Solvothermal synthesis of octahedral and magnetic CoFe$_2$O$_4$–reduced graphene oxide hybrids and their photo-Fenton-like behavior under visible-light irradiation†

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We report a facile solvothermal synthesis of novel octahedral CoFe$_2$O$_4$–reduced graphene oxide (RGO) hybrid and pure CoFe$_2$O$_4$ that were used as heterogeneous photo-Fenton catalysts for the degradation of organic dyes in water. We investigated the structures, morphologies and catalytic activity of both the CoFe$_2$O$_4$ nanoparticles and CoFe$_2$O$_4$–RGO hybrids. The morphology of CoFe$_2$O$_4$ nanoparticles displays size-dependent shapes changing from granular (or sheet) to octahedral shapes with the introduction of RGO. Compared with bare CoFe$_2$O$_4$, the octahedral CoFe$_2$O$_4$–RGO hybrids serve as novel bifunctional materials displaying higher saturation magnetization values and excellent heterogeneous activation of H$_2$O$_2$ at nearly neutral pH. The high saturation magnetization (41.98 emu g$^{-1}$) of CoFe$_2$O$_4$–RGO hybrids aids their separation from the reaction mixture. In addition, the remarkable enhancement in the photo-Fenton activity of the CoFe$_2$O$_4$–RGO hybrids under visible light irradiation was attributed to the graphene/CoFe$_2$O$_4$ heterojunction, which aided the separation of excited electrons and holes. Furthermore, the CoFe$_2$O$_4$–RGO hybrids exhibited better removal efficiency for cationic methylene blue (MB) dye than for anionic methyl orange (MO) dye. Meanwhile, the CoFe$_2$O$_4$–RGO hybrids displayed acceptable photocatalytic stability, and we proposed an activation mechanism of H$_2$O$_2$ by the octahedral CoFe$_2$O$_4$–RGO hybrids.

1. Introduction

It is well-known that the inappropriate disposal of organic effluents from industries raises environmental concerns. Among many methods used for the treatment of industrial effluents, graphene-based adsorption is a highly efficient and less time-consuming method owing to its exceptionally large surface area, open porous structure, and excellent chemical stability. However, the separation and recycling of graphene used for purification of industrial wastewater remain a challenge, and this limits its further application.

Nowadays, the integration of magnetic particle technology to solve water pollution problems has attracted the interest of many researchers. Firstly, the dispersion of magnetic nanoparticles (NPs) on graphene sheets enables the development of materials with promising separability and reusability. This is because the graphene-based hybrids containing magnetic NPs do not introduce foreign pollutants into the wastewater system during the whole purification process. Furthermore, the magnetic NPs can enhance the functionalities of graphene-based hybrids; therefore, the potential of magnetic NPs–graphene hybrids is huge, and they can be widely applied in catalysis and water purification. In view of these, several studies have been conducted to investigate the adsorption and photocatalytic performance of graphene-based hybrids containing magnetic NPs. Yao et al. fabricated magnetic Fe$_3$O$_4$@graphene hybrids. The methylene blue (MB) and Congo red (CR) absorption capacities of these hybrids good and separation of the hybrids after absorption was easy. Fu et al. synthesized a set of graphene-based nanocomposite photocatalysts including MFe$_2$O$_4$–graphene (M = Zn, Ni, Mn, Co) hybrids. These photocatalysts show high visible-light-responsive photo-degradation efficiency towards dyes. They were also stable and the photocatalyst can be conveniently separated by the use of an external magnetic field.

Among these magnetic NPs, cobalt ferrite (CoFe$_2$O$_4$) nanocrystals with a spinel structure have attracted a great deal of attention for environmental remediation owing to their outstanding properties, such as their size- and shape-dependent magnetic behavior, and their relatively lower bandgap that perfectly matches with the solar spectrum. To date, several studies have been conducted using CoFe$_2$O$_4$ as...
heterogeneous catalysts for activating hydrogen peroxide\textsuperscript{19,20} or peroxymonosulfate (PMS). \textsuperscript{21–24} When heterogeneous PMS is activated for the degradation of dyes, the reaction mechanism is represented by the following equations:\textsuperscript{19,21–26}

\[
\begin{align*}
\text{CoFe}_2\text{O}_4 + h\nu &\rightarrow \text{CoFe}_2\text{O}_4 (h\nu^+ + e_{\text{CB}}^-) \quad (1) \\
\text{CoFe}_2\text{O}_4 (e_{\text{CB}}^-) + \text{HSO}_4^- &\rightarrow \text{CoFe}_2\text{O}_4 + \text{OH}^- + \text{SO}_4^{2-} \quad (2) \\
\text{CoFe}_2\text{O}_4 (h\nu^+) + \text{OH}^- &\rightarrow \text{CoFe}_2\text{O}_4 + \text{HO}^- \quad (3) \\
\text{Co}^{2+} + \text{HSO}_4^- &\rightarrow \text{Co}^{3+} + \text{SO}_4^{2-} + \text{OH}^- \quad (4) \\
\text{Fe}^{3+} + h\nu + \text{OH}^- &\rightarrow \text{Fe}^{2+} + \text{HO}^- \quad (5) \\
\text{Fe}^{2+} + \text{HSO}_4^- &\rightarrow \text{Fe}^{3+} + \text{SO}_4^{2-} + \text{OH}^- \quad (6) \\
\text{Dye} + \text{SO}_4^{2-} \text{ and } \text{HO}^- &\rightarrow \text{intermediates} \rightarrow \text{CO}_2 + \text{H}_2\text{O} \quad (7)
\end{align*}
\]

Interestingly, CoFe\textsubscript{2}O\textsubscript{4} possesses obscure Co leaching properties owning to its strong Fe–Co linkages. Moreover, Fe helps to enrich the hydroxyl group content on the surface of Fe–Co catalysts, and CoFe\textsubscript{2}O\textsubscript{4} are magnetically recyclable owing to the unique ferromagnetic properties. A number of studies have shown that when immobilization of CoFe\textsubscript{2}O\textsubscript{4} nanoparticles onto graphene sheets, the Fenton-like catalysis efficiency of the resultant hybrid catalyst improved remarkably. Moitra \textit{et al.}\textsuperscript{19} used CoFe\textsubscript{2}O\textsubscript{4}–reduced graphene oxide (RGO) nanocomposites for heterogeneously activating H\textsubscript{2}O\textsubscript{2}. The CoFe\textsubscript{2}O\textsubscript{4}–RGO exhibited good visible-light-responsive photocatalytic properties for the degradation of various synthetic dyes. \textit{Lu}\textsuperscript{27} and \textit{He}\textsuperscript{27} \textit{et al.} have studied the effects of RGO contents, initial solution pH, H\textsubscript{2}O\textsubscript{2} concentration and operation parameters (\textit{e.g.} dosage and dye concentration) on the catalytic activity. Their results showed that the incorporation of the graphene significantly enhanced the photocatalytic performance of the CoFe\textsubscript{2}O\textsubscript{4} in which the graphene acted as a charge carrier to capture the delocalized electrons. In addition, the CoFe\textsubscript{2}O\textsubscript{4}–graphene hybrids exhibited good reusability and were easily separated from the reaction mixture. To the best of our knowledge, the effect of particle size and shape of CoFe\textsubscript{2}O\textsubscript{4} on catalytic activity and magnetic properties of CoFe\textsubscript{2}O\textsubscript{4}/graphene nanocomposites, has been rarely studied. Moreover, there are few reports on the relationship between the adsorption behavior and surface charge of CoFe\textsubscript{2}O\textsubscript{4}/graphene nanocomposites. Furthermore, very few reports about the underlying growth mechanism of octahedral CoFe\textsubscript{2}O\textsubscript{4} on graphene sheets are available.

In the present study, we synthesized octahedral CoFe\textsubscript{2}O\textsubscript{4}–RGO hybrids \textit{via} interface engineering. These hybrids served as excellent heterogeneous catalysts for activating H\textsubscript{2}O\textsubscript{2} at nearly neutral pH. In addition, the catalytic performance of CoFe\textsubscript{2}O\textsubscript{4}–RGO hybrids was evaluated for the degradation of diverse dyes, \textit{i.e.} MB and methyl orange (MO), under visible-light irradiation. Also, we have presented the underlying photo-Fenton-like behavior and have investigated the remarkable effect of particle size and shape of CoFe\textsubscript{2}O\textsubscript{4} NPs in CoFe\textsubscript{2}O\textsubscript{4}/RGO hybrids to promote the Fenton chemistry and tuning of the magnetic property.

2. Experimental

2.1. Chemicals and reagents

Cobaltous nitrate hexahydrate (Co(NO\textsubscript{3})\textsubscript{2}·6H\textsubscript{2}O, 99%), ferric nitrate nonahydrate (Fe(NO\textsubscript{3})\textsubscript{3}·9H\textsubscript{2}O, 99.9%), isopropanol (IPA, AR), methylene blue (MB, AR), and methyl orange (MO, AR) were obtained from Sinopharm chemical reagent Co., Ltd (Shanghai, China). Hydrogen peroxide (H\textsubscript{2}O\textsubscript{2}, 30\%), ethanol, and sodium hydroxide (NaOH, AR) were purchased from Xilong Chemical Co., Ltd (Shantou, China). A graphene suspension (thickness, ~0.8 nm; 1 mg mL\textsuperscript{-1}) was supplied by XFNANO Materials Tech Co. Ltd (Nanjing, China). Deionized water was supplied by a Milli-Q ultrafiltration system (Billericia, USA).

2.2. Synthesis of materials

The reduced graphene oxide loading in the CoFe\textsubscript{2}O\textsubscript{4}–RGO hybrids was adjusted by changing the initial volume (x) of graphene suspension. And the sample was denoted as CoFe\textsubscript{2}O\textsubscript{4}–RGO(x) (where x = 35 and 70 mL). As a typical procedure, 35 mL of the graphene suspension was dissolved in 265 mL of ethanol while it constantly stirred using a magnetic stirrer, and it was then sonicated for 30 min. Secondly, 4.3654 g of cobaltous nitrate hexahydrate and 12.1164 g of ferric nitrate nonahydrate were directly added to the above graphene solution while it was vigorously stirred for 30 min. The concentration of Co\textsuperscript{2+} and Fe\textsuperscript{3+} in the resulting solution (marked as solution A) was 0.05 mol L\textsuperscript{-1} and 0.1 mol L\textsuperscript{-1}, respectively. Subsequently, a 4 M NaOH solution was added dropwise to a uniform solution A until a pH of ~13 was attained, and this resulted in complete precipitation of the metal cations. The resultant mixture was then transferred into four Teflon-lined stainless-steel autoclaves (volume: 100 mL, degree of filling: ~80\%), and heated in an oven at 180 °C for 20 h. The sample was allowed to naturally cool down to room temperature. The products were centrifuged and directly collected, then they were several washed with deionized water and ethanol and dried at 60 °C for 24 h. For comparison, pure CoFe\textsubscript{2}O\textsubscript{4} NPs were also fabricated following a similar solvothermal process without the introduction of graphene.

2.3. Characterization

The X-ray diffraction (XRD) diffractograms were obtained using a D/Max-2550 diffractometer with Cu Kz radiation source (\(\lambda = 0.15406\) nm). The Raman spectra were obtained using a HORIBA JY LabRAM HR evolution microspectrometer at an excitation wavelength of 532 nm. The Fourier transform infrared spectroscopy (FT-IR) spectra were recorded using a Spectrum Two FT-IR spectrophotometer with samples pressed in KBr. The exact amount of RGO in the as-obtained products was measured using a thermal gravimetric analyzer (TGA/DSC 3+, Mettler-toledo, Switzerland). The UV-visible diffuse reflectance spectra (DRS) were obtained using a UV-vis spectrophotometer (UV-2550, Shimadzu, Japan). The surface morphologies of products were observed using a scanning electron microscope (GeminiSEM 500, Zeiss, Germany). The X-ray photoelectron spectroscopy (XPS) was carried out using a ThermoFisher
Scientific ESCALAB250Xi photoelectron spectrometer using with Al X-ray source. All binding energy values were calibrated using the carbon (C 1s) peak observed at 284.0 eV. The specific surface area and the pore-size distribution were measured using Autosorb-iQ2 (Quantachrome, USA) based on the nitrogen adsorption–desorption isotherms. The hysteresis loops were determined using a vibrating sample magnetometer (PPMS DynaCool Cryogen-free System 9 T, Quantum Design, USA). The zeta potential measurements were conducted on a zeta potential analyzer JS94K2 (Shanghai, China).

2.4. Catalytic test procedure

The photocatalytic activities of the as-obtained samples were measured using the photocatalytic degradation of water-soluble dyes (MO or MB) under visible light irradiation together with H2O2. A 500 W xenon lamp (cut-off wavelength 420 nm) was used as the visible-light source in the photo-Fenton reaction. The distance between the surface of the reaction solution and the visible-light source was approximately 11 cm. The experiment was conducted as follows: 60 mg of the photocatalyst was dispersed into 60 mL of the 10 mg mL−1 dye aqueous solution (MO or MB) while it was constantly stirred. Before irradiation, the solution was stirred using a magnetic stirrer in a dark room for 1.5 h to ensure the establishment of an adsorption/desorption equilibrium. Next, a certain amount of 30% H2O2 was introduced into the above resultant mixture, which was then exposed to the visible-light source while the mixture was continuously stirred. The concentrations of MO and MB in the reacting mixtures were monitored by measuring the characteristic absorbance at wavelengths of 464 and 664 nm, respectively.

The removal rate (rev (%)) of the dyes (MO and MB) was determined using eqn (8):

\[
\text{Removal rate (\%)} = \left( \frac{A_0 - A_t}{A_0} \right) \times 100\% = \left( \frac{C_0 - C_t}{C_0} \right) \times 100\%
\]

where \(A_t\) is the absorbance of dye solution at \(t\) time, \(A_0\) is the initial absorbance of dye solution, \(C_0\) is the initial concentration of aqueous dye, and \(C_t\) represents the concentration of the aqueous dye at \(t\) time.

A trapping experiment, which is similar to the photodegradation procedure, was conducted to identify the dominant reactive species upon the addition of 0.1 M IPA. Moreover, the photocatalytic stability of CoFe2O4–RGO hybrids was investigated via recycling tests. After each cycle, the catalyst was separated from the reaction solution using an external magnet, and it was washed several times with deionized water and ethanol. Then the same volume of fresh MB solution was added for the next cycle.

3. Results and discussion

3.1. Characterization and the synthesis mechanism

The XRD diffractograms of the bare CoFe2O4 and CoFe2O4–RGO hybrids are shown in Fig. 1(a). The XRD diffractograms of CoFe2O4–RGO hybrids are identical to that of the bare CoFe2O4, which matches well with cubic spinel-type CoFe2O4 (JCPDS 22-1086). Moreover, the observed diffraction peaks were sharp, implying that the as-prepared CoFe2O4 particles were highly crystalline. Peaks attributed to impurities were not observed on the XRD diffractogram, indicating that the CoFe2O4 samples were highly pure. In addition, the characteristic peak that was observed at 2θ = 26.5 corresponding to RGO (002) was not detected in the XRD diffractogram of the CoFe2O4–RGO hybrids. This may be ascribed to the proper separation and decoration of the RGO sheets by excess CoFe2O4 nanoclusters, resulting in a low diffraction intensity of graphene in the XRD diffractogram of the CoFe2O4–RGO hybrids.

Fig. 1(b) shows the Raman spectra of bare CoFe2O4 and the CoFe2O4–RGO hybrids. Three major Raman scattering peaks can be observed at about 312, 468, and 687 cm−1, which are attributed to the specific characteristics \(A_{1g}, E_{2g}\), and \(3T_{2g}\) of CoFe2O4 ferrite. CoFe2O4–RGO hybrids and pristine CoFe2O4 show similar peaks in the range of 250 to 750 cm−1. In addition, the CoFe2O4–RGO hybrids exhibit two Raman peaks at approximately 1347 (D band) and 1589 cm−1 (G band), where the D and G bands correspond to the RGO. These results indicate that graphene has been incorporated into the CoFe2O4–RGO hybrids.

The hybridization of the CoFe2O4 NPs with RGO was further verified by observing the FT-IR spectra (shown in Fig. 1(c)). For the CoFe2O4 sample, the peaks at approximately 3446 cm−1 and 1628 cm−1 were respectively ascribed to the stretching and bending vibrations of O–H in the presence of water molecules absorbed on the surface. In addition, tetrahedral Co–O bonds were observed at approximately 593 cm−1, and the octahedral Fe–O stretching vibrations were observed at approximately 458 cm−1. When compared with the bare CoFe2O4, some new peaks were observed in the CoFe2O4–RGO hybrids. The band at approximate 1630 cm−1 may correspond to the C=C skeletal vibration of graphitic domains of RGO. Furthermore, the peaks observed at 1398 and 1594 cm−1 can be attributed to the deformation vibrations of the tertiary C–OH groups and the unoxidized carbon backbone on the surface of RGO, respectively. Therefore, we can infer that the target CoFe2O4–RGO hybrids had been synthesized.

TG–DSC analysis was carried out under air atmosphere to confirm the amount of RGO present in the CoFe2O4–RGO hybrids (Fig. 1(d)–(f)). For pure CoFe2O4, one obvious weight loss region was observed at 90–110 °C, which was attributed to the desorption of adsorbed water. Meanwhile, the heat absorption peak at 100 °C can be observed on the DSC curve of the pure CoFe2O4 sample. Furthermore, another obvious heat absorption region was observed at 800–900 °C, which may be ascribed to intrinsic magnetic transformation of CoFe2O4. But this variation of magnetic structure does not lead to obvious weight loss. However, two obvious weight loss steps were observed at 90–110 °C and 215–460 °C for the CoFe2O4–RGO hybrids, which were ascribed to the loss of adsorbed water and the burnt-out of RGO, respectively. On the basis of the present TG–DSC analysis results, the RGO content in the CoFe2O4–RGO(35) and CoFe2O4–RGO(70) samples was ~1.97 and 3.80 wt%, respectively. Moreover, the UV-vis DRS of pure CoFe2O4 and CoFe2O4–RGO are shown in Fig. S1.† The pure
CoFe$_2$O$_4$ and CoFe$_2$O$_4$–RGO hybrids all exhibit strong absorption in the range of 350–800 nm, which may be due to their black color and narrow bandgap, implying that these materials can act as excellent visible-light-responsive photocatalysts.

To investigate the morphology and composition of the as-fabricated products, the field emission scanning electron microscope (FESEM) and the elemental mapping of the pure CoFe$_2$O$_4$ and CoFe$_2$O$_4$–RGO hybrids were obtained. Two morphological features were observed on the FESEM of the bare CoFe$_2$O$_4$ NPs, i.e. granular and sheet-like, and the individual pure CoFe$_2$O$_4$ NPs were withing a size range of 4–98 nm; however, aggregation of some granular particles occurred, and this resulted in the formation of a large amount of void space (Fig. 2(a)).

Furthermore, the SEM micrographs of CoFe$_2$O$_4$–RGO hybrids contained three distinct morphological features: (i) intimate interfacial connections exist between the individual CoFe$_2$O$_4$ and the RGO nanosheets, (ii) the CoFe$_2$O$_4$ particles were mainly composed of regular octahedrons and irregular subunit particles with a broad size distribution. The edge lengths of octahedrons are in the range of 60 to 480 nm, while the sizes of the irregular subunit particles are in the range of 10 to 40 nm, (iii) part of the irregular subunit particles were dispersed on the octahedral CoFe$_2$O$_4$ and the wrinkled

Fig. 1  The crystal structure of the as-synthesized pure CoFe$_2$O$_4$ and the CoFe$_2$O$_4$–RGO hybrids with different amounts of graphene: (a) XRD diffractograms, (b) Raman spectra, (c) FTIR spectra, and (d–f) TG–DSC curves.
graphene planes, the other part of subunit particles agglomerated to form nanoclusters (Fig. 2(b, h, and i)). Compared with the pristine CoFe₂O₄, the average particle size of CoFe₂O₄ in the CoFe₂O₄–RGO hybrids was larger than that of pure CoFe₂O₄ owing to the introduction of graphene sheets. Also, the morphologies of the CoFe₂O₄ particles in the CoFe₂O₄–RGO hybrids were different from that of the pristine CoFe₂O₄, i.e. it contained octahedrons and irregular subunit particles. This indicates that the hierarchical flexible graphene sheets facilitate the quick growth and Ostwald Ripening of the CoFe₂O₄ nanocrystals. In addition, the corresponding EDX mapping micrographs of CoFe₂O₄–RGO(35) hybrid clearly showed that Co, Fe, and O elements were uniformly distributed in the CoFe₂O₄ particles, implying the formation of bimetallic oxides but not the formation of simple mixtures of the two corresponding single-metal oxides. To further understand the synergistic effects of the CoFe₂O₄–RGO hybrid in photo-Fenton behavior, XPS was used to determine the chemical state of the corresponding transition metal and oxygen in the samples. From the survey spectrum in Fig. 3(a), Fe, Co, O, and C were present in the sample, and no other impurity was detected in the CoFe₂O₄–RGO hybrid. The high-resolution Fe 2p spectra of the pure CoFe₂O₄ and CoFe₂O₄–RGO(35) are shown in Fig. 3(b) and (c). For the CoFe₂O₄–RGO(35) sample, the doublets binding energies of Fe 2p₃/2 at 710.3 eV and Fe 2p₁/₂ at 723.3 eV are due to the contributions from Fe³⁺ ions in the octahedral sites (A-site) of the spinel-type CoFe₂O₄, while the doublets binding energies of Fe 2p₃/2 at 712.7 eV and Fe 2p₁/₂ at 724.9 eV are due to the contributions from Fe³⁺ ions in the tetrahedral sites (B-site). The Fe 2p spectrum of the CoFe₂O₄–RGO(35) hybrid is nearly identical to that of the high-resolution Fe 2p spectrum of the pristine CoFe₂O₄ (Fig. 3(b)), except for a negative shift in the binding energy (Fig. S2†). Also, the high-resolution Co 2p spectrum of CoFe₂O₄–RGO(35) hybrid is nearly identical to that
of the pure CoFe$_2$O$_4$, except for a negative shift in the binding energy. These results indicate that efficient charge transfer between the CoFe$_2$O$_4$ NPs and graphene sheets occurred owing to the intimate interfacial connections, which may be beneficial for the separation of charge carriers.

The high-resolution Co 2p spectrum of CoFe$_2$O$_4$–RGO(35) hybrid is similar to its Fe 2p spectrum (Fig. 3(d)); the peaks observed at 779.2 eV (Co 2p$_{3/2}$) and 794.6 eV (Co 2p$_{1/2}$) can be assigned to the contributions from Co$^{2+}$ ions in the octahedral sites (A-site), while the doublets binding energies of Co 2p$_{3/2}$ at 781.2 eV and Co 2p$_{1/2}$ at 795.7 eV are ascribed to the contributions from Co$^{3+}$ ions in tetrahedral sites (B-site). Compared with the Fe 2p spectrum, there are two obvious shake-up satellite peaks at 785.8 eV and 802.2 eV in the Co 2p spectrum. The presence of these intense satellite signals indicates that a large number of high spin Co$^{2+}$ ions occupy the octahedral sites (A-site) in spinel-type CoFe$_2$O$_4$. The deconvoluted C 1s spectra of CoFe$_2$O$_4$–RGO(35), as shown in Fig. 3(e), reveals the presence of oxygen-containing groups (C-OH).
The specific surface area and pore-size distribution of the as-prepared samples were determined using the nitrogen sorption technique. The nitrogen adsorption–desorption isotherms are shown in Fig. 4(a). All the samples exhibit typical II-type curves with steep uptakes \((P/P_0 > 0.95)\) and distinct hysteresis loops \((0.40 < P/P_0 < 0.95)\), indicating the co-existence of non-ordered mesopores \((2–50 \text{ nm})\) and numerous macropores \((>50 \text{ nm})\). The primary CoFe\(_2\)O\(_4\) NPs are randomly stacked and this results in the creation of macropores. The Brunauer–Emmett–Teller surface areas of CoFe\(_2\)O\(_4\), CoFe\(_2\)O\(_4\)–RGO(35), and CoFe\(_2\)O\(_4\)–RGO(70) were 71.20, 35.11, and 40.42 m\(^2\) g\(^{-1}\), respectively. Furthermore, the CoFe\(_2\)O\(_4\)–RGO(35) and CoFe\(_2\)O\(_4\)–RGO(70) exhibit classical type H3 hysteresis loops, arising from the capillary condensation and slit-shaped type pore structure of the RGO nanosheets. In addition, both O\(_{\text{latt}}\) and O\(_{\text{surf}}\) play an important role in the radical’s generation process.\(^{46}\)

The fabrication process of the octahedral CoFe\(_2\)O\(_4\)–RGO hybrids in this study is depicted in Fig. 5. In our study, the growth conditions used to fabricate the octahedral CoFe\(_2\)O\(_4\)/RGO hybrids were a high pH \((\sim 13)\) and a relatively long solvothermal time, which is briefly described as follows:

\[
\text{Co}^{2+} + 2\text{OH}^- \rightarrow \text{Co(OH)}_2 \quad (9)
\]

\[
\text{Fe}^{3+} + 3\text{OH}^- \rightarrow \text{Fe(OH)}_3 \quad (10)
\]

\[
\text{Co(OH)}_2 + \text{Fe(OH)}_3 + \text{graphene oxide} + \text{EtOH} \rightarrow \text{CoFe}_2\text{O}_4/\text{RGO} + \text{H}_2\text{O} \quad (11)
\]

This indicates that the formation of this unique nanohybrid can be controlled in two stages: (i) the formation of Co(OH)\(_2\) and Fe(OH)\(_3\) colloidal particles; it is well known that GO possesses negative charge owing to the existence of numerous oxygen-containing functional groups, such as carboxyl, carbonyl, and hydroxyl.\(^{49}\) In addition, the oxygen-containing groups on the GO sheets may serve as heterogeneous nucleation sites.\(^{48}\) When Co(NO\(_3\))\(_2\)·6H\(_2\)O and Fe(NO\(_3\))\(_3\)·9H\(_2\)O are added into the GO suspension, the positively charged metal ions (Co\(^{2+}\) and Fe\(^{3+}\)) would anchor onto the surface of the GO sheets via electrostatic attraction. After increasing the pH of the mixed solution by the introduction of sufficient OH\(^-\), the alkaline conditions promotes the precipitation of Co\(^{2+}\) and Fe\(^{3+}\) ions in ethyl alcohol and this result in the formation of nucleation precursors, i.e. Co(OH)\(_2\) and Fe(OH)\(_3\) colloidal particles (eqn (1) and (2)).\(^{51,52}\) (ii) \textit{In situ} growth of the CoFe\(_2\)O\(_4\) crystals and the confining effects of graphene; during the solvothermal treatment, the Co(OH)\(_2\) and Fe(OH)\(_3\) colloidal particles gradually combine to form CoFe\(_2\)O\(_4\) nuclei. Meanwhile, the GO sheets were reduced during the solvothermal reactions (eqn (3)).\(^{53}\)

To reduce the surface energy, aggregation of many CoFe\(_2\)O\(_4\) nanocrystals further occurred.\(^{54,55}\) It is well known that the CoFe\(_2\)O\(_4\) nanocrystals with a broad size distribution may result from the quick nucleation and growth.\(^{56}\) The CoFe\(_2\)O\(_4\) nanocrystals which are tightly anchored onto the graphene sheets

**Fig. 4** (a) Nitrogen adsorption–desorption isotherm curves and (b) the corresponding Barrett–Joyner–Halenda pore size distribution of the as-prepared CoFe\(_2\)O\(_4\) and CoFe\(_2\)O\(_4\)–RGO.
grew quickly owing to the abundant growth units near the surface of the RGO sheets. Then the octahedral CoFe$_2$O$_4$ particles, which are formed under high autogenous pressure at 180 °C, are enclosed by the {111} surfaces to achieve the lowest Gibbs free energy. Compared with CoFe$_2$O$_4$ nanocrystals which tightly anchors itself onto graphene sheets, the growth of the CoFe$_2$O$_4$ nanocrystals in the mixture solution was slower and it formed irregular CoFe$_2$O$_4$ nanoclusters. Consequently, the CoFe$_2$O$_4$ particles within the CoFe$_2$O$_4$–RGO hybrids exhibited a more obvious difference in morphology, i.e. octahedrons and irregular subunit particles, and the CoFe$_2$O$_4$–RGO hybrids were obtained. However, further experimental evidence is required to verify the present observations.

3.2. Catalytic activity
Control experiments were conducted at near-neutral pH conditions (adjusted with phosphate buffer solution) to compare the degradation efficiency of MB dye by the as-prepared sample and CoFe$_2$O$_4$–RGO/H$_2$O$_2$ at various conditions. MB dye was selected as the targeted pollutant owing to its stability under the visible light. Fig. 6(a) and (b) shows the removal efficiency of MB in the presence of either the catalysts

![Fig. 5 Schematic of the fabrication process of the octahedral CoFe$_2$O$_4$–RGO hybrids.](image)

![Fig. 6 Photocatalytic degradation of MB by the as-prepared samples alone (a), or the catalysts in conjunction with H$_2$O$_2$ (b) under visible-light irradiation. Unless otherwise stated, the reaction conditions are: [MB] = 10 mg L$^{-1}$, 60 mL MB solution, [catalyst] = 1 mg mL$^{-1}$, visible light irradiation ($\lambda > 420$ nm, 500 W), $T = 25$ °C, initial pH = 6.4.](image)
CoFe2O4 and CoFe2O4–RGO) alone or the catalysts together with 1.0 mL 30 wt% H2O2 under visible-light irradiation. The adsorption process was conducted by dispersing the products in the MB solution in the dark for 1.5 h and the photodegradation step began at \( t = 0 \). Clearly, the adsorption ability of bare CoFe2O4 was much higher than that of the CoFe2O4–RGO hybrids owing to the large specific surface area of the CoFe2O4 sample and the MB molecules having an acceptable molecular size. Interestingly, for bare CoFe2O4, the MB molecules were easily re-released into an aqueous solution from its surface after light irradiation or on addition of the H2O2 co-catalyst, which is consistent with a previous report.\(^{53}\) The explanation for this observation is that a weak interaction exists between the CoFe2O4 NPs and MB dye. In addition, the degradation efficiency of the MB solution was rather low in the absence of the photocatalysts.

Fig. 6(a) shows the total removal efficiency of three samples (i.e., CoFe2O4, CoFe2O4–RGO(35), and CoFe2O4–RGO(70)) without H2O2 co-catalyst for MB dye. The order of the removal efficiency is CoFe2O4–RGO(70) > CoFe2O4–RGO(35) > CoFe2O4. The elimination of MB dye by only the catalysts was dominated by the adsorption process due to the \( \pi-\pi \) stacking between the aromatic dye molecules and the hexagonal regions of graphene.\(^{56-57}\) However, with the simultaneous introduction of the as-prepared catalysts and the H2O2 co-catalyst, the photodegradation of MB was fast and efficient owing to the reaction of photogenerated carriers and H2O2.\(^{28}\) The order of the photodegradation efficiency is CoFe2O4–RGO(35) > CoFe2O4–RGO(70) > CoFe2O4 > H2O2. During the photodegradation process, CoFe2O4–RGO(35) exhibited the highest MB photodegradation efficiency (~43%) in 3.5 h among all the samples, and the total MB removal efficiency (~73%) was achieved when the CoFe2O4–RGO(35)/H2O2 system was used (Fig. 6(b)).

The adsorption reaction is often related to the surface charge of the NPs. The effect of solution pH on zeta potentials of the CoFe2O4 and CoFe2O4–RGO hybrids is shown in Fig. 6(c). It can be seen that zeta potentials of the as-prepared particles decreased as the pH increases and that of the CoFe2O4 and CoFe2O4–RGO hybrids were negatively charged at pH > 4.3. In addition, the estimated zeta potentials observed for CoFe2O4, CoFe2O4–RGO(35), and CoFe2O4–RGO(70) were about -19.1, -33.7, and -42.9 mV at pH = ~6.3, respectively. Therefore, the adsorption equilibrium rate toward the cationic MB dye of the CoFe2O4–RGO hybrids was more rapid compared to that of the pure CoFe2O4. This is because of the strong electrostatic interactions between MB and CoFe2O4–RGO hybrids surfaces. Also,
desorption of the MB molecules from the surface of the CoFe2O4-RGO hybrids into the aqueous solution was difficult after light irradiation or addition of the H2O2 co-catalyst.

Fig. 6(d) depicts the effect of hydrogen peroxide dose on MB degradation using CoFe2O4-RGO hybrids. We observed that the concentration of MB reduced slowly under light irradiation for the “visible light with H2O2” system. The introduction of H2O2 weakly improves the photocatalytic effects by increasing the amount of MB dissolved in current experiment, e.g., from 1 mL to 2 mL and from 2 mL to 3 mL under visible light irradiances for 3.5 h, the degradation rate of MB can be slightly increased to 22.2% and 25.14%, respectively. This result indicates that the degradation rate of MB could not be significantly improved by simply increasing the content of H2O2. Compared with the “visible light with H2O2” system, the heterogeneous photo-Fenton mechanism of the CoFe2O4-RGO hybrids was discussed in a later section of this paper.

To identify the main active species for the degradation of dyes in the “CoFe2O4-RGO(70) + H2O2” system, IPA was separately added into the system as a scavenger of hydroxyl radicals (‘OH). As shown in Fig. 7(a), the dye removal efficiency of the CoFe2O4-RGO(70) hybrid for MB was higher than that of MO. In addition, the adsorption and photodegradation rate of the CoFe2O4-RGO(70) hybrid for the anionic dye (MO) was lower when compared with that of the cationic dye (MB), and this was attributed to the repulsive forces that exist between MO and the –OH/R–COO groups on the surface of CoFe2O4-RGO(70) hybrid. Furthermore, the original photodegradation rate is high, i.e. without adding a quenching agent into the reaction solution, and the degree of MB photodegradation was as high as 53.3%. However, the degradation of MB was obviously inhibited by the introduction of 0.1 M IPA, which resulted in a decrease in the degree of MB photodegradation to 4.4% within 3.5 h. This result clearly proved that ‘OH is the dominating reactive species in the “CoFe2O4-RGO + H2O2” system, and it contributes to the decolorization of MB.

The recyclability and stability of the catalyst are important parameters for its commercialization. To evaluate the reusability of our catalytic system, three rounds of the photodegradation reaction for MB dye were conducted with the recyclable CoFe2O4-RGO hybrids under the same reaction conditions. After each cycle, the catalyst was separated from the reaction solution using an external magnet, and it was washed several times with deionized water and ethanol. As shown in Fig. 7(b), the removal efficiencies for MB in the third cycle were nearly the same as that of the first two cycles. Although the activity of the CoFe2O4-RGO(70) hybrid slightly decreased in the third cycle, the removal efficiencies for MB was still 79.2%. The slight loss of the photodegradation activity may be due to the loss of the catalysts during the recovery process or the weak decoration of CoFe2O4 and RGO. All these analyses indicate that the CoFe2O4-RGO hybrid is still highly efficient, arising from the graphene/CoFe2O4 heterojunction and the lowest energy {111} facets of octahedral CoFe2O4, when reused.

The magnetic hysteresis loops of the pure CoFe2O4 and CoFe2O4-RGO hybrids were conducted using a vibrating sample magnetometer (VSM) from −14 000 to 14 000 Oe at room temperature, as shown in Fig. 7(c). All the samples exhibit ferromagnetic behavior. The saturation magnetization (Ms) value of the pure CoFe2O4 nanocrystals is approximately 31.15 emu g−1, which is lower than that of the CoFe2O4-RGO hybrids (41.98–45.14 emu g−1). Compared with the bare CoFe2O4 NPs, the enhancement of the Ms for CoFe2O4-RGO hybrids is mainly due to the larger average particle size of the octahedral CoFe2O4 nanocrystals.

Hence, more efficient separation of CoFe2O4-RGO hybrids is possible. As shown in Fig. 7(d), the first bottle (marked with 0 s) is the initial MB (10 mg L−1) solution without CoFe2O4-RGO(70) catalyst. The CoFe2O4-RGO hybrids were uniformly dispersed in MB solution to form a dark brown suspension (marked with 30 s). However, when an external magnet was placed close to the bottle, the synthesized samples which had absorbed MB from the aqueous solution were rapidly separated from the aqueous media (marked with 60 s). Thus, the CoFe2O4-RGO catalyst is quite acceptable owing to its good separation ability. In addition, the Ms values of the samples were lower than that of the bulk counterpart (74 emu g−1), which may be ascribed to the possible effect of surface spin canting on the CoFe2O4 NPs as well as the weight-fraction of the RGO. Moreover, the field-dependent magnetization curves of the CoFe2O4 and CoFe2O4-RGO hybrids exhibit non-negligible remanence (Mr = 14.92–25.42 emu g−1) and high coercivity (Hc = 4.32–6.41 kOe), indicating that they are hard magnetic materials (Table S1†).

3.3. Activation mechanism

Based on the information above, we propose an activation mechanism of H2O2 by the CoFe2O4-RGO hybrids, as presented in Fig. 8. Clearly, the main active sites of CoFe2O4-RGO hybrids are the cobalt and iron ions (Co2+/3+ and Fe3+/2+) in a spinel-type CoFe2O4 surface lattice. When the CoFe2O4 particles are excited using a daylight lamp, electron–hole pairs are produced on the surface of the CoFe2O4 particles (eqn (14)). The presence of highly conductive RGO in the CoFe2O4-RGO hybrids can inhibit the recombination of electrons and holes because electrons are quickly transferred to the RGO sheets due to a percolation mechanism (eqn (15)). Then the photogenerated electrons within the graphene sheets could be captured by H2O2 to

\[
\text{’OH + H}_2\text{O}_2 \rightarrow \text{H}_2\text{O} + \text{’OOH} \tag{12}
\]

\[
\text{HOO’ + ’OH} \rightarrow \text{H}_2\text{O} + \text{O}_2 \tag{13}
\]
These catalysts exhibited excellent H$_2$O$_2$ activity for the degradation of MB dye at nearly neutral pH and room temperature. On screening the catalyst, we observed that CoFe$_2$O$_4$–RGO(35) and CoFe$_2$O$_4$–RGO(70) hybrids exhibited superior photo-Fenton activity for the removal of MB when compared with pure CoFe$_2$O$_4$. We demonstrated that ‘OH is the dominating reactive species in the “CoFe$_2$O$_4$–RGO + H$_2$O$_2$” heterogeneous photo-Fenton system, and it is responsible for the decolorization of MB. In addition, the high catalytic activity of CoFe$_2$O$_4$–RGO(70) hybrids was attributed to the synergistic effect of redox Co$^{2+/3+}$ and Fe$^{3+/2+}$ pairs. Based on the results obtained from the experiment conducted, we proposed an activation mechanism of H$_2$O$_2$ by the CoFe$_2$O$_4$–RGO hybrids. In view of the excellent catalytic activity, the outstanding reusability and stability, as well as the magnetic recyclability of CoFe$_2$O$_4$–RGO hybrid, it can potentially be applied in advanced oxidation processes as a heterogeneous photo-Fenton catalyst.

### Conflicts of interest

There are no conflicts to declare.

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