Cyclic monoterpenes trapped in a polyaromatic capsule: unusual selectivity, isomerization, and volatility suppression†

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Cyclic monoterpenes (CMTs) are intractable natural products with high volatility and strong odors so that there has been no molecular receptor capable of selectively and tightly trapping CMTs in both solution and the solid state. We herein report that a polyaromatic capsule acts as a functional nanoflask for CMTs with the following five features: (i) the capsule can selectively bind menthone from mixtures with other saturated CMTs in water. In contrast, (ii) treatment of the capsule with mixtures of menthone and π-conjugated CMTs gives rise to ternary host–guest complexes with high pair-selectivity. Notably, (iii) the encapsulated menthone displays unusual isomerization from a typical chair conformer to otherwise unstable conformers upon heating. (iv) The selective binding of volatilized CMTs is demonstrated by the capsule even in the solid state at atmospheric pressure. Furthermore, (v) the volatilities of CMTs are significantly suppressed at elevated temperatures by the capsule upon encapsulation in solution as well as in the solid state.

Introduction

Cyclic monoterpenes (CMTs) are simple yet important natural products, selectively biosynthesized from an acyclic monoterpene derivative in plants (Fig. 1a).1 The bioproducts are widely used as common fragrances, essences, and perfumes, due to their characteristic odors, and as key starting materials for drug synthesis.2,3 The relatively small structures (i.e., ~0.7 nm and ~200 Da) and the absence of strongly interactive substituents make CMTs highly volatile. Thus, these intractable features cause great difficulty in their molecular-level recognition and microscale separation from CMT mixtures, besides the existence of plenty of derivatives with only tiny structural differences.4 There have been many reports on molecular receptors (e.g., covalent tubular and bowl-shaped compounds,5 hydrogen-bonding and coordination capsules,6 and π-stacking capsules7) capable of binding CMTs such as camphor (CMP), menthol (MTL), and p-cymene (CMN; Fig. 1a). However, the majority of the previous receptors possess open and/or flexible cavities surrounded by aliphatic and/or small aromatic

†Electronic supplementary information (ESI) available. CCDC 2058897 and 2059097. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/d1sc01987b
frameworks\textsuperscript{5–8} so that the selective recognition and high volatility suppression of CMTs have been hardly accomplished by the artificial receptors so far. What type of molecular receptor is the most suitable for these bioproducts? We herein focused on polyaromatic capsule 1 (Fig. 1b and c) to strategically utilize multiple CH–π and hydrogen-bonding interactions with CMTs in the closed cavity (Fig. 1d).

Capsule 1, with a simple M\textsubscript{2}L\textsubscript{4} composition, provides a spherical cavity with relatively high rigidity and a diameter of \(~1\) nm (Fig. 1b),\textsuperscript{9} in contrast to biological cavities with high flexibility and complex shapes. In water, the polyaromatic cavity of 1 binds various biomolecules, e.g., relatively large androgens (\(~1.4\) nm in length) and medium-sized xanthines and disaccharides (\(~1.0\) nm in length), with a high selectivity through efficient host–guest interactions.\textsuperscript{10} Here we disclose that, unlike the previous receptors,\textsuperscript{5–8} capsule 1 acts as a functional nanoflask toward relatively small CMTs with the following novel characteristics. (i) The capsule binds menthone (MTO; Fig. 1a) with high efficiency and selectivity from mixtures with other saturated CMTs (e.g., MTL and CMP) in water at room temperature. The host–guest structure is fully confirmed by NMR, MS, ITC, and X-ray crystallographic analyses. On the other hand, (ii) MTO is pair-selectively bound by 1 with π-conjugated CMTs (i.e., CMP, TPN, and TBM) under the same conditions. In the confined cavity, (iii) unusual thermal isomerization of MTO from a typical diequatorial chair conformer to otherwise unstable diaxial chair and twist-boat conformers can be observed in a quantitative and reversible fashion. The cavity-directed, conformational stabilization of the resultant isomers is supported by theoretical calculations. (iv) The selective binding of volatilized MTO is also accomplished by solid 1 at ambient temperature and pressure. Finally, we reveal that (v) capsule 1 significantly suppresses the volatilities of CMTs upon encapsulation in solution as well as in the solid state. Combined with the crystal structures, the present behaviour demonstrates the presence of highly efficient CH–π and hydrogen-bonding interactions between the polyaromatic host cavity and the CMT guests.

Results and discussion
Selective trap of CMTs by the capsule in water
Capsule 1 bound various CMTs to give the corresponding 1 : 1 host–guest complexes in a quantitative fashion and displayed a distinct binding preference. When a mixture of (−)-menthone (MTO), \(p\)-menthane (MTA), and (−)-menthol (MTL; 10 equiv. each) was suspended in a \(1\) \(\text{D}_2\)O solution (0.5 mL) of 1 (0.26 \(\mu\)mol) at room temperature for 1 h, host–guest complexes 1·MTO, 1·MTA, and 1·MTL were formed in 90, 10, and 0% yields, respectively (Fig. 2a, left).\textsuperscript{11} In the \(^1\text{H}\) NMR spectrum, the three methyl and six methylene signals of encapsulated MTO were observed in the range of \(~2.65\) to \(~0.71\) ppm. These guest signals are outstandingly shifted upfield, due to the efficient shielding effect (\(\Delta\rho_{\text{max}} = \sim 4.09\) ppm) of the polyaromatic host (Fig. 2b and c). Selective formation of 1·MTO was further confirmed by the ESI-TOF MS spectrum, showing a series of prominent peaks derived from the [\(1\cdot\text{MTO}–n\cdot\text{NO}_3\)]\textsuperscript{−} species (\(n = 4\) to 2; Fig. S4†). A similar binding study clarified that 1 binds MTA stronger than MTL (in a 67 : 33 ratio; Fig. S17†).\textsuperscript{11} Therefore, the present competitive experiments elucidated the binding preference in the order of MTO \(\gg\) MTA > MTL. The binding constant (\(K_b\)) of 1 toward MTO was estimated to be 6.6 \(\times\) \(10^5\) \(\text{M}^{-1}\) by isothermal titration calorimetry (ITC) analysis (Fig. S20 and Table S1†).\textsuperscript{11} The obtained thermodynamics indicate that the guest binding is an enthalpically favorable process (\(\Delta H = \sim 84.5\) \(\text{kJ mol}^{-1}\), \(\Delta S = \sim 51.5\) \(\text{kJ mol}^{-1}\) (25 °C)), through multiple host–guest interactions.

Whereas bicyclic monoterpene such as (−)-camphor (CMP) and (−)-borneol (BNL) were also quantitatively bound by 1 in water (Fig. S9–S13†), their binding affinities were lower than those of the monocyclic ones. Treatment of a mixture of CMP and BNL (10 equiv. each) with 1 in \(\text{D}_2\)O gave rise to 1·CMP and 1·BNL in a 9 : 1 ratio (Fig. 2a, right and 2d).\textsuperscript{12} Since 1 bound MTL more efficiently (65% selectivity) than CMP from a 1 : 1 MTL/CMP mixture (Fig. S19†),\textsuperscript{12} the present competitive studies established the binding affinities of 1 for the tested CMTs to be MTO \(\gg\) MTA > MTL > CMP \(\gg\) BNL. This order stems not from the guest solubility in water (Fig. S23†) but from the identity of the guest shape and substituents. High recognition ability of 1
toward MTO most probably results from its planar structure (0.9 nm in the longest length) with three methyl groups and one carbonyl group (Fig. 1d). Eventually, the X-ray crystallographic analysis of 1·MTO indicated that the functional groups of MTO, adopting a typical diequatorial chair conformer, efficiently interact with the polyaromatic host cavity through seven CH–π and four CH–O=C hydrogen-bonding interactions (Fig. 2e).10,13 The number of the observed host–guest interactions within 1·MTO is larger than that within 1·CMP (i.e., one CH–π and four CH–O=C interactions; Fig. 2f), which is consistent with the ITC data in solution (∆H = −84.5 and −14.9 kJ mol⁻¹, respectively).11,12 The spherical structure of CMP (~0.6 nm in diameter) is too small to efficiently contact with the spherical cavity (~1.0 nm) of 1.

**Pairwise encapsulation of CMTs by the capsule**

Pairwise encapsulation of CMTs exceptionally took place, when 1 was combined with MTO and π-conjugated CMTs, such as p-cymene (CMN) and π-terpinene (TPN), in water.14 Although there is no electrostatic interaction between MTO and CMN, ternary host–guest complex 1·(MTO-CMN) was generated exclusively from a 1 : 1 mixture of MTO and CMN (10 equiv. each), being stirred in a D₂O solution of 1 (0.52 mM) at room temperature for 1 h (Fig. 3a, left). The ¹H NMR spectrum of the product displayed new signals derived from the host and the two guests (Fig. 3c), which are completely different from those of 1·MTO and 1·(CMN)₂, prepared separately (Fig. 3b and d, respectively). The formation of the ternary complex was indicated by ESI-TOF MS analysis, where the MS peaks corresponding to [1·(MTO-CMN)–n·NO₃⁻]⁺ (n = 4 to 2) species were observed at m/z = 976.6, 1322.4, and 2014.6 (Fig. 3e).14 The 1 : 1 ratio of guests MTO and CMN within 1 was further confirmed by ¹H NMR analysis after extraction of the encapsulated guests with CDCl₃. Notably, the selective formation of 1·(MTO-CMN) (~98%) was found even using a 1 : 5 : 50 mixture of 1, MTO, and CMN in water at room temperature (Fig. S28†).

The observed pair-selectivity is ascribed to structure-based complementary guest uptake. The optimized structure of 1·(MTO-CMN) suggested that the size and shape of a stacked heterodimer of planar MTO (diequatorial chair form) and CMN nicely fit those of the polyaromatic cavity of 1 (Fig. 3f).14 The π-conjugated framework of CMN is essential for the pair-selective encapsulation with MTO so that 1·(MTO-TPN) also formed in high selectivity (83%) by the treatment of 1 with MTO and TPN, bearing a cyclohexadiene ring, under the same conditions (Fig. 3a, right). In contrast, the formation of a ternary complex was undetected from a mixture of 1, MTO, and TPA under various conditions (Fig. S15f).

**Thermal isomerization of menthone within the capsule**

In the confined cavity of 1, a unique conformation change of MTO was observed from a typical chair form to otherwise labile forms by thermal stimuli. In the ¹H NMR spectra of 1·MTO in D₂O, unexpectedly, the aromatic host signals and the aliphatic guest signals were fully shifted upon heating at 100 °C for 8 h, without the appearance of empty host signals (Fig. 4a and b, left). Particularly, the downfield shift of the inner protons (H₁) of 1 and its splitting into three signals suggested the generation of three new species derived from MTO in the cavity.19 The IR analysis of 1·MTO showed the C==O stretching band of the encapsulated MTO being also shifted by 10 cm⁻¹ through the heating (Fig. S36†). On the other hand, the ESI-TOF MS peaks of the host–guest complex were unchanged before and after the treatment (Fig. 4a and b, right). The new host and guest NMR signals gradually went back to the initial positions at 30 °C (t₁/₂ = 22 h; Fig. S37d†).11,17 The signal shifts could be repeated at least three times without the decomposition of the host–guest complex (Fig. S37e†). These findings indicate the thermal isomerization of MTO in the cavity of 1 in water.

Theoretical calculation studies suggested the formation of diaxial-type chair isomer ACI and twist-boat isomers TBI-a (major) and TBI-b (minor) upon heating and their large conformational stabilization within 1 (Fig. 4c).11,18 Whereas the energy of ACI is 13.6 kJ mol⁻¹ higher than that of MTO (diequatorial-type chair form) without 1 (Fig. 4c), the energy difference (ΔE) between 1·ACI and 1·MTO decreased to be 5.7 kJ mol⁻¹ (Fig. 4d). Accordingly, the diaxial chair conformer was stabilized by 7.9 kJ mol⁻¹ upon encapsulation, as compared with the diequatorial chair conformer. Conformers TBI-a and TBI-b are also stabilized by 10.0 and 6.0 kJ mol⁻¹ within 1, respectively. In the optimized structures, roughly bowl-shaped ACI and TBI are suitably accommodated in the spherical cavity of 1 (Fig. 4d), displaying multiple host–guest CH–π and hydrogen-bonding interactions (Fig. S40b†). To the best of our knowledge, there has been no report on the spectroscopic observation of the diaxial to diequatorial19 and chair to twist-
boat isomerizations of CMTs as well as other cyclohexane derivatives in solution.

Selective trap of volatilized CMTs by the capsule solid

Inspired by the solution-state binding ability of 1 toward CMTs, we next examined the binding of volatilized CMTs using solid 1. To our surprise, a solid of 1 could efficiently and selectively bind MTO from a mixture with volatilized CMTs at room temperature and atmospheric pressure. Each of MTO, MTA, and MTL (100 equiv.) was placed in a closed glass vessel (50 mL) including 1 (0.26 \( \mu \)mol) without direct host-guest contact (Fig. 5a). After standing for 14 h, the resultant solid was exposed to reduced pressure (480 Pa, 1 h) to remove CMTs outside the solid. The guest binding and its selectivity were elucidated by \(^1\)H NMR analysis in D\(_2\)O. The NMR spectrum indicated that 1 binds the guests quantitatively to yield host-guest complexes 1-MTO, 1-MTL, and 1-MTA in 70, 30, and 0%, respectively (Fig. 5b). Solid 1 also bound volatilized MTO and CMN pair-selectively to give 1-(MTO-CMN) in 95% yield after standing for 2 d (Fig. 5c).

It is noteworthy that each of the tested CMTs was quantitatively encapsulated by 1 in water, as described above, whereas solid 1 displayed largely different affinity toward each of the volatilized CMTs, e.g., quantitative binding to MTO yet weak binding to CMP (4%) under the same conditions (Fig. S45†). The rigid and sterically demanding features of CMP are major reasons for the selectivity.

Volatility suppression of CMTs by the capsule

Notably, the volatility of CMTs was largely suppressed upon encapsulation by 1 in both solution and the solid state. When a mixture of 1-CMP and CMP (≈12 equiv.) in D\(_2\)O was heated at 80 °C, the majority of free CMP (96%) was volatilized after 25 h. On the other hand, CMP within 1 remained fully intact under the same conditions, as confirmed by time-dependent \(^1\)H NMR and ESI-TOF MS analyses (Fig. 6a and b). No peaks derived from empty 1 were detected in the MS spectrum. On the basis of the NMR studies, the kinetic analysis of the guest volatilization revealed that the half-lives (\(t_{1/2}\)) of free and encapsulated CMP are approximately 5 and >690 h, respectively (Fig. 6c), indicating an encapsulation-induced >140 times suppression of the thermal volatility. In a similar way, the volatilization of MTL was suppressed by a factor of >70 upon encapsulation by 1 (Fig. S47†).

Moreover, encapsulated CMTs (i.e., MTL and MTO) in the cavity of solid 1 hardly volatilized even under high vacuum at elevated temperature (i.e., 480 Pa and 80 °C) for 1 h, as confirmed by \(^1\)H NMR analysis (Fig. 6d and e). In contrast, free MTL and MTO volatilized instantly (<4 min) under the same conditions. Thermal gravimetric (TG) analysis of solid 1-MTO showed 4.4% weight loss in the range of 88–200°C, assignable to the release of the encapsulated MTO (Fig. S49†). The present results demonstrated that 1 acts as an excellent molecular
storage for volatile CMTs in solution as well as in the solid state.\(^{24}\)

**Conclusions**

We have revealed the new functions of a polyaromatic capsule acting as a nanoflask for cyclic monoterpenes (CMTs) both in solution and in the solid state. Although the recognition of CMTs is highly challenging due to their simple structure and high volatility, the capsule preferentially bound menthone from a mixture of CMTs in water and in the gas phase under ambient conditions. Unusual pair-selective encapsulations of two different CMTs (e.g., menthone and \(p\)-cymene) were also accomplished by the capsule. Moreover, the capsule prompted the thermal isomerization of menthone and largely suppressed the volatilities of CMTs in the confined cavity. This level of control over CMTs has not been accomplished with previous molecular receptors\(^3\)\(^–\)\(^5\) and solid materials.\(^{25}\) We hope that the present work will open the door to new host–guest chemistry of polyaromatic nanoflasks\(^8\) toward not only natural but also synthetic volatile compounds.

**Author contributions**

R. S., Y. T., K. N., S. T., and M. Y. designed the work, carried out research, analyzed data, and wrote the paper. Y. S. contributed to crystallographic analysis. M. Y. is the principal investigator. All authors discussed the results and commented on the manuscript.

**Conflicts of interest**

There are no conflicts to declare.

**Acknowledgements**

This work was supported by JSPS KAKENHI (Grant No. JP18H01990/JP19H04566) and “Support for Tokyo Tech Advanced Researchers (STAR)”. The computations were performed using computers in Research Center for Computational Science, Okazaki, Japan.

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MS condition and thus MS peaks derived from 1-MTO
were observed at m/z = 942.8 and 1278.0 (Fig. 3e and
S29†). Theoretical studies of the host–guest complexes
indicated that the formation of 1-MTO-CMN is more
favorable (25 kJ mol⁻¹) than that of 1-MTO or 1-(CMN)₂
in the gas phase (Fig. S31†).
16 The 1H NMR, 1H–1H COSY, and HSQC analyses of MTO
after the thermal treatment indicated the formation of
three isomers of MTO, besides the diequatorial chair form
(Fig. S35a–c†).
17 The activation energy (ΔG‡) of the thermal isomerization of
MTO within 1 was estimated to be 92.8 kJ mol⁻¹ at 303 K,
by the Eyring plots (Fig. S37†).
18 The conformational isomers of MTO (20 structures) were
initially generated through automatic conformational analysis
using Conflex 8, Rev. C program and subsequently
their optimized structures and energies were obtained by
DFT calculations (CAM-B3LYP/6-31G(d) level of theory,
gas phase; Fig. S40a†). Furthermore, the optimized structures
and energies were obtained from the host–guest complexes,
using 1 (R = –OCH₃) and the most stable four
conformers of MTO, by DFT calculations (CAM-B3LYP/
LanL2DZ (for Pt), 6-31G(d) (for C, H, N, O) level of theory,