Stable luminescent iridium(III) complexes with bis(N-heterocyclic carbene) ligands: photo-stability, excited state properties, visible-light-driven radical cyclization and CO₂ reduction, and cellular imaging†

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A new class of cyclometalated Ir(III) complexes supported by various bidentate C-deprotonated (C^N) and cis-chelating bis(N-heterocyclic carbene) (bis-NHC) ligands has been synthesized. These complexes display strong emission in deaerated solutions at room temperature with photoluminescence quantum yields up to 89% and emission lifetimes up to 96 µs. A photo-stable complex containing C-deprotonated fluorenyl-substituted C^N shows no significant decomposition even upon irradiation for over 120 h by blue LEDs (12 W). These, together with the strong absorption in the visible region and rich photo-redox properties, allow the bis-NHC Ir(III) complexes to act as good photo-catalysts for reductive C–C bond formation from C(sp³/sp²)–Br bonds cleavage using visible-light irradiation (λ > 440 nm). A water-soluble complex with a glucose-functionalized bis-NHC ligand catalysed a visible-light-driven radical cyclization for the synthesis of pyrroline in aqueous media. Also, the bis-NHC Ir(III) complex in combination with a cobalt catalyst can catalyse the visible-light-driven CO₂ reduction with excellent turnover numbers (>2400) and selectivity (CO over H₂ in gas phase: >95%). Additionally, this series of bis-NHC Ir(III) complexes are found to localize in and stain endoplasmic reticulum (ER) of various cell lines with high selectivity, and exhibit high cytotoxicity towards cancer cells, revealing their potential uses as bioimaging and/or anti-cancer agents.

Introduction

Luminescent organometallic complexes of 3rd row transition metals, such as Ir, Pt, and Au are currently receiving burgeoning interest due to their profound applications in materials science,20–22 biology,23,24 and organic synthesis.25–27 In particular, the favourable emission properties of cyclometalated Ir(III) complexes have been harnessed for applications by many research groups with examples such as the development of high performance OLEDs by Thompson and co-workers3–5,28–31 and the design of bioimaging and cellular probes by Lo and co-workers.32–36 Over the past decade, as a result of extensive work most notably by MacMillan,37–39 Yoon,40–41 and Stephenson and co-workers,42–45 luminescent cyclometalated iridium(III) complexes are widely used in photo-redox catalysis, which has been shown to have useful applications in organic synthesis.

In 2008, Yoon and co-workers46 reported [2 + 2] enone cycloadditions, and MacMillan and co-workers48 published the alkylation of aldehydes, both of which were catalysed by triplet metallo-ligand-charge-transfer (MLCT) excited state of [Ru(bpy)₃]²⁺ generated upon visible-light irradiation. Subsequently, Stephenson and co-workers achieved the reductive dehalogenation of activated alkyl halides catalysed by [Ru(bpy)₃]²⁺,49 and reductive dehalogenation of alkyl, alkenyl and aryl iodides by using fac-Ir(ppy)₃ as photo-redox catalyst.50 Compared to [Ru(bpy)₃]²⁺ and fac-Ir(ppy)₃, the application of luminescent platinum(II) photo-catalysts is still nascent. Recently, our group and Wu’s group demonstrated that pincer Pt(II) complexes are capable of catalysing light induced C–C bond formation.51,52

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The important features that allow luminescent transition metal complexes to act as useful photo-redox catalysts or photo-sensitizers for light induced reactions include: (1) the long lifetime of their electronic excited states, thus allowing bimolecular reaction to proceed in solution; (2) their electronic excited states as both strong reducing and oxidizing reagents with reduction potentials systematically varied by the auxiliary ligands.41 For the photo-catalysis to have practical interest, the design of highly stable photo-redox catalysts with long-lived electronic excited states in solution is desirable.

Our endeavour to develop transition metal photo-catalysis is to use visible light for activation of small molecules such as CO2, described in this work, and for C–X bond functionalization. A major consideration is to utilize visible light, which falls within the solar spectrum and avoids deleterious high-energy UV-initiated photochemical side reactions.38,49 In the literature, platinum group metal complexes and semi-conductors are usually used as photo-redox catalysts for the photochemical reduction of CO2.46 Earlier examples of transition metal photocatalysts used for the photochemical reduction of CO2 include cobalt porphyrins,31 Re[bpy](CO)3Cl53 and Ir(terpy)[ppy]Cl3.34 More recently, systems comprising fac-Ir(ppy)3 in conjunction with [Ni(P2bimiq1)]2+,55 Fe[porphyrin],56 [Co(TPA)Cl]3+,57 (TPA = tris(2-pyridylmethyl)amine) and [Co(N5)]2+ (N5 = 2,13-dimethyl-3,6,9,12,18-pentaazaacyclo[12.3.1]octadeca-1(18),2,12,14,16-pentaene)38 were reported for photochemical reduction of CO2.

The stability of photo-catalysts is an important issue for the practical application of transition metal photochemistry. Numerous studies revealed that the photochemically active excited states of [Ru(bpy)]32+,59-61 [Ir(ppy)2(bpy)]3+,60,62 and fac-Ir(ppy)3 are not stable under light irradiation for a long period of time as a result of dissociation of coordinated ligand(s) presumably via low lying d–d excited state(s). To address the photo-stability issue, we considered the use of N-heterocyclic carbene (NHC) ligands which have been receiving burgeoning attention in coordination chemistry due to their strong σ-donor strength to develop robust metal photo-sensitizers and photo-catalysts.64-70 Also, the N-substituent of NHC ligands can be used to tune both the physical and chemical properties of the resultant photo-active transition metal complexes such as their solubility in various solvents including water. NHC ligands functionalized with carboxylate, sulfonate, amine/ammonium, and alcohol motifs have been reported for the development of water-soluble transition metal catalysts for Suzuki coupling, hydrosilylation, hydrogenation, olefin metathesis and CO2 reduction.65,71

Compared with Ir(m) complexes supported by bidentate acetylacetone39 and/or 2,2’-bipyridine1 ligands, the photo-physical and application studies of the related bis-NHC Ir(m) complexes are relatively scarce. In 2010, cationic bis-NHC Ir(m) complexes were reported by De Cola and co-workers72 to have application in blue-light emitting electrochemical cells; subsequently, a number of bis-NHC Ir(m) complexes were reported for photophysical and biological studies.59,67,73-74 In this work, a series of strongly luminescent Ir(m) complexes (Chart 1) containing bis-NHC ligands and visible light absorbing C-deprotonated (C=N) ligands was synthesized and their photophysical and electrochemical properties were examined. These complexes display high photo-stability and are strongly emissive with long lifetimes of up to 96 μs in solution at room temperature. The water-soluble luminescent Ir(m) complexes, containing the glucose-functionalyzed NHC ligand, were found to be active photo-catalysts for radical cyclization leading to the formation of 5-membered pyrrole rings in aqueous media with high substrate conversions and yields. One of the photo-stable Ir(m) complexes was utilized as a photo-sensitizer and in conjunction with a recently reported catalyst [Co(TPA)Cl]Cl to convert CO2 into CO with a turnover number (TON) > 2400, selectivity in gas phase > 95% and yield of 5.6% (1 mL out of 18 mL of CO2 was converted into CO at 5 μM concentration of Co complex). Some of the complexes were also demonstrated as potential bioimaging and/or anti-cancer agents.

**Results**

**Synthesis, characterization and photo-stability of [(C=N)2Ir(NHC)]X complexes**

The structures of 18 bis-NHC Ir(m) complexes synthesized in this work (1a and 2a, 2b, 3–9), together with previously reported Ir(m) complexes 1a,73,74 1b,73 1d,28 and 2d,31 are depicted in Chart 1. Complexes 1–9 were synthesized in good yields by refluxing [(C=N)2Ir(μ-Cl)]X with bis(imidazolium) salts in the presence of silver(i) oxide in 2-methoxyethanol. Details of synthesis and characterization data of ligands and complexes are provided in the ESI.† Notably, 1H NMR spectra of 7b and 7c show poorly resolved peaks in the aromatic region (6.7–8.2 ppm) at ambient temperature, and hence 1H NMR spectra are recorded from 238 K to 300 K (Fig. S1, ESI†) in order to assure the purity of 7b and 7c.

Complexes 1a–4Mea, 1b–4Meb, 4Meo–9b with N-methyl or N-butyl substituent on bis-NHC ligands are soluble in most common aprotic solvents, but not in protic solvents e.g. methanol (MeOH) or water. Complexes 1c, 4MeC, 6Hc, 7c and 9c with glucose functionalized bis-NHC ligand are soluble in MeOH, ethanol (EtOH) and water.

The photo-stability of these Ir(m) complexes with bis-NHC ligands was examined by using 4Meb and 6Hb as representative examples. 4Meb and 6Hb in degassed deuterated MeCN were irradiated using blue light (12 W blue LEDs) for 5 days. The photolysis was monitored by 1H NMR spectroscopy. As depicted in Fig. 1a, in the case of 4Meb, less than 5% of the complex was observed to undergo photochemical decomposition after irradiation of 120 h, revealing its outstanding photo-stability. Under the same conditions, Ru(bpy)3Cl2, fac-Ir(ppy)3 and [(dFCF3ppy)Ir(btbpy)]PF6 (dFCF3ppy = 3,5-difluoro-2-[5-(tri-fluoromethyl)-2-pyridinyl]phenyl; dbtbpy = 4,4’-bis(1,1-dimethyl-ethyl)-2,2’-bipyridine) were found to decompose after 10 h irradiation as revealed by the changes of their 1H NMR spectra (Fig. 1c).

Interestingly, irradiation of 6Hb using blue LEDs for 10 h led to a clean and quantitative conversion to a new species which did not show further changes upon subsequent irradiation for another 90 h (Fig. 1b, this species is noted as cis-6Hb). The yield of scale synthesis of cis-6Hb from the irradiation of degassed
MeCN solution of $6^{\text{Hb}}$ (100 mg in 4 mL MeCN) was 94%. \textit{cis}-$6^{\text{Hb}}$ was found to be stable upon standing in solution in the dark for another 40 h, or exposed to air for another 20 h (Fig. S2, ESI†).

To verify that the transformation of $6^{\text{Hb}}$ to \textit{cis}-$6^{\text{Hb}}$ was caused by visible-light irradiation, a negative control experiment was conducted by keeping $6^{\text{Hb}}$ in deuterated MeCN in the dark for another 40 h.

Fig. 1 $^1$H NMR spectra for (a) $4^{\text{Meb}}$; (b) $6^{\text{Hb}}$; (c) Ru(bpy)$_3$Cl$_2$, fac-Ir(ppy)$_3$ and [(dFCF$_3$ppy)$_2$Ir(dtbbpy)]PF$_6$ in deuterated MeCN solution for irradiating by blue light (12 W, $\lambda_{\text{max}} = 462$ nm) (all solutions are degased by nitrogen gas for 10 min); (d) crystal structure diagrams showing the photoinduced transformation of coordination for Ir metals in $6^{\text{Hb}}$ and \textit{cis}-$6^{\text{Hb}}$ (butyl group on bis-NHC ligand and hydrogen atoms are omitted for clarity); (e) UV/Vis absorption (dotted line), excitation (dashed line) and emission (solid line) spectra of solutions of $6^{\text{Hb}}$ and \textit{cis}-$6^{\text{Hb}}$ (concentration of 2.0 $\times$ 10$^{-5}$ M) in degassed DCM at 298 K.

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(Fig. S2, ESIf), and no structural changes were detected by $^1$H NMR spectroscopy.

The UV-Vis absorption, emission and excitation spectra of cis-6b measured in DCM solution at 298 K are different from those of 6b with hypsochromic shifts in peak maxima (Fig. 1e). The emission band for cis-6b displays vibronic spacings of 1245 cm$^{-1}$ while that of 6b shows spacings of 908 cm$^{-1}$. Taking into consideration the ESI-MS and spectroscopic data of cis-6b, cis-6b is likely a structural isomer for 6b. The exact structure of cis-6b has been determined by X-ray crystallography.

**X-Ray crystallography**

Crystals of 4Meb (with PF$_6$ counter anion), 6b, cis-6b and 7b suitable for X-ray crystallographic analysis were obtained by slow diffusion of diethyl ether into DCM (6b, MeCN (4Meb and 7b) and chloroform (cis-6b)) solution of these complexes, respectively. Perspective views of 6b and cis-6b are shown in Fig. 1d (those of 4Meb, 7b are shown Fig. S3 in ESIf). Selected bond lengths and angles are compiled in Table S1 (ESI†). Similar to the reported examples, 6b adopts a distorted octahedral geometry (Fig. 1), the iridium atom is coordinated by two cyclometalated C$^\equiv$N ligands and one bis-NHC ligand. The two C$^\equiv$N ligands adopt a mutually eclipsed configuration, with the two N atoms (N1 and N2) trans to each other and with Ir–N bond length of 2.067(2) and 2.071(2) Å. The two substituted phenyl rings are oriented cis to each other with Ir–C bond length of 2.054(2) and 2.077(2) Å. The C$^{\equiv}$NHC atoms of bis-NHC ligand are trans to two C atoms from C$^\equiv$N ligands. The Ir–C$^{\equiv}$NHC distances of 2.105(2) Å and 2.128(2) Å are comparable to the literature values of Ir–C$^{\equiv}$NHC bonds trans to phenyl groups. The C$^{\equiv}$NHC–Ir–C$^{\equiv}$NHC bite angles of 4Meb (84.7(3)°), 6b (85.10(9)°), 7b (85.58(6)°) and that of reported [(dfppy)$_2$Ir(NHC$^{\equiv}$M)PF$_6$ (85.28(15)°)] (dfppy = 4,6-difluoro-phenylpyridine) are consistent. The structures of 4Meb and 7b are in line with that of 6b. Interestingly, for cis-6b the N atoms from the two C$^\equiv$N ligands are cis to each other (Fig. 1d), indicating that 6b and cis-6b are structural isomers. The crystallographic refinement parameters for 4Meb, 6b, cis-6b and 7b are summarized in Table S2 (ESI†).

**Electronic absorption and emission spectroscopy**

Photophysical data of 1–9 including their UV/Vis absorption and emission maxima, emission lifetimes and photo-luminescent quantum yields are tabulated in Table 1. All the complexes show intense high-energy absorptions ranging from 250 to 380 nm and lower-energy absorptions at 416–494 nm (Fig. 2, 3 and S4 in ESIf). The high-energy absorptions are attributed to the $^3\pi \rightarrow \pi^*$ transition of the ligands, whereas the low energy absorption bands should be ascribed to the admixture of metal-to-ligand charge transfer (MLCT) transitions and ligand centered (LC) $\pi$–$\pi^*$ transitions. The charge transfer (CT) assignment is supported by the energy trend observed in the complexes with different C$^\equiv$N ligands. For example, the low-energy absorptions of 5b and 5bf reveal significant red-shifts from those of 1bf and 4Meb respectively. This can be rationalized by the extended $\pi$-conjugation of the C$^\equiv$N ligand in 8b and 5bf, leading to the lowering of the $\pi^*$C$^\equiv$N orbitals and hence decreases in CT energy. These red-shifts in the UV-Vis absorptions, together with the large molar absorptivities in the visible region (e.g., 5bf: 7.2 × 10$^4$ M$^{-1}$ cm$^{-1}$), allow the complexes to show strong absorption in the visible region. This is believed to be crucial for harvesting visible light in the solar spectrum as well as avoiding UV-initiated side reactions due to the use of UV in the photo-catalysis by the Ir(III) complexes. As revealed by the UV absorption spectra of 6b and 6bf (Fig. S4b, ESIf), a bathochromic shift about 500 cm$^{-1}$ is observed for the lower-energy absorption band(s). These could be attributed to the electron-withdrawing group of CF$_3$ on the cyclometalated ligand which lowers the C$^\equiv$N based LUMO level.

On the other hand, the N-alkyl substituents on the bis-NHC ligand are found to have only minor effects, if any, on the UV/Vis absorption spectra of the Ir(III) complexes, as revealed by the overlaid spectra of the group of 1a, 1b, 1c (Fig. S4a, ESIf). Complex 1a absorbs weakly at the wavelength from 430 nm to 500 nm with molar absorptivity less than 500 M$^{-1}$ cm$^{-1}$ in contrast to the high values of 2500 M$^{-1}$ cm$^{-1}$ for 1d with the same C$^\equiv$N luminophore. The calculated wavelength for ground state HOMO → LUMO (S$_0 \rightarrow$ S$_1$ transition) is 369 nm and 414 nm for 1a and 1d, respectively (Fig. S4e and f, ESIf). These calculated values are in reasonable agreement with the corresponding experimental absorption $\lambda_{max}$ values 410 and 460 nm respectively.

Complexes 1–9 display strong phosphorescence in deaerated solution at room temperature (Fig. 2, 3 and S4 (ESI†); Table 1). All the complexes show vibronic-structured emissions, and their emission lifetimes are found to be in the microsecond regime. For example, 4Meb exhibits structured emission bands with vibronic spacings of $\sim$1380 cm$^{-1}$ and a long emission lifetime of 28.2 µs. Only small negative solvatochromic effects on the emission at 524 nm are found (±5 nm; Fig. S4d, ESIf). These findings, together with TD-DFT calculations, suggest that the photoluminescence of the complexes is derived from triplet metal-perturbed ligand-centred ($^3$LC) $\pi$–$\pi^*$ excited states. On the other hand, the structureless emission of 1d (the acetylacetonate (acac) analogous of 1a) at 516 nm should be ascribed to the $^3$MLCT/LLCT emission.$^2$ Interestingly, complex 3a ($\Phi = 0.2\%$; $\tau = 4.8$ µs) displays dual emission (Fig. 2) in contrast to the single emission of the acac analogue 3d,$^7$ suggesting the possibility of modulation of photophysical properties of Ir(III) complexes by the NHC ligands.

Changes in chemical structure of the cyclometalated ligands also result in a profound effect on the photophysical properties of the Ir(III) complexes. For example, 5bf displays a significant red-shift in emission maximum (576 nm) when compared to 4Meb and 4Heb (525 and 527 nm respectively). This can be rationalized by the extended $\pi$-conjugation of the C$^\equiv$N ligands, leading to the decrease in energy of metal-perturbed $^3$LC emission. Moreover, for 4Mea, 4Meb and 4Heb, their photoluminescence quantum yields in solution are affected by the substituents on the fluorenly moiety as well as the N-alkyl groups of the bis-NHC ligands, e.g. complex 4Heb exhibits a higher photoluminescence quantum yield than 4Mea and
Table 1  Photophysical data of complexes 1–9

| Complex | Medium$^a$ | Absorption ($\lambda_{\text{max}}$/nm) (10$^{-3}$ $\varepsilon$/M cm$^{-1}$) | Emission $\lambda_{\text{em}}$/nm ($\tau$/µs) | $\Phi^d$/% |
|---------|------------|-------------------------------------------------|--------------------------------|-------------|
| 1a      | CH$_2$Cl$_2$ | 255 (26.8), 267 (24.9), 311 (10.6), 342 (6.1), 380 (3.7), 416 (1.2) | 470 (2.1), 499, 534 | 89 |
| 1b      | CH$_2$Cl$_2$ | 254 (36.8), 266 (35.5), 311 (15.5), 342 (9.1), 381 (5.82), 416 (2.16) | 470 (2.1), 500, 534 | 89 |
| 1c      | H$_2$O      | 254 (36.8), 266 (35.5), 311 (15.5), 342 (9.1), 381 (5.82), 416 (2.16) | 469 (2.0), 498, 534 | 89 |
| 2a      | CH$_2$Cl$_2$ | 252 (13.5), 290 (21.7), 332 (12.6), 369 (6.07), 405 (4.92) | 530, 547, 570 (3.03) | 11.1 |
| 3a      | CH$_2$Cl$_2$ | 252 (15.5), 282 (15.1), 313 (20.5), 335 (23.1), 375 (21.9), 389 (20.9), 440 (20.9) | 472, 674 (4.8), 824 | 0.2 |
| 4$^{4a}$b | CH$_2$Cl$_2$$^b$ | 260 (31.4), 297 (30.0), 316 (33.5), 334 (32.3), 360 (22.4), 378 (16.8), 421 (10.5) | 524, 564 (28.6), 614 | 65 |
| 4$^{4b}$b | CH$_2$Cl$_2$ | 267 (52.2), 297 (46.8), 316 (50.9), 335 (49.0), 358 (36.3), 376 (27.3), 421 (16.2) | 525, 566 (28.7), 614 | 75 |
| 4$^{4c}$b | CH$_2$Cl$_2$ | 268 (42.1), 299 (42.3), 320 (50.4), 338 (49.2), 364 (32.6), 379 (25.6), 421 (14.6) | 527, 568 (32.7), 616 | 82 |
| 5b      | CH$_2$Cl$_2$ | 260 (54.1), 273 (42.5), 310 (32.9), 323 (34.8), 377 (67.2), 398 (97.0), 436 (71.6) | 576, 961 (6.4), 625, 683 | 22.6 |
| 6$^{6a}$b | CH$_2$Cl$_2$ | 271 (23.0), 321 (25.0), 377 (9.57), 407 (6.81), 436 (3.38) | 531, 560 (6.4) | 78 |
| 6$^{6c}$c | H$_2$O      | 268 (21.4), 321 (22.1), 377 (8.06), 407 (5.49), 436 (2.42) | 536, 560 (5.0) | 66 |
| 6$^{6b}$b | CH$_2$Cl$_2$ | 270 (23.0), 323 (32.1), 383 (10.7), 415 (7.88), 447 (4.08) | 530, 558 (3.2) | 68 |
| 7b      | CH$_2$Cl$_2$ | 296 (26.8), 333 (18.5), 388 (9.31), 437 (7.76), 467 (4.60) | 582, 623 (6.2) | 36 |
| 7c      | H$_2$O      | 293 (25.9), 330 (17.4), 386 (9.63), 432 (6.92), 464 (3.71) | 583, 622 (5.9) | 32 |
| 8b      | CH$_2$Cl$_2$ | 260 (80.1), 290 (48.5), 315 (35.4), 376 (34.5), 429 (7.52), 485 (3.47), 494 (2.88) | 618, 659 (5.2) | 3 |
| 9b      | CH$_2$Cl$_2$ | 313 (24.2), 334 (23.8), 384 (12.7), 428 (10.7), 457 (8.20), 485 (3.47) | 617, 667 (7.4) | 13 |
| 9c      | H$_2$O      | 315 (18.4), 380 (9.69), 420 (7.6), 454 (4.95), 475 (2.15) | 620 (4.5) | 668 |

$^a$ Measured in degassed CH$_2$Cl$_2$ and water (2.0 × 10$^{-5}$ M) at 298 K. $^b$ 4$^{4a}$b, 4$^{4b}$b and 4$^{4c}$b were measured at the concentration of 5 × 10$^{-6}$ M. $^c$ 5b was measured at the concentration of 1 × 10$^{-6}$ M. $^d$ Phosphorescence quantum yields were measured by using [Ru(bpy)$_3$]($PF_6$)$_2$ (Φ = 0.062 in MeCN) as standard.

4$^{4c}$b. This is probably attributed to the fact that the long hexyl and N-butyl chains disfavor intermolecular stacking interactions among the planar C$^N$N ligands, leading to reduced triplet–triplet annihilation and a higher photoluminescence quantum yield.

In addition, 5b shows a significantly longer emission lifetime (96.1 µs) than 4$^{4c}$b (28.7 µs). This might be due to the reduced metal character in the electronic excited state of 5b, and hence slower triplet radiative decay. Similarly, with a smaller parentage of metal character in the frontier molecular orbitals, 6$^{6b}$ (6.4 µs) shows a longer emission lifetime than 6$^{6b}$ (3.2 µs). These long-lived triplet excited states allow the Ir(m) NHC complexes to undergo a variety of photochemical reactions, notably for visible-light-driven photo-catalytic reductive C–Br bond cleavage and CO$_2$ reduction which will be illustrated later.

**Electrochemistry**

The electrochemical data of 1b–9b are summarized in Table 2 (all values versus Ag/AgNO$_3$, scan rate of 100 mV s$^{-1}$, 0.1 M $^n$Bu$_4$NPF$_6$ in MeCN as supporting electrolyte) and Table S3$^†$ (values vs. Cp$_2$Fe$/^+$/0, ESI$^†$). These complexes display one irreversible oxidative wave at $E_{pa}$ = 0.62–1.04 V and one irreversible reductive wave at $E_{pc}$ = −2.52 to −1.94 V (vs. Ag/AgNO$_3$). The cyclic voltammograms of 1a have been reported elsewhere$^{24}$ with the oxidation potential of +1.16 V (vs. Ag/AgNO$_3$) in CH$_2$Cl$_2$. By comparing with the electrochemical data of 1d and 2d$^2$, both of which have the ancillary acac ligand, the first oxidation wave of 1a should originate from Ir(m)-centred/C$^N$N ligand-centred oxidation. This assignment is further supported by the observation of the more positive oxidation potential of 6$^{6b}$ (+1.04 V) than 6$^{6b}$ (+0.98 V). The electron-withdrawing –CF$_3$ group of 6$^{6b}$...
will lower the energy levels of the \( \pi(C^\text{N}) \) orbitals, leading to the more positive oxidation potential of \( 6\text{b} \). On the other hand, \( 4\text{Me}\text{b} \) and \( 5\text{b} \) have more extended \( \pi \)-conjugated \( C^\text{N} \) ligands than \( 1\text{b} \). This will result in an increase in the energy level of \( \pi(C^\text{N}) \) orbitals, thus \( 4\text{Me}\text{b} \) and \( 5\text{b} \) can undergo oxidation more readily than \( 1\text{b} \) and hence less positive oxidation potentials are found.

Compared with \( 1\text{b} \), \([\text{dfppy}]_2\text{Ir}([\text{bis-NHC}\text{Bu}])\text{PF}_6\) (ref. 72) displays a more anodic oxidation potential of \( E_{pa} = 1.04 \text{V} \) and a similar reduction potential of \( E_{pc} = -2.37 \text{V} \) (vs. \( \text{Cp}_2\text{Fe}^{+/0} \)), which is attributed to stabilization of \( \text{HOMO} \)s as a result of the presence of highly electron-withdrawing F substitution on the \( C^\text{N} \) ligand (dfppy). For the irreversible reduction wave, a less negative reduction potential is found for \( 6\text{Hb} \) (\( E_{pc} = -2.14 \text{V} \) vs. \( \text{Ag/AgNO}_3 \)), Table 2, Fig. S4†), as compared to that of \( 1\text{b} \) (\( E_{pc} = -2.52 \text{V} \) vs. \( \text{Ag/AgNO}_3 \), Table 2, Fig. 4) which has a less extended \( \pi \)-conjugation of the \( C^\text{N} \) ligand. The reductive wave is further anodically shifted in the case of \( 6\text{b} \) (\( E_{pc} = -1.94 \text{V} \) vs. \( \text{Ag/AgNO}_3 \), Table 2, Fig. S5a, ESI†). Therefore, the reduction process should be localized on the \( C^\text{N} \) ligands.

**Excited state properties**

Nano-second transient absorption and emission spectroscopic (tr-abs and tr-em) measurements were undertaken for the highly emissive bis-NHC Ir[m] complexes with long-lived electronic excited states i.e. \( 4\text{Me}\text{b} \) (Fig. 5) and \( 4\text{Me}\text{c} \) (Fig. 6a) in MeCN and water respectively. These two complexes have the same lumiphore but different functionalized NHC groups, leading to a different solubility in solvents, such as one is soluble in organic solvents and the other in water. The tr-abs and tr-em spectra of the other bis-NHC Ir[m] complexes e.g. \( 5\text{b} \) in MeCN (Fig. S6a, ESI†), and \( 6\text{Hc} \) (Fig. S6b, ESI†) in water were also recorded and the results can be found in the ESI†.

The time-resolved absorption and emission spectra of \( 4\text{Me}\text{c} \) recorded at various time intervals after excitation at 355 nm are depicted in Fig. 5. The kinetic decay analysis of bleaching of ground-state of \( 4\text{Me}\text{b} \) (\( 5\text{a} = 15.4 \mu\text{s} \), Fig. 5a, left inset) matches well with the growth of the absorption of triplet excited state (\( 4\text{Me}\text{c} \) (Fig. 5a, right inset)) as well as the emission lifetime (\( 5\text{a} = 16.5 \mu\text{s} \), Fig. 5b, right inset) in MeCN at 298 K.

The transient absorption spectra of \( 4\text{Me}\text{c} \) in aqueous solution recorded at different energies of laser beams (355 nm) reveal different spectral changes. As depicted in Fig. 6a, in addition to the growth of the absorption of triplet excited state of \( 4\text{Me}\text{c} \) from 380 to 700 nm, the emergence of an absorption band from 650 to 730 nm was observed at laser pulse energy \( \geq 50 \text{mJ} \) (beam area: 0.5 cm²). This absorption band could be quenched upon addition of acetone (Fig. 6b and S7a and b, ESI†), but no quenching of transient absorption at 495 nm and emission at 524 nm were observed in the presence of acetone (Fig. 6c and S7, ESI†). The decay rate constant monitored at 720 nm of \( 4\text{Me}\text{c} \)

![Table 2](https://example.com/table2.png)

**Table 2** Electrochemical data\(^a\) of bis-NHC Ir(III) complexes

| Complex | \( E_{pa} \) \( / \text{V} \) | \( E_{pc} \) \( / \text{V} \) |
|---------|-----------------|-----------------|
| 1b      | 0.87            | -2.52           |
| 2a      | 0.64            | -2.24           |
| 4Mea    | 0.74            | -2.41           |
| 4Meb    | 0.79            | -2.40           |
| 4Heb    | 0.74            | -2.44           |
| 5b      | 0.72            | -2.25           |
| 6Hb     | 0.98            | -2.14           |
| 6b      | 1.04            | -1.94           |
| 7b      | 0.69            | -2.09           |
| 8b      | 0.62            | -1.94           |
| 9b      | 0.62            | -2.16           |

\(^a\) Supporting electrolyte: 0.1 M \( \text{Bu}_4\text{NPF}_6 \) in MeCN and values are recorded vs. \( \text{Ag/AgNO}_3 \) (0.1 M) in MeCN; \( \text{Cp}_2\text{Fe}^{+/0} \) occurs at the range of 0.05–0.08 \( \text{V} \) vs. \( \text{Ag/AgNO}_3 \). \(^b\) Values refer to oxidation peak potential (\( E_{pa} \)) at 25 °C for irreversible couples at a scan rate of 100 mV s\(^{-1}\). \(^c\) Values refer to reduction peak potential (\( E_{pc} \)) for the irreversible reduction waves.

**Fig. 4** Cyclic voltammograms of \( 1\text{b} \), \( 2\text{a} \), \( 4\text{Me}\text{b} \), \( 5\text{b} \) and \( 6\text{b} \) in MeCN with \( \text{Bu}_4\text{NPF}_6 \) (0.1 M) as supporting electrolyte. Conditions: glass-carbon, working electrode, scan rate: 100 mV s\(^{-1}\).

**Fig. 5** Time-resolved spectra of \( 4\text{Me}\text{b} \) (a) tr-abs (insets: decay of tr-abs at \( \lambda = 350 \text{nm} \) and 480 nm); (b) tr-em (decay of tr-em at \( \lambda = 524 \text{nm} \)) spectra recorded at specified times after laser pulse excitation (355 nm) in degassed MeCN at 298 K.
example is \( \text{Ir}^{IV} \text{(Ir}^{III}/\text{Ir}^{IV} \rangle = -1.51 \text{ V vs. SCE} \) (here \( \text{Ir}^{IV} \) is a simple notation to denote the oxidized \( \text{Ir}^{III} \) species, the site of the oxidation can be the metal and/or C-N ligand), which is a stronger reductant than \( [\text{Ru(bpy)}_3]^{2+} \). As a result, it is anticipated that the bis-NHC Ir(m) complexes described herein, upon photoexcitation in the visible-light region, can catalyse a number of reactions which are not feasible by the widely used \( [\text{Ru(bpy)}_3]^{2+} \).

**Photo-catalysis**

With good photo-stability, long excited state lifetime, favourable absorption in the spectral region of blue LED and tunable photo-redox properties, these bis-NHC Ir(m) complexes have been investigated for photo-redox reactions (Chart 2), examples of which are described in the following section.

We considered a recent work by Lee and co-workers on visible-light-induced reductive cyclization of aromatic iodides and bromides to form indoline using \( [\text{[ppy]}_2\text{Ir(dtbppy)}] \) as photo-catalyst (PC). Yet, aryl bromides are found to be less reactive than iodides. Barriault and co-workers addressed this issue by using dinuclear gold(I) complexes \( [\text{Au}_2(\text{dppm})_2\text{OTf}_2] \) as photo-catalysts. However, this gold complex only shows absorption in the high-energy UV region \( (\lambda < 300 \text{ nm}) \), which may result in destructive effects on the products and/or lead to undesired side reactions.

As the present bis-NHC Ir(III) complexes show strong absorption at \( \lambda > 400 \text{ nm} \), they were used to photo-catalyse the reductive cyclization of aryl bromides using blue LEDs. Among the bis-NHC Ir(m) complexes examined, \( 1\text{b}, 4\text{Meb} \) and \( \text{d}^6\text{b} \) displayed good photo-catalytic activity for reductive cyclization of aryl bromide (A1) in terms of both substrate conversion and product yield (Table 3).

In the case of another aryl bromide substrate A2, \( 4\text{Meb} \) showed both good substrate conversion and product yield (Table 4, entries 4–6). It was chosen as photosensitizer for optimization of the reaction conditions. A number of control experiments were performed, and no reaction was observed in the absence of amine or light (entries 13 and 14). Lowering the

(Fig. 6d, absence of acetone) was much faster than that measured at 495 nm in Fig. 6c/S7/c and decay of emission at 524 nm (Fig. S7d, ESIF). In view of these different kinetic behaviours, the transient absorption from 650 to 730 nm depicted in Fig. 6b might originate from hydrated electrons \( e^{-}_{\text{aq}} \) which were formed by the photo-induced ionization of \( 6\text{h}^\text{c} \) in aqueous solution upon excitation with high energy laser beams. This is in line with a reported photoionization of \( [\text{Pt}_2(\text{POP})_4]^+ \), as well as the findings in the studies of solvated electron with acetone. Similarly, complex \( 6\text{h}^\text{c} \) in aqueous solution was also observed to undergo photo-ionization as revealed by the increase in transient absorption in the region of 600 to 700 nm (Fig. S8b, ESIF). For \( 4\text{Meb} \), its transient absorption spectra monitored at high laser pulse energy in MeCN exhibit similar profiles as for lower energy (Fig. S8d, ESIF). However, newly generated long-lived species have been observed at high laser pulse energies after 100 \( \mu \text{s} \) (Fig. S8b–e, ESIF), revealing that photo-ionization of \( 4\text{Meb} \) with the generation of \( [4\text{Meb}]^+ \) likely occurs. The accompanying solvated electrons are not observed in this case due to ready quenching by MeCN.

Based on the electrochemical data and the determination of \( E_{\text{irr},0} \) from the spectroscopic measurements, the excited state redox properties of the bis-NHC Ir(m) complexes can be estimated (Table S3, ESIF). The triplet excited states of the complexes are found to be powerful oxidants and reductants, and some of them are even more reactive towards photoredox reactions than \( [\text{Ru(bpy)}_3]^{2+} \) and \( \text{fac-}[\text{ppy}]_3 \). A representative

![Figure 6](https://example.com/figure6.png)

**Fig. 6** Transient absorption spectra recorded (a) under specified laser energy; (b) in presence of specified concentrations of acetone; of a degassed aqueous solution of \( 4\text{Mec} \) (about \( 1 \times 10^{-5} \text{ M} \)). Kinetic studies of (c) \( \lambda_{	ext{tr-abs}} \) (495 nm) and (d) \( \lambda_{	ext{tr-abs}} \) (720 nm) in the absence/presence of specified concentrations of acetone.
loading of PC (Table 3, entry 4), the absence of HCOOH or exposure to air (entries 11 and 12) were observed to decrease the substrate conversion of this reaction. Interestingly, using 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU, entry 10), PC conversion was also led to excellent substrate conversion and good product yield comparable to what was obtained with tetramethylthelylenediamine (TMEDA, entry 9), for 4 h. Determined by ¹H NMR spectroscopy by adding internal standard of 5,5-dimethyl-2,2'-bipyrindine.

The radical cyclization of the alkyl bromides, B1 and B2, catalysed by 1b, 4⁴ᵇᵇ or 6⁴ᵇ proceeded smoothly, with reasonable to excellent substrate conversion and product yields. The yield of cyclization of B1 was improved to up to 90% (entry 5, Table 3) when formic acid was added and 6⁴ᵇ was used as a photo-catalyst (PC). When 2,2,6,6-tetramethyl-1-piperidinolxy (TEMPO) was added, the reaction was totally inhibited, indicating the involvement of a radical intermediate in the reaction (entry 12).

The photo-catalytic reaction could be initiated from the oxidative quenching of 4⁴ᵇᵇ with alkyl halides. This is because the excited state reduction potential of 4⁴ᵇᵇ (E(ºm⁴ᵇᵇ) vs. SCE) was determined by a reductive quenching mechanism to generate Ir(III) species (E(ºm⁴ᵇᵇ) vs. SCE = −1.51 V (vs. SCE)), can allow a direct one electron reduction of alkyl halides by 4⁴ᵇᵇ*, leading to carbon-halogen bond cleavage to give alkyl radical in the case of sp³ carbon or radical anion intermediate for sp² carbon. Subsequent reactions of the alkyl radical or radical anion intermediate with C(sp³)=H bond lead to C-C bond formation. An aminium radical cation generated from the oxidation of amine by [4⁴ᵇᵇ]PF₆ can serve as an electron donor to complete the reductive process. In the case of [[ppy]Ir(dtbppy)] PF₆, its triplet excited state ([(ppy)Ir(dtbppy)]°) reacts with DIPEA via a reductive quenching mechanism to generate Ir(III) species (E(ºm⁴ᵇᵇ) vs. SCE = −1.51 V (vs. SCE)), which initiates the subsequent reducing catalytic reaction. Amines Conversion/\% Yield/\%

| Entry | PC (substrate) | Amines | Conversion/\% | Yield/\% |
|-------|----------------|--------|---------------|---------|
| 1     | 1b (A1)        | DIPEA  | 97            | 67      |
| 2     | 4⁴ᵇᵇ (A1)     | DIPEA  | 97            | 64      |
| 3     | 6⁴ᵇᵇ (A1)     | DIPEA  | 96            | 64      |
| 4     | 4⁴ᵇᵇ (A2)     | DIPEA  | 98            | 54      |
| 5     | 1b (A2)        | DIPEA  | 99            | <15     |
| 6     | 6⁴ᵇᵇ (A2)     | DIPEA  | 84            | 51      |
| 7     | 4⁴ᵇᵇ (A3)     | TEA    | 97            | 79      |
| 8     | 4⁴ᵇᵇ (A1)     | TEA    | 91            | 64      |
| 9     | 4⁴ᵇᵇ (A1)     | TMEDA  | 52            | 32      |
| 10    | 4⁴ᵇᵇ (A1)     | DBU    | 96            | 72      |
| 11    | 4⁴ᵇᵇ (A1)     | DIPEA  | 85            | 54      |
| 12    | 4⁴ᵇᵇ (A1)     | DIPEA  | 62            | 28      |
| 13    | 4⁴ᵇᵇ (A1)     |       | 0             | 0       |
| 14    | 4⁴ᵇᵇ (A1)     |       | 0             | 0       |
| 15    | Ru(bpy)Cl₂ (A1)| DIPEA  | 0             | 0       |
| 16    | fac-Ir(bpy)₂ (A1)| DIPEA | 27           | 16      |

Visible-light-driven CO₂ reduction

There is a surge of interest in developing photo-catalytic CO₂ reduction using earth abundant metal complexes as catalysts...
Table 5 Visible-light-induced C–C bond formation of alkyl bromide

| Entry | PC                  | Amines | Conversion/ (%) | Yield/ (%) |
|-------|---------------------|--------|-----------------|------------|
| 1     | Ru[bpyp]Cl₂        | DIPEA  | 19              | 5          |
| 2     | fac-Ir[ppy]₃      | DIPEA  | 90              | 72         |
| 3     | 1b                  | DIPEA  | 99              | 85         |
| 4     | 4Meb                | DIPEA  | 99              | 76         |
| 5     | 6Hb                 | DIPEA  | 99              | 76         |
| 6     | 4Meb                | TEA    | 99              | 72         |
| 7     | 4Meb                | TMEDA  | 84              | 55         |
| 8     | 4Meb                | DBU    | 99              | 64         |
| 9     | 4Meb                | —      | 0               | 0          |
| 10    | 4Meb                | DIPEA  | 27              | 15         |
| 11    | 4Meb                | DIPEA  | 0               | 0          |
| 12    | 4Meb                | DIPEA  | 0               | 0          |
| 13    | 4Meb                | DIPEA  | 99              | 59         |

*a* Entry 1–12: R = H (substrate B1); procedure: substrate 50 μmol, PC (2 mol%), amine (5 equiv.) and HCOOH (2.5 equiv.) in 4 mL MeCN solution was degassed by nitrogen, and irradiated by blue light (12 W, λmax = 462 nm) at 25 °C. *b* Determined by ¹H NMR spectroscopy by adding an internal standard of 5,5′-dimethyl-2,2′-bipyridine. *c* Absence of HCOOH (20 equiv.). *d* Absence of light. *e* Presence of TEMPO (radical trapping reagent, 2 equiv.). *f* Entry 13: R = Me (substrate B2), and luminescent cyclometalated Ir(m) complexes particularly fac-Ir[ppy]₃ as PC. Nevertheless, several recent reports drew attention to the instability of fac-Ir[ppy]₃, which is a challenge for efficient light-driven CO₂ reduction in the long run.⁵⁵,⁵⁷,⁶₃ In view of the good photo-stability of 4McB, we investigated the visible-light-driven CO₂ reduction by utilizing 4McB in combination with the recently reported [Co(TPA)Cl]Cl complex.⁶⁷

A CO₂-saturated MeCN/triethylamine solution (4 : 1, v/v; 4 mL) containing catalytic amounts of 4McB and [Co(TPA)Cl]Cl was irradiated by blue LEDs (12 W) for a specified time period, and the evolved gases were separated and identified by GC-TCD equipped with a molecular sieve column. The volume of H₂ and CO gases were calculated by using CH₄ as the internal standard.

As shown in Fig. 7a and b, the amount of CO and H₂ generated from the reaction mixture is found to show strong dependence on the concentrations of 4McB and Co(n) catalysts. Particularly, the highest TON value of over 5000 can be accomplished at 0.5 μM of Co(n), while no generation of gases is found at low concentration of Co(n) (50 nM; Fig. S11, ESI†). Similarly, only negligible amount of product gases can be detected after irradiation for 24 h when [Ir] (0.005 mM) is lower than [Co] (0.05 mM). With the representative system containing 4McB (0.5 mM) and Co(n) (0.005 mM), (Fig. 7c) the visible-light-driven CO₂ reduction in three parallel reaction runs gives TON (CO) > 2400 (conversion of about 18 mL of CO₂ into 1 mL of CO) with excellent selectivity in generating CO over H₂ in the gaseous phase (>95%) after reaction for 72 h. This result is better than for the system utilizing fac-Ir[ppy]₃ as PC, which reveals only TON (CO) > 900 and selectivity (CO) of 85% under similar reaction conditions.⁶⁷

In order to confirm the roles of catalysts in photo-driven CO₂ reduction, several control experiments were performed (Table S5, ESI†). Firstly, in the absence of Co(n) complex, no CO gas was observed in a CO₂-saturated MeCN/TEA (4 : 1, v/v; 4 mL) solution after irradiation for 24 h. On the other hand, in the absence of light, sacrificial amine or PC, the reaction mixture only gives negligible amounts of CO. To ascertain the catalytic role of the Co(n) complex in the reaction, mercury was added to the reaction mixture in order to exclude the possibility of CO generation from heterogeneous Co nanoparticles. To specify, 29.2 μmol of CO (0.714 mL, TON 146) could be generated from the solution with [Ir] (0.5 mM), [Co] (0.05 mM) and elementary Hg (1 mL) after irradiation for 18 h, and this result is comparable with that using the solution mixture without Hg (28.4 μmol of CO formed, TON 142).

**Cellular imaging and cell viability assay**

We have a long-standing interest in luminescent transition metal complexes, particularly those supported by NHC ligands that display anti-cancer activities.⁹,¹⁰,¹⁶,⁵⁷,⁶⁷–⁶⁹ The luminescence would allow direct monitoring of cellular uptake and tracking of cellular location in cancer cells by fluorescence microscopy, and such properties of transition metal NHC complexes have been demonstrated to be useful in the elucidation of the anti-cancer mechanism of action. In this work, in view of their favourable
ligand are more cytotoxic than those with glucose units. Complexes. Those complexes with butyl groups on the bis-NHC depending on the lipophilicity and structures of the Ir(III) complexes. CO2 in a CO2-saturated MeCN/TEA (4/1, v/v, 4 mL in total) solution as an outstanding photophysical properties, the cellular imaging of these complexes in HeLa cells was performed. A assay. As shown in Table S6 (ESI†), the Ir(III) complexes display potent cytotoxicity, with IC50 values ranging from 0.5 to 56.1 μM depending on the lipophilicity and structures of the Ir(III) complexes. Those complexes with butyl groups on the bis-NHC ligand are more cytotoxic than those with glucose units.

Since complexes 1b, 4Meb, 6a, 6Hb and 6Hc demonstrate outstanding photophysical properties i.e. high quantum yield with long-lived electronic excited states, cellular imaging of these complexes in HeLa cells were performed. After treatment of human cervical cancer cells (HeLa) with the Ir(III) complexes for 15 min, strong green/yellow luminescence was observed in the cytoplasm (but not nucleus) of cancer cells, as revealed by fluorescence microscopy (Fig. 8). Co-localization analyses indicate that the emission of these complexes are mainly localized in the endoplasmic reticulum (ER), which is stained by red-emissive ER-specific ER-Tracker®; a high Pearson’s correlation coefficient (R) between the emissions of complexes and ER-Tracker® is found (for example, 6Hb shows a high R value of 0.80). Consistently, these complexes do not accumulate in other organelles such as lysosome or mitochondria, as shown by the poor co-localization of the emissions of the complexes with the emission of Lyso-Tracker® and Mito-Tracker® respectively (Fig. S13, ESI†). Noticeably, 1b+ (both the counter anions of triflate and chloride) was found in our laboratory to be specifically localized in the ER of cancer cells (Fig. S14†), but not, as reported elsewhere, in mitochondria. With the specific accumulation of the complexes in ER, the cytotoxic properties of the complexes may originate from the induction of ER stress and immunogenic cell death.†

Discussion

This work reveals the potential impact and usefulness of bis-NHC ligands for the development of robust metal photo-catalysts. Compared with the well-known complexes fac-Ir(ppy)3, this series of bis-NHC Ir(III) complexes exhibits: (i) outstanding photo-stability under visible light irradiation; (ii) good photocatalytic performance for several photo-electrochemically reactions; and (iii) long-lived emissive excited states.

Stability of NHC metal complexes

The most important feature of metal–ligand bonding between transition metals and NHCs can be rationalized as the σ-coordination (sp2-hybridized lone pair electron) from NHCs. The contribution of both π-back-bonding into carbene p-orbital and π-donation from the carbene p-orbital might be considered as less significant for NHC metal complexes. These behaviours are similar to the coordination characteristics of phosphines, but NHCs are in general better electron-donors than phosphines. Thus, the stronger metal–ligand interaction renders NHC–metal coordination less labile than metal–phosphine bonding and the NHC complexes are more thermally stable. The distinct electronic properties and coordination chemistry of NHCs can also lead to improved catalytic activity of the metal complexes, owing to the increased catalyst stability and consequently lower rates of catalyst decomposition. In the previous section, the studies on the photo-stability of 4Meb and fac-Ir(ppy)3 have revealed the outstanding stabilization contribution from bis-NHC carbene ligands. Thus this might be one reason that explains the better performance for photo-catalytic reactions using the present bis-NHC Ir(III) complexes as PC.

Role of bis-NHC ligands in tuning emission energy and lifetimes

A comparison between the photophysical data of bis-NHC Ir(III) complexes with two notable Ir(III) complexes (fac-Ir(ppy)3 and...
(ppy)_2Ir(acac) (1d)) will be helpful in understanding the role of NHC on the strong absorbivities, high emission efficiencies and long lifetimes of the present bis-NHC Ir(III) complexes. On the basis of DFT calculations on 1a as the representative example (Fig. S4, ESI†), the lowest-energy absorption bands of 1a, 1d and fac-Ir(ppy)_3 are ascribed to the HOMO → LUMO transitions, the energy of which is sensitive to the charge of ancillary ligands. The transition energy of fac-Ir(ppy)_3 and 1d with the negatively charge ancillary ligands (C^N and acac) are quite close with wavelengths of 416 and 414 nm respectively. However, for 1a with the neutral ancillary ligand (bis-NHC), the HOMO → LUMO transition is markedly hypochromically shifted to 369 nm, which is consistent with the experimental observations. Fig. S10 (ESI†) shows the surfaces and the energies of HOMO and LUMO of 1a, fac-Ir(ppy)_3, and 1d. On the whole, the HOMO contains comparable components of iridium and C^N ligand, and the LUMO is mainly localized on the C^N ligand. Thus, the transition can be assigned as an admixture of metal-to-ligand charge transfer (MLCT) and ligand centered (LC) π→π* transition, which is in accordance with the assignments reported in the literature.

The different amounts of metal character in the frontier molecular orbitals of 1a, 1d and Ir(ppy)_3, as deduced by TD-DFT calculations, can account for the photophysical properties and long emission lifetimes of the Ir(III) NHC complexes. As shown in Fig. S10 (ESI†), fac-Ir(ppy)_3 and 1d have similar energy levels in HOMO and LUMO (around −4.5 eV and −2.0 eV). By simple substitution of the negatively charged ancillary ligands (C^N and acac) in Ir(ppy)_3 and 1d by neutral NHC ligand, 1a is found to show lower energy levels of HOMO and LUMO (about −5.2 eV and −2.4 eV). This can be explained by the less electron-donating effect of the neutral NHC ligand compared to the negatively charged auxiliary ligands, as well as a certain contribution of the stabilization of dπ(Ir) by the π-acceptor orbitals of NHC ligand of 1a. As a result, the energy level of the HOMO of 1a is 700 mV lower than that of 1d, which coincides with the experimental observation that the oxidation potential of 1a is 410 mV more positive than that of 1d. On the other hand, the transition energy of HOMO → LUMO in 1a is estimated to be 0.3 eV larger than in Ir(ppy)_3 and 1d, which may account for the blue shift of low-energy absorption of 1a (369 nm vs. 410 nm), as compared to that of 1d and Ir(ppy)_3, in UV-Vis absorption spectra. Notably, TD-DFT calculation reveals that frontier molecular orbitals of 1a show a smaller metal character than those of 1d (Ir character in the HOMO of 1a and 1d are found to be ~30 and ~40% respectively), and hence the triplet excited states of 1a are likely to have smaller metal character. This would probably slow down the spin–orbit coupling, resulting in slower radiative and non-radiative T_1 → S_0 decay for 1a. As a result, 1a shows a longer emission lifetime than 1d.

**Triplet excited states for photo-catalytic reactivity**

In addition to the photo-stability and strong absorbivity in the visible light region, the long-lived triplet excited states, especially their propensity to lose an electron, are to a large extent believed to be crucial for the photo-catalytic properties of the complexes. The long triplet excited state lifetime of 4Me^b is beneficial for bi-molecular photochemical reactions allowing the electron-transfer pathways to have sufficient time to take place prior to the decay of the excited state to the ground state.

Considering the fact that the excited state of [Ru(bpy)_3]^{2+} ([Ru(bpy)_3]^{2+}) = 0.77 V vs. SCE) has a more powerful oxidative potential than for 4Me^b ([Ir(ppy)_3]^3+) = 0.51 V vs. SCE, Table S3) and no radical cyclization products are obtained when using [Ru(bpy)_3]^{2+}, the photo-catalytic route via reductive quenching...
cycle is not feasible. The photo-catalytic reaction would possibly be initialized via an oxidative quenching cycle of excited states of the photo-catalyst. By carefully examining the transient-absorption spectra of triplet state of \(4\text{Me}b\) in MeCN, newly generated long-lived species were observed by using higher energy laser beams (Fig. S8b-g). These long-lived species were found to be increased in the presence of substrate A1, and could be quenched by DIPEA. This long-lived species might be Ir(IV), which is generated by single electron transfer from \([4\text{Me}b]^+\) to A1. The calculated excited-state reduction potential of \(4\text{Me}b\) \((E(\text{Ir}^\text{IV/III}) = -1.51 \text{V vs. SCE}, \text{Table S3}^\dagger)\) reveals that \(4\text{Me}b\) is not a stronger photo-reductant than fac-Ir(ppy)_3, \((E(\text{Ir}^\text{IV/III}) = -1.73 \text{V vs. SCE})\). However, the performance of \(4\text{Me}b\) in visible-light-driven radical cyclization and CO2 reduction prove to be better than fac-Ir(ppy)_3. For example, the conversion of substrates A1/B1 to indoline/pyrrolidine by \(4\text{Me}b\) (97%/99%) are higher than those by fac-Ir(ppy)_3 (27%/90%). Therefore, there should be other reasons for the good performance of \(4\text{Me}b\) in photocatalysis. Plausible reasons could be: (i) the generated Ir(IV) \([4\text{Me}b]^\dagger\] \((E(\text{Ir}^\text{IV/III}) = 0.96 \text{V vs. SCE}, \text{Table S3})\) are more easily reduced by amines than \([\text{fac-Ir(ppy)}_3]^\dagger\] \((E(\text{Ir}^\text{IV/III}) = 0.77 \text{V vs. SCE})\); and (ii) the higher photo-stability of \(4\text{Me}c\) compared with that of fac-Ir(ppy)_3.

Consequently, the excellent photo-stability, strong absorptivity, long lifetimes and the photo-ionization behaviours of our bis-NHC Ir(m) complexes possibly enable the photolysis of long-lived excited states of Ir(m)* to Ir(m) more easily and thus promote the radical cyclization in a catalytic cycle.

**Diversities of biological activities of Ir(m) complexes with different N-substituents on NHC ligands**

Although N-substituents of NHC ligands are found to show negligible effects on the luminescent properties of the bis-NHC Ir(m) complexes, they have a determinant role on the physical properties and hence the anti-cancer activities of the complexes. For example, complexes with N-butyl substituents on the bis-NHC ligands generally display lower \(\lambda_{50}\) values than those with N-methyl substituents (Table S6 in ESI†). This is probably attributed to the increase in lipophilicity of the complexes, resulting in better permeability through cellular membrane and higher accumulation of the complexes in cancer cells. The fast cellular uptake of the complexes is supported by the strong luminescence observed in HeLa cells after incubation of the cells with the complexes for 15 min (Fig. 8 and S12 in ESI†). On the other hand, the glucose-functionalized NHC ligands give the complexes good aqueous solubility and a more hydrophilic nature, leading to likely slower cellular uptake as well as reduced cytotoxicity toward cancer cells. It is noteworthy that the complexes accumulate in cellular ER as revealed by fluorescence microscopy images of the costaining experiments. This can consequently induce ER stress and immunogenic cell death, which probably accounts for the high cytotoxicity of the complexes. Further manipulations of the functionalities on NHC ligands of the Ir(m) complexes may realize the development of a new class of diagnostic and/or therapeutic agents.

**Conclusions**

A new series of cyclometalated Ir(m) complexes bearing bis-NHC ligands has been demonstrated as strongly luminescent materials and promising photo-catalysts for visible-light-driven radical cyclization and CO2 reduction and as biological theranostic agents. Owing to the high stability of theIr–C\(^{3\text{NHC}}\) bond, these bis-NHC Ir(m) complexes show excellent photo-stability compared with widely-used PC, such as fac-Ir(ppy)_3 and \([\text{Ru(bpy)}_3]^\dagger\). With long-lived triplet excited states and rich photoredox properties, \(4\text{Me}c\) can undergo photo-ionization in aqueous media as supported by transient absorption experiments, while \(4\text{Me}b\) is found to be a more effective catalyst for radical cyclization and CO2 reduction than fac-Ir(ppy)_3 under visible-light irradiation. Interestingly, water-soluble \(6\text{H}c\), which contains a glucose moiety on the bis-NHC ligands, has been demonstrated as the first PC for the synthesis of pyrrolidine in aqueous media. In addition, through modulations of the chemical structures of the cyclometalated ligands and/or N-substituents on the bis-NHC ligands, the Ir(m) complexes have been found to show different luminescent properties as well as anti-cancer activities, indicating the potential of the complexes as theranostic agents.

**Author contribution**

Chen Yang, Faisal Mehmood, Tsz-Lung Lam, Yuan Wu, Chi-Shun Yeung, Xiangguo Guan, Kai Li, Clive Yik-Sham Chung, Cong-Ying Zhou and Tao Tao Zou carried out all the experiments and performed data analysis. Chen Yang, Chi-Ming Che and Sharon Lai-Fung Chan designed the experiments, analysed the data and wrote the manuscript. All authors reviewed the manuscript.

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