Transition Metal Oxide-Based Perovskite Structures as a Bifunctional Oxygen Electro catalysts: Fe Doped LaCoO3 Nanoparticles

Shaik Mahammad Rafi

Abstract: In this report, we have investigated lanthanum cobalt iron (LaCo1−xFexO3) perovskite nanoparticles synthesized by combining metallic nitrates, deionized water, and citric acid by using sol-gel method and subsequently calculated at 400 °C for 1h and 900 °C for 7h, respectively. The formation of single-phase perovskite structure is a series of LaCo1−xFexO3 (x = 0, 0.2, 0.4, 0.6, 0.8, 1). The crystal structure, mean particle, and morphology properties of the prepared LaCo1−xFexO3 perovskite oxide nanoparticles were examined by X-ray diffraction (XRD), field emission scanning electron microscopy (FESEM). The perovskite structure has shown special performance for oxygen reduction reaction (ORR) and oxygen evolution reaction (OER) catalytic activity in alkaline medium. As the combined valence transition metal oxides are rising capable candidates for bifunctional electrocatalysts, the electrochemical performance of the LaCo1−xFexO3 catalyst was thoroughly investigated. Koutecky-Levich results on the ORR polarization curves of all compounds shows that the four-electron pathway is favorable on these perovskite oxides. In this paper, we report B-site Fe doping in perovskite structure is a sufficient strategy to improve ORR and OER catalytic activity for application in metal-air batteries.

Keywords: Metal-air batteries, Bifunctional catalyst, Perovskite oxides, Electrochemical behavior.

I. INTRODUCTION

Due to technological developments, people’s demand, safety, clean, and renewable energy sources have been required considerable attention over the years. Also, rechargeable metal-air batteries are considered one of the most exciting technologies owing to its low price, high power density, environmentally friendly [1]. The development of rechargeable metal-air batteries is mostly restricted by two slow mechanisms, notably, the ORR and OER kinetics of oxygen catalyst [2]. Nowadays, Pt and IrO2 are considered as the standard electrocatalysts about ORR and OER. Nevertheless, the high expensive, disappointing lifetime and limited availability badly limit their practical applications [3,4]. Therefore, the current research on the alternative capable catalysts based on non-precious metal/metal composites, developing low-cost catalyst materials, practically both ORR and OER are active bifunctional electrocatalyst with high demand [5].

Among the different bifunctional catalysts, lanthanum-based perovskite structure is composed of many oxides which are represented by a chemical formula ABO3. In this case perovskite structure, the A-cation is generally alkaline-earth elements or rare-earth elements on the concerns of the lattice. B-cation on the lattice’s concern, which may be transition metal elements. Usually, the metallic A-cation is coordinated by 12 neighboring oxygen ions, similarly altering the insignificant B-cation is coordinated by 6 oxygen ions only. Appropriate substitution of A or B cations can guide to oxygen deficiencies without alternating the basic perovskite structures. ABO3 perovskite oxides have been applicable as cathode electrocatalysts due to their excellent stability in alkaline solution and high catalytic activity for ORR/OER. In this present work LaCo1−xFexO3 perovskite oxide nanoparticles was synthesized by sol-gel method [6]. LaCoO3 based materials contain very interesting electrocatalytic and electrical properties due to their good ionic conductivity and with high electronic conductivity [7]. Zhu et.al introducing the effect of A-cation defect into LaFeO2 enhances the best ORR and OER catalytic activity and, therefore, the best bifunctionality. This can be ascribed to the formation of abundant oxygen gaps and a certain amount of the Fe3+ species [8]. B-site cation with a reducible primitive transition metal like Mn, Fe, Co provides the activity of the catalytic reaction. Coupling at the B-site leads to the synergistic effect of two different ions, with enhanced electrocatalytic activity. The perovskite oxides activity was associated with redox properties and the high oxygen mobility of the transition metal used [9-11]. However, there has been a significant contribution to the investigation of lanthanum-based perovskite oxide (LaMO3, M = Co, Fe, Mn) in the application of a bifunctional air cathode [12]. This is because, in the present work, we summarized comprehensive research for the sol-gel synthesis of lanthanum-based perovskite oxide LaCo0.9Fe0.1 nanoparticles and their electro-catalytic activities characterization for ORR and OER reactions.

II. EXPERIMENTAL SECTION

A. Materials

Lanthanum nitrate hexahydrate (La(NO3)3·6H2O, 99+%, Aldrich), cobalt nitrate hexahydrate (Co(NO3)2·6H2O, 97.7%, Alfa Aesar), iron nitrate nonahydrate (Fe(NO3)3·9H2O, 98+%, Alfa Aesar), citric acid anhydrous (C6H8O7, 99.5+%, Alfa Aesar), potassium hydroxide (KOH, 85%, Alfa Aesar). These chemicals are used for received without further method.

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3658
B. Synthesis of LaCo$_{1-x}$Fe$_x$O$_3$ Nanoparticles

For the synthesis of Lanthanum-based nanoparticles, in a series of [LaCo$_{1-x}$Fe$_x$O$_3$] (x = 0, 0.2, 0.4, 0.6, 0.8, 1) were synthesized by the sol-gel method. Lanthanum, cobalt, and iron nitrate compounds were dissolved in deionized water. Then added citric acid to the mixed solution. The molar ratio of the metal nitrates and citric acid was set to 1:2. After that, all mixed compound solution was kept stirring with a magnetic stirrer at the temperature of 80 °C. Continue this stirring process until the brown gel formation. Subsequently, these prepared gel samples were placed in a heated furnace at 180 °C to form amorphous precursor, then the dried powder was transferred to a crucible and calcined at 400 °C for 1 h and 900 °C for 7 h. The obtained perovskite structure nanoparticles are LaCo$_3$, LaCo$_{0.8}$Fe$_{0.2}$O$_3$, LaCo$_{0.6}$Fe$_{0.4}$O$_3$, LaCo$_{0.4}$Fe$_{0.6}$O$_3$, LaCo$_{0.2}$Fe$_{0.8}$O$_3$, and LaFe$_3$O$_4$.

![Flow Chart](Image)

Figure 1. The flow chart shows the fabrication of the LaCo$_{1-x}$Fe$_x$O$_3$ compound by the sol-gel method.

C. Characterization Techniques

a. X-ray diffractometer

The crystal structure of synthesized LaCo$_{1-x}$Fe$_x$O$_3$ perovskite oxide nanoparticles were characterized by X-ray diffraction (Rigaku Cu Kα radiation with a wavelength of λ = 1.54 Å) technique and 2θ range from 10° to 80°.

b. Field Emission Scanning Electron Microscopy (FESEM)

The mean particle size and morphology of LaCo$_{1-x}$Fe$_x$O$_3$ perovskite oxide nanoparticles were characterized by field emission scanning electron microscopy (FESEM) (SIGMA 500, ZEISS) and the distribution of elemental analysis was characterized by energy dispersive spectroscopy (EDS) mapping (Bruker, XFlash 6130).

c. Electrochemical Measurements

The synthesized LaCo$_{1-x}$Fe$_x$O$_3$ perovskite oxide nanoparticles for ORR and OER activities, the electrochemical measurements were examined using a rotating ring disk electrode (RRDE) technique connected to the workstation (BAS Inc.). An RRDE is a saturated three-electrode test system for an electrochemical cell, where includes a glassy carbon (GC) substrate diameter of 3.0 mm, a saturated calomel electrode (SCE), and a graphite rod diameter of 6.0 mm. This three-electrode system play a role of reference, working and counter electrodes. The molar ratio of 0.1 M aqueous KOH solution is used as the electrolyte. The catalyst ink solution was prepared by careful mixing of LaCo$_{1-x}$Fe$_x$O$_3$ perovskite oxide (2 mg) and Vulcan-XC72 carbon (2 mg) were diluted in 1 mL of ethanol solution mixed with 0.15 mL of Nafion® (perfluorinated resin solution, 5 w%, Aldrich) by using an ultrasonic cleaner (model: SD200H) to form a homogeneous ink. RRDE glassy carbon disk electrode was earlier polished with a 0.05 μm alumina abrasive suspension (R&B Inc.) on a clean polishing cloth (R&B Inc.) subsequently rinsed with distilled water, and then dried in the oven. After completion of ultrasonication for 1h, 0.36 μL of catalyst ink was coated on to the glassy carbon electrode. Then RRDE experiments and high purity oxygen gas (99.99%) were purged to allow the electrolyte to saturate with oxygen for 30 min. The polarization curves for the oxygen reduction reaction of LaCo$_{1-x}$Fe$_x$O$_3$ was tested at the rotation speeds of 500, 1000, 1500, 2000, and 2500 rpm, and oxygen evolution reaction (OER) of LaCo$_{1-x}$Fe$_x$O$_3$ at the rotational speed of 1500 rpm, electrolyte using a CHI electrochemical workstation with a scan rate is 0.005 V s$^{-1}$. The measurements of the electrochemical impedance were carried out with a voltage bias of 5 mV amplitude in the frequency range between 0.1 Hz to 100 kHz. The electron transfer number (n) is the overall reaction was to understand the reaction mechanism of ORR by using the Koutecky-Levich equation:

\[
\frac{1}{I} = \frac{1}{I_k} + \frac{1}{I_d} + \frac{1}{nFAlCo_2} + \frac{0.62nFAD_0^{2/3} \omega^{1/2}v^{-1/2}cCo_2}{1}
\]

Where I is the measured current, $I_k$ and $I_d$ are the kinetic and diffusion-limiting current densities, respectively n is the number of electrons exchanged in the reaction, F is the Faraday constant, A is the geometrical surface area (1.3428 × (1.5 mm)$^2$), k is the rate constant for electron-transfer, $C_o$ is the bulk concentration of oxygen dissolved in the electrolyte (1.14×10$^{-6}$ mol cm$^{-3}$), $D_o$ is the diffusion coefficient of oxygen in the bulk solution (1.73×10$^{-5}$ cm$^2$ s$^{-1}$), $\omega$ is the rotation rate (rad/s), $\nu$ is the kinetic viscosity of the electrolyte (0.01 cm$^2$ s$^{-1}$).

III. RESULTS AND DISCUSSION

The crystal structure of the synthesized perovskite oxide nanoparticles LaCo$_{1-x}$Fe$_x$O$_3$ calcined at 400 °C for 1 h and 900 °C for 7 h were examined by using X-ray diffraction (XRD), as shown in Fig.2(a).
All the samples exhibited the single-phase perovskite cubic crystal structure. Conformed from pure LaCoO$_3$ JCPDS-ICDD file no: 048-0123. No characteristic peaks of lanthanum oxide or cobalt oxide were observed, indicating that the variation of the La/Co molar ratio did not impair the phase structure. Interestingly, we found that there was a clear peak change towards lower 2θ angle (left shift) with increasing of Fe content doping into LaCoO$_3$ such as (0, 0.2, 0.4, 0.6, 0.8, 1) [13].

The surface morphology of the synthesized LaCo$_{1-x}$Fe$_x$O$_3$ perovskite oxide nanoparticles was examined by using field emission scanning electron microscopy (FESEM) as shown in Fig. 3(a-f). The particle size of the components is shown in Fig. 3(a-f) which are 441.3 nm, 641.53 nm, 433.86 nm, 393 nm, 327.83 nm, and 280.53 nm. The particles exhibit random shapes and non-uniform distribution. The modification in sizes of particle probably perovskite attributed to a combination of changes. Fig. 3(g) shows the EDS elemental mapping of perovskite oxide LaCo$_{0.8}$Fe$_{0.2}$O$_3$ under FESEM mode which represents the occupancy of 1:1 atomic ratio for La: M (Co, Fe) catalysts. Every chemical element is gradually distributed in the perovskite oxide sample.

Figure 2. XRD patterns of the (a) LaCo$_{1-x}$Fe$_x$O$_3$ over a 2θ range of 10-80°, (b) enlarged XRD peak between 30° and 35°.

In order to evaluate the activities of the LaCo$_{1-x}$Fe$_x$O$_3$ perovskite oxide of the network structure for ORR and OER. Oxygen saturated 0.1 M KOH electrolyte was evaluated in linear sweep voltammetry (LSV) with rotating ring disk electrode (RRDE) measurements with different rotational speeds of 500, 1000, 1500, 2000, and 2500 rpm. Fig. 4(a) shows the LaCo$_{1-x}$Fe$_x$O$_3$ perovskite nanoparticles in the ORR current densities of the six samples with a scan rate is 0.005 V s$^{-1}$ under a rotation speed of 1500 rpm in potential range from -1.0 V to 0.0 V vs. SCE. It indicates the maximum diffusion-limited current density for LaFeO$_3$ that decreases in the sequence of LaCo$_{0.6}$Fe$_{0.4}$O$_3$ > LaCo$_{0.8}$Fe$_{0.2}$O$_3$ > LaCo$_{0.2}$Fe$_{0.8}$O$_3$ > LaCo$_{0.4}$Fe$_{0.6}$O$_3$ > LaCo$_{0.8}$Fe$_{0.2}$O$_3$ > LaFeO$_3$. The onset potential of six samples is nearly -0.30 V. The OER operations for models prepared to evaluate their application as bifunctional oxygen electrocatalysts. LSV curves recorded for OER activity in Fig. 4(b) shows an OER explanation of every catalyst in the potential between 0.0 V
Transition Metal Oxide-Based Perovskite Structures as a Bifunctional Oxygen Electrocatalysts: Fe Doped LaCoO3 Nanoparticles

Figure 4. (a) Linear sweep voltammetry (LSV) test on ORR polarization curves at 1500 rpm, (b) OER current densities in the potentials range from 0.0 V to 1.0 V, (c) ORR current density of LaFeO3 at different rotation speeds of 500 to 2500 rpm, (d) Koutecky-Levich analysis of all samples based on ORR curves at -0.80 V.

IV. CONCLUSIONS

In summary, LaCo1-xFexO3 was synthesized by using a sol-gel method that decreases the valence state of the surface Co and increases the valence state of the surface Fe. The experimental results indicate the LaFeO3 perovskite oxide with cubic crystal structure, there is a clear peak shift towards lower 2θ angle with increasing of Fe content doping into LaCoO3 such as (0, 0.2, 0.4, 0.6, 0.8, 1), and particle size of the corresponding single or double layers has shown several hundreds of nanometers. All the synthesized perovskite oxides were evaluated in 0.1 M KOH electrolyte for electrochemical tests towards ORR and OER. Based on the LSV results, LaFeO3 shows the maximum diffusion-limited current density for ORR. The OER explanation of every catalyst in the potentials range from 0.0 V to 1.0 V vs. SCE. At 0.65 V the kinetics of LaCo0.4Fe0.6O3 is abruptly increasing. Similarly, according to the Koutecky-Levich results on the ORR polarization curves of all compounds indicates a four-electron pathway. Fe-doped LaCoO3 provides a low-cost electrocatalyst for air cathode in metal-air batteries.

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Transition Metal Oxide-Based Perovskite Structures as a Bifunctional Oxygen Electrocatalysts: Fe Doped LaCoO3 Nanoparticles