Chiroptical inversion of a planar chiral redox-switchable rotaxane†

Marius Gaedke,a Felix Witte,a Jana Anhäuser,b Henrik Hupatz,a Hendrik V. Schröder,a† Arto Valkonen,c Kari Rissanen,c Arne Lützen,b Beate Paulusa and Christoph A. Schalley* a

A tetrathiafulvalene (TTF)-containing crown ether macrocycle with $C_s$ symmetry was designed to implement planar chirality into a redox-active [2]rotaxane. The directionality of the macrocycle atom sequence together with the non-symmetric axle renders the corresponding [2]rotaxane mechanically planar chiral. Enantiomeric separation of the [2]rotaxane was achieved by chiral HPLC. The electrochemical properties – caused by the reversible oxidation of the TTF – are similar to a non-chiral control. Reversible inversion of the main band in the ECD spectra for the individual enantiomers was observed after oxidation. Experimental evidence, conformational analysis and DFT calculations of the neutral and doubly oxidised species indicate that mainly electronic effects of the oxidation are responsible for the chiroptical switching. This is the first electrochemically switchable rotaxane with a reversible inversion of the main ECD band.

Introduction

Evidenced by the homochirality in our biosphere,1-3 chirality is a fundamental principle, which governs the molecular recognition and activity of virtually all biomolecules. Therefore, gaining control over the preferred isomer of a molecule or an assembly by carefully designing a molecular system is a worthwhile endeavour.

The term “chiroptical switch” has been used by Canary to refer to molecules, which are capable of “changes in their interaction with polarized light”.4 Potential applications are information processing, data storage and sensing. In this context, the ground breaking work of Feringa and co-workers5-7 on overcrowded alkenes, which act as light triggered chiroptical switches was awarded with the Nobel Prize in chemistry 2016 “for the design and synthesis of molecular machines”8 and underlines the general interest in this topic.

Mechanically interlocked molecules (MIMs)9-12 consist of parts that can move relative to each other guided by intramolecular forces. Therefore, we envisioned them to be ideal candidates for chiroptical switches in which conformational or even configurational changes in the MIM occur.

An achiral wheel with directionality in its atom sequence forms a chiral [2]rotaxane, when threaded onto a directional

Scheme 1 (a) Reversible one-electron oxidations of the TTF moiety, (b) reversible oxidation of a directional crown ether wheel bearing a TTF unit, (c) chiroptical switching of the planar chiral [2]rotaxane enantiomers.

† Electronic supplementary information (ESI) available: Synthetic procedures including full characterisation of new compounds, electrochemical data, crystallographic data and mass spectrometry data. CCDC 1910670. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c9sc03694f
Results and discussion

Synthesis and characterisation

The prerequisite for rotaxane formation is a sufficiently high binding constant between the crown ether and the ammonium axle. ITC experiments revealed an association constant of \( K_a = (3.6 \pm 0.3) \times 10^5 \text{M}^{-1} \) and a 1 : 1 stoichiometry for pseudorotaxane formation from \( \text{dTTFC8} \) and axle \( \text{A1} \) (Scheme 2). The binding constant is very similar to that of our previous non-directional TTF-decorated wheel \( \text{TTF8} \) \( \left( K_a = (4.4 \pm 0.4) \times 10^5 \text{M}^{-1} \right) \) for thermodynamic parameters see ESI,† Section 4), which indicates the positional change of the TTF unit not significantly affect the binding properties of the wheel.

As for the non-chiral \( [2] \)rotaxane 1, rotaxane formation was achieved with nitrile-oxide stopper \( \text{St1} \) using a catalyst-free end-capping protocol established by Takada and co-workers\(^{39} \) yielding a racemic mixture of rotaxane \( (\text{rac})-2 \) (73%). The non-ionic version \( (\text{rac})-2\text{Ac} \) (95%) was obtained through \( N \)-acylation with \( \text{Ac}_2\text{O} \) (Scheme 2). The \( ^1\text{H} \) NMR spectra of \( (\text{rac})-2 \) and \( (\text{rac})-2\text{Ac} \) (Fig. 1) reveal a diastereotopic splitting of the macrocycle’s methylene protons as well as of the axle methylene protons \( \text{H}_i \).\(^{7,28} \) The splitting of both macrocycle and axle protons is characteristic for the formation of a chiral, yet racemic \( [2] \)rotaxane. Isoxazole formation during stopper attachment leads to a strong downfield shift of 3.88 ppm for proton \( \text{H}_i \).

In \( (\text{rac})-2 \), the S-methyl protons on \( \text{dTTFC8} \) split into two singlets of the same intensity. Comparable rotaxanes also showed this behaviour on the same position,\(^{27,28} \) HR-ESI mass and tandem MS experiments support the interlocked architecture (Fig. S1†).

For non-ionic \( (\text{rac})-2\text{Ac} \), the shift of \( \text{H}_i (\Delta \delta = +0.28 \text{ppm}) \) and \( \text{H}_i (\Delta \delta = +0.76 \text{ppm}) \) relative to \( (\text{rac})-2 \) suggests that the wheel translates towards the isoxazole moiety in the absence of attractive interactions with the ammonium ion. Two sets of signals are observed due to the \( \text{cis} - \text{trans} \) isomerism of the amide bond in \( (\text{rac})-2\text{Ac} \). Variable temperature NMR experiments (Fig. S3†) in DMSO-\( d_6 \) reveal the same barrier \( (\Delta \delta^2 = 74 \pm 2 \text{kJ mol}^{-1}) \) for amide \( \text{cis} - \text{trans} \) isomerisation as observed for a similar acetylated rotaxane.\(^{41} \)
Optoelectronic properties

Photometric titrations of (rac)-2 and (rac)-2Ac with Fe(ClO₄)₃ (Fig. 2a and b) show similar bands for the three redox states (TTF, TTF⁺ and TTF²⁺) of both rotaxanes with distinct isosbestic points. These findings are consistent with structurally related rotaxanes featuring a non-directional TTF-decorated wheel.38

Cyclovoltammetric (CV) experiments were conducted with dTTFC8, (rac)-2 and (rac)-2Ac in dichloromethane (Fig. 2c). The potentials for (rac)-2 (116 mV and 407 mV) are considerably higher for both oxidation steps as compared to dTTFC8 (64 mV and 362 mV). Both oxidations are thus energetically disfavoured because of the charge repulsion between the TTF cation radical as well as the TTF dication and the ammonium station. In case of (rac)-2Ac (18 mV and 392 mV) the first oxidation is more easily accomplished in comparison to the free macrocycle and the second oxidation is disfavoured. We attribute this behaviour to a stabilising interaction with the isoxazole moiety on the axle for the first oxidation.

For the second oxidation, the limited accessibility of the TTF²⁺ by counterions caused by the steric demand of the axle needs to be taken into account. Again these trends were already observed for the non-directional macrocycle and rotaxanes thereof.38 The reversibility of the redox-waves of (rac)-2 and (rac)-2Ac strongly indicated that the interlocked structures remain intact during the redox switching, however it is reasonable to assume conformational changes to occur due to charge repulsion and charge stabilisation. The data does not show any significant change in the electrochemical properties by introducing directionality into the TTF decorated wheel.

Enantiomer separation on chiral HPLC and CD spectroscopy

The two enantiomers of (rac)-2Ac could be separated using HPLC with a CHIRALPAK® IA stationary phase. The optical purity was determined (>99% ee; Fig. 3a) and mirror-image CD spectra were obtained for the neutral enantiomers with bands at 242 nm and 325 nm (Fig. 3b). We assigned the absolute configuration based on the computational results (see below).
The oxidised species $2\text{Ac}^{2+}$ and $2\text{Ac}^{2+}$ show bands at the same wavelengths. While no sign inversion occurs at 325 nm, the band at 242 nm exhibits a sign inversion during the oxidation step. To exclude decomposition to be responsible for the switching, $2\text{Ac}^{2+}$ was reduced back to the neutral state using Zn dust and then showed the initial CD spectrum again (Fig. 3d and e dashed lines). Surprisingly, no other CD signals are observed at a higher wavelength, although the change in UV/Vis absorption is most pronounced at 460 nm and 844 nm for the radical cation and at 703 nm for the dication (Fig. 2b). The absorption is most pronounced at 460 nm and 844 nm for the radical cation and at 703 nm for the dication (Fig. 2b). The reason for the sign change remains ambiguous. In fact, conformational changes were observed to induce transitions in CD spectra of non-interlocked TTF derivatives with centrochiral elements earlier. Other examples show varying intensities or shifts of the maxima upon oxidation of the TTF attached.

Nevertheless, no TTF derivative is reported that shows a sign inversion in the ECD spectra of $2\text{Ac}$ due to a change in its electronic properties or to a (co-)conformational change, density functional theory (DFT) calculations were performed at the TPSS-D3(BJ) and B97X-D3 (ref. 76) levels. Conformational analyses reveal the structure depicted in Fig. 4 (left) to be the most stable one for $(R_{mp})_{2}\text{Ac}$. It is at least 18 kJ mol$^{-1}$ more favourable than any other possible conformation found by theory (see Table S2†). For $(R_{mp})_{2}\text{Ac}^{2+}$, there are two conformations relatively close in electronic energy: Conformer A (Fig. 4 middle) and B (Fig. 4 right) with a flipped naphthalene unit, ca. 9 kJ mol$^{-1}$ more stable than A. This conformational change is explained by the oxidation of $(R_{mp})_{2}\text{Ac}$ occurring fairly localised at the TTF unit. The emerging charge of the oxidised TTF moiety is then stabilised by the naphthalene that moves into close proximity of the TTF. Additionally, an atoms-in-molecules (AIM) analysis suggests that the electrostatic attraction between the naphthalene and TTF moieties outweighs all other non-covalent interactions for $(R_{mp})_{2}\text{Ac}^{2+}$, while the maximisation of non-covalent interactions (C–H⋯π and π–π stacking) is the most important factor in the neutral state (see ESI† for details).

The simulated CD spectra in Fig. 4 were obtained using simplified time-dependent DFT at the B97X-D3 level. The spectrum of $(R_{mp})_{2}\text{Ac}$ shows a deviation of around 40–50 nm, while that of $(R_{mp})_{2}\text{Ac}^{2+}$ is off by less than 20 nm compared to experiment. The experimentally detected sign inversion at 242 nm is reproduced well by the calculations. The conformational change of $(R_{mp})_{2}\text{Ac}$ upon oxidation, however, hardly influences the shape of the CD spectra as both conformations yield very similar CD spectra in the region between 230 and 400 nm. Therefore, we exclude the conformational change as the prime origin of the sign inversion.

To rationalise the optical behaviour of $(R_{mp})_{2}\text{Ac}$ and $(R_{mp})_{2}\text{Ac}^{2+}$, we examined its valence electronic structure, which is, as expected, dominated by orbitals localised at the TTF moiety (see Fig. S20†). Analysing the electronic transitions in the spectral region between 230 and 400 nm reveals that practically every excitation involves the TTF unit to some extent. While many transitions are of local nature, $i.e.$, between orbitals in close proximity, quite a few display a charge-transfer-like behaviour (insets Fig. 4 and S21†). For neutral $(R_{mp})_{2}\text{Ac}$, the vast majority of these transitions can be described by advancing an electron from an orbital centred at the TTF core, usually the HOMO, into an orbital located in another part of the rotaxane $(e.g. \text{the dimethoxy-phenyl moiety})$. For $(R_{mp})_{2}\text{Ac}^{2+}$, the corresponding transitions progress from some orbital in the molecule into an orbital localised at the TTF moiety, usually the LUMO or LUMO+1. This induces differently oriented magnetic dipole transition moments leading to different signs in the CD spectrum. Hence, we conclude that the sign inversion in the CD spectra upon oxidation can be exclusively attributed to the change of the electronic structure.
Conclusions

In conclusion, electrochemically switchable crown ether/ammonium [2]rotaxanes bearing a directional wheel are reported. The wheel features a redox-switchable TTF unit. The directionality had no observable impact on the electrochemical and optical properties of the racemic mixtures determined by UV/Vis spectroscopy and CV measurements. Instead, the pure enantiomers of the acetylated non-ionic derivatives display a redox-induced reversible inversion of the sign in the ECD spectrum without a change of absolute configuration. The mechanism and the absolute configuration of this chiroptical switch has been examined by computational methods. While co-conformational changes have hardly any impact on the ECD spectra, the changes in electronic structure induced by oxidation play a pivotal role. These results underline the impact of the mechanical bond, which allows the construction of intriguing switchable chemical assemblies with unexpected properties. This is the first in class example of a redox-controlled chiroptical switch with a complete sign reversal based on a mechanically planar chiral rotaxane. In the future, these properties could be employed in materials science to construct novel optoelectronic building blocks.

Conflicts of interest

There are no conflicts to declare.

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