TEMPO-Functionalized Nanoporous Au Nanocomposite for the Electrochemical Detection of $\text{H}_2\text{O}_2$

Dongxiao Wen, Qianrui Liu, Ying Cui, Huaixia Yang, and Jinming Kong

1Pharmacy College, Henan University of Chinese Medicine, Zhengzhou 450008, China
2School of Environmental and Biological Engineering, Nanjing University of Science and Technology, Nanjing 210094, China

Correspondence should be addressed to Huaixia Yang; yangh uaixia886@163.com and Jinming Kong; j.kong@njust.edu.cn

Received 8 March 2018; Accepted 26 April 2018; Published 10 June 2018

A novel nanocomposite of nanoporous gold nanoparticles (np-AuNPs) functionalized with 2,2,6,6-tetramethyl-1-piperidinyloxy radical (TEMPO) was prepared; assembled carboxyl groups on gold nanoporous nanoparticles surface were combined with TEMPO by the "bridge" of carboxylate-zirconium-carboxylate chemistry. SEM images and UV-Vis spectroscopies of np-AuNPs indicated that a safe, sustainable, and simplified one-step dealloying synthesis approach is successful. The TEMPO-np-AuNPs exhibited a good performance for the electrochemical detection of $\text{H}_2\text{O}_2$ due to its higher number of electrochemical activity sites and surface area of 7.49 m$^2$ g$^{-1}$ for load bigger amount of TEMPO radicals. The TEMPO-functionalized np-AuNPs have a broad pH range and shorter response time for $\text{H}_2\text{O}_2$ catalysis verified by the response of amperometric signal under different pH and time interval. A wide linear range with a detection limit of $7.8 \times 10^{-7}$ M and a higher sensitivity of 110.403 $\mu$A M$^{-1}$ cm$^{-2}$ were obtained for detecting $\text{H}_2\text{O}_2$ at optimal conditions.

1. Introduction

Hydrogen peroxide ($\text{H}_2\text{O}_2$) is the smallest and simplest peroxide, which could be generated in many biological processes and cause severe oxidative damage [1–3]. It has been applied in many biosynthetic reactions and also plays an important role in various fields, especially in immune cell activation, vascular remodelling, apoptosis, stomatal closure, root growth, and so on [1, 4]. Because of this, many practical explorations have been done in the detection and monitoring of $\text{H}_2\text{O}_2$ in pharmaceutical [1–6], biological [5, 6], clinical [1, 3, 7], chemical, textile [1, 2, 8–10], and food industries [1–6, 10]. The dynamic equilibrium of the production and consumption of $\text{H}_2\text{O}_2$ is closely interlinked with the quality of our life; therefore, the development of new materials and techniques for quantitative detection of $\text{H}_2\text{O}_2$ at a trace level has presented a significant role in the fundamental studies of diagnostic and monitoring applications. However, most existing techniques suffer from many drawbacks, such as inherent instability, time consuming, poor selectivity, low activity, complicated, and costly immobilization procedures [1–3, 9–11]. Thus, it is necessary to develop new electrochemical sensors overcoming those drawbacks for accurate and sensitive detection of $\text{H}_2\text{O}_2$.

A number of new materials such as metals nanomaterials, carbon nanotubes, quantum dots, nanocomposites, and redox substances have been employed for the detection of $\text{H}_2\text{O}_2$ [1–9]. In recent studies, nanoscale hollow materials have attracted scientists’ attention for their unique porous structural and versatile properties. There have been some reports already about its applications in fuel cell and remarkable catalytic oxidation activities to CO, methanol, hydrazine, and so on [12–15]. Particularly, nanoporous gold nanoparticles (np-AuNPs), one of the most important nanoscale hollow materials, combining their high surface areas with tunable surface plasmon resonance (SPR) features, can serve larger immobilized surface and excellent catalytic sites in chemical reactions and also provide a substrate material to manufacture functionalized nanocomposite [12, 15]. Meanwhile, TEMPO plays a key role in chemistry and biology as an organic redox catalyst for alcohol, aldehydes, and ketones [16–21] and has shown its catalytic oxidation potential to
H₂O₂ due to the electrochemical oxidation of stable nitroxyl radical [22, 23]. Recently, the nitroxyl radical of TEMPO has been reported to attach to solid surfaces by chemical routes to form supramolecular assemblies with high coverage of catalytic radicals [24, 25]. Based on considerations above, in order to take full advantage of their inherent catalytic oxidation activity of the np-AuNPs and the TEMPO, we explore the synergistic effect of TEMPO-functionalized np-AuNPs for the peroxidase-like activity response and use this feature for H₂O₂ detection.

Most nanoporous materials are prepared by using traditional dealloying, template, electrochemical methods, and directed self-assembly [9,10,26], and the dealloying approach is mostly adopted to synthesize various nanoporous hollow structures [12–15], using AgCl templates to get the nanoporous gold [12, 26, 27]. However, this traditional dealloying approach is limited by its harsh cumbersome and high-heat conditions [12, 28]. In this work, we prepared the zero-dimensional hollow np-AuNPs via an improved one-step dealloying synthesis in moderate conditions instead of traditional two-step dealloying methods. To improve the peroxidase-like activity of the np-AuNPs and TEMPO, we assembled mercaptoacetic acid (MA) on the surfaces of np-AuNPs/GCE, and then the 4-carboxy-TEMPO was connected to the mercaptoacetic acid through Zr⁴⁺ of np-AuNPs/GCE, and then the 4-carboxy-TEMPO was washed the MA-np-AuNPs/GCE with ultrapure water to 5% retention. A platinum wire auxiliary electrode, a saturated calomel reference electrode, and a modified glassy carbon electrode (GCE, Φ=3 mm, as the working electrode) are included in a standard three-electrode cell. All ultrapure water (≥18.25 MΩ) for experiments was obtained from a Millipore Milli-Q water purification system.

2.3. Preparation of Nanoporous Gold Nanoparticles. The nanoporous gold nanoparticles (np-AuNPs) were prepared according to the literature with some modification [12]. Briefly, 160 µL solution of hydroquinone (28 mM) and 10 mM AgNO₃ aqueous solution (60 µL) were mixed firstly into 4.5 mL PVP solution (90 mM) in sequence. Then the mixed system came to an equilibrium stirring for 5s, 650 r.m.p. Next, 40 mM HAuCl₄ (100 µL) was added dropwise into the system at room temperature under gentle stirring. After 3 min standing, the color of the reaction liquid turned to stable (reddish brown), and then the residual AgCl, which is formed during the reaction, was removed by the additional concentrated NH₄OH (1 mL). Finally, the resulting solution was centrifuged and washed repeatedly with ultrapure water, 7,000 r.p.m. 5 min, to collect the np-AuNPs. Finally, 1.0 mg of np-AuNPs sample was fully suspended in a mixed solution (1.0 mL ethanol, 1.0 mL of 90 mM PVP) by ultrasonication.

2.4. Electrode Fabrication of Tempo-Np-Au NPs/GCE. The bare glassy carbon electrode (GCE) was polished with 0.30µM Al₂O₃ slurries on the chamois leather until a mirror-like surface is obtained. Next, it was ultrasonically cleaned with absolute ethanol and ultrapure water, dried with N₂, Then, the electrode was subjected to cyclic voltammetry (CV) in 0.1 M KCl with the potential of -0.4 V and 1.6 V, at a scan rate of 100 mVs⁻¹, until the reproducible cyclic voltammograms were obtained.

A quantity of 1µL of the prepared np-AuNPs suspension was dropped on the surface of freshly pretreated bare GCE above. After the GCE was air dried at room temperature, 20 µL of 0.2 mM MA solution was added onto the np-AuNPs/GCE surface and incubated for 1.0 h at room temperature; by so doing, a MA self-assembled monolayer formed on the surface of np-AuNPs (MA-np-AuNPs/GCE) via chemisorption and the chemistry of formation of MA-SAM on the np-AuNPs surface; a procedure probably involves the oxidative addition to form the S-Au bond by losing the hydrogen as H₂ or H₂O [29, 30]. Subsequently, the TEMPO-np-AuNPs/GCE was prepared by carboxylate- zirconium-carboxylate chemistry [31–33]. After washing the MA-np-AuNPs/GCE with ultrapure water to remove the remaining MA, the electrode was immersed in ZrOCl₂-8H₂O 60% ethanol solution (5.0 mM) for 30 min. The modified electrode was taken out and then washed with absolute ethanol, dried with N₂. Next, the electrode was incubated in 20 µL of 0.2 mM 4-carboxy-TEMPO solution for 30 min, followed by rinsing with ultrapure water to remove remaining reactants. TEMPO-functionalized nanoporous Au nanocomposite electrode was prepared. The establishment of this sensor for H₂O₂ detection is depicted in Figure 1.
3. Results and Discussion

3.1. Characterization of Modified Electrode by SEM and UV-Vis. By following procedures above, we successfully synthesized the free-standing np-AuNPs under moderate conditions by using a one-step aqueous solution-based approach, which circumvents the limits of stringent and harsh multistep protocols of traditional dealloying approaches. As shown in Figures 2(a) and 2(b), SEM images indicate that the as-synthesized np-AuNPs have a shape of spherical and exhibit an extremely roughened surface, which is consistent with the result of Srikanth’s report [12].

As it has been known to all, AgNPs and AuNPs including np-AuNPs are attractive due to their surface plasmon resonance (SPR) properties. Ag\(^{+}\) AuNPs and np-AuNPs often exhibit different spectral absorptions in the UV-Vis wavelength region [26, 34]. Therefore, we can use absorption spectroscopy to monitor the process of reaction; the UV-Vis spectral absorptions of the as-prepared np-AuNPs are shown in Figure 2(c). A peak appeared at 324 nm, which is the typical UV-Vis absorbance peak of Ag\(^{+}\) [34]. After 100\(\mu\)L H\(_{2}\)AuCl\(_4\) solution was added in the system, a significant increase of signal appeared at 530 nm, which is caused by the grown of AuNPs on the surface of AgCl templates in the kinetic-controlled process [35]. Subsequently, with NH\(_4\)OH added to stop the reaction, the surface plasmon resonance (SPR) peak of 530 nm underwent a red shift to 596 nm-700 nm with longer and broader profile, which belongs to the characteristic peak of porous gold nanoparticles. The typical SPR peak of AuNPs always exhibited around 530 nm; the Ag\(^{+}\) broad peak maximum usually occurs at 410 nm, in solution [34]. It was found that, after the centrifuge step to remove residual AgCl, the absorbance peak at 324 nm disappeared, indicating the formation of the np-AuNPs.

3.2. Characterization of Modified Electrode by Cyclic Voltammetry. To characterize the modified electrode TEMPO-np-AuNPs, the typical cyclic voltammetry was performed. The cyclic voltammogram (CV) curves of different modified electrodes in the absence of oxygen in 0.1 M PBS supporting electrolyte (pH = 7.0), at 50 mVs\(^{-1}\), are shown in Figure 3(a). For the bare GCE, there are no redox peaks (A). With the np-AuNPs added on the GCE surface, a broad peak arose at 1.2-1.5V (B) due to the increased effective electroactive area of np-AuNPs. What is more, when MA self-assembled on the np-AuNPs, an obvious peak increase was observed due to the S-Au bond formation (C). By comparing the CVs of C and D, a new pair of well-behaved redox peaks of 0.38 V (cathodic peak) and 0.82 V (anodic oxidation peak) was obtained after 4-carboxy-TEMPO was finally attached to the former-electrode surface. In addition, the decreases of a pair of peaks at -0.1 V/1.2-1.5 V due to the steric hindrance also indicated that 4-carboxy-TEMPO was successfully connected with MA-np-AuNPs via carboxylate-zirconium-carboxylate chemistry.

Further, we characterized the formed np-AuNPs with electroactive activity through CV at different scan rates from 10 mVs\(^{-1}\) to 200 mVs\(^{-1}\). As shown in Figure 3(b), the intensities of the electric current increased linearly along with the increase of scan rates \((k_1 = 3.934 (+ 0.122))\], which confirmed that the electroactive np-AuNPs attached on the GCE electrode surface by adsorption according to the theory of homogeneous redox catalysis [22], rather than other interactions. As mentioned above, the CV of TEMPO-np-AuNPs has a new pair of redox peak (0.38 V/0.82 V), while the CV of np-AuNPs did not. As displayed in Figure 2(c), with the scan rates increasing, the current intensity of this redox couple enhanced with a slight redshift; both the \(I_{pa}\) (black spots) and \(I_{pc}\) (red spots) peak currents were linearly proportional to
3.3. Electrochemical Measurement of Np-AuNPs Active Surface Areas. The effective areas of different surface modification were estimated by CV method. As shown in Figure 3(d), the anodic oxidation current of the three curves all rose at about 1.2 V and had a typical reduction peak around 0.75 V, which is caused by the reversible redox reaction in 0.5 M H$_2$SO$_4$. However, the MA-np-AuNPs (B) exhibit a higher peak at 0.75 V and sustain a large redshift as compared with np-AuNPs (A), indicating that it is hard to be oxidized with H$_2$SO$_4$ due to the increased impedance of charge transfer after MA is immobilized on the surface of np-AuNPs. When the TEMPO is chemically modified on MA-np-AuNPs (C), the cathodic peak (around 0.75 V) had an obvious enhancement, which may be caused by the diffusion layer of TEMPO- and is a three-dimensional steady-state. Besides, the rough porous surface of np-AuNPs contributed to generate the multimodal of the curve (e.g., 0.20 V-0.45 V of curve C).

For each experiment, the amount of np-AuNPs used was the same; CV of np-AuNPs in 0.5 M H$_2$SO$_4$ (A) were measured to calculate their electroactive surface areas via integrating the area of the gold oxide reduction curve. An electroactive surface area of 7.49 m$^2$g$^{-1}$ for np-AuNPs is obtained by Randles-Sevcik equation and assuming a specific charge of 450 $\mu$C cm$^{-2}$ for the gold oxide reduction [12, 44]. Notably, the active surface areas of np-AuNPs are higher than that of the commercial Au electrodes, Au nanoparticle, Au
Figure 3: (a) Different modified GCE in 0.1 M PBS (pH = 7.0) under N₂. Scan rate: 50 mV s⁻¹. (A) Bare GCE, (B) np-AuNPs/GCE, (C) MA-np-AuNPs/GCE, and (D and inset) TEMPO-np-AuNPs/GCE. CVs of np-AuNPs (b) and TEMPO-np-AuNPs (c) on the GC electrode surface under N₂, 0.1 M PBS (pH = 7.0) at different scan rates from 10 mV s⁻¹ to 200 mV s⁻¹. Inset: (b) the linear relationship between the scan rate and the currents at a potential of 0.55 V. (c) The linear relationship between anodic (black spots, at 0.95 V) and cathodic (red spots, at 0.38 V) peak currents and scan rate. (d) CVs of different electrodes in 0.5 M H₂SO₄ under N₂, 0.1 M PBS (pH = 7.0) at a scan rate of 100 mV s⁻¹. (A) np-AuNPs, (B) MA-np-AuNPs, and (C) TEMPO-np-AuNPs.

nanocoral, and other kinds of Au electrodes [45], near 49 times high according to report. The large surface-to-volume ratio of metal nanoparticles is closely related to its high electrical conductivity, catalytic ability, and surface reaction activity such as for the detection of H₂O₂ [2, 45].

3.4. Electrochemical Detection of H₂O₂. As we have known already, the electrochemical behavior of np-AuNPs modified GCE electrode for H₂O₂ was studied by CV. As shown in Figure 4(a), when 3 mM H₂O₂ was added to 0.1 M phosphate buffer saline (pH=7.0) under N₂, compared to the bare GCE, distinct increases of the AuNPs/GCE and np-AuNPs/GCE response currents were observed. Interestingly, the np-AuNPs/GCE had a significant advantage in the difference of the current intensity (ΔI) under the same condition in the presence of H₂O₂. As one of the most famous catalytic materials, Au nanomaterials with different shapes and structures have been suggested good responses for electrocatalytic H₂O₂ [1–3]. The current responses toward H₂O₂ concentration over the range of 0.5 μM-100 μM on the np-AuNPs/GCE (Figure 4(b)) were studied. The current intensity had a steady rise with the increasing of H₂O₂ concentration. The HO⁻ radical resulted from H₂O₂ would
AuNPs sharply rose to 403 μA (Figure 5(a)). When H₂O₂ was put into 0.1 M PBS electrolyte buffer, an obvious increase of the peak current at 0.95 V in the CV of TEMPO-np-AuNPs is observed (R² = 0.9973) in the range of 26 μM to 100 μM. Applied potential: 0.55 V. Inset: plot of electrocatalytic current of H₂O₂ versus its concentrations.

be stabilized by the np-AuNPs [3, 46]. The surface property of np-AuNPs may influence the catalytic ability of H₂O₂ and the charge-transfer processes. When the H₂O₂ concentration is below 26 μM, it exhibited a linear correlation of I (μA) = 0.14749C (μM) +14.77624 (R² = 0.99057, RSD = 4.2%) between 0.5 μM and 26 μM. With the increase of stable surface charge transfer for the continuous regeneration of the charge-transfer processes, this sensor can advance the application of electrocatalytic electron transfer way to detect H₂O₂.

3.5. Factors Influencing Detection. We all know that time and pH can directly affect the stability of the reaction proceeding and catalytic of enzymes. The analytical performance of the sensor is usually closely linked with the stability of the materials on electrode, which was partially influenced by the time and pH of electrolyte solution. So the pH and time were tested for their electrical catalytic activity to an optimal condition in this work. Firstly, the effect of pH on the potential of electrocatalytic activity on the TEMPO-np-AuNPs at the range of 5.0 to 8.5. Compared with the pH decided enzyme-modified electrode detection methods, this sensor can advance the application of TEMPO-np-AuNPs to detect H₂O₂ in vivo measurements [1].

The surface charge transfer for the continuous regeneration of the charge-transfer processes, this sensor can advance the application of electrocatalytic electron transfer way to detect H₂O₂.

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Further, we investigated the time factor in the presence of 5 mM H$_2$O$_2$. Figure 5(d) shows TEMPO-np-AuNPs can catalyze H$_2$O$_2$ to produce O$_2$, immediately, which then comes to an equilibrium state with a maximum current strength in 40 min. By comparing the current strength at 1.6 V, we found that the floating electric potential of 40 min only takes 13.64% in the whole catalytic oxidation process. It is possible that this remaining increase is closely related to the irregular gaps and surface of the np-AuNPs, which may lead to the time retardation to the reversible one-electron behavior of the redox couple TEMPO/TEMPO$^+$ [46]. So we selected 40 min stirring constantly after H$_2$O$_2$ was added to the 0.1 M PBS with the buffer system of pH=7.0 to ensure the sufficient current response in the following measurements.

3.6. Steady-State Amperometric Response of H$_2$O$_2$. To evaluate the applied potential of TEMPO-MA-np-AuNPs/GCE as peroxidase-like. We investigated the response of the amperometric signal under the optimal condition with the 0.95 V as the applied potential ($E_{app}$). As shown in Figure 6, the TEMPO-np-AuNPs/GCE not only can achieve a quick steady-state current within 10 s, but also has a stronger response (C) compared to the np-AuNPs/GCE (B).

Moreover, with the addition of same amount H$_2$O$_2$ in every interval of 100 s, the TEMPO-np-AuNPs nanocomposite on the GCE exhibited a good linear chronoamperometric response to H$_2$O$_2$ from 2.0 μM to 500 μM, which is shown in Figure 6. As what we have seen, when the concentration of H$_2$O$_2$ is below 10.0 μM, the slope of the

**Figure 5:** (a) CVs of different GCE under the absence (A and B) and presence (C and D) of 5 mM H$_2$O$_2$ in N$_2$-saturated 0.1 M PBS (pH = 7.0) at a scan rate of 50 mVs$^{-1}$. (A) Bare GCE (B and C) np-AuNPs/GCE, and (D) 4-carboxy-TEMPO-np-AuNPs/GCE. Optimization of experimental conditions: (b) CVs of the TEMPO-np-AuNPs/GCE in different pH values at the range of 2.0 ∼ 9.0 and (c) the relationship between the net current at 0.95V with different pH. (d) CVs of TEMPO-np-AuNPs/GCE in 0.1 M PBS (pH = 7.0), in the presence 5 mM H$_2$O$_2$, in different times from 0 min to 60 min, N$_2$-saturated, 50mVs$^{-1}$. Inset: the changes of the electric current (1.6 V) with the time increasing.
Mechanism process:

Scheme 1: Equilibrium of single electron between TEMPO-np-AuNPs and H$_2$O$_2$.

linear ($K_b = 0.66608, R^2 = 0.98185$) is lower than that in a higher concentration range. In other words, the sensitivity of electrocatalytic oxidation activity to H$_2$O$_2$ is improved with the concentration of H$_2$O$_2$ increased in the solution and there are no substrate inhibition effects occurring at high concentration of H$_2$O$_2$. With successive addition of H$_2$O$_2$ (n = 5) as shown in Figures 6(b) and 6(c), we can obtain two linear regression equations: (1) below 10.0 µM, $y = 0.66608x + 4.51516$ ($R^2 = 0.98185$); (2) 10.0 µM to 500 µM, $y = 0.1078x + 0.6047$ ($R^2 = 0.9998$). Error bars represent the standard deviations of five independent measurements. The repeatability of the system is assured by a relative standard deviation (RSD) of 2.8%. We use conventional three times the standard deviation of [LOD = 3(RSD/slope)] to estimate the limit of detection (LOD) of H$_2$O$_2$, which is 0.78 µM in our work and lower than some Au nanomaterials and nitroxide derivatives for H$_2$O$_2$ detection as listed in Table 1. The results of the stability measurements indicated that np-AuNPs still keep its stable original roughened surface and porous structure after 30 days of storage at 4°C. Moreover, compared with previous results, the TEMPO-np-AuNPs/GCE can retain 98.6% of its initial current response results in the same measurement conditions. Therefore, the TEMPO-np-AuNPs nanocomposite has a remarkable superiority for the electrochemical detection of H$_2$O$_2$ over the conventional electrochemical sensing materials and most of the reported Au nanomaterials and nitroxide derivatives probes (Table 1).

3.7. Interference Study. Some coexisting potential electroactive species may affect the sensor response, such as sucrose (SC), glucose (GC), dopamine (DA), and ascorbic acid (AA) [3, 48]. Good selectivity is crucial to ensure and facilitate the accurate assessment for biosensor in a particular application. For better detection in vivo, the interference study of TEMPO-np-AuNPs/GCE for the electrochemical detection of H$_2$O$_2$ was carried out to evaluate its practical feasibility. Results were shown in Figure 6(d); a discernible slight fluctuation is hard to see after the abundant successive addition of each interfering species (SC/GC/DA/AA), while no obvious interference signal was observed. Notably, the TEMPO-np-AuNPs/GCE had an 18µA response ($E_{app} = +0.95$ V) as soon as another 0.1mM H$_2$O$_2$ was injected into the complex system of interference. The initial small responses caused by SC, GC, DA, and AA belong to normal current fluctuations, which is a deductible interference compared with that caused by H$_2$O$_2$. TEMPO-np-AuNPs/GCE exhibited an acceptable selectivity towards the practical in vivo electrochemical detection of H$_2$O$_2$. 
Table 1: Comparison of recent Au nanomaterials and nitroxide derivatives for H$_2$O$_2$ detection.

| Electrode design                | L. R. (M)      | D.L. (µM)  | Stability          | Reference |
|---------------------------------|----------------|------------|--------------------|-----------|
| HRP/Cys/AuNP/ITO                | 8.0×10$^{-6}$–3.0×10$^{-3}$ | 2.00     | 83% (12 weeks)     | [6]       |
| HRP/ CaCO$_3$-AuNPs/ATP/Au     | 5.0×10$^{-7}$–5.2×10$^{-3}$ | 0.10     | 96.4% (30 days)    | [36]      |
| HRP-nano-Au                     | 1.2×10$^{-5}$–1.1×10$^{-3}$ | 6.10     | 75% (5 weeks)      | [37]      |
| HRP/AuNPs/poly(St-co-AA)        | 8.0×10$^{-6}$–7.0×10$^{-3}$ | 4.00     | 97.8% (60 days)    | [38]      |
| Au/CeO$_2$ nanocomposite        | 0–3.0×10$^{-4}$ | 5.00     | ---                | [9]       |
| AuNPs-N-GQDs/GC                 | 2.5×10$^{-5}$–1.3×10$^{-2}$ | 0.12     | 89%, 3 weeks       | [10]      |
| Poly(BCB)/Au-NPs/GCE           | 6.0×10$^{-5}$–1.0×10$^{-2}$ | 0.23     | 95% (2 week)       | [39]      |
| GC/MTMOS-Au$_{73}$Ag$_{27}$     | 1.0×10$^{-5}$–7.0×10$^{-5}$ | 1.00     | ---                | [40]      |
| Hb/Au nanoflowers/CNTs/GCE     | 1.0×10$^{-6}$–6.0×10$^{-4}$ | 7.30     | ---                | [41]      |
| TEMPO/GCE                       | 1.0×10$^{-7}$–1.0×10$^{-3}$ | 0.05     | 100%, 3 months     | [22]      |
| ChOx/GCE                        | 2×10$^{-3}$–2.5×10$^{-3}$ | 20.0     | ---                | [23]      |
| TEMPO-MA-np-AuNPs/GCE          | 2.5×10$^{-6}$–5.0×10$^{-4}$ | 0.78     | 98.6% (30 days)    | This work |

Figure 6: (a) Chronoamperometric responses observed at (A) bare GCE, (B) np-AuNPs/GCE, (C) TEMPO-np-AuNPs/GCE, and (D) TEMPO-np-AuNPs/GCE after successively injecting H$_2$O$_2$, in 0.1 M PBS (pH=7.0). Applied potential: 1.10 V and the calibration plot between the oxidation current and the H$_2$O$_2$ concentration. (b and c) Amperometric response of H$_2$O$_2$ and interferants at TEMPO-np-AuNPs/GCE at 1.10V in PBS (0.1 M, pH = 7.0). Injection sequence: 0.1mM H$_2$O$_2$, 100 mM SC, 100 mM GC, 100 mM DA, 100 mM AA, and 0.2 mM H$_2$O$_2$. 

R$^2$=0.98185

y=0.66608x+4.51516

R$^2$=0.99982

y=0.10778x+0.60468

R$^2$=0.99468
4. Conclusions

In this work, we demonstrated a strategy for producing large specific surface area nanoporous gold nanoparticles and manufactured the TEMPO-functionalized np-AuNPs nanocomposite; the electrooxidation to $\text{H}_2\text{O}_2$ with a high density of radicals on the TEMPO-np-AuNPs surface was also investigated. It is worth noting that this new nonenzymatic $\text{H}_2\text{O}_2$ probe is prepared under a gentle, secure, low-cost, and simple procedure. When the np-AuNPs are combined with 4-carboxy-TEMPO, the advantages of their unique properties for the electrochemical detection of $\text{H}_2\text{O}_2$ come to a double effective enhancement. Compared the peroxidase-like activity based on the direct electron transfer of the TEMPO-np-AuNPs/GCE to the electrochemical hydrogen peroxide biosensors of TEMPO-based ligand, we obtained a wide linear range for $\text{H}_2\text{O}_2$ detection. The enzyme-like activity with low detection limit, high sensitivity, low-cost, anti-interference, good reproducibility, and stability of this nanocomposite with TEMPO-based ligand make a contribution to improve the detection current signal of $\text{H}_2\text{O}_2$. Furthermore, this TEMPO-functionalized np-AuNPs nanocomposite study plays a significant role in facilitating the research of biosensors, gold nanomedicine, catalysis, or cancer therapy.

Data Availability

All the data used to support the findings of this study are included within the article or are available from the corresponding author upon request. These data are available for unrestricted use, distribution, and reproduction in any medium, provided the original work is properly cited.

Conflicts of Interest

The authors declare that they have no conflicts of interest.

Acknowledgments

This work was supported by the National Natural Science Foundation of China (Grant no. 21575066).

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