Enhanced photocatalytic removal of NO over titania/hydroxyapatite (TiO$_2$/HAp) composites with improved adsorption and charge mobility ability†

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Effective photocatalysis processes with production of fewer toxic intermediates are highly desirable for air purification. In this study, titania/hydroxyapatite (TiO$_2$/HAp) composites were synthesized by a facile hydrothermal method and employed to decontaminate nitric oxide (NO) in air under simulated solar light irradiation for the first time. It was found that the photocatalytic activity of the as-prepared TiO$_2$/HAp composite (44.61%) was superior to those of the pristine components (TiO$_2$: 38.57%, HAp: 36.73%) and mechanically mixed samples (35.36%). The TiO$_2$/HAp composite with mass ratio of 3 : 1 (75% TiO$_2$/HAp) exhibited the highest NO removal efficiency among them. Moreover, the toxic intermediate NO$_2$ production was significantly inhibited over TiO$_2$/HAp. These synergistically improved properties can be ascribed to the high separation efficiency and faster transfer of the photo-generated charge carriers as evidenced by the experimental results from photocurrent tests and electrochemical impedance spectroscopy (EIS). The results from temperature programmed desorption (TPD) confirmed that the 75% TiO$_2$/HAp sample had stronger chemisorption for NO due to the increased concentration of surface OH groups. Furthermore, the electron spin resonance (ESR) characterization suggested that ‘O$_2^-$’ and ‘OH radicals were the major species involved for NO removal over TiO$_2$/HAp composites. The five recycling tests suggested that the TiO$_2$/HAp has superior photocatalytic stability. This study suggests that the combination of TiO$_2$ with HAp is an effective approach for air purification.

1. Introduction

Nitric oxides (NO$_x$) generated from the combustion of fossil fuels have attracted lots of attention because they can lead to various environmental problems, such as acid rain, haze, photochemical smog, ozone depletion and destruction of animal and plant species. The control of NO$_x$ in the atmosphere has become one of the greatest challenges in environmental protection. Several techniques have been developed for removing NO$_x$, such as electrochemical methods, a chemical absorption-biological reduction integrated system, selective catalytic reduction and semiconductor photocatalysis. Among these methods, semiconductor photocatalysis is a promising technology for low concentration air purification, because it allows for the utilization of renewable, safe and clean solar light to solve environmental problems.

Titanium dioxide (TiO$_2$) is one of the most popular photocatalysts in view of its non-toxicity, cheapness and excellent photochemical stability. However, the drawbacks of low adsorption capacity for pollutants and high recombination of photo-generated electron-hole pairs impair its photocatalytic performance. In order to solve these issues, several strategies have been adopted to enhance the photocatalytic efficiency. For example, modification with noble metals on the surface of TiO$_2$ is considered as an efficient way for improving the photocatalytic activity. Besides, the combination of TiO$_2$ with other semiconductor materials of suitable band structure (e.g. C$_3$N$_4$, Ag$_2$CO$_3$, Fe$_2$O$_3$, and Bi$_2$WO$_6$, etc.) can improve the charge separation efficiency and prolong the life of charges, enhancing the photocatalytic activity of TiO$_2$ eventually. However, the above two strategies can not solve the drawback of low adsorption capacity for pollutants. Recently, some literatures reported that TiO$_2$ supported materials, such as zeolite, silica, purity natural diatomite (rich in SiO$_2$ 87–91%), activated carbon, hydroxyapatite (HAp), can offer abundant active adsorption sites, which resulted in quick mass transfer and fast catalytic reactions.
Among these supported materials, hydroxyapatite (HAp), Ca₁₀(PO₄)₆(OH)₂, with mechanical stability, non-toxicity, good biocompatibility, and low cost has been widely studied. Yuichi Komazaki et al. reported that NO₂ was effectively collected on HAp in the annular diffusion scrubber. Due to its inherent hexagonal crystal, HAp possess hydroxyl ions (OH⁻) channels, which can enhance the electrical conductivity and charge carrier migration. The surface OH groups or H₂O molecules adsorbed can react with the h⁺ to produce hydroxyl (OH⁻) radicals. Fu et al. reported that the ZnSn(OH)₄ provided a mass of OH groups and facilitated OH radicals formation to enhance the photocatalytic performance. In addition, the electron state change of the PO₄³⁻ groups in the surface of HAp during the photocatalytic process can result in the generation of O₂⁻ radicals. Dai et al. also found that the oxygen vacancies were formed on the surface of HAp and acted as electrons receiver to achieve the electron–hole pairs separation.

Hence, it is expected that combining the merits of HAp with TiO₂ can not only improve the adsorption capacity of TiO₂ for pollutants, but also decrease the recombination of photo-generated electron–hole pairs. So the combination of TiO₂ and HAp may be an effective approach for air purification. In this paper, the TiO₂/HAp composites were synthesized by a hydrothermal method. The as-prepared TiO₂/HAp composites showed excellent photocatalytic activity on NO₂ removal under the simulated solar light. The physical and chemical properties of the TiO₂/HAp were deeply discussed and the possible photocatalytic mechanism was proposed.

2. Experimental section

2.1. Materials
Titanium tetrachloride (TiCl₄, analytical pure), ammonium hydroxide solution (NH₃·H₂O solution, 25%–28wt%, analytical pure), ammonium phosphate ((NH₄)₂HPO₄, analytical pure), ammonium nitrate (NH₄NO₃, analytical pure), calcium nitrate (Ca(NO₃)₂·4H₂O, analytical pure), absolute ethyl alcohol (C₂H₆O, analytical grade) were purchased from Sino-pharm (Shanghai, China) and used without further treatment.

2.2. Sample synthesis
The TiO₂/HAp composites were prepared by adding a certain amount of as-prepared TiO₂ powder into the precursor solution of HAp through a facile hydrothermal treatment. Firstly, 18 mL NH₃·H₂O solution (25%–28 wt%) was added dropwise into 10 mL TiCl₄ solution and simultaneously stirred for 2 hour. Next, the white precipitate was collected by centrifugal separation, and washed thoroughly with deionized water and absolute ethyl alcohol for several times, and then dried at 60 ºC overnight. The precursor powder was dispersed in 60 mL deionized water with vigorously stirring. Then it was put in 100 mL Teflon-lined autoclave and kept at 190 ºC for 8 h to obtain the TiO₂ powder. Secondly, (NH₄)₂HPO₄ and Ca(NO₃)₂ were dissolved in water with a molar ratio of Ca/P = 1.67. Then its pH value was adjusted by ammonia to 9–10. The prepared TiO₂ powder was added into the above solution and stirred for 2 h to get a suspension. Next, the suspension was transferred into a Teflon-lined autoclave and subsequently heated at 190 ºC for 8 h. The obtained products were centrifuged and washed with deionized water and absolute ethyl alcohol, and then dried at 60 ºC overnight.

To get TiO₂/HAp composites with different mass ratios, the amount of TiO₂ powder was varied. The samples with mass ratios of TiO₂ : HAp = 1 : 1, 3 : 1, 5 : 1 were labeled as 50%, 75% and 83% TiO₂/HAp (calculating by the ratio of TiO₂), respectively. For comparison, the pure HAp was prepared under the same procedure without adding TiO₂ powder.

2.3. Characterization
The XRD patterns were collected on a Philips X’pert PRO SUPER diffractometer using Cu Kα radiation (λ = 1.5406 Å) at a scan rate of 0.05°/2θ per s. FTIR (Magna-IR 50, Nicolet, America) spectra was obtained using a Fourier transform infrared spectroscopy absorption spectrometer from 4000 to 400 cm⁻¹ at room temperature. X-ray photoelectron spectroscopy (XPS) were collected on ThermoFisher Scientific photoelectron spectrometer (Escalab 250Xi). The morphologies of as-prepared samples were analyzed by transmission electron microscopy (TEM, JEM-2100, Japan), and high-resolution transmission electron microscopy (HRTEM) microscopy. The specific surface areas of the as-synthesized samples were determined by N₂ sorption using a Micromeritics Gemini 2390 analyzer and the pore size distribution was calculated using the Barrett–Joyner–Halenda (BJH) model. UV-vis diffuse reflectance spectra (DRS) was recorded by a UV-vis spectrophotometer using BaSO₄ as the reference standard. Temperature programmed desorption (TPD) was carried out on a chemisorption analyzer (Bjbuilder, PCA 1200, China). The sample for electron spin resonance spectroscopy (ER200-SRC, BRUKER, Germany) was prepared by mixing 0.05 prepared samples in a 25 mM DMPO solution with 50 mL aqueous dispersion for DMPO·OH or 50 mL alcohol dispersion for DMPO·O₂⁻, respectively. The light irradiation source was a 300 W Xe arc lamp (PLS-SXE 300, Beijing).

2.4. Evaluation of photoelectrochemical properties
The photoelectrochemical properties of pure TiO₂, HAp and TiO₂/HAp composites were carried out using a Parstat 4000 electrochemical workstation (USA) in a conventional three-electrode cell. In this three-electrode cell system, Ag/AgCl electrode and a platinum plate were used as reference electrode and counter electrode, respectively. In order to fabricate the working electrode, 0.09 wt% ethyl cellulose and 0.73 wt% terpineol were dissolved in ethanol, then 0.18 wt% sample was added in the above solution to obtain homogeneous suspension by bath sonication. Then the samples were coated on thefluorine doped tin oxide (FTO) glass through dip coating method. In order to remove the additives, the prepared electrodes were heat-treated in a furnace at 500 ºC for 30 min. The current–time curves were measured at 0.2 V vs. Ag/AgCl in 0.1 mol L⁻¹ Na₂SO₃ solution under a 300 W Xe arc lamp irradiation. Electrochemical impedance spectroscopy (EIS) was analyzed at a frequency range of 0.1 Hz to 100 kHz with a 5 mV voltage amplitude under
open-circuit voltage in 1 mmol L\(^{-1}\) K\(_2\)Fe(CN)\(_6\) and K\(_4\)Fe(CN)\(_6\) solution.

2.5. Evaluation of photocatalytic activity

The photocatalytic activity of the as-synthesized samples was investigated through the removal of NO at ppb level in a continuous flow reactor (30 cm \(\times\) 15 cm \(\times\) 10 cm) under simulated solar light irradiation at room temperature. The dispersed photocatalyst (0.1 g) was coated on a glass dish (\(D = 9.0\) cm) as shown in Fig. S1†. Then the dishes were treated at 70 °C until complete removal of water in the suspension. The simulated solar light source was a commercial 300 W Xe arc lamp (PLS-SXE 300, Beijing), which was vertical placed 20 cm above the reactor. The light intensity at the surface of the as-synthesized samples was controlled at 0.75 W cm\(^{-2}\) as measured by a photometer (THORLABS PD130, USA). The NO gas at ppb level was obtained from a compressed gas cylinder at air generator. The NO concentration of 50 ppm of NO (N\(_2\) balance). The 50 ppm of NO gas at ppb level was obtained from a compressed gas cylinder at 25 °C. As shown in Fig. 1a, the NO concentration was measured by a photometer (THORLABS PD130, USA) at a sampling rate of 0.6 L min\(^{-1}\). The NO concentration was calculated according to the following equation:

\[
R (\%) = (1 - C/C_0) \times 100\%
\]

where \(C\) is the outlet concentration of NO each access gases, and \(C_0\) is the inlet concentration before irradiation.

3. Results and discussion

3.1. Phase structure and chemical composition

Fig. 1a show the XRD patterns of TiO\(_2\), HAp and TiO\(_2\)/HAp composites. For the pure TiO\(_2\), the diffraction peaks at 25.2°, 37.8°, 48.0°, 53.9°, 55.2°, 62.7°, 68.5°, 70.3° are consistent with the characteristic peaks of anatase-TiO\(_2\) (JCPDS, 04-0477).\(^{25}\) As for pure HAp, all diffraction peaks and their relative intensities are agreement with the standard power diffraction patterns of hexagonal phase HAp (JCPDS, 09-0432). Fig. 1b shows the crystal structure model of HAp. It is clearly seen that the hydroxyapatite surface is rich in OH\(^-\) groups, PO\(_4\)\(^{3-}\) groups. No other impurity peak is observed, which indicates that pure TiO\(_2\) and HAp are successfully synthesized. Moreover, the diffraction patterns of the other three samples (50%, 75% and 83% TiO\(_2\)/HAp composites) show that both anatase-TiO\(_2\) and hexagonal phase HAp are existed. It can be concluded that the TiO\(_2\)/HAp composites were successfully synthesized by the hydrothermal method. Obviously, the intensities of TiO\(_2\) diffraction peaks have corresponding positive relationship with its relative content in the composites (50%, 75% and 83% TiO\(_2\)/HAp).

The FTIR spectra of pure TiO\(_2\), HAp and TiO\(_2\)/HAp composites are shown in Fig. 1c. The peaks at 3429 and 1626 cm\(^{-1}\) are caused by the lattice water of samples.\(^{26}\) The typical broad absorption bands of TiO\(_2\) are observed at 800–450 cm\(^{-1}\), which is attributed to the stretching vibration of Ti–O–Ti bonds.\(^{27}\) For HAp sample, the characteristic bands of PO\(_4\)\(^{3-}\) present at 472 cm\(^{-1}\) (\(v_3\)), 562 and 604 cm\(^{-1}\) (\(v_4\)), 966 cm\(^{-1}\) (\(v_1\)) and 1031 and 1095 cm\(^{-1}\) (\(v_2\)).\(^{28}\) The peaks appearing at 3573 and 633 cm\(^{-1}\) are attributed to stretching and bending vibration modes of OH bands of HAp, respectively.\(^{29}\) For the TiO\(_2\)/HAp composites, well- preserved characteristic peaks of TiO\(_2\) and HAp in the spectra of TiO\(_2\)/HAp (50%, 75% and 83%) composites, indicating no change of structure or phase transformation of TiO\(_2\) and HAp after composites synthesis. This observation demonstrates that the TiO\(_2\)/HAp composites are synthesized successfully, which is consistent with the XRD results.

XPS was used to further determine surface chemical composition and chemical states of pure TiO\(_2\), HAp and TiO\(_2\)/HAp composites. Fig. 2 shows the XPS spectra of TiO\(_2\), HAp and 75% TiO\(_2\)/HAp. In Fig. 2a Ti, O, Ca and P elements are observed in the survey spectra of TiO\(_2\)/HAp. The additional C element is present at 283.9 eV which is attributed to the adventitious carbon as shown in Fig. 2b, the Ti 2p peaks of TiO\(_2\) located at 459.2 eV and 464.7 eV are attributed to the Ti–O bonds.\(^{30}\) However, the Ti 2p peaks of 75% TiO\(_2\)/HAp move to 458.7 eV and 464.3 eV respectively, the lower energy levels, which could result from the existence of the Ti–O–Ca bonds in TiO\(_2\)/HAp. Similarly, Duo
et al. reported that the shift of the Ti 2p peaks in the TiO2/BiOCl composite was ascribed to the formation of Ti–O–Bi bonds. In Fig. 2c the Ca 2p peaks of HAp located at 346.6 eV and 350.1 eV are observed, while the Ca 2p peaks of 75% TiO2/HAp shift to 346.9 eV and 350.4 eV. Fig. 2d shows the corresponding high-resolution O 1s spectra of pure TiO2, HAp and TiO2/HAp. The two peaks at 529.7 eV and 530.9 eV in TiO2 are assigned to the Ti-O bonds and -OH groups, respectively. The two peaks at 530.8 eV and 532.4 eV in HAp are attributed to phosphate group (PO43-) and adsorbed water, respectively. The three peaks at 529.5 eV, 530.6 eV and 532.2 eV in 75% TiO2/HAp are assigned to lattice oxygen species TiO4 (Ti–O bonds), phosphate group (PO43-) and chemically adsorbed oxygen species of water, respectively.

3.2. Morphological analysis

The microcomposition and crystal structure of TiO2, HAp and TiO2/HAp were further analysed by TEM microscopy. Fig. 3a shows that the as-prepared TiO2 contains of intersection nanoparticles with an irregular shape. The size range of TiO2 is from 11 nm to 39 nm. In Fig. 3b the HAp power is composed of rod-like nanocrystals with the length of 32.17–127.23 nm and the width of 14.96–40.49 nm. Fig. 3c shows the TEM image of 75% TiO2/HAp composite. It is clearly seen that 75% TiO2/HAp possesses both irregular TiO2 nanoparticles and rod-like HAp. In particular, the size of HAp in 75% TiO2/HAp is the length of 49.32 nm and width of 19.51 nm. Fig. 3d shows the HRTEM image of the 75% TiO2/HAp. The two peaks at 0.344 and 0.351 nm is ascribed to the (101) crystal plane of anatase-TiO2 nanoparticles (JCPDS, 04-0477) and the (002) crystal plane of rod-like hexagonal phase HAp (JCPDS, 09-0432), respectively. The distinct intersections of lattice fringes are observed in Fig. 3d, indicating an intimate contact is formed between TiO2 and HAp.

3.3. Photocatalytic activity and stability

The photocatalytic degradation NO in air of the as-prepared samples is investigated. Fig. 4a shows the NO concentration vs. irradiation time curves of TiO2/HAp composites, mechanically mixed TiO2/HAp (mass ratio of TiO2 and HAp = 3:1), the same ratio to 75% TiO2/HAp composite), pure TiO2 and HAp under solar light irradiation. All TiO2/HAp (50%, 75% and 83%) composites show better NO removal performance as compared to pure TiO2 and HAp. The NO removal ratio of 75% TiO2/HAp is the highest of 44.61% among the six samples. In order to make a clearly quantitative comparison, we used the Langmuir–Hinshelwood model in the initial time period to evaluate the rates of NO photodegradation. According to previous study, the linear plot between ln(C0/C) and irradiation time (t) indicates that photocatalytic NO degradation over the as-prepared samples follows first-order order kinetics. The rate constant (kNO) for TiO2, HAp, and TiO2/HAp composites are listed in Fig. 4b. The kNO (0.1462 min-1) of 75% TiO2/HAp is the highest among the samples. The kNO follows the order of 75% TiO2/HAp > 83% TiO2/HAp > 50% TiO2/HAp > TiO2 > HAp, which is different from the BET result, and their specific surface area follows the order of 83% TiO2/HAp > 75% TiO2/HAp > 50% TiO2/HAp > HAp. It indicates that the specific surface area is not the only reason accounting for the improved photocatalytic activity.

Table 1 Summary of surface area, and pore diameter of the pure and composite materials

| Sample       | TiO2 | TiO2/HAp | TiO2/HAp | TiO2/HAp |
|--------------|------|----------|----------|----------|
| S_BET (m² g⁻¹) | 73   | 98       | 80       | 57       | 54       |
| Pore diameter (nm) | 12.9 | 11.2     | 14.9     | 30.5     | 39.9     |
Therefore, 75% TiO2/HAp composite has better photocatalytic generation of NO2. The above results demonstrate that 75% process. Fig. 4c shows that the amount of NO2 produced by TiO2 stored simultaneously in the NO photocatalytic oxidation same condition, which is lower than that of 75% TiO2/HAp. through the mechanically mixed method is 35.36% under the sample under simulated solar light irradiation for the removal of NO.

In addition, the toxic intermediate NO2 generation is monitored simultaneously in the NO photocatalytic oxidation process. Fig. 4e shows that the amount of NO2 produced by TiO2 and TiO2/HAp (50%, 75% and 83%) composites was about 12.1–15.8 ppb, much lower than that of 87.5 ppb produced by HAp. It indicates that the addition of TiO2 is beneficial to inhibit the generation of NO2. The above results demonstrate that 75% TiO2/HAp shows the best photocatalytic activity.

The photocatalytic stability of 75% TiO2/HAp composite was evaluated. In Fig. 4d, the removal ratio of NO in the five cycle experiments was not significantly decreased, the maximum of 44.61% and the minimum of 41.02%. It suggests that the TiO2/ HAp composites possess favorable stability and durability property.

3.4. Mechanism of photocatalytic activity improvement

Generally, the improved photocatalytic activity of catalysts is related to their optical absorption ability, target pollutant-catalyst adsorption/desorption ability, and the charge carriers’ mobility. Fig. 5a shows the UV-vis diffuse reflectance spectroscopy to investigate the optical properties of the samples. The optical absorption edges and band gap energies ($E_g$) for the as-prepared samples are shown in Table S1.† It clear seen that the absorption edge of pure TiO2 is located at 413 nm, corresponding to band gap ($E_g$) of 3.01 eV, while the absorption edge of pure HAp is less than 250 nm. Compared with pure HAp, the absorption edge of the TiO2/HAp composites shows a significant red shift. Moreover, the absorption edge of the TiO2/HAp composites shows a slight blue shift compare with TiO2, which indicates that the TiO2/HAp composites possess higher oxidation capacity than pure TiO2.

In order to investigate the adsorption/desorption process of NO on the synthesized photocatalysts, temperature programmed desorption (TPD) tests were carried out. Fig. 5b shows that in the TPD curves of pure TiO2, there appear two desorption peaks located at around 249 °C and 655 °C, which were assigned to the desorption of NO and NO2–O2, respectively. While the 75% TiO2/HAp composite have three main peaks located at around 276 °C, 474 °C and 702 °C. Among them, the stronger desorption peak of NO at 474 °C was assigned to the stronger chemisorption bond between NO and surface OH groups of HAp.21

Photocurrent response was employed to investigate the separation efficiency of photo-generated electrons and holes. Fig. 6a shows the photocurrent–time curves of TiO2, HAp and TiO2/HAp composites under the simulated solar light irradiation. There is no photocurrent over pure HAp.24 The photocurrent density of 75% TiO2/HAp composites is 1.2 times as high as the pure TiO2, which indicates that more efficient separation of the photo-generated electron–hole pairs over 75% TiO2/HAp. According to the TEM and XPS results, the transfer and separation of the photo-generated carriers relate to the intimate interfacial contact between TiO2 and HAp, which improves photocatalytic activity of the TiO2/HAp.

Moreover, in order to explain the interfacial charge transfer and charge carriers recombination, electrochemical impedance
TiO2/HAp composite, ESR was investigated to determine the photo-generated carriers. It was found that the pure TiO2 and HAp, which indicates that 75% TiO2/HAp improve electron migration and promote the separation of photo-generated carriers.

In order to further reveal the photocatalytic mechanism of TiO2/HAp composite, ESR was investigated to determine the active reaction groups. As shown in Fig. 7, the characteristic peaks of the DMPO-’O2’− adducts (originated from ‘O2’− radicals) are observed from HAp, while signals of DMPO-’OH (four characteristic peaks of DMPO-’OH with an intensity ratio of 1 : 2 : 2 : 1 originated from ‘OH radicals) adducts are weak, which is conformity with previous study. For pure TiO2, the appearance of DMPO-’OH and the absence of DMPO-’O2’− signal indicate that NO is removed through oxidation by ‘OH radicals via pure TiO2. The band positions of TiO2 can be theoretically calculated by the following equations: $E_{CB} = \chi - E^e + 0.5 E_g$, $E_{VB} = E_{CB} - E_g$, where $E_{CB}$ is the valence band edge potential, $E_{CB}$ is conduction band edge potential, $\chi$ represents the electronegativity of the semiconductor, which is the geometric mean of the electronegativity of the constituent atoms, where the $\chi$ of TiO2 calculated to be 5.81 eV. $E^e$ is the energy of free electrons on the hydrogen scale (~4.5 eV). $E_g$ is the band gap energy of the semiconductor, which is calculated from the results of UV-vis diffuse reflectance spectroscopy. The VB and CB potentials of pure TiO2 are calculated to be 2.81 eV and ~0.20 eV, respectively. Because the CB potential of TiO2 is more positive than the redox potential of $\text{O}_2’/\text{O}_2$ (−0.33 eV), the ‘O2’− radicals could not be formed, which is consistent with the ESR results. The VB of TiO2 is more positive than that of OH/’OH (1.99 eV), thus the photo-generated hole can oxidize OH into ‘OH radicals. The signal intensity of ‘OH radicals in 75% TiO2/HAp is significant stronger than that in pure TiO2 and HAp, which demonstrates that the combination of TiO2 and HAp can promote the generation of ‘OH radicals. Additionally, 75% TiO2/HAp have similar signal intensity of ‘O2’− radicals to HAp, while it is significantly stronger than that in pure TiO2. It may be because the electron state change of the PO43− groups in the surface of HAp in photocatalytic process under UV irradiation can result in the generation of ‘O2’− radicals.

4. Conclusion

Based on the above analysis, the photocatalytic mechanism of TiO2/HAp composite under simulated solar light irradiation is showed in Fig. 8. TiO2 photocatalyst is excited by the simulated solar light irradiation to generate photoinduced electrons which are then transferred to the oxygen defect level from the electron state change of the PO43− groups in the surface of HAp. Then the electron captured by HAp reacts with the surrounding oxygen on the vacancies to form ‘O2’− radicals. Next, a fraction of ‘O2’− reacts with H2O and H+ to form H2O2, and then the electrons combines with H2O2 to produce ‘OH radicals. This transfer behavior of the electrons from TiO2 transfer to the vacancy on HAp enhances the separation and lengths the lifetime of carriers, and inhibits the recombination of photo-generated electrons−holes pairs, and result in the generation of more ‘OH radicals. Simultaneously, the photoinduced hole on the VB of TiO2 reacts with H2O or OH− to form ‘OH radicals. At last, the ‘O2’− and ‘OH radicals can further react with NO or NO3− to produce NO3−.

In summary, TiO2/HAp composite has been successfully synthesized by a facile hydrothermal method. The NO removal ratio of 75% TiO2/HAp is 44.61%, which is the highest among the pure TiO2, HAp and TiO2/HAp composites. The results of TPD analysis, photocurrent generation and EIS measurements demonstrated that the enhanced photocatalytic activity of TiO2/HAp composites can be ascribed to stronger chemisorption for NO, higher separation efficiency and faster transfer of the photo-generated electron−holes pairs. The ESR results revealed that ‘O2’− and ‘OH radicals as the major active species in the photocatalytic reaction process. The photocatalytic removal mechanism of NO over the TiO2/HAp composites is illustrated in detail. This study suggests that 75% TiO2/HAp composites with the highest photocatalytic activity is considered to be a promising air-cleaning material.
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