First Principles Studies on Boron Sites in Zeolites
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ABSTRACT:
A systematic computational investigation on protonated and nonprotonated boron-containing zeolites (boralites), performed by using different periodic density functional theory approximations, is presented. Both minimum energy structures and finite temperature behavior of model boron sodalites were analyzed. All of the adopted computational schemes agree in predicting an acid site composed of a silanol Si−OH group loosely linked to a planar BO3 structure in the protonated system and a BO4 tetrahedral site in the sodium-containing zeolite. Calculated structural and vibrational properties are in line with experimental data. Comparisons of the protonated boralite site with Al and Ga zeolitic acid sites are discussed as well. Results indicate that this class of mild acid catalysts is characterized by significant framework flexibility and pronounced thermal effects due to the loosely bound acid site.

INTRODUCTION

A widely adopted procedure to tailor physicochemical properties of zeolites relevant for catalysis is the isomorphous substitution of Si atoms with trivalent metal cations TIII.1 In particular, the acidity of protonated zeolites may be tuned by appropriately choosing the TIII cation replacing Si in the tetrahedral T site of the zeolitic framework. The acid strength decreases in the sequence Al > Ga > Fe ≈ B.2 This order was rationalized on the basis of the different chemical structures formed by acid protons in zeolites, namely, bridge acid sites Si−(OH)−TIII and silanol groups weakly interacting with the trivalent cation, Si−OH⋯TIII.3−5 Bridge sites are characterized by a three-coordinate protonated oxygen, resulting in very reactive protons. However, silanol sites are significantly weaker acids and may therefore be exploited in applications requiring mild conditions.3 In recent years, much attention has been devoted to the synthesis and characterization of borosilicate zeolites, also known as boralites,3 because their lower acidity makes them suitable catalysts for reactions such as double-bond isomerization of linear olefins,6 MTBE cracking,7 and Beckmann rearrangement of cyclohexanone oxime.8 Their ability to work as mild solid acid catalysts may also be of relevance in the context of the design of new industrial processes compatible with environmental and sustainable development issues. Acid boralites are also interesting for their change in the coordination geometry at the boron center, which occurs reversibly upon dehydration and rehydration.1 Specifically, both 11B solid-state NMR9−13 and IR studies14−22 have shown that B is characterized by a trigonal geometry in dehydrated samples, whereas upon rehydration a tetrahedral structure of the boron site is recovered. However, upon being washed with water, trigonal boron can be readily hydrolyzed to a defective B-site and then completely extracted from the framework (de-boronation process).1,11,12,18 The trigonal-to-tetrahedral coordination change is also observed in the presence of other bases such as pyridine and ammonia.23,24 Such a transformation, like in the case of water, is accompanied by a proton transfer from the SiOH2− group to the probe molecule. Structural information on B-sites in boralites has been obtained by single-crystal X-ray diffraction studies on as-synthesized B-ferrierite25 and B-MFI26,27 containing ethylenediamine divalent cations as extraframework species. Boron was reported to be in a tetrahedral geometry, statistically occupying different T sites. This picture has been essentially confirmed in B-MFI synthesized by using tetrapropyl ammonium as a templating molecule.28

Structure, acidity, and relative stability of B-sites have been the subject of several theoretical studies based on cluster models. Calculations on cluster models of acid B-ZSM5,29−31 B-mazzite,32,33 and B-MCM24 agree in predicting a trigonal structure for the B-site, with a silanol group formed on the adjacent Si atom. Properties such as proton affinities, stretching frequencies, and NMR chemical shifts have also been calculated on model sites, with an overall agreement with experimental trends.10,29,32−34 However, to the best of our knowledge, no computational study on the properties of the B-site in a zeolite crystal at conditions close to the operative ones has been performed to date. Moreover, establishing to what extent the insertion of B, characterized by an ionic radius much smaller than Si, could affect the framework structure at long range, for example, by deforming pore openings, may be extremely useful for synthesis and application of new boron-containing framework types. This relevant and still-open issue might be properly addressed by taking into account the periodicity of the zeolitic framework.

The unit cell content of sodalite [Si12O40] enables us to perform an extensive computational study on these systems. Beside gaining a thorough theoretical characterization of zeolitic B-sites, the aim of the present work is also to investigate the effect of the choice of the type and size of the basis set, density functional method, and trajectory sampling approach on the quality of the results.
In the first part of the study, different basis sets and density functional approximations were compared by calculating minimum energy structures of B-SOD models containing either H⁺ or Na⁺ as counterions. The effect of framework relaxation on the B acid site structure was also investigated by optimizing the geometry of protonated B-SOD crystals in which atoms beyond the third coordination shell were kept fixed.

In the second part of the work, the room-temperature behavior of B-sites in a zeolite framework was studied by means of the Car–Parrinello approach. Trajectories characterized by different simulation parameters were performed to systematically assess the reproducibility of results on B-zeolites. In the final part of the work, calculated physicochemical properties were compared with available experimental data on different B-zeolites and with results of previous computational studies on Al- and Ga-containing sodalites.

2. Computational Methods

Sodalite, which has unit cell stoichiometry of [Si12O24] in its all-silica form, is characterized by six-membered rings as maximum pore openings. Its unit cell is formed by two cuboctahedral β-cages. The β-cage is the building block of several larger zeolites and makes sodalite one of the choice model systems for carrying out accurate and extensive computational studies. Hydrothermal synthesis of B-SOD has been reported; however, cell parameters and unit cell content were not available. Models of the B-SOD crystal were built from the all-silica sodalite structure by replacing with a B atom one of the 12 equivalent Si atoms in the unit cell. The model crystal is characterized by a B molar fraction of 0.0833, which is of the order of those commonly found in standard borates. The unit cell contraction due to the smaller size of the B cation with respect to Si was taken into account by applying Vegard's law

\[ V_x = V_{Si} - V_{Si} \left(1 - \frac{d_B}{d_{Si}} \right)^3 \]  

where \( V_{Si} \) is the unit cell volume of the all-silica phase, \( d_B \) and \( d_{Si} \) are the tetrahedral B–O and Si–O bond distances, and \( x = \frac{[B]/[B]}{[Si]/[Si]} \) is the molar fraction of B. Through the use of a \( d_0/d_{Si} \) ratio of 0.9125 as in ref 25 and the unit cell of the all-silica sodalite (8.836 Å), a unit cell parameter of 8.777 Å was obtained for the B-SOD models.

Two systems were considered: an acid B-sodalite model, \( H^+\)-B-SOD, and its Na-substituted counterpart, \( Na^+\)-B-SOD, where the extraframework species balancing the negative charge of the framework were \( H^+ \) and \( Na^+ \), respectively. In both models, the starting configuration was built by assuming a tetrahedral geometry for the B-site. In \( H^+\)-B-SOD, the proton was placed at a bonding distance from one of the oxygen atoms bonded to B, while in \( Na^+\)-B-SOD the Na nucleus was placed close to the center of the six-membered ring containing B.

Geometry optimizations (GOs) and first principles molecular dynamics (FPMD) simulations were performed by using the plane waves (PWs) DFT code CPMD. Wave functions were expanded in plane waves (Γ point only), and norm-conserving pseudopotentials were adopted for the ionic core–valence electron interactions. Nonlocality up to \( J = 2 \) was adopted for all of the atomic species with the exception of H, for which a local pseudopotential was used. The pseudopotentials, in numerical form, were obtained via the Troullier and Martins approach. GOs on \( Na^+\)-B-SOD and \( H^+\)-B-SOD were carried out by using quasi-Newton methods and adopting several gradient-corrected density functional approximations. In particular, the following functionals or combinations of functionals were employed: Becke–Perdew (BP), Becke–Lee–Yang–Parr (BLYP), generalized gradient approximation (GGA), Perdew–Burke–Emzerhof (PBE), and Hamprecht–Cohen–Tozer–Handy (HCTH/120). Convergence of the results with respect to the kinetic energy cutoff in the PW expansion was also checked.

The performance of PW versus localized basis sets was investigated as well. Localized basis sets were used in the QUICKSTEP DFT-based approach. QUICKSTEP (QS) is an electronic structure program, available in the CP2K code, which computes energy and forces through the Gaussian plane wave (GPW) method, a hybrid scheme of Gaussian and plane wave functions. In particular, the Kohn–Sham orbitals are expanded in a linear combination of atom-centered Gaussians while the electron density is expanded in a PW basis set. In analogy with PW-based methods, in the GPW approach only valence electrons are accounted for, and the atomic core–valence electron interactions are represented by pseudopotentials. A detailed review of the method can be found in ref 49. The PBE and pseudopotentials of the Goedecker–Teter–Hutter (GTH) type optimized for PBE were adopted in the GOs performed with the GPW method. The convergence of atomic basis set size was analyzed by comparing results obtained by using a DZVP and a TZV2P basis set. A 300 Ry cutoff for the PW expansion of the density was adopted. A tight convergence criterion upon the wave function was enforced during the optimization runs (convergence criterion for the electronic gradient: \( \epsilon \leq 10^{-7} \) a.u.). The convergence criterion for nuclear gradients in GOs was \( \delta \leq 1 \times 10^{-7} \) a.u. as in the PW calculations.

The effect of framework relaxation on the structure of the B acid site was studied by fixing the positions of the framework atoms beyond the third shell from the B–O(H)–Si oxygen in the \( H^+\)-B-SOD system to the all-silica sodalite X-ray positions. Relaxation effects were also investigated on three model sodalites \( H[T^III][Si_{11}O_{24}] \), with \( T^{III} = Al, Ga, \) or Fe. Norm-conserving pseudopotentials (d-nonnlocality) were used for these atoms as well, and in the case of Fe, a restricted open-shell formalism, with Fe^III in the high-spin state, was adopted.

Car–Parrinello (CP) FPMD simulations on \( Na^+\)-B-SOD and \( H^+\)-B-SOD were performed at room-temperature conditions by using the PBE functional with a 70 Ry cutoff (PBE/70). Simulations were carried out also with the BP functional at a 60 Ry cutoff (BP/60) to (i) evaluate whether and up to what extent the choice of DFT approximation/PW cutoff might affect average structural properties and (ii) allow comparison with results previously obtained by FPMD on \( H^+\)-Al-SOD and \( H^+\)-Ga-SOD with the BP/60 computational scheme. The validity of the BP/60 approach has been already proven in several examples dealing with many aspects of framework minerals and zeolite chemistry: structure, reactivity, and behavior at high pressure and temperature conditions. Here its accuracy is further tested by cross-checking the BP/60 results on \( H^+\)-Al-SOD with those obtained on the same system with PBE/70. The equations of motion were integrated by using a time step of 0.121 fs (5 a.u.) and adopting an inertia parameter of 500 a.u. for the electronic coefficients. Such values have been shown to ensure adiabatic separation of ionic and electronic degrees of freedom and have been adopted in all FPMD runs here presented. In the CP approach both ionic and electronic degrees of freedom are propagated by equations of motion derived from an extended Lagrangian in which an inertia parameter is associated to the wave
TABLE 1: First Principles Molecular Dynamics Simulations of H–B-SOD and Na–B-SOD

| initial configuration | t_{\text{equil}} | t_c | \omega_i | \omega_e |
|-----------------------|-----------------|-----|---------|---------|
| H–B-SOD               |                 |     |         |         |
| BP/60(1)              | 2.6             | 6.0 | 3200    |         |
| BP/60(2)              | 16.0            | 20.0| 3200    |         |
| BP/60(3)              | 16.0            | 18.0| 3200    |         |
| PBE/70(1)             | 2.7             | 7.3 | 3200    |         |
| PBE/70(2)             | 11.0            | 64.1| 3200    |         |
| PBE/70(3)             | 2.0             | 39.0| 3500    | 12 000  |
| BO/PBE/70             | 3.0             | 8.5 | 3200    |         |
| Na–B-SOD              |                 |     |         |         |
| BP/60                 | 5.0             | 13.0| 1200    |         |
| PBE/70                | 5.0             | 19.3| 1200    |         |

\(t_{\text{equil}}\) is the total equilibration time (in ps), \(t_c\) is the data collection time (in ps), and \(\omega_i\) and \(\omega_e\) are the characteristic frequencies of the N\(\ddot{o}\)s\(\ddot{e}\)--Hoover thermostats on the ions and on the electrons, respectively (in cm\(^{-1}\)). The N\(\ddot{o}\)s\(\ddot{e}\)--Hoover target ionic temperature and electronic kinetic energy were 300 K and 0.0035 a.u., respectively.

function coefficient motions, while in Born–Oppenheimer (BO) simulations the wave function is quenched at each time step to the BO surface corresponding to the instantaneous ionic configuration. The influence of the CP electronic inertia parameter was investigated by performing a PBE/70 BO simulation on the H–B-SOD system in the NVT ensemble.

Both BP/60 and PBE/70 simulations were performed in the NVT ensemble with N\(\ddot{o}\)s\(\ddot{e}\)--Hoover chain thermostats and a target ionic temperature of 300 K. Different starting conditions, equilibration times, and thermostat parameters were chosen (Table 1). Equilibrations were performed by alternating 0.1 ps bins in the NVE ensemble by 0.1 ps bins with velocity rescaling (target temperature rising from 100 to 300 K, with tolerance from 32 to 37 K). For some simulations, an additional equilibration period in the NVT ensemble at 300 K was carried out. For the H–B-SOD system, data were collected for periods longer than those typical of FPMD simulations to guarantee a converged description of the B-site properties at equilibrium conditions.

3. Results and Discussion

3.1. Optimized Structures of H–B-SOD and Na–B-SOD.

The boron site is found to be trigonal in H–B-SOD and tetrahedral in Na–B-SOD, independently of the adopted DFT functional and basis set type and size and in line with both available experimental data and results of previous calculations on model clusters. Relevant structural parameters are reported, for comparison, in Tables 2–5. Graphical representations of the optimized structures are shown in Figure 1.

In Na–B-SOD, the Na cation, which is coordinated to six oxygen atoms of the six-membered ring containing B, is found at a distance of 3.0 Å from B. The tetrahedral geometry of the B-site slightly deviates from ideality; in particular longer B–O distances are found for the two oxygen atoms coordinated to Na, namely, O* and O\(^\circ\) (see Figure 1 for atom labels).

The H–B-SOD system is characterized by a silanol group Si\(^*\)--O\(^*\)H and a trigonal B-site in a nearly planar BO\(_3\) geometry. These groups are loosely interacting, as indicated by the large separation between the silanol oxygen O* and the B atom (> 2.5 Å). The B–O bonds in the BO\(_3\) structure are shorter than those of the tetrahedral B-site in Na–B-SOD. Interatomic separations and B–O\(^*\)--Si\(^*\) angles are well-converged with respect to basis set size (Tables 2 and 3). In particular, the largest variation in the geometrical parameters of the B-site is found for the B–O\(^*\) nonbonding distance in H–B-SOD and amounts to 0.7% in passing from 60 to 90 Ry with BP and to 0.5% from 60 to 110 Ry with PBE. The corresponding B–O\(^*\)--Si\(^*\) angle variations amount to 1.6% and 0.6%, respectively. Cutoff-dependent changes affecting the geometry of the B-site in Na–B-SOD are even smaller.

The dependency of the geometry on the DFT approximation and on the type of basis set is reported in Tables 4 and 5. It can be noticed how all of the adopted functionals provide structural descriptions of the B-site remarkably in agreement among each other. On this basis, the effect of localized versus plane wave basis sets on the calculated structural parameters was studied with the PBE functional only.

For both H– and Na–B-SOD, the GPW approach provides results well in line with PW calculations, confirming a planar trigonal and a tetrahedral B-site geometry, respectively. It is also worth pointing out that the nonbonding distances calculated with the GPW method are very close to the PW ones, thus demonstrating reproducibility of calculated structural parameters between the two approaches. In particular, the GPW B–O\(^*\) and B–Si\(^*\) distances in H–B-SOD (Table 4) indicate that the large separation between B and the silanol group found in the model acid B-zeolite is not dependent on the locality of the basis set. Basis set size slightly affects geometrical parameters; in particular the GPW/TZV2P geometries are closer to those obtained with PW at 70 Ry.

3.2. Framework Relaxation in H–B-SOD.

The minimum energy structure in H–B-SOD is characterized by a trigonal boron site weakly interacting with a silanol group, in keeping with previous calculations performed on cluster models of proton-exchanged B-zeolites.\(^{29,31–34}\) The nonbonding B–O\(^*\) distance (2.57 Å with PBE/70) is however significantly larger than the calculated values reported in the literature, which range from 2.02 to 2.37 Å depending, beside the electronic structure treatment, on the size and extent of relaxation of the adopted cluster model. In this respect, it was shown that converged structural properties of zeolitic acid sites required cluster models built by positioning around O* at least 3–4 shells of atoms, the first two of which must be fully relaxed.\(^{29}\) However, even though framework relaxation appeared to greatly affect the optimized geometry, extrapolation of these observations to the limiting case of a fully relaxed periodic crystal structure has not been performed yet. To shed light on this issue and establish whether differences among predicted B–O\(^*\) distances should be attributed to geometrical constraints or to the calculation approach, GO results on the fully relaxed periodic H–B-SOD model were compared to those obtained on an identical model in which only the first three shells of atoms around O* were allowed to relax. The starting configuration for the constrained GOs was the same as that adopted in the full GOs on H–B-SOD, with the outer-shell atoms kept fixed to the SOD crystallographic positions. Results obtained with the PBE functional using both PW and GPW approaches are reported in Table 6. Comparison with data in Table 4 shows that the predicted B-site geometry is significantly altered by the imposed geometrical constraint, with a more pronounced effect in the case of localized basis sets. Both constrained optimizations led to structures characterized by significantly shorter B–O\(^*\) and B–Si\(^*\) separations and a larger Si\(^*\)--O\(^*\)--B angle with respect to the fully relaxed structure. Such a picture was also confirmed by constrained GOs performed with the other functionals. Interestingly, B–O\(^*\) distances (2.06 and 1.96 Å for PW and GPW, respectively) compare well with the values of 2.02 and
2.023 Å obtained recently from B3LYP/6-311+g(d,p) calculations on cluster models of B-MCM-22 and B-ZSM5 where a similar extent of framework atom relaxation was adopted. The effect of the applied geometrical constraints on the resulting minimum energy structure is pictured in Figure 2, where the partially frozen PBE/70-optimized geometry is...

**Figure 1.** Top panels: Graphical representations of the optimized structures of H–B-SOD (left panel) and Na–B-SOD. Bottom panels: Ball-and-stick representations of the boron site in H–B-SOD (left panel) and Na–B-SOD (right panel). Atom color code: Si, gray; O, red; B, yellow; H, white; Na, blue. The images refer to optimizations performed with PWs at the PBE/70 level.

**TABLE 2: Optimized Geometrical Parameters of the B-Site in H–B-SOD Calculated with the BP and PBE Functionals and PW Basis Sets as a Function of the PW Cutoff**

|       | B–O* | B–O1 | B–O2 | B–O# | O*–H | Si*–O* | B–Si* | B–O*–Si* |
|-------|------|------|------|------|------|--------|-------|----------|
| BP/60 | 2.556| 1.377| 1.373| 1.363| 0.976| 1.648  | 3.918 | 136.4    |
| BP/70 | 2.572| 1.372| 1.370| 1.357| 0.975| 1.639  | 3.953 | 138.6    |
| BP/90 | 2.576| 1.370| 1.370| 1.355| 0.971| 1.636  | 3.956 | 138.7    |
| PBE/60| 2.562| 1.377| 1.374| 1.364| 0.975| 1.649  | 3.923 | 136.3    |
| PBE/70| 2.574| 1.373| 1.371| 1.358| 0.974| 1.639  | 3.938 | 137.3    |
| PBE/90| 2.573| 1.370| 1.368| 1.355| 0.971| 1.637  | 3.937 | 137.4    |
| PBE/110| 2.577| 1.371| 1.368| 1.356| 0.970| 1.637  | 3.939 | 137.2    |

a Distances are given in angstroms, and angles are given in degrees.

**TABLE 3: Optimized Geometrical Parameters of the B-Site in Na–B-SOD Calculated with the BP and PBE Functionals and PW Basis Sets as a Function of the PW Cutoff**

|       | B–O* | B–O1 | B–O2 | B–O# | O*–Na | Si*–O* | B–Si* | B–O*–Si* |
|-------|------|------|------|------|-------|--------|-------|----------|
| BP/60 | 1.528| 1.454| 1.506| 1.448| 2.218 | 1.620  | 3.017 | 146.8    |
| BP/70 | 1.527| 1.451| 1.502| 1.446| 2.213 | 1.612  | 3.016 | 147.9    |
| BP/90 | 1.522| 1.447| 1.503| 1.443| 2.223 | 1.609  | 3.016 | 149.0    |
| PBE/60| 1.525| 1.453| 1.507| 1.450| 2.208 | 1.620  | 3.014 | 146.8    |
| PBE/70| 1.524| 1.450| 1.506| 1.446| 2.209 | 1.612  | 3.018 | 146.5    |
| PBE/90| 1.520| 1.447| 1.503| 1.443| 2.216 | 1.608  | 3.018 | 149.5    |
| PBE/110| 1.522| 1.447| 1.503| 1.443| 2.209 | 1.610  | 3.018 | 149.0    |

a Distances are given in angstroms, and angles are given in degrees.
TABLE 4: Optimized Geometrical Parameters of the B-Site in H–B-SOD Calculated with Different Functionals at a 70 Ry PW Cutoff

|         | B–O* | B–O1 | B–O* | B–O2 | O*–H | Si*–O* | B–Si* | B–O*–Si* |
|---------|------|------|------|------|------|--------|-------|----------|
| PW      | 2.574 | 1.373 | 1.371 | 1.358 | 0.974 | 1.639  | 3.938  | 137.3    |
| PBE/70  | 2.572 | 1.372 | 1.370 | 1.357 | 0.975 | 1.639  | 3.953  | 138.6    |
| BP/70   | 2.571 | 1.371 | 1.370 | 1.358 | 0.973 | 1.638  | 3.950  | 138.5    |
| GGA/70  | 2.609 | 1.366 | 1.361 | 1.349 | 0.962 | 1.630  | 3.985  | 139.0    |
| HCTH/70 | 2.576 | 1.374 | 1.372 | 1.358 | 0.974 | 1.638  | 3.954  | 138.4    |
| BLYP/70 | 2.603 | 1.389 | 1.388 | 1.377 | 0.982 | 1.660  | 3.929  | 133.1    |
| PBE/DZVP| 2.547 | 1.388 | 1.382 | 1.371 | 0.971 | 1.641  | 3.914  | 137.3    |

a Results obtained by adopting the GPW method with the PBE functional as a function of the basis set size are also reported. Distances in angstroms, and angles are given in degrees.

TABLE 5: Optimized Geometrical Parameters of the B-Site in Na–B-SOD Calculated with Different Functionals at a 70 Ry PW Cutoff

|         | B–O* | B–O1 | B–O* | B–O2 | O*–Na | Si*–O* | B–Si* | B–O*–Si* |
|---------|------|------|------|------|-------|--------|-------|----------|
| PW      | 1.524 | 1.450 | 1.506 | 1.446 | 2.209 | 1.612  | 3.018  | 146.5    |
| PBE/70  | 1.527 | 1.451 | 1.502 | 1.446 | 2.213 | 1.612  | 3.016  | 147.9    |
| BP/70   | 1.519 | 1.443 | 1.504 | 1.445 | 2.209 | 1.609  | 3.018  | 149.5    |
| GGA/70  | 1.516 | 1.441 | 1.507 | 1.439 | 2.209 | 1.603  | 3.020  | 150.9    |
| HCTH/70 | 1.527 | 1.449 | 1.511 | 1.447 | 2.213 | 1.608  | 3.024  | 149.4    |
| BLYP/70 | 1.518 | 1.458 | 1.506 | 1.457 | 2.203 | 1.603  | 3.024  | 151.0    |
| PBE/DZVP| 1.519 | 1.459 | 1.500 | 1.456 | 2.223 | 1.610  | 3.028  | 150.8    |
| PBE/TZV2P| 1.519 | 1.384 | 1.411 | 1.378 | 0.973 | 1.650  | 3.407  | 141.6    |

a Results obtained by adopting the GPW method with the PBE functional as a function of the basis set size are also reported. Distances in angstroms, and angles are given in degrees.

TABLE 6: Geometrical Parameters of the B-Site in H–B-SOD from Constrained Geometry Optimizations with the PW and GPW Approaches

|         | B–O* | B–O1 | B–O* | B–O2 | O*–H | Si*–O* | B–Si* | B–O*–Si* |
|---------|------|------|------|------|------|--------|-------|----------|
| PW      | 2.059 | 1.365 | 1.394 | 1.363 | 0.976 | 1.638  | 3.507  | 142.8    |
| PBE/TZV2P| 1.957 | 1.384 | 1.411 | 1.378 | 0.973 | 1.650  | 3.407  | 141.6    |

a Distances are given in angstroms, and angles are given in degrees.

The partially frozen framework are in line with the corresponding results on cluster models.31,34

The energetic stabilization associated with framework relaxation in the H–TIII-SOD systems may be estimated from the energy differences between the fully and partially relaxed structures. The framework relaxation energy per unit cell amounts to 44.6, 52.1, 65.8, and 57.2 kcal/mol for H–B-SOD, H–Al-SOD, H–Ga-SOD, and H–Fe-SOD, respectively, indicating that full relaxation of the crystal structure has a large stabilizing effect on the acid site structure.

3.3. Equilibration Process in Acid B Sodalite. Several simulations characterized by different parameters have been carried out on the H–B-SOD system (Table 1). The first two simulations, namely, BP/60(1) and PBE/70(1), in which different electronic structure treatments, starting configurations, and equilibration procedures were adopted, provided a similar average structure of the zeolite matrix. However, a significant difference was found for the mean values of the nonbonding B–O* distance, which amounted to 2.59 and 2.71 Å in BP/60(1) and PBE/70(1), respectively. This difference, affecting the loosely interacting atoms, might depend on the different electronic structure treatment but could also be related to the simulation length. The equilibration period and/or the data collection time might have been insufficient to properly sample the system under study and specifically the weak interaction between the acid site atoms O* and B. That this could be the case was suggested by the two PBE/70 and BP/60 simulations.
on Na–B-SOD that converged to the same structural description in simulation times on the order of those adopted in standard FPMD runs of zeolites (see section 3.4.2). Therefore, other simulations characterized by longer equilibration and sampling times were performed on H–B-SOD. Care was taken to check whether these simulations achieved adiabatic sampling. Indeed, the drifts of the fictitious electronic kinetic energy and of the conserved quantity in the longest trajectory (64 ps) were $5 \times 10^{-4}$ and $4.5 \times 10^{-5}$ a.u., respectively, thus confirming adiabatic sampling. As a further check, a BO simulation was carried out using the same initial configuration of the PBE/70(2) run.

Remarkably, the longer simulations converged to a closer description of the B acid site, which is characterized by a B–O* equilibrium separation of ~2.7 Å and large thermal fluctuations. The differences in the results of the two shortest simulations are therefore not related to the choice of DFT approximation, basis set size, or thermostats but rather to the time required to equilibrate the weakly interacting units composing the B acid site, namely, the silanol group and the trigonal BO3 moiety. By taking into account that the presence of the B-site perturbs the zeolite framework, a long thermalization period is needed by the zeolite lattice to reorganize and find the optimal structural arrangement (i.e., the equilibrium geometry) for simulations at ambient temperature conditions.

### 3.4. Finite Temperature Behavior of Boron Sites

#### 3.4.1. Protonated Site

In view of the previous results, the behavior of the B acid site will be thoroughly discussed only in the case of the longest simulation (that is, PBE/70(2) in Table 1). However, comparison with results obtained with other DFT approximations and PW cutoffs, as well as with different sampling approaches, will be shortly presented. Indeed all of the trajectories longer than ~20 ps, or with a long equilibration period (more than 10 ps), were found to provide well-converged results.

Figure 3 shows the B–O running coordination numbers $N(r)$ (top) and the radial distribution functions $g(r)$ (bottom) calculated for the H–B-SOD system from FPMD simulations with different functionals and PW cutoffs (namely, PBE/70(2) and BP/60(2)). Results from BO simulations at the PBE/70 level are also shown for comparison.

The similarity of the curves indicates a very good agreement between the simulations. Moreover, the structural properties obtained from the two CP trajectories are consistent with the BO simulation as well, thus confirming that both the time step and the inertia parameter here selected for the CP equations (i.e., 5 and 500 a.u., respectively) are appropriate for the description of these systems.

The B–O $g(r)$’s are characterized by two maxima in the nearest neighbor region. The first peak is found at 1.4 Å; correspondingly, $N(r)$ increases from 0 to 3 for distances comprised between 1.2 and 1.4 Å, remaining equal to 3 up to 2.4 Å. This is a clear indication of the fact that the B atom retains the trigonal geometry at room temperature. Relevant mean geometrical parameters along with their relative standard deviations are reported in Table 8. The distances between B and O1, O*, and O2 do not differ significantly (less than 0.2%) from those of the minimum energy structure, indicating that the temperature has a negligible effect on the internal geometry of the BO3 moiety. Inspection of these values also reveals that both at 0 K and at room temperature the B–O2 distance is shorter than the B–O1 and B–O* ones. This can be due to the fact that, while the B–O1 and B–O* bonds belong to an unbroken four-membered ring, the B–O2 bond is part of the relaxed four-membered ring, which is broken due to the formation of the silanol group. Also the average O–B–O angles in the BO3 site are practically identical to the minimum energy structure values and correspond to a planar trigonal geometry.

The second peak in the B–O $g(r)$ is located at ~2.7 Å, while $N(r)$ reaches a value of 4 at 3.0 Å. This peak is much broader than the first one and corresponds to the B–O* contacts. Here, the thermal effect on the structural arrangement of the weakly interacting groups Si*–O*H and BO3 is evident. In fact, by comparing the B–O* average distance (Table 8) with the
corresponding value obtained from GOs (Table 2) an increase of about 0.15 Å (5%) is found. In addition, the trigonal B–O bond distances are characterized by standard deviations significantly smaller than that of the B–O* mean distance, indicating a large thermal fluctuation of this latter contact. Also the mean B–O*–Si* angle and the B–Si* distance are different from the minimum energy values; however their change in percent (2–3%) is lower than that found for the B–O* separation. These findings suggest that the silanol oxygen O* is the most mobile framework atom. Indeed, being less constrained by the framework, O* can explore at finite temperature a wide range of configurations characterized by different separations from the trigonal BO3 unit.

These results suggested further analysis of the behavior of the silanol group during the longest run. Actually, after 11 ps, a structural rearrangement of the protonated B-site was observed. Such a reorientation, which was accompanied by concerted rigid rotations of the BO3 unit and of SiO4 tetrahedra connected to the site, led to the geometrical arrangement represented in Figure 4. Such a transition, detected in the 64 ps long simulation, brings the O*H silanol group, originally pointing to the center of a six-membered ring, in an adjacent six-membered ring. This reorientation of the silanol group, which occurs in less than 1 ps, does not affect the average structural properties of the system. Indeed, as clearly demonstrated by Figure 4, the final geometrical arrangement of the protonated B-SOD system may be considered as the mirror image of the starting structure, obtained by rigid unit rotations.61

3.4.2. Unprotonated Site. Information on the average structure of the B-site in Na–B-SOD can be extracted from the BO–Og(r) and N(r), shown in Figure 5, while average bond distances and angles are reported in Table 9. A first sharp peak in the g(r) is located at 1.49 Å, while the N(r) integrates to 4 at 1.65 Å for both BP/60 and PBE/70 simulations. This is a clear indication that in the Na–B-SOD system the B atom remains tetracoordinated also at room temperature. The thermally averaged structure of the B-site in Na–B-SOD is characterized by O–B–O angles typical of a tetrahedron and is very close to the minimum energy geometry. As found in the optimized structure, the longest B–O bond involves the O* atom, which is the one closest to the Na extraframework cation. The parameter affected by the largest oscillations around the six-membered ring in an adjacent six-membered ring. This reorientation of the silanol group, which occurs in less than 1 ps, does not affect the average structural properties of the system. Indeed, as clearly demonstrated by Figure 4, the final geometrical arrangement of the protonated B-SOD system may be considered as the mirror image of the starting structure, obtained by rigid unit rotations.61

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equilibrium value is the O\textsuperscript{*}–Na separation, as expected on the basis of the higher mobility of the extraframework species Na\textsuperscript{+}.

The calculated standard deviations of the four B–O bonds are 1 order of magnitude smaller than that found for the B–O\textsuperscript{*} separation in H–B-SOD. Also, the Si\textsuperscript{*}–B distance is shorter and its oscillation smaller with respect to the H–B-SOD system. This indicates a lower flexibility of a zeolite framework containing a tetrahedral B-site with respect to the case of the silanol–BO\textsubscript{3}–containing zeolite.

Differences in the framework structure of the protonated and nonprotonated model boralites may be detected by inspection of the \(g(r)\)'s calculated for the \(\text{Si}–\text{Si}\) and \(\text{B}–\text{Si}\) atoms (Figure 6). The \(\text{Si}–\text{Si} g(r)\)'s show a first maximum peaked at 3.18 Å, while the positions of the second peak are 4.4 and 4.3 Å for H–B-SOD and Na–B-SOD, respectively. As expected, differences are even more evident in the \(\text{B}–\text{Si} g(r)\)'s: The peak corresponding to the first neighbors is found at shorter distances in H–B-SOD because the B–O bonds in the BO\textsubscript{3} units are shorter than those in the BO\textsubscript{4} tetrahedron. Moreover, also the second and third peaks are displaced at shorter distances in H–B-SOD due to the breaking of the four-membered ring containing the B acid site. These results suggest therefore that the presence of different local geometries of the B-sites in Na–B-SOD and H–B-SOD affects the framework structure also at long range.

3.5. Acid Site in B, Al, and Ga Zeolites. Significant flexibility of protonated boralites, due to the nonbonding interactions characterizing the B acid site, emerged from the previous discussion. In this respect, comparison with results obtained with the same approach on other zeolitic acid sites may be of relevance to gain further insight on this issue.

Al and Ga protonated model sodalites were studied by FPMD at the BP/60 level.\(^3\)–\(^5\) As confirmed by a PBE/70 simulation on the H–Al-SOD system (Figure 7), structural properties predicted by the BP/60 scheme are well in line with the PBE/70 description, thus allowing comparison between results obtained at finite temperature on T\(\text{III}\)-substituted zeolites (T\(\text{III}\) = Al, Ga, or B).

The T\(\text{III}\)–O, Si–Si, and T\(\text{III}\)–Si \(g(r)\)'s for the B, Al, and Ga acid sodalites are shown in Figures 7 and 8, while the distributions of the T\(\text{III}\)–O\textsuperscript{*} distances and O–T\(\text{III}\)–O and T–O–T angles are represented in Figure 9.

In the system containing the strongest acid site, namely, H–Al-SOD, the Al–O \(g(r)\) is characterized by a first maximum with a peak at 1.77 Å and a shoulder at 1.95 Å. The peak
corresponds to the most probable bond distances of Al with the three unprotonated oxygens. The shoulder represents the Al-O* bond, which is elongated with respect to the other Al-O distances, but still is part of a tetrahedral AlO_4 unit. Such a tetrahedron is however distorted as indicated by the O-Al-O angle distribution, which is characterized by two maxima at 102° and 114°.3

In the H-Ga-SOD case, the Ga-O g(r) first maximum is split into two peaks: The first one, which integrates to 3, is found at 1.85 Å, while the second is located at 2.20 Å and corresponds to the Ga-O* contacts. Such a value suggested that the Ga-O* bond is weaker than the Al-O* one and that the Ga acid site has silanol character.4 Also, the O-Ga-O angle distribution indicates that the geometry of the Ga site is significantly distorted from an ideal tetrahedron and might be considered close to a trigonal arrangement. However, the limiting case of a planar trigonal geometry of the TIIIsite and a nonbonded silanol group is achieved only in the case of the B acid site, as indicated by the O-B-O angle distribution. The peaks at 90° and 120° correspond respectively to the three angles involving O* and to the three angles of the BO_3 unit.

By comparing the distributions of the TII-O* distances, another interesting feature emerges. In the case of H-B-SOD, the full width at half-height (0.6 Å) is about twice that found for the Al- and Ga-containing systems. Such a result is a consequence of the weak interaction between O* and the BO_3 group and suggests that the high flexibility of the acid site, detected in the H-B-SOD simulations, is indeed a peculiarity of protonated boralites. In addition, this large local flexibility is also at the basis of significant structural deformations of the framework, which, as shown by the TII-Si g(r)'s, are much larger than in the Al and Ga zeolites. Moreover, the average T-O-T angle distribution is broader and tails toward larger values of the angles in the case of H-B-SOD.

Figure 9. Distributions of selected geometrical parameters calculated for the H-TII-SOD systems (TII = B, Al, or Ga). Top panel: TII-O* distance distribution. Middle panel: Distribution of the six O-TII-O angles. Bottom panel: T-O-T angle distribution. T represents any framework cation.

Figure 10. Representation of the average geometries of the TIIIO_4(H) polyhedra calculated for the H-TII-SOD systems (TII = B, Al, or Ga). The average geometry of the SiO_4 unit, calculated for the all-silica sodalite structure, is also shown for comparison: (A) AlO_4(H); (B) GaO_4(H); (C) BO_4(H); (D) SiO_4. The distances of the TII atom from the plane formed by the three unprotonated oxygens are 0.37, 0.27, and 0.02 Å for Al, Ga, and B, respectively. The corresponding distance in SiO_4 is 0.53 Å. Atom color code: Al, light blue; Ga, green; B, yellow; Si, gray; O, red; H, white.

Figure 11. Calculated vibrational spectra of Na-B-SOD, H-B-SOD, and H-Al-SOD.

The O-Si-O angle distributions calculated for all 11 Si sites show a single maximum at 109.5°. In the AlO_4 unit, the splitting of the O-Al-O distribution in two peaks indicates that the tetrahedral geometry is distorted to a trigonal pyramid structure. Indeed, while Si occupies the center of the SiO_4 tetrahedron, Al is found closer to the three unprotonated oxygens (Figure 10). Such oxygens may be considered as forming a pyramid base, while the fourth and more distant protonated oxygen O* may be considered as the pyramid vertex. The pyramidal character becomes more pronounced in the case of GaO_4, as indicated by a comparison of the TII-O* distance distributions and by the broadening and further splitting of the O-Ga-O distribution. However, the transition to a trigonal pyramid is fully accomplished only in the case of B, because the TII cation lies in the same plane of the three oxygen atoms forming the basis of the pyramid (Figure 10).
TABLE 9: Selected Mean Geometrical Parameters and Relative Standard Deviations for the Na–B-SOD System

|                  | B–O* | B–O1 | B–O# | B–O2 | O*–Na |
|------------------|------|------|------|------|-------|
| BP86/60          | 1.529 (0.05) | 1.455 (0.04) | 1.514 (0.04) | 1.454 (0.04) | 2.264 (0.11) |
| PBE/70           | 1.527 (0.05) | 1.450 (0.04) | 1.513 (0.05) | 1.450 (0.04) | 2.262 (0.12) |
| Si–O*           |      |      |      |      |       |
| B–Si*           |      |      |      |      |       |
| BP86/60          | 1.619 (0.03) | 3.021 (0.06) | 147.9 (6) | 109.4 (3) |
| PBE/70           | 1.613 (0.03) | 3.021 (0.06) | 149.0 (7) | 109.4 (3) |
| B–O*–Si*        |      |      |      |      |       |
| (O–B–O)         |      |      |      |      |       |

* Distances are given in angstroms, and angles are given in degrees. (O–B–O) indicates the average of the six angles in BO₄.

3.6. Simulated Vibrational Spectra. IR spectroscopy has been widely used for the characterization of borates. Most studies were performed on B-MFI-type zeolites; therefore only a qualitative comparison with the present simulation results is possible.

In IR studies on evacuated B-MFI, a band at 1380 cm⁻¹ was assigned to three-coordinate B, while bands in the 860–920 cm⁻¹ range and at 670 cm⁻¹ were assigned to B–O–Si symmetric stretching and bending modes, respectively. A band at 700 cm⁻¹ and a shoulder at 970 cm⁻¹ were detected as well. Further studies reported the presence of two bands at 1380 and 1405 cm⁻¹, attributed to trigonal B occupying different crystal sites. Moreover, other authors reported bands at 1390 and 1150 cm⁻¹, ascribed to the simultaneous presence of trigonal and tetrahedral B-sites, respectively. The latter band may be partially hidden by the crystalline silica matrix modes.

B-CHA has been recently synthesized and characterized by IR spectroscopy. Upon calcination two main changes are reported: A typical signal of a trigonal BO₃ site appears at 1390 cm⁻¹, and a shoulder at 810 and 912 cm⁻¹. The 1390 cm⁻¹ band disappears when B-CHA is contacted with water.

In general, while the presence of BO₃ is easily detectable from the 1390 cm⁻¹ band, no definite attribution of a specific band to BO₄ modes has been possible by IR spectroscopy in evacuated borates. This might also be due to the possible simultaneous presence of both kinds of sites in the same sample. Moreover, preparation, heating, and dehydration procedures undergone by the samples may affect band position and intensity.

Vibrational spectra were obtained from the Fourier transform (FT) of T–O–T oscillation autocorrelation functions calculated from the longest trajectory of each system. Partial spectra due to bending, symmetric, and asymmetric stretching modes of the T–O–T bridges were also calculated by the FT of the relative correlation functions. Such an analysis allowed us to single out the contributions to the total spectrum of specific atoms or groups of atoms, thus enabling the assignments to the bands.

Due to the use of a fictitious inertia parameter in the CP equations, simulated vibrational frequencies are systematically red-shifted. Correction of this friction effect by application of a proper scaling factor is a well-established procedure, already applied to the study of other zeolitic systems. In the present case, the scaling factor has been determined from the ratio of the H₂O asymmetric stretching frequencies calculated from 30 ps long BO (3690 cm⁻¹) and CP (3532 cm⁻¹) trajectories of an isolated water molecule at the PBE/70 level.

Calculated vibrational spectra of Na–B-SOD and H–B-SOD are reported in Figure 11. The most significant differences between the two spectra are found in the 3600 and 1300 cm⁻¹ regions. A peak at 3643 cm⁻¹ is present only in the H–B-SOD spectrum and is due to the silanol hydroxyl group. The double band in the H–B-SOD spectrum, with peaks at 1338 and 1290 cm⁻¹, is missing in the Na–B-SOD one. Such a band has to be attributed to the trigonal B-site. The splitting into two peaks derives from the different B–O bond lengths in H–B-SOD (Table 8). Actually, by calculating the distinct contributions of the three B–O–Si subsystems to the total spectrum, it is found that the higher wavenumber component is due to the B–O–Si structure characterized by the shortest B–O bond.

In the T–O–T asymmetric stretching region, the Na–B-SOD spectrum presents a shoulder at 1184 cm⁻¹ that is missing in the H–B-SOD system. This feature should be ascribed to B in tetrahedral sites and in particular to the B–O–Si asymmetric stretching component. At lower wavenumbers, a signal at 910 cm⁻¹ is present in both systems and is associated with the symmetric B–O–Si stretching modes. In the so-called silica window region (800–1000 cm⁻¹) of the Na–B-SOD spectrum, other bands are present with peaks at 850, 944, and 974 cm⁻¹, while a peak at 860 cm⁻¹ is found in the case of H–B-SOD. Such peaks are associated with B–O–Si symmetric stretching modes. The presence of a multiple band is due to the different interactions of Na with the oxygen atoms of the BO₄ tetrahedron. Indeed only two of them are coordinated to Na with different distances (Table 9).

Concerning the OH stretching mode, the one calculated for H–B-SOD is blue-shifted by 85 cm⁻¹ with respect to the corresponding value found in H–Al-SOD (3558 cm⁻¹). This difference can be considered a spectroscopic index of the different strengths of acid sites in Al- and B-zeolites.

It may be relevant to point out that, as opposed to both boron-containing models, in the H–Al–SOD system no band is found either in the silica window or in the ~1300 cm⁻¹ region. Therefore, in the commonly occurring situation of the simultaneous presence of B and Al, framework B-sites can be experimentally detected by bands in the silica window, while the signature of trigonal B is found in the 1300 cm⁻¹ region. However, the shoulder at 1184 cm⁻¹ calculated for the tetrahedral B-site in Na–B-SOD may overlap with a similar signal calculated at 1179 cm⁻¹ in the case of H–Al-SOD.

4. Summary and Conclusions

Extensive simulations on boron-containing sodalites have been performed, both at 0 K and at room temperature, by using periodic DFT level approaches. These studies allowed us to highlight a strict correlation between the framework flexibility and the local geometry of the B acid site. On the basis of the results presented here it emerges that, to achieve a proper
structural description of these systems, the use of a computational scheme allowing the full relaxation of the zeolite framework is recommended. As opposed to Na\(^+\)-exchanged B-SOD systems, characterized by tetrahedral B-sites, anhydrous protonated borolites are characterized by BO\(_x\) polyhedra with a pyramidal geometry, where the base is a BO\(_3\) planar structure and the vertex is an oxygen bearing a proton. Such a picture is confirmed by all of the different computational schemes here adopted. Comparison among T\(^{1\text{H}}\)-containing acid sites has shown a progressive transition of the T\(^{1\text{H}}\)O\(_4\) polyhedron from a tetrahedral to a pyramidal structure. A calculation of the O\(^*\) distance elongation, from 2.57 Å at 0 K to \(\geq 2.7\) Å at 300 K, due to the silanol group mobility. Last but not least, the presence of a protonated B-site in a SiO\(_4\) tetrahedra network induces a relevant framework flexibility, which may also allow long-range structural rearrangements through concerted rotations of rigid polyhedral units. Such a flexibility may be responsible for the peculiar behavior of acid borolites toward basic molecules and might also be related to the facile nature of the de-boronation process.

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