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Deactivation Kinetics of Solid Acid Catalyst with Laterally Interacting Protons

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ABSTRACT: Differences in catalyst deactivation kinetics in solid acid catalysis are studied with catalyst models that allow for lateral interaction between protons. Deactivation of a solid acid catalyst with laterally interacting protons induces inhomogeneity of proton reactivity that develops with time. As a consequence, product selectivity changes and deactivation will accelerate. This is demonstrated by simulations of the deactivation kinetics of the alkylation reaction of propylene with isobutane. The effect of lateral interactions between protons arises because initial catalyst deactivation is not caused by pore blocking or coke deposition but by a molecular mechanism where protons are consumed due to the formation of stable nonreactive carbenium ions. High selectivity to alkylate requires a catalyst with protons of high reactivity. When protons become consumed by formation of stable deactivating carbenium ions, initially reactive protons are converted into protons of lower reactivity. The latter only catalyze deactivating oligomerization reactions. Simulations that compare the deactivation kinetics of a catalyst model with laterally interacting protons and a catalyst model that contains protons of similar but different reactivity, but that do not laterally interact, are compared. These simulations demonstrate that the lateral interaction catalyst model is initially more selective but also has a lower stability. Catalyst deactivation of the alkylation reaction occurs through two reaction channels. One reaction channel is due to oligomerization of reactant propylene. The other deactivation reaction channel is initiated by deprotonation of intermediate carbenium ions that increase alkene concentration. By consecutive reactions, this also leads to deactivation. The hydride transfer reaction competes with oligomerization reactions. It is favored by strongly acid sites that also suppress the deprotonation reaction. Within the laterally interacting proton catalyst model, when reactive protons become deactivated, weakly reactive protons are generated that only catalyze deactivating oligomerization and consecutive reactions. This rapid formation of the weakly reactive protons is the cause of decreasing selectivity with reaction time and increased rate of deactivation. Solutions of the mean field kinetic equations as well as stochastic simulations are presented. Comparative simulations with a reduced number of neighbors of the protons illustrate decreased deactivation rates when the proton density decreases. Island formation of adsorbed reaction intermediates on the catalyst surface is observed in stochastic kinetics simulations. When alkylselectivity is high, this island formation increases the rate of catalyst deactivation in comparison to the rate of deactivation according to the mean field studies. A nonlinear dynamics model of proton dynamics is provided, which shows that the differences between stochastic and mean field simulations are due to frustrated proton state percolation.

KEYWORDS: alkylation catalysis, solid acid catalysis, catalyst deactivation, laterally interacting protons, kinetics simulations, nonlinear dynamics, site percolation

1. INTRODUCTION

Surface chemical reactivity is strongly affected by lateral interactions between chemisorbed reaction intermediates. Then, for a catalytically reactive surface, sites of activity cannot be considered to behave independently. On transition metals, it may lead to collective phenomena such as surface island overlayer formation or dynamic phenomena such as time-dependent oscillatory reaction kinetics.¹

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The physical chemistry of lateral interactions of protons in solid acid catalysis has been much less investigated. However, it has become understood that the close presence of protons near each other tends to enhance their reactivity.

Classical ion exchange experiments demonstrate a reduction in intrinsic proton reactivity when protons become exchanged by cations. It induces a strengthening of the proton bond zeolite framework mainly caused by the negative charge that is generated on the zeolite framework. This was also demonstrated by earlier quantum-chemical calculations with cluster models of the zeolite framework. A recent experimental study on methanol activation shows enhanced reactivity with high proton concentration. Enhanced catalyst deactivation when the proton concentration is greater is found in a model study of the methanol to olefin (MTO) reaction.

For the alkylation reaction of propylene and isobutane that is catalyzed by solid acid catalysts, we will computationally study catalyst deactivation as a function of proton concentration. Proton reactivity differences can be caused by changes in framework composition or structure of the zeolite. Instead, in the simulations that we present, zeolite framework composition is not altered and proton reactivity differences are due to a proton-removing reaction caused by a nonspecific deactivating reaction.

Whereas often pore or site blocking, which may cause mass transfer limitations, are the cause of accelerated deactivation, for the low-temperature alkylation reaction such deactivation is initially not dominant. A molecular mechanism then initiates catalyst deactivation.

Alkylation of propylene or n-butene with isobutane is a widely used refinery process, based on neat sulfuric acid or hydrogen fluoride as catalyst. In this low-temperature reaction, branched C₃ or C₄ alkanes (the alkylate) are formed as useful high-octane fuel. Replacement of this homogeneous process by a heterogeneous process is highly desirable. The main drawback of the use of a heterogeneous solid acid catalyst is its short catalyst lifetime of approximately 10 h.

Experimental studies of deactivation kinetics of this reaction show, after an initial period of selective alkylate formation, accelerated deactivation and a delayed production of oligomerization product. Research on the recently developed Alkyl Clean process has demonstrated that the initial deactivation is due to the formation of deactivating hydrocarbons that can be readily removed in a successive low-temperature hydrogenation reaction step.

The delayed production of alkene oligomers and high reactivity of deactivating adsorbed coproduct molecules indicate that the deactivation mechanism has a molecular chemical origin. When pore blocking and mass transfer limitations would cause deactivation, oligomerization would not continue after alkylate production has decayed. We will present model deactivation kinetics simulations that show such an accelerated decline of alkylate production and overshoot of oligomer production. The low-temperature deactivation of the alkylation reaction is different from the deactivation kinetics of high-temperature solid acid catalytic reactions such as catalytic cracking and the MTO reaction, where deactivation is dominated by pore or site blocking of the zeolite catalyst micropores by deposition of a carbonaceous residue of low reactivity, which has to be oxidatively removed. An additional reason for accelerated decay is the often inhomogeneous reactant and product distribution in catalytic plug flow reactors. This is critical to alkylation deactivation kinetics where alkene concentration has to be kept low. It is this reason that the alkylation reaction is preferentially performed in a slurry reactor or continuously stirred tank reactor (CSTR). Comparison of plug flow reactor data with CSTR data shows an increase in lifetime from 30 min to 15 h.

The fundamental reason that in many solid acid catalyzed reactions catalyst deactivation is rapid is the usual presence of alkenes as the reactant, reaction intermediate, or product. Even when saturated molecules are converted as, for example, alkanes in the catalytic cracking reaction or methanol in the MTO reaction, alkenes will be formed as reaction intermediates. Alkenes are highly reactive and will readily oligomerize. They are often desirable reaction intermediates, but apart from leading to formation of desired product, they will, through successive reactions, also produce deactivating carbonaceous residue or coke. In the alkylation reaction, this problem is even more severe since propylene or n-butene is used as a reactant.

Different from high-temperature catalysis, where catalyst deactivation is mainly due to coke deposition, catalyst deactivation of the alkylation reaction is initially dominated by proton consumption. It is the result of a competitive chain of reactions of unsaturated reaction intermediate molecules that is related to the paring reaction. Proton consumption then is due to the formation of stable substituted cyclopentadienyl positively charged carbenium ions. This happens when the Lewis basicity of the deprotonated zeolite framework reaction center is less than the reaction intermediate that becomes protonated. Not only will the proton consumption reduce the surface concentration of the protons that are left but also these protons will also have reduced intrinsic reactivity. This is, among other things, due to the previously mentioned negative charge that proton removal generates on the zeolite lattice.

We will conclude this introductory section with a short summary of the current understanding of the molecular mechanism of the alkylation reaction. It will provide an opportunity to introduce the two reaction channels that are the main cause of catalyst deactivation. We will, at the close of this section, also introduce the laterally interacting proton model that is the core topic of the paper.

Several catalytic reaction cycles combine to convert alkene and isobutane into alkylate. The complex network of the alkylation reaction is reasonably well understood. We will base our model on the catalytic mechanism of the reaction previously presented and the earlier foundling work by Schmerling.

In a previous paper, we have extensively studied the full reaction mechanism of the alkylation reaction of propylene and isobutane to give C₇ and C₈ isomers using quantum chemically calculated elementary reaction rate constants of the many elementary reaction steps that constitute the corresponding catalytic cycles. These elementary reaction rate constants have been applied in microkinetic simulations of the complete reaction system. The information deduced from these microkinetic simulations will be used here as input to the lumped kinetics simulations. In order to make the analysis of the deactivation kinetics tractable, the kinetic simulations will be based on a simplified, but useful, version of the alkylation reaction cycle.

A schematic representation of this cycle is shown in Figure 1, which is useful to discuss the kinetic conditions that determine high alkylate selectivity and slow catalyst deactivation.

The reaction consists of an initiation cycle and a propagation cycle. As we will see, no deactivation will occur when the
propagation cycle has no feed-forward loop connection with the reaction initiation cycle.

The alkylation reaction starts with adsorption of propylene to a proton. This results in a propyl carbenium cation that becomes adsorbed to the surface as an alkoxy intermediate. The propyl carbenium ion initiates the alkylation reaction by consecutive reaction with isobutane that gives (undesirable) propane as a product and the desirable isobutyl cation as reaction intermediate. It will depend on the specific zeolite whether this carbenium ion will remain adsorbed to the surface or will be a carbenium ion intermediate that nearly freely rotates and weakly interacts with the surface.27,28

In the initiation reaction, the reaction of the propyl cation with another propylene molecule that gives a protonated oligomer competes with the formation of the isobutyl cation by reaction of isobutane. The oligomerization reaction will initiate consecutive reactions that deactivate the catalyst.

The reaction of isobutane with the propyl cation is an example of a hydride transfer reaction. Transfer of the hydride ion converts the propyl cation into propane. The isobutyl cation intermediate that is cogenerated from isobutane now carries a positive charge. As long as the rates of the respective hydride transfer reactions are fast, deactivation will be suppressed and alkylate selectivity will be high.

The hydride transfer reaction has been quantum chemically well investigated.26,29,30 Especially when larger carbenium ions are involved, the corresponding reaction intermediates can be considered as nearly freely moving. Their formation has relatively low activation barriers when proton reactivity is large.19 Transition states of the alkene oligomerization reaction have more constrained mobility and a stronger interaction with the proton reaction center and are hence less proton reactivity demanding.19,31 Therefore, competition between hydride transfer and oligomerization reactions is in favor of the hydride transfer reaction when it is catalyzed by highly reactive proton sites.10,26

Once isobutyl cations are formed, the propagation reaction cycle begins. The isobutyl cation reacts with a propylene molecule to give a C7+ cation. In the next reaction step, hydride transfer with isobutane will produce the C7 alkylate molecule and the isobutyl cation will be regenerated. The propagation cycle continues by a following reaction of the isobutyl cation with another propylene molecule etc.

The isobutyl cation can be considered the organo-catalyst of the propagation reaction cycle. In this propagation reaction cycle, the proton does not play an explicit role.

The role of the proton becomes different when parasitic deprotonation reactions occur that convert the iC4+ cation and C7+ cation to their respective olefins. Then a proton is back-donated to the solid. The rate of deprotonation reactions becomes suppressed by competitive hydride transfer reactions of the C7+ or iC4+ cations with isobutane.

The occurrence of the deprotonation reactions opens a second deactivation channel next to that propylene oligomerization. The olefins produced by deprotonation will also initiate consecutive oligomerization reaction that deactivate the catalyst.

Additionally, a feed-forward loop with the reaction initiation cycle opens due to regeneration of protons. This will reinitiate the initiation reaction channel that competes with the deactivating propylene oligomerization.27,28

Figure 1. Scheme of the alkylation reaction cycle including deactivation reaction paths. Desirable rate relations are indicated for high selectivity of alkylate and slow deactivation rate.

Figure 2. Schematic illustration of the laterally interacting two proton reactivity catalyst model: the dual interacting proton catalyst model. (a) Illustration of the three proton states: reactive state H1+ that catalyzes alkylate production, deactivated state H2+ with no reactivity, and the state H3+, where the proton has lower reactivity and will only catalyze alkene oligomerization. (b) Dynamics of proton deactivation in the dual interacting proton catalyst model. Deactivation of the reactive H1+ proton state by catalytic reactions is slow, but once a deactivated proton state H2+ is a neighbor, a proton in its H3+ state is rapidly converted into a proton in the less reactive H1+ state. The latter also deactivates slowly by catalysis but only catalyzes oligomerization and consecutive deactivating reactions.
Theoretically, the alkylation catalyst will be infinitely stable as long as deprotonation of intermediate carbenium ions does not occur, since then the initiation reaction cycle will not be reactivated.

In the following sections, detailed kinetic simulations will be presented on the basis of the mechanistic principles discussed here and illustrated in Figure 1. Catalyst models will be compared with laterally interacting and noninteracting protons. For further comparison, we will also present kinetics simulations as a function of proton concentration. The laterally interacting proton catalyst model that we will use is illustrated in Figures 2. This dual interacting proton catalyst model simplifies differences in reactivity between protons to only two kinds: a strongly reactive and a weakly reactive proton. Within the interacting proton model, three different proton site states are defined: the strongly reactive proton state \( H_1^+ \), a deactivated proton state \( H_2^- \), and the weakly reactive proton state \( H_3^- \). The strongly reactive proton state \( H_1^+ \) catalyzes alkylation formation as well as olefin oligomerization. It deactivates, by deactivating side reactions of the alkylation reaction cycle, to the state \( H_2^- \). Once a proton in the state \( H_1^+ \) has a nonreactive \( H_2^- \) proton state as a neighbor, due to the now present negative charge on the zeolite framework in comparison to the deactivation time of the protons, it will nearly instantaneously (faster than a picosecond) convert to a proton state of lower reactivity, \( H_3^- \), that cannot catalyze alkylation formation. Protons in the \( H_3^- \) state will only catalyze olefin oligomerization. They will also deactivate to proton state \( H_2^- \).

Experimental evidence of protons of different reactivity that give only selective alkylation versus alkene oligomerization is taken from the reports of Lercher et al. For zeolites of the faujasite structure, they demonstrated that differences in alkylation selectivity between La-promoted X and Y zeolites are related to the location of the reactive proton near La positions in cavities of the zeolite framework. Microkinetics simulations of La-containing and non-La-containing zeolites confirm the large selectivity difference of such protons, which can be distinguished by their large difference in ammonia adsorption energy.

In the next section, for later reference, we will present kinetics simulations of deactivation of the alkylation reaction for a single alkylation reactive proton state catalyst model. This will be followed by analogous simulations of the dual interacting proton catalyst model of Figure 2. In the course of deactivation, it generates next to a reactive proton also the less reactive proton. Results will be compared with kinetic simulations of a catalyst model that also contains two kinds of protons with different reactivity, but in this case they do not laterally interact and both are present at the start of the reaction. The highly reactive proton is alkylation selective, and the weakly reactive proton only catalyzes oligomerization.

In the discussion section that follows, we will analyze the nonlinear dynamics of the dual interacting proton model by comparing mean field and stochastic simulations. It will appear that the high coverage of the proton sites with reaction intermediates leads to deviations from the mean field kinetics simulations. Surface percolation of proton dynamics affects the dependence of deactivation on proton concentration. A simple three proton state model, which is not explicitly coupled with kinetics that we introduced previously, will not suffice. The paper is concluded with a short summarizing conclusion section.
conditions, where changes in reactant concentration by reactions can be ignored. The product accumulates as in a batch CSTR reactor, the preferred reactor for alkylation. Proton concentration dependence is calculated through definition of dimensionless reaction rate constants as discussed in section SI 2.

Since, as we will see for the dual interacting and noninteracting proton catalyst models, deactivation times of the respective products will be different, accumulated product cannot be used as a measure of product selectivity. A steady-state alkylation selectivity that measures the fraction of propylene incorporation into alkylate can be defined in the intermediate reaction time regime where surface intermediate concentrations are finite, and stationary. It is calculated from the expression $S(C_i) = \frac{C_i}{(C_2 + C_3 + C_{-i} + xC_m)} \times 100\%$. Since we do not consider the oligomer length explicitly, $x$ is set equal to 1. The rates of production, $C_i$, are deduced from the slopes of the $C_i(t)$ vs $t$ curves. In alkylation catalysis, next to the selectivity, catalyst lifetime is usually the measure by which the activities of different catalysts are compared. The catalyst lifetime is defined as the time when the rate of production starts to decline exponentially. The catalyst decline rate is defined as the inverse slope of the logarithmic plot of $C(t)$ vs $t$ when the catalyst deactivation rate has become exponential. The unit of time is $k_i^{-1}$, which is maintained the same in the simulations. In the respective figures (or legends), we mention alkylation selectivities as well as lifetimes and deactivation rates.

We will compare kinetics simulations when alkylate is the major product and when instead oligomerization of alkenes dominates. The differences are defined by respective default elementary reaction rate parameters that remain the same throughout the paper, unless mentioned specifically.

First-principles microkinetics simulations based on quantum chemically calculated elementary reaction rate data representative for these two cases are available. In these simulations, faujasite zeolites with reactive and less reactive protons were compared. The default reaction rate parameters selected in the kinetics simulations of this paper have been chosen to give approximate agreement with the data of the microkinetics simulations. The default reaction rate constant of the initiating hydride transfer reaction $k_1$ (Figure 3), which is scaled to 1, has been chosen to be equal to the rates of the reactions of the propagation cycle ($k_4 = k_1$, $k_5 = k_2$ = 1). Reaction rate constants of deprotonation of $iC_4^+$ and $C_{7-}$ are chosen 1 order of magnitude less ($k_3 = k_5 = 0.1$). Reaction rate constants of proton deactivation by alkenes or alkene oligomers again are chosen 1 order of magnitude smaller ($k_6 = k_7 = k_8 = 0.1$). The reaction rate of propylene protonation to form intermediate carbenium ion is fast ($k_9 = 10$). $k_{10}$ the lumped elementary reaction rate constant of reactant alkene oligomerization, has been chosen to vary. For high alkylation selectivity case, $k_{10}$ is chosen equal to $k_1$, the elementary reaction rate constant of hydride transfer. For the low alkylation selectivity case, $k_{10}$ is chosen as 10 times $k_1$. This increased reaction rate constant can be considered to be due to an increased reactant concentration of propylene. There is no $k_{11}$. This is reserved for the elementary reaction rate constant of proton dynamics in the dual site interacting proton catalyst model to be discussed in section 3.

Figure 4. Deactivation kinetics: the case when alkylate production dominates ($k_1 = k_{10} = 1$, $k_2 = k_4 = 1$, $k_3 = k_5 = 0.1$, $k_6 = k_7 = k_{11} = 0.01$, $k_9 = 10$). For kinetic symbols, refer to Figure 3. The rate of hydride transfer is comparable to that of propylene oligomerization. (a) Change in reactant surface concentration at short time scale. (b–d) Longer time scales for the change in surface concentration, product formation, and rate of product formation, respectively, as a function of time. The output concentrations of C3 and C7− and their rates of formation are the same (overlap of green and brown curves in (c) and (d)). In addition, the output concentrations of iC4+ and C7− are the same (overlap of black and blue curves (c) and (d)).
2.1. Deactivation Kinetics for High and Low Alkylation Selectivity. We will compare deactivation kinetics when alkylation selectivity is relatively high and low. It will mainly depend on the reaction rate ratios of hydride transfer reactions ($k_1$ and $k_4$) versus the reaction rates of propylene oligomerization ($k_{10}$) and deprotonation ($k_3$, $k_5$), respectively. We will also consider the effect of proton concentration on selectivity and deactivation rate of the reaction.

In Figure 4, we consider the case where alkylate production dominates. We select the rate of reaction initiation by hydride transfer to the $C_3^+$ cation to be comparable to the rate of olefin addition ($k_1 = k_{10} = 1$). For this case, Figure 5 compares deactivation rate as a function of proton concentration. Figure 6 gives the results of kinetics simulations where the elementary reaction rate of olefin oligomerization is much faster ($k_1 = 1$, $k_{10} = 10$).

For high alkylation selectivity, the relative concentration of reactant propylene has to be low in comparison to that of isobutane. Figure 4a,b shows surface coverage in the initiation and deactivation time regimes, respectively. Product distributions and their respective rates of deactivation are shown in Figure 4c,d.

In Figure 4 and later figures, we will plot the surface concentrations of reaction intermediates, $\theta_{\text{int}}$, the amount of product produced per unit proton, $C(t)$, and the rate of change of product produced also normalized per unit proton, $C'(t)$. Figure 4a shows rapid initial proton consumption and overshoot of $C_3^+$ formation. After a relatively short time, $C_3^+$ intermediate formation is overtaken by $iC_4^+$ and $C_7^+$ formation and a steady-state situation evolves. The steady-state alkylate selectivity is mentioned in the legend and equals 62%. Figure 4c illustrates the accumulation of reaction products with time. As the catalyst starts to deactivate, the accumulated product reaches its maximum. In this case, steady-state selectivity and accumulated product selectivity are similar. In the dual interacting and noninteracting proton catalyst models that we will discuss in the section 3, this will not be the case, since the deactivation times of the different products will be different.

Figure 4b,d illustrates the rate of catalyst deactivation that starts at time 40 in the simulations. Then the deactivated proton state $H_0^+$ starts to appear in Figure 4b. The main reaction product is $C_7$ alkylate. The next coproducts are propane $C_3$ and propylene oligomer $C_n^+$, followed by $iC_4^+$ and $C_7^+$ formation.

The lifetime of the catalyst is 200 time units as deduced from the logarithmic plots of product rates of formation, when the slope starts to deviate from zero (see Figure S1a in section SI 1). When the deactivation rate has become exponential, the catalyst deactivation rate becomes equal to 0.0075 time$^{-1}$. In alkylation kinetics, differences in steady-state selectivity multiplied by catalyst lifetime are the determinants that define their productivity.

Whereas the elementary reaction rate constant of propylene oligomerization is 10 times larger than the deprotonation reaction rate constants, in the simulations the respective apparent reaction rates are nearly the same (the factor of 2 difference observed in Figure 4c,d is derived because protons are...
generated by the two deprotonation reaction ($i$C$_4^+$ and C$_7^+$). The reduced relative rate of the oligomerization reaction is because oligomerization has to compete with hydride transfer in the initiation reaction.

Experiments of the alkylation reaction of n-butene and isobutane show, in the regime of high alkylation selectivity, coreaction products due to addition of butenes to the C$_8$ olefins. They can be considered signatures of the deprotonation reaction of the intermediate carbenium ions.

It is interesting to compare the variation in decay time with initial concentrations of protons. Although the protons do not laterally interact, proton concentration affects the selectivity and deactivation rate of the reaction, because reaction intermediates will have different concentrations in the reaction medium. This is illustrated in Figure 5.

In the steady-state regime, respective selectivities are the same but deactivation times are different. This makes the selectivity of accumulated products different. In Figure 5a, the delay of deactivation when surface concentration is reduced is illustrated from the different time dependencies of the surface intermediate. As shown in Figure S1b in section SI 1, a reduction in proton density by half increases the catalyst lifetime time from 200 time steps to 260 time steps. The rate of decline (see Figure 5c) decreases by a factor of 30% (the alkylation deactivation rate is 0.005 time$^{-1}$). When the proton density was reduced by half (Figure 5b) a higher product formation rate per proton occurs. This happens because in the initiation cycle at reduced proton density there is less deactivating propylene oligomerization. In Figure 5d, this is reflected in the increased C$_3^+$ surface concentration when the proton density is reduced.

When the proton density is decreased, two competing phenomena occur. First, the product formation per unit surface area becomes reduced because of less reactive sites. Then the deactivation of H$_1^+$ by readsorption of oligomers ($k_9$, $k_7$, $k_{11}$) produced via the reaction steps $k_3$, $k_5$, and $k_{10}$ also becomes less.

Figure 6 shows the results of kinetic simulations when the reaction rate of propylene oligomerization is large in comparison to that of the hydride transfer reaction rate ($k_1/k_{10} = 0.1$). Now the rate of propylene oligomer formation, C$_n^+$, is high in comparison to that of alkylate C$_7$ production and the rate of deactivation has become substantially faster. Steady-state alkylate selectivity is calculated where the surface coverage of the reaction intermediate is maximum. It has dropped to 12.3%. The catalyst lifetime equals 50 time steps, and the deactivation rate becomes 0.03 time$^{-1}$.

The kinetics simulations apply to a batch reactor and differential conditions where the reactant concentration does not change during the reaction. Experimentally, the preferred reactor for the alkylation reaction is a CSTR, because the reactant concentration in it is uniform and the reactant alkene concentration can be kept low. When the reaction is executed with initial excess isobutane and 100% conversion of reactant alkene, the alkylation selectivity is high and the rate of catalyst deactivation is relatively slow. The catalyst lifetime can be on the order of 10 h or more. Then the reaction rate of hydride transfer is large in comparison to the apparent reaction rate of alkene oligomerization and intermediate carbenium ion deprotonation. The kinetics can be considered similar to that shown in Figure 4.

Of course then, as long as in the experiment conversion of propylene remains 100%, no change in propylene conversion is
observed. However, the rates of alkylate and light alkane production decrease gradually and there is a gradual increase of oligomers due to the deprotonation reaction.\(^{8,10,11}\) Once alkene conversion drops below 100%, then due to the suddenly increased alkene concentration, the relative rate of alkene oligomerization increases sharply, alkylate selectivity drops sharply, and the rate of catalyst deactivation increases.\(^{10,18}\) Now the kinetic simulations of Figure 6, with the increased rate constant of propylene oligomerization (\(k_{10}\)), apply.

### 2.2. Feed-Forward Relation between the Deprotonation and Oligomerization Reactions

As illustrated in Figure 1, catalyst deactivation goes through two deactivation channels. One deactivation channel competes with the initiation reaction cycle. Then propylene oligomerization competes with the hydride transfer reaction of isobutane with propyl cation. The other deactivation channel results from the deprotonation reactions of iC\(_4\)+ and C\(_7\)+. This reaction is a parasite on the propagation reaction cycle. At the same time, it reinitiates the initiation reaction cycle. This can be demonstrated by a simulation where the elementary reaction rate constants of deprotonation \(k_5\) and \(k_6\) are put equal to zero (see Figure 7a).

Figure 7a shows a short transient period of initial propylene oligomer formation, which is rapidly taken over by constant production of alkylate with 100% selectivity. After the initial transient period, the initiation reaction cycle is taken over by the propagation reaction cycle that cannot deactivate because of the absence of the deprotonation reactions.

The feed-forward relation between deprotonation and propylene oligomerization is illustrated in Figure 7b. In this figure, a comparison is made between the two deactivation channels by suppression of the propylene oligomerization reaction. This case is compared with both deactivation channels being operational (default case). This illustrates that when alklylation selectivity is high the dominant deactivation channel is carbenium ion deprotonation. The catalyst lifetime is not significantly affected. It is only reduced by half and the rate of catalyst deactivation is increased by a factor of 2 when both deactivation channels contribute. Alklylation selectivity is slightly higher when propylene oligomerization is suppressed.

### 3. DUAL INTERACTING PROTON CATALYST MODEL

Here we will compare deactivation kinetics according to the dual interacting proton catalyst model of Figure 2 with catalyst deactivation by a dual noninteracting proton catalyst model. This dual noninteracting proton catalyst model contains two kinds of protons of different reactivities as the dual interacting proton model, but these protons have no lateral interaction. They are permanently present during the reaction.

In the dual interacting proton catalyst model, the three proton states H\(_1\)+, H\(_2\)+, and H\(_3\)+ are defined. Proton states H\(_1\)+ and H\(_3\)+ are strongly reactive and weakly reactive, respectively, and proton state H\(_2\)+ is the deactivated proton state. Proton state H\(_1\)+ is converted to proton state H\(_3\)+ when a proton is present on a neighboring site in the deactivated proton state H\(_2\)+. Protons H\(_1\)+ produce alkylate as well as alkene oligomers. Protons H\(_3\)+ only catalyze propylene oligomer formation. In the dual interacting proton model, at the start of the reactions only protons in the state H\(_1\)+ are present.

In contrast to the dual noninteracting proton catalyst model, protons in the reactive proton state H\(_1\)+ are not converted into protons of the weakly reactive proton state H\(_3\)+. Protons in respective proton states H\(_1\)+ and H\(_3\)+ are present from the start of the reaction. Different from the dual interacting proton catalyst model, their deactivated states will not affect the reactivity of protons in either proton state H\(_1\)+ or H\(_3\)+.

The comparison of the two catalyst models is relevant, since it has been suggested\(^{10,11}\) that the experimentally observed delayed production of alkene oligomerization is due to the additional presence of weakly reactive protons that only catalyze alkene oligomerization. This continues after the decline of alkylate production, which is catalyzed by strongly reactive protons.

Here we will show that an increase in oligomer production once alkylate production has declined is only consistent with a laterally interacting proton catalyst model. We will also see that alkylate production has a reduced lifetime when it is catalyzed by the dual interacting proton catalyst model in comparison with the lifetime when isolated protons catalyze the reaction. This is due to a decrease in neighbor deactivated proton sites next to it. In section 4, the effect of surface vacant proton sites on catalyst lifetime will be discussed in more detail.
Figure 8 shows the catalytic reaction cycle of the dual interacting proton catalyst model. Parts of the elementary reactions are similar to those in Figure 3, the catalytic reaction cycle of the single proton case, except that proton state H1+ is converted to the less reactive state H3+ when a deactivated proton state H2+ becomes a neighboring site. Since this conversion has an electronic cause, the corresponding rate constant $k_8$ is fast in comparison to chemical reaction rates. Its default value is chosen as fast such that the kinetics is not affected by its further increase. The proton state H3+ will not initiate the alkylation reaction but only catalyze the oligomerization reaction of propylene. Default parameters of the reaction rate constants of oligomerization by protons in the state H3+ are chosen the same as those of oligomerization by the protons in the state H1+.

The ordinary differential equations that calculate the time dependence of the dual interacting proton catalyst model are solved in a fashion analogous to that for the kinetics equations that correspond to the reaction cycle of the single proton reactivity model of Figure 3. Details are given in section SI 3.

Figure 9 shows the reaction cycles that correspond to the dual noninteracting proton catalyst model. It shows two independent reaction cycles. One reaction cycle is the same as in Figure 3, and the other concerns the oligomerization of propylene only and the corresponding proton deactivation. The solutions of the corresponding kinetics equations are given in section SI 4.

In Figure 10, the deactivation patterns of the dual interacting proton catalyst model and the noninteracting proton catalyst model are given. For the dual interacting proton catalyst model, in Figure 10a, b changes in production formation and rate of product deactivation are shown with time. Default kinetic parameters have been chosen that, for the single proton catalyst model, give high alkylate selectivity.

Different from the single proton catalyst model, now deactivation times of reaction products are different. The selective alkylation production lifetime is now 100 time steps, 0.5 times shorter, and the decline rate of alkylation production 0.065 time$^{-1}$, 8.33 times faster, in comparison to those of the single proton catalyst model. After the decline of selective alkylation production, there is an overshoot of oligomer production. This overshoot of oligomer production is also apparent from Figure 10, which shows the time dependence of production rates of products.

Similar to the case in alkylation batch experiments, this catalyst model will only show high alkylate selectivity when the reaction is stopped in time before oligomer overshoot production sets in. Steady-state selectivities are mentioned in the legends. The steady-state alkylate selectivity is now 56%, which is, as expected, slightly less than that for the single proton catalyst model.

As we will explain in detail in section 4, enhanced deactivation of alkylation production and overshoot of oligomer production derive from the rapid conversion of proton state H1+ to the less reactive proton state H3+ once a deactivated proton state H2+ appears next to it. Protons in proton state H3+ are initially absent.

Figure 8. Catalytic reaction cycle of the alkylation reaction of propylene and isobutane according to the dual interacting proton catalyst model. Conservation of surface species: H1+ + H2+ + H3+ + C3+i + C4+i + C7 = 1.

Figure 9. Dual noninteracting proton catalyst model mechanism. Protons of the two catalyst cycles are not shared, but oligomer production Cn interacts with both reaction cycles. Protons in the state H1+ catalyze alkylation and oligomerization, and protons in the state H3+ only catalyze oligomerization. Conservation of surface species: H1+ + H2+ + iC4+i + C3+i + C7+i = 1/2, H2+ + H3+ + C7+i = 1/2.
in the dual interacting proton catalyst model. Their later formation is the reason for the overshoot in oligomerization production after alkylation decline. In Figure S3, complementary to Figure 10, the time evolution of surface intermediates is shown. The product evolution closely follows the occupation of H1⁺ and H3⁺ protons.

Figure 10c,d shows, for the same elementary reaction rate default parameters, comparable data for the dual noninteracting proton catalyst model. In comparison to the proton concentration in the dual site interacting proton catalyst model, in this figure, the concentrations of respective proton states H1⁺ and H3⁺ have been initialized to cover half the surface each.

As we have explained in Section 2, this reduction in proton density causes the deactivation of alkylate production to be delayed by a factor of 1.3 in comparison to the single proton case (Figure 5c).

The now high initial relative rate of oligomerization production is dramatic in comparison to that in the other two proton catalyst models. It is due to the high oligomerization rate of the H3⁺ protons. In the dual noninteracting proton catalyst model, different from the other catalyst models, oligomerization by the H3⁺ protons does not compete with the hydride transfer reaction. The lifetime of alkylate production is 250 time steps, which is slightly longer than that of the single proton catalyst model of 200 time steps but the lifetime of oligomer production has decreased to 130 time steps. Now, after a decline of oligomerization, alkylate production dominates. In comparison to the single proton catalyst model, the steady-state alkylate selectivity has decreased by 61%. Because of its longer lifetime,
accumulated alkylate production is higher than that of oligomer production.

Kinetics as shown in Figures 10c,d will sensitively depend on the rate of deactivation of the H3+ protons by the oligomer molecules. The oligomerization catalysis by the weakly reactive protons in the proton state H3+, which are not able to catalyze alkylation, should be slower than that of the H1+ protons that are able to catalyze this reaction. Such a decreased rate of deactivation of the less reactive protons H3+ is consistent with their known lower rate of propylene oligomerization. Sarazen et al.⁷ indicate a decrease by at least a factor of 10 in oligomerization rate when the reactivity of highly reactive and less reactive protons in faujasite zeolite are compared.

Figure 10e,f presents oligomer product formation and deactivation rates of deactivation, when accordingly their respective reaction rate constants catalyzed by protons in proton state H1+ (k12, k13) are reduced by half. In comparison to Figure 10c,d, now steady-state alkylate selectivity increases, but oligomer production also dominates at longer times. The lifetime of oligomer production is 260 time steps, and its deactivation rate of 0.01 time−1 is now close to that of alkylation (0.009 time−1). Importantly, different from the dual interacting proton catalyst model, beyond deactivation of alklylation, the slope of oligomerization rate change is always negative.

In the dual noninteracting proton catalyst model, the reactivity of proton states H1+ and H3+ is coupled through deactivation by oligomers that are produced by both proton states. When it is normalized to the same initial H1+ proton concentration, the alkylation lifetime of the dual noninteracting proton catalyst model is reduced by 5% in comparison to the single proton catalyst model.

In this section, we presented simulations with default elementary reaction rate constants that give, for the single proton catalyst model, relatively high alkylate selectivity. In section 4, for the dual interacting proton catalyst model, we will also analyze simulations when alkylate selectivity is low.

This section has provided evidence that alkylate production lifetime is longest when only reactive protons H1+ are initially present and these protons are isolated. The deactivation rate increases by 1 order of magnitude when lateral interactions between protons H1+ are present. Because of the generation of weakly reactive protons H3+, later in time, delayed production of oligomers occurs after a decline of alklylation production.

Different from the dual interacting proton catalyst model, the dual noninteracting proton catalyst model has a lower initial selectivity of alklylate production in comparison to that of oligomer formation. As previously mentioned, this selectivity difference is due to the apparent higher rate of oligomer formation by the also initially present protons H3+ that cannot catalyze alkylate formation.

In the next section, we will analyze the proton dynamics that is fundamental to the kinetic differences induced by the lateral interactions of the protons. The rate of alkylate production follows the dynamics of protons H1+, and the rate of oligomer production follows the dynamics of protons H3+.

4. DISCUSSION OF DEACTIVATION KINETICS

Kinetics modeling, discussed in section 3, has demonstrated that deactivation kinetics changes nonlinearly when protons interact laterally. In the section 4.1, we will analyze the dynamics of a three proton state model that shows deactivation dynamics similar to that discussed above for the dual site interacting proton catalyst model, but without coupling to the full reaction kinetics scheme of Figure 3. The advantage of this three proton state model is that analytical solutions of its dynamics can be found. We will use nonlinear dynamics to deduce the relation between the rates of deactivation of protons in respective H1+ and H3+ states and their rate of interconversion.

In section 4.2, we will return to the full dual interacting proton catalyst model. We will compare stochastic simulations with solutions of the corresponding mean field equations. The stochastic simulations will be used to study deactivation kinetics of laterally interacting protons as a function of proton coordination.

4.1. Dynamics of the Three Proton State Model. In the first part of this section, we will present a mean field analysis of a three proton state dynamics model that simulates deactivation without explicit consideration of the full kinetics of the alkylation reaction. The section will be concluded with a comparison of mean field and stochastic results that includes an analysis of the dependence on proton coordination to other protons.

The mean field ordinary differential equations that describe the time evolution of probabilities ni of the respective proton states H1+ are given by eqs 1a–1c:

\[ \frac{d n_1}{d t} = -k_{12}n_1 - mk_{13}n_1n_2 \]  
(1a)

\[ \frac{d n_2}{d t} = k_{12}n_1 + k_{13}n_3 \]  
(1b)

\[ \frac{d n_3}{d t} = -k_{32}n_3 + mk_{13}n_1n_2 \]  
(1c)

The rate constants k12 and k13 refer to the respective proton deactivation rates of proton states H1+ and H3+. As in the dual site interacting proton catalyst model, proton state H1+ represents the reactive proton state that catalyzes the alkylation reaction and proton state H3+ represents the proton state that only catalyzes propylene oligomerization. The rate constant k13 is the rate of conversion of proton state H1+ to proton state H3+ when it gets as a neighbor the deactivated H1+ proton state. Since the rates of local surface atom rearrangement and electronic changes will be several orders of magnitude faster than the deactivation rates of proton states H1+ and H3+, respectively, one expects k13 to be large in comparison to k12 and k32. m in eq 1a is the number of neighbors of a proton.

If in eq 1a k13 is set equal to zero, eqs 1a and 1c decouple and the probabilities ni correspond to that of the dual noninteracting proton catalyst model. Each proton will decay exponentially with decay constants 1/k12 and 1/k32, respectively.

In the simulations of the dual interacting proton catalyst model shown in Figure 10a,b, we observe a delayed oligomer production after alkylation production has declined. Since this reflects the respective dynamics of protons H1+ and H3+, we are interested to know for which relationship of the rate constants in eq 1a, deactivation of protons in state H3+, that catalyze alkylation, occurs quickly and protons in state H1+ that only catalyze oligomerization remain active after deactivation of the H1+ protons.

The condition for this to happen is that proton state probability n3 crosses state probability n1 (see Figure 11). One can deduce an approximate condition that is given in eq 2 (for the proof, refer to section SI 5.2):

\[ k_{12} + 2mk_{13} > k_{32} \]  
(2)
4.1.1. Stochastic Solution of the Three Proton State Model: Proton Coordination Dependence. In the stochastic simulations, the protons are considered to be located on a lattice as indicated in Figure 12. A proton can have as a neighbor another proton in the same state. In addition, we will consider also the possibility of inert site vacancies that will have state probability $n_4$. Method details on the stochastic simulations are provided in section SI 5.3.

Stochastic simulations and mean field equations give very similar results except for the chemically relevant case of Figure 11 (see section SI 5). We will limit the discussion here to this case, where proton state $H_3^+$ deactivation is delayed beyond deactivation of proton state $H_1^+$. As we will see, the difference between the mean field and stochastic simulations is due to pattern formation of the respective proton states in the latter. When the mean field approximation is used, as in section 4.1, the implicit assumption is made that the distribution of proton states is uniform. Stochastic simulations make the conversion of proton states $H_1^+$ to proton states $H_3^+$ faster. The decay of proton state probability $n_1$ is faster, and there is more delay of proton state probability $n_3$.

Here, we will show this for the three proton state model. In section 4.2, the consequences of this pattern formation will be investigated. Then, in simulations, proton dynamics and kinetics are coupled.

Mean field and stochastic simulations of the respective proton states are shown in Figure 13. The dashed lines show the time evolution for the stochastic case. The solid lines give the corresponding mean field calculated values. We compare a surface without vacancies ($n_4 = 0$) with two surface configurations in which vacancies are present ($n_4 = 0.3, n_4 = 0.7$).

Significant differences between mean field solutions and stochastic simulations can be observed in the simulations of Figure 13. For convenience of comparison, they have been done with same parameter values as used in Figure 11. Stochastic simulations slow the rates of deactivation. The differences in the fast decay time of proton state $H_1^+$ and the delayed deactivation time of proton state $H_3^+$ decrease.

When vacancy concentration $n_4$ increases, the decreased coupling of proton states $H_1^+$ and $H_3^+$ causes the decay rate of proton states $H_1^+$ and deactivation delay of proton states $H_3^+$ to become less. This confirms the previous conclusion that the alkylation deactivation rate becomes less when protons become

According to this relation, when the conversion rate of proton state $H_1^+$ to proton state $H_3^+$ is fast, after decay of alkylation, oligomer formation still increases and it decays later. This delay becomes independent of $k_{13}$ when it exceeds a maximum value.

When the intrinsic rate of deactivation of proton in state $H_1^+$ ($k_{12}$) is small, a high rate of conversion of proton state $H_1^+$ to proton state $H_3^+$ (strong coupling between the protons; $2mk_{13}$) is necessary to overcome the rate of deactivation of the protons in state $H_3^+$ ($k_{12}$). A low rate of proton state $H_1^+$ deactivation implies dynamics that corresponds to high initial alkylate selectivity. As Figure 10c illustrates, when proton states are decoupled, the rate of deactivation of proton $H_1^+$ (reflected in the deactivation rate of oligomerization) is greater than the deactivation rate of proton state $H_3^+$ (reflected in the deactivation rate of alkylation). The proton dynamics that corresponds to the kinetics of dual interacting proton catalyst model kinetics of Figure 10b and dual noninteracting proton catalyst model of Figure 10d, both calculated with the same default parameters, is shown in Figure 11.

Figure 11 shows the proton dynamics of the interacting three proton catalyst model, where the condition of eq 2 is satisfied. A comparison is made with the independent proton dynamics.

The rate constant of proton state $H_1^+$ deactivation has been chosen to be slower than that of proton state $H_3^+$. The strong coupling between proton states causes, in the coupled system, deactivation of proton state $H_1^+$ to be faster than that of proton state $H_3^+$. Slow exponential decay of proton state $H_1^+$ is converted into fast nonexponential accelerated decay. This is even the case when the deactivation rate constant of reactive proton state $H_1^+$ ($k_{12} = 0.01$) is initially slower than that of the less reactive proton state $H_1^+$ ($k_{12} = 1$).

For a detailed nonlinear dynamics analysis of eqs 1a, where four different deactivation rate regimes are identified, we refer to section SI 5. It appears that only one deactivation rate regime shows dynamics as found in Figure 11.

Before returning to the full kinetics of the alkylation reaction, we present stochastic simulations of the dynamics of the three proton state model and compare these with the mean field solutions.

Figure 11. Dynamics of the three proton state model. Rate parameter values: $k_{12} = 0.01$, $k_{13} = 1.0$. Comparison of time evolution of proton state probabilities $n_i$ of protons $H_i^+$, respectively, with strong coupling of the protons ($k_{13} = 10.0$, solid lines) and absence of coupling of the protons ($k_{13} = 0$, dashed lines). The time evolution of respective proton states $H_1^+$, $H_2^+$, and $H_3^+$ are simulated with rate constants such that state probability $n_3$ crosses state probability $n_1$. This is, within dual interacting proton catalyst model, the condition of delayed oligomer production (see Figure 10b).

Figure 12. 2D lattice representation of the laterally interacting three proton catalyst model.
isolated. According to the stochastic simulations, $H_1^+$ sites already behave as if they are completely isolated when the vacancy concentration $n_4$ equals 0.5 (see Figure S13 and section SI 5.3 for more details).

Pattern formations of the respective proton states are responsible for this isolated site behavior. As Figure 14 illustrates, the distribution of proton state probability $n_1$ is no longer homogeneous around proton state probability $n_2$. Proton state $H_3^+$ regions grow around deactivated proton state $H_2^+$ regions. Since the deactivation rate of proton state $H_1^+$ is slower than that of proton state $H_3^+$, the system has to wait for the transient proton state $H_4^+$ layer that surrounds the deactivated proton state $H_2^+$ regions to decay. After decay, the proton state $H_3^+$ generation process is restarted. Due to island formation, percolation of state probabilities $H_2^+$ has become reduced. This occurs when $n_4$ is 0.5, since this is near the percolation threshold^42 of a hexagonal lattice.

4.2. Stochastic Simulations of the Dual Interacting Proton Catalyst Model. Here we extend the analysis of section 4.1 with simulations for the dual interacting proton catalyst model as a function of surface vacancy concentration. The full kinetics according to the reaction cycle of Figure 8 is coupled to stochastic dynamics of the protons (for method details see section SI 5.4).

Stochastic simulations differ again from mean field simulations. However, remarkably, when the alkylation
selectivity is high (Figure 15), in contrast to the three proton state model stochastic simulations show now enhanced rates of deactivation and short catalyst lifetimes. Analogous to the three proton state model, delay of deactivation of oligomerization when alkylate production has deactivated becomes less.

When alkylation selectivity is low (Figure 17), one finds that the deactivation rate is decreased in comparison to the mean field simulation with longer catalyst lifetimes. The differences between mean field and stochastic simulations relate to proton mobility limitations caused by surface overlayer patterns of adsorbed reaction intermediates.

The stochastic simulations of Figure 15a,b show a shorter catalyst lifetime of alkylate and oligomerization production in comparison to the mean field simulations. As was mentioned, this is due to pattern formation of reaction intermediates adsorbed on the surface lattice. Differently from the three proton state model, the surface becomes now also occupied by C3+, iC4+, and C7+ cations (see Figure 16a). Once a deactivated proton state H2+ is generated, it will have a reduced probability to meet a proton in state H1+. This will reduce the initial rate of proton state H3+ generation. Because proton state H3+ dominates the oligomer removal rate, oligomer production now is increased. This intermediate oligomer concentration then causes quicker deactivation of state H1+. Similar to the dual noninteracting proton catalyst model, the presence of slowly deactivating H3+...
protons increased the deactivation rate of the H$_1^+$ protons (compare Figure 10d with Figure 10f).

With an increase in surface vacancies (Figure 15c−f), the mean field and stochastic simulations show an increasingly longer alkylation lifetime and slower deactivation time. In the stochastic simulations, the relative lifetime of alkylation increases more than in the mean field simulations, but the relative increase in lifetime of oligomerization is less, as is the case for the respective deactivation rates. The initial steady state selectivity of alkylation does not change. When $n_3 = 0.7$, protons in stochastic simulations start to behave as isolated protons. This is due to the loss of H$_2^+$ percolation, as also discussed in section 4.1.1 (for more details, refer to section SI 6). Then in the mean field simulation, the delay in peak oligomer formation is still 40 time steps.

Figure 17 shows that, when steady state alkylation selectivity is fast, oligomerization product formation always dominates. A double peak in oligomer production is observed. The first peak is
due to the $H_1^+$ protons that rapidly deactivate, and the latter maximum is from oligomer production by the proton in the $H_1^+$ state.

Different from what was observed in Figure 15, the stochastic simulations in Figure 17 now show a delay in the rate of deactivation of oligomer production by the $H_1^+$ proton states. The deactivation rate of proton state $H_1^+$ by $C_n^-$ is so rapid that the meeting probability with proton state $H_1^+$ dominates. A representative surface overlayer intermediate pattern is shown in Figure 16b. Because oligomer concentration is already high, a further change in oligomer concentration will have no significant effect. In the stochastic simulations, delayed formation of proton state $H_1^+$ is now the cause of the slower rate of decay.

5. CONCLUSION

In this paper, we have presented kinetic simulations of deactivation rates of the alkylation reaction of propylene with isobutane catalyzed by a variety of surface models of the solid acid catalyst. The question is addressed of how lateral interactions between protons will change catalyst stability and product distribution as a function of time.

Catalyst deactivation by protons that laterally interact has been computationally studied with a dual interacting proton catalyst model. In the dual interacting proton catalyst model initially only reactive protons are present that selectively produce alkylate. In time, they become deactivated. This induces still present reactive protons to convert to less reactive protons that now only catalyze propylene oligomerization. The latter reaction also deactivates in time.

The existence of protons selective to alkylation, different from protons that only oligomerize alkene oligomerization, is consistent with experiments and theory. Strongly acidic protons promote alkylation production over alkene oligomerization. This is because reactive protons favor the hydride transfer reaction between isobutane and carbenium ion versus the propylene oligomerization reaction. Catalyst deactivation occurs through two deactivation reaction channels that do not operate independently of each other.

Deprotonation of reaction intermediate carbenium ions will initiate catalyst deactivation by generation of alkenes. This leads to deactivation because of consecutive deactivating oligomerization reactions. This deactivation reaction channel is mainly operational at high alkylate selectivity.

This second deactivation reaction channel, which dominates when alkylate selectivity is low, is reactant alkene oligomerization. This reaction competes with the hydride transfer reaction that initiates the alkylation reaction.

The presence of these two deactivation channels is consistent with the two experimental selectivity regimes of high and low alkylate selectivity observed in the CSTR experiment. Selective alkylation catalyzed by solid acid catalysts has to be done with high isobutane to alkene ratio under integral conditions in a CSTR that enables initially 100% conversion of reactant alkenes. This minimizes alkene concentration and provides a uniform distribution of reactant and product in the reactor. Then alkylate selectivity is high. There is initially no observable catalyst deactivation. Catalyst deactivation becomes observable once catalyst protons have become deactivated such that the conversion of reactant alkene drops below 100%.

Experimentally the change in product selectivity as a function of reaction time can be followed also in the 100% propylene conversion regime. In this reaction regime where alkylation selectivity is high, a gradual decline in alkylate and alkane production rate is observed. In addition, a steady production rate increase of oligomer molecules happens that is derived from deprotonated carbenium ions. This shows that when alkylate selectivity is high the main deactivation is due to the deprotonation reactions of intermediate carbenium ions.

Because of ongoing deactivation in the experiment, conversion of alkene will at some point decrease to less than 100%. Due to the increased propylene reactant concentration, the rate of the propylene oligomerization reaction increases and the hydride transfer reaction rate will no longer be able to compete with it. Then, selective alkylate production decreases sharply and reactant alkene oligomers become the main product.

These observations agree with kinetics simulations of the deactivation of the alkylation reaction. When alkylate selectivity is high, the main deactivation is due to disruption of the propagation reaction cycle through deprotonation of intermediate carbenium ions. When oligomer production dominates, the deprotonation reaction becomes replaced by deactivating alkene oligomerization as the main cause of catalyst deactivation.

Kinetics simulations with the dual interacting proton catalyst model show that lateral interactions strongly and negatively affect the lifetime of the alkylation catalyst. Additionally, it is found that, due to the increased rate of alkylate deactivation, oligomer production continues after deactivation of alkylate production.

Whereas, as a function of time, the CSTR experiment shows initially high alkylation activity and delayed oligomer production when alkylate selectivity declines, this is no indication that lateral interactions as discussed in this paper play a role. In the experiment that initially converts alkene 100%, the decline of alkylate production followed by an increase of oligomer production is the result of the sudden increase in alkene concentration. This happens when alkene conversion decreases to less than 100% due to gradual loss of reactive protons.

In the simulations, we have compared kinetics of the dual interacting proton catalyst model with two other catalyst models: a single proton catalyst model and a dual catalyst noninteracting proton catalyst model. In the absence of lateral interactions, it is found that an increased local proton concentration will also increase catalyst deactivation rate. It increases locally the concentration of alkene intermediates that rapidly deactivate the catalyst. Therefore, a reduced surface concentration of protons is beneficial to catalyst lifetime. Experiments by Mores et al. on the deactivation of the MTO reaction report a related proton concentration effect that extends catalyst lifetime. In reference to a study by Schüßler et al., which deals with the alkylation reaction, it is also reported that a decreased proton concentration enhances the lifetime of the catalyst.

A comparison of the dual interacting proton catalyst model and a dual catalyst noninteracting catalyst model shows that, even in the absence of lateral interactions, the additional presence of protons that only catalyze propylene oligomerization has a large negative effect on catalyst selectivity and also reduces catalyst life. When alkylation is ongoing, oligomer production by protons that are not selective to alkylate production is much faster than that by the alkylate-producing protons. Experimental results confirm the conclusion that the additional presence of weakly reactive sites decreases catalyst life.

Nonlinear dynamics analysis suggests that the deactivation kinetics of laterally interacting protons as simulated with the
dual interacting proton catalyst model is not necessarily specific to the alkylation reaction. It may occur in reactions catalyzed by laterally interacting protons that require highly reactive protons for selective production of a desirable product. Examples of such reactions in addition to the alkylation reaction are the MTO reaction or the catalytic cracking reaction. Interestingly, this implies that the effect is expected to be absent in hydrocracking or hydroisomerisation reactions, which are less sensitive to proton reactivity since conversions of intermediate olefins are then reaction rate controlling.20

Stochastic simulations have been compared with mean field simulations. Stochastic simulations find differences with mean field simulations since adsorbed reaction intermediates are not homogeneously distributed on the working catalyst. When these simulations are applied to the dual interacting proton catalyst model, one finds that coverage of the surface by reaction intermediates inhibits additional conversion of reactive protons into weakly reactive protons. Because of the large difference in overall oligomer production rates between reactive protons, which catalyze alkylation, and weakly reactive protons, which only catalyze oligomerization, this affects a selective alkylation catalyst differently from a catalyst that dominantly produces oligomers. In comparison to mean field simulations, it reduces the lifetime of a selective alkylation catalyst but makes it slightly longer when alkylation selectivity is low. The stochastic simulations show that laterally interacting protons behave kinetically as isolated protons, when on a hexagonal lattice the proton vacancy concentration is in excess of 50%.

**ASSOCIATED CONTENT**

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acscatal.8b01511. Mean-field solutions of kinetic equations for single reactivity and dual interacting and dual noninteracting proton catalyst models, three proton state model with stochastic simulations, and stochastic simulations of the dual interacting proton catalyst model (PDF)

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Notes

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