Article

Structure, Spectra and Photochemistry of 2-Amino-4-Methylthiazole: FTIR Matrix Isolation and Theoretical Studies

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Abstract: The structure, tautomerization pathways, vibrational spectra, and photochemistry of 2-amino-4-methylthiazole (AMT) molecule were studied by matrix isolation FTIR spectroscopy and DFT calculations undertaken at the B3LYP/6-311++G(3df,3pd) level of theory. The most stable tautomer with the five-membered ring stabilized by two double C=C and C=N bonds, was detected in argon matrices after deposition. When the AMT/Ar matrices were exposed to 265 nm selective irradiation, three main photoproducts, N-(1-sulfanylprop-1-en-2-yl)carbodiimide (fp1), N-(1-thioxopropan-2-yl)carbodiimide (fp2) and N-(2-methylthiiran-2-yl)carbodiimide (fp3), were photoproduced by a cleavage of the CS–CN bond together with hydrogen atom migration. The minor photoreaction caused by the cleavage of the CS–CC bond and followed by hydrogen migration formed 2-methyl-1H-azirene-1-carbimidothioic acid (fp15). We have also found that cleavage of the CS–CN bond followed by disruption of the N–C bond produced cyanamide (fp11) and the C(CH3)=CH–S biradical that transformed into 2-methylthiirene (fp12) and further photoreactions produced 1-propyne-1-thiole (fp13) or methylthioketene (fp14). Cleavage of the CS–CC bond followed by disruption of the N–C bond produced propyne (fp22) and the S–C(NH2)=N biradical that transformed into 3-aminethiazirene (fp23); further photoreactions produced N-sulfanylcarbodiimide (fp25). As a result of these transformations, several molecular complexes were identified as photoproducts besides new molecules in the AMT photolysis process.  

Keywords: thiazole; argon matrices; DFT  

1. Introduction  

Thiazoles are basic heterocyclic compounds with five-membered ring including one nitrogen and one sulfur atom. Many compounds containing the thiazole system in the molecule exhibit a broad spectrum of pharmacological activity. Sulfathiazole, Ritonavir, Abafungin, Tiazofurin, or Bleomycin are often used antimicrobial, antiretroviral, antifungal, antineoplastic or antitumor drugs with the thiazole ring in the structure, respectively [1–3]. An important naturally occurring compound containing a thiazole ring is thiamine (i.e., vitamin B1) consisting of thiazole and pyrimidine rings, connected by a methylene group. Thiamine is an important cofactor in the enzymes of glycolysis, the Krebs cycle and the pentose phosphate pathway. It also takes part in the biosynthesis of neurotransmitters—acetylcholine and gamma-aminobutyric acid [4,5]. The latest review papers on the biological and pharmaceutical properties of 2-aminothiazoles concern both the synthesis and anti-cancer properties of these compounds [6–8]. Another interesting aspect in the research on thiazoles is their role in prebiotic chemistry on the early Earth [9,10]. Recently, 2-aminoazoles have been considered as intermediates in nucleotide synthesis or nucleotide activating or/and selective agents in prebiotic chemistry [11]. 2-aminothiazole
selectively crystallizes the needed sugars in the presence of mixtures of aldoses and ketoses playing the role of chaperone for ribonucleotide synthesis [12].

Photochemistry of some five-membered ring containing two hetero atoms has been investigated but the complete reaction pathways have not been established yet [13]. Flash photolysis of thiazole in the gas phase gave initially the highly excited NCS molecule which acted as a precursor to the rotationally and vibrationally excited CN radical [14]. VUV photodissociation of thiazole molecule investigated by TOF-MS and photoelectron photoion coincidence spectroscopy produced several ionic species, such as C₂H₅S⁺, CHS⁺, CH₂N⁺ or S⁺ [15]. Recently, the UV induced photoreactions of thiazole isolated in low temperature argon matrices were examined [16]. The major ring-opening photoreactions caused by the cleavage of the S1–C2 bond gave syn-2-isocyanoethenethiol or 2-isocyanothiirane as the initial photoproducts that changed into secondary photolysis products after prolonged irradiation time. When the N3–C4 bond was broken and, on one side, was followed by the cleavage of the S1–C2 bond, the CH₂=C=S or HC≡CSH molecules were formed, and when on the other side, the C5–Si final cleavage took place, the HN=C=S or N≡C–SH products were identified. Other minor ring-opening photoreactions were also observed.

In the present paper we reveal a combined spectroscopic and theoretical study of the structure, vibrational spectra and photochemical behavior of 2-amino-4-methylthiazole (AMT). The structure of AMT is presented in Scheme 1.

Scheme 1. Schematic representation of the AMT molecule and the adopted atom numbering.

2. Results and Discussion

2.1. Structure and Energetics

The 2-amino-4-methylthiazole molecule (AMT) can exist in several tautomeric forms, related to the migration of a proton within adjacent atoms. In the case of this heterocyclic compound there is a hydrogen atom transfer within the three-atom systems: three-carbon, imino-enamine and amidine. One may expect formation of five different types of tautomers which, following the numbering shown in Scheme 1 are denoted as AMT1, AMT2, its isomer AMT2′, AMT3, the isomeric structure AMT3′, AMT4, its isomer AMT4′ and AMT5. Full geometry optimization at the B3LYP/6-311++G(3df,3pd) level revealed eight minima. Figure S1 presents the optimized structures of the AMT minima, their energy parameters are gathered in Table 1; the selected calculated geometrical parameters of the AMT tautomers are collected in Table S1. The structures of the relevant transition states are shown in Figure S2 with the energy parameters given in Table 1. The tautomers of AMT differ by the presence of hydrogen at the N and C atoms, respectively, in both substituents and in the ring. In the case of tautomeric transformations AMT1 ↔ AMT5, AMT2 ↔ AMT4 and AMT2′ ↔ AMT4′, hydrogen transfer occurs only between carbon atoms, HC=C=C=C=CH. In turn, in the case of AMT1 ↔ AMT2 and AMT4 ↔ AMT5 transformations, the hydrogen atom migrates within the amidine system, HN–C=N ↔ N≡C–NH. On the other hand, the tautomeric transformations of AMT3′ ↔ AMT4 and AMT3 ↔ AMT4′ take place in the imino-enamine system, HN–C=C ↔ N≡C–CH. The consequence of the migration of the hydrogen atom is the different position of the two double bonds within the structures of the energy minima of 2-amino-4-methylthiazole, which affects its aromaticity, ring stabilization and the stability of individual tautomers.
Table 1. Relative energies ($\Delta E$), relative zero-point corrected energies ($\Delta E_{ZPE}$), relative Gibbs free energies ($\Delta G$) (in kJ mol$^{-1}$) and abundances at 298 K for 2-amino-4-methylthiazole tautomers and transition states calculated at the B3LYP/6-311++G(3df,3pd) level.

| Minima    | $\Delta E$ | $\Delta E_{ZPE}$ | $\Delta G_{298.15}$ | %$_{298}$ |
|-----------|------------|------------------|----------------------|-----------|
| AMT1      | 0.0        | 0.0              | 0.0                  | 100       |
| AMT2      | 31.3       | 31.1             | 30.8                 | 0         |
| AMT2'     | 45.3       | 44.1             | 43.8                 | 0         |
| AMT3      | 42.5       | 41.3             | 40.3                 | 0         |
| AMT3'     | 46.1       | 45.6             | 44.5                 | 0         |
| AMT4      | 54.5       | 55.7             | 54.6                 | 0         |
| AMT4'     | 68.5       | 68.9             | 68.0                 | 0         |
| AMT5      | 44.4       | 44.5             | 42.6                 | 0         |

| TS         | $\Delta E$ | $\Delta E_{ZPE}$ | $\nu_i$/cm$^{-1}$  |
|------------|------------|------------------|--------------------|
| tsAMT(1-2) | 224.8      | 208.6            | 1965i              |
| tsAMT(1-5) | 379.3      | 360.3            | 1509i              |
| tsAMT(2-2')| 77.9       | 72.1             | 850i               |
| tsAMT(2-4) | 363.0      | 343.6            | 1429i              |
| tsAMT(2'-4')| 405.8     | 387.1            | 805i               |
| tsAMT(3'-3) | 87.8      | 86.0             | 1008i              |
| tsAMT(3'-4')| 342.0      | 324.8            | 2032i              |
| tsAMT(4-3') | 285.0      | 266.9            | 2046i              |
| tsAMT(4-4') | 88.0       | 80.8             | 951i               |
| tsAMT(5-4) | 211.0      | 195.4            | 1969i              |

The obtained relative energies ($\Delta G_{298.15}$) (see Table 1) indicate that the AMT1 structure is the most stable one, which is mainly related to its aromaticity and the most effective stabilization of the five-membered ring by two double bonds C=C and C=N. AMT2, AMT2', AMT3, AMT5, AMT3' and AMT4, AMT4' tautomers are characterized by much lower stability. The relative abundances of the studied AMT tautomers were calculated using equation $\Delta G = -RT \ln K$, where $\Delta G$ denotes the difference between Gibbs free energy for two tautomeric forms and $K$ is the equilibrium constant for these species. Large energy differences induce that only AMT1 tautomer is populated to a significant extent under thermodynamic equilibrium. The populations of other tautomers would be negligibly low.

Figure 1 shows the ZPE corrected potential energy diagram for the tautomerization of the AMT structures. The direct transformation possible for the most stable AMT1 species include AMT1 $\rightarrow$ AMT5 tautomerization through tsAMT(1-5) proceeding by hydrogen transfer between carbon atoms of the substituent CH$_3$ group and CH group of the thiazole ring with a very high energy barrier of 360.3 kJ mol$^{-1}$. Next transformations in this reaction path AMT5 $\rightarrow$ AMT4, AMT4 $\rightarrow$ AMT3' and AMT3 $\rightarrow$ AMT4' include hydrogen transfer between nitrogen atoms and between nitrogen and carbon atoms via tsAMT(5-4), tsAMT(4-3') and tsAMT(3-4') transition states with high energy barriers of 150.9, 211.2 and 283.5 kJ mol$^{-1}$, respectively. Hydrogen transfer processes between CH$_3$ group and CH group of the thiazole ring with the very high energy barriers of 343.0 and 312.5 kJ mol$^{-1}$ arise from AMT2' $\rightarrow$ AMT4 and AMT2 $\rightarrow$ AMT4 reactions, respectively. The lowest energy barriers occur for transformations AMT2 $\rightarrow$ AMT2', AMT4 $\rightarrow$ AMT4', AMT3' $\rightarrow$ AMT3 and are equal to 28.0, 25.1, 40.4 kJ mol$^{-1}$, respectively. For these transformations there is no migration of hydrogen between different atoms within the molecule, but the so-called rotation on the double bond, also described in the literature as the reflection of the hydrogen atom in the isomerization reaction (One Bond Flip—OBF) [17,18]. In 2-amino-4-methylthiazole, as a result of this transformation, the position of the hydrogen atom changes only in relation to the adjacent nitrogen atom, which requires much less energy than in the case of other transformations. The most stable AMT1 may also undergo a tautomerization through tsAMT(1-2) to form AMT2 by the direct hydrogen transfer between two N atoms. The calculated energy barrier for this reaction is very high and equals 208.6 kJ mol$^{-1}$. 
2.2. Matrix Isolation Spectra of AMT

The estimated relative abundance of AMT tautomers in the gas phase based on the calculated values of Gibbs free energy is presented in Table 1. According to these results only the most stable AMT1 structure is expected to be present in low temperature matrices directly after deposition. Figure 2 shows the FTIR spectrum of AMT/Ar matrix deposited at 15 K (10 K for measurement) compared to the simulated theoretical spectrum of the AMT1 tautomer. Calculated spectra of all tautomers are presented in Figure S3 for comparison. As can be seen, a satisfying reproduction of the band positions of the experimental data was obtained for the B3LYP method of calculation, therefore it was used for further calculations. In the case of the MP2 method, the values of the calculated wavenumbers differed to a greater extent from the corresponding experimental wavenumbers. However, one can see that the spectrum simulated by the MP2 method better approximates the intensity of the bands in the range below 700 cm\(^{-1}\). Most likely, the distribution of potential energy between these vibrations differs between methods. Moreover, if there are different contributions to different internal coordinates in a given vibration, this translates into changes in the dipole moment, and thus in the vibration intensity. The positions and intensities of the bands observed in the spectra of AMT/Ar matrices together with the wavenumbers predicted by the calculations for AMT1 tautomer are listed in Table 2.
Figure 2. (a) Experimental spectrum of AMT isolated in an argon matrix deposited at 15 K and measured at 10 K; (b) spectrum of AMT1 tautomer calculated by the MP2/aug-cc-pVTZ method; (c) spectrum of AMT1 tautomer calculated by the B3LYP/6-311++G(3df,3pd) method.

Table 2. Observed vibrational wavenumbers in an argon matrix (cm$^{-1}$) compared with the calculated wavenumbers of the AMT1 tautomer of 2-amino-4-methylthiazole$^1$.  

| Harmonic | B3LYP | MP2 | AMT/Ar | Approximate Description$^4$ |
|----------|-------|-----|--------|-----------------------------|
|          | Calc  | Scaled$^2$ | Int | Harmonic | Calc  | Scaled | Int | Exper | Int |                     |
| 3672     | 3540  | 34$^3$ |       | 3681     | 3508  | 38     |      | 3525.5 | w | $\nu_{as}NH_2$     |
| 3564     | 3436  | 42    |       | 3563     | 3396  | 39     |      | 3419.5 | m | $\nu_{s}NH_2$     |
| 3257     | 3140  | 1     |       | 3278     | 3124  | 1      |      | n.o.    | n.o. | $\nu CH$          |
| 3116     | 3004  | 15    |       | 3171     | 3022  | 9      |      | 3009.0 | vVW | $\gamma_{as}CH_3$ |
| 3081     | 2970  | 12    |       | 3151     | 3003  | 8      |      | 2971.5 | vVW | $\gamma_{as}CH_3$ |
| 3031     | 2922  | 27    |       | 3071     | 2927  | 22     |      | 2938.5 | vVW | $\gamma_{s}CH_3$  |
| 1639     | 1603  | 156   |       | 1641     | 1597  | 117    |      | 1600.5 | vs  | $\delta NH_2$    |
### Table 2. Cont.

| Harmonic Calc | Scaled Calc | Int | Harmonic Calc | Scaled Calc | Int | Exper | Int | Approximate Description |
|---------------|-------------|-----|---------------|-------------|-----|-------|-----|-------------------------|
| 1581          | 1546        | 69  | 1567          | 1525        | 41  | 1547.5| m   | Ring stretch            |
| 1557          | 1523        | 109 | 1559          | 1517        | 92  | 1525.0| vs  | Ring stretch            |
| 1483          | 1450        | 14  | 1499          | 1459        | 12  | 1448.0| w   | δasCH₃                  |
| 1476          | 1443        | 6   | 1492          | 1452        | 7   | 1383.0| w   | δasCH₃                  |
| 1417          | 1386        | 10  | 1417          | 1379        | 10  | 1358.0| w   | δsCH₃                   |
| 1355          | 1326        | 117 | 1368          | 1331        | 70  | 1329.5| vs  | Ring stretch, δCH       |
| 1301          | 1272        | 36  | 1316          | 1280        | 57  | 1286.0| vV  | Ring stretch, ρNH₂       |
| 1164          | 1139        | 7   | 1163          | 1132        | 5   | 1142.5| vV  | δCH                     |
| 1116          | 1092        | 20  | 1121          | 1091        | 28  | 1100.5| vV  | ρNH₂                    |
| 1064          | 1040        | 1   | 1065          | 1036        | 1   | 1037.5| vV  | ρCH₃                    |
| 1015          | 992         | 6   | 1011          | 984         | 6   | 995.0 | vV  | ρCH₃                    |
| 976           | 954         | 14  | 976           | 950         | 14  | 963.0 | vV  | Ring def                |
| 844           | 825         | 7   | 876           | 852         | 7   | 851.5 | vV  | Ring def, γS-C          |
| 705           | 705         | 23  | 711           | 711         | 17  | 701.5 | vV  | γCH                     |
| 675           | 675         | 8   | 678           | 678         | 198 | 692.5 | m   | Ring def, υS-C          |
| 657           | 657         | 31  | 699           | 628         | 20  | 657.0 | vV  | Ring breath              |
| 614           | 614         | 31  | 628           | 597         | 27  | 629.0 | vV  | Ring def                |
| 592           | 592         | 37  | 597           | 551         | 3   | 607.0 | vV  | Ring def                |
| 550           | 550         | 5   | 551           | 539         | 58  | 444.0 | w   | ωNH₂                    |
| 509           | 509         | 198 | 539           | 539         | 58  | 444.0 | w   | ωNH₂                    |
| 402           | 402         | 1   | 398           | 398         | 2   | 398   |     | Skeletal def             |
| 299           | 299         | 14  | 255           | 255         | 27  | 255   |     | Skeletal def             |
| 277           | 277         | 10  | 285           | 285         | 3   | 285   |     | Skeletal def             |
| 267           | 267         | 8   | 272           | 272         | 4   | 272   |     | Skeletal def             |
| 205           | 205         | 13  | 205           | 205         | 11  | 205   |     | Skeletal def             |
| 144           | 144         | 0   | 146           | 146         | 0   | 146   |     | τCH₃                    |

1 B3LYP/6-311++G(3df,3pd) and MP2/aug-cc-pVTZ methods were used in calculations. 2 The following scaling factors were used: 1.0/1.0 (below 750 cm⁻¹), 0.978/0.973 (in the range 750–2500 cm⁻¹), 0.964/0.953 (above 2500 cm⁻¹) for B3LYP/MP2 calculations, respectively. 3 The IR calculated intensities expressed in km mol⁻¹. 4 Abbreviations: υ, bond stretching; δ, bending, deformation in plane; ω, wagging; ρ, rocking; γ, out of plane bending, torsion; τ, twisting.

#### 2.3. Photolysis

The AMT molecules isolated in argon matrices have been irradiated using UV photons delivered by an OPO type tunable laser. The irradiations started at a radiation wavelength of 300 nm and proceeded with gradual decrease of the output wavelength. After each irradiation, an infrared spectrum of the matrix was taken. The observation of the photolysis began at 275 nm and led to the consumption of the AMT1 tautomer. However, the rate of the reaction was relatively low. Further irradiation at 265 nm led to a pronounced decrease of the intensity of the AMT1 bands and concomitantly to the growth of new bands with simultaneous appearance and increase of new bands in the spectra. The behavior of the most representative bands during photolysis is shown in Figure S4. These experiments allowed for interesting photo-transformations to be detected. Literature reports on photochemical rearrangements of thiazole, thiazole derivatives and five-membered ring heterocycles in general [13,16,19–22], allowed to propose the photoreaction pathways of 2-amino-4-methylthiazole. All proposed paths of photolysis of AMT are presented in Figure S5 and Figure S6. The structures of all photoproducts, named fp1-fp41, were optimized.
using the B3LYP/6-311++G(3df,3pd) method; their harmonic wavenumbers were also calculated and gathered in Table S2. The analysis of the most intense and characteristic bands for these molecules and comparing their position with new product bands appearing in the experimental spectrum allowed for rejection of the following structures from further considerations: fp4-fp10, fp16-fp21 and fp26-fp41. The photoproduct bands of AMT were identified and assigned to fp1, fp2, fp3, fp11-fp15 and fp22-fp25 structures. Scheme 2 presents the identified photoreaction pathways of AMT.

Scheme 2. The identified photoreaction pathways of AMT.

Ring-opening reactions by cleavage of the S1–C2 bond

Identification of new carbodiimides containing sulfur atom. Figure 3 and Figure S7 present the most characteristic regions of the IR spectrum where new bands can be observed. The most prominent photoproduct absorptions appeared in the 2300–2000 cm\(^{-1}\) region. The intense, broad band at ca. 2135 cm\(^{-1}\) was identified after 5 min of irradiation at 265 nm and its intensity grew during 180 min of photolysis at 265 nm and then slightly diminished up to 300 min of further irradiation, which could be the result of other AMT photoreactions leading to the sample decomposition in the studied matrices. The behavior of this band during photolysis is shown in Figure S8. This intense band is characteristic of structures containing the isocyano, cyano or carbodiimide group [16,21–24]. Considering the cleavage of the S1–C2 bond as the mechanism of initial photoproduct formation postulated for the thiazole molecule, similar reaction can be proposed for AMT. The experimental wavenumbers compared with calculated harmonic and anharmonic wavenumbers of the photoproduct molecules observed after photolysis of AMT/Ar matrix are gathered in Table 3. Optimized structures of the photoproduct molecules are presented in Figure 4. Anharmonic wavenumbers calculated for the identified photoproducts are collected in Table S3.
Figure 3. Selected regions of the AMT/Ar spectra: (a) after deposition at 15 K/10 K (black traces) and after 300 min irradiation at 265 nm (red traces), (b) B3LYP/6-311++G(3df,3pd) stick spectra of the identified photoproducts.

Table 3. Experimental and B3LYP/6-311++G(3df,3pd) calculated wavenumbers (cm\(^{-1}\)) and intensities (km mol\(^{-1}\)) of the photoproduct molecules observed after photolysis of AMT/Ar matrix.

| fp1 | Assignmt 1 |
|-----|------------|
| fp1a | fp1s |
| \(\nu\) experim | \(\nu\) harm | \(\nu\) anharm | \(\nu\) experim | \(\nu\) harm | \(\nu\) anharm | \(\nu\) harN-CN | \(\delta\) NH |
| 2134 br | 2217 (1338) | 2171 | 2142 br | 2221 (1332) | 2182 | \(\nu\) harN-CN |
| 887.5 | 922 (443) | 868 | 885.0 | 912 (406) | 864 | \(\delta\) NH |
| 644 br | 649 (24) | 641 | 631 (15) | 621 | \(\delta\) NCN |

| fp2 | Assignmt |
|-----|-----------|
| fp2a | fp2s |
| \(\nu\) experim | \(\nu\) harm | \(\nu\) anharm | \(\nu\) experim | \(\nu\) harm | \(\nu\) anharm | \(\nu\) harN-CN | \(\delta\) CH | \(\rho\) CH\(_3\) | \(\nu\) C-C(H\(_3\)) | \(\delta\) NH | \(\delta\) CCN |
| 2130 br | 2221 (1185) | 2158 | 2137 br | 2219 (949) | 2177 | \(\nu\) harN-CN |
| 1080.5 | 1085 (58) | 1071 | 1057.5 | 1077 (30) | 1058 | \(\delta\) CH |
| 1000.0 | 1025 (25) | 998 | 998.5 | 1012 (27) | 991 | \(\rho\) CH\(_3\) |
| 873.0 | 932 (312) | 855 | 898.0 | 927 (313) | 898 | \(\nu\) C-C(H\(_3\)) |
| | 595 (53) | 565 | 644 br | 655 (33) | 638 | \(\delta\) NH |

| fp3 | Assignmt |
|-----|----------|
| \(\nu\) experim | \(\nu\) harm | \(\nu\) anharm | \(\nu\) harN-CN | \(\delta\) NH | \(\delta\) CCN |
| 3443.5 | 3586 (91) | 3424 | \(\nu\) NH |
| 2160 br | 2233 (1035) | 2193 | \(\nu\) harN-CN |
| 1373.5 1375.5 | 1426 (48) | 1390 | \(\delta\) NH |
| 865.0 | 913 (217) | 856 | \(\delta\) CCN |

| fp15 | Assignmt |
|-----|----------|
| \(\nu\) experim | fp15syn/syn | fp15anti/syn | fp15syn/anti | \(\nu\) harN-CN | \(\delta\) NH | \(\nu\) C=N |
| 1618.0 | 1653 (175) | 1612 | 1652 (173) | 1602 | 1677 (226) | 1636 | \(\nu\) C=N |
| 1250.0 | 1255 (304) | 1210 | 1268 (266) | 1220 | 1262 (166) | 1223 | \(\delta\) NH |
| 1103.0 | 1130 (68) | 1101 | 1120 (52) | 1096 | 1092 (47) | 1066 | \(\nu\) C=N |

1 Abbreviations: \(\nu\), bond stretching; \(\delta\), bending, deformation in plane; \(\rho\), rocking.
After disruption of the S1–C2 bond the hydrogen atom can migrate from N6 to S1 giving N-(1-sulfanylprop-1-en-2-yl)carbodiimide (fp1) as photoproduct (see Scheme 2). The optimization resulted in two conformers of fp1 with E and Z conformation around C=C bond (fp1a and fp1s, respectively), with the small relative energy difference $\Delta E_{\text{ZPE}} = 3.9 \text{ kJ mol}^{-1}$. The comparison of the calculated spectral pattern of the fp1a and fp1s with the experimental spectrum allowed for assignment of 2134, 887.5 and 644 cm$^{-1}$ bands to fp1a and 2142, 885 cm$^{-1}$ band to fp1s products, respectively. When the hydrogen atom migrates from N6 to C5 after the cleavage of the S1–C2 bond the N-(2-methylthiiran-2-yl)carbodiimide molecule (fp3) can be formed. A careful analysis of the spectrum revealed bands of fp3 photoproduct at 3443.5, 2160, 1373.5, 1375.5 and 865.0 cm$^{-1}$. Additionally, several absorptions due to N-(1-thioxopropan-2-yl)carbodiimide (fp2) can be also observed. On the basis of calculations broad absorption at 2130 cm$^{-1}$ and 1080.5, 1000.0 and 873.0 cm$^{-1}$ bands were attributed to fp2a photoprodut, while 2137, 1263.0, 1057.5, 998.5, 898.0 and 644 cm$^{-1}$ bands were assigned to the fp2s structure. Both conformers of fp2 differ in position of the H–C=S group in relation to the carbodiimide group, and thus they vary in the values of the NCCS torsion angle ($-138.3^\circ$ for fp2a, $-3.9^\circ$ for fp2s), which gives a very small electronic energy difference $\Delta E_{\text{ZPE}} = 0.2 \text{ kJ mol}^{-1}$ for these structures.

Formation of cyanamide complexes. Cleavage of the S1–C2 bond followed by disruption of the N3–C4 bond (see Scheme 2) produces cyanamide molecule (fp11) and C(CH$_3$)=CH-S biradical which changes into 2-methylthiirene (fp12). Further photoreactions can produce 1-propyne-1-thiole (fp13) or methylthioketene (fp14). Indeed, the spectral analysis revealed several new vibrational features present in the typical regions of cyanamide (H$_2$N–C≡N) absorptions. The performed experiments allowed to distinguish three groups of bands, labeled as fp11-12, fp11-13 and fp11-14, (see Figures 3 and S7) which respond in different
ways to the photolysis process. The absorptions of group fp11-12 appeared after 2 min of irradiation at 265 nm as very weak bands, very slowly increasing its intensity up to 300 min of irradiation. The bands of group fp11-13 were observed after 1 min of irradiation at 275 nm, they visibly grew during irradiation at 265 nm. The absorptions of group fp11-14 appeared after 1 min of irradiation at 275 nm and grew clearly as well defined bands during irradiation at 265 nm. The absorptions attributed to group fp11-12 appeared at 3165, 2256.0, 2252.0 and 665.0 cm$^{-1}$. The bands due to group fp11-13 were observed at 3475.0, 3365 and 2260.0 cm$^{-1}$. The absorptions of group fp11-14 were identified at 3466.0, 3314.0, 2267.0, 1777.0 and 638.0 cm$^{-1}$. As will be discussed below the bands due to group fp11-12 can be assigned to the H$_2$N–C≡N + CH$_3$–CSC–H complex, the bands of group fp11-13 can be attributed to the H$_2$N–C≡N + CH$_3$C≡CSH complex and the bands of group fp11-14 to the H$_2$N–C≡N + S=C=C(H)CH$_3$ complex.

Formation of H$_2$N–C≡N···CH$_3$-CSC-H complex. The calculations resulted in two local minima on the potential energy surface of the cyanamide-2-methylthiirene system that correspond to the stable structures presented in Figure 5. In the more stable structure, fp11-12a ($\Delta E_{CP} = -30.8$ kJ mol$^{-1}$), the NH group of cyanamide serves as a proton donor toward the sulfur atom and as a proton acceptor for the CH$_3$ group of 2-methylthiirene. In turn, in the fp11-12b configuration, ($\Delta E_{CP} = -26.2$ kJ mol$^{-1}$), similar NH···S and CH···N bonds are present with the CH group of 2-methylthiirene interacting with cyanamide. The full sets of harmonic vibrational wavenumbers of the optimized structures are presented in Table S2. The comparison of the experimental spectra (bands assigned to the group fp11-12) with the calculated for the two structures evidences that the complex has the structure fp11-12a. In Table 4 the theoretical wavenumber shifts, $\Delta \nu_{\text{theor}} = (\nu_{\text{comp}} - \nu_{\text{mon}})_{\text{theor}}$ for both optimized structures are compared with the experimental ones, $\Delta \nu_{\text{exp}} = (\nu_{\text{comp}} - \nu_{\text{mon}})_{\text{exp}}$. The experimental wavenumber for the H$_2$N–C≡N molecule was taken from reference [24,25]. The formation of the structure fp11-12a in the matrix is reflected by the appearance of the bands due to perturbed $\nu_s$NH$_2$, $\nu$C≡N and $\omega$NH$_2$ modes of cyanamide. The 3165 and 2256.0, 2252.0 cm$^{-1}$ absorptions assigned to the perturbed NH$_2$ symmetric stretching and C≡N stretching vibrations are 235 and ca. 10 cm$^{-1}$ red shifted and the 665.0 cm$^{-1}$ band attributed to the NH$_2$ wagging mode is 251 cm$^{-1}$ blue shifted with the corresponding modes of the H$_2$N–C≡N monomer, respectively. As one can see in Table 4 the observed pattern of the band shifts is better reflected by the shifts of the more stable complex of structure fp11-12a, particularly the observed shift of the $\nu$NH$_2$ band ($\Delta \nu_{\text{obs}} = +251$ cm$^{-1}$) shows better agreement with the one calculated for this structure ($\Delta \nu_{\text{calc}} = +206$ cm$^{-1}$). Unfortunately, none of the 2-methylthiirene band was identified in the spectra but it seems to be justified by the low intensities of the absorptions of this molecule: the calculated intensity of the most intense band of CH$_3$-CSC-H does not exceed 40 km mol$^{-1}$ for both the monomer and the complexes. So, most probably absorptions of 2-methylthiirene are overlapped by other bands.

![Figure 5. B3LYP/6-311++G(3df,3pd) optimized structures of the H$_2$N–C≡N···CH$_3$-CSC-H complex. The selected bond distances (in Å) and interaction energies $\Delta E_{CP}$ (in kJ mol$^{-1}$) are given.](image-url)
Table 4. Wavenumbers and wavenumbers shifts ($\Delta \nu = (\nu_{\text{comp}} - \nu_{\text{mon}})$) (in cm$^{-1}$) calculated and observed for the H$_2$N–C≡N···CH$_3$–CSC–H complex in solid argon$^1$.

| Assignment     | Experimental | Calculated Harmonic | Assignment |
|----------------|--------------|---------------------|------------|
|                | v            | fp11-12a            | fp11-12b   |
| $\nu_{sNH_2}$ | 3165 br      | 3279                | 3294       | $-235$     |
|                | 2256.0       | 2335                | 2333       | $-8$       |
|                | 2252.0       |                     | $-16$      | $-12$      |
| $\omega_{NH_2}$| 665.0        | 695                 | 672        | $+251$     |

$^1$ In all calculations the B3LYP/6-311++G(3df,3pd) method was used. The experimental wavenumbers for the H$_2$N–C≡N monomer were taken from refs. [24,25].

**Formation of the H$_2$N–C≡N···CH$_3$C≡CSH complex.** Figure 6 presents two structures corresponding to the stationary points calculated for the cyanamide complex with 1-propyne-1-thiol. In both structures the NH group of cyanamide acts as a weak proton donor towards the C≡C bond. In structure fp11-13a ($\Delta E_{\text{CP}} = -15.1$ kJ mol$^{-1}$) H$_2$N–C≡N plays also a role of proton acceptor of the CH$_3$ group of CH$_3$CCHS and in slightly more stable structure fp11-13b ($\Delta E_{\text{CP}} = -15.5$ kJ mol$^{-1}$) 1-propyne-1-thiol forms the SH···N bond with cyanamide. The full sets of vibrational wavenumbers of the optimized structures are presented in Table S2. In Table 5 the theoretical wavenumber shifts for both optimized structures are compared with the experimental ones. The comparison of the experimental spectra (bands assigned to the group fp11-13) with the calculated for the two structures indicates that the complex trapped in the matrix has the structure fp11-13a. Three distinct bands due to the H$_2$N–C≡N···CH$_3$C≡CSH complex appeared in the spectral regions of cyanamide: at 3475.0 cm$^{-1}$, ($\Delta \nu_{\text{exp}} = -11$ cm$^{-1}$), in the $\nu_{sNH_2}$ region; at 3365 cm$^{-1}$, ($\Delta \nu_{\text{exp}} = -35$ cm$^{-1}$), in the vicinity of $\nu_{asNH_2}$ and at 2260.0 cm$^{-1}$, ($\Delta \nu_{\text{exp}} = -4$ cm$^{-1}$) in the region of the $\nu_{C\equiv N}$ vibration, respectively. The observed wavenumber shifts match reasonably well the values calculated for the fp11-13a structure: $\Delta \nu_{\text{calc}} = -22$, $-54$ and $-8$ cm$^{-1}$, respectively. In spite of the very careful spectral analysis, no bands attributed to the perturbed 1-propyne-1-thiole molecule were identified, which may be caused by the overlapping of its bands with other absorptions as the calculations predict low intensity bands for this compound (see Table S2).

Figure 6. B3LYP/6-311++G(3df,3pd) optimized structures of the H$_2$N–C≡N···CH$_3$C≡CSH complex. The selected bond distances (in Å) and interaction energies $\Delta E_{\text{CP}}$ (in kJ mol$^{-1}$) are given.
Table 5. Wavenumbers and wavenumbers shifts ($\Delta \nu = (\nu_{\text{comp}} - \nu_{\text{mon}})$) (in cm$^{-1}$) calculated and observed for the H$_2$N–C≡N···CH$_3$C≡CSH complex in solid argon $^1$.

| Assignment | Experimental | Calculated Harmonic |
|------------|--------------|---------------------|
|            | $\nu$ | $\Delta \nu$ | $\nu$ | $\Delta \nu$ | $\nu$ | $\Delta \nu$ |
| $\nu_{\text{as}}$NH$_2$ | 3475.0 | -11 | 3615 | -22 | 3607 | -30 |
| $\nu_{\text{s}}$NH$_2$ | 3365 sh | -35 | 3493 | -54 | 3471 | -76 |
| $\nu_{\text{C}}$≡N | 2260.0 | -4 | 2343 | -8 | 2341 | -10 |

$^1$ In all calculations the B3LYP/6-311++G(3df,3pd) method was used. The experimental wavenumbers for the H$_2$N–C≡N monomer were taken from ref. [24,25].

Formation of the H$_2$N–C≡N···S=C=C(H)CH$_3$ complex. The calculations resulted in one global minimum on the potential energy surface of the cyanamide complex with methylthioketene ($\Delta E_{CP} = -15.6$ kJ mol$^{-1}$) presented in Figure 7. In this structure the NH···S bond is formed and a weak interaction of the CH$_3$ group of S=C=C(H)CH$_3$ with the nitrogen atom of CN group of H$_2$N–C≡N is also present. In Table 6 the theoretical wavenumber shifts for the fp11-14 structure are compared with the experimental ones. The cyanamide and methylthioketene monomer wavenumbers were taken from references [24–26]. The full set of harmonic vibrational wavenumbers of the optimized structure is presented in Table S2. In the region of the $\nu_{\text{C}}$≡N vibration a distinct band grows after matrix irradiation (at 2267.0 cm$^{-1}$) which suggests formation of one structure. The suggestion is supported by appearance of one intense band for the $\nu_{\text{as}}$C≡S vibration of S=C=C(H)CH$_3$ (1777.0 cm$^{-1}$) and for the $\nu_{\text{as}}$NH$_2$ and $\nu_{\text{s}}$NH$_2$ vibrations of H$_2$N–C≡N (3466.0 and 3314.0 cm$^{-1}$, respectively). The 638.0 cm$^{-1}$ band due to the perturbed $\gamma$CH vibration of methylthioketene was also observed. As one can see in Table 6 the experimental shifts are in accord with the theoretical ones both for the cyanamide and for methylthioketene molecules.

Figure 7. B3LYP/6-311++G(3df,3pd) optimized structure of the H$_2$N–C≡N···S=C=C(H)CH$_3$ complex. The selected bond distances (in Å) and interaction energies $\Delta E_{CP}$ (in kJ mol$^{-1}$) are given.
Table 6. Wavenumbers and wavenumbers shifts ($\Delta\nu = (\nu_{\text{comp}} - \nu_{\text{mon}})$) (in cm$^{-1}$) calculated and observed for the H$_2$N–C≡N⋯S=C=C(H)CH$_3$ complex in solid argon $^1$.

| Assignment          | Experimental | Calculated Harmonic | Assignment          |
|---------------------|--------------|---------------------|---------------------|
| $\nu_{\text{as}}$NH$_2$ | 3466.0       | 3610                | $\nu_{\text{as}}$NH$_2$ | 3314.0       | 3453                | $\nu_{\text{as}}$C≡C=S |
| $\nu_s$NH$_2$       | 3314.0       | 3453                | $\gamma$C≡N         | 2267.0       | 2342                | $\gamma$CH           |
| $\nu_{\text{as}}$C≡N | 1777.0       | 1837                | $\nu_C=C=S$         | 638.0        | 656                 |

In all calculations the B3LYP/6-311++G(3df,3pd) method was used. The experimental wavenumbers for the monomers were taken from refs. [24–26].

Ring-opening reactions by cleavage of the S1–C5 bond

Identification of sulfur compound with the azireno group. After cleavage of the S1–C5 bond the hydrogen atom can migrate from N6 to S1 giving 2-methyl-1H-azirene-1-carbimidothioic acid (fp15) as photoproduct (see Scheme 2). As one can see in Figure S7 only three absorptions were observed for fp15, they were identified at 1618.0, 1250.0 and 1103.0 cm$^{-1}$. These bands appeared after 2 min of irradiation at 270 nm and their intensity grew during 300 min of photolysis at 265 nm. The performed optimization resulted in four stable structures of fp15 with different conformations of both SH and NH groups around the C=N double bond. The relative energy differences ($\Delta E_{\text{ZPE}}$) for these structures are: 0.0, 2.4, 3.2 and 7.2 kJ mol$^{-1}$ for the fp15$^{\text{anti(SH)}}$/fp15$^{\text{syn(NH)}}$, fp15$^{\text{syn/syn}}$, fp15$^{\text{syn/syn}}$ and fp15$^{\text{anti/anti}}$ forms, respectively. The comparison of the experimental bands with the absorptions calculated for the conformers of fp15 suggests that the $\text{syn/syn}$ form of 2-methyl-1H-azirene-1-carbimidothioic acid is present in the matrix (see Table 3).

Formation of propyne complexes. Cleavage of the S1–C5 bond followed by disruption of the N3–C4 bond (see Scheme 2) produces the propyne molecule (fp22) and S-C(NH$_2$)≡N biradical which changes into 3-aminethiazirene (fp23). Further photoreactions can produce thiaziridin-3-imine (fp24) or N-sulfanylcarbodiimide (fp25). Based on the spectra calculated for the possible propyne complexes, the complex between propyne and N-sulfanylcarbodiimide (fp22-25) was identified in the AMT/Ar matrix after photolysis since the spectral analysis revealed several new vibrational features characteristic for the carbodiimide modes. However, the formation of the 3-aminethiazirene molecule (in fp22-23) in the matrix cannot be completely ruled out due to the presence of a broad shoulder at 1792 cm$^{-1}$ on an intense 1777 cm$^{-1}$ band observed for the fp11-14 complex (see the discussion above). The 1792 cm$^{-1}$ absorption might be assigned to the strongest band of the $\nu_{\text{as}}$NCN vibration in the fp22-23 complex for which the calculated intensity exceeds 200 km mol$^{-1}$.

Formation of the H$_3$C–C≡CH⋯HN=C=S–SH complexes. The calculations resulted in two local minima on the potential energy surface of the propyne—N-sulfanylcarbodiimide system that correspond to the stable structures presented in Figure 8. In the more stable structure, fp22-25a ($\Delta E_{\text{CP}} = -9.8$ kJ mol$^{-1}$), the NH group of N-sulfanylcarbodiimide serves as a proton donor toward the C≡C triple bond and as a proton acceptor for the CH$_3$ group of propyne. In turn, in the fp22-25b configuration, ($\Delta E_{\text{CP}} = -6.0$ kJ mol$^{-1}$), the SH group plays the role of proton donor toward the C≡C triple bond and the nitrogen atom is the proton acceptor of the CH$_3$ group of propyne. The full sets of harmonic vibrational wavenumbers of the optimized structures are presented in Table S2. The propyne monomer wavenumbers were taken from references [27,28]. The experimental spectra of HN=C=N–SH have not been reported so far and the $\Delta\nu_{\text{exp}}$ values for this molecule could not be estimated. So, in this case the theoretical anharmonic wavenumbers of HN=C=N–SH were taken to $\Delta\nu_{\text{exp}}$ calculations instead of the missing monomer experimental wavenumbers. The anharmonic wavenumbers calculated for HN=C=N–SH are presented in Table S3.
Table 7 the theoretical wavenumber shifts for both optimized structures are compared with the experimental ones. As one can see, a reasonable agreement can be noticed between the observed shifts and calculated for the structure fp22-25a. Some inconsistency in these values may be due to the lack of experimental wavenumbers for N-sulfanylcarbodiimide. The anharmonic wavenumber of the $\nu_{as}$NCN vibration in the monomer (2163 cm$^{-1}$) is much higher than the identified wavenumber of the perturbed $\nu_{as}$NCN vibration in the complex (2117.5 cm$^{-1}$). Most likely the anharmonicity of this vibration is greater than the calculations predict, which may be due to the coupling of the stretching vibrations of the NCN group and the NH group present in the immediate vicinity. A similar situation occurs for fp1 or fp2 photoproducts, for which the $\nu_{as}$NCN vibration bands were observed at ca. 2135 cm$^{-1}$, and the calculations predicted the anharmonic wavenumbers between 2182 and 2158 cm$^{-1}$ for this mode. The 898 cm$^{-1}$ band was only tentatively assigned for the complex since this wavenumber coincides with the $\delta$NH bands in other carbodiimide molecules present in the matrix. The 635.0 cm$^{-1}$ band due to the perturbed $\delta$CH/$\gamma$CH vibration of propyne was also observed. As can be seen in Figure 3, there is some evidence that the complex of configuration fp22-25b is also present in the matrix: the 3400.0 and 2112.0 cm$^{-1}$ bands can be tentatively assigned to the perturbed $\nu$NH and $\nu_{as}$NCN vibrations in this complex, respectively.

![Figure 8. B3LYP/6-311++G(3df,3pd) optimized structure of the H$_3$C–C≡CH···HN=C=N–SH complex. The selected bond distances (in Å) and interaction energies $\Delta E^{CP}$ (in kJ mol$^{-1}$) are given.](image)

**Table 7.** Wavenumbers and wavenumbers shifts ($\Delta \nu = (\nu_{\text{comp}} - \nu_{\text{mon}})$) (in cm$^{-1}$) calculated and observed for the H$_3$C–C≡CH···HN=C=N–SH complex in solid argon $^1$.

| H$_3$C–C≡CH···HN=C=N–SH | Assignment |
|--------------------------|------------|
| **Experimental** | **Calculated Harmonic** | |
| | fp22-25a | fp22-25b | |
| $\nu$ | $\Delta \nu$ | $\nu$ | $\Delta \nu$ | $\nu$ | $\Delta \nu$ |
| 3337.0 | -40 | 3464 | -93 | 3555 | -2 | $\nu$NH |
| 2117.5 | -44.5 | 2211 | -1 | 2208 | -4 | $\nu_{as}$NCN |
| 898 tentatively | +28 | 976 | +31 | 949 | +4 | $\delta$NH |
| 635 | +4 | 675 | +15 | 668 | +2 | $\delta$CH |

$^1$ In all calculations the B3LYP/6-311++G(3df,3pd) method was used. The experimental wavenumbers for the propyne monomer were taken from ref. [27,28]; for the HN=C=N–SH monomer the anharmonic wavenumbers were used for $\Delta \nu_{\text{exp}}$ calculations (Table S3). $^2$ The magnitude of this shift is explained in the text.
3. Materials and Methods

2-amino-4-methylthiazole was commercially available (98%, Sigma-Aldrich, Merck, KGaA, Darmstadt, Germany). The AMT/Ar matrices were prepared by passing of high purity argon through the glass U-tube with AMT kept at room temperature and situated outside the cryostat chamber. The AMT/Ar gaseous mixtures were deposited onto a cold CsI window kept at a temperature 15 K/10 K (deposition/measurement) by means of a closed cycle helium refrigerator (ARS-2HW, APD-Cryogenics). Infrared spectra were recorded between 4000–400 cm$^{-1}$ in a transmission mode with a resolution of 0.5 cm$^{-1}$ by means of a Nicolet iS50 FTIR spectrometer equipped with a liquid N$_2$ cooled MCT detector.

After the infrared spectra of the initially deposited matrices were recorded the samples were irradiated with the tunable UV radiation provided by the frequency doubled signal beam of a pulsed (7 ns) optical parametric oscillator Vibrant 355 (Opotek Inc., Carlsbad, CA, USA) (repetition rate 10 Hz, average pulse energies ~7.0 mJ (325 nm) and ~2.5 mJ (250 nm)) pumped with a pulsed Nd:YAG laser (Quantel).

All calculations were performed with the Gaussian 09 suite of programs [29]. Structures of the minima and transition states were optimized at the B3LYP/6-311++G(3df,3pd) level of theory. On the basis of the force constant matrixes calculated at the same level, harmonic frequencies were evaluated as well as zero-point vibrational energy (ZPE) corrections and the corresponding thermodynamic functions. The transition states linking the tautomers were localized on the PES and verified by the presence of one imaginary vibrational frequency. The corresponding minima were connected to each TS by following intrinsic reaction coordinate (IRC). For the most stable tautomer AMT1 the structure and vibrational spectrum were also calculated at the MP2/aug-cc-pVTZ level. In order to account for anharmonic effects, the calculated vibrational frequencies were scaled by a factor of 1.0/1.0 (for wavenumbers in the range from 400 to 750 cm$^{-1}$), 0.978/0.973 (for wavenumbers in the range from 750 to 2500 cm$^{-1}$) or 0.964/0.953 (for wavenumbers above 2500 cm$^{-1}$) for B3LYP/MP2 calculations, respectively. The vibrational spectra were simulated using SYNSPEC program [30]. Anharmonic wavenumbers were calculated for the selected monomeric species at the B3LYP/6-311++G(3df,3pd) level of theory. The structures of the complexes were also optimized at the B3LYP level and their binding energies were corrected by the Boys-Bernardi full counterpoise procedure (CP) [31]. According to the reviewer’s suggestion, we re-computed the electronic energy values at the B3LYP-D3(BJ) level to get $\Delta E_{\text{CP}}$ (in kJ mol$^{-1}$) for the structures of complexes obtained after optimization at B3LYP/6-311++G(3df,3pd). The new energy values are presented in Table S4, Supplementary Materials. As can be seen, introduction of D3 dispersion correction to calculations reduces the binding energy in complexes, but overall the order of the structures in terms of energetics does not change.

4. Conclusions

Matrix isolation FT-IR spectroscopy, supported by B3LYP/6-311++G(3df,3pd) calculations, allowed for the first time to observe and characterize spectroscopically the AMT1 tautomer of 2-amino-4-methylthiazole which owes its stability to aromaticity and effective stabilization of the five-membered ring by C=C and C=N double bonds. The narrow band UV irradiation technique was applied to study the photochemistry of 2-amino-4-methylthiazole. There are two main mechanisms that initiate photochemical reactions. The major ring-opening photoreactions caused by the cleavage of the S1–C2 bond lead to the formation of C(NH$_2$)=N–C(CH$_3$)=CH–S biradical. The hydrogen atom of the NH$_2$ group of the radical migrates to S1 to form N-(1-sulfanylprop-1-en-2-yl)carbodiimide (fp1a, fp1s) or to C5 to produce N-(2-methylthiiran-2-yl)carbodiimide (fp3); subsequent photochemical reactions may transform both fp1, fp3 photoproducts into N-(1-thioxopropan-2-yl)carbodiimide (fp2a, fp2s) as shown in the Scheme 2. Cleavage of the S1–C2 bond followed by disruption of the N3–C4 bond produces the cyanamide molecule and the ‘C(CH$_3$)=CH–S‘ biradical that transforms into 2-methylthiirene (fp12); further photoreactions can produce 1-propyne-1-thiole (fp13) or methylthioketene (fp14).
The minor ring-opening photoreaction is caused by the cleavage of the S1–C5 bond upon UV irradiation. The hydrogen atom migration from N6 to S1 in S–(NH2)=N–C(CH3)=CH forms 2-methyl-1H-azirene-1-carbimidothioic acid (fp15) as photoproduct. Cleavage of the S1–C5 bond followed by disruption of the N3–C4 bond produces the propyne molecule and S–C(NH2)=N biradical that transforms into 3-aminethiazirene (fp23). Further photoreactions produce N-sulfanylcarbodiimide (fp25). As the above studies showed, apart from new photoproduct molecules obtained in the AMT photolysis process, several molecular complexes were also identified as photoproducts. As the two co-products of photodissociation channel of a molecule are trapped in one cage they may form a complex. The fact why most of the co-products formed in the same matrix cage create complexes but some stay as monomers was discussed widely in our earlier papers [32–35]. It is worth mentioning that cyanamide [36] and propyne [37] molecules have been identified in space, so their molecular complexes presented in this paper may have potential significance for astrophysics, astrochemistry or even astrobiology.

Supplementary Materials: The following supporting information can be downloaded online. Table S1: Selected structural parameters calculated for 2-amino-4-methylthiazole (AMT) tautomers at the B3LYP/6-311++G(3df,3pd) level of theory; Table S2: Harmonic wavenumbers, scaled wavenumbers (cm\(^{-1}\)) and intensities (km mol\(^{-1}\)) of 2-amino-4-methylthiazole photoproducts calculated at the B3LYP/6-311++G(3df,3pd) level of theory; Table S3: Anharmonic and harmonic wavenumbers (in cm\(^{-1}\)) calculated for the fp1, fp2, fp3, fp15, fp25, fp23 molecules at the B3LYP/6-311++G(3df,3pd) level of theory; Table S4: Binding energy and CP-corrected energy (in kJ mol\(^{-1}\)) calculated at B3LYP-D3(BJ) for the structures of complexes obtained after optimization at B3LYP/6-311++G(3df,3pd); Figure S1: B3LYP/6-311++G(3df,3pd) optimized structures of the tautomers of 2-amino-4-methylthiazole; Figure S2: B3LYP/6-311++G(3df,3pd) optimized structures of the transition states; Figure S3: Comparison of the AMT/Ar matrix spectrum with the spectra calculated for all of the tautomers of AMT molecule at the B3LYP/6-311++G(3df,3pd) level of theory; Figure S4: The 2300–2000 cm\(^{-1}\) region in the AMT/Ar matrix spectra during irradiation at wavelengths between 295 and 265 nm; Figure S5: Photoreaction pathways of AMT. Part 1; Figure S6: Photoreaction pathways of AMT. Part 2; Figure S7: 3600–2900 cm\(^{-1}\) (a), 2300–1200 cm\(^{-1}\) (b), 1200–420 cm\(^{-1}\) (c) regions of the spectra of AMT/Ar matrix after deposition (bottom spectra) and after 300 min irradiation at 265 nm of the deposited matrix (upper spectra); Figure S8: The 2135 cm\(^{-1}\) band in the AMT/Ar matrix spectra during irradiation at 265 nm.

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30. Irikura, K.K. Multiphoton Absorption Strengths from Simple Ab Initio Wavefunctions; National Institute of Standards and Technology: Gaithesburg, MD, USA, 1995.

31. Boys, S.F.; Bernardi, F. The Calculation of small molecular interactions by the differences of separate total energies. Some procedures with reduced errors. *Mol. Phys.* **1970**, *19*, 553–566. [CrossRef]

32. Sałdyka, M.; Mielke, Z. Photodecomposition of formohydroxamic acid. Matrix isolation FTIR and DFT studies. *Phys. Chem. Chem. Phys.* **2003**, *5*, 4790–4797. [CrossRef]

33. Sałdyka, M. Photodecomposition of N-hydroxyurea in argon matrices. FTIR and theoretical studies. *RSC Adv.* **2013**, *3*, 1922–1932. [CrossRef]

34. Sałdyka, M.; Mielke, Z. Photochemistry of acetohydroxamic acid in solid argon: FTIR and theoretical studies. *J. Phys. Chem. A* **2018**, *122*, 60–71. [CrossRef]

35. Sałdyka, M.; Coussan, S. Infrared spectra and photodecomposition of benzohydroxamic acid isolated in argon matrices. *J. Mol. Struct.* **2020**, *1219*, 128506. [CrossRef]

36. Ligterink, N.F.W.; El-Abd, S.J.; Brogan, C.L.; Hunter, T.R.; Remijan, A.J.; Garrod, R.T.; McGuire, B.M. The Family of Amide Molecules toward NGC 6334I. *Astrophys. J.* **2020**, *901*, 1–23. [CrossRef]

37. Fuente, A.; Garcia-Burillo, S.; Gerin, M.; Teyssier, D.; Usero, A.; Rizzo, J.R.; de Vicente, P. Photon-dominated Chemistry in the Nucleus of M82: Widespread HOC+ Emission in the Inner 650 Parsec Disk. *Astrophys. J.* **2005**, *619*, L155–L158. [CrossRef]