Redox-ligand sustains controlled generation of CF₃ radicals by well-defined copper complex†

Jérémy Jacquet, Sébastien Blanchard, Etienne Derat, Marine Desage-El Murr* and Louis Fensterbank*

A well-defined copper complex bearing iminosemiquinone ligands performs single electron reduction of an electrophilic CF₃⁺ source into CF₃ radicals. This redox behavior is enabled by the ligand which shuttles through two different redox states (iminosemiquinone and iminobenzoquinone) while the copper center is preserved as a Cu(II). This system was used in the trifluoromethylation of silyl enol ethers, heteroaromatics and in the hydrotrifluoromethylation of alkynes. This is the first example of cooperative redox catalysis for the controlled generation of CF₃ radicals.

Introduction

Cooperative redox catalysis steps away from the established organometallic paradigm involving redox events occurring at the metal center by introducing ligand-based electronic participation. Drawing inspiration from the redox relays involved in enzymatic processes, this area of catalysis hinges upon a participative role of the non-innocent ligand in catalytic elementary steps in order to achieve multielectronic transformations with base metals.2–8

While redox non-innocence in biological settings has long been familiar to bioinorganic chemists, the use in a broader context of “synthetic” redox ligands as surrogates of the original biological radicals is emerging as a catalytic tool of its own. These privileged molecular scaffolds have been shown to actively participate in catalytic events through reversible delocalization of spin density.4 This redox interplay between ligand and metal can provide attractive alternative mechanistic venues for catalyst development. Among expected benefits, the enabling potential of these ligands towards first-row transition metals or redox silent metals could enlarge their chemistry and unravel unprecedented reactivities. This area of catalysis is currently on the fast track for the development of alternative catalytic approaches circumventing the use of precious noble metals.3

Seeking to enlarge the scope of applications, we have been focusing on a well-known family of metal complexes exhibiting non-innocent behavior: copper complexes with iminosemiquinone ligands. Originally developed as enzyme mimics,9 these complexes are of particular interest due to the versatile ligand scaffold which can accommodate two successive monoelectronic oxidation steps through a redox chemical interplay between two oxidation states: amidophenolate (AP), iminosemiquinone (SQ) and iminobenzoquinone (BQ). These systems are now being developed in broader (catalytic) contexts9 with several metals including iridium,10 cobalt,11,12,13 palladium14,15 and copper.16,17 Previous work allowed us to establish that the interaction of complex 1 [Cu(II)(LBQ)2] with an electrophilic source of CF₃ triggers CF₃ uptake by the complex through ligand-based bi-electronic redox participation while preserving the metal Cu(II) oxidation state.18 The resulting [Cu(n)(L₈Q)₂CF₃]⁺ complex 2 exhibits a nucleophilic intramolecular CF₃ reactivity, thus suggesting that redox involvement of ligands can sustain formal umpolung of the CF₃ moiety (Fig. 1).

Transfer of trifluoromethyl groups through metal-catalyzed processes is currently a topic of acute interest as this group belongs to the privileged moieties in synthetic chemistry. Its introduction in a biologically active scaffold enhances metabolic stability and favors permeation of drugs through the blood

Fig. 1 Ligand-based redox reactivity of complex 1 with an electrophilic CF₃⁺ source.
The widespread benefits of the introduction of trifluoromethyl groups also pervade through materials chemistry and agrochemistry. Accordingly, numerous strategies have been devised by chemists for the introduction of this chemical functionality. 

The wide range of available nucleophilic, electrophilic and radical trifluoromethylation sources provides varied catalytic manifolds opportunities, which plays no mean part in the success of this flourishing area. The wealth of CF₃ sources available has allowed flexibility in the development of a variety of approaches differing from the electronics involved. In this matter, several metals have proven valuable to the chemists among which copper and palladium. Strategies involving the addition of CF₃ moieties across multiple bonds have also been developed and these efficient approaches allow building of molecular complexity.

Recently, among these strategies, major focus has been placed on synthetic application of systems implying CF₃ radicals and recent elegant contributions rely on photo-redox catalytic manifolds as means to generate such radicals in a clean and controlled fashion. Radical trifluoromethylation of unsaturated moieties is inherently challenging because of competing radical-atom transfer and several methods have sought to take advantage of this reactivity in order to develop hetero- and carbotrifluoromethylation. However, direct hydrotrifluoromethylation of multiple bonds, thus generating simple trifluoromethylated alkenes and alkanes, has been less explored and is an emerging application. While these methods are an elegant and efficient way to produce CF₃ radicals in a controlled fashion, they mostly imply expensive and less sustainable noble metals such as ruthenium and iridium and/or also often imply the use of additives playing the role of sacrificial electron donors and redox relays.

On the cheaper end of the metal spectrum, copper is a metal that has enjoyed recent exciting applications in trifluoromethylation of various organic substrates. Notably, an efficient Cu(1)-photocatalyzed trifluoromethylation of alkenes was recently reported. Another generation of CF₃ radicals by Cu(i) salts from an electrophilic CF₃ source (5) in the trifluoromethylation of unsaturated olefins to generate allylic trifluoromethlated products has been reported by Wang et al. using CuCl and by Buchwald et al. using [(MeCN)₄Cu]PF₆. A similar transformation was reported by Fu and Liu using electrophilic source 4 and CuTC ((thiophene-2-carbonyloxy) copper) and the authors postulated the involvement of a Cu(n)–CF₃ species. The use of simple copper salts in these reactions is convenient but the lack of ligands that could stabilize the active copper species generates less controlled reactive intermediates and is a drawback for mechanistic elucidation. Moreover, these protocols often require between 10 and 20 mol% catalyst loading in order to provide good results.

Here, we show that well-defined complex 1 can be used in conjunction with an electrophilic CF₃ source (Umemoto 4 or Togni II 5 reagents) in the presence of TEMPO (2,2,6,6-tetramethylpiperidine-N-oxyl radical, Scheme 1) led to the formation of the TEMPO–CF₃ adduct, as evidenced by NMR ¹⁹F spectroscopy (δ = −55.7 ppm). This reactivity was evaluated in CH₂Cl₂ and NMP with two CF₃ sources (4 and 5) and was found to proceed in fairly good yields (up to 69%). These results suggest that complex 1 can promote the generation of CF₃ radical from a CF₃ source in the reaction medium. An independent equimolar mixture of TEMPO and complex 1 was followed by UV-vis and no reaction was found to occur, thus ruling out the possibility of direct electronic transfer between complex 1 and TEMPO and subsequent ionic CF₃ transfer. Also, no formation of the TEMPO–CF₃ adduct occurs in the absence of complex 1, which is thus mandatory. This reactivity is in sharp contrast with that of the fully oxidized complex 2 [Cu(n)[L₃Sn(CF₃)]] which was found to be inert to reaction with TEMPO as no transfer of the CF₃ moiety was observed.

During the trifluoromethylation of TEMPO, the reaction medium turns from dark green to dark purple (source 4) or dark orange (source 5). We thus decided to investigate the locus of the electronic transfer by UV-vis spectroscopy. The initial green colour of the solution mainly corresponds to that of [Cu(L₃Sn)] complex 1, as attested by the presence on the spectra of its two characteristic bands around 300 nm and 800 nm and as can be seen in Fig. 2 (green curve). At the end of the reaction, these two characteristic bands have disappeared, while broad new features appear with maxima at 425 nm, 525 nm and 725 nm (Fig. 2, brick red curve). Interestingly, very similar bands have been reported for the electrochemically synthesized [Cu(L₃Sn)][(I₈O)] complex 3, arising from single electron oxidation of complex 1. We therefore independently synthesized complex 3 via amidation of an equimolar mixture of [Cu(L₃Sn)] complex 1 and its related dicationic complex 6 [Cu(L₃Sn)]₂⁺ (see ESI† for preparation) and recorded its absorption spectrum (Fig. 2, sky blue curve). Indeed, an

![Scheme 1 TEMPO trapping experiments for the generation of CF₃ radicals with complexes 1 and 2. NMP: N-methylpyrrolidinone.](image-url)
excellent match between this reference spectrum and that of the reaction medium at the end of the reaction is observed. An almost perfect match is even obtained upon addition of one equivalent of dibenzothiophene, which is released in the reaction medium upon trifluoromethylation with Umemoto’s reagent 4 (Fig. 2, purple curve).

These results strongly suggest that [Cu(LSQ)(LBQ)]⁺ complex 3, the monooxidized product of [Cu(LSQ)₂] complex 1, is formed during the reaction and point towards a SET process involved in the reaction and sustained by the redox ligand.

Since complex 3 seems to be a cornerstone for the reactivity examined here, its electronic structure was evaluated through DFT calculations. These calculations were performed with Turbomole v6.4, using the B3LYP functional complemented by the D3 dispersion scheme and with the Def2-SV(P) basis set. In Fig. 3 can be seen the optimized structure together with the total spin density. It appears that one unpaired electron remains on the metallic center (in the dx²−y² orbital) while the second one is dispatched over the two ligands, with an antiferromagnetic coupling between the two. Thus, while we formally write that complex 3 is bearing two different ligands (one SQ and one BQ), the picture emerging from DFT calculations shows that the unpaired electron is fully delocalized on both ligands. The corresponding UV-vis spectrum of complex 3 was calculated using the same DFT level (Fig. 4). In the 400–800 nm range, the experimental and calculated spectra appear to be very similar. Two groups of transitions, centered around 425 nm and 525 nm bands, while an additional band around 790 nm may be associated with the 725 nm shoulder. Overall, the calculated spectrum fits well with the experimental data. Thus, all the data collected allow us to confirm the nature of complex 3 as [Cu(LSQ)(LBQ)]⁺.

When using Togni’s reagent 5, the spectrum at the end of the reaction displays maxima around 300 nm, 390 nm and 500 nm (Fig. 5, brick red curve), which differs notably from the spectrum of [Cu(LSQ)(LBQ)]⁺ complex 3. We wondered if the iodo-benzoate released from Togni’s reagent might interact with the final copper complex and to our delight, adding the corresponding carboxylate to independently generated [Cu(LSQ)(LBQ)]⁺ complex 3 induces a change in its UV-vis spectrum, which then displays a very good match with the spectrum observed at the end of the reaction (Fig. 5, purple curve).

Thus, in the case of Togni’s reagent 5, a SET from one of the iminosemiquinone ligands of [Cu(LSQ)₂] complex 1 to form [Cu(LSQ)(LBQ)]⁺ complex 3 also appears very likely. These experiments strongly support a SET occurring between complex 1 and the CF₃⁺ source, in which the CF₃⁺ source is reduced to form a CF₃ radical while the concomitant oxidation of complex 1 into complex 3 is ligand-based and sustained by oxidation of an iminosemiquinone ligand into an iminobenzoquinone ligand.

These observations led us to probe the reactivity of this system and evaluate the possibility of CF₃ uptake in various
families of radical acceptors. Photoredox\textsuperscript{a,4,5} and copper-catalyzed\textsuperscript{6} catalytic trifluoromethylation of silyl enol ethers to generate α-trifluoromethyl ketones have been reported and this family of substrates was used as benchmark reactivity for our system.

We were pleased to see that reaction of silyl enol ethers 7a–d with reagent 5 in the presence of 5 mol% of complex 1 yielded the corresponding trifluoromethylation adducts 8a–d in promising to good yields from 47 to 94% (Table 1). However, substituting source 4 for 5 provided lower yields. The reaction proceeds within 18 hours at room temperature, which compares well with reported conditions (24 h at rt\textsuperscript{15} and 12 h at rt to 50 °C\textsuperscript{6}).

Introduction of CF\textsubscript{3} motifs in heteroaromatics is also a topic of interest and several radical-based methods have been reported.\textsuperscript{6,9,10} Efficiency of the redox catalyst 1 was probed with indole, pyrrole and furane derivatives (Table 2, 9a–f). The yields range from 59 to 87% and the regioselectivity obtained for 10b is consistent with reported radical trifluoromethylation conditions.\textsuperscript{40} Lowering the catalyst loading to 2 mol% resulted in decreased yields. Furthermore, the structure of compound 10c was confirmed by X-ray crystallography. Successful preparation of 10c and 10e also demonstrates the mildness of our reaction conditions towards electrophiles such as aryl iodides, which are typically reactive under photoredox conditions\textsuperscript{37,38} and could in our approach be used as synthetic handles for further functionalization by classic methods.

Having established the catalytic activity of complex 1 with well-known radical acceptors, we turned our attention towards hydrotrifluoromethylation of alkynes or alkenes. Since a seminal study from Kitazume in 1985 reporting the ultrasound-promoted hydrotrifluoromethylation of alkynes with trifluoroalkylcuprates, generated \textit{in situ} from CF\textsubscript{3}I or CF\textsubscript{3}Br, zinc powder and substoichiometric amounts of CuI,\textsuperscript{39} only a handful of strategies for selective hydrotrifluoromethylation have been reported.\textsuperscript{7,7a,7b} These involve inorganic electrode,\textsuperscript{7,2a} silver-catalyzed\textsuperscript{2a} or organic\textsuperscript{53,54} and organometallic\textsuperscript{55,56} photoredox systems.

Cho and co-workers have reported the controlled trifluoromethylation of alkynes with CF\textsubscript{3}I, \textit{fac-[Ir(ppy)\textsubscript{3}]} as catalyst and DBU (10 equiv.) as reductive quench and H source.\textsuperscript{76} The Gouverneur group has reported a method for alkynes and alkenes using photoredox catalyst [Ru(bpy)\textsubscript{3}]\textsuperscript{2+} in conjunction with Umemoto reagent 4 as CF\textsubscript{3}I source and methanol as H source.\textsuperscript{55} An organic photoredox system was reported for alkynes by Scaino and co-workers using reagent 5 along with organic photocatalyst Methylene Blue and DBU (2 equiv.).\textsuperscript{54} We selected alkyne 11a as a model for optimization studies and established that the best conditions involved the use of complex 1 and reagent 5 in NMP (Table 3, entry 2). Introduction of additive such as CHD (1,4-cyclohexadiene) as H atom donor did not improve the yields (entries 3 and 4). The scope of the reaction was evaluated and provided the corresponding trifluoromethylated alkenes in yields between 39 and 61%. These results prove competitive towards the literature with 68%.

**Table 1** Trifluoromethylation of silyl enol ethers

| Yields | Reaction conditions: 1 equiv. silyl enol ether, 1.5 equiv. reagent 5. Reaction was performed with 4 instead of 5. Isolated yields in brackets. |
|--------|--------------------------------------------------------------------------------------------------|
| 8a 79% (71%) | | 8b 47% | | 8c 94% (88%) | | 8d 56% (53%) |

Yields were determined by \textsuperscript{19}F NMR analysis, reaction conditions: 1 equiv. heteroaromatic, 1.5 equiv. reagent 5. Reaction conditions: 5 mol% complex 1, 4 equiv. heteroaromatic, 1 equiv. reagent 5. 2 mol% catalyst loading. Isolated yields in brackets.

**Table 2** Trifluoromethylation of heteroaromatics

| Yields | Reaction conditions: 5 mol% complex 1, 4 equiv. heteroaromatic, 1 equiv. reagent 5. |
|--------|----------------------------------------------------------------------------------|
| 10a 72% (67%) | | 10b 59% | | 10c 57% (87%) |
| 10d 74% | | 10e 30% | | 10f 62% (60%) |

Yields were determined by \textsuperscript{19}F NMR analysis, reaction conditions: 1 equiv. heteroaromatic, 1.5 equiv. reagent 5. Reaction conditions: 5 mol% complex 1, 4 equiv. heteroaromatic, 1 equiv. reagent 5. 2 mol% catalyst loading. Isolated yields in brackets.
Table 3 Hydrotrifluoromethylation of alkynes

| Entry | CF₃⁺ source | Solvent | Additive | 12a<sup>a</sup> |
|-------|-------------|---------|----------|-----------------|
| 1     | 4           | NMP     | —        | 11%             |
| 2     | 5           | NMP     | —        | 61%             |
| 3     | 5           | NMP     | CHD (1 equiv.) | 35%    |
| 4     | 5           | CH₂Cl₂  | CHD (1 equiv.) | 23%    |

CHD: 1,4-cyclohexadiene. Reaction conditions: alkyne (4 equiv.), CF₃⁺ source (1 equiv.), solvent, rt, 18 h. <sup>a</sup> Yields were determined by <sup>19</sup>F NMR analysis. Isolated yields in brackets.

Table 4 Attempted hydrotrifluoromethylation of alkynes and trifluoromethylation of allenes

| Entry | CF₃⁺ source | 14a | 14b | 14c |
|-------|-------------|-----|-----|-----|
| 1     | 4           | Traces | Traces | Traces |
| 2     | 5           | 19%     | 29%      |       |

1,4-CHD = 1,4-cyclohexadiene. Reaction conditions: alkyne (4 equiv.), CF₃⁺ source (1 equiv.), 1,4-CHD (1 equiv.), NMP, rt, 18 h. Yields were determined by <sup>19</sup>F NMR analysis. <sup>a</sup> Reaction conditions: 5 mol% complex 1, 1.5 equiv. reagent 5, CH₂Cl₂, r.t., 18 h.

Scheme 2 Proposed mechanism for the controlled generation of CF₃ radicals sustained by redox ligand. [NMP–H] corresponds to the species resulting from H atom abstraction, more likely at the CH₂ position adjacent to the nitrogen.⁷⁶

The alkene (Scheme 2) is easier to oxidize than its vinylic counterpart – generated upon addition of CF₃ on the alkene – and thus undergoes oxidation to the alkyl cation followed by proton loss to generate the allylic CF₃ product more quickly than H transfer. This observation also extends to allene 15 which was found to undergo oxytrifluoromethylation to yield product 16 presumably arising from trapping of the cationic intermediate by 2-iodobenzoate generated in the reaction upon reduction of reagent 5.⁷¹

Based on literature precedents, mechanistic possibilities for this transformation include organometallic, ionic or radical pathways. An organometallic route was proposed by Beller and co-workers⁷² and by Sanford and co-workers⁷³ for the copper-catalyzed trifluoromethylation of vinyl boronic acids with CF₃ radicals generated by t-BuOOH and CF₃SO₂Na or CF₃I and photocatalyst Ru(bpy)₃Cl₂·6H₂O respectively. Both involved a Cu<sup>III</sup>–CF₃ intermediate releasing the final product through reductive elimination. Investigating the copper-catalyzed allylic trifluoromethylation of alkynes with reagent 4, Fu and Liu postulated a similar mechanism based on DFT calculations.⁷⁴ This mechanism can be ruled out in our system since the corresponding Cu–CF₃ complex 2 resulting from CF₃ uptake by complex 1 has been isolated and cannot transfer its CF₃ moiety.¹⁸ An ionic pathway seems unlikely as the presence of the catalyst is mandatory and this route would be in contradiction with the TEMPO trapping experiments (Scheme 1). Indeed, adding TEMPO to a solution of complex 1 does not change the UV-vis profile, therefore suggesting that no electronic transfer occurs between TEMPO and complex 1.⁷³

In light of these considerations, we propose a mechanism (Scheme 2) implying the generation of CF₃ radicals by SET from complex 1 to the CF₃⁺ source and subsequent addition of this radical onto the unsaturated moieties. The resulting alkyl–CF₃ radical 17 (Scheme 2, black pathway) could then undergo another SET thereby regenerating complex 1 along with the oxidized cation 18 and closing the catalytic cycle. Proton loss provides the expected product 19. In the case of alkynes (Scheme 2, grey pathway), vinyl–CF₃ radical 20 obtained presumably undergoes hydrogen transfer with NMP,⁷² thereby
providing trifluoromethylated alkene 21. The resulting NMP radical is then oxidized to the cation, thus regenerating complex 1. Also, performing the reaction of complex 3 instead of complex 1 with TEMPO in the presence of reagent 5 only provides TEMPO-CF₃ adducts in trace amounts (<3%) thus indicating that only reduced complex 1 is capable of reducing the CF₃⁺ source.⁷⁵

**Conclusions**

We have shown that a well-defined copper(II) complex can promote generation of CF₃ radicals by reduction of an electrophilic CF₃⁺ source. This unusual behavior is enabled by a redox dialogue between the CF₃⁺ source and the complex, mediated by the redox-active ligand which enables SET through ligand-centered oxidation. This strategy offers milder conditions and complementarity to photoredox catalytic manifolds as well-defined molecular complexes performing controlled generation of CF₃ radicals. This method circumvents changes in the redox state of the copper center by promoting electronic transfer at the ligand instead, thus limiting uncontrolled reactivities at the metal center, and was successfully applied to the trifluoromethylation of silyl enol ethers, heteroaromatics and hydrotrifluoromethylation of alkynes. To the best of our knowledge, this is the first example of hydrotrifluoromethylation of alkynes catalyzed by a well-defined copper complex and overall of controlled generation of CF₃ radicals by cooperative redox catalysis. This reactivity opens the way towards new developments in the field of redox ligand-based cooperative catalysis as possible mild alternatives to expensive noble metals such as iridium and ruthenium. Among other foreseeable broad catalytic perspectives, the participative nature of these ligands could allow them to stabilize high and/or disfavored metallic oxidation states, which is a prospect we are currently pursuing.

**Acknowledgements**

The authors thank MENRT (JJ), UPMC, CNRS, ANR for funding (grant ANR-11-JS07-004-01), IUF (LF) and LabEx MiChem. The authors also wish to thank Lise-Marie Chamoreau for X-ray analysis.

**Notes and references**

1 W. Kaim and B. Schwederski, *Coord. Chem. Rev.*, 2010, 254, 1580.
2 J. I. van der Vlugt, *Eur. J. Inorg. Chem.*, 2012, 363.
3 S. Blanchard, E. Derat, M. Desage-El Murr, L. Fensterbank, M. Malaire and V. Mouriès-Mansuy, *Eur. J. Inorg. Chem.*, 2012, 376.
4 V. Lyaskovskyy and B. de Bruin, *ACS Catal.*, 2012, 2, 270.
5 P. J. Chiirik and K. Wieghardt, *Science*, 2010, 327, 794.
6 O. R. Luca and R. H. Crabtree, *Chem. Soc. Rev.*, 2013, 42, 1440.
7 V. K. K. Praneeth, M. R. Ringenberg and T. R. Ward, *Angew. Chem., Int. Ed.*, 2012, 51, 10228.
8 H. Grützmacher, *Angew. Chem., Int. Ed.*, 2008, 47, 1814.
9 P. Chaudhuri, M. Hess, J. Müller, K. Hildenbrand, E. Bill, T. Weyhermüller and K. Wieghardt, *J. Am. Chem. Soc.*, 1999, 121, 9599.
10 D. L. J. Broere, R. Plessius and J. I. van der Vlugt, *Chem. Soc. Rev.*, 2015, 44, 6886.
11 M. R. Ringenberg, S. L. Kokatam, Z. M. Heiden and T. B. Rauchfuss, *J. Am. Chem. Soc.*, 2008, 130, 788.
12 A. L. Smith, K. I. Hardcastle and J. D. Soper, *J. Am. Chem. Soc.*, 2010, 132, 14358.
13 For a highlight, see: W. I. Dzik, J. I. van der Vlugt, J. N. H. Reek and B. de Bruin, *Angew. Chem., Int. Ed.*, 2011, 50, 3356.
14 D. L. J. Broere, B. de Bruin, J. N. H. Reek, M. Lutz, S. Dechert and J. I. van der Vlugt, *J. Am. Chem. Soc.*, 2014, 136, 11574.
15 D. L. J. Broere, L. L. Metz, B. de Bruin, J. N. H. Reek, M. A. Siegler and J. I. van der Vlugt, *Angew. Chem., Int. Ed.*, 2015, 54, 1516.
16 M. W. Bezpalko, B. M. Foxman and C. M. Thomas, *Inorg. Chem.*, 2013, 52, 12329.
17 Z. Alaji, E. Safaei, L. Chiang, R. M. Clarke, C. Mu and T. Storr, *Eur. J. Inorg. Chem.*, 2014, 6066.
18 J. Jacquet, E. Salanouve, M. Orio, H. Vezin, S. Blanchard, E. Derat, M. Desage-El Murr and L. Fensterbank, *Chem. Commun.*, 2014, 50, 10394.
19 K. Müller, C. Faeh and F. Diederich, *Science*, 2007, 317, 1881.
20 S. Purser, P. R. Moore, S. Swallow and V. Gouverneur, *Chem. Soc. Rev.*, 2008, 37, 320.
21 M. G. Campbell and T. Ritter, *Chem. Rev.*, 2015, 115, 612.
22 T. Yamazaki, T. Taguchi and I. Ojima, *Unique Properties of Fluorine and Their Relevance to Medicinal Chemistry and Chemical Biology*, Wiley-Blackwell, Chichester, 2009.
23 J. Wang, M. Sánchez-Roselló, J. L. Aceña, C. del Pozo, A. E. Sorochnisky, S. Fustero, V. A. Soloshonok and H. Liu, *Chem. Rev.*, 2014, 114, 2432.
24 C. Alonso, E. Martinez de Marigorta, G. Rubiales and F. Palacios, *Chem. Rev.*, 2015, 115, 1847.
25 T. Furuya, A. S. Kamlet and T. Ritter, *Nature*, 2011, 473, 470.
26 C. Hollingworth and V. Gouverneur, *Chem. Commun.*, 2012, 48, 2929.
27 O. A. Tomashenko and V. V. Grushin, *Chem. Rev.*, 2011, 111, 4475.
28 X.-F. Wu, H. Neumann and M. Beller, *Chem.–Asian J.*, 2012, 7, 1744.
29 J.-A. Ma and D. Cahard, *J. Fluorine Chem.*, 2007, 128, 975.
30 S.-M. Wang, J.-B. Han, C.-P. Zhang, H.-L. Qin and J.-C. Xiao, *Tetrahedron*, 2015, 71, 7949.
31 T. Liu and Q. Shen, *Eur. J. Org. Chem.*, 2012, 6679.
32 E. J. Cho, T. D. Senecal, T. Kinzel, Y. Zhang, D. A. Watson and S. L. Buchwald, *Science*, 2010, 328, 1679.
33 R. J. Lundgren and M. Stradiotto, *Angew. Chem., Int. Ed.*, 2010, 49, 9322.
34 Y. Ye and M. S. Sanford, *Synlett*, 2012, 23, 1696.
35 E. Merino and C. Nevado, *Chem. Soc. Rev.*, 2014, 43, 6598.
36 G. Han, Q. Wang, Y. Liu and Q. Wang, *Org. Lett.*, 2014, 16, 5914.
37 Y. Li, Y. Lu, G. Qiu and Q. Ding, *Org. Lett.*, 2014, 16, 4240.
Possible pathways for catalyst deactivation could arise from conversion of complex 1 into complex 2 (Fig. 1) upon CF\textsubscript{3} uptake, which would result in a catalytic dead-end as complex 2 does not perform CF\textsubscript{3} transfer (ref. 18). Deactivation of a photoredox catalyst by ligand functionalization with the generated radical has recently been reported (J. J. Devery III, J. J. Douglas, J. D. Nguyen, K. P. Cole, R. A. Flowers II and C. R. J. Stephenson, Chem. Sci., 2013, 6, 537). This type of deactivation pathway could also be at work in our system. We thank one referee for this suggestion.

Hydrogen abstraction from NMP has been reported and can occur at two sites: the methyl group on the nitrogen or the \textit{CH}_2 position adjacent to the nitrogen (S. M. Aschmann and R. Atkinson, Atmos. Environ., 1999, 33, 591; C. Li, T. Takanohashi, I. Saito and M. Iino, Energy Fuels, 2003, 17, 1399; G. Solignac, I. Magnier, F. Molton, C. Duboc, A. Togni, M. Martin Reviejo and K. Wirtz, J. Atmos. Chem., 2006, 54, 89).