2D TiVCT\(_x\) layered nanosheets grown on nickel foam as highly efficient electrocatalysts for the hydrogen evolution reaction

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Exploring highly efficient and durable catalysts for the hydrogen evolution reaction (HER) is crucial for the hydrogen economy and environmental protection issues. Numerous studies have now found that transition metal carbide MXenes are ideal candidates as catalysts for the hydrogen evolution reaction. However, MXenes are inclined to easily undergo lamellar structure agglomeration and stacking, which impedes their further applications. Besides, most of the extant research has focused on single transition metal carbides, and the investigation of double transition metal carbide MXenes is rather rare. In this research work, a three-dimensional (3D) TiVCT\(_x\)-based conductive electrode was constructed by depositing 2D TiVCT\(_x\) nanosheets on 3D network structured nickel foam (NF) to synthesize a hybrid electrode material (abbreviated as TiVCT\(_x\)@NF). TiVCT\(_x\)@NF exhibits efficient electrochemical properties with a low overpotential of 151 mV at 10 mA cm\(^{-2}\) and a small Tafel slope of 116 mV dec\(^{-1}\). Benefitting from the open layer structure and strong interfacial coupling effect, compared to the pristine structure, the resulting TiVCT\(_x\)@NF has greatly increased active sites for the hydrogen evolution reaction (HER) and encounters less resistance for charge transfer. In addition, TiVCT\(_x\)@NF exhibits better stability in long-term acidic electrolytes. This work provides a tactic to prepare three-dimensional network electrode materials and broadens the application of single transition metal carbide MXenes as water splitting electrodes in the HER, which is beneficial to the application of noble metal-free electrocatalysts.

1. Introduction

In recent years, due to the combustion and emission of a variety of chemical fuels, energy shortage and environmental pollution have become global problems. The development of various new and renewable energy sources is an important issue at present. As a secondary energy source, hydrogen energy is clean, efficient, and can be easily stored and transported, and is thus regarded as an ideal energy carrier.\(^1\)\(^2\) Water splitting is one of the most promising technologies for industrial hydrogen production.\(^3\) The best electrocatalysts for the HER are platinum electrodes or noble metals; however, due to the rising price of precious metals, the improvement of the catalytic activity and cost-effectiveness of existing metal hydrogen evolution electrodes has become a primary issue for manufacturers.\(^4\)\(^5\)

MXenes are two-dimensional transition metal compounds, which can be expressed as M\(_{n+1}\)X\(_n\)T\(_x\) (where M represents a transition metal element, X stands for C or N, and T represents –OH, –F and other functional groups). Many studies\(^6\)\(^9\) have shown that MXenes are excellent materials for water electrolysis due to their high specific surface area, good electrical conductivity and stability, which has aroused the extensive interest of scholars. Tang \etal.\(^10\) proposed that due to the features of the MXene Ti\(_3\)C\(_2\) such as its large surface area, excellent conductivity and adjustable terminations, it can be applied in the hydrogen evolution reaction (HER). Huang \etal.\(^11\) prepared MoS\(_2\)/Ti\(_3\)C\(_2\) composite heterostructures by the hydrothermal method. When used as a HER electrocatalyst, MoS\(_2\)/Ti\(_3\)C\(_2\) exhibits excellent electrochemical activity with a low overpotential of \(\sim 280\) mV at \(10\) mA cm\(^{-2}\). Compared to MoS\(_2\) nanosheets, the catalytic current density induced by MoS\(_2\)/Ti\(_3\)C\(_2\) at an overpotential of \(\sim 400\) mV is more than 6.2 times that of the MoS\(_2\) nanosheets. Le \etal.\(^12\) used ammonia heat treatment to enhance the hydrogen evolution catalysis of Ti\(_3\)C\(_2\) MXenes by modification with a nitrogen heteroatom. As shown in the experimental results, compared to the Ti\(_3\)C\(_2\) MXene, the nitrogen-doped Ti\(_3\)C\(_2\) annealed at \(600\) °C exhibited superior HER electrocatalytic performance with an overpotential as low as \(198\) mV at \(10\) mA cm\(^{-2}\) and a much smaller Tafel slope of \(92\) mV dec\(^{-1}\). Liu \etal.\(^13\) mixed H\(_8\)MoN\(_2\)S\(_4\) and Ti\(_3\)C\(_2\)T\(_x\) to prepare MoS\(_2\)/Ti\(_3\)C\(_2\)T\(_x\) layered nanomaterials. The active sites of the two-
dimensional nanomaterials are greatly increased, the exchange current density is more than 25 times that of MoS2, and the electrochemical performance is significantly improved.

With the deepening of research, scholars have tried to introduce transition metal elements into MXenes to obtain different electrochemical properties. According to a previous study, compared to Ti3C2, the TiVCT MXene has some unique electrochemical properties due to the addition of the transition metal element V. Yazdanparast et al. successfully obtained thin films of TiVCT by etching and layering. The obtained TiVCT nanosheets possess a large specific surface area and multiple active sites, and the functional group structure on the TiVCT MXene has excellent stability. Li et al. investigated the electrochemical properties of the TiVCT MXene, and the results showed that the TiVCT MXene exhibited metallic properties with high electronic conductivity, and its comprehensive electrochemical performance is better than that of the Ti3C MXene. Li et al. prepared TiVCT MXenes and investigated their electrochemical properties, and discovered that the performance of the TiVCT MXenes did not yield the desired results. They then used 6-phenylenediamine (o-PD) as a metamer for oxidant-free polymerization and mixed it with the prepared TiVCT to obtain three-dimensionally stable TiVCT/poly-o-PD hybrids. Electrochemical studies have shown that the TiVCT/poly-o-PD hybrid possessed a high specific surface area, excellent cycling stability, and improved comprehensive electrochemical performance. Nevertheless, there are few reports on the hydrogen evolution performance of the TiVCT MXene in the electrolysis of water.

Ni is an abundant transition metal element with the features of low cost and high catalytic activity in hydrogen production by water electrolysis. Previous studies have shown that the hybridization of nickel foam (NF) is beneficial to the improvement of the electrochemical performance of nanomaterials. Herein, a new strategy is proposed, and the novel hybrid TiVCT@nickel foam (NF) is obtained by compounding nickel foam with layered MXene nanosheets. The surface element distribution, morphology and structure of TiVCT@NF are systematically studied, and the electrochemical hydrogen evolution performance and stability of the hybrid nanomaterials are further investigated.

2. Experimental

2.1 Preparation of TiVAIC powders

In a preliminary work, we have successfully prepared TiVAIC powders using TiH2 (99.6% pure), V (99.6% pure), Al (99.5% pure) and C (99.8% pure) powders with an atomic ratio of 1 : 1 : 1.3 : 1 by the activation reaction sintering method at a sintering temperature of 1500 °C, and the specific details about the preparation of the MAX TiVAIC are shown in Zou’s study. Then, the fabricated TiVAIC powders were broken by a crushing machine (GJ500-1, Hangzhou, China) and sieved through a mesh 325 sieve.

2.2 Preparation of TiVCT nanosheets

Layered TiVCT was synthesized by liquid etching of the as-prepared TiVAIC powders. The etching was carried out by immersing the TiVAIC powders in a 40% HF solution with a solid-to-liquid ratio of 1 : 20 and stirred at 500 rpm for 48 h at 45 °C. Then, the suspension was centrifuged 7–8 times (3 min at a time) at 3500 rpm to separate the multilayered MXene. The centrifuged suspension supernatant was washed and filtered with distilled water until the pH of the supernatant was equal to ~7. After that, the multilayered TiVCT MXene was obtained. The obtained multilayered TiVCT was stirred with 25% TMAOH solution for 24 h at room temperature. Then, the mixture was centrifuged at 8500 rpm for 5 min, and the supernatant was collected. The TiVCT nanosheets were obtained by washing the supernatant with distilled water until the pH of the supernatant was neutral and filtered. The final TiVCT nanosheet powders were dried at 50 °C for 12 h in a vacuum oven (DZF-6020, Shanghai, China).

2.3 Preparation of TiVCT nanosheets@NF

After sonication for 1 h, 5 mg of the TiVCT nanosheets and 20 μL Nafion solution (5 wt%) were dispersed in a 980 μL deionized water/ethanol mixed solution with a volume ratio of 1 : 1. Then, 5 μL of the mixed solution was evenly added to commercial NF (1 cm × 1 cm). After that, the NF loaded with the catalyst was put into a vacuum sintering furnace and sintered in vacuum at 300 °C for 2 h, and TiVCT nanosheets@NF was obtained.

2.4 Characterization

The sample powders were analyzed by X-ray diffraction (XRD, D/ MAX-rA) using an X-ray diffractometer equipped with Cu Kα radiation (λ = 1.541 Å). The morphologies of the samples were characterized by scanning electron microscopy (SEM, JSM-5600LV) and transmission electron microscopy (TEM, F200X). The elemental composition of the samples was characterized by scanning electron microscopy equipped with energy dispersive X-ray spectrometry (EDS). The sample surface was characterized by X-ray photoelectron spectroscopy using Mg Kα as the excitation source.

2.5 Electrocatalytic performance evaluation

All electrochemical measurements were carried out using a typical three-electrode configuration with an electrochemical workstation (CHI660E, Shanghai, China). TiVCT nanosheets@NF was used as the working electrode (the loading weight is 0.025 mg cm−2), and an Ag/AgCl electrode and glassy carbon electrode were used as the reference electrode and counter electrode, respectively. The HER testing was performed in 0.5 mol L−1 H2SO4.

For the HER, the potential calibration of the reversible hydrogen electrode (RHE) was calibrated with respect to the following equation:

$$E(\text{RHE}) = E(\text{Ag}/\text{AgCl}) + 0.0591 \text{pH} + 0.196$$ (1)
Cyclic voltammetry (CV) was used to measure the double-layer capacitance ($C_{dl}$), which was recorded from $-6$ to $94$ mV vs. RHE at scan rates from 20 to 120 mV s$^{-1}$ at an interval of 20 mV s$^{-1}$. Linear sweep voltammetry (LSV) was performed at a scan rate of 10 mV s$^{-1}$ with 90% IR compensation by sweeping the potential from 0 to $-0.6$ V vs. RHE. Electrochemical impedance spectroscopy (EIS) was performed in a frequency range of 100 kHz to 0.01 Hz at an overpotential of open-circuit potential (OCP). The long-term stability test was performed by successive CV at a scan rate of 100 mV s$^{-1}$ for 1000 cycles. Lastly, chronoamperometric measurements at the requisite overpotential to 10 mA cm$^{-2}$ were used to evaluate the long-term durability.

3. Results and discussion

In this study, multi-layered TiVCT$_x$ MXenes with different abundant functional groups were successfully prepared through the etching method of removing Al atoms from TiVAIC powder using HF. Ultrathin nanostructured TiVCT$_x$ was obtained by exfoliation. Next, the TiVCT$_x$ nanosheet powders were loaded on commercial nickel foam (NF), and then placed in a vacuum furnace for annealing at 300$^\circ$C for 2 h to obtain TiVCT$_x$@NF hybrids (Fig. 1). Fig. 2(a) shows an SEM image of the etched products. It can be clearly seen that a lamellar morphology appears, and these nanosheets are produced by HF etching away the Al in TiVAIC. In order to acquire the micro-morphological and structural information of the TiVCT$_x$ MXene in more detail, Fig. 2(b) shows a TEM image of multilayer TiVCT$_x$ and Fig. 2(c) shows a TEM image of few-layer TiVCT$_x$. The multilayered TiVCT$_x$ nanosheets exhibit clearly visible ripples, indicating the ultrathin feature of the TiVCT$_x$ nanosheets. The multilayered TiVCT$_x$ nanosheets present an unordered vertical arrangement. By calculation, the average nanolayer spacing is 2.2 nm (Fig. 2(d and e)).

Fig. 3(a) shows the XRD patterns of TiVAIC and the TiVCT$_x$ MXene after the HF solution etching. It can be seen that the XRD patterns before and after TiVAIC etching are obviously different, indicating the successful synthesis of the TiVCT$_x$ MXene. After the etching of TiVAIC, the diffraction peak intensity at 40.38$^\circ$ is significantly weakened. Additionally, in the spectrum of the TiVCT$_x$ MXene, a diffraction peak appears at a small angle of 6.85$^\circ$, indicating the introduction of other molecules (according to a study by Li et al.$^{17}$). Fig. 3(b) shows the surface element distribution of the TiVCT$_x$ MXene. It can be seen that the Ti, V, C, F, and O elements are well and uniformly dispersed on the nanosheet structure. The presence of the element O is due to the intercalation of oxygen-containing groups between the nanosheets.

Fig. 4 shows the micromorphology and the results of the elemental distribution of the TiVCT$_x$@NF hybrid materials. Fig. 4(a) exhibits the microstructure of the porous nickel foam framework. In Fig. 4(b), it can be clearly seen that plenty of white nanoparticles are attached to the surface of the nickel foam after hydrothermal treatment. In Fig. 4(c), it can be obviously seen that TiVCT$_x$@NF is composed of Ni, Ti, O, C, F, and V, and the six elements are uniformly distributed, indicating the successful synthesis of TiVCT$_x$@NF.
As shown in Fig. 5, XPS was used to analyze the element valence composition of TiVCTₓ, and PeakFit was used to perform peak processing on the XPS data, and the corresponding characteristic peaks were fixed according to the binding energy. As shown in Fig. 5(b), the spectrum of Ti 2p contains two singlets and one doublet; the peaks at 454.8 eV and 455.36 eV correspond to Ti–C; the peaks at 456.3 eV, 459.2 eV and 464.70 eV correspond to Ti–O; and the peak at 461 eV corresponds to Ti–F. As shown in Fig. 5(c), the binding energy peaks identified for TiVCTₓ at 281.6, 282.1, 285.0, 285.3, and 286.3 eV correspond to C═O, C–O, graphitic C–C, C–V, and C–Ti. As shown in Fig. 5(d), the spectrum of F 1s contains two singlets and one doublet. The spectrum is easily fitted by one doublet peak of F–C, which contains two peaks at binding energies of 687.5 eV and 688.5 eV. The peak at 685.4 eV is related to the characteristic F–V, and the peak at 686.4 eV corresponds to C–Ti–Fₓ. The results suggest the embedding of –F in the MXene layered structure. As shown in Fig. 5(e), the spectra of V 2p and O 1s contain V–C, VₓOᵧ, V–F, O–Ti, O═V and –OH peaks, indicating the successful etching of Al and the combination of O. In general, it can be clearly seen from the above results that the Al atomic layer is completely etched, and new functional groups (–F, –OH, –O) are introduced. The addition of these functional groups (–F, –OH, –O) may increase the HER catalytic
activity of the MXene nanosheets. With the increased adsorption energy of protons, the charge transfer resistance decreases. Due to the surface functional groups, the TiVCT nanosheets are negatively charged. Combining the analysis results, the following equations are obtained:

$$2\text{TiVAIC} + 6\text{HF} \rightarrow 2\text{TiVC} + 2\text{AlF}_3 + 3\text{H}_2$$ \hspace{1cm} (2)

$$\text{TiVC} + 2\text{HF} \rightarrow \text{TiVCF}_2 + \text{H}_2$$ \hspace{1cm} (3)

$$\text{TiVC} + 2\text{H}_2\text{O} \rightarrow \text{TiVCO}_2 + 2\text{H}_2$$ \hspace{1cm} (4)

$$\text{TiVC} + 2\text{H}_2\text{O} \rightarrow \text{TiVC(OH)}_2 + \text{H}_2$$ \hspace{1cm} (5)

Electrochemical double-layer capacitance ($C_{dl}$) is an electrochemical parameter that can evaluate the difference in the electrochemically active surface area of different samples and can be determined by cyclic voltammetry (CV). The number of active sites is one of the important factors determining the HER activity, and the electrochemically active surface area depends on the number of active sites. Fig. 6(a), (c) and (e) present the cyclic voltammetry curves of TiVCT, @NF, TiVCTx, and commercial Pt/C, respectively, measured at scan rates of 20–120 mV s$^{-1}$ with an interval of 20 mV s$^{-1}$. Fig. 6(b), (d) and (f) present the curves of capacitive current versus scan rate for TiVCT, @NF, TiVCTx and commercial Pt/C at 0.15 V, respectively. It can be seen from the above results that the $C_{dl}$ of TiVCT, @NF is 26.9 mF cm$^{-2}$, which is much larger than that of TiVCTx (3.8 mF cm$^{-2}$), indicating that through the combination with nickel foam, the specific surface area is increased, the active sites are increased, and the hydrogen evolution catalytic performance is enhanced. Besides, the $C_{dl}$ of TiVCT, @NF is close to the $C_{dl}$ of the Pt/C electrode (31.8 mF cm$^{-2}$).

In order to explain the good electrochemical behavior of the TiVCNT nanosheets, the relevant tests on their hydrogen evolution performance were carried out. Using a standard three-electrode electrochemical system, the electrochemical activity of TiVCTx and its hybrid products for the HER was tested in a 0.5 mol L$^{-1}$ H$_2$SO$_4$ electrolyte. Besides, the commercial platinum–carbon catalyst (10 wt% Pt/C) with extremely high hydrogen evolution activity and pure nickel foam (NF) were used as references in this test. Fig. 7(a) shows the HER polarization curves of TiVCTx, NF, TiVCT, @NF and commercial platinum–carbon in 0.5 mol L$^{-1}$ H$_2$SO$_4$. It can be seen from the figure that the hydrogen evolution overpotential of the TiVCTx nanosheets obtained after exfoliation at 10 mA cm$^{-2}$ is 583 mV, while the hydrogen evolution overpotential of TiVCT, @NF after modification is 151 mV, which is close to that of the platinum–carbon electrode (33 mV). The hydrogen evolution overpotential of NF at 10 mA cm$^{-2}$ is 261 mV, indicating that the hydrogen evolution catalytic activity of NF is stronger than that of TiVCTx, and the hydrogen evolution catalytic activity is further improved after the combination of TiVCTx and NF. TiVCT, @NF can be selected as a candidate for excellent hydrogen evolution catalyst materials.

The Tafel slope reveals the rate-determining steps of the hydrogen evolution catalytic reaction. Generally, when the Tafel slope reaches 30 mV, 40 mV or 120 mV, the rate-determining step of the hydrogen evolution catalytic reaction is identified as a Tafel, Heyrovsky or Volmer step, respectively. To understand the kinetics of the hydrogen evolution catalytic reaction of the hybrid TiVCT, @NF in the HER, the Tafel slope was used to reveal its reaction rate-determining steps. Fig. 7(b) shows the Tafel slope curves of the four samples in 0.5 mol L$^{-1}$ H$_2$SO$_4$. It can be seen that the Tafel slopes of TiVCT, @NF, TiVCT, @NF and commercial platinum–carbon are 213 mV dec$^{-1}$, 180 mV dec$^{-1}$, 116 mV dec$^{-1}$ and 29 mV dec$^{-1}$, respectively. The electrochemical test results show that the water splitting of the hybrid TiVCT, @NF in the hydrogen evolution catalytic process is based on the Volmer–Heyrovsky reaction mechanism, which involves the rapid hydrogen adsorption Volmer reaction (H$_2$O$^+$ + e$^-$ → H$_{ads}$ + H$_2$O) and the slow hydrogen desorption Heyrovsky reaction (H$_{ads}$ + H$_2$O$^+$ + e$^-$ → H$_2$ + H$_2$O). Previous
studies have shown that the active site of the HER is generally the –OH group, and the two basic steps of water splitting are as follows: (1) H$_3$O$^+$ ions are first adsorbed on the –OH group and then combined with electrons to generate a hydrogen atom. (2) The hydrogen atom generated in step 1 combines with an H$_3$O$^+$ ion and an electron to generate H$_2$. The hydrogen desorption process in step 2 is the rate-determining step of the hydrogen evolution catalytic reaction. In addition, other scholars have reported the excellent electrochemical performance of MXene electrodes, as shown in Fig. 7(c) and (d); with the same electrolyte conditions, the TiVCT$_x$@NF electrodes are superior to MXene electrodes (Ti$_3$C$_2$T$_x$, RuS$_2$-Ti$_3$C$_2$T$_x$ (ref. 30) and Mo$_2$TiC$_2$T$_x$@NF$^{31}$).

Furthermore, electrochemical impedance spectroscopy (EIS) was used to reveal the electron transfer kinetics of TiVCT$_x$@NF during the hydrogen evolution reaction. Fig. 8(a) presents a Nyquist plot of the four samples. In the Nyquist diagram, the impedance value represents the resistance encountered in the transfer of charges in the HER, and the smaller the arc radius, the smaller the impedance value of the test sample, indicating that the hydrogen evolution catalytic reaction is more likely to occur. The size of the semicircle is related to the electrode surface properties and determines the size of the charge transfer resistance ($R_{ct}$). It is obvious that the impedance value of the hybrid TiVCT$_x$@NF is significantly lower than that of the TiVCT$_x$ MXene. The electrochemical impedance spectrum is fitted using the Z-view software; the fitted resistance circuit diagram is shown in Fig. 8(b) and the related resistance parameter fitting results are presented in Table 1. $R_1$ represents the equivalent series resistance (ESR), which consists of three parts: the ionic resistance of the electrolyte, the resistance inside the active material and the current collector, and the interfacial contact resistance between the electrode and the electrolyte. $R_2$ represents the
diffusion resistance, which refers to the resistance caused by the transport of ions in a thin layer, which is related to the current changes, concentration changes, and electrode materials. $R_3$ represents the charge transfer resistance ($R_{ct}$). As shown in Table 1, the fitted resistance of TiVCTₓ@NF is smaller than that of the other two electrode materials, and the
$R_i$ of TiVCT$_x$@NF is only 186.5 $\Omega$. This indicates that the TiVCT$_x$@NF electrode system has a smaller internal resistance and larger electroactive surface, and it is easier for ions to penetrate between the sheets of the electrode material. These results indicate that TiVCT$_x$@NF has better electrochemical performance.
To evaluate the comprehensive performance of a hydrogen evolution catalyst, it is necessary not only to consider the catalytic activity of the hydrogen evolution reaction but also to test its stability. In order to examine the catalytic stability of the TiVCTx@NF composites in 0.5 mol L\(^{-1}\) H\(_2\)SO\(_4\) solution, long-term CV cycling tests for 1000 cycles were carried out. As shown in Fig. 9(a), the time-dependent electrical density curve remains almost unchanged for the HER after 1000 test cycles. As shown

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**Fig. 7**  (a) Polarization curves and (b) corresponding Tafel plots of TiVCTx, TiVCTx@NF, NF and commercial Pt/C at a scan rate of 10 mV s\(^{-1}\) in a 0.5 M H\(_2\)SO\(_4\) solution. (c) Polarization curves and (d) corresponding Tafel plots of Ti3C2Tx, TiVCTx@NF, RuSx-Ti3C2Tx and Mo2TiC2Tx@NF at a scan rate of 10 mV s\(^{-1}\) in a 0.5 M H\(_2\)SO\(_4\) solution.

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**Fig. 8**  (a) Electrochemical impedance spectra of Pt/C, NF, TiVCTx and TiVCTx@NF for the electrocatalytic HER at an overpotential of open-circuit potential. (b) Equivalent circuit diagram of the TiVCTx@NF composites.
in Fig. 9(b), the catalytic activity for hydrogen evolution of TiVCT\textsubscript{x}@NF remains almost unchanged at 0.15 V for 12 h. From the above results, it can be seen that the hybrid TiVCT\textsubscript{x}@NF exhibits excellent HER catalytic activity and good stability during long-term electrolysis, indicating that TiVCT\textsubscript{x}@NF is a highly stable electrocatalyst for HER.

4. Conclusion

In summary, TiVCT\textsubscript{x}@NF was successfully prepared using the annealing method. Selecting a typical three-electrode system, the comprehensive electrochemical performance of TiVCT\textsubscript{x}@NF was tested. The experimental results show that the composite TiVCT\textsubscript{x}@NF exhibits excellent electrochemical properties with a low overpotential of 151 mV at 10 mA cm\textsuperscript{-2} and a small Tafel slope of 116 mV dec\textsuperscript{-1}. In a 0.5 M H\textsubscript{2}SO\textsubscript{4} solution, TiVCT\textsubscript{x}@NF exhibits excellent HER catalytic activity after 1000 test cycles and long-term stability over 12 h. Compared to the TiVCT\textsubscript{x} MXene, TiVCT\textsubscript{x}@NF has more active sites and less resistance encountered in the transfer of charges in the HER. Based on these excellent electrochemical properties, TiVCT\textsubscript{x}@NF is expected to be applied as an electrocatalytically active catalyst for water splitting.

Conflicts of interest

The authors declare no competing financial interest.

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References

1. S. M. Fernández-Valverde, in Hydrogen: The Ecological Fuel for Mexican Future, Springer, 2007.
2. S. H. Jensen, P. H. Larsen and M. Mogensen, Hydrogen and synthetic fuel production from renewable energy sources, Int. J. Hydrogen Energy, 2007, 32, 3253–3257.
3. M. G. Walter, E. L. Warren, J. R. McKone, S. W. Boettcher, Q. Mi, E. A. Santori and N. S. Lewis, Solar water splitting cells, Chem. Rev., 2010, 110, 6446–6473.
4. Y. Wei, R. A. Soomro, X. Xie and B. Xu, Design of efficient electrocatalysts for hydrogen evolution reaction based on 2D MXenes, J. Energy Chem., 2021, 55, 244–255.
5. Z. Li, Z. Qi, S. Wang, T. Ma, L. Zhou, Z. Wu, X. Luan, F. Lin, M. Chen and J. T. Miller, In situ formed Pt\textsubscript{x}Ti nanoparticles on a two-dimensional transition metal carbide (MXene) used as efficient catalysts for hydrogen evolution reactions, Nano Lett., 2019, 19, 5102–5108.
6. H. Wang, Y. Wu, J. Zhang, G. Li, H. Huang, X. Zhang and Q. Jiang, Enhancement of the electrical properties of MXene Ti\textsubscript{3}C\textsubscript{2} nanosheets by post-treatments of alkalinization and calcination, Mater. Lett., 2015, 160, 537–540.
7. Z. Guo, L. Gao, Z. Xu, S. Teo, C. Zhang, Y. Kamata, S. Hayase and T. Ma, High electrical conductivity 2D MXene serves as

Table 1 EIS fitting parameters of the TiVCT\textsubscript{x}, NF and TiVCT\textsubscript{x}@NF electrodes

| Working electrode | $R_1$  | $R_2$  | $R_3$  | $C_1$  | CPE\textsubscript{T1} | CPE\textsubscript{T2} |
|-----------------|--------|--------|--------|--------|----------------|----------------|
| TiVCT\textsubscript{x} | 15.62  | 115.1  | 2334   | $1.02 \times 10^{-5}$ | 6.39 $\times 10^{-6}$ | 0.8533 |
| NF              | 15.34  | 9.447  | 442.8  | $2.89 \times 10^{-5}$ | 5.17 $\times 10^{-6}$ | 0.8693 |
| TiVCT\textsubscript{x}@NF | 1.33   | 3.695  | 186.5  | 0.0079 | 0.0021         | 0.8247 |

Fig. 9 (a) Polarization curves of TiVCT\textsubscript{x}@NF initially and after 1000 CV cycles. (b) Current density versus time for TiVCT\textsubscript{x}@NF at a 150 mV potential.

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additive of perovskite for efficient solar cells, *Small*, 2018, 14, 1802738.
8. Y. Zhang, X. Zha, K. Luo, Y. Qin, X. Bai, J. Xu, C. Lin, Q. Huang and S. Du, Theoretical study on the electrical and mechanical properties of MXene multilayer structures through strain regulation, *Chem. Phys. Lett.*, 2020, 760, 137997.
9. Z. W. Seh, K. D. Fredrickson, B. Anasori, J. Kibsgaard, A. L. Strickler, M. R. Lukatskaya, Y. Gogotsi, T. F. Jaramillo and A. Vojvodic, Two-dimensional molybdenum carbide (MXene) as an efficient electrocatalyst for hydrogen evolution, *ACS Energy Lett.*, 2016, 1, 589–594.
10. R. Tang, S. Xiong, G. Deng, Y. Wang, L. Su, C. Ding, L. Yang and C. Liao, Ti$_{3}$C$_{2}$ 2D MXene: recent progress and perspectives in photocatalysis, *ACS Appl. Mater. Interfaces*, 2020, 12, 56663–56680.
11. L. Huang, L. Ai, M. Wang, J. Jiang and S. Wang, Hierarchical MoS$_2$ nanosheets integrated Ti$_3$C$_2$ MXenes for electrocatalytic hydrogen evolution, *Int. J. Hydrogen Energy*, 2019, 44, 965–976.
12. T. A. Le, Q. V. Bui, N. Q. Tran, Y. Cho, Y. Hong, Y. Kawazoe and H. Lee, Synergistic effects of nitrogen doping on MXene for enhancement of hydrogen evolution reaction, *ACS Sustainable Chem. Eng.*, 2019, 7, 16879–16888.
13. J. Liu, Y. Liu, D. Xu, Y. Zhu, W. Peng, Y. Li, F. Zhang and X. Fan, Hierarchical “nanoroll” like MoS$_2$/Ti$_3$C$_2$T$_x$ hybrid with high electrocatalytic hydrogen evolution activity, *Appl. Catal., B*, 2019, 241, 89–94.
14. Z. He, T. Rong, Y. Li, J. Ma, Q. Li, F. Wu, Y. Wang and F. Fang, Two-Dimensional TiVC Solid-Solution MXene as Surface-Enhanced Raman Scattering Substrate, *ACS Nano*, 2022, 16, 4072–4083.
15. S. Yazdanparast, S. Soltanmohammad, A. Fash-White, G. J. Tucker and G. L. Brennecke, Synthesis and surface chemistry of 2D TiVC solid-solution MXenes, *ACS Appl. Mater. Interfaces*, 2020, 12, 20129–20137.
16. Y. Li, L. Li, R. Huang, Y. Zhang and Y. Wen, Computational screening of pristine and functionalized ordered TiVC MXenes as highly efficient anode materials for lithium-ion batteries, *Nanoscale*, 2021, 13, 2995–3001.
17. Y. Li, J. Zhang, Y. Cheng, K. Feng, J. Li, L. Yang and S. Yin, Stable TiVCT$_x$/poly-o-phenylenediamine composites with three-dimensional tremella-like architecture for supercapacitor and Li-ion battery applications, *Chem. Eng. J.*, 2022, 34578.
18. J. R. McKone, B. F. Sadtler, C. A. Werlang, N. S. Lewis and H. B. Gray, Ni-Mo nanopowders for efficient electrochemical hydrogen evolution, *ACS Catal.*, 2013, 3, 166–169.
19. G. Wang, J. Huang, S. Chen, Y. Gao and D. Cao, Preparation and supercapacitance of CuO nanosheet arrays grown on nickel foam, *J. Power Sources*, 2011, 196, 5756–5760.
20. W. Kong, C. Lu, W. Zhang, J. Pu and Z. Wang, Homogeneous core-shell NiCo$_2$S$_4$ nanostructures supported on nickel foam for supercapacitors, *J. Mater. Chem. A*, 2015, 3, 12452–12460.
21. T. Hussain, L. Truong, S. Hussain, S. A. Pital, S. Chun and M. Kaseem, Novel core-shell structured electrocatalyst with 1D-NiTe$_2$ over 3D metal structure for efficient hydrogen evolution reactions, *J. Alloys Compd.*, 2022, 165797.
22. T. Hussain, M. Hussain, S. Hussain and M. Kaseem, Microwave-assisted synthesis of NiTe$_2$ photocatalyst as a facile and scalable approach for energy-efficient photocatalysis and detoxification of harmful organic dyes, *Sep. Purif. Technol.*, 2022, 282, 120025.
23. H. Zou, X. Li, C. Zhang, Y. Wen, Y. Fan, Y. Liu, L. Xiong, X. Zheng and J. Yang, Reactive synthesis for porous TiVAlC ceramics by TiH$_2$, V and graphite powders, *Ceram. Int.*, 2021, 47, 28288–28295.
24. S. Li, P. Tuo, J. Xie, X. Zhang, J. Xu, J. Bao, B. Pan and Y. Xie, Ultrathin MXene nanosheets with rich fluorine termination groups realizing efficient electrocatalytic hydrogen evolution, *Nano Energy*, 2018, 47, 512–518.
25. X. Wang, J. Wang, J. Qin, X. Xie, R. Yang and M. Cao, Surface charge engineering for covalently assembling three-dimensional MXene network for all-climate sodium ion batteries, *ACS Appl. Mater. Interfaces*, 2020, 12, 39181–39194.
26. P. Sharma and T. S. Bhatti, A review on electrochemical double-layer capacitors, *Energy Convers. Manage.*, 2010, 51, 2901–2912.
27. S. A. Vilekar, I. Fishtik and R. Datta, Kinetics of the hydrogen electrode reaction, *J. Electrochem. Soc.*, 2010, 157, B1040.
28. A. Lasia, Mechanism and kinetics of the hydrogen evolution reaction, *Int. J. Hydrogen Energy*, 2019, 44, 19484–19518.
29. A. D. Handoko, K. D. Fredrickson, B. Anasori, K. W. Convey, L. R. Johnson, Y. Gogotsi, A. Vojvodic and Z. W. Seh, Tuning the basal plane functionalization of two-dimensional metal carbides (MXenes) to control hydrogen evolution activity, *ACS Appl. Energy Mater.*, 2017, 1, 173–180.
30. V. Ramalingam, P. Varadhan, H. C. Fu, H. Kim, D. Zhang, S. Chen, L. Song, D. Ma, Y. Wang and H. N. Alshareef, Heteroatom-mediated interactions between ruthenium single atoms and an MXene support for efficient hydrogen evolution, *Adv. Mater.*, 2019, 31, 1903841.
31. J. Wang, P. He, Y. Shen, L. Dai, Z. Li, Y. Wu and C. An, FeNi nanoparticles on Mo$_2$TiC$_2$Tx MXene@nickel foam as robust electrocatalysts for overall water splitting, *Nano Res.*, 2021, 14, 3474–3481.