Photocatalytic activity of ZnO/GaP\textsubscript{1−x}N\textsubscript{x} for water splitting

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The phosphidation of a ZnO/GaN solid solution photocatalyst enhanced significantly its activity for water splitting. The photocatalysts were heated with phosphorus in a vacuum-sealed quartz tube. Activation due to phosphidation was restricted within the narrow temperature range of 823–873 K, and varying amounts of P were added to the solid solution. In addition to X-ray diffraction peaks due to ZnO/GaN, active phosphide ZnO/GaN provided a single GaP peak with diffraction angles higher than normal GaP by 2θ = 0.20–0.44°, indicative of the formation of a GaP\textsubscript{1−x}N\textsubscript{x} alloy system. The diffraction peaks were simulated using first principles Ab initio calculations on molecular models of GaP\textsubscript{1−x}N\textsubscript{x}. The comparison with experimental shifts showed that the highest activity was induced in an x range of 0.034–0.074 of GaP\textsubscript{1−x}N\textsubscript{x}, which the transition of photon-induced electrons from the valence to the conduction band is indirect. Many reports have shown that incorporating small amounts of N into GaP induces a change in the fundamental optical transition from an indirect-gap (Γ\textsubscript{15} − X\textsubscript{1}) transition of GaP\textsubscript{0.98}N\textsubscript{0.02} to a direct-gap transition and lowering the band gap with increasing N concentration.\textsuperscript{11–27} The N-induced decrease in the band gap is a result of an interaction between the nitrogen atoms and the GaP matrix, which forms a GaP\textsubscript{1−x}N\textsubscript{x} alloy system.\textsuperscript{16–19,22} These results indicate that GaP\textsubscript{1−x}N\textsubscript{x} compounds, or at least GaP\textsubscript{1−x}N\textsubscript{x} layer formation on either GaN or GaP, efficiently produce visible-light driven photoexcited carriers, and thus it is of interest to apply the GaP\textsubscript{1−x}N\textsubscript{x} system as a photocatalyst for water splitting. Previously, GaP\textsubscript{1−x}N\textsubscript{x} (x = 0–0.035) was grown using a metal organic chemical vapor deposition (MOCVD) technique on GaP and then employed as an electrode for water-splitting reactions. The band edges were shown to be too negative to drive water-splitting, although the GaP\textsubscript{0.98}N\textsubscript{0.02}/GaP had a band gap of ~2 eV optimal for water splitting applications.\textsuperscript{28} However, photocatalysis of GaP\textsubscript{1−x}N\textsubscript{x} for water splitting is still ambiguous, because of its electronic properties sensitive to P and N compositions, and thus it has to be examined from the viewpoint of the introduction of a small amount of nitrogen atoms that is a key to obtain the active GaP\textsubscript{1−x}N\textsubscript{x} alloy system.

There are two ways to prepare Ga–N–P-compounds. The nitridation of GaP has mostly been used. Another strategy, which is scarcely examined, is phosphidating the GaN surface. The present study describes the results of GaP nitridation and GaN phosphidation, which revealed that the latter using ZnO/GaN produced highly active photocatalysts for the water splitting reaction. The photocatalytically active GaP\textsubscript{1−x}N\textsubscript{x} structures...
are discussed based on cluster simulation for a new GaP phase produced by phosphidation.

**Experimental section**

A solid solution of ZnO/GaN, that is (Ga$_{x}$Zn$_{1-x}$)(N$_{x}$O$_{1-x}$), was prepared by nitriding metal oxides under a NH$_3$ flow according to a method reported previously.$^{1,4}$ The starting material was a powder containing 1 : 2 molar ratio of Ga$_2$O$_3$ (99.98% pure reagent from Nacalai Chemical Co. Ltd) and ZnO (extra grade from Nacalai Chemical Co. Ltd), which was mixed in an agate mortar and placed in a quartz tube. The quartz tube was set in a rotary kiln furnace, where the powder was slowly rotated to undergo uniform nitridation with gaseous NH$_3$. The flow rate of NH$_3$ was 300–400 mL min$^{-1}$. The nitridation temperature and time were 1123 K and 14 h, respectively. For phosphidation, the ZnO/GaN prepared was mixed with red phosphorus (99.9%; Aldrich) in a quartz tube, sealed under vacuum, and then heated at temperatures from 773 to 1073 K. The phosphidation time was 7–28 h. The ratio of P to ZnO/GaN in the sealed quartz tube is an important factor for phosphidation, therefore the molar ratio of P to ZnO/GaN was adjusted from 0 to 0.10. The ZnO/GaN prepared by phosphidation is referred to as ZnO/GaN-P ($T, R$) for simplicity, where $T$ and $R$ represent the temperature and molar ratio of P to ZnO/GaN upon phosphidation, respectively. For example, ZnO/GaN-P (823, 0.07) indicates phosphidation at 823 K using a molar ratio of 0.07.

For the preparation of GaP$_{1-x}$N$_x$ photocatalysts, GaP (99.999%; Kojundo Chemicals Co. Ltd) was nitrided for 3–16 h between 873 and 1123 K at an NH$_3$ flow rate of 300 mL min$^{-1}$. In addition, ZnO was added at a molar ratio of 0.5–5% to GaP and then subjected to a similar nitridation process.

For the loading of RuO$_2$ as a cocatalyst, the nitrides and phosphides prepared were dispersed in THF containing the Ru carbonyl complex Ru$_3$(CO)$_{12}$ at room temperature, dried under vacuum, and then oxidized at 623 K for 1.5 h to convert the complex to RuO$_2$ particles.$^{1,2}$ The loading of the cocatalyst was 3.5 wt% based on the amount of Ru metal. A 3.5 wt% RuO$_2$-loaded ZnO/GaN photocatalyst was denoted as RuO$_2$-ZnO/GaN, unless otherwise specified.

Water splitting was conducted in a closed gas-circulating apparatus equipped with a Pyrex glass reactor and gas chromatograph for analyzing H$_2$ and O$_2$ gas, as described elsewhere.$^{1,2}$ The photocatalyst powder (0.2 g) was dispersed in ion-exchanged water using a magnetic stirrer followed by irradiation of incident light intensity (50 mW cm$^{-2}$) using a 300 W Xe lamp (Eagle Engineering LX300) from the outside. The reactor was cooled to 293 K by water circulating through a water jacket during the photocatalytic reaction.

In the simulation of XRD peak patterns, we have used a supercell of 2 $\times$ 4 $\times$ 1 for a 64 atom structure with the composition Ga$_{32}$P$_{32}$. The model was chosen to give a symmetric substitution for the dopant nitrogen. The central strand was taken to minimize the interaction of the dopant arising from periodic boundary consideration. All the calculations were performed using the CASTEP (Cambridge Serial Total Energy Package) and the associated programs for symmetry analysis, which has been described elsewhere.$^{29-31}$ CASTEP is a pseudopotential total energy code that employs Perdew and Zunger$^{23}$ parameterization of the exchange-correlation energy, super cells and special point integration over the Brillouin zone and a plane wave basis set for the expansion of wave functions. Becke-Perdew parameterization of the exchange-correlation functional, which includes gradient correction (GGA)$^{33,34}$ was employed. The pseudopotentials are constructed from the CASTEP database. Reciprocal space integration over the Brillouin zone is approximated through careful sampling at a finite number of $k$ points using the Monkhorst-Pack scheme.$^{35}$ Here we used symmetric 5 $k$-points in all calculations. The basis cutoff used is 300.0 eV. The energy tolerance is $1.0 \times 10^{-5}$ eV per atom, the force tolerance 0.3 eV nm$^{-1}$, maximum stress 0.05 GPa and the displacement tolerance 0.0001 nm.

**Results**

The X-ray diffraction patterns of GaP nitried between 873 and 1123 K consisted of strong GaP peaks, and no significant peak shifts were observed (ESI, Fig. S1†). Very small peaks attributed to GaN appeared. For RuO$_2$-loading on the nitried GaP, there was no significant production of H$_2$ and O$_2$, although a very small amount of H$_2$ production occurred. In nitridation of a ZnO and GaP mixture, X-ray diffraction patterns were nearly the same as those of GaP, and slight photocatalytic activity for H$_2$ production was found, but no oxygen evolution was observed. The procedure for obtaining GaP$_{1-x}$N$_x$ by nitriding GaP and a mixture of ZnO and GaP failed to produce active photocatalysts.

Fig. 1 shows results of water splitting using RuO$_2$-ZnO/GaN and RuO$_2$-ZnO/GaN-P (823, 0.07), respectively. No gas-phase products were observed for either photocatalyst without light irradiation. For RuO$_2$-ZnO/GaN, light irradiation resulted in the immediate production of both H$_2$ and O$_2$, which was nearly proportional to the irradiation time. Re-running the reaction after evacuating the gas-phase products resulted in a similar production of H$_2$ and O$_2$. Even after the fourth run, gas production occurred steadily. The production ratio of H$_2$ to O$_2$ was 2.1 : 1. For RuO$_2$-ZnO/GaN-P (823, 0.07), both H$_2$ and O$_2$ evolved immediately upon light irradiation, with the initial photocatalytic activity nearly 8-fold greater than that of original RuO$_2$-ZnO/GaN. Upon repeating the photocatalytic reaction, production gradually was reduced, and became constant after the third repeat. Interestingly, this production was still nearly 6-fold greater than that using RuO$_2$-ZnO/GaN. The production ratio of H$_2$ to O$_2$ was 2.2 : 1.0.

The effects of phosphidation on the catalytic activity of RuO$_2$-ZnO/GaN were examined in the temperature range of 823–1073 K at different molar ratios of ZnO/GaN to P. The photocatalytic activity of RuO$_2$-ZnO/GaN-P (823, x) with $x = 0–0.1$ was too low to detect. Fig. 2 shows the results using RuO$_2$-ZnO/GaN-P (823, x) with $x = 0–0.1$. The activity increased with the value of $x$, reached a maximum at $x = 0.07$, and then decreased slightly. In RuO$_2$-ZnO/GaN-P (873, x) at $x = 0–0.1$, the activity was enhanced as the value of $x$ increased, passed through a maximum at $x = 0.05$, and then decreased significantly. The maximum activity frequently occurred when phosphidation was
conducted between 823 and 873 K, although the P ratio at the maximum was different. The results using RuO$_2$–ZnO/GaN-P (1073, $x$) with $x = 0$–0.05 showed that the activity increased at $x = 0.005$, followed by a monotonous decrease with increasing $x$. It is to be noted that no activation occurred when ZnO/GaN was heated under similar conditions in a vacuum-sealed quartz tube without phosphorus.

Fig. 3 shows the effects of time on ZnO/GaN upon phosphidation at 823 K. Phosphidation for 7 h resulted in little activation, with activity nearly the same as the original ZnO/GaN. A sharp increase in activity occurred upon phosphidation for 14 h, followed by the appearance of a maximum at 21 h, and then a slight decrease for 28 h. Note that the phosphidation of ZnO/GaN led to the formation of photocatalysts with significantly greater activity than the original ZnO/GaN when RuO$_2$ was loaded as the cocatalyst. The phosphidation temperatures suitable for high activity generation remained in the narrow range of 823–873 K. The SEM observation showed that there were no significant changes in the morphology of ZnO/GaN particles and their aggregates by phosphidation (ESI Fig. S2†). Thus, the activity enhancement is not directly associated with crystal growth of specific planes but the effect of N and P atom combination.

Fig. 4 shows the UV-Vis diffuse reflectance spectrum of ZnO/GaN-P (823, 0.07), together with ZnO/GaN. Light absorption started at around 500 nm, then increased sharply and reached a maximum at approximately 350 nm. A shift to a shorter wavelength by ~7 nm was observed in comparison with ZnO/GaN.
peak position from the original GaP. These results indicate that phosphidation producing GaP takes place in surface regions, and the interior remains as GaN. Fig. 6 shows the X-ray diffraction patterns of the $2\theta = 28–29^\circ$ region for ZnO/GaN-P (873, $x$) as a function of $x$. Below $x = 0.01$, no significant diffraction peaks were observed in this region. A single peak appeared at $2\theta = 28.50^\circ$ for $x = 0.05$, which shifted to $2\theta = 28.68^\circ$ for $x = 0.07$. For $x = 0.1$, two peaks appeared at $2\theta = 28.32^\circ$ and $2\theta = 28.68^\circ$. The peak shifts of GaP to a higher angle indicates that the N atoms are involved in a fundamental unit of the tetrahedral GaP$_4$, producing GaP$_1.x$.N$_x$, because the N$_3^-$ size is smaller than that of P$_3^-$: Therefore, the shift of the GaP diffraction peak represents a change in the active site conformation induced by P introduction.

Fig. 7 shows the X-ray diffraction patterns of ZnO/GaN-P (823, 0.07) at different phosphidation times. A single peak appeared at $2\theta = 28.62^\circ$ upon heating for 7 h. The peak increased in size at 14 h, and shifted slightly to a lower diffraction angle by 0.07° at 21 h. A very slight shift to a lower diffraction angle occurred with 28 h of phosphidation. To
visualize the active structures, the photocatalytic activity obtained under various conditions such as different phosphidation temperatures, time and P amounts was consolidated and plotted against diffraction peak positions due to GaP in the range \(2\theta = 28^\circ - 29^\circ\). As shown in Fig. 8, the photocatalytic activity at 823 K increased as the diffraction angle increased from \(2\theta = 28.32^\circ\) (GaP crystal), producing a maximum at approximately \(2\theta = 28.6^\circ\), and decreases with further increase in the diffraction angle. A similar trend can be seen for 873 K phosphidation.

These correlations demonstrate that a high photocatalytic activity depends on the structures of GaP\(_{1-x}\)N\(_x\) formed through the gas–solid reaction between vaporized P and GaN. To determine the active conformation, the diffraction peak shifts were simulated by taking the geometry optimized model of the Ga\(_n\)P\(_{32}\)N\(_y\) into account. Fig. 9 shows the GaP diffraction peaks calculated by the powder diffraction method as implemented in the Reflex Module of the Dassault Systemes BIOVIA method for \(n = 32\) and \(y = 1, 3, 4, 6, 7\) and 8. A \(2\theta\) range of 10 to 90° was used with a step size of 0.05. Peak position errors can result from an incorrect zero point or from errors in the placement or transparency of the sample. In this case we have used Bragg–Brentano correction.\(^{36}\) A peak due to Ga\(_{32}\)P\(_{32}\) appeared at \(2\theta = 27.65^\circ\), shifted to a higher angle with increasing N atom content, and attained at \(2\theta = 28.44^\circ\) for \(y = 4\), \(2\theta = 28.75^\circ\) for \(y = 6\) and \(2\theta = 29.03^\circ\) for \(y = 8\). The diffraction angle shifts against the percentage of P atoms in the total of P and N atoms \(P/(P + N) \times 100\%\) showed a good linear relationship (ESI Fig. S3†). In comparison with the experimental value which provides the greatest photocatalytic activity for shifts by \(2\theta = 0.20\)–0.44° to a higher diffraction angle from GaP (Fig. 8), the linear relationship provided the result that the optimal conditions for the highest activity are the involvement of N at the percentage of 3.4–7.4% in GaP.

Fig. 8 Photocatalytic activity vs. diffraction peak angles of GaP produced by phosphidation. The diffraction peak angle of normal GaP is shown for comparison. The photocatalytic activity for H\(_2\) production only was shown for simplicity. Numbers in figure shows the molar percentage of P added to ZnO/GaN and the time for phosphidation. No expression of time means 14 h phosphidation. Open and filled triangles: ZnO/GaN phosphidated at 823 K. Open rhombus: at 873 K. Open square: at 1073 K. Filled circle: normal GaP.

Fig. 9 Simulated X-ray diffraction peaks of GaP for Ga\(_{32}\)P\(_{32}\)N\(_y\) (\(y = 1–8\)).
around $-14$ eV, P 3s AO at around $-10$ eV, and the valence band, in which the lower energy part is mainly due to Ga 4s + P 3p + N 2p orbitals. The HOMO is composed of Ga 4p + P 3p + N 2p AOs. In the unoccupied bands, LUMO consists of Ga 4s4p + P 3p + N 2s AOs. The band gap is estimated to be 0.793 eV.

**Discussion**

The phosphidation of ZnO/GaN increased the photocatalytic activity for water splitting by a factor of 6–8. The quantum efficiency was estimated to be in the range 1–2% from comparison with photocatalytic activity previously obtained for RuO$_2$–ZnO/GaN without phosphidation. The phosphidation conditions are restricted in a very narrow range in temperatures and the P amounts introduced, which indicates that the activation is very sensitive to the ratio of P to N. The appearance of band structures of hybridized Ga 4s4p + P 3p + N 2s AOs, which forms intermediate electronic levels. The valence bands of ZnO/GaN-P consist of the hybridized Ga 4s4p + P 3p + N 2s AOs, which forms intermediate electronic levels. The valence bands of ZnO/GaN-P consist of the hybridized Ga 4s4p + P 3p + N 2s AOs, which forms intermediate electronic levels.

The phosphidation process occurred, and also the diffusion of GaP crystal appeared, as shown in Fig. 6. Thus, the photocatalytic activity against GaP peak shifted progressed, larger shifts of the GaP diffraction peak occurred, and also the diffraction peak assigned to the original GaP crystal appeared, as shown in Fig. 6. Thus, the photocatalytic activity that decreases with larger diffraction peak shifts are due to not only the dilution of N atom concentrations but the formation of inactive GaP phases.

The UV-Vis diffuse reflectance spectrum showed that the light absorbing wavelength of ZnO/GaN-P was slightly (~7 nm) shorter than that of ZnO/GaN, opposite to those reported for the formation of the GaP$_{1-x}$N$_x$ phase, in which the N atom inclusion in GaP caused shifts to a longer wavelength. The present ZnO/GaN system had characteristics of absorbing longer wavelength light than GaN, because of the solid solution effects of ZnO. It seems likely that the removal of the ZnO component from ZnO/GaN surfaces during phosphidation in vacuum has larger effects than the phosphidation, which results in the light absorption edge shifts to shorter wavelengths.

The band edges of GaP$_{1-x}$N$_x$/GaP(x = 0–0.035) were previously reported to be negative by more than 300 mV. This might be associated with the structures of thin films grown in mismatch between the GaP$_{1-x}$N$_x$ film and substrate GaP. The flat band potentials of GaP at pH = 0 are in the range of 1.2–0.93 V, and thus the band edge for the oxidation of water is negative or very slightly positive, which makes it difficult for water splitting to proceed efficiently. On the other hand, ZnO/GaN has a valence band level considerably lower than the water oxidation level. The valence bands of ZnO/GaN-P consist of the hybridized Ga 4p + P 3p + N 2p AOs, which forms intermediate electronic band structures of GaP and GaN. Thus, the N atom inclusion to GaP lowers the valence band edge enough for the oxidation to proceed. This has positive effects on the oxidation of water, although the conduction band edge shifts to a lower energy level. The disadvantage of not only GaP but possibly many metal...
phosphides is the valence band edge being higher than or nearly the same as the oxidation level of water. Schematic band structures and energy levels are schematically depicted in Fig. 11. The presented results indicate that the consolidation of metal phosphides with N atoms is able to extend the availability of meal phosphides to photocatalysts for water splitting.

In conclusion, the phosphidation of active ZnO/GaN photocatalysts significantly enhances the photocatalytic activity for water splitting. The phosphidation of ZnO/GaN successfully permits to achieve optimal nitrogen conditions in GaP$_{1-x}$N$_x$, which leads to the generation of high photocatalytic performance. The present approach shows promise for application to other metal nitride systems that will be advantageous for efficient photocatalysts for water splitting.

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