SORPTION CHARACTERISTICS OF COOPER SORPTION ON THE CLINOPTILOLITE MEASUREMENT

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1 INTRODUCTION

The sources of toxic metals that enter into water are various human activities, such as operations in the automotive industry, engineering industry, mining and processing of ores, textile industry, waste production [Panda 2014a, Panda 2016b, Panda 2016c, Prisulpcak 2016, Rimar 2016, Straka 2013, Zelenak 2012, Olejarova 2016, Zaborowski 2007]. Many methods are used to reduce the metallic element in water, such as oxidation, reduction, chemical precipitation. Progress in science and technology is also evident in the field of water protection as modern, economically viable and environmentally acceptable cleaning processes and materials such as carbonaceous materials, dendrimers and zeolites are developed [Van Bekkum 2001]. All of them have very varied physicochemical properties which are particularly suitable for separating pollutants from water. Very promising for cleaning processes are specific nanometric nanomaterials applied as membranes or other structured materials [Lehman 2014]. These substances have a higher adsorption surface and may be modified by, for example, incorporating specific functional groups into a carrier. Natural zeolites also belong to such materials.

2 ZEO LITE PROPERTIES

Natural zeolites are hydrated aluminosilicates characterized by their ability to adsorb heavy metals from aqueous solutions. In addition to chemisorption, the metal ions on the zeolite are also immobilized by an ion exchange [Sabova 2010, Chmielewska 2010, Bailey 1999, Macala 2017]. The lattice structure allows the zeolite to act as ion exchange and selective adsorbent. Adsorption and ion exchange depend on their charge and size. The more the ion size coincides with the size of the input pores in the clinoptilolite grid, the easier it will be captured and maintained. The inlet pore diameter is 0.4 nm, which corresponds to the ammonium ions of N\textsubscript{H}_4\textsuperscript{+}, H\textsubscript{2}O\textsuperscript{2-}, 13\textsuperscript{Cs} and 13\textsuperscript{Cs}. These substances have the greatest affinity for binding to clinoptilolite but which act as a selective adsorbent for a wide range of pollutants. As in other microporous materials, as well as in zeolites, the ion exchange sorption between the components of the liquid and solid phases is carried out according to the surface diffusion and the internal diffusion mechanism [Chmielewska 2014, Qiu 2009].

An important mineralogical and chemical parameter of each zeolite is the Si / Al ratio. This value indicates the acidity and alkalinity of the zeolites and also their thermal and chemical stability. Zeolites with a higher Si / Al ratio are more acidic and more stable. The thermal stability limits differ from the different types of zeolites. In addition to the Si / Al ratio, this is also affected by the cationic composition. Surface and ion exchange properties of natural zeolites can be influenced by structural modification. Dealumination is followed by modifying the zeolites by removing aluminum from the tetrahedral coordination of the zeolite grid, increasing the proportion of silicon to aluminum. For organic sorption, sorbent with a higher content of silicon, a sorbent with higher acidity, is more suitable. The action of dilute acids results in partial dealumination and alkali leaching, thereby increasing the number of active centers. This modification of the natural zeolites is carried out under conditions in which the zeolite is not destroyed. The possibility of improving the sorption properties of the zeolites is thus observed [Celiscev 1987]. The action of dilute mineral acids leads to two stages of acid decatination. The first step is the exchange of ions expressible by the following equation:

\[
M^n+(AlO_2)_x(SiO_2)_{n-x}^m+x[H_2O+H^+]->[(H_2O)_x(AlO_2)_x(SiO_2)_{n-x}]^m+xM^n. (2.1)
\]

In the second stage, the dealumination is represented by the equation:

\[
[(AlOOH)_x(SiO_2)_{n-x}]^3xH^+->[H_2O_2x(SiO_2)_{n-x}]^x+xAl^{3+}. (2.2)
\]

One of the other ways of modifying zeolites is boranation. The boronation is due to the fact that B\textsubscript{3}H\textsubscript{4} has an electron deficit and the zeolite in its structure contains sites with increased
electron density. The most reactive sites in the H zeolite structure are:

a) hydroxyl group Si-O-H,
b) Si-O-Si oxygen bridges where oxygen has a high electron density.

Synthetic zeolites are produced as a primary product in the sodium form. Natural zeolite monoole sodium forms can be prepared by treatment with NaOH on the zeolite. Modification of the structure of the zeolites is based on the action of certain chemical solutions on the crystalline lattice of the zeolite. By diluting the alkali hydroxides, silicon is extracted from the structure of the zeolite. Depending on the concentration of the hydroxide-modifying solution, the surface or even internal pores of the zeolite may be partially altered. This change is based on the fact that silicon, which is part of the crystal lattice of clinoptilolite, reacts with the anions (OH) which are formed by the equations [Földesová 2003, Földesová 2007]:

\[
\begin{align*}
\text{Si}^{4+} + 2\text{OH}^- + \text{H}_2\text{O} & \rightarrow 2\text{SiO}_3^{2-} + 2\text{H}_2 \\
\text{SiO}_3^{2-} + 2\text{H}_2\text{O} & \rightarrow 2\text{SiO}_3^{2-} + 2\text{OH}^- 
\end{align*}
\]

Silicon is passed into the solution in the form of soluble silicic acid. In the clinoptilolite increases the proportion of aluminum, and silicate extraction results in a new mesopore structure. At the same time ion exchange of cationic compensating cations in the original clinoptilolite (K+ , Ca2+, Mg2+) takes place under sodium cations. Monocatonic Na+ form zeolite is formation.

3 EXPERIMENTAL

To observe the sorption kinetics, the natural zeolite clinoptilolite with the granuality 2,5 to 5 mm, weighing 30 g, was washed with deionized water and treated in to the mono-ion form with NaOH. This was covered by a model sample of copper cations with a volume of 0,20 dm³. The model sample contained copper cations with an initial concentration of 2,4 g·dm⁻³. At the exact time intervals, the same volume was taken from the model sample during the treatment with clinoptilolite, in which the content of copper cations was determined photometrically. When contacting the solid phase with the solution, the original copper concentration dropped to 0,95 g·dm⁻³. At the equilibrium state after the third hour, the concentration of copper cations in the solution was reduced to 0,95 g·dm⁻³.

When measuring the copper cation adsorption to clinoptilolite, the values given in the Table 1.

| t [min] | 60 | 120 | 180 | 240 |
|--------|----|-----|-----|-----|
| c [g·dm⁻³] | 1,20 | 1,016 | 0,95 | 0,95 |

Table 1. Dependence of copper concentration in solution on time during sorption

The efficiency of clinoptilolite with respect to the copper cations reduction, was calculated according to formula (3.1), its value being 60,4%. The sorption capacity calculated according to (3.2) [Sabová 2010, Inglezakis 2002] was 9,7 mg · dm⁻³.

\[
\eta = \frac{c_0 - c_f}{c_0} \cdot 100 \% \quad (3.1)
\]

\[
a_0 = \frac{(c_0 - c_f)}{m} \cdot V 
\]

The quantitative distribution parameter of the substance between the solid and the liquid phase is the Kp partition coefficient [Korkmaz 2012].

\[
K_p = \frac{c_s}{c_f} \quad (3.3)
\]

The value of this parameter was 1,5.

The kinetic course of sorption is illustrated graphically in Figure 1.

Figure 1. Dependence of copper cation sorption over time

Assuming adsorption takes place according to the Langmire mechanism, and that the sorbed material is sorbed by molecular diffusion on the surface of the adsorbent, the concentration of the adsorbed substance at time t can be expressed by the equation [Pitter 1983]:

\[
c = \frac{\lambda c_f - \beta e^{-\beta t}}{\lambda - e^{-\beta t}} 
\]

For the search relation c = f (t) we calculate the constants of equation (3.1) λ, β, ρ of the measured concentration values for individual time intervals.
According to equation (3.2) for the experimental values of concentration, the constant \( \rho \) and the average of the experimental values were calculated:

\[
\rho = \frac{1}{c} \ln \left( \frac{(c - c_0)(c - \beta)}{(c_0 - \beta)(c - c_0)} \right) \quad (3.5)
\]

Measured and calculated \( \rho \) datas are listed in Table 2.

| t [min] | c [g dm\(^{-3}\)] | \( \rho \) [min\(^{-1}\)] | \( c_0 \) [g dm\(^{-3}\)] |
|---------|-----------------|----------------|-----------------|
| 60      | 1.20            | 0.02           | 2.3             |
| 120     | 1.016           | 0.01           | 2.1             |
| 180     | 0.95            | 0.01           | 1.03            |
| 240     | 0.95            | 0.01           | 1.03            |

\( \rho \) average 0.01

Table 2. Values needed to calculate the equation constants

4 CONCLUSION

From the measured values we have calculated individual values \( \rho \) and \( \rho \) average for individual measurements according to equation (3.5). If the assumption is that the adsorbed material gets molecular diffusion on the surface of the adsorbent then the \( \rho \) values for all measured datas \( c (t) \) have approximately the same values, which is also confirmed.

Measured average value \( \rho \) average = 0.01 min\(^{-1}\). Search equation \( c = f(t) \) is then according to equation (3.4) using the calculated constants:

\[
c_v = 0.95 \lambda - 3.3 \cdot e^{-0.01 \lambda} / \lambda - e^{-0.01}.
\]

(4.1)

Frequent methods of metals removing from solutions, such as precipitation processes, are effective especially at higher concentrations. In such a way, it is not always possible to achieve the limits set by law, and therefore the use of zeolite appears to be a suitable method for purifying waste water after treatment, e.g. precipitation. Natural zeolites are also more cost-effective than synthetic ion exchangers. The efficiency of natural zeolites as sorbents and ion exchangers as well as sorption capacity are influenced by modification.

At present, the annual world production of natural zeolite is about 3 million tons. Due to its good availability and the low cost of its extraction, natural zeolites are still profitable and their mining has great potential in the future. There are still some reserves in their use for environmental protection.

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