Hydrophosphination using [GeCl{N(SiMe₃)₂}₃] as a pre-catalyst†

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Transformations catalyzed by germanium are scarce, with examples mainly limited to widely catalyzed processes such as polymerisation of lactide and hydroboration of carbonyls. Reported is the first example of hydrophosphination using a germanium pre-catalyst, yielding anti-Markovnikov products when diphenylphosphine is reacted with styrenes or internal alkynes at room temperature.

Over the last two decades, the use of main group compounds as catalysts for organic reactions has seen a vast increase in focus.¹⁻³ Amongst this work, significant progress has been made in the field of germanium catalysis. A fundamental challenge lies in unlocking the catalytic potential of this group 14 element, which is far behind tin in terms of its exploitation in catalysis, and such work may prove instrumental to furthering understanding and harbouring new discoveries in main group catalysis. It is not surprising, based on the ability of Sn(IV) complexes to act as polymerization catalysts,⁴ that initial disclosures in the literature demonstrated the ability of Ge(IV) complexes to act as initiators for the ring-opening polymerization of lactide (Scheme 1a).⁵⁻⁷ Subsequently, there have been numerous reports on low-coordinate Ge(n) species acting as efficient pre-catalysts for the hydroboration of carbonyls (Scheme 1b).⁸⁻¹¹ This reactivity has also been extended to the catalytic cyanosilylation of aldehydes.¹²,¹³

Despite this progress, lactide polymerization and carbonyl hydrofunctionalization are commonly reported catalytic transformations,¹⁴,¹⁵ therefore a desire remains to uncover use for germanium (both Ge(n) and Ge(n)) in the catalysis of a wider range of transformations.¹⁶ It is clear from the current literature that, in the presence of at least one labile co-ligand, Ge(IV) complexes can catalyze non-redox, σ-bond metathesis type transformations such as lactide polymerization. This indicates that opportunities exist for Ge(IV) catalysts, particularly in hydrofunctionalization reactions with pertinent substrates.

The hydrophosphination (HP) of unsaturated hydrocarbons is an atom economical process of forming P–C bonds.¹⁷ Using a main group element to transform another is an area of growing interest, amongst which examples of main group catalyzed HP
have emerged.\textsuperscript{18} Reaction mechanisms for these transformations involving main group, and lanthanide,\textsuperscript{19–23} complexes are often postulated to proceed via $\sigma$-bond metathesis type steps at the catalyst center,\textsuperscript{24} leading us form the hypothesis that a Ge catalyst may have success in this area. To the best of our knowledge, no examples of Ge-catalyzed HP have been reported. However, the heavier group 14 element, Sn, has been shown to be active catalytically in the HP of alkenes and alkynes by Waterman and co-workers, who, in 2014, reported a modest scope of HP reactions using Ph₂PH.\textsuperscript{25} Subsequent studies on the same Sn(ν) system using Ph₃PH in 2016 expanded the scope to further substrates.\textsuperscript{26} However, these transformations required an atmosphere of H₂ due to the catalyst’s propensity to undergo homodehydrocoupling of the phosphine substrate over the desired HP reaction. This, along with the elevated temperatures necessary (65 °C) and high catalyst loading (10 mol%), makes the reaction only marginally more competitive than the analogous catalyst-free, thermal process.\textsuperscript{27,28} We were therefore keen to investigate if a more active catalyst could be found within group 14 in the form of a Ge(IV) species. Rather than start our studies with an over-engineered catalyst structure, and thus be in a strong position to build in ligand design at a later stage, based on understanding. We herein present the first HP of alkenes using a structurally simple Ge(ν) pre-catalyst under mild conditions (Scheme 1c).

Reaction of GeCl₄ with 10 equiv. of Na(HMDS) yields [GeCl[N(SiMe₃)₂]₂]₂, \textsuperscript{1} in good yield (63%, Scheme 2), the structure of which was confirmed by X-ray diffraction.\textsuperscript{29} In an analogous reaction, Thomas and co-workers propose the formation of tetra(amido) THF adduct [Ge(HMDS)₄(THF)₂]:\textsuperscript{7} in our synthesis $^1$H DOSY analysis indicates lack of persistent THF coordination in solution.\textsuperscript{29}

Scheme 2 Synthesis and single crystal X-ray structure of 1 displaying the inherent C$_{3v}$ symmetry of the molecule. Ellipsoids are represented at 30% probability and hydrogen atoms have been omitted for clarity. Symmetry operations: \textit{1} $+ y - x$, $-1 - y$, $z$; \textit{2} $- x + 1$, $y$, $z$; \textit{3} $-1 - y$, $x$, $z$. Bond distance (Å) and angles (°): Ge1–Cl1, 2.2154(11); Ge1–N1, 1.8542(16); Cl–Ge–N1, 102.12(6); N1–Ge–N1’, 115.71(4).

Table 1 Reaction optimization for the hydrophosphination of styrene using pre-catalyst 1

| Entry | Catalyst loading/mol% | Solvent | Spectroscopic yield $^a$/% |
|-------|------------------------|---------|--------------------------|
| 1     | 5                      | C₆D₆    | 91                       |
| 2     | 2.5                    | C₆D₆    | 81                       |
| 3     | 1                      | C₆D₆    | 38                       |
| 4$^c$ | 5                      | THF     | 99                       |
| 5$^c$ | 5                      | C₆D₆    | 94                       |

$^a$ Conditions: styrene (0.5 mmol, 57 µL), Ph₂PH (0.5 mmol, 87 µL), 1 (5 mol%, 0.025 mmol, 14.7 mg), C₆D₆ (0.6 mL). $^b$ Spectroscopic yield calculated via integration against 1,3,5-trimethoxybenzene internal standard. $^c$ 0.6 mmol styrene.

to the anti-Markovnikov HP product 2a after 18 hours at RT with a catalyst loading of 5 mol%, using C₆D₆ as solvent (Table 1, entry 1). Only trace quantities of homodehydrocoupling by-product P₂Ph₄ is formed, therefore any concerns over a competing reaction of this type prove insignificant. Markovnikov product is never observed. Catalyst loadings of 2.5 mol% and 1 mol% result in the reduction of spectroscopic yield to 81% and 38% respectively (entries 2 and 3). Notably, no reaction is observed when GeCl₄ is used as the catalyst. Change of solvent to THF gives similar conversion compared to benzene (entry 4). Other solvents (MeCN (vide infra), DCM, CDCl₃) show reactivity with 1 and are therefore not suitable for the HP reaction.\textsuperscript{29} A modest excess of styrene (1.2 equiv.) gives 94% conversion in an inexpensive NMR solvent (entry 5).

With the optimized conditions in hand (Table 1, entry 5) we then set out to investigate the substrate scope for the reaction. The reaction proceeds with a wide range of styrenes (Table 2). Biphenyl-substituted phosphine 2b is formed in excellent yield (entry 2), p-Me-, p-MeO- and p-CF₃-substituted phosphines (2c–2e) are formed in reduced conversions (entries 3 to 5). Halide substituents are well tolerated, with phosphine products 2f to 2h formed in good to moderate conversion (entries 6 to 8). m- and o-substituted phosphines (2i to 2m) are also formed in good yield (entries 9 to 13).

Naphthyl substituted phosphine 2n is formed in excellent yield (entry 14) and 1,1-diphenylethylene performs comparably well, albeit under slightly more forcing conditions (50 °C, 24 h, entry 15). Phenyl acetylene does not react, but internal alkynes (entries 16 to 19) react at RT: conditions that are highly competitive compared to the leading literature, for example Schmidt’s RT, La-catalyzed process in pyridine solvent\textsuperscript{30} and Cui’s Ca- and Y-mediated chemistry that uses temperatures ranging from RT-75 °C.\textsuperscript{11} Unfortunately, reactions of 2- and 4-vinylpyridine and acrylonitrile with 1 all result in the formation of unidentified polymers irrespective of order of addition of Ph₂PH. Methyl acrylate, butyl acrylate, cis- and trans-stilbene, along with β-methyl styrene do not react. HP reactions using unactivated alkenes (e.g. cyclohexene, allyl benzene) result in no conversion while no stoichiometric reaction of 1
and HPCy₂, or catalytic HP of styrene using 1 and HPCy₂, is observed.²⁹ However, scale-up synthesis of 2a is successful: using 5 mmol Ph₂PH (0.93 g/0.87 mL), 6 mmol styrene, 2.5 mol% Ge, at RT gives 94% after 24 h. Catalyst robustness was tested through an iterative addition protocol: 0.5 mmol Ph₂PH, 0.6 mmol styrene, 5 mol% Ge, at RT gives 94% after 24 h. Catalyst robustness was observed.²⁹ However, scale-up synthesis of 1 requires further optimization.

Focus was then turned to investigating the mechanism of the reaction. A kinetic study of the reaction to form 2a was undertaken via variable time normalization analysis,³² implicates the reaction is approximately first order with respect to 1 and Ph₂PH and a more complex relationship with styrene (≤ 1, ≥ 0.5), in-line with a complex reaction where multiple species can propagate from one Ge-center.²⁹ There is no reaction when a stoichiometric quantity of styrene is added to 1. Conversely, reaction of Ph₂PH with 1 results in the formation of a bright yellow color, which is also observed during catalysis.¹H DOSY analysis of this species shows fast diffusing HN(SiMe₃)₂ peaks and slower diffusing aryloxyl resonances, consistent with protonation of Ge-bound N(SiMe₃)₂ by HPPH₂.²⁹ Monitoring catalysis (15 mol% 1) as a function of time by ³¹P NMR spectroscopy shows the Ph₂PH doublet decreasing in intensity as product 2a, and a singlet at ~36 ppm, increase. The signal at ~36 ppm reaches a maximum when all Ph₂PH is consumed (~60% 2a formed), and then shows a slow decrease of this signal into 2a for the remainder of the reaction. There is also a very small singlet at ~27 ppm that decreases in intensity slowly over the course of the reaction. Our data is consistent with bis- and tris-phosphido complexes (~27 and ~36 ppm respectively) due to the loss of P–H, the fact that integration of this signal into 2a

### Table 2 Substrate scope for HP using pre-catalyst 1<br>

| Entry | Alkene | Product | Spec. yield (%) |
|-------|--------|---------|----------------|
| 1     | R = H  | 2a      | 94             |
| 2     | 4-Ph   | 2b      | 98             |
| 3     | 4-Me   | 2c      | 94             |
| 4     | 4-OMe  | 2d      | 56             |
| 5     | 4-Cl   | 2e      | 83             |
| 6     | 4-F    | 2f      | 76             |
| 7     | 4-Br   | 2g      | 94             |
| 8     | 3-Me   | 2i      | 92             |
| 9     | 3-Br   | 2j      | 82             |
| 10    | 2-Me   | 2k      | 89             |
| 11    | 2-OMe  | 2l      | 89             |
| 12    | 2-Br   | 2m      | 84             |

*Conditions: styrene (0.6 mmol), Ph₂PH (0.5 mmol, 87%), 30 °C, 24 h, 18 h, RT. *⁵ Spectroscopic yield calculated via integration of product peak against a trimethoxybenzene internal standard. ³² 30 °C, 24 h, 18 h, RT, 72 h. Spec. yield and E:Z ratio calculated via ³¹P inverse gated NMR spectroscopy against a P(OEt)₃ internal standard. ³³ 30 °C, 24 h, 18 h, RT, 48 h, Spec. yield and E:Z ratio calculated via ³¹P inverse gated NMR spectroscopy against a P(OEt)₃ internal standard.

During solvent screening, we also observed that 1 could catalytically induce the trimerization of MeCN to 2,6-dimethylpyrimidin-4-amine (Scheme 4). Base-mediated

![Scheme 3 Proposed method by which 1 catalyzes the HP of styrenes. 3bis is only observed in trace quantities.](image)
cascade reactions\textsuperscript{34,35} and an iridium hydride catalyst\textsuperscript{36} have previously been used to obtain the same product under more forcing conditions (e.g. Murahashi’s Ir-catalyzed method uses IrH\textsubscript{5}(P\textsubscript{3}Pr\textsubscript{3})\textsubscript{2} at 140 °C for 12 h giving TON = 12). This reactivity represents a new avenue of reactions this germanium species is capable of catalyzing and will be explored in due course.

In summary, we have disclosed a rare example of a Ge(IV)-catalyzed (non-polymerization) reaction. An operationally simple Ge(IV) tris(amido) complex has shown excellent ability as a pre-catalyst for the hydrophosphination of a range of styrenes and internal alkynes. The transformations proceed at room temperature over a time period competitive with transition metal catalyzed analogues. Preliminary mechanistic studies implicate the formation of a Ge-(tris)phosphido complex as the active catalyst and a redox-neutral mechanistic pathway was postulated. We hope that this work, using such a simple pre-catalyst, serves as a platform for future research in germanium-catalyzed hydrofunctionalization reactions.

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Conflicts of interest

There are no conflicts to declare.

Notes and references

1. P. P. Power, \textit{Nature}, 2010, \textbf{463}, 171–177.
2. T. J. Hadlington, M. Driess and C. Jones, \textit{Chem. Soc. Rev.}, 2018, \textbf{47}, 4176–4197.
3. C. Weetman and S. Inoue, \textit{ChemCatChem}, 2018, \textbf{10}, 4213–4228.
4. J. Wu, T.-L. Yu, C.-T. Chen and C.-C. Lin, \textit{Coord. Chem. Rev.}, 2006, \textbf{250}, 602–626.
5. A. Finne, Reema and A.-C. Albertsson, \textit{J. Polym. Sci., Part A: Polym. Chem.}, 2003, \textbf{41}, 3074–3082.
6. A. J. Chmura, C. J. Chuck, M. G. Davidson, M. D. Jones, M. D. Lunn, S. D. Bull and M. F. Mahon, \textit{Angew. Chem. Int. Ed.}, 2007, \textbf{46}, 2280–2283.
7. J. Guo, P. Haquette, J. Martin, K. Salim and C. M. Thomas, \textit{Angew. Chem. Int. Ed.}, 2013, \textbf{52}, 13584–13587.
8. T. J. Hadlington, M. Hermann, G. Frenking and C. Jones, \textit{J. Am. Chem. Soc.}, 2014, \textbf{136}, 3028–3031.
9. Y. Wu, C. Shan, Y. Sun, P. Chen, J. Ying, J. Zhu, L. Liu and Y. Zhao, \textit{Chem. Commun.}, 2016, \textbf{52}, 13799–13802.
10. J. Schneider, C. P. Sindlinger, S. M. Freitag, H. Schubert and L. Wesemann, \textit{Angew. Chem. Int. Ed.}, 2017, \textbf{56}, 333–337.
11. S. Sinhababu, D. Singh, M. K. Sharma, R. K. Siwatch, P. Mahawar and S. Nagendran, \textit{Dalton Trans.}, 2019, \textbf{48}, 4094–4100.
12. R. K. Siwatch and S. Nagendran, \textit{Chem. – Eur. J.}, 2014, \textbf{20}, 13551–13556.
13. R. Dasgupta, S. Das, S. Hiwase, S. K. Pati and S. Khan, \textit{Organometallics}, 2019, \textbf{38}, 1429–1435.
14. O. Dechy-Cabaret, B. Martin-Vaca and D. Bourissou, \textit{Chem. Rev.}, 2016, \textbf{114}, 6147–6176.
15. C. C. Chong and R. Kinjo, \textit{ACS Catal.}, 2015, \textbf{5}, 3238–3259.
16. T. Sugahara, J. D. Guo, T. Sasamori, S. Nagase and N. Tokitoh, \textit{Angew. Chem. Int. Ed.}, 2018, \textbf{57}, 3499–3503.
17. I. Wauters, W. Debrouwer and C. V. Stevens, \textit{Beilstein J. Org. Chem.}, 2014, \textbf{10}, 1064–1096.
18. M. S. Hill, D. J. Liptrot and C. Weetman, \textit{Chem. Soc. Rev.}, 2016, \textbf{45}, 972–993.
19. M. R. Douglass, C. L. Stern and T. J. Marks, \textit{J. Am. Chem. Soc.}, 2001, \textbf{123}, 10221–10238.
20. A. M. Kawaoka, M. R. Douglass and T. J. Marks, \textit{Organometallics}, 2003, \textbf{22}, 4630–4632.
21. I. V. Basalov, S. C. Rosca, D. M. Lyubov, A. N. Selikhov, G. K. Fukin, Y. Sarazin, J.-F. Carpenter and A. A. Trifonov, \textit{Inorg. Chem.}, 2014, \textbf{53}, 1651–1661.
22. I. V. Basalov, V. Dorcet, G. K. Fukin, J.-F. Carpenter, Y. Sarazin and A. A. Trifonov, \textit{Chem. – Eur. J.}, 2015, \textbf{21}, 6033–6036.
23. J. Yuan, H. Hu and C. Cui, \textit{Chem. – Eur. J.}, 2016, \textbf{22}, 5778–5785.
24. L. Rosenberg, \textit{ACS Catal.}, 2013, \textbf{3}, 2845–2855.
25. K. A. Erickson, L. S. H. Dixon, D. S. Wright and R. Waterman, \textit{Inorg. Chem. Acta}, 2014, \textbf{422}, 141–145.
26. J. P. W. Stelmach, C. A. Bange and R. Waterman, \textit{Dalton Trans.}, 2016, \textbf{45}, 6204–6209.
27. F. Alonso, Y. Moglie, G. Radiyov and M. Yus, \textit{Green Chem.}, 2012, \textbf{14}, 2699–2702.
28. Y. Moglie, M. J. Gonzalez-Soria, I. Martin-Garcia, G. Radiyov and F. Alonso, \textit{Green Chem.}, 2016, \textbf{18}, 4896–4907.
29. See ESI.
30. M. M. I. Basiouny, D. A. Dollard and J. A. R. Schmidt, \textit{ACS Catal.}, 2019, 9, 7143–7153.
31. H. Hu and C. Cui, \textit{Organometallics}, 2012, \textbf{31}, 1208–1211.
32. J. Burés, \textit{Angew. Chem. Int. Ed.}, 2016, \textbf{55}, 16084–16087.
33. J. K. West and L. Stahl, \textit{Organometallics}, 2012, \textbf{31}, 2042–2052.
34. R. Bossio, S. Marcaccini, V. Parrini and R. Pepino, \textit{J. Heterocycl. Chem.}, 1986, \textbf{23}, 889–891.
35. A. Verma, S. Jana, C. Durga Prasad, A. Yadav and S. Kumar, \textit{Chem. Commun.}, 2016, \textbf{52}, 4179–4182.
36. H. Talaya, T. Naota and S.-I. Murahashi, \textit{J. Am. Chem. Soc.}, 1998, \textbf{120}, 4244–4245.