation methods of catalysts lead to different structures of catalysts, thus affecting the activity of catalysts.\textsuperscript{8,9} The Cu/ZnO catalytic system is the basic copper-based catalyst, wherein zinc oxide can improve the dispersion and stability of copper, and lattice oxygen vacancies and paired electrons in zinc oxide can promote the methanol synthesis. Other catalysts, such as Cu/ZnO/SiO\textsubscript{2}, Cu/ZnO/ZrO\textsubscript{2}, Cu/ZnO/Al\textsubscript{2}O\textsubscript{3} and Cu/ZnO/TiO\textsubscript{2}, are generally modified on this basis. ZrO\textsubscript{2} is considered as a good carrier of Cu/ZnO as it provides active sites for CO and CO\textsubscript{2} for the insertion into Zr–OH, and Cu provides active hydrogen for ZrO\textsubscript{2} to hydrogenate formate and hydroxycarbonate to methanol on the ZrO\textsubscript{2} surface. Simultaneously, the addition of ZrO\textsubscript{2} can also improve the dispersion of active component copper. Moreover, ZrO\textsubscript{2} has low hydrophilicity, which can promote the catalytic activity.

Low temperature plasma is also widely used in the field of catalysis, including for catalysts used to synthesize ultrafine particles, surface treatment of catalysts and regeneration of catalysts. Catalysts prepared by a low temperature plasma-assisted synthesis generally have the advantages such as high dispersion, small catalyst particle size, large specific surface area and lattice defects. Bagherzadeh\textsuperscript{10} et al. prepared CuO/ZnO/Al\textsubscript{2}O\textsubscript{3} nanocatalysts by a hydrothermal synthesis and coprecipitation method, and then exposed to glow discharge plasma for 45 min at 1000 V. The results showed that the methanol conversion of CuO/ZnO/Al\textsubscript{2}O\textsubscript{3} catalysts synthesized by plasma-assisted coprecipitation was the highest. Seyedi\textsuperscript{11} et al. synthesized a series of CuO–ZnO–Al\textsubscript{2}O\textsubscript{3}–ZrO\textsubscript{2}...
nanocatalysts via hydrothermal, thermo-chemical and coprecipitation-plasma methods. The results showed that the catalysts treated by argon glow discharge plasma had a smaller particle size, better dispersibility and larger specific surface area, which could improve the hydrogen yield and methanol conversion.

In this study, CuZnO–ZrO2 catalysts were prepared via the plasma-assisted coprecipitation method. The catalysts were characterized and analyzed via the XRD, Nitrogen physical adsorption, TEM, XPS and H2–TPR techniques. The methanol synthesis performance of the catalysts before and after the glow discharge plasma modification was studied. The effects of the reaction temperature and stability of the catalyst were also investigated.

2. Experimental section

2.1 Preparation of catalysts

A mixed solution of metal nitrate was prepared by completely dissolving Zn(NO3)2·6H2O, Zr(NO3)4·5H2O and Cu(NO3)2·3H2O (Cu/Zn/Zr = 1/4/15) in deionized water. The solution with metal nitrates and 0.5 mol L−1 Na2CO3 solution was added into a reaction vessel at 70 °C. The final pH of titration was about 7.5. After the titration, the solution was stirred at 70 °C for 1 h, followed by ageing for 3 h. After cooling to room temperature, the solution was centrifuged and the precipitate was washed three times with deionized water to remove Na+ ions. The precipitate was dried in a vacuum drying chamber for 12 h to obtain the catalyst precursor.

The sample was then calcined in muffle furnace at 500 °C for 3 h and was denoted as CuZnO–ZrO2–C. This sample of the catalyst precursor was treated with a glow discharge plasma equipment for 15 min and was denoted as CuZnO–ZrO2–PC. Similarly, the sample of the catalyst precursor calcined in the muffle furnace at 500 °C for 3 h and then treated by plasma for 15 min and was denoted as CuZnO–ZrO2–CP.

2.2 Characterization techniques

The XRD patterns of the catalysts were obtained on a PANalytical X’Pert Pro X-ray diffractometer at 40 kV and 40 mA over 20° to 80° 2θ range with Cu Kα radiation (λ = 0.15416 nm). N2 adsorption–desorption isotherms for the catalysts were recorded on a Beckman coulter SA 3100 type equipment for 15 min and roasted for 3 h in a muffle furnace at 500 °C and was denoted as CuZnO–ZrO2–PC. Similarly, the sample of the catalyst precursor calcined in the muffle furnace at 500 °C for 3 h and then treated by plasma for 15 min and was denoted as CuZnO–ZrO2–CP.

The basicity of the catalyst was measured via temperature-programmed desorption of CO2 (CO2-TPD) techniques on an AutoChem II 2910 instrument (Micromeritics, USA). The reduction behavior of the catalysts was studied via hydrogen programmed temperature-reduction (H2–TPR) using a Micromeritics AutoChemII 2920 instrument equipped with a thermal conductivity detector (TCD).

2.3 Photocatalytic conversion of CO2

The photocatalytic conversion of CO2 to methanol was performed in an HP-WF51 fixed bed reactor (the inner diameter of a stainless steel reaction tube was 10 mm). Typically, 0.5 g of each catalyst (40–80 mesh) was pretreated at a pressure of 0.2 MPa in a flow of 50% H2/N2 (20 mL min⁻¹) at 280 °C for 4 h. Then, syngas (V(H2) : V(CO2) : V(N2) = 72 : 24 : 4) was applied, and the reaction was begun with the desired reaction parameters.

CO2, N2 and CO were analyzed via gas chromatography (6820, Agilent) with a thermal conductivity detector (TCD) equipped with a Porapak Q column. Methanol and other hydrocarbons were analyzed using a flame ionization detector (FID) equipped with a Porapak Q column.

Conversion of CO2, product selectivity and yield of methanol were calculated by the following eqn (1)–(4).

$$X(\text{CO}_2) = \frac{f(\text{CO}_2 \text{CON} + f(\text{CH}_3\text{OH} \text{ACH}_2\text{OH})/f(\text{CO}_2 \text{CON} + f(\text{CH}_3\text{OH} \text{ACH}_2\text{OH} + f(\text{CO}_2 \text{CON}))}{(1)}$$

$$S(\text{CH}_3\text{OH}) = \frac{f(\text{CH}_3\text{OH} \text{ACH}_2\text{OH})/f(\text{CH}_3\text{OH} \text{ACH}_2\text{OH} + f(\text{CO}_2 \text{CON}))}{(2)}$$

$$S(\text{CO}) = \frac{f(\text{CO}_2 \text{CON} + f(\text{CH}_3\text{OH} \text{ACH}_2\text{OH})}{(3)}$$

$$Y(\text{CH}_3\text{OH}) = \frac{X(\text{CO}_2) \times S(\text{CH}_3\text{OH})}{(4)}$$

$A_i$ represents the peak area of various products in the GC analysis and $f_i$ represents their correction factors.

3. Results and discussion

3.1 Catalyst characterization

3.1.1 XRD analysis. XRD was used to analyze the crystalline structure of the catalysts, and the XRD patterns of CuZnO–ZrO2–C, CuZnO–ZrO2–PC and CuZnO–ZrO2–CP catalysts are shown in Fig. 1. It could be seen that the catalysts exhibited four characteristic diffraction peaks at 2θ = 30.5°, 35.2°, 50.9°, and 60.3°, which are assigned to the (011), (110), (112) and (121) planes of tetragonal phase t-ZrO2. In addition, the typical diffraction peaks of monoclinic CuO are observed only in the CuZnO–ZrO2–C catalyst, which indicated that copper species in the CuZnO–ZrO2–C catalyst was CuO. No obvious diffraction peaks of ZrO2 during calcination, Zn2+ ions might have entered the lattice of zirconia and stabilized tetragonal zirconia.

As shown, the diffraction peaks of CuO disappeared after the plasma treatment (CuZnO–ZrO2–PC and CuZnO–ZrO2–CP). It might be that CuO lost its crystallinity and become amorphous, and highly dispersed after the plasma treatment. Three forms of
copper species were observed in solid solution: lattice doped, surface dispersed and bulk. The highly dispersive CuO on the catalyst surface is considered to be the most active form compared to the bulk CuO and lattice CuO. The intensity of the diffraction peaks of t-ZrO$_2$ decreased after the plasma treatment, indicating that the crystallinity of catalysts after the plasma treatment decreased compared with that of the catalyst directly roasted, and the catalyst modified by plasma had good dispersion.

3.1.2 **N$_2$ adsorption–desorption analysis.** The N$_2$ adsorption–desorption isotherms of the catalysts are shown in Fig. 2. Regardless of the treatment method used, each of the catalyst showed a type IV isotherm, together with an H3 hysteresis loop based on the IUPAC classification, which was characterized as a mesoporous structure. The summary of N$_2$ physisorption data of catalysts are reported in Table 1. The specific surface area of CuZnO–ZrO$_2$–C, CuZnO–ZrO$_2$–PC and CuZnO–ZrO$_2$–CP catalysts as calculated by BET equation was 24.48 m$^2$ g$^{-1}$, 24.57 m$^2$ g$^{-1}$ and 25.38 m$^2$ g$^{-1}$, respectively. The results also showed that the specific surface area of the catalyst treated with plasma did not change. The pore size distributions (PSDs) determined by the BJH method based on the desorption branch are shown in Fig. 2. A peak centered at 4 nm was observed for CuZnO–ZrO$_2$–C. CuZnO–ZrO$_2$–CP showed the same peak maximum at about 4 nm. The peak maxima of the PSDs for CuZnO–ZrO$_2$–PC clearly shifted to smaller size, suggesting a higher homogeneous dispersion of metal oxides. According to the pore size distribution curve, the radius corresponding to the highest point was about 2–5 nm, which indicated that the pore diameter distribution of catalysts was mainly mesoporous.

3.1.3 **Structural morphology analysis.** To investigate the structural morphology of the catalysts, TEM and HRTEM images were recorded, as shown in Fig. 3. It can be seen that the CuZnO–ZrO$_2$–C catalyst had no obvious edge (Fig. 3A), and the agglomeration of catalyst particles was obvious. However, the particle distribution of the catalysts (CuZnO–ZrO$_2$–PC and CuZnO–ZrO$_2$–CP) treated by plasma was more uniform, and the agglomeration phenomenon was obviously improved, particularly in the case of plasma modification before calcination. The HRTEM images revealed that the lattice fringes with the d-spacing of 0.293 nm corresponded to the (011) lattice plane of t-ZrO$_2$. The SEM image of CuZnO–ZrO$_2$–PC was is shown in Fig. 4. Each particle arranged tightly to form a granular layer and Zn$^{2+}$ ions completely entered the lattice of zirconia. It can be seen that CuO might exist in the amorphous form and is highly dispersed, which was consistent with the results of XRD.

| Catalyst            | Specific surface area (m$^2$ g$^{-1}$) | Pore volume (cm$^3$ g$^{-1}$) | Mean pore size (nm) |
|---------------------|---------------------------------------|-------------------------------|---------------------|
| CuZnO–ZrO$_2$–C     | 24.48                                 | 0.055                         | 9.05                |
| CuZnO–ZrO$_2$–PC    | 24.57                                 | 0.035                         | 5.98                |
| CuZnO–ZrO$_2$–CP    | 25.38                                 | 0.056                         | 8.90                |

![Fig. 1 XRD patterns of catalyst before and after modification by plasma.](image1)

![Fig. 2 N$_2$ adsorption–desorption and pore diameter distribution isotherms of the catalysts.](image2)
Fig. 4B showed that the distribution of the Cu, Zn, Zr, and O elements was uniform throughout the sample.

3.1.4 XPS analysis. To further investigate the electronic structures and chemical environment of Zn, Zr, Cu, O, XPS spectra were recorded, as shown in Fig. 5. Fig. 5A revealed a wide-range XPS spectra of CuZnO–ZrO$_2$–C, CuZnO–ZrO$_2$–PC, CuZnO–ZrO$_2$–CP catalysts. It was clear that catalysts exhibited Zn, Zr, Cu, O, C and Na elements. Amongst them, C 1s was the external pollution peak, which might have been due to the hydrocarbon present in the conductive adhesive on the instrument used during XPS analysis. The Na 1s peak appeared because of Na$_2$CO$_3$ solution as precipitating agent during Fig. 3 TEM images (A) of CuZnO–ZrO$_2$–C, (B) of CuZnO–ZrO$_2$–PC, (C) of CuZnO–ZrO$_2$–CP. HRTEM images (D) of CuZnO–ZrO$_2$–C, (E) of CuZnO–ZrO$_2$–PC, (F) of CuZnO–ZrO$_2$–CP.

Fig. 4 SEM images of CuZnO–ZrO$_2$–PC (A) and distribution diagram of Cu, Zn, Zr, O elements (B).
preparation of catalyst and the incomplete washing up of Na⁺ during precipitation washing. The high resolution XPS spectra of Zn 2p (Fig. 5B) exhibited two peaks at binding energies of 1021 and 1044 eV, corresponding to Zn 2p₃/2 and Zn 2p₁/2, respectively, which proved that Zn element existed in the form of Zn²⁺ in the catalyst.

The binding energy of Zn 2p of the plasma-treated catalyst decreased by about 4.5 eV, which indicated that the density of the electron cloud around Zn 2p changed significantly and the chemical environment was altered after the plasma treatment of the catalyst. The high resolution XPS spectra of Zr 3d (Fig. 5C) revealed the presence of two peaks located at the binding energies of 182 and 184 eV, corresponding to Zr 3d₅/2 and Zr 3d₃/2, respectively. Compared with the untreated catalyst, the binding energy of Zr 3d basically did not change, indicating that the chemical environment and energy state around Zr 3d of the plasma-treated catalyst did not change. The peak position of Zn 2p and Cu 2p moved to a lower binding energy level, which is responsible for improving the reduction capacity of the catalyst.

3.1.5 CO₂-TPD analysis. Fig. 6 displays the CO₂-TPD profiles of the catalysts, the CO₂ desorption curve of CuZnO-ZrO₂-C and CuZnO-ZrO₂-CP catalysts showed two strong peaks. Peaks located at the low temperature are ascribed to weakly adsorbed carbon species at weakly basic sites. These weakly basic sites are related to the hydroxyl group (Zr-OH). Peaks located at the high temperature corresponded to the strongly basic sites, such as coordination oxygen atoms. The CuZnO-ZrO₂-PC catalyst showed another peak at 294.3 °C, corresponding to the moderately basic sites, which is related to

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**Fig. 5** XPS survey spectra (A), high resolution XPS spectra of Zn 2p (B), high resolution XPS spectra of Zr 3d (C), high resolution XPS spectra of Cu 2p (D).
the metal–oxygen pairs.\textsuperscript{22} Considering that the reaction temperature was at 210–290 °C, moderately basic sites seemed to be more correlated to the active sites for CO\textsubscript{2} hydrogenation compared to the other basic sites. Therefore, it was speculated that CuZnO–ZrO\textsubscript{2}–PC would exhibit the best CO\textsubscript{2} conversion and CH\textsubscript{3}OH selectivity, which was consistent with the results of catalytic performance.

3.1.6 H\textsubscript{2}-TPR analysis. H\textsubscript{2}-TPR could be used to analyze the reduction performance of metal oxide catalysts. The H\textsubscript{2}-TPR profiles of the catalysts are shown in Fig. 7. The CuZnO–ZrO\textsubscript{2}–C catalyst showed a hydrogen consumption peak at 310.0 °C, which is attributed to the reduction of CuO species.\textsuperscript{23} The reduction temperature of CuO on the CuZnO–ZrO\textsubscript{2}–PC and CuZnO–ZrO\textsubscript{2}–CP catalysts were both lower than that of the CuZnO–ZrO\textsubscript{2}–C catalyst, particularly the CuZnO–ZrO\textsubscript{2}–PC catalyst, which could be due to the strong interaction between metal oxides in the catalyst.\textsuperscript{24} The results show that the plasma-modified catalyst enhanced the catalytic performance and improved the catalytic activity of the catalyst. Considering that the reduction temperature was 280 °C, CuO on the CuZnO–ZrO\textsubscript{2}–PC catalyst is speculated to completely reduce to Cu\textsuperscript{0}. Hence, CuZnO–ZrO\textsubscript{2}–PC would exhibit the best CO\textsubscript{2} conversion and CH\textsubscript{3}OH selectivity, which was consistent with the results of catalytic performance.

3.2 Catalytic performance

3.2.1 Effect of plasma modification on catalytic performance. The performance of CuZnO–ZrO\textsubscript{2}–C, CuZnO–ZrO\textsubscript{2}–PC and CuZnO–ZrO\textsubscript{2}–CP catalysts at 2 MPa, 250 °C, H\textsubscript{2}/CO\textsubscript{2} = 3/1 and GHSV = 12 000 mL g\textsuperscript{-1} h\textsuperscript{-1} are shown in Table 2. Compared with the CuZnO–ZrO\textsubscript{2}–C catalyst, the CO\textsubscript{2} conversion rate of the plasma-treated catalysts significantly increased; the CO\textsubscript{2} conversion of CuZnO–ZrO\textsubscript{2}–PC and CuZnO–ZrO\textsubscript{2}–CP catalysts increased by 39.8% and 33.9%, respectively, while methanol selectivity changes were not obvious. The results showed that the CO\textsubscript{2} conversion over the plasma-modified catalyst significantly increased, and the catalytic activity of the catalyst effectively improved, but the selectivity of methanol did not significantly improve.

3.2.2 Effect of the reaction temperature on catalytic performance. The effect of reaction temperature (the range of 210–290 °C) on catalytic performance is shown in Fig. 8. As shown in Fig. 8A, with the increase in the reaction temperature, the conversion of CO\textsubscript{2} increased, accompanied by a decrease in methanol selectivity, which was the result of a competition between the methanol synthesis and reverse water gas reaction. The methanol synthesis reaction is an exothermic reaction, while reverse water gas reaction is an endothermic reaction.\textsuperscript{25} As shown in Fig. 8B, the methanol yield of all catalysts first increased with the increase in temperature; however, after reaching a certain value, the methanol yield decreased with the rise in temperature. At the maximum methanol yield of the CuZnO–ZrO\textsubscript{2}–C catalyst, the corresponding temperature was about 250 °C, while when the methanol yield of CuZnO–ZrO\textsubscript{2}–PC catalyst was the maximum, the corresponding temperature was about 270 °C, which indicated that the catalyst treated by plasma still had a higher catalytic activity at higher temperature values. This suggests that CuZnO–ZrO\textsubscript{2}–PC had a better high-temperature resistance. Obviously, the catalytic effect of CuZnO–ZrO\textsubscript{2}–PC and CuZnO–ZrO\textsubscript{2}–CP catalysts was better than that of the CuZnO–ZrO\textsubscript{2}–C catalyst.

3.2.3 Stability test of catalysts. In order to study the influence of plasma modification on the stability of the catalyst, the CuZnO–ZrO\textsubscript{2}–C and CuZnO–ZrO\textsubscript{2}–PC catalysts were tested in a fixed bed reactor under the following condition of P = 2 MPa, T = 250 °C, H\textsubscript{2}/CO\textsubscript{2} = 3/1 and GHSV = 12 000 mL g\textsuperscript{-1} h\textsuperscript{-1}. As shown in Fig. 9, after the catalysts were used in a fixed bed

| Catalysts        | X(CO\textsubscript{2})| S(CH\textsubscript{3}OH)| S(CO)| Y(CH\textsubscript{3}OH) |
|------------------|----------------------|-----------------------|------|-----------------------|
| CuZnO–ZrO\textsubscript{2}–C | 10.3                 | 74.1                  | 25.9 | 7.6                   |
| CuZnO–ZrO\textsubscript{2}–PC  | 14.4                 | 78.2                  | 21.8 | 11.3                  |
| CuZnO–ZrO\textsubscript{2}–CP   | 13.8                 | 74.2                  | 25.8 | 10.2                  |

\textsuperscript{a} Condition: P = 2 MPa, T = 250 °C, H\textsubscript{2}/CO\textsubscript{2} = 3/1 and GHSV = 12 000 mL g\textsuperscript{-1} h\textsuperscript{-1}.
reactor for a long time of 100 h, the CO2 conversion of the CuZnO–ZrO2–C catalyst decreased by 7.8% (from 10.3% to 9.5%), and the methanol selectivity decreased by 5.8% (from 74.1% to 69.8%). However, the CO2 conversion of the CuZnO–ZrO2–PC catalyst decreased by 6.3% (from 14.4% to 13.5%), and the methanol selectivity decreased by 3.8% (from 78.2% to 75.2%).

4. Conclusions

The CuZnO–ZrO2 catalysts were pretreated by the glow discharge plasma before and after calcination. Compared with the CuZnO–ZrO2–C catalyst, the performance of the catalyst modified by plasma was better, the CO2 conversion of CuZnO–ZrO2–PC increased by 39.8%, and the methanol yield increased by 48.7%. The XRD characterization showed that the plasma-modified catalyst had a lower crystallinity and a better dispersion, and the characterization of CO2-TPD and H2-TPR provided a basis for the better catalytic performance of plasma-modified catalysts.

Conflicts of interest

There are no conflicts to declare.
9  J. Ye, C. Liu, D. Mei, et al., Active oxygen vacancy site for methanol synthesis from CO₂ hydrogenation on In₂O₃ (110): a DFT study, ACS Catal., 2013, 3(6), 1296–1306.
10  S. B. Bagherzadeh and M. Haghighi, Plasma-enhanced comparative hydrothermal and coprecipitation preparation of CuO/ZnO/Al₂O₃ nanocatalyst used in hydrogen production via methanol steam reforming, Energy Convers. Manage., 2017, 142, 452–465.
11  A. M. Seyedi, M. Haghighi and N. Rahemi, Significant influence of cutting-edge plasma technology on catalytic properties and performance of CuO–ZnO–Al₂O₃–ZrO₂ nanocatalyst used in methanol steam reforming for fuel cell grade hydrogen production, Ceram. Int., 2017, 43(8), 6201–6213.
12  J. Wang, G. Li, Z. Li, et al., A highly selective and stable ZnO–ZrO₂ solid solution catalyst for CO₂ hydrogenation to methanol, Sci. Adv., 2017, 3(10), e1701290.
13  K. Cheng, B. Gu, X. Liu, et al., Direct and highly selective conversion of synthesis gas into lower olefins: design of a bifunctional catalyst combining methanol synthesis and carbon–carbon coupling, Angew. Chem., 2016, 55(15), 4725–4728.
14  W. Shan, W. Shen and C. Li, Structural Characteristics and Redox Behaviors of Ce₁ₓCu₂O₃ Solid Solutions, Chem. Mater., 2003, 15(25), 4761–4767.
15  W. T. Yao, S. H. Yu, Y. Zhou, et al., Formation of uniform CuO nanorods by spontaneous aggregation: Selective synthesis of CuO, Cu₂O, and Cu nanoparticles by a solid–liquid phase arc discharge process, J. Phys. Chem. B, 2005, 109(29), 14011–14016.
16  J. Y. Luo, M. Meng, X. Li, et al., Mesoporous Co₃O₄–CeO₂ and Pd/Co₃O₄–CeO₂ catalysts: synthesis, characterization and mechanistic study of their catalytic properties for low-temperature CO oxidation, J. Catal., 2008, 254(2), 310–324.
17  Y. Yu, Y. M. Chan, Z. Bian, et al., Enhanced performance and selectivity of CO₂ methanation over g-C₃N₄ assisted synthesis of NiCeO₂ catalyst: kinetics and DRIFTS studies, Int. J. Hydrogen Energy, 2018, 43(32), 15191–15204.
18  Y. Yu, Z. Bian, F. Song, et al., Influence of Calcination Temperature on Activity and Selectivity of Ni–CeO₂ and Ni–Ce₀.₆Zr₀.₄O₂ Catalysts for CO₂ Methanation, Top. Catal., 2018, 61(15–17), 1514–1527.
19  J. Xiao, D. Mao, X. Guo, et al., Effect of TiO₂, ZrO₂, and TiO₂–ZrO₂ on the performance of CuO–ZnO catalyst for CO₂ hydrogenation to methanol, Appl. Surf. Sci., 2015, 338, 146–153.
20  S. T. Hossain, E. Azeeva, K. Zhang, et al., A comparative study of CO oxidation over Cu–O–Ce solid solutions and CuO/CeO₂ nanorods catalysts, Appl. Surf. Sci., 2018, 455, 132–143.
21  H. Zhu, R. Razzaq, C. Li, et al., Catalytic Methanation of Carbon Dioxide by Active Oxygen Material CeₓZr₁₋ₓO₂ Supported Ni–Co Bimetallic Nanocatalysts, AIChE J., 2013, 59, 2567–2576.
22  T. Witoon, N. Kachaban, W. Donphai, et al., Tuning of catalytic CO₂ hydrogenation by changing composition of CuO–ZnO–ZrO₂ catalysts, Energy Convers. Manage., 2016, 118, 21–31.
23  H. Lei, R. Nie, G. Wu and Z. Hou, Hydrogenation of CO₂ to CH₃OH over Cu/ZnO catalysts with different ZnO morphology, Fuel, 2015, 154, 161–166.
24  G. Wang, D. Mao, X. Guo, et al., Enhanced performance of the CuO–ZnO–ZrO₂ catalyst for CO₂ hydrogenation to methanol by WO₃ modification, Appl. Surf. Sci., 2018, 456, 403–409.
25  E. Moretti, L. Storaro, A. Talon, et al., Effect of thermal treatments on the catalytic behaviour in the CO preferential oxidation of a CuO–CeO₂–ZrO₂ catalyst with a flower-like morphology, Appl. Catal., B, 2011, 102(3–4), 627–637.