Transformation of the coordination nanostructures of 4,4′,4″-(1,3,5-triazine-2,4,6-triyl) tribenzoic acid molecules on HOPG triggered by the change in the concentration of metal ions†

Sihao Li, Caimei Gong, Yuyang Zhang, Shizhang Fu, Zhongping Wang, Yan Lu, Siyi Gu, Xiaoqing Liu* and Li Wang*†

The modulation effects of Cu2+/Fe3+ ions on the hydrogen-bonded structure of 4,4′,4″-(1,3,5-triazine-2,4,6-triyl) tribenzoic acid (TATB) on a HOPG surface have been investigated at the liquid–solid interface by scanning tunneling microscopy (STM). STM observations directly demonstrated that the self-assembled honeycomb network of TATB has been dramatically transformed after introducing CuCl2/FeCl3 with different concentrations. The metal–organic coordination structures are formed due to the incorporation of the Cu2+/Fe3+ ions. Interestingly, a Cu2+ ion remains coordinated to two COOH groups and only the number of COOH groups involved in coordination doubles when the concentration of Cu2+ ions doubled. A Fe3+ ion changes from coordination to three COOH groups to two COOH groups after increasing the concentration of Fe3+ ions in a mixed solution. Such results suggest that the self-assembled structures of TATB molecules formed by metal–ligand coordination bonds can be effectively adjusted by regulating the concentration of metal ions in a mixed solution, which has rarely been reported before. It explains that the regulatory effect of concentration leads to the diversity of molecular architectures dominated by coordination bonds.

Introduction

Metal–organic coordination structures are materials comprising reticular metal centers and organic linkers in which the two constituents bind with each other via metal–ligand coordination interactions. The coordination-driven molecular self-assembly of nanostructures on solid surfaces has attracted much attention owing to the potential applications of such structures in nanodevices and nanopatterning in the future.1–4 Recently, coordination interaction has been used in the direct preparation of novel nanostructures ranging from clusters,5,6 1D chains or polymers,7–10 to 2D coordination frameworks,11–14 and even 3D structures.15–17 Moreover, their composition, morphology and properties have been well proved, and have opened up specific and targeted application in the areas of chemical and materials sciences.18,19 However, few studies have focused on regulating the coordination structures which involves changing the metal species for coordination, and more importantly, the concentrations of metal ions used to form the coordination structures.

A variety of studies have demonstrated the coordination nanostructures by altering organic ligands, such as carboxylate,3,5,12,18 hydroxyl,14,20,21 pyridine,9,22–25 and triazine.26–30 The triazine derivatives are important in a range of areas as ligands for supramolecular self-assembly of more complex mixed layers, hence it may be interesting to study the coordination properties of the triazine derivatives. In this article, we fabricated different coordination nanostructures by 4,4′,4″-(1,3,5-triazine-2,4,6-triyl) tribenzoic acid (TATB) molecules coordinating with Cu2+ or Fe3+ ion with different concentrations on highly oriented pyrolytic graphite (HOPG) surface. Scanning tunneling microscopy (STM) images directly demonstrated the transformation of the nanostructure from self-assembled honeycomb networks formed by TATB molecules via hydrogen bonding into coordination nanostructures formed through the TATB molecules coordinated with metal ions. Moreover, by varying the concentrations of Cu2+ or Fe3+ ions, coordination structures with different shapes and sizes have been fabricated, which illustrates the concentrations and valence states of the metal ions plays important roles in the coordination nanostructures.

Experimental

Sample preparation
4,4′,4″-(1,3,5-Triazine-2,4,6-triyl) tribenzoic acid (TATB) molecules used in this work weren’t further purified after purchased from Ark Pharm. Ferric chloride and copper chloride were...
purchased from DAMAO. Stock solutions of 30 mL each of TATB and metal chloride were prepared by dissolving these reagents in nonanoic acid purchased from Aladdin. The concentrations of TATB, CuCl$_2$ and FeCl$_3$ were $2.2 \times 10^{-3}$ M, $4.4 \times 10^{-3}$ M and $5.8 \times 10^{-6}$ M, respectively. The molar ratio of TATB to metal chloride in the final product was controlled by the deposited volumes. STM samples were prepared by dropping about 3.0 µL of the solutions on a freshly cleaved HOPG substrate. Then, the samples were kept at room temperature for about 0.5 h. Finally, the samples were investigated by STM at the liquid/solid interface.

**STM measurements**

The STM measurements were performed under atmospheric conditions, ambient temperature of about 25 °C. STM tips were made by cutting Pt/Ir (80/20) wire (California Fine Wire Co., Grover Beach, CA). All the STM images were acquired on a Pico-SPM (Molecular Imaging, Agilent Technology) scanning tunneling microscope operating in constant-current mode, the imaging parameters are stated in the figure captions, and extracted from raw data files using Pico Image Basic 6.2 program (Agilent, USA). To avoid possible experimental artifacts, each STM sample was repeatedly measured with different tips. The drift was calibrated using the underlying graphite lattice as a reference. High-resolution STM (HR-STM) images were low-pass-filtered.

**Constructed and computational details**

The above structures are calculated by first-principles density functional theory (DFT), as implemented in the Vienna *Ab initio* Simulation Package (VASP), with the PBE generalized-gradient exchange and correlation functional. A plane-wave basis set with a kinetic energy cutoff of 400 eV is adopted and the van der Waals interaction is considered with the optB86b-vdW method. A $1 \times 1 \times 1$ k-mesh was adopted for geometry optimization and total energy calculations. For structure relaxations, all atoms were fully relaxed until the residual force per atom was less than 0.02 eV Å$^{-1}$.

**Results and discussion**

After depositing a droplet of nonanoic acid solution containing TATB molecules onto a freshly cleaved HOPG surface, STM images show that TATB molecules form a uniform monolayer with molecules lying flat on the surface. Fig. 1a shows the typical honeycomb network of TATB on HOPG. The high resolution STM image in Fig. 1b revealed that each TATB molecule has a three-clover structure, where each clover can be assigned to a benzoic acid moiety on the bridge arm of TATB. The group composed of six TATB molecules forms a hexagonal network, and the unit cell of the assembly is marked in Fig. 1b (rhomboid in white solid line). The vectors for the unit cell, a and b, is respectively of $(3.68 \pm 0.2)$ Å and $(3.68 \pm 0.2)$ Å and the angle $\alpha$ between the vectors is $(60 \pm 1)^\circ$. The proposed model of the supramolecular honeycomb network shown in Fig. 1c indicates that each TATB molecule can interact with three adjacent molecules by the O–H⋯O hydrogen bonds (as illustrated by yellow dashed lines in Fig. 1c). Simulated STM image in the Fig. 1d matches well with the STM result. We measured the distance of O–H⋯O hydrogen bonds is approximately 2.64 Å, which are in good agreement with the lengths as previously reported in organic crystals.$^{31}$

When the nonanoic acid containing CuCl$_2$ and TATB (the molar ratio was adjusted to 1 : 2) is deposited on an HOPG surface, a well-ordered assembling structure composed of parallel chains labeled by white dotted line is found as shown in Fig. 2a, which is quite different from the honeycomb network composed of pure TATB molecules. Close examination in Fig. 2b reveals that these chains consist of some herringbone-like features and some dots. Since this structure only emerged after the introduction of CuCl$_2$, these herringbone-like features and dots were attributed to TATB molecules and Cu$^{2+}$ ions, respectively. The arrangement diagram of TATB molecules and Cu$^{2+}$ was superimposed on Fig. 2b, in which the red graphics mark the distribution of TATB molecules and the light blue circles mark the arrangement of Cu$^{2+}$ ions. Fig. 2c illustrates the corresponding optimized structure model. Two TATB molecules are bridged by Cu$^{2+}$ ion to form coordination dimers, and the dimers are connected by two C–H⋯O hydrogen bonds as illustrated by green dashed lines in Fig. 2c. Simulated STM images in Fig. 2d demonstrate that the C–H⋯O hydrogen bond lengths are $(3.0 \pm 0.2)$ Å and the O–Cu coordination bond lengths are $(4.0 \pm 0.3)$ Å, which is consistent with the STM result.

If we tune the molar ratio of CuCl$_2$/TATB from 1 : 2 to 1 : 1, another parallel chain labeled by white dotted line were...
and Cu\(^{2+}\) is realized on HOPG by increasing the concentration of metal ions. The three-clover graphics label TATB molecules, and the blue circle represents the copper ions. (c) Schematic model of 1:2 Cu\(^{2+}/\)TATB chain. (d) Simulated STM image of (c). (e) STM image of Cu\(^{2+}/\)TATB (Cu\(^{2+}:\) TATB = 1:1) self-assembled structures on HOPG surface. Set point: \(I = 0.57\) nA, \(V = -0.6\) V. (f) High resolution STM image of 1:1 Cu\(^{2+}/\)TATB chain superimposed with the schematic model. (g) Simulated STM image of (f). Blue dotted line: C–H···O hydrogen bond; yellow dotted line: O–H···O hydrogen bond.

obtained as shown in Fig. 2e. An overview image of the new chains reveals that such nanostructure is different with that above. Close examination in Fig. 2f reveals that these chains consist of some three-clover features and some dots. Similar to the above, these three-clover features and dots are designated as TATB molecules and Cu\(^{2+}\) ions, respectively. The optimized structure model superimposed on Fig. 2f illustrates two TATB molecules are bridged by Cu\(^{2+}\) ion to form coordination dimers, and the dimers on the same chain are connected by two O–H···O hydrogen bonds while the dimers on adjacent chains interact through O–Cu–O coordination bond. The corresponding simulated STM image in Fig. 2g coincides well with the experimental STM image in Fig. 2f and shows that O–H···O hydrogen bond lengths are approximately 2.6 ± 0.2 Å and O–Cu–O coordination bond lengths are approximately 4.0 ± 0.2 Å, which confirms that the O–Cu–O coordination bonds are the main driving force within the structure. Compared with the above chains, the dimers are also formed by O–Cu–O coordination bond. However, the driving force between the dimers on the adjacent chain changes from C–H···O hydrogen bond into O–Cu–O coordination bond, which results to the transformation of self-assembly structure of the chains. The result shows that the polymorphic self-assembly of TATB molecules and Cu\(^{2+}\) is realized on HOPG by increasing the concentration of metal ions.

Similar structural transformation was obtained by depositing a mixture of TATB molecules and FeCl\(_3\) with a molar ratio of 1:1 on the HOPG surface. The estimated coverage of iron on the surface is about 1.67 ion per nm\(^2\). It can be seen that the building blocks for the structure are the trimers of TATB molecules in which one bright dot is surrounded by three bright clover-shaped structures with a local threefold symmetry as marked by a green dotted circle in Fig. 3a. As the proposed model superimposed on the STM image in Fig. 3b, the central small bright spot is attributed to Fe\(^{3+}\) and the three clover-shaped bright spots represents three TATB molecules. In this model of TATB/Fe\(^{3+}\) trimers, three carboxyl groups linked with one Fe\(^{3+}\) ion through metal coordination, and the three carboxyl groups are connected to each other through O–H···O hydrogen bonds. The corresponding simulated STM image in Fig. 3c is consistent with the experimental STM image in Fig. 3b and shows that the length of three Fe–O coordination bonds between three oxygen atoms within the carboxyl groups of TATB molecules and the centered Fe\(^{3+}\) ion are 2.110 Å, 2.089 Å and 2.106 Å respectively, which is in good agreement with typical length of Fe–O bond in previous studies.

After adjusting the molar ratio of FeCl\(_3\) : TATB to 3:2, the self-assembly structure of TATB molecules and Fe\(^{3+}\) ions have undergone an obvious change from trimer structure to chain structure (Fig. 3d). Close examination in Fig. 3e allows us to propose the structure model: TATB molecules between adjacent chains interact via O–Fe–O coordination bond, while TATB molecules on the same chain are contact with each other by O–H···O hydrogen bonds. In the Fig. 3e, Fe\(^{3+}\) ions are marked by small brown dots and red circles, and TATB molecules are presented with optimized structural formulas. Obviously, O–Fe–
O coordination bond plays a dominant role in the assembly structure. It is worth noting that one Fe\(^{3+}\) ion coordinates with two carboxyl groups instead of three carboxyl groups mentioned above, as a consequence, the self-assembly structure of TATB and FeCl\(_3\) changes from network into chain. Fig. 3f shows the corresponding simulated STM image obtained by DFT calculation.

The DFT calculation provides an insight into the adsorption of Cu\(^{2+}\) or Fe\(^{3+}\) ion in the self-assembled structures of TATB molecules via metal–organic coordination bonding. Fig. 4 illustrates seven potential sites of metal adsorbed on TATB dimer and the DFT results for seven adsorption sites are listed in Table 1. The corresponding optimized structure models and STM simulation images are shown in ESI Fig. 1 and 2.

In the system of Cu\(^{2+}/Fe^{3+}\) and TATB dimer, the lowest total energy is \(E_{S7}\) and \(E_{S5}\), respectively, indicating that the best adsorption site of Cu\(^{2+}\) is S7 and that of Fe\(^{3+}\) is S5. On this basis, DFT calculations are performed on four systems with multiple potential configurations according to our experimental system mentioned above as shown in Fig. 5. The DFT results demonstrate that the configurations S4, S8, S12 and S16 have the lowest total energy of \(-1359.54\) eV, \(-1362.14\) eV, \(-686.99\) eV and \(-692.16\) eV in the system with the molar ratio Cu\(^{2+}/TATB\) of 1 : 2, Cu\(^{2+}/TATB\) of 1 : 1, Fe\(^{3+}/TATB\) of 1 : 1 and Fe\(^{3+}/TATB\) of 3 : 2, respectively. Comparing several configurations in each system, it is easy to find that Cu\(^{2+}/Fe^{3+}\) ions are all coordinated with TATB molecules in configurations S4, S8, S12 and S16 rather than one or more Cu\(^{2+}/Fe^{3+}\) ions only adsorbed on TATB molecules in other configurations. Significantly, the configuration in S4, S8, S12 and S16 is in good agreement with the corresponding model proposed in each system, which further proves the correctness of our proposed model.

| Table 1 | Calculated total system energy for single Cu\(^{2+}\) or Fe\(^{3+}\) ion adsorbed on TATB dimer at different sites. Energy unit: eV |
|---------|--------------------------------------------------|
| Energy  | Cu\(^{2+}\) | Fe\(^{3+}\) |
| \(E_{S1}\) | \(-678.26\) | \(-681.18\) |
| \(E_{S2}\) | \(-678.35\) | \(-679.2\) |
| \(E_{S3}\) | \(-678.73\) | \(-680.86\) |
| \(E_{S4}\) | \(-678.410\) | \(-682.35\) |
| \(E_{S5}\) | \(-678.00\) | \(-682.38\) |
| \(E_{S6}\) | \(-678.33\) | \(-681.63\) |
| \(E_{S7}\) | \(-679.18\) | \(-681.64\) |

Conclusions

The self-assembled hydrogen-bonded molecular network of TATB has been dramatically transformed after introducing CuCl\(_2/FeCl_3\) with different concentrations. The deposited monolayer of TATB molecules shows honeycomb structure. The Cu\(^{2+}/Fe^{3+}\) ions can incorporate with TATB molecule via metal–organic coordination to form new structures, and the self-assembled structures of TATB molecules on HOPG formed by metal–ligand coordination bond can be effectively adjusted by tuning the concentration of metal ions in mixed solution.

Author contributions

X. L. and L. W. conceived and designed the experiment, discussed and analyzed data, and wrote the manuscript. S. L. and X. L. performed sample preparation and STM. S. L., S. G., X. L., Y. Z. and Z. W. analyzed the data. S. L., S. F., C. G., and Y. L. performed the DFT calculations and theoretical analyses. All authors discussed the results and commented on the manuscript.

Conflicts of interest

There are no conflicts to declare.
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