Research Article

Removal of Congo Red by Silver Carp (*Hypophthalmichthys molitrix*) Fish Bone Powder: Kinetics, Equilibrium, and Thermodynamic Study

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1. Introduction

Industries, such as textile, leather, pulp and paper, cotton, paint, plastic, and cosmetics, use dyes to color their products and thus consume substantially a large volume of synthetic dye as well as process water for their production. As a result, they produce a considerable amount of colored wastewater. The textile industry consumes more than ten thousand tonnes of dye per annum out of which approximately one hundred tonnes of dye are discharged annually into the water stream [1]. Although dye-containing effluent contains aesthetic contaminants, it has negative impact on overall aquatic inhabitants through reduction in phytoplankton photosynthesis [2, 3]. Toxicity, contamination by color, nonbiodegradability, and complex structure of dye can cause enormous threat to human health and the aquatic environment as a whole [4]. Several classes of dyes are reported in the literature that include anionic dyes (e.g., all sort of direct dye, acid dye, and reactive dye), cationic dyes (all basic dyes), and nonionic dye (e.g., dispersed dyes) [5].

Congo red (CR) is known to be an anionic benzidine-based diazo synthetic dye, which is characterized by large molecular structure with two azo groups that are bound to aromatic rings having recalcitrant molecular structure [6]. The effluent from the above-mentioned industries, if not treated properly, can cause severe damage to aquatic habitants. Moreover, Congo red is responsible for causing cancer to mammals. However, it is a challenging task to biodegrade Congo red due to the stability of its structure [1]. Hence, effective removal of this dye form wastewaters has to be performed before discharging into the water bodies.

Removal of dyes from water and wastewater can be done by means of any of the methods, namely, biological, chemical, and physical methods. Several wastewater
treatment techniques have been reported to treat dyes from wastewaters, which include but not limited to, biological degradation [7], membrane separation [8], electrolysis [9], and oxidation [10]. These methods are having limitations such as high cost of plant, higher operational costs, interference by another constituents, and excessive sludge generation. Adsorption, on the other hand, is a much better process for contaminants removal from wastewater because of the availability of adsorbents, efficiency of the process, uncomplicated operation, and low-cost of the adsorbent. Some conventional adsorbents were used for treating the dye-containing wastewater that includes ash, nanomaterials, and activated carbon. Nowadays biomaterials and agricultural waste materials (such as pine bark, orange peel, peanut shell, sawdust, and sunflower) are getting more priority from the researchers throughout the world [11–13]. In this context, adsorption of CR by using waste biomaterials would have been a promising process in the field of wastewater treatment.

Numerous investigations have been conducted by using various bioadsorbents in the view to remove Congo red from industrial wastewater [14]. Parvin et al. reported considerably high adsorption capacity of chemically treated eggshell membrane (117.65 mg g⁻¹) for Congo red removal through adsorption [15]. Rejected tea and tea dust, on the other hand, were reported to be effective adsorbents in case of methylene blue (MB) dye removal with high adsorption capacity of 242.11 mg g⁻¹ and 175.4 mg g⁻¹, respectively [16,17]. Ghanizadeh and Asgari conducted an experiment for the quest of MB removal by using cattle and sheep bone charcoal as adsorbent where a maximum sorption capacity was found to be 5 mg g⁻¹ [18]. Furthermore, a different study was conducted for the removal of methylene blue dye from the wastewater using NaOH modified fish bone charcoal as adsorbent where the extent of sorption capacity was reported to be 605.82 ± 9.09 mg g⁻¹ [19]. Although different bioadsorbents were reported to be used in wastewater treatment, bone of silver carp fish, in particular, as an adsorbent in Congo red removal from simulated effluent has not been reported to the best of our knowledge. Silver carp is a very common and locally available fish in Bangladesh. Although different income groups of the country have different levels of fish consumptions, more than 50% of consumption of fish belongs to carp species that includes silver carp with few other species [20]. Huge consumption of silver carp fish generates a considerable number of bones as environmental waste. So, the conversion of this waste to potential adsorbent for dye removal from aqueous solution would be a promising technique to keep environment clean [21].

The aim of this research was to develop an adsorbent from bone of silver carp fish (Hypophthalmichthys molitrix) through some simple steps and to find out the potentiality of this adsorbent in removing Congo red from synthetic aqueous solution. In this regard, the influence of several essential operating parameters (e.g., solution pH, dosages of adsorbent, initial concentration of adsorbate, and contact time) on Congo red removal using fish bone powder (FBP) adsorbent was investigated. In addition, adsorption kinetics and equilibrium were analyzed by applying mathematical models to elucidate a plausible adsorption mechanism.

2. Materials and Methods

2.1. Adsorbent Preparation. The fish bone of silver carp (Hypophthalmichthys molitrix) was collected from a local fish market (Jashore, Bangladesh). The bones were separated from the fish meat, and the separated fish bones were washed with hot distilled water (343 K) for quite a few times. After that, the washed fish bones were dried under 423 K for 40 min using a drier (DGH-9030) and then cooled down to room temperature. The dried bones were pulverized by using a locally fabricated ball mill, and a particle size of 60 to 100 mesh was taken. The powdered form of fish bone was kept in an airtight container for conducting adsorption studies [19].

2.2. Adsorbate. Congo red (chemical formula: C₃₂H₂₇N₄Na₂O₅S₂; molar mass: 696.664 g mol⁻¹; λ_max = 500 nm), an anionic dye, was used in this study as an adsorbate. As-received analytical grade Congo red (Merck, Germany) was used in this study. A synthetic stock solution of 1000 mg L⁻¹ was prepared by dissolving a required amount of Congo red into a certain amount of distilled water. Varying concentration of Congo red dye solution was prepared by diluting synthetic stock solution of Congo red with distilled water.

2.3. Sample Characterization. Solution pH was measured over the range of 2–8 by a pH meter. Concentrations of dye were measured by a double beam ultraviolet-visible (UV-VIS) spectrophotometer (SHIMADZU UV-1800) at its characteristic wavelength. The functional group of FBP before and after adsorption experiment were determined from infrared (IR) spectra, which were recorded by Perkin Elmer (Spectrum 100) Fourier Transform Infrared (FTIR) spectrophotometer, with a wavelength range of 4,000 cm⁻¹ to 400 cm⁻¹.

2.4. Batch Adsorption Study. Batchwise adsorption experiments were conducted to investigate the influence of different parameters, that is, pH (2 to 8), adsorbent dosage (100 mg to 700 mg), contact time (0 to 240 min), and initial dye concentration (50–300 mg L⁻¹) on the adsorptive removal of CR. At a certain concentration, measured amount of Congo red solution (100 mL) was taken with a specific amount of FBP in a 250 mL Erlenmeyer flask. The initial pH of each solution was properly adjusted by adding either 0.1 M hydrochloric acid (HCl) or 0.1 M sodium hydroxide (NaOH). Proper mixing of the solutions was done by using a continuous stirrer with a constant speed of 150 rpm and at 303 K for each experiment (unless otherwise stated). After a certain time of stirring, the samples were taken out of the shaker and then filtered. The filtrates were taken for their absorbance measurement. However, removal percentage (% R) of CR dye and the equilibrium amount of dye adsorption (mg g⁻¹) was determined by utilizing the following equations [21]:

\begin{align*}
\text{% R} &= \frac{C_0 - C_f}{C_0} \\
V_{\text{eq}} &= \frac{C_0 V - C_f V_f}{V_f}
\end{align*}
\[ \%R = \frac{C_0 - C_e}{C_0} \times 100 \]  
\[ q_e = \frac{C_0 - C_e}{w} \times v, \]  
where the symbols \( C_0 \) and \( C_e \) are defined as the initial and final concentration of the Congo red solution (mg L\(^{-1}\)), respectively, \( q_e \) is denoted as the amount of Congo red adsorbed (mg g\(^{-1}\)), \( w \) is symbolized as the weight of FBP adsorbent, and \( v \) is termed as the solution quantity in liter (L).

2.5. Kinetics Study. To analyze the CR adsorption mechanism onto FBP, the adsorption kinetics were investigated by using both Lagergren’s pseudo-first-order equation (equation (3)) and pseudo-second-order rate equation (equation (4)). The equations can be expressed as follows [21]:

\[ \log (q_e - q_t) = \log q_e - k_1t, \]  
\[ \frac{t}{q_t} = \frac{1}{k_2q_e^2} + \frac{t}{q_e}, \]  
where \( q_e \) and \( q_t \) are symbolized as the adsorption capacity (mg g\(^{-1}\)) at equilibrium and that at time \( t \) (min), respectively, while \( k_1 \) (min\(^{-1}\)) and \( k_2 \) (g mg\(^{-1}\) min\(^{-1}\)) are termed as the reaction rate constant for Lagergren’s pseudo-first-order model and pseudo-second-order model, respectively.

In order to examine the inherent mechanism of the Congo red adsorption onto FBP, the intraparticle diffusion model, as well as Bangham’s equation, was also investigated. Intraparticle diffusion can be well explained according to Weber–Morris model, which is expressed as follows [22]:

\[ q_t = k_3q_e^{1/2} + I, \]  
where \( q_e \) is denoted as the amount of CR dye adsorbed per an unit mass of FBP adsorbent (mg g\(^{-1}\)) at time \( t \) and \( k_{int} \) is termed as the rate constant for intraparticle diffusion model (mg g\(^{-1}\) min\(^{-1}\)).

On the other hand, Bangham’s pore diffusion model can be equated as follows [23,24]:

\[ \log \left( \frac{C_0}{C_0 - q_t \cdot m} \right) = \log \left( \frac{k_0 \cdot m}{2.303 \cdot v} \right) + \sigma \cdot \log t, \]  
where \( m \) (g L\(^{-1}\)) is adsorbent concentration and \( k_0 \) and \( \sigma \) are constants (often called Bangham’s parameters), while \( C_0, q_0, \) and \( v \) possess their usual meaning as stated earlier in this article.

2.6. Isotherm Study. To explain the equilibrium distribution of dye particles onto the surfaces of FBP, the three most common adsorption isotherm models were studied. In this study, the Langmuir isotherm, the Freundlich isotherm, and the Temkin isotherm models were analyzed, which were mentioned correspondingly in the following equations [25]:

\[ \frac{C_e}{q_e} = \frac{1}{bq_m} + \frac{C_e}{q_m}, \]  
\[ \log q_e = \log k_F + \frac{1}{n} \log C_e, \]  
\[ q_e = BlnK_T + BlnC_e, \]  
where \( C_e \) (mg L\(^{-1}\)) signifies the equilibrium adsorbate (CR) concentration, while \( q_e \) (mg g\(^{-1}\)) and \( q_m \) (mg g\(^{-1}\)) indicate the equilibrium sorption capacity and the maximum capacity of sorption of the adsorbent, respectively, \( b \) (L mg\(^{-1}\)) indicates the Langmuir adsorption constant related to free energy of adsorption, \( k_F \) (mg g\(^{-1}\)) designates the adsorption capacity at unit concentration, \( 1/n \) denotes the adsorption intensity, \( K_T \) signifies the Temkin equilibrium binding constant (L mg\(^{-1}\)) corresponding to the maximum binding energy, and \( B \) (kcal mol\(^{-1}\)) is the heat of adsorption \((B = RT/b)\).

2.7. Thermodynamic Study. Thermodynamic investigations for adsorption of CR on FBP were conducted by considering four dissimilar temperatures (298, 308, 318, and 328 K) where all other key parameters (e.g., solution pH, solution concentration, and adsorbent dose) were kept constant. The thermodynamic parameters, such as Gibbs energy (\( \Delta G \)), enthalpy of adsorption (\( \Delta H \)), and entropy of adsorption process (\( \Delta S \)), were calculated from the following equations [26]:

\[ \Delta G = -RT \ln q_e, \]  
\[ \ln q = \frac{\Delta S}{R} \frac{\Delta H}{RT}, \]  
where \( R \) (kJ mol\(^{-1}\)), \( T \) (K), and \( q \) (mg g\(^{-1}\)) indicate the ideal gas constant, temperature, and the uptake capacity related to Langmuir constant, respectively.

3. Results and Discussion

3.1. Characterization of the FBP Adsorbent

3.1.1. General Characterization. The bulk density of the fish bone powder was measured to be 0.67 g cm\(^{-3}\). The moisture content and the ash content of the adsorbent were found to be 3.1% and 36%, respectively. The pH of the adsorbent was determined by mixing fish bone powder with distilled water, and the value of pH was found as 7.8.

3.1.2. Determination of pH\(_{pzc}\). The pH of the point of zero charge (pH\(_{pzc}\)) is an important parameter that gives idea to understanding the surface chemistry of an adsorbent. It is one of the essential physicochemical parameters that indicate the net charge of an adsorbent when it comes in contact with water [27]. It is defined as pH of the solution at which the surface charge of the adsorbent has zero value. The global surface charge of the adsorbent is negative at pH higher than
3.1.3. FTIR Analysis. Adsorbent surface properties play a vital role in the entire adsorption process [1]. Therefore, the Fourier transform infrared (FTIR) spectrum is very important to understand the functional groups available in an adsorbent. Fish bone powder is supposed to consist of different functional groups on its surface. These functional groups may have acidic or alkaline properties. The FTIR spectra of FBP in case of before Congo red adsorption and after Congo red adsorption are presented in Figure S1 (Supplemental File 1). The spectrum of FBP before adsorption shows several peaks at different wavelengths, which indicates that various functional groups are present in the adsorbent. The FTIR spectroscopic study showed strong absorption bands at 3478.85 cm\(^{-1}\) and 3415.00 cm\(^{-1}\) for hydroxyl group (O-H) stretching of alcohol group for intermolecular bonding, while the peak at 3236.10 cm\(^{-1}\) is assigned by O-H stretching of carboxylic group [29]. The medium band at 1638.54 cm\(^{-1}\) and 1617.81 cm\(^{-1}\) commonly appeared due to N-H bending of amine [29]. The band at 1112 cm\(^{-1}\) is due to O-H vibration of methanol. The absorption bands at 613.85 cm\(^{-1}\) and 470.15 cm\(^{-1}\) signify the presence of PO\(_3^{2-}\) group on the adsorbent surface. It is obvious from Figure S1 that the absorption bands (after adsorption) at lower range slightly shifted due to adsorption of Congo red. This can be attributed to the fact that the functional groups present in the adsorbent might be involved in complexation with adsorbate.

3.2. Effect of Solution pH. There are several factors that affect adsorption of anionic dyes. Solution pH is one of them because of its impact on both the active sites of the adsorbent and the ionization process of the dye molecule in the solution. The effect of solution pH on the adsorption process of Congo red is shown in Figure 2. The experiments were conducted by pouring 100 mL Congo red (150 mg L\(^{-1}\)) into 100 mg of fish bone adsorbent. It was seen that pH of the adsorbate provides a noteworthy effect on the adsorption process of CR onto FBP surface. Maximum adsorption (96.3%) was found to occur at pH 2, while the adsorption decreased with further rise in pH. The Congo red removal decreased from 96.3% to 17% when the pH was changed from 2 to 8. This phenomenon can be explained by different interactions between adsorbate and adsorbent surface, such as electrostatic force, degree of ionization, hydrogen bonding, and speciation. The sulfonate moiety of Congo red contains negative sulfonic groups (−SO\(_3^-\)), which dissociates to polar groups (R−SO\(_3^-\)) at acidic medium [30]. Hence, the acidic medium favors the adsorption of Congo red onto FBP surface. According to the point of zero charge (pH\(_{\text{pzc}}\) = 7) of fish bone powder shown in Figure 1, the predominant functional groups of the adsorbent surface are positive at pH < pH\(_{\text{pzc}}\) and negative at pH > pH\(_{\text{pzc}}\). In acidic condition, the hydrophilic adsorption (electrostatic forces as well as hydrogen bonding) become dominant in the adsorption mechanism. However, the decrease in Congo red removal at higher pH may be due to the competition between excess OH\(^-\) ions and dye anions for the adsorption sites. Similar results have been reported for Congo red adsorption on pine bark and activated carbon [11,30]. However, for carrying out further studies using different parameters, the pH value 2 was considered as an optimum value.

3.3. Effect of Adsorbent Dosage on CR Dye Adsorption. The optimum adsorbent dosage is a significant parameter that affects the amount of adsorbed CR by fish bone powder. Varying amounts of FBP (from 100 mg to 700 mg) were used in the present study keeping all other parameters constant. Percent removal of CR is depicted in Figure 3 from where it is obvious that, irrespective of the initial concentrations, removal of Congo red increased from about 91% to 97% with pH\(_{\text{pzc}}\) (i.e., pH > pH\(_{\text{pzc}}\)) and positive at pH lower than pH\(_{\text{pzc}}\) (i.e., pH < pH\(_{\text{pzc}}\)). Published reports revealed that adsorption of cationic dye occurred at pH > pH\(_{\text{pzc}}\) due to the existence of some functional groups, such as COO\(^-\) and OH\(^-\). On the other hand, the adsorption of anionic dye occurred at pH < pH\(_{\text{pzc}}\) [28]. In this investigation, pH\(_{\text{pzc}}\) of the fish bone powder was determined to be 7 as shown in Figure 1. This signifies that at pH < pH\(_{\text{pzc}}\) the adsorbent surface carries positive charge and at pH > pH\(_{\text{pzc}}\) the negative charge exists on the adsorbent surface.
increasing adsorbent dosages at the given condition. With
the increase in the amount of adsorbent, the total number of
active sites of the FBP increases. Therefore, significantly large
amount of CR dye can be adsorbed by the FBP resulting in
higher dye removal efficiency. Similar results were reported
to occur for the Congo red removal by utilizing eggshell
membrane [15].

3.4. Effect of Contact Time for Adsorption Reaction at Varying
Dye Concentration. In adsorption process, contact time is
supposed to be an important parameter to consider because
it gives an idea about how long it will take to reach equi-
librium. To demonstrate the effect of contact time on ad-
sorption, a series of experiments were conducted by varying
the initial concentration of Congo red (50–300 mg L−1)
keeping all other parameters constant. The obtained results
are depicted in Figure 4, which demonstrates that the dye
uptake (mg g−1) was very fast at the initial stage (e.g., up to
45 min) after which it became slow and formed plateau.
However, for the sake of ensuring a complete reaction,
60 min was selected for performing all other experiments. It
is also obtained from Figure 4 that the adsorbed amount of
dye increased with the increase in initial concentration. It
signifies that the more adsorption sites of the adsorbent were
occupied with the increased amount of Congo red. The result
is consistent with the findings reported by other research
groups [17, 31].

3.5. Adsorption Kinetics. For determining the adsorptive
removal mechanism of CR onto FBP surface, two types of
kinetic models were studied to analyze the experimental
data. The models were (i) pseudo-first-order and (ii) pseudo-
second-order kinetic models. Considering the extent of the
correlation coefficient ($R^2$), it was obvious that the process
of adsorption could not be well matched by the pseudo-first-
order kinetic model. On the other hand, pseudo-second-
order kinetic model ($\frac{t}{q_t}$) versus $t$ according to (4) was
analyzed at different initial CR concentrations, which is
shown in Figure 5. It was found that the experimental data
fitted very well with pseudo-second-order kinetic model.
However, the values of reaction rate constant ($k_2$), equi-
librium uptake capacity ($q_e$), and correlation coefficient ($R^2$)
were tabulated in Table 1 along with other parameters de-
termined from different models. The values of correlation
coefficient were between 0.9999 and 1, while the values of
sorption capacities determined from this model ($q_{e,cal}$) were
closer to that obtained experimentally ($q_{e,exp}$). These out-
comes articulate that the pseudo-second-order model can
describe the sorption kinetics very well. Similar results were
reported for the adsorptive removal of Congo red on various
adsorbents, such as pine cone [32], Bengal Gram Seed Husk
[33], and eggshell membrane [15].

To reconfirm the best fit kinetic model, the data was
further analyzed by squared sum of errors (SSE) method. It is
presumed that the lower the value of SSE, the more fit the
model. The SSE values for different concentrations were
calculated by using the following equation [34]:

$$SSE = \sum \left( \frac{q_{e,exp} - q_{e,cal}}{q_{e,exp}} \right)^2,$$

where $q_{e,exp}$ (mg L$^{-1}$) is the experimental sorption capacity at
equilibrium and $q_{e,cal}$ (mg L$^{-1}$) is the calculated sorption
capacity obtained from corresponding kinetic model. The
SSE values for different initial concentrations obtained were
tabulated in Table 1. The lower SSE values for pseudo-
second-order kinetic model indicated that the sorption ki-
netics of Congo red onto fish bone powder could be better
described by pseudo-second-order model.

3.6. Mass Transfer Mechanism. There are several steps in-
volved in the solid-liquid sorption process. The transmission
of the adsorbate initially occurs from the bulk solution to the
exterior of the adsorbent. Then, the adsorbate diffuses
through the boundary layer and at the end it moves through the pores of the adsorbent. In the course of the transmission of adsorbate, it is adsorbed at the available active sites. Although the pseudo-second-order kinetic model was found to be the best fit model, it did not provide any information regarding mass transfer mechanism. In this study, the mass transfer mechanism was analyzed by using a well-known diffusion model, namely, Weber–Morris intraparticle diffusion model [35]. Diffusion control is supposed to be the most consistent explanation for almost all reported results in which the rate equation of pseudo-second order for adsorption has been found to fit the data well [36]. Intraparticle diffusion model was studied by plotting $q_t$ versus $t^{0.5}$ according to equation (5) and is depicted in Figure 6 where two distinct linear regions are noticed. The first linear region signifies the rapid external surface loading or film diffusion, while the second linear region represents the pore diffusion. The rate constant of pore diffusion ($k_{ud2}$) was determined from the slope of each line, while intercept of each line represented the boundary layer thickness during adsorption process. A schematic diagram of Congo red transport is presented in Figure 7. The calculated values of $k_{ud}$ and $I$ are tabulated in Table 1. Similar outcomes were reported for activated carbon as well [37].

The kinetic data for Congo red adsorption on FBP was further interpreted by using Bangham’s diffusion model (equation (6)) in order to examine whether or not the pore diffusion was the only rate controlling step in the sorption process. For determining Bangham’s parameters, a plot of $\log[\log(C_t/(C_0 - q_t \cdot m))]$ versus $\log t$ for different initial concentration was drawn as shown in Figure S2 (Supplemental File 2). The plot was found to be linear for each initial concentration. Bangham’s parameters ($k_0$ and $\sigma$) were calculated from the intercept and slope, respectively, and were tabulated in Table 1. The correlation coefficients for different initial concentrations were in a range from 0.9644 to 0.7588. According to the above-mentioned analysis, it could be concluded that both film diffusion and pore diffusion are involved in different stages of the sorption process [24].

### 3.7. Adsorption Isotherm Studies

Isotherms of adsorption are the indicative of adsorbate partitioning among the solid phase (adsorbent) as well as the liquid phase (adsorbate). The isotherms find the extent of adsorption of certain adsorbate onto a specific adsorbent at defined conditions. In this present study, two common isotherm models were used to discuss the adsorption isotherms for CR removal onto FBP at pH 2. The discussed models were Langmuir as well as Freundlich isotherm models by which the data were analyzed.

Langmuir isotherm model considers a monolayer adsorption process where the sorption occurs at homogeneous surfaces of the adsorbent. In this case, by using equation (7), $(1/q_m)$ versus $(1/C_0)$ is plotted, which is shown in Figure 8(a). The maximum sorption capacity $q_m$ (mg·g⁻¹) and the Langmuir constant $b$ (L·mg⁻¹) are determined from the slope and intercept of the straight line, respectively.

Freundlich isotherm model (equation (8)), on the other hand, is supposed to be applied for adsorption of adsorbent on heterogeneous surfaces. This model can deal with
Figure 6: Intraparticle diffusion analysis.

Figure 7: Schematic diagram of Congo red transport and sorption mechanism.

Figure 8: Isotherm analysis using (a) Langmuir and (b) Freundlich model.
multilayer adsorption process too. However, the related parameters such as $K_F$ and $n$ can be determined from the intercept and slope of the straight line of a plot log $C_e$ versus log $q_e$, which is depicted in Figure 8(b). Temkin isotherm parameters such as Temkin isotherm constant ($K_T$) and heat of adsorption ($B$) are determined from the slope and intercept of the plot of a plot $q_e$ versus ln $C_e$ (for simplicity figure not shown) by using equation (9).

The extent of different isotherm parameters obtained from the above-mentioned analyses is tabulated in Table 2. On the basis of the extent of the correlation coefficients, it can be summarized that the adsorption process can be better explained by using the Langmuir isotherm model ($R^2 = 0.9929$) compared to the Freundlich and Temkin isotherm models. Thus, the result indicated that a monolayer adsorption occurred during the discussed adsorption process. Similar findings were reported by various researchers for different adsorbents [15, 33, 38]. It is noteworthy to mention that the adsorption capacity of the current adsorbent (FBP) is considerably high compared to other biomaterial-derived adsorbents as shown in Table 3.

### Table 2: Linearized isotherm coefficients for Congo red adsorption onto FBP surface.

| Isotherms          | Parameters       | Value          |
|--------------------|------------------|----------------|
| Langmuir isotherm  | $q_m$ (mg·g⁻¹)   | 666.67         |
|                    | $b$ (L·mg⁻¹)     | 0.0020         |
|                    | $R^2$            | 0.9929         |
| Freundlich isotherm| $K_F$ (L·mg⁻¹)   | 1.358          |
|                    | $N$              | 0.848          |
|                    | $R^2$            | 0.9895         |
| Temkin isotherm    | $B$ (kcal·mol⁻¹) | 0.05           |
|                    | $K_F$ (L·mg⁻¹)   | 0.031          |
|                    | $R^2$            | 0.8971         |

### Table 3: Comparison of sorption capacities of low-cost adsorbents for various dyes.

| Materials                  | Dye     | $q_m$ (mg·g⁻¹) | References |
|----------------------------|---------|---------------|------------|
| Fish bone powder (FBP)     | CR      | 666.67        | In this work |
| Fish bone                  | CR      | 37.36         | [4]         |
| Eggshell membrane           | CR      | 117.65        | [15]        |
| Modified fishbone charcoal  | MB      | 605.82        | [19]        |
| Untreated fishbone charcoal | MB      | 70.42         | [19]        |
| Raw pine cone               | CR      | 32.65         | [32]        |
| Acid-treated pine cone      | CR      | 40.19         | [32]        |
| Bone char                   | CR      | 73.91         | [39]        |
| Tuna fish bone              | CR      | 329.0         | [40]        |

#### 3.8. Effect of Temperature and Thermodynamic Studies.

The effect of temperature on Congo red adsorption from aqueous solution using fish bone powder was tested in this study. Batch experiments were carried out at varying temperatures keeping all other parameters constant. The results are depicted in Figure 9 where it is evident that the removal of Congo red increases from 141 mg·g⁻¹ to 381 mg·g⁻¹ with the increase in temperature from 298 K to 328 K. An increase in temperature leads to fast diffusion of Congo red molecules through the external boundary layer and the internal pores of the adsorbent due to less resistance offered by viscous forces in the aqueous phase [41]. Moreover, the increase in sorption capacity at elevated temperature may happen due to the pore size enlargement of the adsorbent [42].

Thermodynamic parameters are the essential parametric tools to understand whether the adsorption process is physical or chemical, spontaneous or nonspontaneous, and exothermic or endothermic. To determine these parameters, ln $q$ versus $(1/T)$ was plotted using equation (11), which is shown in Figure 10. The values of change in enthalpy ($\Delta H$) and entropy ($\Delta S$) can be determined from the slope and intercept of the graphical plot, while the values of Gibbs energy ($\Delta G$) are calculated by using equation (10). However, the thermodynamic parameters calculated from this study are presented in Table 4. From the table, it is obvious that the values of $\Delta G$ become successively more negative with the temperature rise. Such phenomenon signifies the process of CR adsorption onto FBP surface is spontaneous and feasible. The positive value of change in enthalpy indicates that the adsorption is endothermic in nature, which can be substantiated by the increase in adsorption capacity with increase in temperature. To know about the binding method between Congo red and fish bone powder, the activation energy ($E_a$) of the adsorption process was determined as well. Physical adsorption (physiosorption) generally has low activation energy ($<40$ kJ·mol⁻¹), while chemical adsorption (chemisorption) has considerably high activation energy ($>40$ kJ·mol⁻¹) [43]. The activation energy was determined to be 29.84 kJ·mol⁻¹, which indicated that the adsorption process was of physisorption type. On the other hand, the affinity and increased randomness or disorderliness at the adsorbent-adsorbate interface during the period of adsorption of Congo red by FBP were corroborated by the positive value of $\Delta S$ [15,44].
4. Conclusion

This research investigated the applicability of fish bone as an adsorbent in the view to the anionic dye (Congo red) removal from synthetic wastewater solution. The isoelectric point (pH\textsubscript{PZC}) indicated that the adsorbent surface is positive below the value of pH 7, which resulted in a favorable Congo red adsorption at low pH. From the analysis of kinetic data, it was obvious that the adsorption of Congo red followed a pseudo-second-order reaction model. Experimental data was analyzed with different models of isotherm, and it was perceived that the Langmuir isotherm model fitted very well with the experimentally obtained data compared with Freundlich and Temkim isotherm models. The maximum sorption capacity was evaluated to be 666.67 mg·g\(^{-1}\). The extent of activation energy (29.84 kJ·mol\(^{-1}\)) confirmed that the binding between Congo red and fish bone powder occurred through a physisorption process. In addition to this, the adsorption process was observed to be endothermic and spontaneous in nature. The findings of this research proclaimed that fish bone powder derived from a locally available fish (silver carp) can be an effective adsorbent for Congo red removal from wastewater.

Data Availability

The data used to support the findings of this study are available from the corresponding author upon request.

Conflicts of Interest

The authors declare that they have no known conflicts of interest or personal relationships that could have appeared to influence the work reported in this paper.

Authors’ Contributions

Shahanaz Parvin was responsible for conceptualization, methodology, validation, and supervision. Manjur Hussain and Farhana Akter were responsible for investigation, formal analysis, and writing of the original draft. Biplob Kumar Biswas was responsible for conceptualization, supervision, and review and editing of the manuscript.

Supplementary Materials

Figure S1: FTIR spectra of the fish bone powder (a) before and (b) after adsorption. Figure S2: Bangham’s diffusion model for Congo red adsorption by fish bone powder adsorbent. (Supplementary Materials)

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