Spectrophotometric determination of the stability constant of vanadium(v)-p-chloro-N-phenyl-p-chlorobenzo-hydroxamic acid complex

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Abstract

The stability constant of the complex vanadium(v) p-chloro-N-phenyl-p-chlorobenzo-hydroxamic acid was determined spectrophotometrically. The ligand p-chloro-N-phenyl-p-chlorobenzo-hydroxamic acid abbreviated in this paper as PCNPPCBHA was synthesized by the coupling of equimolar solutions of p-chlorobenzoic chloride and freshly prepared p-chloro-ß-phenyl hydroxylamine, in which the later was prepared by the reduction of p-chloronitrobenzene with zinc powder. The pH at maximum extraction of the ligand PCNPPCBHA to the metal vanadium(v) and the ratio of the ligand: metal, were determined.

Keywords: Spectrophotometric determination, Vanadium(V), p -chloro - N - phenyl - p - chlorobenzo-hydroxamic acid.

1. Introduction

Hydroxamic acid have been known since 1869 with the discovery of oxalohydroxamic acid by Lossen (Rappoport & Liebman 2009). Despite this, researches on these compounds were lacking until the 1980, after which an enormous amount of information has accumulated with respect to their biomedical applications, synthesis, determination of their metal complexes(Mamion.et al . 2004), their pharmaceutical (Musser et al .1990) and industrial applications(Odake et al.1991), so the chemistry of hydroxamic acid has received considerable attention (Miller 1989). Particularly important is their role as siderophores, as well as a model system for natural siderophores (Matzanke et al .1989). The ability of hydroxamic acids to form stable transition metal complexes is the basis of their usefulness as an analytical reagent; this is the reason why hydroxamic acids have been used as chelating agents for determination of metals in samples of environmental and industrial importance (Rasuljan et al .1996) .

2. Experimental

All chemicals used were AnalAR grade. All solutions were prepared according to the usual analytical procedures. using distilled water.

A series of buffer solutions ranging from pH 1.0 to pH 7.0 were prepared from a mixture of 0.2 M hydrochloric acid and 0.2 M potassium chloride in different proportions; and also a mixture of 0.2M acetic acid and 0.2M sodium acetate in different proportions.

2.1. Synthesis of p-Chloro-ß -phenyl hydroxylamine (Vogel 1951)

p-chloro-ß-phenyl hydroxylamine was prepared by reduction of 0.4 mole p-chloronitrobenzene with 0.9 mole zinc powder in present of 0.4 mole of ammonium chloride at temperature of 55°C. The products were recrystallized from petroleum benzene.

2.2. Synthesis of the ligand p- chloro – N – phenyl – p – chloro-benzo-hydroxamic acid (PCNPPCBHA) (Usha & Tand on 1967):

The ligand PCNPPCBHA was prepared by the coupling of 0.1 mole p-chlorobenzoyle chloride (dissolved in diethyl ether) and

[Chemical structure and reaction scheme are shown in the text and diagrams]
0.1 mole of freshly prepared p-chloro-β-phenyl hydroxylamine (dissolved in diethyl ether) at temperature of 0°C and in present of 0.2 mole sodium hydrogen carbonate. The product was recrystallized from a mixture of benzene and petroleum ether (1:1).

2.3. Extraction and spectrophotometric determination of vanadium (V): (determination of the pH at maximum extraction):

5 cm³ portions of 100 ppm vanadium solution ((ammonium Meta-vandate NH₄VO₃, A.R.) were pipetted into a series of 25 cm³ volumetric flasks and completed to the mark with acetate buffer solutions ranging from pH 1.0 to pH 7.0 to give 20 ppm V(v) in each. The solutions were transferred to a series of 100 cm³ separatory funnels; few drops of 0.001M potassium permanganate were added to each until the solution became faint pink color. 25 cm³ of 0.02M PCNPPCBHA in chloroform were added to each separatory funnel and the mixture were shaken gently for two minutes; the layers were allowed to separate; (I) The organic layer (lower phase) and (II) The aqueous layer (upper phase). The aqueous layers were taken for further analysis to determine the pH at maximum extraction.

2.4. Analysis of the aqueous layer by spectrophotometric procedure (Abdalla, 1980, Ryan 1960):

The standard (calibration curve) for vanadium(v) and the sample (aqueous layer) were carried out at the same time for spectrophotometric procedure. Results are shown in Table (6) and (7) and represented graphically in figure (1).

3. Results and discussion

The synthesized ligand p-chloro-N-Phenyl-p-chloro benzohydroxamic acid (PCNPPCBHA) was identified by its melting point table (1); infra-red spectroscopy table (2) appendix (A); ultra-violet absorption in 95% ethyl alcohol table (3) appendix (B); elemental analysis of its nitrogen (Abdalla, 1992) and chlorine (Openshow1995) contents table (4) and table(5) and its reactions towards acidic solutions of vanadium(v) and iron (III) to give respectively; violet and blood- red colored complexes in the chloroform layers.

Table 4: The Melting Point of the Ligand PCNPPCBHA Using Mettler, Melting Point Determination Apparatus.

| Ligand | Melting point of PCNPPCBHA in °C | Found |
|--------|---------------------------------|-------|
| PCNPPCBHA | 158 - 159                      |       |

Table 2: Characteristic Functional Groups of the Ligand PCNPPCBHA Using Infra-Red Spectrophotometer; Perkin Elmer Model 1330.

| Functional group | O - H | C = O | C - N | N - O | C - Cl |
|------------------|-------|-------|-------|-------|--------|
| Found in cm⁻¹    | 3185  | 1620  | 1370  | 825   | 720    |

Table 3: Maximum Absorption of the Ligand PCNPPCBHA In 95% Ethyl Alcohol Using UV/VIS Spectrophotometer; Perkin Elmer 550S.

| Ligand      | Wave length in nm | λ - λ transition of C=O | Aryl ring |
|-------------|-------------------|--------------------------|-----------|
| PCNPPCBHA   |                   | 205                      | 233       |

Table 5: Chlorine Content of the Ligand PCNPPCBHA Using Folin’s Method (Openshow1995).

| Percentage | Chlorine | Calculated | Found |
|------------|----------|------------|-------|
|            | 25.13    | 25.2       |       |

Table 6: Absorbance of the Standard Solutions for the Calibration Curve of Vanadium(v) at λ 520 nm.

| Concentration of Vanadium (V) in ppm | Medium | Absorbance |
|--------------------------------------|--------|------------|
| 1                                   | 5M HCl | 0.084      |
| 2                                   | -      | 0.175      |
| 3                                   | -      | 0.259      |
| 4                                   | -      | 0.348      |
| 5                                   | -      | 0.429      |

* pH at maximum extraction implies the lowest value of the of the absorbance i.e. at pH = 3.00

Table 7: Absorbance of the Aqueous Layer from the Extraction of Vanadium(v) With PCNPPCBHA.

| Solution | 1 | 2 | 3 | 4 | 5 | 6 | 7 |
|----------|---|---|---|---|---|---|---|
| Vol. of V(v) 5x10⁻³ in cm³ | 2 | 3 | 4 | 5 | 6 | 7 | 8 |
| Vol. of Acetate pH 3 in cm³ | 2 | 2 | 2 | 2 | 2 | 2 | 2 |
| Vol. of Dist H₂O in cm³ | 6 | 5 | 4 | 3 | 2 | 1 | 0 |
| Vol. of ligand 5x10⁻³ in cm³ | 8 | 7 | 6 | 5 | 4 | 3 | 2 |
| Vol. of chloroform added to the ligand before added the metal in cm³ | 2 | 3 | 4 | 5 | 6 | 7 | 8 |

Fig 1: Calibration curve of λ (nm) for extraction with PCNPPCBHA

Spectrophotometric studies on Vanadium(v) p-chloro–N-phenyl – p-chlorobenzohydroxamic acid complex:

I) Continuous Variation Method:
The composition of Vanadium (v) – PCNPPCBHA complexes was determined by application of the continuous variation method (Job 1928, Warren et al. 1941). Results are shown in Table (8) and represented graphically in figure (2). The blank used was chloroform.

Table 8: Continuous Variation Method for Vanadium (V) – PCNPPCBHA System.

| Solution | 1 | 2 | 3 | 4 | 5 | 6 | 7 |
|----------|---|---|---|---|---|---|---|
| Vol. of M ≈ L | 0.20 | 0.30 | 0.40 | 0.50 | 0.60 | 0.70 | 0.80 |
| Absorbance at λ 520 nm | 0.301 | 0.383 | 0.455 | 0.530 | 0.486 | 0.418 | 0.369 |

The synthesized ligand p-chloro-β-phenyl hydroxylamine (PCNPPCBHA) was identified by its melting point table (1); infra-red spectroscopy table (2) appendix (A); ultra-violet absorption in 95% ethyl alcohol table (3) appendix (B); elemental analysis of its nitrogen (Abdalla, 1992) and chlorine (Openshow1995) contents table (4) and table(5) and its reactions towards acidic solutions of vanadium(v) and iron (III) to give respectively; violet and blood- red colored complexes in the chloroform layers.
II) Slope Ratio Method:
The ratio of the metal Vanadium (v) to the ligand PCNPPCBHA was determined by application of the slope ratio method (Harvey & Manning 1952) which was found 1:1. Results are shown in table (9) and represented graphically in figure (3).

III) Calibration Curve of Vanadium (v) – PCNPPCBHA system:
A standard calibration curve of Vanadium(v)-PCNPPCBHA system was carried out against chloroform as blank. Results are shown in table (10) and represented graphically in figure (4).

### Table 9: Slope - Ratio Method for Vanadium (V) – PCNPPCBHA System.

| Solution | Vol. of V(V) 5x10^{-2} in cm^3 | Vol. of Acetate pH 3 in cm^3 | Vol. of Dist.H_2O in cm^3 | Vol. of ligand 5x10^{-2} in cm^3 | Vol. of chloroform added to the ligand before added to the metal in cm^3 | [V(V)] x 10^{-3} M | Absorbance at \(\lambda\) 520 nm |
|----------|-------------------------------|------------------------------|--------------------------|----------------------------------|-------------------------------------------------|---------------------|-------------------------|
| Set (A) : | 1 2 3 4 5 6                   | 1 2 3 4 5 6                  | 1 2 3 4 5 6               | 1 2 3 4 5 6                       | 1 2 3 4 5 6                                                   | 0.50 1.00 1.50 2.00 2.50 3.00 | 0.243 0.255 0.274 0.304 0.315 0.331 |
| Set (B) : | 1 2 3 4 5 6                   | 5 5 5 5 5 5                  | 5 5 5 5 5 5               | 5 5 5 5 5 5                       | 5 5 5 5 5 5                                                   | 0.50 1.00 1.50 2.00 2.50 3.00 | 0.218 0.234 0.255 0.275 0.304 0.325 |

### Table 10: Calibration Curve for Vanadium (v) – PCNPPCBHA System.

| Solution | Vol. of V(V) 5x10^{-2} in cm^3 | Vol. of Acetate pH 3 in cm^3 | Vol. of Dist.H_2O in cm^3 | Vol. of ligand 5x10^{-2} in cm^3 | [V(V)] x 10^{-3} M | Absorbance at \(\lambda\) 520 nm |
|----------|-------------------------------|------------------------------|--------------------------|----------------------------------|---------------------|-------------------------|
| 1        | 1 2 3 4 5 6                   | 1 2 3 4 5 6                  | 1 2 3 4 5 6               | 1 2 3 4 5 6                       | 0.50 1.00 1.50 2.00 2.50 3.00 | 0.218 0.234 0.255 0.275 0.304 0.325 |

### Synthesis of the Ligand:
The ligand PCNPPCBHA was prepared by a coupling reaction which took place under ether process (Usha & Tandon 1967) which is favored over the aqueous process (Tandon & Bhattacharya 1962) due to the relative purity of the product obtained. The use of sodium hydrogen carbonate in the coupling reaction is to keep the coupling medium slightly alkaline; also it serves to neutralize hydrochloric acid evolved throughout the reaction. The infra-red spectroscopy for the ligand reveals the absorbance of the most important characteristic functional groups ((Table (2)appendix (A)) such as O-H at 3185 cm^{-1}, the lower shift of O-H instead of 3190 cm^{-1} (Agrawal 1967) is due to the intra molecular hydrogen bonding of the type –OH…C=O.

The ultra-violet spectra of the ligand reveal two absorption maxima (\(\lambda_{\text{max}}\)) when dissolved in 95% ethyl alcohol; The first band absorbs at 233 nm referring to the absorption due to the benzene ring and the other band absorbs at 205 nm referring to the absorption due \(\pi\) — \(\pi^*\) transition of the carbonyl group (C=O). The substitution of the chloride at the para-position of the two phenyl rings to the parent acid decreases the absorption band due to the carbonyl \(\pi\) — \(\pi^*\) transition when compared with the parent acid alone (Bowie et al.1969).

p-Chloro-0-phenyl hydroxylamine which is needed for the synthesis of the ligand is an intermediate and of low stability which decompose upon staying for a long time; so it must be freshly prepared and recrystallized for the coupling reaction.
Calculation of the stability constant for Vanadium (V) PCNPPCBHA complex (Korthoff 1963):

The formation (stability) constant can be determined by job's method.

\[ V^5+ + PCNPPCBHA \leftrightarrow [V \cdot PCNPPCBHA]^5^- \]

\[ K_s = \frac{[V \cdot PCNPPCBHA]}{[V^5+][PCNPPCBHA]} \]  \hspace{1cm} (1)

The concentration of the species \([V \cdot PCNPPCBHA]^5^- ; V^5+ \) and \([PCNPPCBHA]^- \) should be calculated from the maximum absorbance (A) observed in job’s curve figure (2).

\[ A_{520nm} = 0.530 \]

Volume of \( V^5+ = 5\text{cm}^3 \) from table (8)

Volume of PCNPPCBHA= 5 cm\(^3\)

\[ C_{V^5+} = \frac{[V^5+]}{5\text{cm}^3} \]

\[ C_{PCNPPCBHA} = \frac{[PCNPPCBHA]}{5\text{cm}^3} \]

\[ \therefore [V \cdot PCNPPCBHA] = C_{V^5+} \times C_{PCNPPCBHA} \]

Also ; \[ C_{PCNPPCBHA} = [PCNPPCBHA] + [V \cdot PCNPPCBHA]^5^- \]

\[ C_{PCNPPCBHA} = C_{PCNPPCBHA} + [V \cdot PCNPPCBHA]^5^- \]

\[ C_{PCNPPCBHA} = 0.005X5 \]

\[ C_{PCNPPCBHA} = 0.025M \]

\[ A_{ex} \] is the extrapolated absorbance i.e. the absorbance which would be observed if the complex didn’t dissociate at all. When this happens the equilibrium concentration of the complex \([V \cdot PCNPPCBHA]^5^- \) would be equal to the analytical concentration of the vanadium \(C_{V^5+} \).

From job’s curve figure (2):

\[ A_{ex} = 0.561 at 520 \text{ nm} \]

\[ [V \cdot PCNPPCBHA]^5^- = C_{V^5+} \]

\[ A_n = [V \cdot PCNPPCBHA]^5- \]

\[ A = [V \cdot PCNPPCBHA]^5- \]

\[ \therefore \]

\[ \frac{A}{A_n} = \frac{C_{V^5+}}{C_{V^5+}} \]

\[ \therefore [V \cdot PCNPPCBHA]^5- = \frac{A_n}{A} \times C_{V^5+} \]  \hspace{1cm} (4)

Substituting equation (2) and (3) in equation (1) we get:

\[ K_s = \frac{[C_{V^5+} - [V \cdot PCNPPCBHA]^5-]}{[C_{PCNPPCBHA} - A] \times \frac{C_{V^5+}}{A_n} \times C_{V^5+}} \]

Substituting equation (4) in the above equation we get :

\[ = \frac{A}{A_n} \times C_{V^5+} \]

\[ = \frac{0.530}{0.561} \times \frac{0.025}{0.561} \]

\[ = 12.4 \times 10^4 \]

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