A Quasi-Intramolecular Solid-Phase Redox Reaction of Ammonia Ligands and Perchlorate Anion in Diamminesilver(I) Perchlorate

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Abstract: The reaction of ammoniacal AgNO₃ solution (or aq. solution of [Ag(NH₃)₂]NO₃) with aq. NaClO₄ resulted in [Ag(NH₃)₂]ClO₄ (compound 1). Detailed spectroscopic (correlation analysis, IR, Raman, and UV) analyses were performed on [Ag(NH₃)₂]ClO₄. The temperature and enthalpy of phase change for compound 1 were determined to be 225.7 K and 103.04 kJ/mol, respectively. We found the thermal decomposition of [Ag(NH₃)₂]ClO₄ involves a solid-phase quasi-intramolecular redox reaction between the perchlorate anion and ammonia ligand, resulting in lower valence chlorine oxyacid (chlorite, chlorate) components. We did not detect thermal ammonia loss during the formation of AgClO₄. However, a redox reaction between the ammonia and perchlorate ion resulted in intermediates containing chlorate/chlorite, which disproportionated (either in the solid phase or in aqueous solutions after the dissolution of these decomposition intermediates in water) into AgCl and silver perchlorate. We propose that the solid phase AgCl-AgClO₄ mixture eutectically melts, and the resulting AgClO₄ decomposes in this melt into AgCl and O₂. Thus, the final product of decomposition is AgCl, N₂, and H₂O. The intermediate (chlorite, chlorate) phases were identified by IR, XPS, and titrimetric methods.

Keywords: ammine; perchlorate; vibrational spectroscopy; redox reaction; thermal decomposition; disproportionation; chlorite; chlorate

1. Introduction

Continuing our research on thermally initiated solid-phase, quasi-intramolecular redox reactions of compounds containing redox-active cationic and anionic parts [1–4], including ammonia complexes of transition metal with tetraoxometallate anions [5–10], we prepared some diamminesilver(1) complexes (e.g., [Ag(NH₃)₂]XO₄, X = Mn and Cl) and studied them in detail. In this paper, we present our results on the investigation of monoclinic and orthorhombic diamminesilver(I) perchlorate polymorphs (1-M and 1-O).
The hydrated [Ag(NH₃)₂]ClO₄·H₂O (1·H₂O) complex was described by Scagliari and Marangoni [11], but Bruni and Levi [12] raised issues against its existence. The occurrence of [Ag(NH₃)₃]ClO₄ (compound 2) in the solid state and the quantum-chemical results about [Ag(NH₃)₂(H₂O)]⁺ ion [13] prompted us to study the existence of hydrated forms of compound 1 [Ag(NH₃)₂(H₂O)]ClO₄ (compound 3) and [Ag(NH₃)₂]ClO₄·H₂O (1·H₂O).

Diamminesilver(I) perchlorate has two known polymorphs [14], a low-temperature monoclinic (1-M) and a high-temperature orthorombic one (1-M); however, the enthalpy of phase transition between them has not been studied yet. Furthermore, no spectroscopic data are available about the low-temperature polymorph. Although IR spectroscopic data are available for 1-O, their peak assignments are questionable. Furthermore, the thermal decomposition of compound 1 (1-O) has not been investigated previously. Therefore, in our present work, we supplied the available spectroscopic data about polymorph 1-O with low-temperature Raman results, compared their measured Raman spectroscopic data, and performed a correlation analysis for both polymorphs. The thermal analysis data, including DSC and TG-MS, were also evaluated to follow the thermal decomposition characteristics of compound 1 (1-O). Based on these results, we propose the possibility of solid-phase quasi-intramolecular redox reactions between ammonia ligand and perchlorate anion.

2. Results and Discussion
2.1. Efforts to Synthesize of Compound 1 and Its Hydrate(s)

Diamminesilver(I) perchlorate was first prepared by E. Mitscherlich [15] as a colorless crystalline material, when silver perchlorate was dissolved in ammonia, and the solution was left to evaporate. He did not determine the composition of the product. However, Miles et al. [16], and later Nockeman and Meyer [14], used the same procedure and identified the crystallized compounds as [Ag(NH₃)₂]ClO₄ (compound 1). Scagliari and Marangoni reacted an ammoniacal silver nitrate solution with aq. sodium perchlorate at +5 °C [11] and analyzed the resulting crystalline mass as [Ag(NH₃)₂]ClO₄·H₂O (compound 1-H₂O). Bruni and Levi repeated this experiment [12] and found that the product of the reaction was anhydrous compound 1. To solve this controversy, we performed a series of experiments to study the effect of reaction conditions (ammonia concentration, temperature) on the composition of products prepared in the reactions of AgClO₄ and ammonia or [Ag(NH₃)₂]NO₃ and NaClO₄. In all cases, only anhydrous compound 1 was formed. Figure 1 shows the XRD of the reaction products according to the original methods of Mitscherlich and Bruni. The XRD patterns of the products of the two reactions were the same; thus, the formation of compound 1-H₂O or compound 2 in these reactions can be excluded (Figure 2). However, several researchers managed to prepare compound 2 from AgClO₄ and dry ammonia [12,17].

Quantum chemical calculations showed that there was no stable [Ag(NH₃)₂(OH₂)]⁺ structure [3 + 0] (three coordinated and zero non-coordinated ligands), and this species always converges into a linear [Ag(NH₃)₂]⁺ ...OH₂ hydrogen-bound structure [2 + 1] (two coordinated and one non-coordinated ligands), although a small interaction between the oxygen (in water) lone pair and the silver cation was detected [13]. The [Ag(NH₃)₃]⁺ [3 + 0] cation, however, has stable configurations [13,18].

\[
[\text{Ag(NH}_3\text{)}_2\text{(OH}_2\text{)}]^+ = [\text{Ag(NH}_3\text{)}_2]^+ \cdot \text{...OH}_2 
\]

\[
[\text{Ag(NH}_3\text{)}_2]^+ + \text{NH}_3 = [\text{Ag(NH}_3\text{)}_3]^+ 
\]

Comparison of the Gibbs energy changes during the substitution of one of the ammonia molecules with water in the [Ag(NH₃)₃]⁺ cation (+7.5 kJ/mol [13]) (Equation (1)) and the addition of the third ammonia to an [Ag(NH₃)₂]⁺ cation (−30.5 kJ/mol [19])
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(Equation (2)) suggest that the formation of hydrated diamminesilver(I) complexes is unfavorable.

Figure 1. Comparison of the XRD curves of the reaction products prepared with the methods of Mitscherlich [15] (blue) and Scagliari [11] (red).

Figure 2. Schematic representation of phases formed in the AgClO₄-NH₃ and AgClO₄-NH₃-H₂O systems.

We managed to isolate [Ag(NH₃)₃]ClO₄ in the solid phase. The heat that formed in the solid phase (−363.7 kJ/mol) and in the aqueous solutions (−318.5 kJ/mol) did not differ much, but the hydration of liberated ammonia in the aqueous solution may change the Gibbs energy of the reaction to be positive, which prevents the direct formation of the hydrated diamminesilver(I) perchlorate (1-H₂O) [20,21].

2.2. Phase Transition of Compound 1

Compound 1 has two polymorphs (monoclinic and orthorhombic), and the phase transition point was between 200 and 210 K [14]. The cryogenic DSC tests performed on the sample (Figure S1) showed that the phase transition temperature was 226 K, and the transition enthalpy was 103.04 kJ/mol.
2.3. General Spectroscopic Characterization and Correlation Analysis of Compounds 1-M and 1-O

The solid-state IR and the solid and aqueous solution-phase Raman spectra of the high-temperature orthorhombic polymorph of compound 1 (1-O) have already been discussed [14,16]. However, the spectral characteristics of the low-temperature monoclinic form (1-M) and the factor group analysis of the modifications have not been done yet. Therefore, correlation analyses were performed for perchlorate-ion, ammonia ligand, and silver ion (site groups were Cs and C2v, and factor groups were C2h and D2h, for all three species in monoclinic (P2/m) and orthorhombic (Pnmm) polymorphs, respectively). The correlation analysis showed the number of bands belonging to normal modes and distinguished them from combination or overtone bands of the spectra. The vibrational spectra of crystals may be quite difficult to interpret in the case of bands due to second-order (and higher-order) transitions. One of the problems is that although the total wave vector of the combined transition, \( k \), must be 0 (\( k = k_1 + k_2 = 0 \)), the individual values of the component vectors could be anywhere within the first Brillouin zone, which prevents the use of the terminology we are familiar with. In order to simplify things (and keeping in mind that this is not a paper that deals with solid-state physics primarily), we will continue to talk about “overtones” and “combinations” in the sense we use these terms for isolated molecules/ions. However, in reality, things are far more complicated.

2.4. Perchlorate Anion Modes

The isolated tetrahedral perchlorate ion has four normal modes, symmetric stretching (\( \nu_1(A_1) \) (R)) and bending (\( \nu_2(E) \) (R)), and antisymmetric stretching (\( \nu_3(F_2) \) (IR, R)) and bending (\( \nu_4(F_2) \) (IR, R)). The total number of factor group modes due to internal perchlorate vibrations is \( 2 \times 18 = 36 \) (2 \( \times \) 9(IR) + 2 \( \times \) 9 (R)) and 36 (18(IR) + 18(R)) vibrational degrees of freedom, for 1-M and 1-O, respectively (the monoclinic polymorph contains 2 crystallographically distinguishable perchlorate ions; thus, two sets of modes are available) (Figure 3). The hindered rotations (Rxyz)(F1) and translations (Txyz (F2)) of perchlorate ions result in \( 2 \times 12 = 24 \) (2 \( \times \) 6(IR) + 2 \( \times \) 6(R)) and 24 (12(IR) + 12(R)) vibrational degrees of freedom, respectively (Figure S2a,b).

![Figure 3. Correlation diagrams of perchlorate ions in (a) monoclinic (1-M) and (b) orthorhombic (1-O) polymorphs of [Ag(NH3)2]ClO4.](image-url)
2.5. Ammonia Ligand Modes

The C3v ammonia ligand has two symmetric (\(\nu_1(A_1)\) (IR, R) stretching and \(\nu_2(A_1)\) (IR, R) bending) and two antisymmetric (\(\nu_3(E)\) (IR, R) stretching and \(\nu_4(E)\) (IR, R) bending) internal normal modes. The total number of factor group modes due to internal vibrations is \(4 \times 12 = 48(4 \times 6(\text{IR}) + 4 \times 6(\text{R}))\) (four crystallographically distinguishable ammonia molecules) and \(2 \times 24 = 48 (2 \times 12(\text{IR}) + 2 \times 12(\text{R}))\) (two crystallographically distinguishable ammonia molecules) (Figure 4).

The external vibrational modes (\(T_z(A_1)\), \(R_z(A_2)\) and \(T_y(E)\) or \(R_y(E)\)) must be quadrupled due to four different ammonia molecules in the structure, which results in \(4 \times 12 = 48(4 \times 6(\text{IR}) + 4 \times 6(\text{R}))\) and \(2 \times 24 = 48 (2 \times 12(\text{IR}) + 2 \times 12(\text{R}))\) vibrational degrees of freedom (Figure S3a,b).

![Figure 4. Correlation diagrams of ammonia ligand in (a) monoclinic (1-M) and (b) orthorhombic (1-O) polymorphs of [Ag(NH3)2]ClO4.](image)

Silver cations, four and two crystallographic types in the 1-M and 1-O polymorphs, have \(4 \times 3 (T_z (A_u), T_x, T_y (2B_u))\) and \(2 \times 6 (T_z (A_u), T_x (2B_u), T_y (2B_u), \text{and } (T_x, T_y (2B_u)))\) vibrational degrees of freedom, respectively. All of these are only infrared active modes. For both monoclinic and orthorhombic polymorphs, among 48 modes of translational origin, there are 3 acoustic modes that belong to \(A_u + 2B_u\) modes, and the further 45 are optical modes (Figure 5). There are also 36 and 84 optical modes of rotational origin and due to internal vibrations, respectively.
2.6. Assignation of Vibrational Modes in Compound 1 Polymorphs

2.6.1. Anion Modes

The assignation of perchlorate IR bands and Raman shifts found in IR and Raman spectra of compound 1 can be seen in Table 1. The IR data were measured only at room temperature (O-1), but Raman shifts were recorded at 123 K (M-1) and 298 K (O-1) as well. The IR inactive symmetric stretching mode of perchlorate ion ($\nu_1$) appears as a single peak at 933 cm$^{-1}$ in the spectrum of 1-O, which unambiguously shows the distortion of symmetric Td geometry. The antisymmetric stretching mode, as the most intense Cl–O band, is centered around 1050 cm$^{-1}$. This band has a strongly asymmetric shape, and the shoulder detected on the left side shows the splitting of triple (F2) degeneracy as the consequence of the distortion of the geometry. The doubly and triply degenerated relatively weak $\nu_2$ (Cl–O) bands (450 cm$^{-1}$) and $\nu_4$ (Cl–O) bands (615 cm$^{-1}$), respectively, were not split in the room temperature spectrum. The band system between 2050 and 1950 cm$^{-1}$ contains $\nu_1 + \nu_3$ combinations and $2\nu_3$ type (a,b,c—each of three bands set from F2 (Au, 2B1u, 2B2u, and B3u)) first overtones (Figure S2).

We recorded the Raman spectra at three different temperatures (123, 223, and 298 K) to study the effect of temperature and phase transformation on the numerical value/splitting of Raman shifts of 1-O (298 and 223 K) and 1-M (123 K). The spectra in the Cl–O stretching/bending modes range (1200–300 cm$^{-1}$) can be seen in Figure 6. At room temperature and at 223 K, the Raman spectra of 1-O show a similar pattern. There is no pronounced splitting of bands, but a shoulder evolves into a band at 910 cm$^{-1}$, on the right at the bottom of the very intensive $\nu_1$ (A1) Raman band (933 cm$^{-1}$ and 929 cm$^{-1}$ for 1-O and 1-M, respectively). This small intensity band (shoulder) might be the first overtone of $\nu_2$ (Cl–O) (457 cm$^{-1}$ at 223 K). This overtone is a wideband system that may contain several components for 1-O and a band consisting of two components for 1-M. It is also plausible that isotopomers ($^{35}$Cl and $^{37}$Cl) may also contribute to this band [22]. Overtones of 2$\delta_{\text{w}}$(ClO)(F2) (Au, B3u) for 1-M appear together with band(s) of $\delta$(NH).
Figure 6. Raman spectra of compounds 1-O (223 and 298 K) and 1-M (123 K) in the Cl–O Raman shift range.

The combined effect of low temperature and phase transformation causes the splitting/appearance of new bands in the Raman spectrum of 1-M at 223 K. The symmetric ν2(Cl−O)(E) appears as a four-component band system, including a shoulder, which proves that two different types of perchlorate ions exist. The intensive and sharp ν1(Cl−O)(A1) band did not split. Overlapping (* in Table 1 and 2) of the ν3(Cl−O)(F2) bands with the combination band system of ν4a,b,c(Cl−O)(F2) (2Ag and Bg for 1-M) or 2A9g, B1g, B2g, and 2B3g for 1-O) and ν2a,b(Cl−O)(E) (Ag, Bg for 1-M and A9g, B1g, B2g, and B3g for 1-O) components resulted in a wide band, which appears at room temperature. There was a poorly separated band system (1107, 1098, 1089, 1084, 1072, 1058, 1054, and 1044 cm⁻¹) at 273 and 223 K for 1-O and 1-M, respectively. Combinations of ν2 + ν4 type seem to be located at the higher Raman shift side of the obtained range, while the ν3(Cl−O)(F2) bands are probably located in the lower or medium Raman shift range. The ν2(Cl−O)(E) band system contains two bands at 624 and 611 cm⁻¹, which might be attributed to the two types of perchlorate ion and the doublet nature of the bands as well.
Table 1. Assignment of perchlorate bands in IR and Raman spectra of compound 1. (*- mixed bands)

| IR, cm⁻¹ (298 K) | Raman Shift, cm⁻¹ (532 nm Excitation), Our Measurement | Raman Shift, cm⁻¹ | Assignation |
|------------------|--------------------------------------------------------|-------------------|-------------|
| **Our (1-O)**    | [16] [14]                                              | 123 K (1-M)       | 223 K (1-O) |
|                  |                                                        | 298 K (1-O)       | aq. soln. [16] | solid [14] |
| 933              | 940 -                                                  | 929 -             | 933         | 934 926 | ν₁(A1) (R) |
| 450              | 432 -                                                  | 482, 465 sh,      | 459         | 460 460 | ν₂(E) (R) |
| 1073 sh, 1053    | 1090 1084                                              | 1109–1044 *       | 1082, 1044  | 1107 - | ν₃(F₂) (IR, R) |
| 615              | 626 626                                                | 624, 612          | 625 625     | 629 - | ν₄(F₂) (IR, R) |
| 2050–1950 wide   | - -                                                    | 915 sh, 910 wide  | - -         | v₁ + ν₂ᵇᶜ and 2ν₄ᵇᶜ |
|                  | - -                                                    | 1263, 1239 wide   | - -         | 2ν₂ᵇ   |
|                  | - -                                                    | 1259             | 1239 1219   | 4ν₄ᵇᶜ   |

2.6.2. Cation Modes

The normal modes that belong to ammonia and the Ag–N linkages of compounds 1-O and 1-M are summarized in Table 2. The antisymmetric and symmetric N–H stretching (3364 and 3286 cm⁻¹, respectively) modes in the IR spectrum of 1-O partly overlap the first overtone of the δₛ(NH) (~2×1653 cm⁻¹, wide). The antisymmetric deformation of ammonia (ν₄) is a doubly degenerated mode, and for four different ammonias one could expect 4 bands in the IR spectrum. In the Raman spectrum, this band is located at 3370 cm⁻¹ (Aₘ, Bₙ), which decomposes into three components at 123 K (1-M, Aₗ, Bₙ, B₁ₕ, B₃ₕ, B₅ₕ, Table 2). The symmetric stretching band has a higher intensity in the Raman spectra than the antisymmetric one (3297 cm⁻¹ for 1-O (Aₙ, Bₙ), and 3299 and 3291sh cm⁻¹ (Aₗ, B₁ₕ) for 1-M). The first overtone bands of δₛ(NH) gave two IR (3212 and 3185 cm⁻¹, B₁ᵤ, B₂ᵤ) and two Raman bands (3191 and 3183 cm⁻¹, Bₗ, B₈) for 1-M. The symmetric δₛ(NH) Raman bands for 1-O and 1-M appear as a singlet band at 1259 cm⁻¹ and 1219 cm⁻¹, respectively. Since the 2δₛ(NH) has only one overtone band at 2439 cm⁻¹, only one band may be associated with this mode, and this is probably the band located at 1219 cm⁻¹. The other two bands located close to this mode (1263 and 1239 cm⁻¹) belonged to 2δₛ(ClO)(F₂) (2Aₗ, B₂ₖ) overtones. Due to the singlet nature of this mode, two IR bands (1244 and 1210 cm⁻¹) (B₉ and B₁₀) found in the IR spectrum of compound 1-O belong to the two different amino environments.

The wide IR band with the antisymmetric shape at 654 cm⁻¹ belongs to ρ(NH) mode, and one of its higher wavenumber components may result in the wide but weak overtone centered at 1353 cm⁻¹. In the Raman spectrum, this mode was found at 658 cm⁻¹ for 1-O, and a doublet appeared in the Raman spectrum of 1-M (679, 666 cm⁻¹).

The silver-nitrogen vibrational modes are in the far-IR region. The far-IR spectra of compound 1-O can be seen in Figure S4. There was a very wide band at 136 assigned as a mixed band of δ₉(OAgO), δ(OAgN), and δ(NAgN) modes, and a lower intensity and less wide band belonging to a lattice mode at 67 cm⁻¹.
### Table 2. Assignment of bands belonging to cation modes in IR and Raman spectra of compound 1. (*- mixed bands)

| IR | Raman | Assignation |
|----|-------|-------------|
| Our, 1-O, 298K [14] | Our, 1-M, 123K [14] | aq. solution [16] |
| 3364 | 3360 | 3370 | 3389, 3362 | \( \nu_{as}(\text{NH}) \) |
| 3285 | 3285 | 3290 | 3299, 3291 | \( \nu_{s}(\text{NH}) \) |
| 3193, 3182 | 3182 | 3185 | 3191, 3183 | \( \delta_{as}(\text{NH}) \) |
| 1610 | 1612 | 1625, 1610 | 1658, 1634 | \( \delta_{s}(\text{NH}) \) |
| 1351 | - | - | - | 2\( \rho(\text{NH}) \) |
| 1213, 1238 | 1250 | 1244 | 1259, 1219, 1219 | 1223 | 2\( \sigma(\text{NH}) \) |
| 653 | - | 650 | 658 | 679, 666, 666 | \( \rho(\text{NH}) \) |
| 430 | - | - | - | - | \( \nu_{s}(\text{AgN}) \) |
| - | - | 395* | 397 | 397 | \( \nu_{s}(\text{AgN}) \) |
| - | - | 379* | 383, 373 | 372 | \( \delta(\text{NAgN}) \) |
| - | - | 269 | - | - | - |
| - | - | 255 | 243 | 240 | - |
| | | 196 | - | - | - |
| 136 | | | | | Lattice, \( \delta(\text{NAgO}), \text{AgO}_2 \), 161 AgN\_2 |
| 67 | | | | | Lattice? |

#### 2.7. UV Spectroscopic Studies of Complex 1

Both the perchlorate and diamminesilver ions are colorless, thus no absorption is expected in the visible spectral region. Due to the closed d\(^{10}\) shell of the silver\(^+\) ion, there are no d-d transitions, but there is a weak Ag\(^{4d-5s}\) transition at 250 nm, and the wide band at 310 nm may be an O–Ag or N–Ag LMCT band (Figure S5) [22-24]. The perchlorate ion in itself is UV silent [10,24]).

#### 2.8. Non-Isothermal Decomposition of Compound 1 in an Inert Atmosphere and Air

Due to the strong oxidation power of the perchlorate anion and the reducing ability of ammonia ligands, there are two possible thermal decomposition routes of diammine-silver(I) perchlorate: (1) a solid-phase redox reaction between the perchlorate anion and the ammonia ligand, as was found in the case of permanganate [23], and (2) the loss of the ammonia ligand without any redox reaction, as was observed in the case of sulfate [25].

The decomposition process in an inert atmosphere starts with a slow mass decrease from about 50 °C. The process turns more intensive around 125 °C and results in a saw-tooth like DTG curve, with at least four decomposition processes, distinguished by their peak maxima at peak temperatures of 132, 173, 207, 251, and 408 °C. The character of the decomposition process is similar in both atmospheres (inert and air, Figure 7), but the mass decrease in argon is significantly higher than in air. The DTG peak maxima in air are also shifted below 300 °C by about 10 °C, and at higher temperatures by about 20 °C.
toward higher temperatures, compared to peak maxima in argon. This result suggests that oxygen in the air hinders the thermal decomposition of 1. The small mass loss of 0.7% in argon and 0.5% in air atmosphere up to ~125 °C most probably corresponds to adsorbed water and ammonia loss. Above this temperature, the decomposition processes overlap. In argon, in the next two steps (DTG maxima at 173 and 207 °C), complex 1 loses 7.6% and 7.8% of its mass, which is somewhat more than a molecule of NH₃ per step (7.06%). It suggests that not only ammonia loss occurs up to ~270 °C. An additional mass loss of 6.7% with a DTG peak at 251 °C is also in accordance with this. In air, the decomposition processes overlap more with DTG peaks at 181 and 219 °C and corresponding mass losses of 11.8% and 4.5%. The overall mass decrease up to ~270 °C in air (16.8%) is significantly less than in argon (22.8%). In argon, based on the TG results, a part of the ammonia evaporates, and the other part reacts with the oxygens of the perchlorate even in the solid phase. As the result of the decomposition reaction, the intermediate (AgClO₂₃) is formally the mixture of AgClO₃ and AgClO₂, in a ratio of approximately 1:1. The decomposition temperatures of supposed intermediates (AgClO₂, AgClO₃, AgClO₄) do not agree with the decomposition temperatures we found for the decomposition intermediates of compound 1. First, AgClO₄ does not decompose at 220 °C into lower chlorine oxyacid salts because the thermal stability of AgClOₙ (n = 2–4) salts increases with an increasing oxidation degree of chlorine in the oxyacid residue [26–28]. According to this result, the thermal decomposition temperatures of these chlorine oxyacid salts, AgClO₂, AgClO₃, and AgClO₄ are 156 °C, 270 °C, and 476 °C, respectively [26]. It means that the formal AgClO₂₃ composition, which decomposes at 441 °C, cannot be a mixture of AgClO₂ and AgClO₃, but only a mixture of AgCl and AgClO₄. On the other hand, in air, the mass of the intermediate at ~285 °C is 83.3%, which is in good agreement with the calculated value for AgClO₄ (85.89%). A possible explanation is that the aerial oxygen takes a role in ammonia oxidation instead of the ClO₄⁻ ion. The mass of the final residue in argon is very low, only 47.2%, somewhat higher than the metallic silver content of 1, while in air, it is 60.3% at 500 °C. This is in good agreement with the theoretical value for AgCl (59.38%).

Figure 7. DTG curve of compound 1 in an inert (□) and air (●) atmosphere.
In an inert atmosphere, the first two decomposition processes are followed by endothermic heat effects (DSC minima at 174 and 211 °C), and the third process by an exothermic effect with a DSC maximum at 256 °C. In air, the tendency of the DSC curve is the same, but with higher peak temperatures and lower enthalpy changes (Figure S6). These heat effects support the theory of partial ammonia oxidation discussed above.

The last decomposition step with a DTG maximum at 408 °C is an exothermic process, with a mass loss of 30.1%.

2.9. The Nature of the Final Decomposition Product Formed in the Thermal Decomposition of Compound 1

To solve the problems listed above, we studied the final decomposition product (P-480, prepared by heating the sample at 480 °C for 2 h) by XRD. The powder X-ray diffractogram of the P-480 sample showed the absence of metallic silver and the presence of silver chloride (Figure S7). Thus, the last decomposition step that produces AgCl is a thermal decomposition reaction of silver(I) perchlorate into silver(I) chloride and oxygen.

\[ \text{[AgClO}_4, \text{AgCl}] (l) \rightarrow \text{AgCl} (s) + 2\text{O}_2 (g) \] (3)

This reaction can explain the unexpected mass loss in the last decomposition step, because silver chloride melts at 452 °C (Figure S6) and its vapor pressure above the liquid and the inert gas stream results in a partial evaporation of AgCl [29]. The decomposition peak temperature (T\text{dec} = 441 °C) of AgClO\textsubscript{4} is lower than expected (T\text{dec} = 476 °C) [26], because AgCl (Mp. = 452 °C) and AgClO\textsubscript{4} (Mp. = 476 °C) form a eutectic melt (Mp. 325 °C; 1:1 AgCl:AgClO\textsubscript{4}), and the decomposition temperature of AgClO\textsubscript{4} in this melt is lower than the decomposition point of pure AgClO\textsubscript{4}. The silver perchlorate decomposes in the liquid melt phase in an autocatalytic process at a lower temperature than in the solid phase [26]. The reaction heat was measured to be -85.68 kJ/mol.

2.10. Redox Interactions in the Thermal Decomposition Steps of Compound 1

The presence of AgCl in the decomposition intermediate formed at 300 °C—deeply below the thermal decomposition temperature of AgClO\textsubscript{4}—confirms that AgCl was formed in the solid phase quasi-intramolecular redox reaction of [Ag(NH\textsubscript{3})\textsubscript{2}]ClO\textsubscript{4} (compound 1). The TG-M5 results (Figure 8, Figure S8) show that in an inert atmosphere, redox reactions take place between 180 and 300 °C, when N\textsubscript{2} (m/z = 28) (N\textsuperscript{2+}) and H\textsubscript{2}O (18 (H\textsubscript{2}O\textsuperscript{+})) as ammonia oxidation products are formed. Supposing that ammonia is only oxidized into N\textsubscript{2} and H\textsubscript{2}O, the six hydrogens of two molecules of ammonia in compound 1 can convert the perchlorate ion into formal “AgClO” (AgCl, AgCl + AgClO\textsubscript{2} = 1:1, 2AgCl + AgClO\textsubscript{3} or 3AgCl + AgClO\textsubscript{4}). This shows that 3/4 of the overall perchlorate ion content can only be reduced into AgCl (Figure 9).

In an inert atmosphere, a small amount of nitric oxide (m/z = 30, NO), oxygen (m/z = 16, O\textsuperscript{2+}), and HCl (m/z = 36, HCl\textsuperscript{+}) were also detected. Nitrous oxide (m/z = 44, N\textsubscript{2}O\textsuperscript{+}) was not produced. The peak at m/z = 16 cannot be an NH\textsubscript{2} fragment of ammonia, because then, a much more intensive m/z = 17 (NH\textsuperscript{+}) peak would be observed; instead of that, only a weak m/z = 17 (OH\textsuperscript{+}) peak was found (Figure S8) [3].

In air, NO is the main oxidation product. HCl did not form, and a small amount of N\textsubscript{2}O (m/z = 44, N\textsubscript{2}O\textsuperscript{+}) was also detected (Figure S9). The lack of HCl and the appearance of N\textsubscript{2}O might be attributed to the concurrent oxidation reaction with aerial oxygen instead of the perchlorate ion. The solid-state reaction product at 300 °C in air contains AgCl, AgClO\textsubscript{4} (cubic and tetragonal), and AgClO\textsubscript{4}.H\textsubscript{2}O, together with a minor amount of metallic Ag (Figure 10). Since AgClO\textsubscript{4} is extremely hygroscopic, AgClO\textsubscript{4}.H\textsubscript{2}O forms during the preparation of the samples for XRD by taking up humidity from the air.

With the use of stoichiometric relations, the formal reaction equations for the main and by-reactions (redox and non-redox) can be summarized according to Figure 9.
Figure 8. TG–MS of compound 1 in an inert atmosphere, the formation of \( m/z = 28 \) (\( \text{N}_2^+ \)) and \( m/z = 18 \) (\( \text{H}_2\text{O}^+ \)).

\[
2\text{Ag} + 2\text{HCl} + 5\text{H}_2\text{O} \rightarrow [1.5\text{O}_2 + 2\text{N}_2 \text{ or } 0.25\text{N}_2 + 3\text{NO}]
\]

\[
2[\text{Ag(NH}_3\text{)}_2]\text{ClO}_4 \rightarrow 1.5\text{AgCl} + 0.5\text{AgClO}_4 + 3\text{H}_2\text{O} + 2\text{N}_2 \rightarrow 2\text{AgCl} + \text{O}_2
\]

\[
2\text{AgClO}_4 + 4\text{NH}_3
\]

Figure 9. The presumed decomposition reactions of compound 1.

2.11. The Nature of Decomposition Intermediates Formed at 240 °C

Compound 1 gradually transforms into its decomposition intermediate between 200 and 300 °C (Figure 10). The peaks that belong to compound 1 can be seen in the curve of the I-240 sample (blue) with smaller intensity, together with new peaks of compounds. These newly formed compound peaks are also present in the I-300 sample, but with the peaks of the new products of consecutive decomposition reactions.
Figure 10. The powder X-ray diffractogram of compound 1 and decomposition intermediate 240 at 300 °C.

The composition of I-240 changed during storage in air for 3 days. The XRD of I-240 recorded as a function of time (Figure 11) showed that the peak intensities of silver perchlorate monohydrate increased, whereas the peak intensities belonging to the anhydrous polymorphs almost completely disappeared with increasing storage time. It means the changes in phase composition can be attributed to the hygroscopic nature of AgClO4; the transformation of the primary decomposition product is fast and decomposes easily into AgClO4 and AgCl even at 240 °C. However, there are some unassigned XRD peaks of the I-240 intermediate, which belong to a compound that contains coordinated ammonia (see the IR spectrum in Figure S10).
Figure 11. The powder X-ray diffractograms of the change of the composition of I-240.

There are important differences in the shape of the perchlorate $\nu_{as}(\text{ClO})$ bands, and band positions belong to the coordination-sensitive ammonia modes ($\delta(\text{NH})$ and $\rho(\text{NH})$) in the IR spectra of I-240 and compound 1 (Figure 12). Aqueous extraction of I-240 and the evaporation of the solution resulted in silver perchlorate monohydrate and an unknown phase. The IR (Figure 12) spectrum showed an important shift (60 cm$^{-1}$) of peak positions and the shape of $\nu_{as}(\text{Cl–O})$. The $\delta_s(\text{NH})$ and $\rho(\text{NH})$ bands can be assigned to a more strongly coordinated NH$_3$ ligand, as in compound 1 (shifting the $\delta(\text{NH})$ towards the higher wavenumber range [30]). The unknown phase may be a silver salt of a chlorine oxyacid, which contains fewer than 2 ammonia per silver (e.g., a polymeric monamine). This intermediate can decompose into the same compounds (N$_2$, H$_2$O, AgCl, AgClO$_4$) as compound 1. We attributed the change in the shape of the antisymmetric Cl–O IR band to the presence of two different chlorine oxyacid compounds in the intermediate I-240. One is a perchlorate, the other is a chlorate compound. For the chlorate component, the $\nu_{as}(\text{Cl–O})$ bands (AgClO$_3$ ($\nu_{as}(\text{ClO}) = 987,967$) (Figure 12)) are expected to appear in the lower wavenumber range of the wide perchlorate $\nu_{as}(\text{ClO})$ band. The presence of diamminesilver(I) chlorate cannot be expected because it melts and decomposes at 100 °C with the formation of AgClO$_3$[31].
Figure 12. The IR spectroscopic study of the decomposition of the I−240.

The relative amount of perchlorate, chlorate, and chloride ions in the samples I−240 and aq. extract of I−240 were found to be 8.25:3.35:1 and 8.02:3.01:1, respectively. The chlorite intermediate contents of the two samples are close to each other, and ca. 1/4 of the compounds containing chlorine proved to be chlorate ions on both sample surfaces. The silver XPS signals (368.3 and 368.5 eV) are the same for all three expected component types (AgCl, AgClO₃, and AgClO₄ or their complexes with ammonia). The oxygen signals of chlorate and perchlorate ions could easily be distinguished (531.2–531.9 eV: chlorate, and 532.8 eV: perchlorate) (Figure 13).
Summarizing the IR, XRD, and XPS results, the reduction process is not a single-step process, but probably different parallel and consecutive reactions take place with intermediates like AgClO, AgClO₂, and AgClO₃, or their NH₃ complexes. Taking into consideration the gaseous and solid reaction products, the relative amounts of these, the IR and XPS results of the intermediates, and the decomposition temperatures of AgClO, AgClO₂, AgClO₃, AgClO₄, and their known ammonia complexes, we suggest the following general decomposition scheme (Figure 14). The stable intermediate complex formed at the beginning of decomposition (180–240 °C) may only be a material containing chlorate.

The decomposition reaction temperature of compound 1 is >180 °C; thus, the reactions on the left side of the scheme (decomposition routes/disproportionation of AgClO and AgClO₂) may only be formed/decomposed in in situ processes. TG-MS shows the presence of O₂ (m/z = 32, O₂⁻) in the temperature range of the decomposition of compound 1, thus the evolution of oxygen below 235 °C (the melting point of AgClO₃) probably belongs to the decomposition of AgClO/AgClO₂ while above this temperature, it can only be a result of the decomposition of AgClO₃ into AgCl and O₂. The AgClO₃ detected by IR and XPS may form directly (Figure 14) by the reduction of perchlorate ion, or from AgClO₂ via disproportionation as well. The consecutive character of the first decomposition steps strongly suggests that the consecutive disproportionation steps (AgClO-AgClO₂-AgClO₃) are the main decomposition steps, and the decomposition with oxygen evolution, the redox reaction with HCl and Ag formation, and the non-redox thermal deammoniation of compound 1 are only by-reactions.
3. Materials and Methods

Deuton-X Ltd. supplied the chemical-grade compounds and analytical reagents (silver nitrate, sodium perchlorate monohydrate, 25% aq. ammonia solution, potassium carbonate, 70% aq. perchloric acid, ethanol, diethyl ether, 0.1 M standard NaOH and HCl, methyl red indicator).

All perchlorate complexes are potential explosives. Thus, they must be handled with great care.

Silver nitrate (3.4 g, 0.02 mol) was dissolved in 35 mL of cc. NH₃ solution, then 2.85 g (0.02 mol) sodium perchlorate monohydrate was added in 1.4 mL water. The mixture was left to evaporate, and the crystalline material was filtered off, then washed with a copious amount of cold water, ethanol, and diethyl ether. It was dried in air for 15 min in the dark and stored in a desiccator that contained CaO and a small amount of NH₄Cl. We attempted to prepare compound 1-H₂O in an analogous way, but by cooling the ammoniacal silver nitrate (prepared according to Mitscherlich [32]) and sodium perchlorate solutions to +5 °C.

Silver(I) perchlorate was prepared by dissolution of Ag₂CO₃ deposited in the reaction of silver(I) nitrate and potassium carbonate in perchloric acid. The solution of AgClO₄ was evaporated to dryness in a drying oven at 140 °C [33]. Silver perchlorate was dissolved in aq. NH₃ solution and compound 1 was crystallized in a fridge.

The elemental analysis for silver and ammonia content was performed with standard gravimetric (AgCl) and titrimetric (0.1 M NaOH solution added to form Ag₂O and liberate ammonia, and measurement of ammonia after distillation with 0.1 M HCl in the presence of methyl red) methods.

Powder XRDs were measured on a PW-1050 Bragg-Brentano (Philips, Eindhoven, Netherlands) parafocusing goniometer equipped with a copper tube. IR and far-IR spectra were recorded on a Bruker Alpha FT-IR spectrometer (Bruker, Ettlingen, Germany) and a Biorad-Digilab FT5-30-FIR spectrometer (Biorad, Austin, USA), respectively. Raman spectra and UV spectra were measured on a Horiba Jobin-Yvon LabRAM-type microspectrometer with external 532 nm Nd-YAG laser source (Horiba Jobin Yvon GmbH, Bensheim, Germany) and a Jasco V-670 UV–Vis spectrometer (JASCO Co., Oklahoma City, USA), respectively. TG-MS analysis was performed with a TA Instrument SDT Q600 thermal analyzer (TA Instrument, New Castle, USA) coupled to a Hiden Analytical HPR-20/QIC mass spectrometer (Hiden Co., United Kingdom, Warrington), whereas
DSC tests were performed with a Perkin Elmer DSC 7 instrument (Perkin-Elmer, Waltham, USA). Measurement details were the same as in our previous publications [25,34].

4. Conclusions

The reaction of ammoniacal AgNO₃ solution or [Ag(NH₃)₂]NO₃ with aq. NaClO₄ resulted in [Ag(NH₃)₂]ClO₄. We did not find a monohydrated form described previously, and the hydration of the diamminesilver(I) cation cannot be expected based on thermochemical considerations.

The temperature and enthalpy of phase change were determined by DSC, and found to be 225.7 K and 103.04 kJ/mol, respectively. Detailed spectroscopic (correlation analysis, IR, Raman, and UV) analyses were performed on [Ag(NH₃)₂]ClO₄, and all the vibrational modes in the IR and Raman spectra were assigned.

The thermal decomposition of [Ag(NH₃)₂]ClO₄ involves a solid-phase quasi-intramolecular redox reaction between the perchlorate anion and the ammonia ligand, resulting in lower valence chlorine oxyacid (chlorite, chlorate) components.

The formal loss of ammonia and the formation of AgClO₄ during heating is not a favored reaction route. The detected chlorate/chlorite intermediates disproportionate both in solid phase or in aqueous solutions (after dissolution of these decomposition intermediates in water) into AgCl and silver perchlorate.

The solid-phase AgCl-AgClO₄ mixture eutectically melts, and the resulting AgClO₄ decomposes in this melt into AgCl and O₂; thus, the final product of decomposition is AgCl, N₂, and H₂O.

The intermediate (chlorite, chlorate) phases were detected by IR, XPS, and titrimetric methods.

Supplementary Materials: The following are available online at www.mdpi.com/2304-6740/9/5/38/s1, Figure S1: Cryogenic DSC results of [Ag(NH₃)₂]ClO₄; Figure S2: (a) and (b) The hindered rotations (Rxyz(F1)) and translations (Txyz (F2)) of perchlorate ion in the polymorphs of [Ag(NH₃)₂]ClO₄; Figure S3(a) and (b): External vibrational modes (Tz(A1), Rz(A2) and Txz(E) or Rxz(E)) for ammonia in the polymorphs of [Ag(NH₃)₂]ClO₄; Figure S4: The far-IR spectra of compound 1-O; Figure S5: The UV spectra of compound 1; Figure S6: TG-DTG-DSC results of compound 1 in air; Figure S7: The powder X-ray diffractogram of P-480; Figure S8: TG-MS of compound 1 in inert atmosphere, formation of m/z = 14, m/z = 16, m/z = 17, m/z = 18; Figure S9: TG-MS of compound 1 in air atmosphere, formation of m/z = 44; Figure S10: IR spectrum of I-240 intermediate.

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