BIOLOGICAL EFFICACY OF MIXED LIGAND COPPER(II) COMPLEXES OF N(4)-SUBSTITUTED THIOSEMICARBAZONE AND DIIMINE CO-LIGANDS AS CHELATING N, N’- DONOR LIGAND

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ABSTRACT

Objective: In modern years, the biggest dare is to develop new chemotherapeutic agents with high efficiency as well as low toxicity. Hence, to defeat this challenge, extensive endeavors were taken on thiosemicarbazone based metal complexes.

Methods: A sequence of nine thiosemicarbazone based diimine copper(II) complexes derived from three sulfur-containing ligands: N-methyl-2-((1-methyl-1H-pyrrol-2-yl)methylene)hydrazinecarbothioamide H(L1), N-ethyl-2-((1-methyl-1H-pyrrol-2-yl)methylene)hydrazinecarbothioamide H(L2), and N-benzyl-2-((1-methyl-1H-pyrrol-2-yl)methylene)hydrazinecarbothioamide H(L3). The molecular structures and coordination possibilities of thiosemicarbazone ligands toward the central metal ions have been validated by analytical and spectral techniques such as molar conductance, elemental analysis, ultraviolet–Vis, Fourier-transform infrared, and electron paramagnetic resonance spectroscopy confirms that the thiosemicarbazone ligands are coordinated to copper through N5 donor and N6 donor of diimine. The antimicrobial activity and DNA cleavage effectiveness of thiosemicarbazone derivatives and copper(II) complexes were assessed by disc diffusion and agarose gel electrophoresis methods.

Results: All the spectral studies authenticated that the square planar geometry of the thiosemicarbazone copper(II) complexes 1–9. From the results of antibacterial activity against selected Gram-positive (Bacillus thuringiensis) and Gram-negative (Escherichia coli) bacterial strains, complex 8 exhibited noteworthy activity. Interestingly, copper(II) complexes bearing 1,10-phenanthroline (phen) moiety displayed outstanding results together with N4-substituted thiosemicarbazone derivatives and causes complete DNA degradation of SC (superciled) pUC18 DNA.

Conclusion: A variety of substitutions at the thiourea nitrogen atoms have shown potential biological activity. Henceforth, N(4)-substituted thiosemicarbazone based copper(II) complexes virtually reach the effectiveness of conventional chemotherapeutic drugs.

Keywords: Thiosemicarbazide, SC pUC18 DNA, Ascorbic acid, 2,2'-bipyridyl, Heterocyclic bases.

INTRODUCTION

Cancer is one of the world’s top killer diseases and it has overtaken heart disease according to 2010 reports, which should be more than higher and causes lots of deaths by 2030. To defeat the drawback of limited usefulness, toxicity and side effects of the commercially available anticancer drug, accordingly good alternatives were made through the discovery of new drugs for cancer therapy [1,2]. However, the mixed N/S donors of thiosemicarbazone ligands derived from the combination of an aldehyde or ketone and these groups of ligands are useful for achieving coordination spheres of transition metal complexes. These ligands and their metal complexes have received the extensive notice due to their availability of variable bonding modes, an electron donor’s property of ligand, the position of a ligand inside the coordination sphere, structural variety, ion sensing capability, potential biological applications, and the stability of ligands and complexes [3,4]. In the last few decades, a huge number of thiosemicarbazone chemotherapeutic anticancer drug are reported and some of the compounds are various phases of clinical trials [5]. Significantly, the biological properties of thiosemicarbazone compounds are comparatively higher than that of the other transition metal complexes probably due to the synergic effect among thiosemicarbazone ligands and the presence of central metal ion [6]. Recently, Zhu et al. found that potential biological properties of novel N(4)-substituted thiosemicarbazide with pyrrole moiety against PC-9, Eca-109, and SGC-7901 cell lines [7]. Moreover, the substituted thiosemicarbazone at N(4) position did not show any effect on the geometry of the complexes, but it influences the biological activity of the corresponding complexes and shows significant mounting evidence from literature, the synthesis of 1-methyl-2-pyrrolocarboxaldehyde with thiosemicarbazide in the presence of ethanol has been reported by Yurttaş et al. [8], and it is planned to modify with substituted thiosemicarbazide derivatives using glacial acetic acid, ethanol, and water. Likewise, phenyl and ethyl substituted thiosemicarbazone ligands were prepared. However, this was the first report which proved the biological properties of the mixed ligand thiosemicarbazone copper(II) complexes. Consequently, the biological potential of those substituted ligands H(L1) – H(L3) and its mixed ligand copper(II) complexes 1–9 with diimine coligands (phen/bpy) have been studied and reported for their future prospects. It is hoped that this work will be afforded some potential biological properties of mixed ligand thiosemicarbazide copper(II) complexes.

MATERIALS

Chemicals used
All the commercially available chemicals such as 1-methyl-2-pyrrolocarboxaldehyde, 4-methyl-3-thiosemicarbazide, 4-phenyl-3-thiosemicarbazide, 4-ethyl-3-thiosemicarbazide, 1, 10-phenanthroline, and 2, 2’-bipyridyl were purchased from Sigma-Aldrich (USA). The chemicals, ethidium bromide (EB), Tris-HCl, agaroze (molecular biology grade), and boracic acid used for the DNA cleavage studies were procured from Merck (USA). The solvents and reagents for the synthesis of ligands and copper(II) complexes were obtained from Sigma-Aldrich (USA), Merck (USA), and Reachem (India), respectively.

METHODS

Instrumentation
The molar conductance of the synthesized thiosemicarbazone ligands and copper(II) complexes was determined using Elico CM 183 EC-TDS analyzer.
in Dimethyl formamide (DMF) (1×10<sup>−3</sup> M) at room temperature. Elemental analyses (C, H, N, and S) of the compounds were attained by Vario EL-III elemental analyzer. In addition, the structures of the thiosemicarbazone ligands were confirmed by Bruker 300 MHz spectrophotometer using CDCl<sub>3</sub> as a solvent and TMS (Tetramethylsilane) as an internal standard. The Fourier-transform infrared (FT−IR) spectra were carried out on a SHIMADZU IR TRACER 100 spectrophotometer using KBr discs in the range 4000–400 cm<sup>−1</sup>. The ultraviolet (UV−Vis) spectra of the ligands and copper(II) complexes were recorded using JASCO V-630 UV-Vis spectrophotometer in the range 200–1000 nm. The X-band electron paramagnetic resonance (EPR) of the copper(II) complexes was recorded using ESR JEOI JES-FA200 in frozen solutions of DMF. The room temperature measurement was made by VARIAN ESR-112 spectrometer.

**BIOLICAL EVALUATION**

**Antibacterial screening**

The minimum inhibitory concentration (MIC) of synthesized thiosemicarbazone ligands and copper(II) complexes (5 and 50 µg/ml) was investigated by agar disc diffusion [9,10] against Pseudomonas aeruginosa (Gram-negative) and Escherichia coli (Gram-negative) and reported in terms of mm. However, the test was performed in the agar medium with the positive (chloramphenicol) and negative controls (DMF). The disc of 6 mm was dipped into the compounds that were placed onto the sterilized nutrient agar medium inoculated with microorganisms. Then, each plate was incubated at 37°C for 24 h, the zone of inhibition was measured.

**Nucleosome activity on pUC18 DNA**

The extent of DNA cleavage activity of SC pUC18 DNA on thiosemicarbazone ligands and copper(II) complexes was monitored using agarose gel electrophoresis [11] in 5 mM Tris–HCl/50 mM NaCl buffer (pH 7.2) in the presence of ascorbic acid and incubated for 1 h. The test samples were electrophoresed with SC pUC18 DNA by varying the concentration (40, 60 µM). It was loaded on 1.2% agarose gel and dried at 90°C, the solution was kept allowed to stand overnight at room temperature until a yellow precipitate was separated out. Thus, the carbothioamide H(L1)

**EXPERIMENTAL**

### Synthesis of thiosemicarbazone ligands H(L1) – H(L3)

1-methyl-2-ppyrrole carboxaldehyde (0.10 g, 10 mmol) in a hot ethanol solution, and 4-methyl-3-thiosemicarbamide (0.10 g, 10 mmol) in water (30 ml) and glacial acetic acid (2 ml) were added dropwise and refluxed. After refluxing the above solution for 2 h on a water bath at 90°C, the solution was kept allowed to stand overnight at room temperature until a yellow precipitate was separated out. Thus, the yellow precipitate obtained was washed thoroughly with cold ethanol and dried in vacuo over anhydrous CaCl<sub>2</sub>. It was then recrystallized from chloroform. Color: Yellow; Yield: yellow; yield: 60%; m.p 116–118°C; Anal. Calcd for: C<sub>10</sub>H<sub>12</sub>N<sub>2</sub>S (H(L1)): C, 48.96; H, 6.16; N, 28.55; S: C, 16.34% found: C, 48.92; H, 6.14; N, 28.61; S, 16.35%.

**Synthesis of 2-((1-methyl-1H-pyrrol-2-yl)methylene)-N-phenylhydrazine carbothioamide H(L2)**

To a hot ethanolic solution of ligands H(L1) – H(L3) (1 mmol), an excess of hydrazine hydrate (50 µl) was added dropwise with constant stirring for 1 h. The resulting green solution was incubated at 37°C for 24 h, the zone of inhibition was measured.

**Synthesis of 2-((1-methyl-1H-pyrrol-2-yl)methylene)-N-phenylhydrazine carbothioamide H(L3)**

To a hot ethanolic solution of ligands H(L1) – H(L3) (1 mmol), an excess of hydrazine hydrate (50 µl) was added dropwise with constant stirring for 1 h. The resulting green solution was incubated at 37°C for 24 h, the zone of inhibition was measured.

### Synthesis of copper(II) bis complexes (1–3)

A hot ethanolic solution of ligands H(L1) − H(L3) (1 mmol), an ethanolic solution of copper(II) chloride dihydrate (2 mmol) was added dropwise with constant stirring for 1 h. The resulting green solution was incuuted to stand overnight. Later, green precipitate formed was collected by filtration, washed thoroughly with ethanol and dried in vacuo over anhydrous CaCl<sub>2</sub>. The similar method was applied for the synthesis of complexes 2 and 3 using H(L2) and H(L3) instead of H(L1). The schematic representation of synthesis of copper(II) bis complexes is given in Scheme 1.

[Scheme 1: Schematic representation of synthesis of copper(II) bis complexes 1–3]
Synthesis of mixed ligand copper(II) complexes (4–9)

To a hot ethanolic solution of ligand H[L1] (1 mmol), an ethanolic solution of dimine coligands (bpy/phen) was added dropwise with constant stirring for 1 h. To this resulting yellow solution copper(II) chloride dihydrate (2 mmol) in ethanol was added and the stirring was continued for about an hour. The resulting green solution was allowed to stand overnight. Later, green precipitate formed was collected by filtration, washed thoroughly with ethanol and dried in vacuo over anhydrous CaCl₂. The schematic representation of synthesis of mixed ligand copper(II) complexes is given in Scheme 2.

[Cu(L)₂(bpy)]Cl₂ (4)
Color: Green; yield: 62%; Anal. Calcd for: C₁₈H₁₉ClCuN₅S₂; C, 48.00; H, 4.25; N, 18.66; S, 7.12%; Found: C, 47.98; H, 4.29; N, 18.74; S, 7.18%.

[Cu(L)₂(bpy)]Cl₂ (5)
Color: Green; yield: 79%; Anal. Calcd for: C₁₈H₁₉ClCuN₅S₂; C, 49.13; H, 4.56; N, 18.09; S, 6.90%; Found: C, 49.24; H, 4.61; N, 18.14; S, 6.83%.

Electronic spectra
The UV-Vis measurement of the synthesized thiosemicarbazone ligands and their copper(II) complexes in the 1×10⁻³–1 M concentration in DMF was performed in the wavelength range 200–1000 nm. The absorption bands of thiosemicarbazone ligands showed two intense bands close to 280 and 350 nm and designated as π → π* and n → π*, respectively. The low energy band at 280 nm associated with the imine group of thiosemicarbazone. This band has undergone a bathochromic shift in the spectra of the thiosemicarbazone copper(II) complexes as a result of the donation of the lone pair of electrons to the central metal ion [15]. While the band at 300–330 nm has suffered a blue shift in copper(II) complexes as a result of thiolation. These shifts in the spectra of the complexes yet again proved the coordination of azomethine nitrogen and thiocarbamide function to copper. Moreover, all the thiosemicarbazone copper(II) complexes exhibited much significant characteristics d → d transitions (Fig. 1) [16]. From the electronic spectral data square planar geometry was proposed for all the thiosemicarbazone copper(II) complexes.

FT-IR spectroscopy
The most useful main functionalities of the free thiosemicarbazone ligands H[L1] – H[L3] and copper(II) complexes 1–9 were established by FT-IR spectroscopy. The FT-IR spectra of uncomplexed thiosemicarbazone ligands showed the band around 820 cm⁻¹ attributable to the presence of v(C=S) in the solid state. In this semicarbazone copper(II) complexes, v(C=S) disappeared as a result of deprotonation of hydrazinic proton of v(N(2)–H) by thiolation.

Scheme 2: Schematic representation of synthesis of mixed ligand copper(II) complexes

Fig. 1: Electronic spectra of complex 9 in DMF
The antibacterial screening showed that following: different bacterial strains when compared to chloramphenicol. On the zone of inhibition for both Gram-positive and Gram-negative bacterial complexes was screened using disc diffusion by the determination of the antibacterial activity of the thiosemicarbazone copper(II) complexes [24].

The antibacterial activity of the thiosemicarbazone copper(II) complexes 4–9 in DMF were recorded at room temperature as well as liquid nitrogen temperature (77 K) on the X-band EPR at frequency 9.1GHz. The room temperature and frozen X-band EPR spectra of the complexes were quite similar and showed axial type g-tensor parameter with $g_x > 2.0023$ designated that the presence of one unpaired electron in $d^2_{1/2}$ orbital of $B_{1g}$ ground state. In addition to this, the exchange interaction parameter of G could be obtained by the following expression: $G = (g_0 - 2)/(g_0 - 2) = 3.7−4.6$. When the G < 4 indicating exchange interaction in solid copper(II) center was negligible, if G > 4 indicating considerable exchange interaction in the copper(II) complexes. In all the copper(II) complexes, the G value found within 3.77−4.23, specified negligible exchange coupling effect. In addition, $g_z$ afforded valuable information regarding the nature of the metal-ligand bonding [18,19]. According to Hathaway [21], for an ionic environment $g_z$ of the copper(II) complexes have been usually ±2.3 but the value was <2.3 suggesting a covalent environment of the copper(II) complexes [20] while $g_z$ value of the synthesized complexes was achieved <2.3 signify the covalent character of the copper-ligand bonding. In the X-band EPR spectra of the complexes showed a quartet hyperfine splitting at the $g_z$ region and corresponding to the nuclear spin ($I=3/2$). Therefore, the obtained $g_z$ (2.260) > $g_x$ (2.069)>2.0023 of complex 4 (Fig. 2), signify that the copper ion lies in $d^2_{1/2}$ which indicates the characteristics of square planar geometry.

In general, CuN₅ chromophore of square-based geometry was expected to show $g_z$ and $A_z$ values around 2.200 and 180−200×10⁻⁴ [22]. However, the synthesized thiosemicarbazone copper(II) bis complexes and mixed ligand copper(II) complexes displayed $g_z$ (2.258−2.264, 2.260−2.268) and $A_z$ (163−164, 163−165), respectively. These increases in the $g_z$ and decreases in the $A_z$ value possibly due to the replacement of two coordinated nitrogen by two molecules of sulfur. Hence, the observed values reliable with the CuN₅S₅ and CuN₅S chromophore of bis complexes and mixed ligand copper(II) complexes. This was further supported by the value of $f g_z / A_z$ quotient (where $f$ is the index of tetragonal distortion) fall within the range of 1130−138×10⁻⁴ cm characteristics of square planar geometry. Here, the obtained $f$ value of the copper(II) complexes from 136−139×10⁻⁴ cm indicating perfect square planar geometry [23]. This was yet again supported by the neutral and 1:1 electrolyte nature of the thiosemicarbazone copper(II) complexes [24].

**Antibacterial activity: Minimum inhibitory concentration**

The antibacterial activity of the thiosemicarbazone copper(II) complexes was screened using disc diffusion by the determination of zone of inhibition for both Gram-positive and Gram-negative bacterial strains. All the test compounds showed different results against two different bacterial strains when compared to chloramphenicol. On the evaluation of the antibacterial screening showed that following:

- The results concerning *E. coli* thiosemicarbazone copper(II) complexes showed moderate activity at 50μg/ml concentration, and some of the mixed ligand complexes 7 and 8 displayed significant activity at higher concentration (50 μg/ml) than the bis copper(II) complexes under experimental conditions. Thus, the presence of phen/bpy considerably changes the antibacterial activity of the complexes.
- At higher concentration, thiosemicarbazone complex 8 was strongly active and showed improved results in *B. thuringiensis*. However, the zone of inhibition of other copper(II) complexes was comparatively lower.
- The positive results signified that on chelation the polarity of the copper(II) complexes reduced by partial sharing of positive charge with the donor groups over the chelating system. This was predominantly increased the lipophilic nature of the copper(II) complexes and accordingly terminating the growth of the microorganism by potent bacteriostatic effects [25,26].
- In view of that the other principle factors including, geometry, stability, and coordinating sites of the tested compounds which also enhances the antibacterial activity.
- Hence, the present investigation revealed that the MIC values of the mixed ligand thiosemicarbazone complex eight against both bacterial strains were as effective at a maximum concentration and they need a higher concentration to wipe out the tested bacterial strains. This may possibly supportive to open up a novel pathway for developing new antibacterial drugs (Fig. 3).
In recent years, many anticancer agents could be reported as a result of induced photocleavage activity by the strong DNA interaction and DNA strand scission. For this reason, the extent of DNA cleavage activity of thiosemicarbazone copper(II) complexes has been studied for the purpose of generating novel anticancer drug by triggering cell death via DNA damage. To identify the extent of DNA cleavage activity of the complexes 1–9 with SC pUC18 DNA was incubated with 5 mM Tris-HCl/50 mM NaCl (pH 7.2) buffer with increasing concentrations 40–60 µM was studied using agarose gel electrophoresis. As seen in Figs. 4a and b, there was no noticeable cleavage in control experiments (DNA and DNA with ascorbic acid). When the concentration-dependent DNA cleavage was performed at 40 µM concentration with ascorbic acid, only a faint and moderate nuclease activity was noted under identical conditions. While complete DNA degradation was observed in complexes 7 and 8 which converts the SC DNA into undetectable minor fragments at 60 µM concentration.

It is interesting to note that as the concentration increases the amount of Form I (SC DNA) decrease, whereas both Form II (NC DNA) and Form III (Linear) increases correspondingly. Moreover, the majority of the mixed ligand complexes exhibit efficient oxidative cleavage higher than copper(II) bis complexes. In addition, complexes with dimine co-ligands act as a DNA recognition element and engage strong DNA interaction with an extended aromatic ring of bpy and phen [27]. It was apparently due to the partial intercalation and deeper insertion of the dimine coligands would predominately increase the DNA cleavage activity [28]. These complexes were proper candidates to make an effort for the development of potential anticancer drugs.

CONCLUSION

In the present study, nine thiosemicarbazone copper(II) complexes 1–9 were synthesized and characterized by various spectro analytical techniques, and it was coordinated through azomethine nitrogen and thiolate sulfur through deprotonation. The EPR spectral studies confirmed that square planar geometry of the bis and mixed ligand copper(II) complexes. The biological activity of the compounds displayed a higher antibacterial activity for mixed ligand copper(II) complex than bis complexes. This possibly due to the geometry and the effect of the existence of phen and N(4)-substituted thiosemicarbazide derivatives. Correspondingly, complex 8 completely degrade the SC pUC18 DNA in the presence of the activating agent. Only fewer amounts of metal complexes with phen were reported as potential nuclease agent. Interestingly, the existing study revealed noteworthy nuclease activity on mixed ligand complex 8 by strong DNA interaction even at the minimum concentration (60 µM). Consequently, the significant biological activities of complex 8 have been further studied for the purpose of effective chemotherapeutic drug toward different cancer cells.

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AUTHOR’S CONTRIBUTIONS

All authors have equally contributed to this work regarding experimental design, performing experiments, data completion, and progress of the manuscript.

CONFLICTS OF INTEREST

There are no conflicts of interest for all the authors regarding this article.

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