A Study on the Condensation Reaction of 4-Amino-3,5-dimethyl-1,2,4-triazole with Benzaldehydes: Structure and Spectroscopic Properties of Some New Stable Hemiaminals

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Academic Editor: Jean Jacques Vanden Eynde

Received: 15 July 2015 / Accepted: 9 September 2015 / Published: 17 September 2015

Abstract: Studies on the stable hemiaminals and Schiff bases formation in the reaction of substituted benzaldehydes with primary 3,5-dimethyl-1,2,4-triazole 4-amine were carried out under neutral conditions. These products were investigated by IR, Raman, MS, 1H- and 13C-NMR spectra as well as by X-ray crystallography. The effect of reaction conditions: temperature, polarity of the solvents utilized, substrate concentration and the ortho and para benzaldehyde substituents on the yield of products was also examined.

Keywords: 3,5-dimethyl-1,2,4-triazole 4-amine; stable hemiaminals; chemical reactivity; Schiff bases; X-ray structures

1. Introduction

1,2,4-Triazoles and their derivatives have attracted significant attention in several different areas. These nitrogen-rich compounds represent one of the most biologically active classes of the chemical
species. This arises from their ability to bind to a variety of enzymes and receptors in biological systems via diverse non-covalent interactions [1–3]. The antimicrobial activity 3,5-dimethyl-1,2,4-triazole derivatives has been reported [4]. The Schiff bases obtained from 4-amino-3,5-dimethyl-1,2,4-triazole inhibit endocytosis [5] and in copper(II) complexes inhibit protein tyrosine phosphatases [6]. Additionally, 4-Amino-3,5-dimethyl-1,2,4-triazole is a very interesting bridging ligand. It coordinates with Mn(II), Co(II) [7], Cu(II) [8] and mixed valence cobalt [9] forming trinuclear coordination compounds which exhibit very interesting magnetic properties. Polymeric Ag(I) complexes with amino-triazole acting as a tridentate-N donor ligand were also obtained [10]. Trinuclear azide complexes of Cu, Co, Ni, Zn, Mn and Cd were investigated as energetic materials [11]. The inhibitive effect of 3-methyl-4-amino-1,2,4-triazole on the corrosion of copper-nickel alloys has been reported [12]. Furthermore, 4-amino-1,2,4-triazoles and its 3,5-dimethyl derivative readily react with alkylated agents at the N-1 position forming low melting points ionic liquid salts [13].

It is well known that the primary and secondary amines react by nucleophilic addition with carbonyl compounds to give intermediate tetrahedral addition products called hemiaminals [14] as a first step of condensation reaction [15]. The next step is the dehydration of that compound which leads to the formation of stable imines, enamines, hydrazones and related compounds [16]. Typically, the hemiaminals are short-lived species. They are sometimes detected by spectroscopic methods: IR [17], NMR [18–20] and by X-ray observation [21]. The tetrahedral carbinolamine group can be stabilized also by metal ions. Nitrogen-containing aromatic heterocyclic aldehydes react with di-(2-picoly)amine in presence of Zn²⁺ salts to form labile tris-(2-picoly) hemiaminal zinc complexes [22]. Rhodium (III) complexes of o-diphenylphosphinobenzaldehyde react with dihydrazones to give ionic species with a new tridentate PNN–hemiaminal type ligand [23].

One of the stable hemiaminals was obtained as a solid in the reaction between trifluoroacetaldehyde and a secondary amine, N-benzyl piperazine. Fluoral acts as an efficient nucleophilic trifluoromethylating agent towards non-enolizable carbonyl compounds under mild conditions [24]. The stable-in-solution hemiaminals were prepared in the reactions of methyl 3,3,3-trifluoropyruvate and benzylic monoamines and diamines. A similar product obtained from aniline was stable only under inert atmosphere but not in solutions [25]. Hexafluoroacetone reaction with 2-(aminomethyl)aniline also results in a stable tetrahedral product with benzylic amino group [26]. Another stable hemiaminal was obtained from 4-cyclohexyl-3-thiosemicarbazide and di-2-pyridyl ketone [27]. In this case, the formation of a carbinolamine was strongly dependent on the intramolecular hydrogen interactions between N–H···N and O–H···N (py) atoms. The addition of LiBH₄ or the Grignard nucleophilic reagents MeMgCl and PhMgBr to amides is another way of making stable hemiaminals. Aromatic, alkyl and α,β-unsaturated N-acylpyrrole derivatives or one-carbon bridged amides allow one to obtain stable products with good yields [28,29]. Recently, the stable hemiaminals coming from the reaction between 4-amino-1,2,4-triazole and nitro-substituted [30] or cyano-substituted [31] benzaldehydes in acetonitrile under neutral conditions were also obtained. Eleven of them were structurally and spectroscopically characterized.

The present paper describes a novel application of 4-amino-3,5-dimethyl-1,2,4-triazole in the preparation of stable hemiaminals. The effects of temperature, solvent, structure and concentration of the reagents on the product yields and stability are considered. Our results can contribute to a better understanding of the mechanism of hemiaminal formation from aminotriazoles and benzaldehydes.
2. Results and Discussion

By using a published method, 4-amino-4H-3,5-dimethyl-1,2,4-triazole (1) was obtained from one-pot solvothermal reaction of acetonitrile with hydrazine hydrate [32]. The syntheses of (aryl)(3,5-dimethyl-4H-1,2,4-triazole-4-ylamino)methanol (2–14) and \(N\)-benzylidene-4H-3,5-dimethyl-1,2,4-triazole-4-amine (15, 16) derivatives were accomplished according to the reaction outlined in Scheme 1. It should be noticed that the stable hemiaminals are formed only from aromatic aldehydes containing electron-withdrawing groups or atoms.

**Scheme 1.** Synthetic pathway for preparation of compounds 1 to 16.

2.1. X-ray Diffraction

Suitable crystals for X-ray diffraction were obtained for compounds 2, 5, 9, 10 and 15 (Figure 1). The molecular structure consists of two phenyl and triazole aromatic rings connected with the C\(\text{Ph}\)–C–N\(_{4\text{Tr}}\)–N\(_{3\text{Tr}}\) sequence. In the hemiaminals, the C and N\(_{4\text{Tr}}\) atoms are tetrahedral with sp\(^3\) hybridization, which enables formation of four stereoisomers (RS, SR, RR and SS). For imine 15, there is a C–N\(_{4\text{Tr}}\) double bond with sp\(^2\) hybridization. The general atom numbering and selected parameters are summarized in Table 1.
Figure 1. Molecular structure of compounds 2, 5, 9, 10 and 15.

Table 1. Selected geometrical parameters for hemiaminals (2, 5, 9, 10) and imine (15).

| R         | C1-C14 | C14-N4 | N4-N3 | C14-O14 | C1t-C3t | N1-N2 | C1-C14-N4-N3 | Phenyl-Triazole |
|-----------|--------|--------|-------|---------|---------|-------|---------------|-----------------|
| 2-NO2 (2) | 1.529(2) | 1.514(3) | 1.425(2) | 1.363(3) | 1.435(2) | 1.360(3) | 177.1(1) | 19.99(1) |
| 2,4-(NO2)2 (5) | 1.526(3) | 1.461(2) | 1.406(2) | 1.397(3) | 1.481(3) | 1.396(4) | -176.3(2) | 21.27(1) |
| 2-Cl (9)  | 1.513(4) | 1.467(4) | 1.410(4) | 1.425(4) | 1.484(4) | 1.403(3) | 179.0(2) | 9.9(2)  |
| 4-CHO (10) | 1.510(2) | 1.482(2) | 1.409(2) | 1.401(2) | 1.478(2) | 1.399(2) | 174.3(1) | 28.8(2) |
| 4-NO2 (15) | 1.465(2) | 1.265(2) | 1.418(2) | - | 1.468(3) | 1.367(2) | 179.1(1) | 70.48(2) |

The hemiaminal molecules in 2, 5, 9 and 10 form a centrosymmetric (RS-SR) dimer linked by a O–H···N1Tr hydrogen bond (see Figure 2). A strong π-π interaction involving pairs of triazole rings additionally stabilizes the dimers (Table 2).
The structures presented above differ from those obtained by us earlier [30,31] where hemiaminals derived from 3,5-unsubstituted triazoles occur in two conformers: stretched (with configuration RS or SR) and twisted (RR or SS). Furthermore, in the title compounds the centrosymmetric dimers are observed whereas hemiaminals described by us previously [30,31] form infinite polymeric chains or noncentrosymmetric dimers.

2.2. Spectral Studies

The characteristic IR and Raman spectral bands of hemiaminals are given in Table 3. Characteristic strong ν(C=O) stretching vibration at about 1700 cm$^{-1}$ observed in the infrared spectra in aromatic aldehydes, as well as bands observed at 3243 cm$^{-1}$, 3152 cm$^{-1}$ and 1650 cm$^{-1}$ which were assigned to the υasNH$_2$, υsNH$_2$ and σ,ωNH$_2$ vibrations [33] respectively for 4-amino-3,5-dimethyl-1,2,4-triazole, disappear after condensation reaction. A comparison between the NH and OH stretching bands, which were observed for hemiaminals, shows that they appear in the same spectral region. In the IR spectra, the strong OH bands sometimes mask the weaker NH absorption, but in the Raman spectra, the OH bands are very weak [34]. The –OH stretching vibration of the hydroxyl group is observed as a distinct peak at about 3200–3300 cm$^{-1}$ in the IR spectrum. Their values increase with decreasing dC-O bond distance (Table 1). The intramolecular hydrogen bonding interactions of C–OH with N$_{2tr}$ observed in the crystal structures are confirmed by an additional broad shallow –OH stretching peak observed at about 3100 cm$^{-1}$. The band appearing in the Raman spectra at about 3100 cm$^{-1}$ is assigned to the stretching vibration of –NH.
Table 3. Selected spectral data of hemiaminals R1C*H(OH)NHR2.

| R1                  | Vibration Frequencies (cm\(^{-1}\)) | \(^{1}\)H-NMR \(\delta\) (ppm), \(J\) (Hz) | \(^{13}\)C-NMR (ppm) |
|---------------------|------------------------------------|------------------------------------------|-----------------------|
|                     | \(v_{\text{OH}}\) \(a\) | \(v_{\text{OH, N}}\) \(a\) | \(v_{\text{NH}}\) \(b\) | \(\delta_{\text{NH}}\) | \(\delta_{\text{OH}}\) | \(\delta_{\text{CH}}\) | \(J_{\text{CH-NH}}\) | \(J_{\text{CH-OH}}\) | \(\delta(C^*)\) |
| 2-NO\(_2\) C\(_6\)H\(_4\) (2) | 3308 | 3114 | 3090 | 7.15 | 6.91 | 5.96 | 8.12 | 5.33 | 79.2 |
| 3-NO\(_2\) C\(_6\)H\(_4\) (3) | 3312 | 3115 | 3090 | 7.19 | 6.79 | 5.62 | 7.17 | 5.78 | 83.2 |
| 4-NO\(_2\) C\(_6\)H\(_4\) (4) | 3304 | 3079 | 3100 | 7.17 | 6.76 | 5.58 | 7.18 | 5.67 | 83.4 |
| 2,4-(NO\(_2\))\(_2\) C\(_6\)H\(_3\) (5) | 3305 | 3106 | 3090 | 7.33 | 7.26 | 6.00 | 8.35 | 5.01 | 79.1 |
| 3-NO\(_2\),4-Cl C\(_6\)H\(_3\) (6) | 3261 | 3105 | 3080 | 7.20 | 6.86 | 5.57 | 7.44 | 5.72 | 82.7 |
| 3-NO\(_2\),4-CH\(_3\) C\(_6\)H\(_3\) (7) | 3309 | 3105 | 3070 | 7.12 | 6.70 | 5.54 | 6.87 | 5.69 | 82.6 |
| 2-Cl, 5-NO\(_2\) C\(_6\)H\(_3\) (8) | 3309 | 3070 | 3081 | 7.26 | 7.02 | 5.77 | 6.99 | 5.48 | 80.0 |
| 2-Cl C\(_6\)H\(_4\) (9) | 3265 | 3124 | 3081 | 7.05 | 6.60 | 5.76 | 6.29 | 5.15 | 80.8 |
| 4-CHO C\(_6\)H\(_4\) (10) | 3291 | 3077 | 3080 | 7.12 | 6.64 | 5.55 | 6.87 | 5.53 | 83.9 |
| 4-CN C\(_6\)H\(_4\) (11) | 3280 | 3082 | 3084 | 7.13 | 6.70 | 5.53 | 7.06 | 5.72 | 83.6 |
| 4-CF\(_3\) C\(_6\)H\(_4\) (12) | 3284 | 3126 | 3082 | 7.11 | 6.66 | 5.55 | 6.87 | 5.53 | 83.7 |
| 3-C\(_5\)H\(_4\)N (13) | 3202 | 3125 | 3078 | 7.13 | 6.25 | 5.47 | 6.87 | 5.91 | 82.2 |
| 4-C\(_5\)H\(_4\)N (14) | 3191 | 3122 | 3107 | 7.15 | 6.70 | 5.47 | 7.44 | 5.72 | 83.2 |

\(^a\): IR; \(^b\): Raman.

The NMR spectra were obtained in the DMSO solution. DMSO is one of the most polar and aprotic solvents with a high dielectric constant and, due to this, properties of the dissolving species do not come together to agglomerate. For that reason, the hydrogen bonds observed in the solid state are not detected in the \(^1\)H-NMR spectra. In the \(^1\)H-NMR spectra of the compounds 2–16, the singlet at \(\delta\) 5.73 ppm, assigned to the NH\(_2\) protons of the starting compound 1, disappeared and additional resonances assigned to the C–NH–N, C–OH and CH–N for 2–14 (Table 3) and –CH=N– (\(\delta\) = 9.17 and 9.16) for 15 and 16 were detected which confirmed the condensation between the amino and the carbonyl groups.

The proton signals of the methyl triazole substituents (1–16) were observed as singlet at \(\delta\) 2.25 ppm for amine and were shifted to a lower field in the order: amine-hemiaminal-schiff base (\(\delta(CH_3) = 2.25; 2.34 \text{ and } 2.59\) ppm for 1, 4 and 15, respectively). The \(^{13}\)C-NMR spectra of R\(_1\)C*H(OH)NHR\(_2\) showed a characteristic signal \(\delta(C^*)\) at 79.1–80.8 ppm for ortho, at 82.2–83.3 ppm for meta and at 83.2–83.9 ppm for para substituted R\(_1\) aromatic ring.

2.3. Hemiaminal Stability in Solution

The hemiaminal under investigation were stable for a long time in the crystalline form. This observation does not apply to the compounds in solution. The time dependent changes in the \(^1\)H-NMR spectra were used to determine the decomposition of the hemiaminals in DMSO solution at room temperature (Scheme 2).

Compounds 2, 4, 5 and 9 decompose slowly mostly to substrates (Figure 3), similar to the hemiaminal obtained from the 4-nitrobenzaldehyde and 4-amino-1,2,4-triazole (4nba, see Figure 4a).

A greater stability in solution is observed for compounds obtained from the 3- and 4-pyridinecarboxaldehyde. Even after one year, 13 and 14 are detected in solution in the amount of about 20% (Figure 4b). In contrast to the 3- and 4-pyridinecarboxaldehyde, the compound obtained by the condensation of 2-pyridine derivative in EtOH give solely the Schiff base 17a. The products obtained from the condensation performed in hexane solution were a mixture of the hemiaminal 17 and...
Schiff base 17a in the molar ratio 1:2. In DMSO solution, hemiaminal 17 quickly converted to the Schiff base 17a.

Scheme 2. Hemiaminal decomposition reactions.

Figure 3. Conversion (%) of the hemiaminals (HA) 2, 4, 5, 9 and 4nba to substrates (A) and Schiff bases (SB) after 120 h in DMSO solution.

Figure 4. Decomposition of the hemiaminals (a) 4nba; (b) 13 and 14 in DMSO solution as the function of time.
These observations agree well with the theoretically examined Schiff base formation mechanism from benzaldehyde and 4-amine-4H-1,2,4-triazole [35]. The reaction takes place in two steps. In the first step, the hemiaminal is formed. The formation of Schiff base through the water molecule elimination requires an internal equilibrium between the twisted conformation of hemiaminals. The \(^1\)H-NMR spectral data for all stable hemiaminals obtained from 4-amino-3,5-dimethyl-1,2,4-triazole showed that they are stretched conformers. The coupling of NH protons with vicinal CH protons is about 7 to 8 Hz (Table 3). The coupling constant \(3J_{(CH-NH)}\) for 2-pyridinyl hemiaminal is smaller (4.96 Hz) which indicates that the twisted isomer dominates in solution [36].

2.4. Hemiaminal Formation—Effect of Substituents.

To gain a better understanding of the substrate structure effect on the hemiaminal formation, a series of 2- and 4-substituted benzaldehydes was examined, focusing on their reaction with 1 in a 1:1 stoichiometry in CH3CN solution. The reaction mixtures were stirred at 50 °C over 9 h. After solvent evaporating, the remaining solids were investigated by \(^1\)H-NMR in DMSO solution.

The good correlation between the imine and hemiaminal formation and electronic effects of the substituents is observed only for para derivatives (Figure 5a). From the theoretical studies [35,37] it is known that the N–H amine bond is broken first and then the hydrogen atom is transferred to the aldehyde O atom forming an O–H bond. Subsequently, the C–N bond is formed. It seems that the hemiaminal formation must be dependent on the carbonyl C atom electrophilicity. Benzaldehydes containing electron-withdrawing (-R) substituents reduce the hemiaminal formation in order: NO2 > CN > CF3 > CHO > H. Opposite to this, in the case of the substituents containing electron-donating groups (+R), the formation yield increases in order: OH < OCH3 < CH3 < F < Cl < Br < H. The next step of reaction is water molecule elimination from hemiaminal.

**Figure 5.** Variation of the hemiaminal (HA •) and Schiff base (SB □) formation from 4-amino-3,5-dimethyl-1,2,4-triazole and (a) para and (b) ortho substituted benzaldehydes as a function of the corresponding substituent constant [38]. All reaction were performed using 0.172 mmol of substrates in CH3CN (2 mL) at 50 °C. Product yields were obtained from the CH3 \(^1\)H-NMR signals in the region of 2.00–2.50 ppm.
The C–OH bond is broken first. Then, the N–H bond is broken and finally an imine and water are formed. It seems that the stability of the C–OH bond is also dependent on the phenyl ring substituent and this relation is opposite to that described above for hemiaminal formation. The C–O bond is being broken more easily for +R than for −R substituents.

The effect of ortho substituents on the condensation product reaction is more complex (Figure 5b) than for para substituents and could not be explained by the differences in electrophilicity of the carbonyl C atom.

2.5. Hemiaminal Formation—Solvent Effect

The condensation reaction of 2-nitrobenzaldehyde with 4-amino-3,5-dimethyl-1,2,4-triazole was studied in 12 different organic solvents. The solvent effect on the reaction rate and efficiency was investigated by the 1H-NMR spectroscopy (Table 4). The results indicate a higher hemiaminal content in apolar aprotic solvents than in dipolar aprotic media. The hemiaminal yield increases with solvent hydrophobicity, whereas a polar solvent shifts the equilibrium towards the Schiff base formation. Although, at first sight, it is surprising that increasing solvent polarity diminishes the hemiaminal content, this is understandable in terms of changing substrates and products dipole moment. The rate of the first step of condensation decreases with increasing solvent polarity because the activated complex must be less dipolar than the reactants. It means that the dipole moment of the activated complex should be less than the sum of the reactant dipole moment [39]. From the theoretical study [35], it is known that the hemiaminal, as an intermediate of the condensation, is non-ionic. On the other side, due to the strong intermolecular hydrogen interaction, the existence of dimers is possible, which can reduce the polarity of a hemiaminal. The strong influence of the solvent on the second step of the formation of Schiff base and elimination of the water molecule was also observed. The rate is slowest in polar aprotic solvents with high dipole moment. It seems that the activated complex, which leads to the Schiff base, appears to be less dipolar and hence less strongly solvated. In the aprotic electron-pair donor solvents with small dipole moments, the rate of this step is faster. In the hydrogen bonding solvent such as water or iso-propanol, the hemiaminal formation can proceed via a zwitterionic intermediate. The calculations of zwitterion formation between methylamine and formaldehyde have been performed [40] and found that two water molecules reduce the reaction barriers of proton-transfer step [41]. As indicated in Table 4, the water role in the 2-nitro hemiaminal formation in acetonitrile solution is not restricted only to solvent effects [42], as water also acts as a reactive species. The catalytic properties of water molecules in this reaction were thought to be essential in order to facilitate the nucleophilic attack of the amine on the carbonyl group and the proton transfers from amine to water molecule and from water to aldehyde oxygen. The rate for the first step of condensation reaction of 1 with 2-nitrobenzaldehyde depends on the water content in acetonitrile and maximum rate acceleration was observed at 15% by volume water in acetonitrile (Table 4).

2.6. Hemiaminal Formation—Benzaldehyde Concentration and Temperature Effect

The 4-amino-3,5-dimethyl-1,2,4,-triazole is in dynamic equilibrium with the reactant aldehydes. In order to determine the experimental conditions that favor the shift of the equilibrium toward the hemiaminal as a product, the effects of temperature and benzaldehyde concentration were determined
using 2- and 4-nitro substituted benzaldehydes. As can be seen (Figure 6), the highest hemiaminal yield was obtained in the upper range of the aldehyde to amine molar ratio.

**Table 4.** Solvent effect on the hemiaminal (HA) and Schiff base (SB) formation from 2-nitrobenzaldehyde and 4-amino-3,5-dimethyl-1,2,4-triazole.

| Solvent          | HA  | SB  | K    |
|------------------|-----|-----|------|
| n-Hexane         | 70  | 4   | 17.5 |
| Cyclohexane      | 47  | 11  | 4.3  |
| CHCl₃            | 47  | 14  | 3.4  |
| Toluene          | 33  | 8   | 4.1  |
| CH₂Cl₂           | 28  | 16  | 1.8  |
| CH₃CN            | 27  | 4   | 6.8  |
| DMSO             | 48  | 12  | 4.0  |
| Pyridine         | 35  | 13  | 2.7  |
| Triethylamine    | 25  | 19  | 1.3  |
| THF              | 30  | 26  | 1.2  |
| 1,4-Dioxane      | 6   | 44  | 0.1  |
| 2-Propanol       | 38  | 12  | 3.2  |
| H₂O              | 33  | 8   | 4.1  |
| H₂O/CH₃CN (V:V)  |     |     |      |
| 1.5:0.5          | 41  | 10  | 4.1  |
| 1.0:1.0          | 42  | 7   | 6.0  |
| 0.5:1.5          | 41  | 6   | 6.8  |
| 0.4:1.6          | 44  | 6   | 7.3  |
| 0.3:1.7          | 51  | 5   | 10.2 |
| 0.2:1.8          | 51  | 6   | 8.5  |
| 0.1:1.9          | 54  | 7   | 7.7  |

a: (Iₐ/Iₐ + IₐHₐ + IₐSB) × 100 where I is integrated peak intensities of the CH₃ signals in ¹H-NMR spectrum (Iₐ-amine, IₐHₐ-hemiaminal and IₐSB-Schiff base); b: K = HA/SB.

**Figure 6.** Variation of the 4-amino-3,5-dimethyl-1,2,4-triazole (MeATR ●), hemiaminal (HA ■) and Schiff base (SB ▲) concentration as a function of the initial 2 and 4-nitrobenzaldehyde concentration in the reaction of MeATR with benzaldehydes in CH₃CN solution at 50 °C after 9 h. Product concentrations were obtained from the CH₃ ¹H-NMR signals in the region of 2.00–2.50 ppm.
In Table 5, the values of molar ratio $K$ calculated for the formation of hemiaminal 2 and respective Schiff base (from the amine 1 and 2-nitrobenzaldehyde) in acetonitrile at different temperatures are presented. The results show that the temperature increase favors the imine formation. However, it must be noticed that the summary yield of products (HA + SB) at all temperatures is about 30%. This probably indicates that the first step of the reaction—the hemiaminal formation—is a reversible and exothermic process [43]. The second step of Schiff base formation is endothermic.

**Table 5.** The molar ratio of hemiaminal to imine ($K$) values calculated at different temperatures for the reaction of 1 with 2-nitrobenzaldehyde in CH$_3$CN.

| Temperature °C | 40  | 50  | 60  | 70  | 80  |
|----------------|-----|-----|-----|-----|-----|
| $K = I_{HA}/I_{SB}$ | 8.3 | 6.8 | 5.3 | 3.9 | 2.4 |

$^a$: $I$ is integrated peak intensities of the CH$_3$ signals in $^1$H-NMR spectrum.

2.7. Hemiaminal-Aldehyde Interchange Reaction

Finally, the aromatic aldehyde interchange reaction in DMSO solution at room temperature was studied by the $^1$H-NMR. The spectra in Figure 7 show that upon addition of 2-nitrobenzaldehyde and 4-nitro substituted hemiaminal 4 (12.5 mM) in molar ratio 2:1, respectively, in DMSO-d$_6$ at 25 °C, a new signal appears in the hemiaminal proton region. The above experiments also show that the metathesis reaction is occurring quite slowly and that the first step of this process is the hemiaminal disintegration to amine and aldehyde (Scheme 3).
3. Experimental Section

3.1. Materials and Physical Measurements

The reagents and solvents employed were commercially available and used as received without further purification. Elemental analyses were carried out with a CHNS Vario EL III analyzer (Elementar Analysensystem GmbH, Hanau, Germany). The NMR spectra were recorded on a Bruker 300 or 500 MHz spectrometer (Bruker, Poznań, Poland) using solvent as an internal standard. The mass spectra of electrospray ionization (ESI)-MS were obtained on MicrOTOF-Q mass spectrometer (Bruker). The Fourier transform IR spectra were recorded from KBr pellets in the range of 400–4000 cm\(^{-1}\) on a Bruker IFS 66 FT-IR (Bruker). The Fourier-Transform Raman Nicolet Magna 860 FTIR/FT Raman spectrometer (Spectro-Lab, Warszawa, Poland) was used for the Raman spectral measurements at room temperature.

4-amino-3,5-dimethyl-1,2,4-triazole (MeATR) was synthesized in accordance with the published procedure and checked with \(^1\)H-NMR spectra and elemental analysis [32].

3.2. X-ray Crystallography

Single crystal X-Ray diffraction data were collected at Xcalibur four-circle diffractometer (Wroclaw, Poland) with graphite monochromated Mo K\(\alpha\) radiation (\(\lambda = 0.71073\) Å) at 298 K (2, 5, 15) and 100 K (9, 10) using an Oxford Cryosystem adapter [44] and CC. Data collection and data reduction CrysAlisPro, Agilent Technologies [45] program used. The structures were solved by direct methods with SHELXS and was refined by a full-matrix least squares method using SHELXL97 programs [46]. CCDC 1412796-1412800 contains the supplementary crystallographic data for this paper. These data
can be obtained free of charge via http://www.ccdc.cam.ac.uk/conts/retrieving.html (or from the CCDC, 12 Union Road, Cambridge CB2 1EZ, UK; Fax: +44-1223-336033; E-Mail: deposit@ccdc.cam.ac.uk).

3.3. Synthesis of Hemiaminals 2–14

Compounds 2–12 were synthesized according to the following general procedure. A mixture of equimolar amounts (0.54 mmol) of MeATR (1) and a suitable aldehyde ArCHO (in molar ratio 1:1) were dissolved in acetonitrile (5 mL) and refluxed for 3 h. After removing volatile components, the raw solid products was washed with cold acetonitrile and dried in air. Crystals of four hemiaminals were obtained upon slow evaporation of the solvent from the reaction mixtures.

\[(\text{4H-3,5-Dimethyl-1,2,4-triazole-4-ylamino})(2\text{-nitrophenyl})\text{methanol (2): Yield 96%}. \]
\text{Anal. Calc. (%) for C}_{12}\text{H}_{13}\text{N}_5\text{O}_3: C, 50.18; H, 4.98; N, 26.61. Found: C, 49.85; H, 4.28; N, 25.61. IR (KBr, cm}^{-1}: 373 \text{ w}, 393 \text{ w}, 874 \text{ m}, 959 \text{ w}, 978 \text{ w}, 1025 \text{ m}, 1040 \text{ s}, 1061 \text{ vs}, 1144 \text{ w}, 1164 \text{ m}, 1193 \text{ s}, 1248 \text{ m}, 1314 \text{ m}, 1361 \text{ vs}, 1417 \text{ vs}, 1445 \text{ s}, 1474 \text{ m}, 1498 \text{ s}, 1531 \text{ vs}, 1566 \text{ s}, 1612 \text{ m}, 2882 \text{ s}, 3114 \text{ s}, 3308 \text{ vs}. Raman (cm}^{-1}: 210 \text{ vw}, 275 \text{ vw}, 307 \text{ vw}, 336 \text{ vw}, 398 \text{ w}, 415 \text{ w}, 590 \text{ w}, 622 \text{ vv}, 625 \text{ w}, 666 \text{ vv}, 705 \text{ w}, 765 \text{ vv}, 857 \text{ s}, 892 \text{ w}, 1040 \text{ s}, 1060 \text{ vv}, 1100 \text{ vv}, 1150 \text{ w}, 1170 \text{ w}, 1190 \text{ w}, 1360 \text{ vs}, 1450 \text{ vw}, 1530 \text{ w}, 1580 \text{ w}, 1610 \text{ w}, 2940 \text{ vw}, 2980 \text{ vw}, 3000 \text{ vv}, 3040 \text{ vv}, 3080 \text{ w}, 3090 \text{ m}, 3310 \text{ vv}. MS (ESI, m/z): 264.1 [M + H]+; 286.1 [M + Na]+; 302.1 [M + K]+, 549.2 [2M + Na]+. 1H-NMR (DMSO-d6, 298 K, ppm, 500 MHz): \(\delta = 7.93 \text{ (m, 1H, Ar-H6); 7.91 (m, 1H, Ar-H3); 7.79 (d, J = 7.55 Hz, 1H, Ar-H3); 7.65 (d, J = 7.74 Hz, 1H, Ar-H4); 7.15 (d, 1H, J_(C-H)_(N-H) = 8.12 Hz, N-H); 6.91 (d, 1H, J_(C-H)_(O-H) = 5.33 Hz, O-H); 5.96 (dd, 1H, J_(C-H)_(N-H) = 8.12 Hz, J_(C-H)_(O-H) = 5.33 Hz, C-H); 2.27 (s, 6H, Tr-CH3.). 13C-NMR (DMSO-d6, 298 K, ppm, 75 MHz): \(\delta = 151.6 \text{ (Tr-C), 148.9 (Ar-C2), 133.6 (Ar-C1), 133.4 (Ar-C4), 130.2 (Ar-C4), 129.1 (Ar-C6), 124.2 (Ar-C3), 79.2 (C-OH), 10.5 \text{ (Tr-CH3).}}

Crystal data (C_{12}H_{13}N_{5}O_{3}): M = 263.26, crystal system: monoclinic, space group: C2/c, \(a = 18.695(5) \AA, b = 10.752(3) \AA, c = 15.422(4) \AA, \beta = 125.09(3)^{\circ}, V = 2536.59(12) \AA^3, Z = 8, \mu = 0.104 \text{ mm, } \theta_{\text{max}} = 29.43^{\circ}, \text{reflections: 5365, independent: 2885, } R_{\text{int}} = 0.0155, R1 = 0.0458, wR2 = 0.1238, \text{GoF} = 1.026.

\[(\text{4H-3,5-Dimethyl-1,2,4-triazole-4-ylamino})(3\text{-nitrophenyl})\text{methanol (3): Yield 48%}. \]
\text{Anal. Calc. (%) for C}_{11}\text{H}_{13}\text{N}_5\text{O}_3: C, 50.18; H, 4.98; N, 26.61. Found: C, 49.85; H, 4.28; N, 25.61. IR (KBr, cm}^{-1}: 503 \text{ w}, 583 \text{ w}, 608 \text{ w}, 629 \text{ w}, 679 \text{ s}, 693 \text{ w}, 730 \text{ m}, 762 \text{ w}, 803 \text{ m}, 866 \text{ w}, 907 \text{ w}, 917 \text{ w}; 944 \text{ w}; 1003 \text{ w}; 1015 \text{ s}; 1093 \text{ m}; 1202 \text{ m}; 1248 \text{ w}; 1353 \text{ vs}; 1379 \text{ m}; 1421 \text{ s}; 1527 \text{ vs}, 1563 \text{ m}, 1585 \text{ w}, 1617 \text{ w}, 1648 \text{ w}, 2931 \text{ m}, 3115 \text{ s}, 3253 \text{ s}, 3312 \text{ s}. Raman (cm}^{-1}: 186 \text{ vv}, 234 \text{ vv}, 348 \text{ vv}, 420 \text{ vv}, 608 \text{ vv}, 632 \text{ vv}, 679 \text{ vv}, 681 \text{ vv}, 726 \text{ vv}, 764 \text{ vv}, 861 \text{ vv}, 1000 \text{ m}, 1090 \text{ vv}, 1160 \text{ w}, 1200 \text{ w}, 1340 \text{ s}, 1350 \text{ vs}, 1440 \text{ vv}. 1540 \text{ vw}, 2940 \text{ vw}, 3090 \text{ vv}, 3320 \text{ vv, MS (ESI, m/z): 264.1 [M + H]+; 286.1 [M + Na]+; 302.1 [M + K]+, 549.2 [2M + Na]+. 1H-NMR (DMSO-d6, 298 K, ppm, 500 MHz): \(\delta = 8.43 \text{ (s, 1H, Ar-H2); 8.26 (dd, 1H, J_{4,5} = 8.09 Hz, J_{4,6} = 1.62 Hz, Ar-H4); 8.03 (d, 1H, J_{5,6} = 7.98 Hz, Ar-H6); 7.74 (t, 1H, J = 7.98 Hz, H3); 7.19 (d, 1H, J_{C-H}+(N-H) = 7.17 Hz, N-H); 6.79 (d, 1H, J_{C-H}+(O-H) = 5.78 Hz, O-H); 557 (t, 1H, J_{C-H}+(O-H)+(N-H) = 6.47 Hz, C-H); 2.31 (s, 6H, Tr-CH3.). 13C-NMR (DMSO-d6, 298 K, ppm, 75 MHz): \(\delta = 151.5 \text{ (Tr-C), 148.1 (Ar-C3), 143.0 (Ar-C1), 134.2 (Ar-C6), 130.2 (Ar-C3), 123.6 (Ar-C4), 121.8 (Ar-C2), 83.2 (C-OH), 10.7 \text{ (Tr-CH3).}}

(4H-3,5-Dimethyl-1,2,4-triazole-4-ylamino)(4-nitrophenyl)methanol (4): Yield 48%. Anal. Calc. (%) for C11H13N6O5: C, 43.96; H, 4.06; N, 23.53. Found: C, 43.96; H, 4.05; N, 23.52. IR (KBr, cm⁻¹): 3482 m, 3432 m, 3337 m, 3233 s, 3139 s, 2962 m, 2896 m, 2864 m, 2759 s, 2729 m, 2635 m, 2032 w, 1998 w, 1965 w, 1923 w, 1694 m, 1658 w, 1619 w, 1590 s, 1542 w, 1508 m, 1499 m, 1467 w, 1387 m, 1304 m, 1259 s, 1240 s, 1207 w, 1160 w, 1118 s, 1092 w, 1046 w, 1028 w, 1007 m, 974 w, 931 w, 841 w, 791 w, 744 w, 708 w, 671 w, 635 w, 597 w, 574 w, 531 w, 500 w, 439 w. 1H-NMR (DMSO-d6, 298 K, ppm, 500 MHz): δ = 8.23 (d, 1H, J2-6 = 1.72 Hz, Ar-H2), 7.89 (d, 1H, J6-2 = 1.72 Hz, Ar-H6), 7.37 (d, 1H, J(C-H)(N-H) = 5.67 Hz, O-H), 6.97 (t, 1H, J(C-H)(O-H)(N-H) = 6.42 Hz, C-H), 6.47 (s, 6H, Tr-CH3). 13C-NMR (DMSO-d6, 298 K, ppm, 75 MHz): δ = 151.5 (Tr-C), 147.9 (Ar-C4), 131.1 (Ar-C1), 128.7 (Ar-C2,6), 123.7 (Ar-C3,5), 83.4 (C-OH), 10.7 (Tr-CH3).

(2,4-Dinitrophenyl)(4H-3,5-dimethyl-1,2,4-triazole-4-ylamino)methanol (5) Yield 85%. Anal. Calc. (%) for C11H12N6O5: C, 42.86; H, 3.92; N, 27.26. Found: C, 42.76; H, 3.43; N, 27.00. IR (KBr, cm⁻¹): 3024 w, 2999 w, 2940 w, 2872 m, 2787 m, 2741 m, 2635 m, 2580 m, 2148 w, 1739 w, 1600 m, 1547 w, 1499 w, 1355 m, 1293 s, 1248 w, 1199 w, 1122 w, 1098 w, 1058 m, 1024 w, 942 w, 850 w, 736 w, 696 w, 650 w, 607 w, 563 w, 500 w, 471 w. 1H-NMR (DMSO-d6, 298 K, ppm, 500 MHz): δ = 8.76 (d, 1H, J2-6 = 2.30 Hz, Ar-H2), 8.62 (dd, 1H, J6-2 = 8.77 Hz, J3-5 = 2.30 Hz, Ar-H3,5); 8.17 (d, 1H, J6-2 = 8.77 Hz, Ar-H6), 7.33 (d, 1H, J(C-H)(N-H) = 3.85 Hz, N-H), 7.26 (d, 1H, J(C-H)(O-H) = 5.01 Hz, O-H), 6.00 (dd, 1H, J(C-H)(O-H) = 8.35 Hz, J(C-H)(O-H) = 5.01 Hz, C-H), 2.23 (s, 6H, Tr-CH3). 13C-NMR (DMSO-d6, 298 K, ppm, 75 MHz): δ = 151.5 (Tr-C), 148.7 (Ar-C2), 147.9 (Ar-C4), 139.7 (Ar-C1), 131.0 (Ar-C3), 127.6 (Ar-C5), 119.8 (Ar-C3), 79.1 (C-OH), 10.7 (Tr-CH3).

Crystal data: (C11H12N6O5·H2O) M = 326.28, crystal system: triclinic, space group: P1̅1, a = 8.054(3) Å, b = 8.078(3) Å, c = 12.459(3) Å, α = 87.60(3)°, β = 81.60(3)°, γ = 76.84(3)°, V = 1191.4(4) Å³, Z = 2, ρc = 1.463 g·cm⁻³, μ = 0.122 mm, θmax = 28.77°, reflections: 12648, independent: 3517, Rint = 0.0393, R1 = 0.0595, wR2 = 0.1553, GoF = 1.035.

(4-Chloro-3-nitrophenyl)(4H-3,5-dimethyl-1,2,4-triazole-4-ylamino)methanol (6): Yield 46%. Anal. Calc. (%) for C11H12ClN6O5: C, 44.38; H, 4.06; N, 23.53. Found: C, 44.35; H, 3.56; N, 23.31. IR (KBr, cm⁻¹): 3482 m, 3432 m, 3337 m, 3233 s, 3139 s, 2962 m, 2896 m, 2864 m, 2759 s, 2729 m, 2635 m, 2032 w, 1998 w, 1965 w, 1923 w, 1694 m, 1658 w, 1619 w, 1590 s, 1542 w, 1508 m, 1499 m, 1467 w, 1387 m, 1304 m, 1259 s, 1240 s, 1207 w, 1160 w, 1118 s, 1092 w, 1046 w, 1028 w, 974 w, 931 w, 841 w, 791 w, 744 w, 708 w, 671 w, 635 w, 597 w, 574 w, 531 w, 500 w, 439 w. 1H-NMR (DMSO-d6, 298 K, ppm, 500 MHz): δ = 8.23 (d, 1H, J2-6 = 1.72 Hz, Ar-H2), 7.89 (d, 1H, J6-2 = 1.72 Hz, Ar-H6), 7.87 (m, 1H, Ar-H3), 7.20 (d, 1H, J(C-H)(N-H) = 7.44 Hz, N-H), 6.86 (d, 1H, J(C-H)(O-H) = 5.72 Hz, O-H), 5.57 (t, 1H, J(C-H)(O-H)(N-H) = 6.58 Hz, C-H), 2.36 (s, 6H, Tr-CH3). 13C-NMR (DMSO-d6, 298 K, ppm, 75 MHz):
\[ \delta = 151.5 \text{ (Tr-C)}, 147.8 \text{ (Ar-C₃)}, 141.8 \text{ (Ar-C₁)}, 132.8 \text{ (Ar-C₆)}, 131.9 \text{ (Ar-C₅)}, 125.0 \text{ (Ar-C₄)}, 124.2 \text{ (Ar-C₂)}, 82.7 \text{ (C-OH)}, 10.7 \text{ (Tr-CH₃)}. \]

**4H-3,5-Dimethyl-1,2,4-triazole-4-ylamino)(4-methyl-3-nitrophenyl)methanol (7):** Yield 53%. Anal. Calc. (%) for C₂₂H₁₈N₇O₅: C 51.98; H, 4.55; N, 25.26. Found: C, 51.99; H, 5.05; N, 24.96. IR (KBr, cm⁻¹): 504 w, 540 vw, 569 m, 615 w, 666 w, 682 m, 731 m, 754 w, 762 w, 777 m, 810 m, 862 m, 912 w, 952 w, 1024 w, 1065 s, 1080 s, 1165 vw, 1198 m, 1246 w, 1331 vs, 1346 vs, 1382 m, 1419 m, 1454 m, 1500 s, 1524 vs, 1564 m, 1572 s, 1576 m, 1623 w, 2723 m, 2873 m, 3105 s, 3309 vs. Raman (cm⁻¹): 190 vw, 282 vw, 340 w, 394 vw, 427 vw, 572 vw, 616 w, 668 vw, 682 vw, 731 vw, 810 w, 862 vw, 952 vw, 1010 vw, 1070 vw, 1200 m, 1330 w, 1390 w, 1450 vw, 1530 w, 1540 vw, 1570 w, 1630 vw, 2940 w, 2990 w, 3060 vw, 3070 vw, 3100 vw. MS (ESI, m/z): 278.1 [M + H]⁺, 300.1 [M + Na]⁺, 316.1 [M + K]⁺. ¹H-NMR (DMSO-d₆, 298 K, ppm, 500 MHz): δ = 8.16 (d, 1H, J₂-6 = 1.18 Hz, Ar-H₂); 7.76 (dd, 1H, J₆-5 = 7.82 Hz, J₆-2 = 1.42 Hz, Ar-H₆); 7.58 (d, 1H, J₆-2 = 1.42 Hz, Ar-H₃); 7.12 (d, 1H, J₆-2 = 1.42 Hz, Ar-H₃); 6.70 (d, 1H, J₆-2 = 1.42 Hz, Ar-H₃); 5.54 (t, 1H, J₆-2 = 1.42 Hz, C-H); 2.56 (s, 3H, Ar-CH₃); 2.35 (s, 6H, Tr-CH₃)). ¹³C-NMR (DMSO-d₆, 298 K, ppm, 75 MHz): δ = 151.0 (Tr-C), 148.6 (Ar-C₂), 139.9 (Ar-C₁,a), 132.6 (Ar-C₃), 131.7 (Ar-C₆), 122.4 (Ar-C₂), 82.6 (C-OH), 19.3 (Ar-CH₃), 10.2 (Tr-CH₃).

**2-Chloro-5-nitrophenyl)(4H-3,5-dimethyl-1,2,4-triazole-4-ylamino)methanol (8):** Yield 74%. Anal. Calc. (%) for C₁₁H₁₂ClN₅O₃: C 44.38; H, 4.06; N, 23.53; Cl, 11.91. Found: C, 44.39; H, 3.65; N, 23.21; Cl, 12.15. IR (KBr, cm⁻¹): 465 w, 502 w, 511 w, 528 m, 570 w, 585 m, 612 m, 630 w, 662 w, 692 w, 745 s, 769 w, 800 m, 842 m, 858 w, 913 m, 950 w, 985 w, 1027 m, 1041 s, 1068 m, 1103 m, 1196 m, 1247 m, 1278 m 1315 vs, 1377 m, 1419 m, 1440 w, 1471 w, 1507 m, 1530 vs, 1567 w, 1598 w, 1620 w, 1703 w, 2942 w, 3081 w, 3309 w. Raman (cm⁻¹): 240 vw, 278 vw, 318 vw, 340 vw, 427 vw, 502 vw, 597 vw, 631 vw, 666 vw, 696 vw, 725 vw, 768 vw, 801 vw, 819 w, 858 w, 949 w, 985 w, 1024 vw, 1067 w, 1103 v, 1196 v, 1348 vs, 1458 w, 1518 w, 1544 v, 1647 v, 1693 w, 2942 w, 3081 w, 3309 w. MS (ESI, m/z): 298.1 [M + H]⁺, 320.0 [M + Na]⁺, 336.0 [M + K]⁺. ¹H-NMR (DMSO-d₆, 298 K, ppm, 500 MHz): δ = 8.50 (d, 1H, J₆-4 = 2.83 Hz, Ar-H₆); 8.27 (dd, 1H, J₆-4 = 2.83 Hz, Ar-H₆); 7.84 (d, 1H, J₆-4 = 2.83 Hz, Ar-H₆); 7.26 (d, 1H, J₆-4 = 2.83 Hz, Ar-H₆); 6.70 (d, 1H, J₆-4 = 2.83 Hz, Ar-H₆); 5.77 (dd, 1H, J₆-4 = 2.83 Hz, Ar-H₆); 2.36 (s, 6H, Tr-CH₃). ¹³C-NMR (DMSO-d₆, 298 K, ppm, 125 MHz): δ = 151.1 (Tr-C), 148.6 (Ar-C₂), 139.9 (Ar-C₁,a), 132.6 (Ar-C₃), 131.7 (Ar-C₆), 122.4 (Ar-C₂), 82.6 (C-OH), 19.3 (Ar-CH₃), 10.2 (Tr-CH₃).

**2-Chlorophenyl)(4H-3,5-dimethyl-1,2,4-triazole-4-ylamino)methanol (9):** Yield 63%. Anal. Calc. (%) for C₁₁H₁₂ClN₄O: C 52.28; H, 5.19; N, 22.17. Found: C, 52.36; H, 4.94; N, 21.95. IR (KBr, cm⁻¹): 428 w, 464 s, 509 m, 588 s, 613 m, 633 s, 667 m, 705 s, 742 vs, 759 vs, 805 m, 884 s, 955 w, 978 w, 994 m, 1016 vs, 1035 s, 1047 s, 1057 s, 1088 m, 1196 m, 1247 w, 1265 w, 1340 w, 1357 w, 1374 m, 1419 s, 1440 s, 1471 s, 1507 m, 1542 m, 1567 m, 1578 m, 1598 w, 2880 m, 3033 m, 3124 m, 3265 s. Raman (cm⁻¹): 179 vw, 218 w, 259 w, 280 w, 323 w, 330 w, 354 w, 431 vs, 590 w, 615 m, 625 m, 634 w, 670 w, 681 w, 705 w, 739 vw, 764 w, 882 w, 997 m, 1037 vs, 1085 s, 1127 vw, 1159 s, 1196 m, 1212 vw, 1274 w, 1287 w, 1318 vw, 1378 w, 1435 w, 1468 w, 1543 vs, 1576 vs, 1592 m, 1598 m, 1611 w, 1696 w, 2932 m, 2991 v, 3059 m, 3070 m, 3081 m, 3265 vw. MS (ESI, m/z): 253.1 [M + H]⁺, 505.2 [2M + H]⁺.
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1H-NMR (DMSO-d$_6$, 298 K, ppm, 500 MHz): $\delta = 7.67$ (m, 1H, Ar-H$_3$), 7.48 (m, 1H, Ar-H$_6$), 7.41 (m, 2H, Ar-H$_4$,H$_5$), 7.05 (d, 1H, $J_{(C-H)-(N-H)} = 6.29$ Hz, N-H), 6.60 (d, 1H, $J_{(C-H)-(O-H)} = 5.15$ Hz, O-H), 5.76 (t, 1H, $J_{(C-H)-(O-H),(N-H)} = 5.72$ Hz, C-H), 2.30 (s, 6H, Tr-CH$_3$).

13C-NMR (DMSO-d$_6$, 298 K, ppm, 75 MHz): $\delta = 151.8$ (Tr-C), 137.9 (Ar-C$_1$), 132.0 (Ar-C$_2$), 130.6 (Ar-C$_4$), 129.6 (Ar-C$_6$), 129.1 (Ar-C$_3$), 127.7 (Ar-C$_5$), 80.8 (C-OH), 10.7 (Tr-CH$_3$).

Crystal data: M = 252.70, crystal system: monoclinic, space group: $P2_1/c$, $a = 10.882(3)$ Å, $b = 14.734(4)$ Å, $c = 8.240(3)$ Å, $\beta = 104.5(3)^\circ$, $V = 1151.7(6)$ Å$^3$, $Z = 4$, $\rho_c = 1.457$ g·cm$^{-3}$, $\mu = 0.321$ mm, $\theta_{max} = 28.66^\circ$, reflections: 8449, independent: 2788, $R_{int} = 0.0228$, $R1 = 0.1481$, $wR2 = 0.08908$, GoF = 0.974.

(4H-3,5-Dimethyl-1,2,4-triazole-4-ylamino)(4-formylphenyl)methanol (10): Yield 39%. Anal. Calc. (%)

for C$_{12}$H$_{14}$N$_4$O$_2$: C, 58.53; H, 5.73; N, 22.75. Found: C, 58.42; H, 5.41; N, 22.28. IR (KBr, cm$^{-1}$): 490 w, 599 m, 699 m, 706 w, 782 s, 843 m, 893 m, 981 w, 1016 m, 1064 vs, 1113 w, 1166 w, 1210 s, 1252 w, 1273 m, 1304 m, 1335 m, 1384 s, 1419 s, 1509 m, 1541 m, 1568 s, 1608 s, 1700 vs, 2716 s, 2819 s, 2915 m, 3077 s. Raman (cm$^{-1}$): 203 vw, 280 vw, 306 vw, 332 vw, 600 vw, 631 vw, 643 vw, 671 vw, 706 vw, 892 vw, 921 vw, 1066 vw, 1112 vw, 1210 vw, 1380 vw, 1509 vw, 1541 vw, 1568 vw, 1608 s, 1700 vs, 2716 s, 2819 s, 2915 m, 3077 s. MS (ESI, $m/z$): 269.1 [M + H]$^+$. 1H-NMR (DMSO-d$_6$, 298 K, ppm, 500 MHz): $\delta = 8.00$ (d, 2H, $J_{2-3} = 8.20$ Hz, Ar-H$_3$,5), 7.79 (d, 2H, $J_{2-5} = 8.20$ Hz, Ar-H$_2$,6), 7.13 (d, 1H, $J_{(C-H)-(N-H)} = 7.06$ Hz, N-H), 6.60 (d, 1H, $J_{(C-H)-(O-H)} = 5.72$ Hz, O-H), 5.73 (t, 1H, $J_{(C-H)-(O-H),(N-H)} = 6.49$ Hz, C-H), 2.34 (s, 6H, Tr-CH$_3$).

Crystal data: M = 246.27, crystal system: monoclinic, space group: $P2_1/c$, $a = 12.759(4)$ Å, $b = 7.395(3)$ Å, $c = 12.768(6)$ Å, $\beta = 96.91(4)^\circ$, V = 1195.81(8) Å$^3$, $Z = 4$, $\rho_c = 1.368$ g·cm$^{-3}$, $\mu = 0.097$ mm, $\theta_{max} = 24.99^\circ$, reflections: 12521, independent: 2113, $R_{int} = 0.0423$, $R1 = 0.0424$, $wR2 = 0.111$, GoF = 1.000.

(4-Cyanophenyl)(4H-3,5-dimethyl-1,2,4-triazole-4-ylamino)methanol (11): Yield 51%. Anal. Calc. (%)

for C$_{12}$H$_{13}$N$_4$O: C 59.25; H, 5.39; N, 28.79. Found: C, 59.23; H, 5.03; N, 28.40. IR (KBr, cm$^{-1}$): 461 w, 481 w, 503 w, 509 m, 603 s, 664 m, 712 w, 751 s, 768 w, 802 m, 817 m, 894 m, 1019 m, 1064 vs, 1198 w, 1249 w, 1272 w, 1333 w, 1352 w, 1385 m, 1408 m, 1457 m, 1544 w, 1569 m, 1610 w, 2233 s, 2704 w, 2855 w, 2914 w, 3020 s, 3028 s, 3280 s, 3475 v, 4191 v, 558 v, 608 v, 646 v, 718 v, 747 v, 818 v, 880 v, 1001 m, 1186 w, 1223 w, 1290 v, 1319 v, 1364 v, 1411 v, 1508 v, 1538 s, 1581 s, 1610 s, 2227 m, 2928 v, 2988 v, 3051 v, 3084 v, 3281 v, MS (ESI, $m/z$): 244.1 [M + H]$^+$, 266.1 [M + Na]$^+$, 509.2 [2M + Na]$^+$. 1H-NMR (DMSO-d$_6$, 298 K, ppm, 500 MHz): $\delta = 7.89$ (d, 2H, $J_{2-3} = 8.20$ Hz, Ar-H$_3$,5), 7.75 (m, 2H, $J_{2-3} = 8.20$ Hz, Ar-H$_2$,6), 7.13 (d, 1H, $J_{(C-H)-(N-H)} = 7.06$ Hz, N-H), 6.70 (d, 1H, $J_{(C-H)-(O-H)} = 5.72$ Hz, O-H), 5.53 (t, 1H, $J_{(C-H)-(O-H),(N-H)} = 6.49$ Hz, C-H), 2.34 (s, 6H, Tr-CH$_3$).

Crystal data: M = 246.27, crystal system: monoclinic, space group: $P2_1/c$, $a = 12.759(4)$ Å, b = 7.395(3) Å, c = 12.768(6) Å, $\beta = 96.91(4)^\circ$, V = 1195.81(8) Å$^3$, $Z = 4$, $\rho_c = 1.368$ g·cm$^{-3}$, $\mu = 0.097$ mm, $\theta_{max} = 24.99^\circ$, reflections: 12521, independent: 2113, $R_{int} = 0.0423$, $R1 = 0.0424$, $wR2 = 0.111$, GoF = 1.000.
(4H-3,5-Dimethyl-1,2,4-triazole-4-ylamino)(4-(trifluoromethyl)phenyl)methanol (12): Yield 39%. Anal. Calc. (%) for C_{32}H_{33}F_{4}N_{4}O: C 50.35; H, 4.58; N, 19.57. Found: C, 50.3; H, 5.03; N, 18.87. IR (KBr, cm⁻¹): 461 w, 502 w, 581 w, 597 m, 654 w, 668 w, 712 w, 754 m, 770 m, 810 m, 829 m, 896 m, 1016 s, 1053 vs, 1070 vs, 1111 vs, 1152 vs, 1205 m, 1247 w, 1283 w, 1332 vs, 1380 m, 1415 s, 1498 m, 1542 m, 1565 s, 1622 w, 2729 w, 2894 m, 3126 s, 3284 s. Raman (cm⁻¹): 228 s, 275 w, 299 w, 315 m, 400 vw, 415 w, 500 w, 579 w, 604 s, 635 vs, 655 w, 670 s, 679 w, 710 m, 740 m, 753 m, 768 m, 801 w, 894 m, 980 w, 1014 vw, 1050 vw, 1067 w, 1081 vw, 1104 vw, 1185 m, 1204 m, 1284 vw, 1324 vs, 1361 vw, 1384 w, 1463 w, 1540 m, 1592 w, 1621 vs, 2739 vw, 2932 s, 2991 w, 3011 w, 3082 s, 3285 vw. MS (ESI, m/z): 287.1 [M + H]⁺, 309.1 [M + Na]⁺, 325.1 [M + K]⁺. ¹H-NMR (DMSO-d₆, 298 K, ppm, 500 MHz): δ = 7.79 (s, 4H, Ar-H 2,3,5,6), 7.11 (d, 1H, J_{(C-H)-(N-H)} = 6.87 Hz, N-H), 6.66 (d, 1H, J_{(C-H)-(O-H)} = 5.53 Hz, O-H), 5.55 (t, 1H, J_{(C-H)-(O-H), (N-H)} = 6.20 Hz, C-H), 2.34 (s, 6H, Tr-CH₃)). ¹³C-NMR (DMSO-d₆, 298 K, ppm, 125 MHz): δ = 151.6 (Tr-C), 145.3 (Ar-C₁), 129.3 (q, 2J_F-C = 30.88 Hz, Ar-C₄), 128.2 (Ar-C₂,6), 125.5 (q, 3J_F-C = 3.63 Hz, Ar-C₃,5), 124.7 (q, 3J_F-C = 272.4 Hz, Ar-CF₃), 83.7 (C-OH), 10.7 (Tr-CH₃).

Compounds 13 and 14 were synthesized and purified according to the following procedure. A solution of suitable pyridinecarboxaldehyde (0.47 mL, 0.5 mM) in 1 mL of ethanol was added to hot solution of equimolar amounts of MeATR (1) (0.56 g, 0.5 mM) in 10 mL of ethanol. The reaction mixture was refluxed for 2 h, cooled and kept overnight in refrigerator. The solvent was then removed in vacuo and the remaining materials were washed with cold ethanol and dried in air.

(4H-3,5-Dimethyl-1,2,4-triazole-4-ylamino)(pyridin-3-yl)methanol (13): Yield 84%. Anal. Calc. (%) for C₁₀H₁₃N₅O: C, 54.78; H, 5.98; N, 31.94. Found: C, 54.80; H, 6.12; N, 32.18. IR (nujol, cm⁻¹): 410 w, 490 w, 503 w, 582 m, 611 w, 637 m, 671 m, 764 m, 784 m, 831 w, 894 m, 955 m, 980 w, 1030 m, 1051 m, 1065 vs, 1086 m, 1205 w, 1252 w, 1300 w, 1336 m, 1353 w, 1380 w, 1511 m, 1525 m, 1542 m, 1570 s, 1582 m, 1596 m, 2741 m, 3069 s, 3125 s, 3202 s. Raman (cm⁻¹): 225 vw, 269 vw, 294 vw, 311 vw, 338 vw, 357 vw, 409 vw, 487 vw, 597 w, 622 w, 653 v, 677 v, 743 v, 764 v, 799 v, 832 v, 891 v, 980 v, 1028 w, 1041 vs, 1062 vw, 1085 v, 1127 v, 1191 w, 1254 v, 1300 v, 1335 v, 1380 v, 1457 v, 1523 v, 1541 v, 1572 v, 1596 w, 1618 w, 2205 v, 2948 v, 3054 v, 3062 w, 3078 w, 3203 v. MS (ESI, m/z): 220.1 [M + H]⁺, 242.1 [M + Na]⁺, 258.1 [M + K]⁺, 461.2 [2M + Na]. ¹H-NMR (DMSO-d₆, 298 K, ppm, 500 MHz): δ = 8.86 (d, 1H, J_{2-6} = 1.89 Hz, Py-H₂), 8.57 (dd, 1H, J_{4-5} = 4.82 Hz., J_{4-6} = 1.61 Hz Py-H₄), 7.93 (dt, 1H, J_{5-6} = 7.90 Hz, J_{2,4-6} = 1.70 Hz), 7.45 (dd, 1H, J_{5-6} = 7.90 Hz, J_{4-5} = 4.82 Hz, J_{2-5} = 0.66 Hz Py-H₂), 7.13 (d, 1H, J_{(C-H)-(O-H)} = 6.99 Hz, N-H), 6.66 (d, 1H, J_{(C-H)-(O-H)} = 5.67 Hz, O-H), 5.54 (t, 1H, J_{(C-H)-(O-H), (N-H)} = 6.33 Hz, C-H), 2.34 (s, 6H, Tr-CH₃). ¹³C-NMR (DMSO-d₆, 298 K, ppm, 125 MHz): δ = 151.1 (Tr-C), 149.4 (Py-C₄), 148.2 (Py-C₂), 139.2 (Py-C₁), 134.2 (Py-C₆), 123.3 (Py-C₅), 82.2 (C-OH), 10.3 (Tr-CH₃).

(4H-3,5-Dimethyl-1,2,4-triazole-4-ylamino)(pyridin-4-yl)methanol (14): Yield 81%. Anal. Calc. (%) for C₁₀H₁₃N₅O: C, 54.78; H, 5.98; N, 31.94. Found: C, 54.79; H, 6.06; N, 32.51. IR (nujol, cm⁻¹): 406 w, 503 w, 582 m, 611 w, 637 m, 671 m, 764 m, 779 m, 803 s, 848 w, 903 s, 977 w, 997 m, 1031 m, 1061 vs, 1105 s, 1197 m, 1216 w, 1232 w, 1245 w, 1289 w, 1320 w, 1339 m, 1415 vs, 1512 m, 1541 m, 1573 s, 1609 m, 2724 m, 3061 s, 3122 s, 3191 s. Raman (cm⁻¹): 226 vw, 269 vw, 294 vw, 313 v, 336 v, 352 v, 405 v, 500 vw, 523 vw, 591 v, 612 v, 636 v, 670 s, 680 vw, 725 w, 762 w, 805 v, 851 w, 901 v,
1001 vs, 1029 vw, 1058 vw, 1094 vw, 1192 w, 1214 vw, 1228 vw, 1287 vw, 1320 vw, 1335 vw, 1380 vw, 1450 vw, 1540 w, 1562 vw, 1607 w, 1620 w, 2739 vw, 2931 w, 2975 vw, 3059 w, 3108 vw, 3196 vw.

MS (ESI, m/z): 220.1 [M + H]^+ , 242.1 [M + Na]^+, 258.1 [M + K]^+. δH-NMR (DMSO-d6, 298 K, ppm, 500 MHz): δ = 8.62 (d, 1H, J2-3 = 5.85 Hz, Py-H3,5), 7.56 (d, 1H, J = 5.67 Hz, Py-H2,6), 7.15 (d, 1H, J = 7.18 Hz, N-H), 6.69 (d, 1H, J = 5.85 Hz, O-H), 5.47 (t, 1H, J = 6.61 Hz, C-H), 2.34 (s, 6H, Tr-CH3). δ13C-NMR (DMSO-d6, 298 K, ppm, 125 MHz): δ = 151.5 (Tr-C), 150.1 (Py-C3,5), 149.0 (Py-C1), 122.3 (Py-C2,6), 83.2 (C-OH), 10.7 (Tr-CH3).

3.4. Synthesis of Imines 15–16

Schiff bases were prepared according to the following general procedure. A mixture of equimolar amounts (0.5 mmol) of MeATR (1) and an appropriate aldehyde ArCHO (in molar ratio 1:1) were dissolved in acetonitrile (3 mL) with presence of catalytic amounts of hydrochloric acid (2 drops, 36%). The reaction mixture was then refluxed for 3 h. After cooling, the precipitate formed was filtered off, washed with small amount of cold acetonitrile and then dried in the air.

(N-(4-Nitrobenzylidene)-4H-3,5-dimethyl-1,2,4-triazole-4-amine) hydrochloride (15): Yield 97%. Anal. Calc. (%) for C11H12N5O2Cl: C, 46.90; H, 4.29; N, 24.86. Found: C, 46.99; H, 4.12; N, 24.52. IR (KBr, cm−1): 438 w; 503 w; 537 w; 599 w; 649 w; 662 w; 671 w; 692 w; 757 m; 776 m; 846 vs; 873 vs; 884 s; 931 s; 980 s; 1006 vs; 1017 s; 1045 s; 1106 s; 1201 m; 1234 s; 1317 vs; 1346 s; 1376 s; 1404 s; 1475 m; 1525 vs; 1567 m; 1589 s; 1618 w; 1827 m; 2367 s; 2926 w; 2982 w; 3062 w. MS (ESI, m/z, M = C11H11N5O2): 246.1 [M + H]^+, 268.1 [M + Na]^+, 284.1 [M + K]^+. δ13C-NMR (DMSO-d6, 298 K, ppm, 75 MHz): δ = 166.9 (CH=N), 150.5 (Ar-C4), 149.1 (Tr C), 137.6 (Ar-C1), 131.0 (Ar-C2,6), 124.7 (Ar-C3,5), 10.8 (Tr-CH3).

(N-(4-chloro-3-nitrobenzylidene)-4H-3,5-dimethyl-1,2,4-triazole-4-amine) hydrochloride (16): Yield 98%. Anal. Calc. (%) for C11H11N5O2Cl: C, 46.90; H, 4.29; N, 24.86. Found: C, 46.99; H, 4.12; N, 24.52. IR (KBr, cm−1): 422 w; 438 m; 503 m; 537 m; 545 w; 599 w; 649 w; 662 w; 671 w; 692 s; 757 s; 776 m; 846 vs; 873 vs; 884 s; 931 s; 980 s; 1006 vs; 1017 s; 1045 s; 1106 s; 1201 m; 1234 s; 1317 vs; 1346 vs; 1376 s; 1404 s; 1475 m; 1525 vs; 1567 m; 1589 s; 1618 w; 1827 m; 2367 s; 2926 w; 2982 w; 3062 w. MS (ESI, m/z, M = C11H10N5O2Cl): 280.1 [M + H]^+, 302.0 [M + Na]^+, 439.6 [3M·HCl + 4H]^2+. δ13C-NMR (DMSO-d6, 298 K, ppm, 75 MHz): δ = 166.9 (CH=N), 150.5 (Ar-C4), 149.1 (Tr C), 137.6 (Ar-C1), 131.0 (Ar-C2,6), 124.7 (Ar-C3,5), 10.8 (Tr-CH3).
(Tr C), 148.3 (Ar-C3), 134.1 (Ar-C6), 133.3 (Ar-C5), 132.3 (Ar-C4), 130.1 (Ar-C1), 126.7 (Ar-C2), 10.8 (Tr-CH3).

3.5. Reaction of MeATR (1) with 2-Pyridinecarboxaldehyde

A mixture of equimolar amounts (0.17 mmol) of MeATR (1) and 2-pyridinecarboxaldehyde (in molar ratio 1:1) were dissolved in hexane (2 mL) and stirred at 50 °C for 9 h. After removing volatile components, raw solid products washed with cold hexane dried and dissolved in DMSO-d6 and analyzed by NMR spectroscopy.

\((4H-3,5-Dimethyl-1,2,4-triazole-4-ylamino)(pyridin-2-yl)methanol \) (17) Yield 12%. \(^1\)H-NMR (DMSO-d6, 298 K, ppm, 500 MHz): \(\delta = 8.62 \) (m, 1H, Py-H3), 7.89 (td, 1H, J\(_{4,5,5,6} = 7.68\) Hz, J\(_{3,5} = 1.81\) Hz Py-H5), 7.59 (d, 1H, J\(_{5,6} = 7.90\) Hz, Py-H6), 7.41 (ddd, 1H, J\(_{4,5} = 7.44\) Hz, J\(_{3,4} = 4.77\) Hz, J\(_{4,6} = 1.14\) Hz, Py-H4), 7.14 (d, 1H, J\(_{C-H}-(N-H) = 4.96\) Hz, N-H), 6.51 (d, 1H, J\(_{C-H}-(O-H) = 6.29\) Hz, O-H), 5.44 (dd, 1H, J\(_{C-H}-(N-H) = 4.96\) Hz, J\(_{C-H}-(O-H) = 6.29\) Hz, C-H), 2.30 (s, 6H, Tr-CH3).

\((N-(Pyridin-2-yl)methylene)-4H-3,5-dimethyl-1,2,4-triazole-4-amine) \) (17a) Yield 23%. \(^1\)H-NMR (DMSO-d6, 298 K, ppm, 500 MHz): \(\delta = 8.81\) (s, 1H), 8.77 (m, 1H, Py-H3), 8.17 (dt, 1H, J\(_{5,6} = 7.87\) Hz, J\(_{3,4} = 1.03\) Hz, Py-H6) 8.03 (m, 1H, Py-H5), 7.62 (ddd, 1H, J\(_{4,5} = 7.49\) Hz, J\(_{3,4} = 4.82\) Hz, J\(_{4,6} = 1.24\) Hz, Py-H4), 2.47 (s, 6H, Tr-CH3).

3.6. Reaction Survey

The effect of substituents on the condensation reaction—aldehyde (0.172 mmol) and amine 1 (0.172 mmol) were dissolved in acetonitrile (2 mL) and stirred for 9 h at 50 °C. After removing volatile components, the solid products were dissolved in DMSO-d6 and \(^1\)H-NMR spectra were measured. The amount of hemiaminal, Schiff base and unreacted amine were determined from integrated peak intensities.

The effect of solvent on the condensation reaction—2-nitrobenzaldehyde (0.172 mmol), amine 1 (0.172 mmol) and 2 mL of solvent were stirred for 9 h at 50 °C. After removing volatile components, the solid products were dissolved in DMSO-d6 and \(^1\)H-NMR spectra were measured. The amount of hemiaminal (HA), Schiff base (SB) and unreacted amine (A) were determined from integrated peak intensities of the \(\delta(C-CH_3)\) signals (A—2.25 ppm, HA—2.27 ppm, SB—2.47 ppm).

4. Conclusions

In this paper, a new group of hemiaminals derived from aromatic aldehydes (benzyl, pyridyl) and 4-amine-3,5-dimethyl-1,2,4-triazole was presented. We found that most of the electron-withdrawing substituents in the aromatic aldehydes can stabilize the creation of stable hemiaminals e.g., compounds 9, 10, 12, 13 and 14 presented in this paper. The presence of two methyl substituents in the triazole ring significantly affects the crystal and molecular structure of hemiaminals, which form centrosymmetric dimers only, while predominantly polymeric structures have been reported previously. The presence of the methyl groups also affects the conformation of molecules which, in solution and in crystalline form, have the stretched geometry. This means that our hemiaminals in solution have the RS/SR configuration. The current study revealed the enormous influence of the environment on the reaction course and its
efficiency. In this respect, the solvent polarity, the presence of water and its catalytic performance are important factors. A simple relationship between temperature and the product yield as well as the metathesis phenomena observed in this work led to the conclusion that the first stage of condensation—the creation of a hemiaminal—is an exothermic process, while the second—a Schiff base formation—is an endothermic process.

Acknowledgments

The research was supported by Wroclaw Research Centre EIT+ under the project “Biotechnologies and advanced medical technologies—BioMed” (POIG 01.01.02-02-003/08-00) was financed from the European Regional Development Fund (Operational Program Innovative Economy, 1.1.2).

Author Contributions

The project was devised by K.W.-H. and Z.C. Experimental results were obtained by K.W.-H., D.P., A.Z., R.W. and K.D. The manuscript was prepared by K.W.-H. with consultation from Z.C. and K.D.

Conflicts of Interest

The authors declare no conflict of interest.

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*Sample Availability: Not available.*

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