New Insights into H$_2$S Adsorption on Graphene and Graphene-Like Structures: A Comparative DFT Study

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Received: 18 September 2020; Accepted: 5 November 2020; Published: 9 November 2020

Abstract: The efficient removal of pollutants from different environments has been one of the great challenges for scientists in recent years. However, the understanding of the mechanisms underlying this phenomenon is still the subject of passionate debates, mainly due to the lack of experimental tools capable of detecting events at the atomic scale. Herein, a comparative theoretical study was carried out to capture the adsorption of H$_2$S on metal oxide surfaces such as zinc oxide (ZnO) and beryllium oxide (BeO), as well as graphene and Ni-decorated graphene. A simulation based on density-functional theory (DFT) was carried out by adopting General Gradient Approximation (GGA) under the Perdew–Burke–Ernzerhof (PBE) function. The calculations quantified H$_2$S adsorption on the considered metal oxide sheets as well as on the non-decorated graphene having a physical nature. In contrast, H$_2$S adsorbed on Ni-decorated graphene sheet gave an adsorption energy of $-1.64$ eV due to the interaction of S and Ni atoms through the formation of a covalent bond, proof of chemisorption. It seems that the graphene sheet decorated with Ni atoms is a more suitable adsorbent for H$_2$S molecules than BeO, ZnO, or non-decorated graphene, providing a theoretical basis for future studies.

Keywords: H$_2$S adsorption; metal oxide; graphene-like structures; Ni-decorated graphene; DFT

1. Introduction

H$_2$S is one of the air pollutants released during various processes such as petroleum refinery, natural gas and biogas processing, coking plants, wastewater treatment units, etc. [1,2]. This highly
toxic compound has raised global concerns in recent decades due to its harmful effects on the environment as well as on human health [3,4]. If inhaled, even at low concentrations, H$_2$S can seriously affect the nervous system, causes a feeling of weakness, coughing, and runny nose. Moreover, at higher concentrations, the consequences would be even worse, which could lead to visceral injury, coma, acidosis, and death [5,6]. In industrial processes, H$_2$S can cause corrosion of process equipment, contaminate pipelines, and consequently, affect product quality [7]. Moreover, excess H$_2$S in the environment and atmosphere can lead to acid rains and cause serious damage to crops and infrastructure [8,9].

Considering all the above-mentioned problems, the detection and adsorption of H$_2$S have recently been the subject of many theoretical and experimental research studies [10–12]. In this respect, different materials such as metal surfaces [13,14], ionic liquids [11], zeolites [15], and activated carbons [16,17] have been examined as potential H$_2$S adsorbents. In addition, the use of low-dimensional nanostructures including carbon nanotubes [18,19], fullerenes [20], pristine graphene [21], metal oxides [22,23], decorated graphene [24–26], and doped graphene [27,28] for the adsorption of gas molecules has opened up a new window on this topic due to the outstanding properties of these structures.

Since the introduction of graphene to the world in the 1990s, this 2D structure has been used in various applications, particularly in gas sensing applications. Graphene, as a $sp^2$-hybridized monolayer of carbon atoms forming a hexagonal honeycomb lattice, provides great sensing properties due to its high specific surface area, good electron mobility, and excellent electronic conductivity [29,30]. However, several studies have reported a weak physisorption interaction between pristine graphene and common gas molecules [31,32]. Therefore, doping or decorating graphene with different elements has been recognized as solutions to the poor electronic properties of pristine graphene, and to create stronger interactions between graphene and gas molecules [33]. Many previous research works have been devoted to this topic, mainly based on the density-functional theory (DFT) approach. Namely, Mn-doped graphene was modeled using DFT by Jia et al. [7], who showed a significant improvement in adsorption energy and charge density, as well as the higher adsorption capacity of Mn-doped graphene toward H$_2$S compared to pristine graphene. In another DFT-based research, Khodadadi [28] reported that decorating graphene with different transition metals could significantly improve its H$_2$S sensing capability. The same kind of result was also observed in the case of graphene nanosheets doped with Fe [27,28] and Ag [34], as these metals considerably improve their adsorption properties with respect to pristine graphene by Lewis acid–base interactions.

The resulting charge transfer between the adsorbent and the adsorbate therefore changes the resistivity of the material, so that a chemiresistive graphene-based sensor system can be imagined. However, the sensitivity of the device would be too low if pristine graphene were used, due to the too weak and long-range interaction of this kind of surface with most gases at room temperature. Doping graphene with metals thus appears to be a very efficient way to adjust its adsorption properties, and hence, its ability to detect and quantify gaseous molecules by changing its conductive properties. In particular, the advantage of such a method is that the adsorption of metallic nanoparticles on graphene does not deteriorate its carrier mobility [35].

Thus, whereas too weak adsorbate–adsorbent interactions are detrimental to the sensitivity of the sensor, especially at room temperature, it can be expected that the recoverability of the sensor, i.e., its ability to release the adsorbed gas within a reasonable time and at a reasonable temperature, would be poor in the case of too strong interactions. An optimum should therefore exist between too low and too high adsorption energy, and the well-known affinity of Ni for S, although not as high as that of Ag or Pt, should be a good compromise for such an application. Moreover, Ni is one of the transition metals that has been widely used to decorate various materials and improve their adsorption properties. For instance, Gaboardi et al. [36] reported that the hydrogen adsorption capacity of graphene increased by 51% after being decorated by Ni. Therefore, Ni appears to be effective as an element for decorating graphene to improve its adsorption properties towards different substances, especially H$_2$S.
The remarkable properties of graphene have motivated researchers to focus on the adsorption of other types of monolayer structures such as metal oxide surfaces, including but not limited to zinc oxide (ZnO) [37,38], beryllium oxide (BeO), manganese oxide (MnO) [38], vanadium pentoxide (V₂O₅) [23], or sulfide-based and nitride-based surfaces such as molybdenum disulfide (MoS₂) [39], aluminum nitride (AlN), boron nitride (BN), and gallium nitride (GaN) [40]. Namely, the acetone adsorption properties of BeO, MgO, and ZnO surfaces were investigated in a DFT study by Mo et al. [38]. They reported satisfactory adsorption energies of all these surfaces interacting with acetone [38]. In the research by Ganji et al. [40], DFT calculations were performed to evaluate the adsorption of H₂S on graphene-like structures of aluminum nitride (AlN) and gallium nitride (GaN), and reported adsorption energies of −0.30 and −0.33 eV for AlN and GaN sheets, respectively. Based on the results, the GaN sheet with a higher adsorption energy was also identified as a relevant H₂S sensing material. Mashhadzadeh et al. [37] used graphene-like ZnO sheets and ZnO nanotubes to adsorb heavy metal atoms, including Ni²⁺, Cu²⁺, Ag⁺, and Cd²⁺, using DFT calculations. Their results demonstrated that the strongest adsorption occurs between Ni²⁺ and ZnO nanotubes or sheets, compared to the other heavy metals studied. The doping of ZnO monolayers with Pd, Au, Fe, and Co was simulated by Zhang et al. [41] and, according to DFT calculations, a higher adsorption capacity of H₂S was reported for doped nanosheets compared to non-doped ones.

Given the importance of evaluating adsorption of hazardous gases on graphene-like surfaces, to the best of our knowledge, no research work has been conducted to date to provide comparison between the H₂S adsorption properties of ZnO and BeO surfaces with pristine and decorated graphene sheets. The main objective of the present study was to investigate the adsorption properties of ZnO, BeO, and pristine graphene sheets interacting with H₂S and to provide comprehensive comparisons between these results and those obtained for a Ni-decorated graphene surface. To achieve this goal, all the nanostructures mentioned were designed and geometrically optimized using DFT. Adsorption energies as well as the density of states (DOS) and partial density of states (PDOS) calculations were also generated to capture the behavior and intensity of H₂S adsorption properties of these typical surfaces. Finally, the results obtained were discussed and compared to be considered in the future development of the corresponding H₂S gas sensors. It must be clearly understood that this work is intended to be confronted with future experimental aspects, particularly in terms of feasibility and performance (efficiency of the adsorption/desorption process and problems of saturation and contamination of the layer) once such a sensor will have been designed. These considerations, therefore, go well beyond the primary objective of this article.

2. Computational Method

A classical simulation algorithm was used in this study. In the first step, the adsorption structures of H₂S molecules on BeO, ZnO, pristine graphene, and Ni-decorated graphene were modeled and geometrically optimized by DFT calculations using the “Spanish Initiative for Electronic Simulations with Thousands of Atoms (SIESTA)” code [42,43]. In order to study the electronic properties and correlation effects, we adopted the General Gradient Approximation. There is evidence that conventional DFT methods fail to precisely address the van der Waals (vdW) forces and cannot capture the dispersion interactions, despite their crucial role in weakly interacting systems [44]. Therefore, the standard DFT in the GGA approximation has been complemented by an ab initio vdW tactic offered by Grimme, known as the vdW-DFT method. It allows the incorporation of dispersive vdW interactions into the DFT. We also used split-valence double-ξ basis sets of the localized numerical atomic orbitals throughout the modeling process including polarization functions (DZPs). The energy shift was set to 50 meV and the split norm was 0.25. Furthermore, a 5 × 5 × 1 Monkhorst–Pack grid was used for the k-point sampling of the Brillouin zone, while the atomic locations were relaxed until the residual forces on the atoms were below 0.02 eV/Å [43,45]. Moreover, to investigate the charge density, the mesh cutoff was set to 120 Ry applying Basis Set Superposition Error (BSSE) corrections by adding ghost atoms at all stages, which allowed us to calculate the isolated adsorbent. The use
of BSSE corrections is necessary for these kinds of calculations in order to provide accurate energies when the atoms of the system interact [46]. A periodic boundary condition was imposed on graphene and graphene-like surfaces such that the vacuum height was chosen to be equal to 20 Å. This value could eliminate slab–slab interactions. The selected supercell contained 50 atoms [47]. The systems comprising adsorbents and adsorbates were initially modeled in different possible configurations and were then allowed to relax throughout the complete optimization process. According to the BSSE corrections, the adsorption energies of the H$_2$S molecule interacting with the surfaces of BeO, ZnO, pristine graphene, and Ni-decorated graphene were calculated using the following equation:

$$E_{\text{BSSE}} = E(\text{surf}/H_2S) - E(\text{surf}_{\text{ghost}}/H_2S) - E(\text{surf}/H_2S_{\text{ghost}})$$  \hspace{1cm} (1)

where $E_{\text{BSSE}}$ represents the interaction energy and $E(\text{surf}/H_2S)$ stands for the total energy of the nanosheet surface (surf) interacting with the H$_2$S molecule. The “ghost” in $E(\text{surf}_{\text{ghost}}/H_2S)$ and $E(\text{surf}/H_2S_{\text{ghost}})$ terms corresponds to the counterpoise corrections that use “ghost” atoms. In other words, these terms correspond to additional basis wave functions centered at the position of the H$_2$S molecule or the nanosheet surface (pure graphene, ZnO, BeO, or Ni-decorated graphene), but without any atomic potential. Therefore, a negative value of $E_{\text{BSSE}}$ reflects an energetically stable adsorption configuration. Since all the calculations in this work were made at the constant temperature of 0 K, whereas in experimental work, temperature must be considered as an effective factor, it is not be reasonable to compare these theoretical results with the experiments. The latter are therefore not provided here.

3. Results and Discussion

3.1. Geometrical Design and Optimization

According to the optimized H$_2$S structure, the H-S bond length and H-S-H bond angle were obtained as 1.360 Å and 91.68°, respectively. The values obtained are in close agreement with those reported in other studies [7,14,22]. A similar process was performed for the graphene-like ZnO and BeO structures, and the optimized structures are presented in Figure 1. The Be-O and Zn-O bond lengths were 1.57 and 1.87 Å, respectively, which is close to the values reported in previous DFT-based studies [48–53].

![Figure 1. Optimized structures and the corresponding density of states (DOS) plots for: (a) BeO and (b) ZnO nanosheets (the Fermi level is set to zero).](image-url)
In addition, Figure 1 shows the DOS diagrams of the two nanostructures. The DOS illustrates the electronic properties of the structures through the bandgap energy \( E_g \), which takes into account the energy difference between the highest occupied molecular orbital (HOMO) and the lowest unoccupied molecular orbital (LUMO). \( E_g \) is the minimum energy required to excite electrons from the valence band to the conduction band, so that a lower \( E_g \) means a higher electrical conductivity and a higher adsorption potential of the structure. As shown in Figure 1, the bandgap energy of the ZnO surface is equal to 2.2 eV, which is much lower than that of the BeO surface, 6.64 eV. The same kind of result has been provided by previous research regarding the bandgap of ZnO \([54,55]\) and BeO \([56,57]\) nanosheets, demonstrating a close agreement with our own results. Thus, ZnO sheets are more conductive than BeO sheet and hence, should exhibit higher adsorption potential when interacting with H₂S.

In the next step, we brought the H₂S molecule closer to the surface of the studied sheets in various situations. Figure 2 shows 10 different configurations as the scenarios reflecting the adsorption behavior of H₂S with H or S atoms on the hexagonal rings of the ZnO or BeO sheets. These configurations are only initial possibilities, and the position of H₂S on the surface varies during the optimization process until the molecule chooses the one that releases the most energy. These 10 initial configurations are thus sufficient to find the optimized structure. Configurations called H correspond to the sites on which the H₂S molecule approaches the surface by one of its H atoms, whereas configurations called S correspond to the sites where H₂S is brought close to the surface by its S atom. It should also be noted that in all configurations, H₂S molecules are parallel to the metal oxide surfaces. During the optimization process, the molecules rotate and select their most stable situation with the lowest equilibrium distance as well as the highest adsorption energy. The adsorption energies of H₂S at different sites of the ZnO and BeO sheets are also given in Table 1. As seen, the maximum energy released to achieve the most stable H₂S configurations is on the H2 site of BeO and the H1 site of ZnO sheets with adsorption energies of −0.144 and −0.376 eV, respectively.

![Figure 2](image_url)

**Figure 2.** Schematic view of H₂S interacting with ZnO (or BeO) nanosheet in ten various adsorption configurations.

### 3.2. Adsorption of H₂S on Graphene-Like Metal Oxide Surfaces

As shown in Table 1, H₂S was adsorbed at all ten adsorption sites on the BeO surface with consistently negative adsorption energies, the strongest adsorption occurring at the H2 site with a value of −0.144 eV. This same site resulted in a positive adsorption energy for H₂S on graphene-like ZnO, unlike the nine other sites.

Side views of H₂S adsorption on BeO and ZnO nanosheets in the most stable configurations are displayed in Figure 3a. This figure indicates a weak interaction of the hydrogen atom of the H₂S molecule with the oxygen of the BeO surface at an equilibrium distance of 2.709 Å. Furthermore, the corresponding adsorption energy of the H1 site is −0.376 eV with the equilibrium distance of 2.820 Å.
between the sulfur atom of H$_2$S and the zinc atom of the ZnO sheet. The DOS plot of the H$_2$S/BeO complex at the H2 site is displayed in Figure 3b, and it shows a reduction in the $E_g$ from 6.64 eV for BeO alone to 5.65 eV after adsorption of H$_2$S. In addition, as clearly seen in Figure 3b, the DOS plot reveals that the bandgap energy is decreased from 2.2 eV for ZnO alone to 2.1 eV after adsorption of H$_2$S on the H1 site. Such a small difference in bandgap shows that the H$_2$S adsorption does not have a significant effect on the electrical properties of the ZnO nanosheet [58]. The results obtained by calculating the adsorption energy and the DOS confirm a physisorption interaction between H$_2$S and BeO as well as ZnO sheets, while the highest energy of adsorption of H$_2$S on the ZnO surface (−0.376 eV) is about three times higher than that of the same molecule on BeO (−0.144 eV).

**Table 1.** Calculated adsorption energies of H$_2$S on ZnO and BeO nanosheets at the different adsorption sites shown in Figure 2.

| Adsorption Site | Adsorption Energy (eV) on BeO Nanosheet | Adsorption Energy (eV) on ZnO Nanosheet |
|----------------|---------------------------------------|---------------------------------------|
| H1             | −0.096                                | −0.376                                |
| H2             | −0.144                                | 0.011                                 |
| H3             | −0.114                                | −0.199                                |
| H4             | −0.101                                | −0.201                                |
| H5             | −0.110                                | −0.239                                |
| H6             | −0.134                                | −0.236                                |
| S1             | −0.127                                | −0.217                                |
| S2             | −0.141                                | −0.041                                |
| S3             | −0.124                                | −0.192                                |
| S4             | −0.112                                | −0.359                                |

**Figure 3.** (a) Side views of the optimized structures and geometrical parameters of the adsorption of the H$_2$S molecule on the surface of BeO and ZnO nanosheets at sites H2 and H1, respectively; (b) DOS calculated for BeO and ZnO nanosheets before and after adsorption of the H$_2$S molecule at sites H2 and H1, respectively (the Fermi level is set to zero).
The general conclusion of the results discussed in this section is that neither BeO nor ZnO seem to be appropriate materials for H$_2$S gas sensors, either because of a too weak interaction between H$_2$S and the surface or because of a too low impact on the electrical resistance of the device. Since the synthesis of such kinds of nanosheets is also an expensive and time-consuming process, one may need to find other alternatives from more available and less expensive materials such as graphene while improving the sensing properties of graphene. Decorating the latter with various metals such as Ni could be an appropriate solution.

3.3. Adsorption of H$_2$S on Pristine and Ni-Decorated Graphene Nanosheets

At this stage, first, we optimized the geometrical parameters of pristine graphene based on the DFT calculation. From the observed results, we found that the length of the C-C bond of graphene in the optimized situation is about 1.42 Å, which corresponds best to the previous DFT calculations [59–61] and is also identical to the well-known length of the C-C bond in the graphite structure [62].

The adsorption of the H$_2$S molecule on pristine graphene was studied by DFT calculation to compare it with the result of Ni-decorated graphene sheet. Figure 4 shows seven different configurations for the adsorption of the H$_2$S molecule on graphene sheet. We should also mention that for all configurations, H$_2$S molecules are parallel to the surface of pristine graphene. Moreover, the calculated adsorption energy of H$_2$S at seven adsorbed sites are listed in Table 2 in order to select the most stable adsorption configuration for the H$_2$S molecule. As it is obvious from Table 2, H4 is the most stable site with the corresponding minimum released energy of −0.119 eV. As can be seen in the DOS graph presented in Figure 5a, the Dirac point located at the Fermi level and the bandgap energy of pristine graphene is zero. The side view of the most stable adsorption structure (H4) of the H$_2$S molecule on the pristine graphene surface and the corresponding DOS plot are presented in Figure 5b. As shown in this figure, 0.011 e was transferred from H$_2$S to the graphene and the equilibrium distance between them after adsorption was 3.564 Å. The charge transfer from H$_2$S to the graphene increased its conductivity, which is also observable as the dispersion that occurred in the DOS plot of graphene after H$_2$S adsorption (see Figure 5b). Similar results were obtained experimentally by Fattah and Khatami, who showed that the electron transfer from H$_2$S to graphene during the adsorption process is due to the polarity of H$_2$S [34]. It should also be noted that the adsorption energy of the graphene/H$_2$S complex (−0.119 eV) is of the same order of magnitude as that reported elsewhere: −0.15 eV [27] by Cortés-Arriagada et al. using the Perdew–Burke–Ernzerhof (PBE) functional as used in this work, and −0.1 eV in the experimental study presented by Ovsianytskiy et al. [34]. However, there are also different results in this respect, so that the results for the highest adsorption energy of the graphene/H$_2$S complex presented by Ganji et al. (0.02 and −0.04 eV with and without BSSE) were different from those of other studies using Perdew–Burke–Ernzerhof (PBE) and Ceperly–Alder (CA) [26]. As for the C-S distance, a value of about 3.6 Å had indeed been calculated in [27,28]. It should be stressed that the eV values depend on the methodology and assumptions involved in the development of the model and the software used. Comparing the highest adsorption energy obtained for the H$_2$S/pristine graphene complex (−0.119 eV) and the highest adsorption energies already calculated for the H$_2$S/BeO (−0.144 eV) and H$_2$S/ZnO (−0.376 eV) complexes, it can be concluded that metal oxide surfaces seem to be more effective as H$_2$S sensing materials than pristine graphene. Furthermore, the larger equilibrium distance of H$_2$S adsorption on pristine graphene (3.564 Å) compared to the distances previously observed for H$_2$S/BeO (2.709 Å) and H$_2$S/ZnO (2.820 Å) can be seen as more evidence. Overall, it seems that pristine graphene fails to give a sensor of H$_2$S at room temperature, which is why decorating graphene with a Ni atom was practiced to improve its adsorption properties. Especially, the shift from physisorption to chemisorption is necessary for adsorbent–adsorbate interactions to occur and to be maintained at room temperature.
Figure 4. Schematic view of H$_2$S interacting with graphene sheet in seven adsorption configurations.

Table 2. Calculated adsorption energies of H$_2$S on pristine graphene at the different adsorbed sites shown in Figure 4.

| Adsorption Site | Adsorption Energy (eV) |
|-----------------|------------------------|
| H1              | −0.113                 |
| H2              | −0.112                 |
| H3              | −0.118                 |
| H4              | −0.119                 |
| S1              | −0.117                 |
| S2              | −0.114                 |
| S3              | −0.116                 |

Figure 5. DOS plots for pristine graphene sheet: (a) before and (b) after adsorption of the H$_2$S molecule at site H4 (the Fermi level is set to zero).

As mentioned above, the H$_2$S molecule adsorbs on the pristine graphene by weak physical interaction. Decorating the graphene surface with Ni atoms might improve the affinity of H$_2$S molecules for the surface through specific interactions with Ni atoms. In this regard, three different
sites were considered for decorating the graphene surface with a Ni atom, called top (i.e., at the top of a carbon atom), hexagonal (i.e., in the hexagonal well formed by an aromatic ring), and bond (i.e., between two nearest neighboring C atoms), as illustrated in Figure 6. According to the literature, the most stable configuration of Ni-decorated graphene can be obtained by adsorption of the Ni atom on the hexagonal site of pristine graphene [63]. For this reason, this configuration was selected in the present work, and the DOS plot of the Ni-decorated structure on the hexagonal site of pristine graphene is also presented in Figure 7a. Nickel substitution in the graphene sheet has not been considered because such a configuration is very unlikely due to the considerable differences in electronic properties between the nickel and carbon elements and their differences in size.

Figure 6. Top, bond, and hexagonal adsorption sites of Ni on the graphene sheet surface.

Figure 7. (a) Optimized structures of Ni-decorated graphene on the hexagonal site and corresponding DOS plot; (b) side view of the optimized structure and geometrical parameters and DOS plot calculated after adsorption of the H$_2$S molecule on the surface of Ni-decorated graphene (the Fermi level is set to zero).

Figure 7a shows that the DOS plot has been compressed around the Fermi level relative to the DOS plot of pristine graphene, demonstrating the formation of chemical bonds between Ni and carbon atoms of the surface. Moreover, Figure 7b displays an optimized structure and charge transfer of the H$_2$S molecule adsorbed on the surface of the Ni-decorated graphene complex. This figure shows that H$_2$S has been adsorbed on the Ni atom of the Ni-decorated graphene. As seen in Figure 7b, the Mulliken population analysis indicates that a 0.044 charge (i.e., 0.175e–0.11e–0.109e) was transferred from the H$_2$S molecule to the Ni/graphene complex. In other words, 0.044 electrons per H$_2$S molecule
is introduced into the Ni-doped graphene, which is almost four times higher than what we had already calculated for the H$_2$S/graphene complex (0.011 eV), thus corresponding to a decrease in the electrical resistance measured as a sensor response. In this configuration, the calculated adsorption energy of the H$_2$S molecule on the Ni-decorated graphene revealed a released energy of −1.645 eV (against −1.19 eV for Fe-doped graphene [27]). The adsorption strength is thus increased by almost 1300% compared to that of pristine graphene, which confirms the formation of a covalent bond in the H$_2$S/Ni-decorated graphene structure, i.e., it corresponds to a chemisorption process. Compared to the results obtained earlier in this study, the Ni-S bond length is much shorter than the equilibrium distance between H$_2$S and other surfaces, in good agreement with the higher adsorption energy for the Ni-decorated surface.

In the research conducted by Ganji et al. [26,40], the adsorption energies of the H$_2$S molecule on Al-, Ga-, and Pt-decorated graphene are −0.13, −0.09, and −6.37 eV, respectively. Comparison of the results of this paper with those of previous studies shows that the Ni-decorated graphene can give appropriate results in terms of absorption of the H$_2$S molecule. Indeed, the adsorption energy is much higher than on Al and Ga, thus allowing a good interaction with the surface (chemisorption) at room temperature, but not as high as with Pt, thus allowing a good recoverability of the sensor.

The partial density of state (PDOS) for the most stable structures of the BeO and ZnO nanosheets, as well as those of pristine and Ni-decorated graphene, were also calculated and the corresponding plots are shown in Figure 8a–d to illustrate the orbital contribution and the change in electronic structure of these surfaces after H$_2$S adsorption. Figure 8a demonstrates a quite weak interaction between the 2p orbital of oxygen and the 1s orbital of hydrogen, which confirms the physisorption process of H$_2$S molecule on the BeO surface. Likewise, in the case of H$_2$S molecule adsorption on the ZnO surface, a weak interaction is obvious in Figure 8b. Figure 8c,d illustrate the orbital hybridization of H$_2$S with pristine and Ni-decorated graphene surfaces, respectively. Figure 8c,d can be compared since there is no significant overlap between the 2p orbital of carbon and the 3p orbital of sulfur, while a strong hybridization of the Ni 3d and S 3p orbitals (especially in the range 1–4 eV; see the inset in Figure 8d) can be observed. Therefore, the strongest chemical adsorption of H$_2$S, which occurred on Ni-decorated graphene, might be attributed to the hybridization of the orbitals with the very good overlap that can be seen in Figure 8d [64].

![Figure 8](image-url). Partial density of states (PDOS) plots of H$_2$S adsorption on: (a) BeO, (b) ZnO, (c) pristine graphene, and (d) Ni-decorated graphene sheets.
4. Conclusions

In this work, the adsorption properties of graphene-like metal oxide surfaces such as BeO and ZnO with respect to the H$_2$S molecule were evaluated by DFT calculations. These properties were compared to the adsorption properties of pristine and Ni-decorated graphene sheets. The results showed that the H$_2$S molecule physically adsorbs on BeO and ZnO surfaces with adsorption energies of $-0.144$ and $-0.376$ eV, respectively. Similarly, in the case of pristine graphene, the H$_2$S molecule physisorbs on the surface, releasing an energy of $-0.119$ eV. In contrast, the DFT calculations demonstrated a chemisorption process of the H$_2$S molecule on the Ni-decorated graphene sheet with an adsorption energy of $-1.64$ eV, which is due to a strong interaction of H$_2$S with the Ni sites of the decorated surface. The PDOS plots also showed an overlap between the 3p and 3d orbitals of S and Ni atoms, respectively, which results from orbital hybridization during the formation of a covalent bond. Based on the results obtained in this study, it can be concluded that graphene sheet decorated with Ni atom has a great potential for the adsorption of H$_2$S molecules, and might be used for developing new gas sensors. Overall, the transition of the adsorption mechanism from physisorption to chemisorption was successfully captured by the model developed for use in future investigations.

Author Contributions: A.H.M. and S.H. designed the case study and carried out the computational calculations. Z.K. designed the figures and tables. A.S. analyzed and rechecked the results for correctness and wrote the original draft of the article. M.R.G., V.V., A.E. and M.R.S. analyzed the data and discussed the results. V.F. and A.C. reviewed and edited the article to give it its final form. All authors have read and agreed to the published version of the manuscript.

Funding: This research received no external funding.

Conflicts of Interest: The authors declare no conflict of interest.

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