Palladium nanoparticles supported on organofunctionalized kaolin as an efficient heterogeneous catalyst for directed C–H functionalization of arylpyrazoles†

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A heterogeneous catalyst system based on the immobilization of Pd0 nanoparticles onto organofunctionalized kaolin is reported with a view to introducing new synthetic routes of directed C–H functionalization of arylpyrazoles. Various characterization techniques revealed that the functional groups, 3-aminopropyltriethoxysilane (APTES) and phenyltrimethoxysilane, become strongly attached to the kaolin surface through Si–O–Si bonds and 3–6 nm sized Pd0 nanoparticles are uniformly decorated and stabilized through the organic amine moieties. The nano-palladium catalyst is durable, undergoing five times reuse with moderate catalytic activity. The XPS analysis of the catalyst before and after reaction suggested that the reaction might be performed via a catalytic cycle that begins with Pd0.

1. Introduction

The transition-metal-catalyzed direct arylation of arenes bearing different directing groups (such as arylpyridines, benzoazides, acetophenone oximes and anilides) represents a direct and promising approach to access ketones.1–11 In this context, pyrazole serves as an effective directing group for C–H functionalization methods involving cyclometalation. But, at present, methods for directed C–H acylation of 1-arylpyrazole12 are relatively scarce while there are numerous reports on directed C–H acylation of 2-phenyl pyridine (see Scheme 1).13–20 And though these homogeneous transition-metal catalysts offer high selectivity and yields under relatively mild operating conditions, their industrial applicability is limited by the inherent problem of catalyst separation from the product and its recycling.21,22 Considering the easy separation and reusability of a heterogeneous system, it would be desirable to develop highly efficient heterogeneous catalysts for these transformations.

In recent years, supported palladium nanoparticles (PdNPs) have attracted the broad interest of chemists due to their high selectivity and efficiency as heterogeneous catalysts for various reactions, such as hydrogenation of conjugated dienes, enantioselective allylic alkylation, carbon–carbon coupling, asymmetric alkylic substitution and electrocatalytic formic acid oxidation.23–27 A wide range of inorganic and organic supports have been utilised for the preparation of supported PdNPs catalysts, such as carbon,28–30 alumina,31 magnetically recoverable Fe3O4,32,33 MgO,34 composite oxide,35 modified silicas,36,37 carbon nanotube,38 polymers39,40 so on. Even though various approaches to PdNPs immobilization have been considered, progress is still required to further improve supported PdNPs catalytic systems for use in industries regarding the green chemistry context. Clay minerals are naturally abundant minerals of hydrous aluminum phyllosilicates and can be obtained from certain regions in high purity. By the organofunctionalization, clay can support various metal nanoparticles.
including Au, Ag, Pd, Pt, etc. General organic modifiers of clay include quaternary ammonium salts and silane coupling agents. Among the different types of clay minerals, kaolin seems to be more studied and used in environmental engineering, cosmetics and medicine because of its physical and chemical properties, crystal structure, and surface chemistry. However, there is rare report on kaolin supported nanocatalysts, while there are numerous literatures on the montmorillonite clay supported nano-metal catalysts.

Kaolin, Al₄Si₂O₇(OH)₄, is a dioctahedral-layered hydrated aluminosilicate clay of the 1 : 1 type with two distinct interlayer surfaces: a gibbsite-like with aluminum atoms coordinated octahedrally to corner oxygen atoms and hydroxyl groups, and a silica-like structure, where the silicon atoms are coordinated tetrahedrally to oxygen atoms. The alternating adjacent layers are linked by hydrogen bonds involving alumino (Al-OH) and siloxane (Si-O) groups. As a consequence of this structure, the silica/oxygen and alumina/hydroxyl sheets are exposed and functionalized kaolin was washed with (30 mL). The suspension was stirred at room temperature for 6 h. Then the solvent was removed by centrifugation and the residue on grinding furnished a black powder, denoted as NH₂,Ph@KL.

In a similar procedure, NH₂@KL and Ph@KL were prepared under vacuum for 12 h. The sample is assigned as NH₂,Ph@KL. In a similar procedure, NH₂@KL and Ph@KL were prepared adding 1.2 mmol of 3-aminopropyltrimethoxysilane or 1.2 mmol of phenyltrimethoxysilane to the KL, respectively.

2.1.2. Synthesis of Pd⁰ loaded organofunctionalized kaolin [Pd/NH₂,Ph@KL]. Pd⁰ loaded organofunctionalized kaolin was prepared by an impregnation-reduction method using the preparation method of Pd/γ-Al₂O₃ as our previous reports. For example, 3 wt% Pd NH₂,Ph@KL catalyst was prepared by the following procedure: 0.97 g of organofunctionalized kaolin NH₂,Ph@KL was dispersed in tetrahydrofuran (20 mL) solvent with constant stirring for 1 h at room temperature. Then the two kinds of aqueous solutions of PdCl₂ (0.01 M, 28.2 mL) and L-Lysine (0.03 M, 1 mL) were added to the mixture consecutively under vigorous stirring for 20 min. Subsequently, 0.1 M NaOH aqueous solution was added into the mixture to adjust the pH to 7. An aqueous solution of NaBH₄ (0.35 M, 4.5 mL) was added gradually in about 10 min to the suspension. Finally, the mixture was left to stand for 24 h and the solid was separated by centrifugation and washed with distilled water (4 times) and ethanol (once) followed by air drying overnight. The dried solid residue on grinding furnished a black powder, denoted as 3 wt% Pd/NH₂,Ph@KL. With the use of a similar procedure, 3 wt% Pd/NH₂@KL and 3 wt% Pd/Ph@KL were prepared using NH₂@KL and Ph@KL as the supports, respectively.

2.1.3. Synthesis of Pd⁰ loaded kaolin [Pd@KL]. For comparison, 3 wt% Pd/KL was prepared according to the above steps presented in this section using neat KL as the support instead of NH₂,Ph@KL.

2.1.4. Synthesis of Pd⁰ loaded CTAB functionalized kaolin [Pd/CTAB@KL]. For comparison, 3 wt% Pd/CTAB@KL was prepared by a modified impregnation-reduction method. (see ESIF)

2.2. Characterization techniques

The TEM study on the samples was recorded with a JEM-2100 transmission electron microscope with an accelerating voltage of 200 kV. The samples were fine powders deposited on a copper microgrid coated with a holey carbon film. X-ray photoelectron spectroscopy (XPS) analysis was performed on an ESCALAB 250Xi X-ray photoelectron spectrometer of ThermoFisher Scientific and Al Kα radiation was used as the X-ray source. All binding energies were referenced to the C₁s hydrocarbon peak at 284.80 eV. X-ray diffraction (XRD) was carried out using a Rigaku Ultimal IV X-ray diffractometer with Cu-Kα radiation (λ = 1.5406 Å) from 2θ = 10° to 80°, at a scan rate of 8° min⁻¹, operating at 40 kV and 40 mA. Elemental analysis was performed by vario EL cube Elementar analyzer of Elementar Analysensysteme GmbH. The specific surface areas of all samples were taken with an ASAP-2020 accelerated surface area and porosity analyzer of Micromeritics company, calculated by the BET method from the data in a P/V range between 0.06 and 0.30. The Pd loadings in the catalysts were measured on TGA-990 flame atomic absorption spectrophotometer (AAS) of the Beijing Purkinje General Instrument Company. The Pd sensitive wavelength is 244.8 nm. The FTIR (Fourier transform
infrared) spectra of the samples were recorded using a Thermo Scientific Nicolet 6700 infrared spectrometric analyzer in KBr matrix in the range 4000–650 cm$^{-1}$. The H$_2$TPR (temperature programmed reduction) of the samples was carried out in a Chem BET TPR/TPD Chemisorption Analyzer. An about 0.05 g sample was taken inside a quartz U tube and degassed at 120 °C for 0.5 h with argon gas flow. The sample was then cooled to 30 °C and at this temperature the gas flow was changed to 5% H$_2$ in argon. It was then heated at a heating rate of 10 °C min$^{-1}$ up to 600 °C and the spectra were recorded. The solid-state NMR experiments were performed on a Bruker Avance III 400 WB spectrometer equipped with a 9.39 T magnet at 297 K. The $^{1}$H–$^{13}$C cross-polarization magic angle spinning (CP/MAS) NMR spectra were recorded using a Bruker 4 mm standard bore MAS probe head with the ZrO$_2$ rotors spinning at 5.0 kHz rate with a Larmor frequency of 100.62 MHz. The spinning sidebands were suppressed by total sideband suppression (TOSS) technique. The $^{29}$Si CP MAS spectra were recorded on a Bruker AVANCE III 400 WB spectrometer equipped with a 4 mm standard bore CP MAS probe head whose X channel was tuned to 79.50 MHz for $^{29}$Si and the other channel was tuned to 400.18 MHz for broad band $^{1}$H decoupling, using a magnetic field of 9.39 T at 297 K. The dried and finely powdered samples were packed in the ZrO$_2$ rotor closed with Kel-F cap which were spun at 8 kHz rate. A total of 4000 scans were recorded with 6 s recycle delay for each sample. All $^{29}$Si CP MAS chemical shifts are referenced to the resonances of 3-(trimethylsilyl)-1-propanesulfonic acid sodium salt (DSS) standard ($d = 0.0$). The $^{27}$Al MAS spectra were recorded on a Bruker AVANCE III 400 WB spectrometer equipped with a 4 mm standard bore CP MAS probe head whose X channel was tuned to 104.27 MHz for $^{27}$Al, using a magnetic field of 9.39 T at 297 K. The dried and finely powdered samples were packed in the ZrO$_2$ rotor closed with Kel-F cap which were spun at 5 kHz rate, with a $\pi$/12 pulse. A total of 300 scans were recorded with 2 s recycle delay for each sample. All $^{27}$Al MAS chemical shifts are referenced to the resonances of [Al(H$_2$O)$_6$]$^{3+}$ standard ($d = 0.00$). Thin layer chromatography (TLC) was performed on pre-coated silica gel GF254 plates. The $^{1}$H NMR and $^{13}$C NMR spectra were obtained on a 500 MHz Bruker Avance III nuclear magnetic resonance spectrometer with CDCl$_3$ as the solvent. All chemical shifts are reported in ppm using TMS the internal standard. The structures of known compounds were further corroborated by comparing their $^{1}$H NMR data with those of literature.

2.3. Activity test

1-phenylpyrazole 1a and benzaldehyde 2a were used as the model reaction. In a typical reaction, 1-phenylpyrazole (27 mg, 0.20 mmol), benzaldehyde (65 mg, 0.60 mmol), tert-butyl peroxo-xybenzoate (TBPB, 156 mg, 0.80 mmol), catalyst (45 mg), and the solvent of cumene (2.0 mL) were charged in a 25 mL oven-dried reaction tube. Reaction was carried out 130 °C for 24 h in an oil bath under air condition. After being cooled to room temperature, the reaction solution was evaporated in vacuo. The residue was purified by flash column chromatography (silica gel, ethylacetate/petroleum ether = 1 : 3–1 : 5 as an eluent) to afford the desired product 3. All the products were also confirmed by comparing the $^{1}$H NMR and $^{13}$C NMR data with authentic samples.

2.4. Catalysts recycle experiment

After each reaction cycle, the substrate, solvent, and products were removed by centrifugal separation; the separated catalyst was washed drastically with 0.1 M NaOH ethanol solution (twice), distilled water (4 times), and then washed twice with ethanol followed by centrifugation and drying at 80 °C for 12 h. The recovered catalyst was used for the next cycle.

3. Results and discussion

3.1. Characterization of catalysts and support

3.1.1. SEM. The SEM image of kaolin, which is exploited from Inner Mongolia Autonomous Region of China, indicated that the kaolin possess smooth surface and flake-like morphology (see Fig. 1). Yang’s work proves that the smaller Pd nanoparticles were more easily formed and more highly dispersed on the flake-like kaolinite than rod-like and tube-like kaolinite.$^{24}$

3.1.2. FTIR. The FTIR spectra of KL, NH$_2$Ph@KL and 3 wt% Pd/NH$_2$Ph@KL are presented in Fig. 2. The grafting of the amine moiety onto the kaolin surface can be observed from the investigation of the CH$_2$ peaks and the symmetric stretching vibrations of –CH$_2$ groups from grafted APTES can be seen at 2955 cm$^{-1}$. Again, the presence of a weak N–H bending vibration at 1528 cm$^{-1}$, suggests the successful grafting of the organosilane on the kaolin surface. Three other distinct peaks around 900 cm$^{-1}$ and 1250 cm$^{-1}$ for $\delta$ (C–H) and 3050 cm$^{-1}$ for $\nu$ (C–H) of the benzene ring suggest the grafting of phenyl groups on the kaolin surface. After Pd$^0$ loading the basic structure of the functionalized kaolin remains unaltered.

3.1.3. Solid state NMR. The grafting of the binary functional groups into the kaolin framework could also be evidenced by solid state NMR spectra. Basically, the solid state $^{29}$Si CP MAS NMR spectra of NH$_2$Ph@KL shows $^{29}$Si NMR shifts assigned as Q$^0$(0Al) units of silica, and $\delta = -93.26$ ppm in natural kaolinite spectra corresponds to the signal of Si nuclei in the Q$^3$ polymerization state, in which three Si atoms bonded

![Fig. 1 SEM of kaolin (Inner Mongolia Autonomous Region of China).](image-url)
to a SiO₄ tetrahedron. On functionalization, apart from the Q³ peaks, another two downfield peaks T³ and T² were observed at −81.44 and −69.61 ppm, respectively, where T⁰ = [RSi(OET)ₙ(O-Si)]ₙ, in the case of NH₂,Ph@KL, confirming the grafting of the organic silane moieties onto the kaolin surface (see Fig. 3).

The ¹³C CP MAS NMR spectra have three distinct resonances at 10.9, 22.2, and 42.8 ppm, which correspond to the carbon atoms in the >Si–CH₂–CH₂–CH₂–NH₂ group in the sequence from left to right, proving the existence of APTES. In the spectra, another two peaks at 133.0 and 126.7 ppm were observed indicating the presence of carbon atoms of the Ph group (see Fig. 4). In ²⁷Al CP MAS NMR spectra, kaolin gives ²⁷Al line at ~0 ppm corresponding to the hexa-coordinated Al (see Fig. 5).

### 3.1.4. Elemental analysis and AAS

The results of elemental analysis of the prepared materials are shown in Table 1. Table 1 suggests that the C or N content increased correspondingly along with the organofunctionalization of kaolin. C and N were grafted onto the kaolin at 1.36 and 1.04 mmol g⁻¹, respectively, for 3 wt% Pd/NH₂,Ph@KL. The amounts of Pd loaded in the samples were determined by an AAS, and the Pd content of the 3 wt% Pd/NH₂,Ph@KL catalyst is approximately 3 wt%. But the Pd content of 3 wt% Pd/KL is only 1.43%. This result indicated that without organofunctionalization, the binding palladium capacity of kaolin is poor.

### Table 1 The Elemental Analysis and AAS Results

| Sample        | C (mmol g⁻¹) | N (mmol g⁻¹) | Pd loading (wt%) |
|---------------|--------------|--------------|------------------|
| KL            | 0.32         | —            | —                |
| NH₂@KL        | 0.34         | 1.08         | —                |
| Ph@KL         | 1.47         | —            | —                |
| NH₂,Ph@KL     | 1.44         | 0.96         | —                |
| 3 wt% Pd/NH₂@KL | 0.40   | 0.98         | 3.1              |
| 3 wt% Pd/Ph@KL | 1.41   | —            | 2.97             |
| 1 wt% Pd/NH₂,Ph@KL | 1.47 | 1.01         | 0.88             |
| 3 wt% Pd/NH₂,Ph@KL | 1.36 | 1.04         | 3.08             |
| 5 wt% Pd/NH₂,Ph@KL | 1.44 | 1.03         | 5.06             |
| Used 3 wt% Pd NH₂,Ph@KL | 1.38 | 0.78         | 2.03             |
| 3 wt% Pd/KL    | 0.35         | —            | 1.43             |

In addition, we did note a slight decrease in the Pd content after being cycled 5 times (2.03%, Table 1), which can decrease the catalytic activity on the basis of available Pd on the support surface.

### 3.1.5. XPS

In order to investigate the electronic state of the Pd, the X-ray photoelectron spectroscopy (XPS) of the fresh and
used (after 5th recycle) 3 wt% Pd/NH₂Ph@KL catalysts was investigated (see Fig. 6). It is shown that PdNPs on the support exist in the metallic state, corresponding to the binding energies of 334.83 and 340.23 eV over fresh catalyst and 335.13 and 340.43 eV over used catalyst, which are characteristic of the 3d₅/₂ and 3d₃/₂ peaks of Pd⁰. The unchanged valence of Pd confirmed that PdNPs-catalyzed directed C–H functionalization of arylpyrazoles was performed via a catalytic cycle that began with Pd⁰.

3.1.6. XRD. The XRD patterns of the fresh and used 3 wt% Pd/NH₂Ph@KL catalysts are shown in Fig. 7. The broad angle XRD patterns of NH₂Ph@KL indicate no structural change in the kaolin after functionalization with the organic groups. Three diffraction peaks originated at 12.32°, 20.22° and 24.86° correspond to the kaolin diffraction peaks respectively. 2θ = 20.81°, 26.59° is characteristic diffraction peaks of silicon dioxide. No evident peaks from palladium crystals were observed compared with kaolin sample. It is obvious that the spacing of the kaolin barely changed after surface modification and introduction of the reduced Pd nanoparticles, which indicated that the structure of kaolin was maintained and Pd nanoparticles were not intercalated in kaolin but present at the exterior of the kaolin layer.

3.1.7. TEM. The morphology and distribution of grown Pd nanoparticles on NH₂Ph@KL and KL were characterized by TEM. As shown in Fig. 8, the PdNPs of fresh and used (after 5th recycle) 3 wt% Pd/NH₂Ph@KL catalysts were both well dispersed and had narrow size distributions, which is quite consistent with the data obtained from the XRD analysis. The HRTEM images depicted clearly visible lattice fringes that evince the formation of crystalline Pd nanoparticles (Fig. 8d).

In contrast, apparent agglomeration phenomenon happened in the case of 3 wt% Pd/KL (Fig. 8c and f). Therefore, the presence of the organic amine moiety on the kaolin surface facilitated dispersion of the PdNPs on the KL surface through coordination and acts as a stabilizing agent.

Fig. 8j showed the selected area electron diffraction (SAED) of TEM images of the fresh 3 wt% Pd/NH₂Ph@KL catalyst. It exhibited four diffused rings due to (111), (200), (220), and (311) reflections of face centered cubic (fcc) Pd and indicated the crystalline nature of nanoparticles.

3.1.8. TPR. The reducibility of the supported Pd catalyst plays an important role in influencing its catalytic property. The H₂-TPR of NH₂Ph@KL and 3 wt% Pd/NH₂Ph@KL are reported in Fig. 9. The organofunctionalized kaolin (NH₂Ph@KL) does
not show any reduction peak below 110 °C. Meanwhile, the TPR curve of 3% Pd/NH₂,Ph@KL displayed one strong negative peak around 104 °C, which is characteristic of the decomposition of the β-hydride PdH₂.

3.2. The catalytic performance for directed C–H functionalization of arylpyrazoles

Initially, to examine the efficiency of different PdNPs catalysts prepared by us in the ortho-directed C–H functionalization reaction, we used 1-phenylpyrazole 1a with benzaldehyde 2a as the model substrates, tert-butyl peroxybenzoate (TBPB) as oxidant and cumene as solvent. As can be observed from Tables 2 and 3 wt% Pd/NH₂,Ph@KL showed higher activity than other catalysts including Pd/C, Pd/KL and Pd/CTAB@KL, affording (2-(1H-pyrazol-1-yl)phenyl)(phenyl) methanone 3aa in excellent yield (entries 1–4). Compared with Pd/NH₂,Ph@KL, Pd/NH₂@KL and Pd/Ph@KL exhibited lower catalytic activity (entries 5 and 6). These results indicated that the synergistic promoting effects from Ph and NH₂ functionalization increased the catalytic performance of supported PdNPs. The NH₂ functionalization helps in the uniform distribution of the active species PdNPs on the kaolin surface. The Ph functionalization increased the surface hydrophobicity, which facilitated the adsorption of organic molecules. We also examined the effect of different Pd loadings on the reaction. It was found that the catalytic efficiency was significantly influenced by palladium loading and that the catalyst with 3 wt% Pd exhibited the best performance (entries 7 and 8). The yield of product was decreased along with the reduction of oxidant’s amount (entry 9). A control experiment was carried out in the absence of TBPB and it failed to give the expected product 3aa, suggesting the importance of the oxidant (entry 10).

The directed acylation of 1-phenylpyrazole 1a with a variety of aldehydes was performed under the optimized reaction conditions, and the results are presented in Table 3. Various functional groups including methyl, methoxyl, chloro, bromo and cyano were compatible and the desired products were achieved in moderate to good yields (entries 1–6). Compared to its para isomers, o-chlorobenzaldehyde 2e delivered a lower yield, which probably resulted from the steric hindrance effect (entries 3 and 4). Naphthaldehyde 2h as well as benzaldehyde derivatives performed the acylation reaction expediently to give the corresponding product 3ah in 73% yield (entry 7). Using furan-2-carbaldelyde 2i as a heteroaromatic aldehyde resulted desired product (2-(1H-pyrazol-1-yl)phenyl)(furan-2-yl) methanone 3ai in comparatively lower yield (entry 8). Encouragingly, apart from aromatic aldehydes, the methodology was also successful in converting various aliphatic aldehydes with different chain lengths to the corresponding products: all of the catalytic reactions were likely to catalyze the formation of Pd/H complexes. The Pd/H complex may be responsible for the rearrangement of the products.

Table 2. Activity of various catalyst toward ortho-directed C–H activation reaction of 1-phenylpyrazole \(^a\)

| Entry | Catalyst | Oxidant | Yield\(^b\) (%) | TON\(^c\) |
|-------|----------|---------|----------------|---------|
| 1     | 3 wt%Pd/NH₂,Ph@KL | TBPB | 81 | 13 |
| 2     | 5 wt% Pd/C | TBPB | 63 | 10 |
| 3     | 3 wt% Pd/KL | TBPB | 52 | 8 |
| 4     | 3 wt% Pd/CTAB@KL | TBPB | 65 | 10 |
| 5     | 3 wt% Pd/NH₂@KL | TBPB | 67 | 11 |
| 6     | 3 wt% Pd/Ph@KL | TBPB | 59 | 9 |
| 7     | 1 wt%Pd/NH₂,Ph@KL | TBPB | 69 | 11 |
| 8     | 5 wt%Pd/NH₂,Ph@KL | TBPB | 70 | 11 |
| 9\(d\) | 3 wt%Pd/NH₂,Ph@KL | TBPB | 26 | 4 |
| 10    | 3 wt%Pd/NH₂,Ph@KL | — | — | — |

\(^a\) Reaction conditions: 1a (0.2 mmol), 2a (0.6 mmol), catalyst (45 mg), TBPB (0.8 mmol), cumene (2 mL), 130 °C, 24 h, under air. \(^b\) Isolated yield. \(^c\) The TON values were calculated on the basis of the amount of Pd metal. \(^d\) Reaction conditions: 1a (0.2 mmol), 2a (0.6 mmol), catalyst (45 mg), TBPB (0.2 mmol), cumene (2 mL), 130 °C, 24 h, under air.

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Fig. 9. H₂-TPR spectra of (a) NH₂,Ph@KL and (b) 3 wt% Pd/NH₂,Ph@KL.
even after recycling five times. Therefore, the kaolin supported PdNPs-catalyzed C–H activation reaction is more aligned with green catalysis than the homogeneous palladium catalysis is.

The radical trapping experiment was carried out to have a good understanding of the reaction mechanism. When the 1-phenylpyrazole 1a with benzaldehyde 2a reaction was carried out in the presence of 0.4 mmol TEMPO [(2,2,6,6-tetramethyl piperidin-1-yl)oxy], a radical scavenger, only a 30% yield of product was observed. And when 0.8 mmol TEMPO was added in model reaction, no product was detected, suggesting a possible radical approach. The comparison experiments of model reaction under air, oxygen and argon atmosphere situations were performed. Under argon atmosphere, the desired product 3aa was isolated in 71% yield. But surprisingly, under oxygen atmosphere, no product was detected. These results indicate that Pd⁰ is not oxidized to Pd^{II} by air (oxygen) and excess oxygen may hinder the catalytic cycle. Based on the previous reports\(^{65-69}\) and our own results, a tentative

Table 3 Substrate scope of ortho-directed C–H activation of arylpyrazoles

| Entry | 2-Arylpyridine | Aldehyde | Product | Yield\(^a\) (%) |
|-------|----------------|----------|---------|----------------|
| 1     | 1a             | 2b       | 3ab     | 62             |
| 2     | 1a             | 2c       | 3ac     | 56             |
| 3     | 1a             | 2d       | 3ad     | 75             |
| 4     | 1a             | 2e       | 3ae     | 41             |
| 5     | 1a             | 2f       | 3af     | 53             |
| 6     | 1a             | 2g       | 3ag     | 51             |
| 7     | 1a             | 2h       | 3ah     | 73             |
| 8     | 1a             | 2i       | 3ai     | 34             |
| 9     | 1a             | 2j       | 3aj     | 71             |
| 10    | 1a             | 2k       | 3ak     | 36             |
| 11    | 1a             | 2l       | 3al     | 73             |
| 12    | 1a             | 2m       | 3am     | 90             |
| 13    | 1a             | 2n       | 3an     | 77             |
| 14    | 1a             | 2o       | 3ao     | 78             |
| 15    | 1a             | 2p       | 3ap     | 72             |
| 16    | 1a             | 2q       | 3aq     | 39             |
| 17    | 1a             | 2r       | 3ar     | 41             |

\(\text{Entry 2-Arylpyridine Aldehyde Product Yield}^{a}(\%)\)

18 1a 2b 3ab 62
19 1a 2c 3ac 56
20 1a 2d 3ad 75
21 1a 2e 3ae 41
22 1a 2f 3af 53
23 1a 2g 3ag 51
24 1a 2h 3ah 73
25 1a 2i 3ai 34
26 1a 2j 3aj 71
27 1a 2k 3ak 36
28 1a 2l 3al 73
29 1a 2m 3am 90
30 1a 2n 3an 77
31 1a 2o 3ao 78
32 1a 2p 3ap 72
33 1a 2q 3aq 39
34 1a 2r 3ar 41

\(\text{Table 3 (Contd.)}\)

| Entry | 2-Arylpyridine | Aldehyde | Product | Yield\(^a\) (%) |
|-------|----------------|----------|---------|----------------|
18     | 1a             | 2b       | 3ab     | NP             |
19     | 1a             | 2c       | 3ac     | NP             |
20     | 1b             | 2a       | 3ba     | 77             |
21     | 1c             | 2a       | 3ca     | 65             |

\(\text{a Reaction conditions: 1 (0.2 mmol), 2 (0.6 mmol), catalyst (45 mg), TBPB (0.8 mmol), cumene (2 mL), 130 °C, 24 h, under air.}^{b}\) Isolated yield.

Fig. 10 Recyclability of catalysts.
Subsequently, the pyrazole group of 1-phenylpyrazole through the reductive elimination of intermediate YbSOH (or PdSOH) releasing benzoic acid. Finally, the releases a Pd0 to continue catalytic cycle.

4. Conclusions

In summary, we demonstrate a novel, highly efficient, and reusable heterogeneous nano-palladium catalyst supported on organofunctionalized kaolin for ortho-directed C–H activation of arylpyrazoles. Under the synergetic promoting effects from Ph and NH2 functionalizations, 3–6 nm size PdNPs were well-dispersed on kaolin surface and durable for five times reuse with moderate catalytic activity. Using supported palladium nanoparticles as the catalyst, a broad scope of aldehydes, including aryl aldehydes, heteroaromatic aldehydes and aliphatic aldehydes, react with arylpyrazoles to synthesize the corresponding aryl ketones in yield up to 90%. Compared with aromatic aldehydes, the steric hindrance effect of aliphatic aldehydes is a determining factor in the reaction. Further studies are in progress to develop kaolin supported palladium nanoparticles catalyst with higher activity and expand further the scope of heterogeneous palladium catalyzed C–H activation reaction.

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Conflicts of interest

There are no conflicts to declare.
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