Insight into the adsorption of a liquid organic hydrogen carrier, perhydro-\(i\)-dibenzyltoluene (\(i = m, o, p\), on Pt, Pd and PtPd planar surfaces†

Cecil Naphtaly Moro Ouma, * Phillimon Modisha and Dmitri Bessarabov

Liquid organic hydrogen carriers (LOHCs) are considered to be safe and efficient hydrogen storage media with high hydrogen storage capacities. Adsorption of the LOHC perhydro-\(i\)-dibenzyltoluene (\(i = meta (m), ortho (o), para (p)) isomers on (100), (110) and (111) planar surfaces of Pd, Pt and a 50 : 50 PtPd alloy were investigated, using density functional theory with van der Waals corrections. The calculated heats of formation of the isomers indicated that all the isomers considered were energetically stable. Surface selectivity to isomer adsorption was investigated, using isomer adsorption preference and energies. The (110) surface was found to be highly preferred by the different isomers, compared with both the (100) and the (111) surfaces. Among the isomers, isomer–surface attachment occurred most often in the case of perhydro-\(m\)-dibenzyltoluene and perhydro-\(o\)-dibenzyltoluene adsorption. The LOHC isomer adsorption on different surfaces was found to be spontaneous, energetically stable and exothermic, with high isomer adsorption preference for Pt and PtPd surfaces, compared with Pd surfaces. This indicates the ease of loading of the LOHC on Pt and PtPd surfaces, for subsequent dehydrogenation.

Introduction

In ongoing efforts towards the goal of achieving sustainable carbon-free energy solutions, hydrogen and its associated technologies are attracting significant attention.\(^1-5\) This is largely because sustainable energy solutions, more often than not, carry the following shortcomings. Generation: some renewable energy sources (RES) are intermittent.\(^6-10\) Cost: the initial cost of setting up RES is high.\(^11-13\) Storage: the most commonly used energy storage medium is batteries; however, with the increasing demand, especially for electric vehicles, mineral sources (such as those of Li, which plays an integral role in batteries) are continually being depleted, hence leading to increasing production costs.\(^1,11,12,15-17\) Transportation: transportation is closely associated with issues of storage; safe transportation requires safe energy storage technologies.\(^17,11,15,18\) Integration: the current energy infrastructure is grid-based and an integration of RES to the grid, so as to utilize the excess energies generated, is challenging.\(^19-22\) Toxicity: some of the elemental materials used, or by-products, are toxic.\(^16,23-25\)

Because of these challenges, there is great need for a sustainable energy solution that will not suffer from the shortcomings mentioned (cost, storage, transportation, integration and toxicity). Hydrogen and its associated technologies is one such solution.\(^12\) This is because hydrogen can be generated from both renewable energy sources and fossil fuels,\(^15,17,26-29\) it is a clean energy carrier,\(^14,23,30\) and it can be used for both stationary and mobile power applications (e.g., H\(_2\) fuel cell vehicles).\(^9,23,30-32\)

However, despite marked improvements in both hydrogen utilization and production, the application of hydrogen-based technologies in the current market landscape is hampered by its storage and delivery.\(^1,30\) Some of the proposed hydrogen storage solutions include physisorption of hydrogen in porous material surfaces, chemisorption of hydrogen on metal and metal alloy surfaces and the use of compressed hydrogen and cryogenic hydrogen.\(^30\) However, these proposed solutions still present challenges.

For onboard vehicular applications, liquid organic hydrogen carriers (LOHCs) are proposed to be a safe and efficient way to store hydrogen because of the reversible hydrogenation and dehydrogenation processes that can be activated using a catalyst.\(^19,30,31,38\) Hydrogen storage in these carriers takes place when the hydrogen covalently binds onto liquid organic compounds that are unsaturated/aromatic (hydrogen lean) in the presence of a catalyst until it is fully saturated/aliphatic (hydrogen rich). This process is referred to as hydrogenation of the liquid carrier. It is then possible to safely transport the hydrogen-rich liquid organic compound. The hydrogen molecule will then be released from the hydrogen-saturated liquid compound, also in the presence of catalyst, resulting in a hydrogen-lean liquid compound that can be hydrogenated once again. This process is

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referred to as dehydrogenation. It is this operating principle of LOHC systems that makes possible its integration in the existing fuel delivery infrastructure.\textsuperscript{34,39,41–42} Several materials have been proposed as possible LOHC candidates. Each has its own hydrogen storage capacity and operates at different (de)hydrogenation temperatures, over different catalysts.\textsuperscript{19,21} In this study, the fully hydrogenated (hydrogen-rich) LOHC, perhydro-dibenzyltoluene is investigated.\textsuperscript{38,41} Perhydro-dibenzyltoluene (chemical formula C\textsubscript{21}H\textsubscript{38}) is marketed by Sasol Performance Chemicals (South Africa) under the name Marlohc. Its derivative, Marlotherm SH isomer con

FIG. 1. Top view and (b) side view of the relaxed atomic structures of perhydro-i-dibenzytoluene (i = m, o, p) derivatives. (The blue and grey balls indicate C and H atoms, respectively.)
Computational details

Electronic structure calculations based on plane-augmented basis sets within the density functional theory formalism as implemented in the Quantum ESPRESSO code were used in this study. The exchange-correlation approximation was approximated using the generalized gradient approximation as parameterized by Perdew–Burke–Ernzerhof (PBE). Ultrasoft pseudopotential approximation was used to describe the valence electrons of Pd, Pt, C and H, and the electronic configurations of the elements were Kr 4d10, [Xe] 4f14 5d9 6s1, [He] 2s2 2p2 and 1s1 for Pd, Pt, C and H, respectively. To account for the weak molecule–surface interactions, dispersion corrections were done using the vdW-DF2-B86R (rev-vdw-df2) scheme of van der Waals interactions implemented in the Quantum ESPRESSO code.

Bulk unit cells of Pd, Pt and PtPd were first optimized to obtain their respective equilibrium properties. This was done using a kinetic energy cut-off of 30 Ry. The Brillouin zone (BZ) was sampled using a 10×10×10 k-point mesh with a 10−6 Ry and 10−5 Ry per Å convergence criterion imposed on the calculated total energies and forces, respectively. The obtained equilibrium lattice constants for Pd, Pt and PtPd were 3.95 Å, 3.98 Å and 3.98 Å, respectively. The different perhydro-i-dibenzyltoluene (i = o, m, p) isomer configurations were then relaxed using the cluster approach by embedding them in a 20×20×20 Å simulation box. For these systems, all the atomic positions were allowed to move. The only difference between these calculations and those of the bulk unit cells of Pd, Pt and PtPd was that Γ-sampling was used to sample the BZ. The relaxed perhydro-i-dibenzyltoluene (i = o, m, p) isomer configurations are presented in Fig. 1.

From the obtained equilibrium Pd, Pt and PtPd bulk structures, the (100), (110) and (111) surfaces consisting of three layers and 20 Å vacuum distances were constructed with only the bottom layer being fixed. The adsorption surface area (length × width) was chosen based on the full length of the relaxed adsorbate. This was done to ensure that the adsorbate is not strained by the surface size, its orientation on the surface, nor its interaction with its periodic image. The length and width of the surface was therefore calculated as $L_{\text{surf}} = L_{\text{ads}} + 4$ Å and $W_{\text{surf}} = W_{\text{ads}} + 4$ Å, respectively, where $L_{\text{ads}}$ and $W_{\text{ads}}$ are both equal to the longest isomer’s horizontal length among the relaxed perhydro-i-dibenzyltoluene (i = o, m, p) isomer configurations. A further 4 Å was added to the adsorbate length (width) to possibly prevent adsorbate–adsorbate interaction with the periodic images of adsorbates. It should be noted, however, that the role of short-range and long-range adsorbate–adsorbate interaction was not part of this study, hence was not tested; the 4 Å addition is only an estimate.

Nine surfaces were considered: each species (Pd, Pt and PtPd) had three planar surfaces ((100), (110) and (111)). Six isomer configurations were considered. See Fig. 1. Therefore,
the adsorption of each of the isomers on each of the surfaces would require 54 calculations. Furthermore, as Fig. 2 shows, it is possible for each of the isomers to assume several adsorption configurations/orientations on the surface.

Thus, to account for all possible orientations on the different surfaces, a computationally less expensive approach was adopted. This was done using the adsorption locator module within Materials Studio. The adsorption locator module uses molecular dynamics (MD; with universal force fields) as the calculation engine of adsorption energies for all the generated surface-adsorbate configurations. The minimum energy surface-adsorbate configurations of a surface-adsorbate system are generated via a Monte Carlo search of the configurational space of the surface-adsorbate system, which includes different adsorbate rotations and orientations. At the end of the search, a minimum energy configuration as well as the preferred adsorbate adsorption site are obtained.

To validate this approach, of selecting the minimum energy configuration obtained using MD so as to determine whether it is consistent with vdW-DFT, a test case was carried out using ten different configurations generated using the adsorption locator module. These configurations were then used as inputs for vdW-DFT calculations. Atomic relaxations were done on the different configurations and adsorption energies were calculated. The calculated adsorption energies were then compared to those obtained using MD results. There was consistency between the MD and the vdW-DFT results in the prediction of the minimum energy configuration. See Fig. S1 in the ESL. This observation is similar to that reported for the adsorption of natural organic matter on nanoparticle surfaces. Thus, for each isomer, only the configurations where the isomer attached to the surface (as predicted using the adsorption locator) were considered as inputs for subsequent DFT calculations. Configurations where the isomer did not attach to the surface were characterized by large values of the surface-adsorbate distance, $d_o$ (see Fig. 2).

**Results and discussion**

**Heats of formation**

After relaxing the atomic positions of the perhydro-$i$-dibenzyltoluene ($i = o, m, p$) isomer configurations, the heats of formation of the respective isomer configurations were calculated as follows:

$$E_i = E_{C_iH_{10}} - (21E_C + 38E_H) \ldots (1)$$

where $E_{C_iH_{10}}$ is the calculated total energy of the isomer, and $E_C$ and $E_H$ are the chemical potentials of C and H, respectively. This was done to confirm whether all the isomer configurations considered in this study would indeed form. The calculated heats of formation were negative for all the isomer configurations (Table 1), which indicated that all the proposed perhydro-$i$-dibenzyltoluene ($i = o, m, p$) isomer configurations would form. This observation is consistent with experimental observations, reported by Do et al. Perhydro-dibenzyltoluene is known to consist as a mixture of different isomers.

**Surface-isomer configurations**

As already mentioned, only the configurations where the isomer attached to the surface were considered in subsequent DFT calculations. Table 2 shows that isomer attachment was only observed in 21 cases of the 54 surface isomer configuration searched. The (110) surface was highly preferred for perhydro-$i$-dibenzyltoluene ($i = o, m, p$) adsorption, with the exception of 03 perhydro-$o$-dibenzyltoluene and 01 perhydro-$p$-dibenzyltoluene on Pd (110) surfaces. The least preferred isomer adsorption surface was the (111) surface; isomer attachment occurred only in the case of 01 perhydro-$m$-dibenzyltoluene on Pd (111) and PtPd (111) surfaces, and 01 perhydro-$o$-dibenzyltoluene and 02 perhydro-$o$-dibenzyltoluene on Pd (111) and PtPd (111), respectively. The deduced order of surface preference by the different isomers was the (110) surface, followed by the (100) surface and then the (111) surface. Perhydro-$o$-dibenzyltoluene only attached to the PtPd (111) surface. Furthermore, the Pt and Pt-derived (PtPd) surfaces were preferred by the different perhydro-$i$-dibenzyltoluene ($i = o, m, p$) isomers to the Pd surface.

This indicated that Pt doping/alloying might be essential in activating Pd surfaces for isomer adsorption.

| Table 1 | Heats of formation $E_f$ of the different ionomers determined using DFT calculations with van der Waals corrections |
|---------|---------------------------------------------------------------------------------------------------------------|
| Isomer configuration | $E_f$ (eV) | $E_f$ (kJ mol⁻¹) |
| 01 Perhydro-$m$-dibenzyltoluene | -8.01 | -772.39 |
| 02 Perhydro-$m$-dibenzyltoluene | -8.03 | -774.52 |
| 01 Perhydro-$o$-dibenzyltoluene | -8.03 | -774.96 |
| 02 Perhydro-$o$-dibenzyltoluene | -8.02 | -773.61 |
| 03 Perhydro-$o$-dibenzyltoluene | -7.95 | -767.01 |
| 01 Perhydro-$p$-dibenzyltoluene | -8.12 | -783.29 |

| Table 2 | Isomer attachment to Pt, Pt and PtPd surfaces (x indicates that the isomer attached to the particular surface) |
|---------|---------------------------------------------------------------------------------------------------------------|
| Isomer configuration | Species | Surface 100 | 110 | 111 |
| 01 Perhydro-$m$-dibenzyltoluene | Pd | x | x | x |
| 02 Perhydro-$m$-dibenzyltoluene | Pt | x | | | |
| 01 Perhydro-$o$-dibenzyltoluene | Pd | x | | | |
| 02 Perhydro-$o$-dibenzyltoluene | Pt | | | |
| 03 Perhydro-$o$-dibenzyltoluene | PtPd | x | | | |
| 01 Perhydro-$p$-dibenzyltoluene | Pd | x | | | |
| 02 Perhydro-$p$-dibenzyltoluene | Pt | x | | | |

**Results and discussion**

**Heats of formation**

After relaxing the atomic positions of the perhydro-$i$-dibenzyltoluene ($i = o, m, p$) isomer configurations, the heats of formation of the respective isomer configurations were calculated as follows:

$$E_i = E_{C_iH_{10}} - (21E_C + 38E_H) \ldots (1)$$

where $E_{C_iH_{10}}$ is the calculated total energy of the isomer, and $E_C$ and $E_H$ are the chemical potentials of C and H, respectively. This was done to confirm whether all the isomer configurations considered in this study would indeed form. The calculated heats of formation were negative for all the isomer configurations (Table 1), which indicated that all the proposed perhydro-$i$-dibenzyltoluene ($i = o, m, p$) isomer configurations would form. This observation is consistent with experimental observations, reported by Do et al. Perhydro-dibenzyltoluene is known to consist as a mixture of different isomers.
Adsorption geometry relaxations

Using DFT calculations with van der Waals corrections, the 21 configurations (above) were then relaxed by allowing all atoms to move, with the exception of the atoms on the bottom layer of the surfaces. The relaxed surface–isomer configurations for the (100), (110) and (111) surfaces are shown in Fig. 3–6. In the case of isomer adsorption on (100) surfaces, the toluene lies horizontal to the surface, except in the cases of 01 perhydro-o-dibenzyltoluene and 02 perhydro-m-dibenzyltoluene adsorption on the Pt (100) surface. The toluene also lies horizontally on the surface in the case of isomer adsorption on (111) surfaces. In the case of (110) surfaces, the following were the exceptions: 01

![Fig. 3](image)

Obtained minimum energy for surface isomer configurations on (100) surfaces. (m: perhydro-m-dibenzyltoluene; o: perhydro-o-dibenzyltoluene.) — here and elsewhere.

![Fig. 4](image)

Obtained minimum energy for surface isomer configurations on (110) surfaces.
perhydro-ortho-dibenzyltoluene on Pt (110) and PtPd (110) surfaces, 02 perhydro-ortho-dibenzyltoluene on the Pd (110) surface and 03 perhydro-ortho-dibenzyltoluene on Pt (110) and PtPd (110) surfaces.

The calculated equilibrium distances ($d_0$) between the isomers and the different surfaces are tabulated in Table 3. For a particular isomer adsorption, the calculated $d_0$ values revealed that Pd had the smallest $d_0$ and Pt the largest. Some of the H atoms on the two benzene rings of the perhydro-ortho-dibenzyltoluene ($i = o, m, p$) isomers were found to be close to the metal surfaces and the change in C–H bond length for these H atoms was found to be increased by about 5%. This indicated that it would be easy to break these C–H bonds and liberate the H from these benzene rings than the toluene ring. This is consistent with experimental observations. A similar observation on C–H bond breaking and H liberation has also been reported in the case of dodecahydro-N-ethylcarbazole adsorption and dehydrogenation on Pd surfaces.

**Adsorption energy**

Furthermore, since in hydrogen evolution reactions (HERs), the adsorption free energy of a molecule, $\Delta G_M$, has been used as a descriptor for HER activity, it is evident that $E_{ads}$ can then also be used to estimate the HER activity of a surface. The following applies to the $G_M$:

$$\Delta G_M = E_{ads} + \Delta E_{ZPE} + T\Delta S_M$$  \hspace{1cm} (2)

where $\Delta E_{ZPE}$ is the zero point energy difference between the molecule M in the adsorbed and gas phase, $\Delta S_M$ is the change in entropy between the molecule M in the adsorbed and gas phase, and $T$ is the temperature in K. In the case of hydrogen, $E_{ads} \equiv 0.24 \text{ eV}$.\textsuperscript{57,71,72} Pt is considered as the optimal catalyst for HER because its $\Delta G_H \equiv 0 \text{ eV}$, and hence thermoneutral. Assuming Pt to be the best catalyst for isomer adsorption in this study, the energy difference from Pt values was then estimated as follows:
Table 3  Equilibrium distance ($d_0$) and adsorption energies ($E_{ads}$) of perhydro-$i$-dibenzyltoluene ($i = o, m, p$) on (100) and (111) surfaces determined using van der Waals calculations

| Isomer configuration                                      | Species | Surface | $d_0$ (Å) | $E_{ads}$ (eV) | $E_{ads}$ (kJ mol$^{-1}$) |
|-----------------------------------------------------------|---------|---------|-----------|----------------|----------------------------|
| 01 Perhydro-m-dibenzyltoluene                             | Pd      | 110     | 2.00      | −1.90          | −183.06                    |
|                                                           | Pt      | 100     | 2.28      | −1.62          | −156.05                    |
| 02 Perhydro-m-dibenzyltoluene                             | Pt      | 100     | 1.98      | −1.72          | −166.43                    |
|                                                           | PtPd    | 100     | 1.96      | −1.64          | −158.61                    |
| 01 Perhydro-o-dibenzyltoluene                             | Pt      | 100     | 2.18      | −1.55          | −149.75                    |
| 01 Perhydro-m-dibenzyltoluene                             | Pd      | 110     | 1.49      | −2.45          | −236.52                    |
| 02 Perhydro-m-dibenzyltoluene                             | Pt      | 110     | 1.75      | −2.89          | −279.21                    |
|                                                           | PtPd    | 110     | 1.59      | −2.29          | −221.07                    |
| 01 Perhydro-p-dibenzyltoluene                             | PtPd    | 110     | 1.73      | −2.69          | −259.98                    |
| 01 Perhydro-m-dibenzyltoluene                             | PtPd    | 110     | 1.42      | −2.45          | −236.40                    |
| 01 Perhydro-o-dibenzyltoluene                             | Pd      | 110     | 1.44      | −2.37          | −228.22                    |
| 02 Perhydro-o-dibenzyltoluene                             | Pt      | 110     | 1.61      | −2.20          | −212.62                    |
| 03 Perhydro-o-dibenzyltoluene                             | Pt      | 110     | 1.50      | −2.22          | −214.05                    |
| 01 Perhydro-p-dibenzyltoluene                             | PtPd    | 110     | 1.69      | −2.63          | −253.86                    |
| 01 Perhydro-m-dibenzyltoluene                             | PtPd    | 110     | 1.53      | −2.57          | −248.45                    |
| 01 Perhydro-o-dibenzyltoluene                             | Pt      | 110     | 1.53      | −2.23          | −214.84                    |
| 01 Perhydro-o-dibenzyltoluene                             | Pt      | 111     | 1.90      | −2.08          | −201.04                    |
| 01 Perhydro-o-dibenzyltoluene                             | PtPd    | 111     | 2.15      | −1.86          | −179.38                    |
| 03 Perhydro-o-dibenzyltoluene                             | Pt      | 111     | 1.79      | −2.02          | −194.98                    |

$$\Delta E_{TN} = E_{ads(PtPd)}^{Pd} - E_{ads}^{Pt}$$  

(3)

where $E_{ads(PtPd)}^{Pd}$ and $E_{ads}^{Pt}$ are the calculated isomer adsorption energies (obtained using eqn (2)) on Pd or PtPd and Pt surfaces, respectively. Table 4 shows that $\Delta E_{TN} < \pm 0.1$ eV in the case of isomer adsorption on PtPd surfaces and $\Delta E_{TN} > \pm 0.1$ eV in the case of isomer adsorption on Pd surfaces. Thus, alloying Pd and Pt is likely to improve the activity of Pd towards isomer adsorption and probably also the catalytic activity.

The catalytic activity of Pt and Pd on alumina supports has been investigated for the dehydrogenation of perhydro-dibenzyltoluene. It was discovered, experimentally, that 5 wt% Pt/Al$_2$O$_3$ has high activity (40% degree of dehydrogenation) compared with 5 wt% Pd/Al$_2$O$_3$ (8% degree of dehydrogenation). A compound chemically similar to perhydro-dibenzyltoluene, cyclohexane, showed similar behaviour, with even lower activity when bimetallic Pt–Pd was used for dehydrogenation. In the dehydrogenation of decalin to naphthalene, DFT calculations revealed the following. The first dehydrogenation step (decalin to tetralin) is energetically more favoured on Pt surfaces than on Pd surfaces, while the reverse applies for the second dehydrogenation step (tetralin to naphthalene). Using the estimate obtained using eqn (3), the activity of the Pt (110) surface towards dehydrogenation is more energetically favourable in the case of 01 perhydro-m-dibenzyltoluene, 02 perhydro-m-dibenzyltoluene and 03 perhydro-o-dibenzyltoluene than either the Pd (110) or PtPd (110) surfaces. This observation is similar to what is reported elsewhere in literature. In the case of 01 perhydro-o-dibenzyltoluene and 02 perhydro-o-dibenzyltoluene, dehydrogenation will be more energetically favoured on the Pd (110) and PtPd (110) surfaces that on the Pt (110) surface.

Table 4  Activity estimation using $\Delta E_{TN}$

| Isomer configuration                                      | Species | Surface | $\Delta E_{TN}$ (eV) |
|-----------------------------------------------------------|---------|---------|----------------------|
| 02 Perhydro-m-dibenzyltoluene                             | Pt      | 100     | 0.00                 |
|                                                           | PtPd    | 100     | 0.08                 |
| 01 Perhydro-m-dibenzyltoluene                             | Pd      | 110     | 0.44                 |
|                                                           | Pt      | 110     | 0.00                 |
|                                                           | PtPd    | 110     | 0.60                 |
| 01 Perhydro-o-dibenzyltoluene                             | Pt      | 110     | 0.00                 |
|                                                           | PtPd    | 110     | −0.12                |
| 02 Perhydro-o-dibenzyltoluene                             | Pd      | 110     | −0.16                |
|                                                           | Pt      | 110     | 0.00                 |
|                                                           | PtPd    | 110     | −0.01                |
| 03 Perhydro-o-dibenzyltoluene                             | Pt      | 110     | 0.00                 |
|                                                           | PtPd    | 110     | 0.06                 |

Charge analysis

Charge analysis was carried out in efforts to better understand how the charges are distributed in the isomer–adsorbate systems. Charge density difference was determined as follows:

$$\rho_{CDD} = \rho_{IS} - (\rho_{I} + \rho_{S})$$  

(4)

where the charge density distribution of the isomer ($\rho_{I}$) and the pristine surface ($\rho_{S}$) were subtracted from the charge density distribution of the isomer–surface system ($\rho_{IS}$). The obtained charge density difference ($\rho_{CDD}$) at ±0.002 iso-surface levels are shown in Fig. 7–9. In the figures, yellow regions indicate regions where $\rho_{CDD}$ is positive while the light blue regions indicate regions where $\rho_{CDD}$ in negative. It is evident that charge...
accumulation occurs at the interface of the surface and the isomer, with negative charge accumulation largely concentrated on H atoms and C–H covalent bonds for the H atoms close to the surfaces. This indicated that there will be charge transfer between the isomer and the surface.

Conclusions
Surface and isomer selectivity was investigated using DFT calculations with van der Waals corrections. Different isomer configurations of perhydro-i-dibenzyltoluene ($i = a$, $m$, $p$) on
planar ((100), (110) and (111)) surfaces of Pd, Pt and PtPd were investigated. Although all the perhydro-i-dibenzyltoluene ($i = o, m, p$) isomers and their associated configurations were found to be stable, adsorption of the perhydro-p-dibenzyltoluene isomer and its configurations on the different surfaces was found to be least likely; perhydro-i-dibenzyltoluene ($i = o, m$) was much more likely to be adsorbed on the surfaces under consideration. Isomer adsorption on the (110) Pd, Pt and PtPd surfaces was highly preferred, followed by the (100) surfaces and the (111) surfaces. This indicates that, among the surfaces considered in this study, the best surface for catalytic reaction will be the (110) surfaces of Pd, Pt and PtPd. Pt surfaces had the highest adsorbate binding energies among the individual perhydro-i-dibenzyltoluene ($i = o, m, p$) isomer configurations, followed by the PtPd alloy and then, finally, the Pd surfaces.

Conflicts of interest

There are no conflicts to declare.

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References

1 D. Viesi, L. Crema and M. Testi, *Int. J. Hydrogen Energy*, 2017, **42**, 27354–27373.
2 D. Bessarabov, G. Human, A. J. Kruger, S. Chiuta, P. M. Modisha, S. P. du Preez, S. P. Oclofe, I. Vincent, J. Van Der Merwe, H. W. Langmi and others, *Int. J. Hydrogen Energy*, 2017, **42**, 13568–13588.
3 D. Geburtig, P. Preuster, A. Bößmann, K. Müller and P. Wasserscheid, *Int. J. Hydrogen Energy*, 2016, **41**, 1010–1017.
4 R. Dittmeyer, T. Boeltken, P. Piernartini, M. Selinske, M. Loewert, F. Dallmann, H. Kreuder, M. Cholewa, A. Wunsch, M. Belimov, S. Farsi and P. Pfeifer, *Curr. Opin. Chem. Eng.*, 2017, **17**, 108–125.
5 M. Schwarz, P. Bachmann, T. N. Silva, S. Mohr, M. Scheuremeyer, F. Späth, U. Bauer, F. Düll, J. Steinhauser, C. Hohner, T. Döpper, H. Noei, A. Stierle, C. Papp, H. P. Steinrück, P. Wasserscheid, A. Görling and J. Libuda, *Chem.–Eur. J.*, 2017, **23**, 14806–14818.
6 M. Eypasch, M. Schimpe, A. Kanwar, T. Hartmann, S. Herzog, T. Frank and T. Hamacher, *Appl. Energy*, 2017, **185**, 320–330.
7 O. Elishav, G. Tvil, B. Mosevitsky, D. Lewin, G. E. Shter and G. S. Grader, *Energy Procedia*, 2017, **135**, 3–13.
8 H. Kreuder, T. Boeltken, M. Cholewa, J. Meier, P. Pfeifer and R. Dittmeyer, *Int. J. Hydrogen Energy*, 2016, **41**, 12082–12092.
9 B. Wang, T. yan Chang, Z. Jiang, J. jia Wei, Y. hai Zhang, S. Yang and T. Fang, *Int. J. Hydrogen Energy*, 2018, **43**, 7317–7325.
10 E. Gianotti, M. M. Taillades-Jacquin, J. Rizi, D. J. Jones, J. Rozie, D. J. Jones, I. Charles, G. Montpellier and U. M. R. Cnrs, *ACS Catal.*, 2018, **8**, 4660–4680.
11 A. Lahnaoui, C. Wulf, H. Heinrichs and D. Dalmazzone, *Appl. Energy*, 2018, **223**, 317–328.
12 W. Peters, A. Seidel, S. Herzog, A. Bößmann, W. Schwieger and P. Wasserscheid, *Energy Environ. Sci.*, 2015, **8**, 3013–3021.
13 Y. Dong, M. Yang, P. Mei, C. Li and L. Li, *Int. J. Hydrogen Energy*, 2016, **41**, 8498–8505.
14 H. Wang, X. Zhou and M. Ouyang, *Int. J. Hydrogen Energy*, 2016, **41**, 18062–18071.
15 M. Aziz, *Energy Procedia*, 2017, **105**, 1495–1500.
16 S. Fei, B. Han, L. Li, P. Mei, T. Zhu, M. Yang and H. Cheng, *Int. J. Hydrogen Energy*, 2017, **42**, 25942–25950.
17 L. A. Prananto, M. K. Biddinika and M. Aziz, *Energy Procedia*, 2017, **142**, 1603–1608.
18 V. Milella, A. Koelpin and E. Schluecker, *Energy Technol.*, 2018, 547–557.
19 M. Amend, S. Schernich, M. Sobota, I. Nikiforidis, W. Hieringer, D. Assenbaum, C. Gleichweit, H. J. Drescher, C. Papp, H. P. Steinrück, A. Görling, P. Wasserscheid, M. Laurin and J. Libuda, *Chem.–Eur. J.*, 2013, **19**, 10854–10865.
20 M. H. Rehman, M. Reisslein, A. Rachedi, M. Erol-Kantarci and M. Radenkovic, *IEEE Trans. Industr. Inform.*, 2018, 1.
21 M. Liserre, T. Sauter and J. Hung, *IEEE Ind. Electron. Mag.*, 2010, **4**, 18–39.
22 R. Loisel, M. Sanchez-Angulo, F. Schoefs and A. Gaillard, *Renewable Energy*, 2018, **126**, 36–48.
23 Q.-L. Zhu and Q. Xu, *Energy Environ. Sci.*, 2015, **8**, 478–512.
24 A. N. Kalenchuk, V. I. Bogdan, S. F. Dunaev and L. M. Kustov, *Fuel Process. Technol.*, 2018, **169**, 94–100.
25 N. A. A. Rusman and M. Dahari, *Int. J. Hydrogen Energy*, 2016, **41**, 12108–12126.
26 M. Aziz, F. B. Juangsa, W. Kurniawan and B. A. Budiman, *Energy*, 2016, **116**, 489–497.
27 M. Aziz, *Int. J. Hydrogen Energy*, 2016, **41**, 104–112.
28 M. Aziz, I. N. Zaini, T. Oda, A. Morihara and T. Kashiwagi, *Int. J. Hydrogen Energy*, 2017, **42**, 2904–2913.
29 A. M. Abdalla, S. Hossain, O. B. Nisfendi, A. T. Azad, M. Dawood and A. K. Azad, *Energy Convers. Manage.*, 2018, **165**, 602–627.
30 Z. Chen, M. Yang, T. Zhu, Z. Zhang, X. Chen, Z. Liu, Y. Dong, G. Cheng and H. Cheng, *Int. J. Hydrogen Energy*, 2018, **1–9**, DOI: 10.1016/j.ijhydene.2018.03.088.
31 C. Chilev and F. D. Lamari, *Int. J. Hydrogen Energy*, 2016, 41, 1744–1758.
32 M. Yang, Y. Sun and H. Cheng, *Int. J. Hydrogen Energy*, 2017, 42, 28647–28648.
33 L. Li, M. Yang, Y. Dong, P. Mei and H. Cheng, *Int. J. Hydrogen Energy*, 2016, 41, 16129–16134.
34 J. Von Wild, T. Friedrich, A. Cooper, B. Toseland, G. Muraro, W. TeGrotenhuis, Y. Wang, P. Humble and A. Karim, in *World Hydrogen Energy Conference*, Essen, 2010.
35 A. Haupt and K. Müller, *Energy*, 2017, 118, 1123–1130.
36 K. Müller, R. Aslam, A. Fischer, K. Stark, P. Wasserscheid and W. Arlt, *Int. J. Hydrogen Energy*, 2016, 41, 22097–22103.
37 R. Aslam, M. Mineeva, K. Müller and W. Arlt, *Sep. Purif. Technol.*, 2016, 163, 140–144.
38 P. M. Modisha, J. H. L. Jordaan, A. Bößmann, P. Wasserscheid and D. Bessarabov, *Int. J. Hydrogen Energy*, 2018, 43, 5620–5636.
39 G. Do, P. Preuster, R. Aslam, A. Bößmann, K. Müller, W. Arlt and P. Wasserscheid, *React. Chem. Eng.*, 2016, 1, 313–320.
40 I. Y. Choi, B. S. Shin, S. K. Kwak, K. S. Kang, C. W. Yoon and J. W. Kang, *Int. J. Hydrogen Energy*, 2016, 41, 9367–9373.
41 C. Krieger, K. Müller and W. Arlt, *Sol. Energy*, 2016, 123, 40–50.
42 J. Haro, R. Gómez and J. M. Ferreira, *J. Catal.*, 1976, 45, 326–331.
43 K. Müller, K. Stark, V. N. Emelyanenko, M. A. Varfolomeev, D. H. Zaitsau, E. Shoifet, C. Schick, S. P. Verevkin and W. Arlt, *Ind. Eng. Chem. Res.*, 2015, 54, 7967–7976.
44 P. Inhetveen, N. S. A. Alt and E. Schlücker, *Vib. Spectrosc.*, 2016, 83, 85–93.
45 A. Fikrt, R. Brehmer, V. O. Milella, K. Müller, A. Bößmann, P. Preuster, N. Alt, E. Schlücker, P. Wasserscheid and W. Arlt, *Appl. Energy*, 2017, 194, 1–8.
46 M. Amende, A. Kaftan, P. Bachmann, R. Brehmer, P. Preuster, M. Koch, P. Wasserscheid and J. Libuda, *Appl. Surf. Sci.*, 2016, 360, 671–683.
47 A. Mehranfar and M. Izadyar, *Int. J. Hydrogen Energy*, 2017, 42, 9966–9977.
48 A. Mehranfar, M. Izadyar and A. A. Esmaeili, *Int. J. Hydrogen Energy*, 2015, 40, 5797–5806.
49 Y. Dong, M. Yang, T. Zhu, X. Chen, C. Li, H. Ke and H. Cheng, *Energy Technol.*, 2018, 6, 558–562.
50 M. Yang, G. Cheng, D. Xie, T. Zhu, Y. Dong, H. Ke and H. Cheng, *Int. J. Hydrogen Energy*, 2018, 43, 8868–8876.
51 F. Sotoodeh and K. J. Smith, *J. Phys. Chem. C*, 2013, 117, 194–204.
52 F. Sotoodeh and K. J. Smith, *J. Catal.*, 2011, 279, 36–47.
53 A. N. Kay Lup, F. Abnisa, W. M. A. Wan Daud and M. K. Aroua, *J. Ind. Eng. Chem.*, 2017, 56, 1–34.
54 S. Sui, X. Wang, X. Zhou, Y. Su, S. Riffat and C. Liu, *J. Mater. Chem. A*, 2017, 5, 1808–1825.
55 M. P. Andersson, T. Bligaard, A. Kustov, K. E. Larsen, J. Grieve, T. Johansson, C. H. Christensen and J. K. Norskov, *J. Catal.*, 2006, 239, 501–506.
56 P. Frontera, A. Macario, M. Ferraro and P. Antonucci, *Catalysts*, 2017, 7, 59.
57 Z. W. Seh, J. Kibsgaard, C. F. Dickens, I. Chorkendorff, J. K. Norskov and T. F. Jaramillo, *Science*, 2017, 355, eaad4998.
58 L. G. Devi, N. Kottam, B. N. Murthy and S. G. Kumar, *J. Mol. Catal. A: Chem.*, 2010, 328, 44–52.
59 R. Michalsky, Y.-J. Zhang and A. A. Peterson, *ACS Catal.*, 2014, 4, 1274–1278.
60 W.-C. Liu, W. T. Ralston, G. Melaet and G. A. Somorjai, *Appl. Catal., A*, 2017, 545, 17–23.
61 E. Antolini, *Energies*, 2017, 10(1), 42.
62 C. Wen, A. Yin and W.-L. Dai, *Appl. Catal., B*, 2014, 160, 730–741.
63 K. Izumi, S. Miyazaki, S. Yoshida, T. Mizokawa and E. Hanamura, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2007, 76, 75111.
64 K. O. Obodo, C. N. M. Ouma, J. T. Obodo and M. Braun, *Phys. Chem. Chem. Phys.*, 2017, 19(29), 19050–19057.
65 X. Shen, Y. Pan, B. Liu, J. Yang, J. Zeng and Z. Peng, *Phys. Chem. Chem. Phys.*, 2017, 19, 12628–12632.
66 J. K. Norskov, T. Bligaard, A. Logadottir, J. R. Kitchin, J. G. Chen, S. Pandelov and U. Stimming, *J. Electrochem. Soc.*, 2005, 152, J23.
67 J. K. Norskov, T. Bligaard, J. Rossmeisl and C. H. Christensen, *Nat. Chem.*, 2009, 1, 37–46.
68 J. D. Benck, T. R. Hellstern, J. Kibsgaard, P. Chakthranont and T. F. Jaramillo, *ACS Catal.*, 2014, 4, 3957–3971.
69 W.-K. Chen, S.-H. Liu, M.-J. Cao, Q.-G. Yan and C.-H. Lu, *J. Mol. Struct.: THEOCHEM*, 2006, 770, 87–91.
70 F. S. Tautz, *Prog. Surf. Sci.*, 2007, 82, 479–520.
71 C. Tsai, K. Chan, J. K. Norskov and F. Abild-Pedersen, *Catal. Sci. Technol.*, 2015, 5, 246–253.
72 S. H. Mir, S. Chakraborty, J. Wärnå, S. Narayan, P. C. Jha, P. K. Jha and R. Ahuja, *Catal. Sci. Technol.*, 2017, 7, 687–692.
73 P. Giannozzi, S. Baroni, N. Bonini, M. Calandra, R. Car, C. Cavazzoni, D. Ceresoli, G. L. Chiarotti, M. Cococcioni, I. Dabo and others, *J. Phys.: Condens. Matter*, 2009, 21, 395502.
74 M. Ernzerhof and G. E. Scuseria, *J. Chem. Phys.*, 1999, 110, 5029.
75 I. Hamada, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2014, 89, 121103.
76 2015 Dassault Systèmes BIOVIA, *BIOVIA Material Studio 2016 [16.1.0.21]*, San Diego, Dassault Systèmes, 2015.
77 N. N. Nyangiwe, C. N. Ouma and N. Musee, *Mater. Chem. Phys.*, 2017, 200, 270–279.
78 K. Müller, K. Stark, V. N. Emel’yanenko, M. A. Varfolomeev, D. H. Zaitsau, E. Shoifet, C. Schick, S. P. Verevkin and W. Arlt, *Ind. Eng. Chem. Res.*, 2015, 54, 7967–7976.
79 N. Brückner, K. Obesser, A. Bößmann, D. Teichmann, W. Arlt, J. Dungs and P. Wasserscheid, *ChemSusChem*, 2014, 7, 229–235.