Photocatalytic properties of zirconium oxide–zinc oxide nanoparticles synthesised using microwave irradiation

Olga Długosz1 · Krzysztof Szostak1 · Marcin Banach1

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Abstract

ZrO2–ZnO nanoparticles with different mass concentrations of ZrO2 (1%, 5%, 10%, 20%) were prepared using a two-stage precipitation method with microwave irradiation. The ZrO2–ZnO with 10% of ZrO2 had the highest photocatalytic activity. The obtained material was characterised using XRD, which confirmed a high crystal structure in the synthesised material, and transmission electron microscopy (TEM) analysis, which depicted that micro and nano needle-shaped particles had been obtained and that irregularly shaped nanoparticles were present on the surface of those particles. TEM–EDX analysis confirmed the presence of both ZnO and ZrO2 in the product. FT-IR analyses showed that the positions of peaks related to Zn–O and Zr–O absorption bands did not change in ZrO2–ZnO NPs. The initial concentration of Methylene Blue (MB), the pH of the solution, and the mass of the photocatalyst were investigated to determine the photocatalytic efficiency of the material and the degree of removal of the MB. The highest efficiency (97%) was obtained in the following conditions: 30 mg/dm3 of the MB solution, pH 9 and 70 mg ZrO2–ZnO after 30 min of reaction time under UV irradiation.

Keywords Zinc oxide · Zirconium oxide · Nanoparticles · Photocatalytic properties · Methylene Blue · Microwave irradiation

Introduction

Oxide semiconductor nanomaterials are gaining increasing interest due to their attractive functional properties related to their small particle size and hence large active surface (Dal Santo et al. 2018; Sowik et al. 2019). Among these features, the photocatalytic properties are of particular interest. This is due to the fact that these kinds of nanomaterials can be used in many fields such as wastewater treatment or coating materials (Ruiz-Hitzky et al. 2019).

Zinc oxide, especially on a nanometric scale, is a low cost, environmentally friendly with potentially high photocatalytic properties (Lakshmi et al. 2012; Lu et al. 2017). ZnO has a broader energy gap (3.2 eV) compared to TiO2, which directly contributes to lower photocatalytic activity (Nagao et al. 2016). The modification of the structure of ZnO nanoparticles (ZnO NPs) with the addition of metal and metal oxide nanoparticles affects the distance between the conduction band and the valence band, which may improve the photocatalytic properties of the material (Karuppaiah et al. 2019).

The main limitations of the photocatalytic properties of ZnO nanoparticles are the relatively broad energy gap and the rapid recombination of electron–hole pairs, which limits the photocatalytic activity of ZnO NPs. It has been noted that the addition of metal oxides improves the photocatalytic properties of, among others, ZnO, TiO2, MgO, WO3 and SnO2 in both UV and visible light (Thirumalai et al. 2017; Reszczężyńska et al. 2015; Al-Hamdi et al. 2015; Kokorin et al. 2018). One material that enhances the photocatalytic properties of ZnO NPs is zirconium oxide (ZrO2).

Zirconium oxide belongs to the group of semiconductor materials. It usually occurs in an amorphous form with a relatively wide energy gap (~ 5.25 eV (Oluwabi et al. 2018)). Pure ZrO2 has very limited photocatalytic properties, but the addition of the oxide in the other oxides, including TiO2, SnO2, ZnO has beneficial effect on their photocatalytic activity (Aghabeygi et al. 2017; Pirzada et al. 2015). One of the mechanisms to improve properties of photocatalytic oxides is the modification their structure by creating numerous
surface defects, developing the surface of the deposition of other oxide nucleation, thus supporting the electron transfer processes accompanying the photocatalytic process (Uribe López et al. 2019). In the case of the ZrO₂–ZnO system, in a two-stage process, ZnO particles are obtained on the surface of the amorphous ZrO₂ particles, which may have a different structure than those obtained in a single-stage process (Ibrahim 2015). Additionally, the structure of ZrO₂ plays an important role in the reduction of the electron–hole recombination of doped systems (Gurushantha et al. 2017). The replacement of Zn²⁺ with Zr⁴⁺ disturbs the charge balance. The disturbed charge balance contributes to the adsorption of more hydroxide ions at the catalyst surface, which additionally blocks the recombination of electron–hole pairs (Karuppaiah et al. 2019). Aghabeygi and Khademi-Shamami (2018) investigated ZnO/ZrO₂ nanocomposites prepared with ultrasonic irradiation. The ZnO nanocomposite was excited by the transfer of the electrons promoted from the conduction band of ZnO–ZrO₂ and simultaneously transferred the holes from the valence band of ZrO₂ to the ZnO.

There are known methods for obtaining zinc oxide nanoparticles modified by ZrO₂ or other compounds. The great limitations in the proposed tests are the long time of the photodegradation process and the low concentration of dyes used in the processes. Sapawe et al. received ZnO–ZrO₂–HY zeolite nanocomposite, for the removal of Methylene Blue. The dye removal efficiency was 99.7%, but the dye concentration was 10 mg/dm³, and the process was carried out for 90 min (Sapawe et al. 2013). Chen et al. obtained ZnO modified with SnO₂ nanoparticles to remove various dye wastewaters. By increasing the initial concentration of Congo Red from 20 to 40 ppm, the photodegradation time increased from 40 to 80 min while reducing the efficiency from 96.8 to 89.3% (Chen et al. 2015). These are the reasons why it is necessary to dilute the dye solution, use larger masses of photocatalyst or extend the processing time.

Metal oxide nanoparticles can be obtained by hydrothermal processes usually carried out by several hours. Qui et al. they received doped ZrO₂–ZnO photocatalyst; however, the process was carried out under pressure reactor for 12 h (Qiu et al. 2018). Aghabeygi and Khademi-Shamami developed an alternative method of obtaining ZnO/ZrO₂ nanocomposite using ultrasound, but after 30 min of ultrasonication, the material was calcined (Aghabeygi and Khademi-Shamami 2018). Insertion of the energy into the system using microwave radiation reduces the synthesis time. In addition, the amount of energy in the process allows obtaining products with a high degree of crystallinity, so that the calcining process can be omitted.

This study aimed to prepare ZrO₂–ZnO nanoparticles using a precipitation method with microwave irradiation. The photocatalytic properties of the obtained material were investigated for the degradation of Methylene Blue (MB) solution under UV irradiation. The effects of the mass concentration of ZrO₂ in ZnO–ZrO₂ NPs, the initial concentration of MB, the pH of the solution, and the mass of photocatalyst were investigated.

**Experimental**

**Materials**

Zinc sulphate ZnSO₄·5H₂O (Sigma Aldrich), zirconyl chloride ZrOCl₂·8H₂O (Sigma Aldrich) and sodium carbonate Na₂CO₃ (Sigma Aldrich) were used in the synthesis of the photocatalytic material. MB was purchased from Sigma Aldrich and used to study the photocatalytic properties of ZrO₂–ZnO. All chemicals were of analytical grade and were used without further purification.

**Synthesis of ZrO₂–ZnO**

A precipitation method was used to produce ZrO₂–ZnO. Two-step synthesis was carried out in a microwave reactor. First, Zr(OH)₄ was synthesised by mixing 20 ml 1 M ZrOCl₂·8H₂O with 20 ml 2 M Na₂CO₃. The suspension was transferred to a Teflon vessel and it was heated in a MAGNUM II microwave reactor (Ertec, Poland) for 10 min at 150 °C. The obtained zirconium oxide(IV) was filtered and dried at 70 °C for 24 h. In the next step, solid ZrO₂ nanoparticles in an amount corresponding to fixed percentage composition, 20 ml of 1 M ZnSO₄·5H₂O and 10 ml of 2 M Na₂CO₃ were added to the beaker. After mixing, the suspension was transferred to a Teflon vessel and heated in the microwave reactor for 10 min at 180 °C. The obtained product was filtered and dried at 70 °C for 24 h. The effect of the ZrO₂ mass on the photocatalytic properties of the final product was also established.

**Instrumental analysis**

The structure of ZrO₂–ZnO before and after the photocatalytic reaction was examined using transmission electron microscopy (TEM) with energy dispersive X-ray analysis (EDX) and elemental mapping using a Tecnai Transmission Electron Microscope, F20 X-Twin, FEI Europe. The Fourier Transform Infrared Spectroscopy (FTIR) method ( Nicolet 380) was helpful for determining the degree of sorption of the dye as well as its decomposition under the influence of UV. XRD analysis (Philips X’Pert camera with monochromator PW 1752/00 CuKα) determined the degree of crystallinity of ZrO₂, ZnO and ZrO₂–ZnO.
Optical studies

The spectrum of MB and its concentration were examined by UV–Vis spectroscopy (Rayleigh UV-1800 spectrophotometer). The calibration curve was prepared at 664 nm and was made in triplicate (Fig. 1).

The absorption, which corresponds to the energy difference between the top of the valence band and the bottom of the conduction band, is defined to designate the band gap using the UV–Vis method (Mexicana De Fisica et al. 2007). The value of the band gap ($E_g$) energy was obtained using the following equation:

$$E_g = \frac{h \cdot c \cdot 10^{19}}{\lambda} \text{[eV]},$$

where $h$—Plank’s constant = $6.626 \cdot 10^{-34}$ Js, $c$—speed of light = $3.0 \cdot 10^8$ m/s, $\lambda$—cut off wavelength/absorption edge in nm.

Photocatalytic degradation experiments

The photodegradation of MB using ZrO$_2$–ZnO was studied using batch mode experiments. This was done by studying the effect of initial concentration of dye, material dose and pH of the solution. The photocatalytic activities of the photocatalyst were investigated by measuring the degradation of MB aqueous solutions under UV light irradiation using a UV lamp (wavelength 360 nm). A certain amount of the material was added while mixing 30 ml of the solution of dye with a magnetic stirrer for a specified period of time. The amounts of the degraded dye $R_D$ (mg/g) (2) and efficiency of photocatalytic degradation of MB (3) were calculated from the following dependences:

$$R_D \text{ (mg/g)} = \frac{(C_0 - C_t)V}{m},$$

$$E(\%) = \frac{(C_0 - C_t)}{C_0} \cdot 100,$$

where $C_0$ and $C_t$ are the initial and final concentrations of the dye solution (mg/dm$^3$) at time $t$, $V$ is the solution volume (dm$^3$), and $m$ is the mass of the material (g).

Statistical analysis

The initial concentration of the dye was tested using a UV–Vis spectrophotometer in the range of 20–60 mg/dm$^3$. The effect of the mass of the photocatalytic deposit was examined in the range of 10–50 mg and the effect of the pH of the solution was tested in the range from 3 to 11. The pH of the solution was adjusted with 15% HNO$_3$ or 0.1 M NaOH. A central composite design (CCD) plan was used to examine the effect of parameters on the photocatalysis process and determine the most favourable parameter values (Table 1). As the output parameters, the degree of dye removal per 1 g of the deposit and photodegradation efficiency were used. The research results were developed using the program STATISTICA (version 10.0).

Results and discussion

Selection of the final deposit based on the photocatalytic activity of the product

In the first stage, to obtain the material with the highest photocatalytic efficiency, the most favourable mass concentration of ZrO$_2$ in the ZrO$_2$–ZnO nanoparticles was determined. For this purpose, a various mass of ZrO$_2$ powder and
a constant volume of ZnSO₄ and Na₂CO₃ solutions were added to the beaker. The mixture was transferred to a Teflon vessel and heated using microwave irradiation for 10 min at 180 °C. After the process, the product was filtered, washed and dried. The photodegradation efficiency and the amount of degraded dye of ZrO₂–ZnO with a mass concentration of ZrO₂ equalling 1%, 5%, 10% and 20% are presented in Fig. 2. The efficiency of pure ZnO and ZrO₂ nanoparticles is also shown. Each material was tested in darkness and in UV light in triplicate. The results of the photocatalytic and sorption studies were used to compare the properties of each deposit.

ZrO₂–ZnO with 10% ZrO₂ showed the highest photocatalytic activity compared to materials with different concentrations of ZrO₂. A further increase in the concentration of ZrO₂ increased the sorption capacity of the material but reduced photocatalytic properties, which was undesirable (Fig. 2).

On the basis of the conducted photodegradation studies of MB on pure ZrO₂, an increase in absorbance at a wavelength of 664 nm was observed during the reaction time of irradiation. Higher absorbance of the solution indicates an increase in the concentration of the dye in the solution. The study of the kinetics of the sorption process confirmed the high impact of physical sorption on the process (Table 2).

The second-order model is a model with the highest determination coefficient, which suggests the chemical nature of sorption. Based on the analysis of the process carried out under UV, the desorption process took place. The second-order model, despite a much lower value of the reaction rate constant is still positive, which would indicate a slow sorption process. In the case of the first-order model, the determination coefficient is lower, especially in the case of the UV process. However, the reaction rate constant is negative, which confirms the desorption process. Both models show some disadvantages in the actual description of the sorption process on ZrO₂, hence the complex mechanism of the process, combining the dye with the material, both by chemical and physical interactions. The process of desorption under the influence of UV is not yet known, therefore, further research should be carried out.

The constant of the reaction rate in the pseudo-first-order model in the process conducted with UV radiation has a negative value, which confirmed the desorption process. In all kinetic models, the degree of regression fit of experimental data to model equations is high, which suggests a complex process of physical sorption combined with chemical sorption. According to the results, after the first minute of irradiation with UV radiation, an increase in the concentration of MB in the solution was observed (Fig. 3).

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**Fig. 2** Photocatalytic efficiency of ZrO₂–ZnO with percentage concentrations of ZrO₂ from 0 to 20% wt.

**Table 2** Kinetic parameters of MB adsorption onto ZrO₂ in darkness and in UV light

| Kinetic model               | Parameters | Sorption in the darkness | Sorption under UV |
|-----------------------------|------------|---------------------------|-------------------|
| First order kinetic model   | k₁ (min⁻¹) | 0.1181 ± 0.05134          | −0.02487 ± 0.003306 |
|                            | Rₘₐₓ (mg/g)| 5.9265 ± 1.1720           | 0.9112 ± 0.8889    |
|                            | R²         | 0.9519                    | 0.7281             |
| Second order kinetic model  | k₄ (g/mg/min)| 0.06232 ± 0.004051        | 1.386 · 10⁻³ ± 2.985 · 10⁻⁴ |
|                            | Rₘₐₓ (mg/g)| 28.9855 ± 0.2198          | 25.8398 ± 0.3849   |
|                            | R²         | 0.9998                    | 0.9991             |
| Elovich                     | β (g/mg)   | 0.5304 ± 0.03424          | 0.6910 ± 0.1976    |
|                            | R²         | 0.9836                    | 0.7535             |
| Experimental data           | Rₑₓᵖ (mg/g)| 28.5712                  | 26.9068            |
The surface energy of adsorption decreases with increasing temperature. In the case of UV radiation, the energy added to the system may increase the energy of the MB particles and thus reduce the surface energy between the dye particles and the ZrO₂ surface. With a significant contribution of physical sorption, the energy of the bond between the bed and the dye is lower compared to chemical sorption. Therefore, after providing additional energy in the form of UV radiation, MB particles can be desorbed (Fig. 4).

Additionally, for the obtained materials, the width of the energy gap was determined (Table 3). A higher mass ratio of ZrO₂ in ZrO₂–ZnO materials caused a decreased width in the energy gap. The width of the energy gap for each material was obtained using the UV–Vis method. Absorption, which corresponds to the energy difference between the top of the valence band and the bottom of the conduction band, was used to designate the band gaps (Fig. 5). Based on the studies, material with 10%wt of ZrO₂ was chosen as the material for further analysis.

**Photocatalytic degradation of dye solutions**

**Effects of parameters processes**

A CCD with three factors and five coded levels was adopted in the optimization study. Table 4 presents the
experimental values of the photocatalytic efficiency of MB degradation and the amounts of the degraded dye under designed conditions. The response variable was fitted using a regression model:

\[ y = a_0 + \sum_{i=1}^{k} a_i \cdot x_i + \sum_{i=1}^{k} a_{ii} \cdot x_i^2 + \sum_{i<j}^{k} a_{ij} \cdot x_i \cdot x_j, \]  

(4)

where \( y \) — efficiency of photodegradation (%), \( a_0 \) — constant term, \( a_i \) — linear coefficients, \( a_{ii} \) — quadratic coefficients, \( a_{ij} \) — coefficient of the interaction parameters, \( x_i, x_j \) — values of the independent variables.

Based on the ANOVA results, parameters that significantly affected the photocatalytic efficiency of \( \text{ZrO}_2-\text{ZnO} \) were determined (Table 5). All selected input variables as well as their interactions had a significant effect on the photocatalytic properties of the material. The \( p \) value of 0.05 was determined as the significance level in the analysis.

The increase in the gradient between the amount of \( \text{ZrO}_2-\text{ZnO} \) and the dye in the system contributed to the increase in the removal rate of MB from the solution, which is associated with an increase in driving force. The initial concentration of the dye as a factor with a negative coefficient estimate caused a decrease in the efficiency of the \( \text{ZrO}_2-\text{ZnO} \) together with an increase in the concentration of dye in the system (Fig. 5). Lee et al. studied the effect of various factors on \( \text{ZnO} \) efficiency, also confirming the negative effect of the initial concentration of dye and a positive coefficient estimate of the pH of the solution on the degradation of phenoxyacetic acid (Lee et al. 2015).

**Effect of initial dye concentration on photocatalytic performance** The photocatalytic degradation of MB was carried out at different initial concentrations of dye in the range of 20–60 mg/dm³. The results are shown in Fig. 7. An increase in the concentration of dye decreases the path length of photons entering the pollutant solution. At high pollutant concentrations, a significant amount of light may be absorbed...
by the pollutant molecules rather than the catalyst, and this reduces catalytic efficiency. For the degree of dye removal, constant bed weight and higher MB concentration increases the concentration gradient of active reagents. A greater driving force improves the degree of dye removal from the solution. Similar results were obtained by investigating the effectiveness of ultrasonicated ZnO–ZrO2 on another dye, congo red (Aghabeygi and Khademi-Shamami 2018), and by testing MB on ZnO modified SnO2 (Lin et al. 2018).

The effect of photocatalyst content on the photocatalytic performance The amount of catalyst was one of the most effective parameters in the degradation studies. This was due to the increase in the number of ZnO particles, which increased the adsorption of photons and adsorption of pollutant molecules. On the other hand, the concentration gradient between the bed and the dye was the highest for 10 mg of the bed. The degree of dye removal, unlike efficiency, is referred to a unit mass of the bed used, hence an increase in the amount of ZrO2–ZnO causes a decrease in the use of the bed surface. It has been observed that with the increase of ZrO2–ZnO (10–90 mg), the efficiency of the deposit increases, which is in contrast to the degree of dye removal. With a higher amount of photocatalyst mass, the number of free electron–hole pairs increases, there-

Table 5 Coefficient of regression model and its significance on deposit efficiency (%) and dye removal rate (mg/g)

| Independent variables | Symbol | \(E (%)\) | Coefficient estimate | Std. deviation | \(F\) value | \(p\) value | \(R_D\) (mg/g) | Coefficient estimate | Std. deviation | \(F\) value | \(p\) value |
|-----------------------|--------|-----------|---------------------|---------------|------------|------------|-------------|---------------------|---------------|------------|------------|
| Mean/constant         |        | 89.41     | 0.3957              | 62.40         | 0.0028     | 21.97      | 0.0665      | 330.4               | 0.0019        |
| \((1) C_0 \text{[mg/dm}^3\text{]} (L)\) \(x_1\) |        | -15.61    | 0.5109              | -30.55        | 0.0208     | 13.30      | 0.0859      | 154.9               | 0.0041        |
| \(C_0 \text{[mg/dm}^3\text{]} (Q)\) \(x_1^2\) |        | -13.63    | 1.043               | -13.07        | 0.0486     | -4.705     | 0.1753      | -26.84              | 0.0237        |
| \((2) m \text{[mg]} (L)\) \(x_2\) |        | 24.21     | 0.5109              | 47.39         | 0.0134     | -27.79     | 0.0859      | -323.7              | 0.0020        |
| \(m \text{[mg]} (Q)\) \(x_2^2\) |        | -28.21    | 1.043               | -27.05        | 0.0235     | 15.04      | 0.1753      | 85.78               | 0.0074        |
| \((3) \text{pH [–]} (L)\) \(x_3\) |        | 31.85     | 0.5107              | 62.36         | 0.0102     | 8.467      | 0.0858      | 98.64               | 0.0065        |
| \( \text{pH [–]} (Q)\) \(x_3^2\) |        | -27.00    | 1.043               | -25.89        | 0.0246     | -7.689     | 0.1752      | -43.86              | 0.0145        |
| \(1L \text{ by 2L} \) \(x_1 \cdot x_2\) |        | 13.02     | 1.123               | 11.60         | 0.0547     | -2.264     | 0.1887      | -12.00              | 0.0529        |
| \(1L \text{ by 3L} \) \(x_1 \cdot x_3\) |        | -11.09    | 1.122               | -9.881        | 0.0642     | 0.6271     | 0.1886      | 3.325               | 0.1860        |
| \(2L \text{ by 3L} \) \(x_2 \cdot x_3\) |        | 15.22     | 1.122               | 13.57         | 0.0468     | -3.138     | 0.1886      | -16.64              | 0.0382        |

Italic values indicate significance of \(p\) value \((p < 0.05)\)

\(L\)—linear main effects of input values
\(Q\)—square main effects of input values
\(1L \text{ by 2L} \) linear interactions effects between \(C_0\) and \(m\)
\(1L \text{ by 3L} \) linear interactions effects between \(C_0\) and \(\text{pH}\)
\(2L \text{ by 3L} \) linear interactions effects between \(m\) and \(\text{pH}\)

Fig. 6 Optimization plot for the removal of MB by ZrO2–ZnO (10% ZrO2)
fore, more particles of the dye are removed. However, not all active spaces of the ZrO$_2$–ZnO are used (Fig. 8). In the case of deposit efficiency, the concentration gradient increases, which results in a higher degree of dye removal without this being converted into a unitary amount of dye in the solution.

**Effect of the initial pH of the dye solution on photocatalytic performance** At the acidic pH of the solution, the lowest degree of oxidation and the lowest degree of dye removal from the solution were observed (Fig. 9). Spathis and Poulious showed that in aqueous solutions, under the influence of UV irradiation, ZnO can undergo partial dissolution, which is known as photocorrosion (Spathis and Poulios 1995). In an acidic environment, it can proceed as follows:

$$2\text{ZnO} + 2\text{H}^+ + 2\text{h}_\text{VB}^+ \rightarrow \text{H}_2\text{O}_2 + \text{Zn}^{2+}. \quad (5)$$

This results in a double inhibition of the photocatalytic activity of the material. First, it reduces the amount of zinc oxide that would generate electron–hole pairs. Additionally, a part of $\text{h}_\text{VB}^+$, instead of participating in the generation of $\text{H}$, can participate in the competitive ZnO oxidation reaction and thus reduce the overall efficiency of the photocatalytic process. In solutions with a higher pH near the $\text{Zn}^{2+}$ ions, mainly $\text{H}_2\text{O}$ and $\text{OH}^-$ ions ($\text{H}_2\text{O}_{\text{ads}}, \text{OH}_{\text{ads}}$) are absorbed. They can be oxidised by $\text{h}_\text{VB}^+$, which is generated under the influence of UV radiation (Aghabeygi and Khademi-Shamami 2018). The $\text{OH}$ radicals can participate in further free radical reactions with dye particles adsorbed on the ZnO surface. This leads to the degradation of organic compounds:

$$\text{dye} + \text{OH}^- \rightarrow \text{decomposition products}. \quad (6)$$

**Kinetic studies, effect of time on photocatalytic performance**

In the general approach, the mechanism of photocatalytic degradation follows the Langmuir–Hinshelwood (L–H) kinetic model (Lee et al. 2015). The L–H model was used to investigate the effect of time on the heterogeneous photodegradation process and the equation of reaction rate ($r$) is represented by

$$r = \frac{K_1K_2C}{1 + K_2C}, \quad (7)$$

where $K_1$ is the rate constant, which includes various parameters such as nanophotocatalyst content, and $K_2$ is the adsorption constant. After assuming that the dye concentration is low, $K_2C \ll 1$, hence the reaction rate is a
first-order equation (Aghabeygi and Khademi-Shamami 2018). Taking into account the boundary conditions \[ C = C_0 \text{ at } t (\text{time}) = 0 \], it is possible to determine the reaction rate constant:

\[
-ln \frac{C}{C_0} = kt,
\]

where \( k \) is the photodegradation rate constant (min\(^{-1}\)).

A series of tests were carried out where the concentration of the dye in the solution was examined at a concentration of MB 40 mg/dm\(^3\) with 500 mg of photocatalyst in 300 ml of mixture. The results are shown in Fig. 10. For the first 30 min, to establish the equilibrium, the process was carried out in darkness. After selected intervals, a sample was taken to examine the concentration of the dye. The convex curve indicates a decreasing amount of dye in the mixture and, consequently, the dye particles are degraded with a lower rate. It can be seen from Fig. 10b that the kinetic experimental data are in good agreement with the kinetic model \( R^2 = 0.9925 \) and the first-order rate constant equals 0.03324 min\(^{-1}\). Belcha et al. examined the effect of catalyst concentration on the MB degradation rate over time. They determined rate constants of 0.0124 min\(^{-1}\) for ZnO NPs obtained using the sol–gel method and 0.0084 min\(^{-1}\) for ZnO NPs obtained using the precipitation method (Balcha et al. 2016). Payra et al. investigated the photocatalytic degradation mechanism of MB over ZIF-derived ZnO, obtaining first-order kinetics and a rate constant of 0.0452 min\(^{-1}\), which confirmed the efficiency of the photocatalytic material (Payra et al. 2019).

**Characterization**

**TEM–EDX investigations** TEM analysis confirms that ZrO\(_2\)–ZnO NPs preserve their original structure after sorption and photocatalytic processes (Fig. 11a–f). Different types of particles were obtained. Long needle-shaped microparticles as well as their nanoscale counterparts were observed (Fig. 11a, c, e). At the same time, irregularly shaped nanoparticles consisting of crystallites of several nanometers appear on the surface of other particles (Fig. 11b, d, f). An EDX analysis was also carried out to confirm the elemental composition of the obtained composite (Fig. 12). The EDX analysis confirmed the presence of all three elements: Zr, Zn and O. It can be seen that in the first area of scanning, there are sharp peaks from zinc and oxygen indicating the presence of ZnO, however, there are no visible reflexes from zirconium. In contrast to the first area, in the second scanning region, clear peaks are visible from both zinc and zirconium as well as oxygen, which confirms the presence of ZrO\(_2\) and ZnO. Depending on the place of sample analysis, the material composition differs from 0.54 to 24.41% wt ZrO\(_2\). Compared to the results, the percentage composition of the XRD analysis was 10–11% wt ZrO\(_2\). A discrepancy in the results indicates the inhomogeneity of the material in the nanoscale, nonetheless, it was expected that both the use of powdered ZrO\(_2\) and the use of only 10% of ZrO\(_2\) on the whole mass would prevent a homogeneous material for the entire volume from being obtained.

**XRD investigations** XRD analysis confirmed that zinc oxide in crystalline form and amorphous zirconium oxide had been obtained. Materials after sorption and photodegradation processes did not change their form. Figure 13 shows the XRD patterns of the synthesized pure ZnO, ZrO\(_2\) and

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**Fig. 10** Absorption of MB in different exposure times under UV radiation (a); kinetics of photocatalyst degradation of MB under UV radiation (b)
Fig. 11 TEM images of the ZrO$_2$–ZnO nanoparticles: pure ZrO$_2$–ZnO (a, b), after the sorption process (c, d), after the photocatalytic process (e, f).

Fig. 12 TEM-EDX images of the ZrO$_2$–ZnO nanoparticles
ZrO₂–ZnO (10%) before and after the sorption process and the photodegradation of the MB. The XRD pattern of the prepared crystalline ZnO is in accordance with the hexagonal zincite crystalline structure of ZnO, with sharp and strong peaks at 2θs of 31.75°, 34.40°, 36.24°, 47.54° and 56.61°, which correspond to (100), (002), (101), (012) and (110), respectively (Garino et al. 2019; Hasanpoor et al. 2015). The XRD pattern for pure ZrO₂ confirmed characteristic planes at 2θs of 30.23°, 34.92° and 50.26° (Singh and Nakate 2014; Manjunatha and Dharmaprakash 2016). The peak centred at 30.23° (101) is characteristic of the tetragonal crystalline phase (Uribe López et al. 2019). In the ZrO₂–ZnO, the corresponding peaks of ZrO₂ are absent, whereas peaks corresponding to the ZnO remain sharp and strong. In addition, the diffractograms representing ZnO–ZrO₂ show less homogeneous patterns compared to pure ZnO, which indicates an increased amorphous structure. The patterns suggest a greater polydispersity of the material surface, thanks to which the material has better photocatalytic properties.

**FTIR results** The FT-IR spectra of pure ZnO, pure ZrO₂ and ZnO–ZrO₂ nanoparticles are presented in Fig. 14. All materials show spectra at a wavenumber of about 500 cm⁻¹, which indicates the metal–oxygen bonding of Zn–O and Zr–O (Janaki et al. 2015). The peaks at around 3400 cm⁻¹ and 1500 cm⁻¹ in all samples are the stretching and bending vibrational bands of the hydroxyl groups adsorbed on the surface. These effects from hydroxyl groups are strongly visible due to the lack of calcination of the obtained products. At the same time, it should be noted that these interactions are more pronounced for samples with ZrO₂, which is related to the amorphous nature of these products and the formation of bonds between O–H and Zr–O (Zinatloo-Ajabshir and Salavati-Niasari 2016). The sharp signals in the region of 1360–1420 cm⁻¹ appear from carbonate anions, which is a direct result of the use of sodium carbonate as a precipitating agent (Khan et al. 2016). There are no visible peaks derived from adsorbed MB at 1390, 1487 or 1590 cm⁻¹ (Bartošová et al. 2017) in the ZnO–ZrO₂ after the sorption process due to the fact that they coincide with the peaks from pure bed in the region of 1360–1600 cm⁻¹, which in combination with a relatively small amount of adsorbed MB, causes them to be totally covered.

**Mechanism of photocatalyst**

The mechanism of dye photodegradation using a ZnO semiconductor initiated by UV can be presented in the following equation:

\[
\text{ZnO} + \text{hv} \rightarrow \text{ZnO}(h_v^+ + e_o^-),
\]

where \(h_v^+\) are holes of the generated valence band of the semiconductor, which have strong oxidising properties, and \(e_o^-\) are electrons generated in the conduction band of the semiconductor, which have high reducing properties (Uribe López et al. 2019).

The addition of semiconductors is one way of improving the photocatalytic properties of ZnO. Different widths of energy bands occurring in both oxides contribute to reducing the distance between the conduction bands and the valence bands in the whole composite. In the ZrO₂–ZnO system, some of the electrons pass from the conduction band of ZrO₂ to the conduction band of ZnO. At the same time, from the valence band of ZnO, some of the holes migrate to the valence band of ZrO₂, leading to the generation of electron–hole pairs (Fig. 15). Electron transfer, and consequently
the transfer of holes, contributes to the reduction of the rate of recombination of the electron–hole pairs, a phenomenon that reduces photocatalytic efficiency (Aghabeygi and Kha- demi-Shamami 2018; Lin et al. 2018).

The photocatalytic decomposition of dye, in this case MB, takes place in stages. Radicals and ions obtained in the photocatalytic process cause the dye to degrade in a multistep process. In the initial stages, the least stable bonds decay, resulting in single ring compounds, which in the final stage become basic inorganic compounds like \( \text{H}_2\text{O} \), \( \text{CO}_2 \), \( \text{NO}_3^- \), and \( \text{SO}_4^{2-} \) (Adeleke et al. 2018).

**Effect of type of dye on photocatalytic performance**

The efficiency of photocatalytic processes depends not only on photocatalyst properties but also on effluent composition, including the type of removed dye. The effect of the composition of the dye mixture on the photocatalytic properties of the deposit was investigated by studies of the photodegradation of a mixture of Methyl Orange (MO) and MB (MB) dye. The UV–Vis spectra of the solutions of the dyes after 30 min of photodegradation are shown in Fig. 16. Based on the UV–Vis spectra, the characteristic wavelengths for each dye were chosen: 464 nm for MO and 664 nm for MB. Analysis of the results showed a much higher amount of degraded MB compared to MO. The amount of degraded MB, which is a cationic dye, remained at 90%, while the amount of degraded MO did not exceed 70% (Table 6).

The higher efficiency of MB removal may result from the presence of an increased amount of \( \cdot \text{OH} \) and \( \text{OH}^- \) groups on the surface of \( \text{ZrO}_2–\text{ZnO} \), which interact with positively charged MB ions and cause their decomposition rather than the anionic dye. Similar conclusions have been reached by Trandafilović et al. among others, who investigated the effect of the dye type on the efficiency of a \( \text{ZnO}–\text{Eu} \) deposit. They confirmed higher MB removal compared to MO. They also confirmed the significant effect of the presence of \( \text{OH} \) groups during the removal of MB, while for MO, the most important factor was the formation of holes (Trandafilović et al. 2017). Beura and Thangadurai studied the degradation efficiency of the Sn–ZnO–graphene for mixed dyes (MO and MB). 17% of MO was degraded under UV, whereas MB was
The amount of the ZrO2 significantly increased the photocatalytic properties of microwave irradiation. Results showed that the addition of anionic dye MO. of cationic dye MB was greater than the photodegradation be over 97% after just 30 min. The photocatalytic activity of ZnO crystallites were successfully synthesised with different proportions of ZrO2 in a precipitation method with the help of microwave irradiation. Results showed that the addition of ZrO2 significantly increased the photocatalytic properties of ZnO. Different widths of energy bands occurring in both oxides contributed to reducing the distance between the conduction bands and the valence bands in the whole composite. An increase in the mass of the bed, an increase in the pH of the solution, and a decrease in the initial concentration of MB caused the photocatalytic efficiency of ZrO2–ZnO to be over 97% after just 30 min. The photocatalytic activity of cationic dye MB was greater than the photodegradation of anionic dye MO.

Conclusions

ZnO crystallites were successfully synthesised with different amounts of ZrO2 via a precipitation method with the help of microwave irradiation. Results showed that the addition of ZrO2 significantly increased the photocatalytic properties of ZnO. Different widths of energy bands occurring in both oxides contributed to reducing the distance between the conduction bands and the valence bands in the whole composite. An increase in the mass of the bed, an increase in the pH of the solution, and a decrease in the initial concentration of MB caused the photocatalytic efficiency of ZrO2–ZnO to be over 97% after just 30 min. The photocatalytic activity of cationic dye MB was greater than the photodegradation of anionic dye MO.

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Compliance with ethical standards

Conflict of interest The authors report no declarations of interest.

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