KINETICS OF PASSIVATION AND PITTING CORROSION OF POLYCRYSTALLINE COPPER IN BORATE BUFFER SOLUTIONS CONTAINING SODIUM CHLORIDE

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Abstract—The pitting corrosion of copper in borate buffer containing sodium chloride is studied by using potentiostatic and potentiodynamic techniques complemented with scanning electron microscopy and EDAX. The breakdown potential shifts towards more negative values as the sodium chloride concentration increases. During pitting both soluble Cu(I) and Cu(II) species are detected. The first stage of pitting is explained through the competition between the passive layer formation and the nucleation and growth of the CuCl layer in equilibrium with Cu(I)-chloride complexes in solution. When salt nuclei reach the metal surface, pit growth under charge-transfer control is found. In the following stage the kinetics of pit growth changes to a diffusion controlled process when the thick CuCl layer is completed. Secondary breakdown of the salt layer results in copper dissolution through Cu(II) soluble species. The corresponding overall process is discussed in terms of a sum of nucleation and growth processes. The reaction model reproduces the potentiostatic current transients of copper in weakly alkaline borate buffer containing sodium chloride.

INTRODUCTION

The corrosion and passivation of copper have been studied over a wide range of experimental conditions[1]. In the absence of aggressive anions, the passivation of copper in alkaline solution is due to an oxide layer whose electrochemical characteristics and structure depends considerably on the solution composition and applied potential[2]. Potentiodynamic polarization curves exhibit various anodic peaks assigned to the formation of Cu(I), Cu(II) and Cu(III) species and a multiplicity of cathodic peaks related to the electroreduction of different soluble and insoluble copper-containing species[3]. The structure of the passive layer as revealed by ESCA and ISS in borate buffers consists of an inner CuO and outer CuO hydrated layers[4]. During the anodization of copper in neutral and alkaline solutions both soluble Cu(I) and Cu(II) species have been detected through the rrd techniques[3, 5].

The electrochemical behaviour of copper in solutions containing aggressive anions has been also extensively investigated in acid and weakly acid solutions[6–8]. Anodic potentiostatic transients for the electrodissolution of shielded horizontal Cu electrodes in aqueous HCl were quantitatively compared to Mueller's theory of passivation[9–11]. The anodic transient exhibits a first current plateau which is related to the following reaction:

\[ \text{Cl}^- + \text{Cu} \rightarrow \text{CuCl} + e^- \quad (1) \]

where the anodic layer of CuCl passivates the electrode. The process is convection free for rapid transients. The anodic transients present also a second plateau which corresponds to a complex ion formation as CuCl is dissolved as CuCl$_2^-$ and CuCl$_3^-$. In this case, the electrodissolution process is controlled by the diffusion of the complex ions out of the electrode surface. This process is followed by CuCl precipitation to form a thick layer on the electrode.

Secondary breakdown of the CuCl layer results in pitting corrosion of copper[12]. At pH values close to 9.00 in the presence of Cl$^-$ ions, the passive oxide film is formed on the copper surface previous to its breakdown and pit growth[13]. Breakdown potentials, corrosion products inside the pits and pH effects on the pitting process are also reported[13, 14].

A general survey of the literature on the corrosion and passivation of copper in aqueous solutions indicates that the information on the processes leading to passivity breakdown and pit growth are considerably more scarce than that available for other metals such as iron, nickel and ferrous alloys.

The present paper deals with the competition between the onset of passivity promoted by water and OH$^-$ ions and the formation of a CuCl layer and its influence on the pitting corrosion of copper in weakly alkaline borate buffer (pH 9.00) in the presence of NaCl. The pitting process is studied by using conventional electrochemical techniques, scanning electron microscopy and EDAX. The kinetics of pit initiation and growth are discussed through the nucleation and growth theory as applied to the interpretation of the pitting corrosion of iron and nickel in aqueous solutions[15, 16].

EXPERIMENTAL

The working electrodes were made of polycrystalline 99.9% copper rods axially embedded in Araldite holders to obtain exposed circular areas of 0.072 cm$^2$ ready to be mounted in a rotating disc electrode device. The pretreatment of each specimen consisted of a
gradual polishing starting with fine-grained emery paper followed with diamond paste and finally with alumina paste (1 μm diam.). The polished specimen was successively rinsed with acetone A.R., twice-distilled water, and finally dried in air at room temperature. A fresh specimen was used for each run. The potential of the specimen was measured against a saturated calomel electrode (sce) provided with a Luggin-Haber capillary tip. A platinum plate was used as counterelectrode. The three electrodes were mounted in a conventional Pyrex glass cell. Measurements were made at 25 ± 0.1°C in the borate-boric acid buffers at pH 9.00 [0.075 Na₂B₄O₇ + 0.15 M H₃BO₃ + x cₖB₃O₅ (0 ≤ x ≤ 1 M)]. Solutions were prepared from twice-distilled water and A.R. chemicals and purged with purified nitrogen for 3 h previous to each run.

Singular triangular potential sweep (stpS) voltammograms between the cathodic (E,,,) and the anodic (E, J switching potential at the sweep rates (u) 0.02, 0.002 and 0.00016 V s⁻¹ were recorded in the apparent current density (i) us potential (E) scale. The ring-disc technique (rrde) was used to detect soluble copper ionic species formed during electrooxidation of copper to either Cu(I) or Cu(II). For this purpose a gold ring (0.035 cm² apparent area) copper disc (0.125 cm² apparent area) rotating electrode was used (collection efficiency 0.25) at rotation speeds (ω) in the 250 < ω < 3000 rpm range. Current transients (I/t) at constant potential (E,) were also recorded. In this case, preceding the application of the potential step at E,, the working electrode was held firstly at -1.2 V (E,) for 90 s and later at -0.10 V (E,) for 90 s to form a thin Cu₂O layer in the order of ca 10 Å thickness[4]. SEM observations and EDAX data were also obtained.

RESULTS

1. Potentiodynamic polarization curves in the borax buffer

Potentiodynamic curves of polycrystalline copper in borax buffer (pH 9.00) recorded at 0.002 V s⁻¹ (Fig. 1) show between E,, = -1.00 V and E, = 0.80 V two anodic peaks at -0.10 V (peak A,) and 0.00 V (peak A,) and at E, > E, a passive region containing a broad hump at 0.40 V (peak A,). After cycling at 0.02 V s⁻¹ until the voltammogram remains unchanged, the stabilized E/j profile shows only a single peak (peak A) at 0.03 V instead of peaks A and A, and peak A at 0.60 V. The potential scan in the negative direction presents two cathodic peaks at -0.30 V (peak C,) and -0.44 V (peak C). Inspection of the voltammogram resulting when E, is stepwise changed from -0.20 to 0.80 V shows that the pairs of peaks A, and C, and A, (A,) and C, are complementary peaks [Fig. 2(a)]. Peak A, is assigned to the electroformation of Cu₂O while peak A, is related to the electroformation of either CuO or Cu(OH)₂. From ESCA and ISS one concludes[4] that the film produced at E, > E, consists of an inner Cu₂O layer and an outer CuO or Cu(OH)₂ layer. This complex structured layer is electroreduced in two consecutive steps to Cu(II) and to copper at the potential ranges of current peaks C, and C, respectively. From the rtps run at 0.10 V s⁻¹ between E,, = -1.00 V and E, = -0.20 V it is concluded that the current contribution at the negative side of peak A involves a small peak (peak A) at -0.36 V whose conjugated cathodic contribution appears at -0.65 V [Fig. 2(b)]. The charge density related to peaks A, and C, which is estimated as 0.3 mC cm⁻², can be assigned to the electroformation and electroreduction of a Cu(OH)₂ monolayer[17] at potentials slightly more negative than the equilibrium potential of the Cu/Cu₂O electrode (Table I).
Table 1. Standard equilibrium potentials (vs sce) of redox processes related to copper in base solutions containing chloride ions \[4,29\]

| Reaction | \( E^*/V \) |
|----------|-------------|
| \(Cu + Cl^- \rightarrow CuCl + e^-\) | -0.105 |
| \(Cu + 2Cl^- \rightarrow CuCl_2 + e^-\) | -0.0324 |
| \(Cu + 3Cl^- \rightarrow CuCl_2 + e^-\) | -0.0442 |
| \(2Cu + H_2O \rightarrow CuO + 2H^+ + 2e^-\) | 0.217 |
| \(CuO + H_2O \rightarrow 2Cu(OH)_2 + 2H^+ + 2e^-\) | 0.487 |
| \(Cu_2O + H_2O = 2CuO + 2H^+ + e^-\) | 0.427 |

2. Potentiodynamic polarization curves in the borax buffer containing NaCl

The addition of NaCl to the borax buffer increases the heights of peaks \(A_1\) and \(A_2\) proportionally to \(c_{\text{NaCl}}\) (0.1 \(\leq c_{\text{NaCl}} \leq 1\) M). For \(c_{\text{NaCl}} \geq 0.5\) M a new current peak (peak \(A_3\)) is observed at \(ca\ 0.00\) V (Fig. 3). At a constant \(c_{\text{NaCl}}\) the heights of peaks \(A_1\), \(A_2\) and \(A_3\) increase according to \(\omega\), although the greatest influence of \(\omega\) appears on \(A_3\). The passive current also increases as \(c_{\text{NaCl}}\) increases. The \(E/J\) profiles exhibit an abrupt increase in the anodic current when a critical potential value (\(E_b\)) is exceeded. The rapid current increase is preceded by current oscillations probably related to the activation–passivation processes occurring at the Cu electrode. The current increase at potentials greater then \(E_b\) is associated with film breakdown and pitting corrosion of copper.

Simultaneously, white–orange spots appear on the electrode surface. The increase in the NaCl concentration shifts \(E_b\) towards more negative values according to the following equation (Fig. 4):

\[ E_b = a - b' \log c_{\text{NaCl}} \]

where \(a\) and \(b'\) are constants whose values depend on \(v\) (Fig. 4). Otherwise as \(v\) decreases \(E_b\) becomes more negative, and as \(\omega\) increases from 0 to 3000 rpm, a shift in \(E_b\) of about 30 mV in the positive potential direction is noticed (Fig. 3). Voltammograms recorded at 0.002 V s\(^{-1}\) between \(E_{\infty} = -0.10\) V and \(E_{\infty} > E_b\) after 90 s at \(E_{\infty}\), to form a Cu–O prepassive layer (Fig. 5), show the typical behaviour of a nucleation and growth process. During potential scan in the negative direction the current is higher than that observed in the reverse scan. At potentials close to 0.0 V pit repassivation is observed.

Voltammograms at 0.02 V s\(^{-1}\) in the borax buffer + 0.5 M NaCl (Fig. 6) run between \(E_{\infty} = -1.00\) V and \(E_{\infty}\) values gradually changed from -0.25 to 0.40 V, exhibit practically no change in peaks \(A_1\) and \(C_1\), but the reaction corresponding to peak \(A_3\) appears at potentials closer to the equilibrium potential of the \(Cu/CuCl\) \((c_{\text{NaCl}} = 0.5\) M) (Table 1). However, the corresponding cathodic contribution is found at potentials more negative than those reported in acid LiCl 1 M[18]. In these cases the charges of peaks \(C_3\) and \(C_4\) overlap to a great extent. This fact suggests that the CuCl electroreduction becomes more difficult when the inner Cu–O layer is already formed.
In the presence of NaCl peaks $A_1$ and $C_1$ appear at potentials more positive than in free NaCl solution. In this case during the negative-direction potential scan when the electroreduction of Cu(OH)$_2$ occurs, an anodic current is detected which suggests that in the NaCl-containing solutions the passive state is mainly due to the electroformation of the Cu(OH)$_2$ layer. Data obtained with the rde technique by applying to the disc a potential sweep of 0.002 V s$^{-1}$ and keeping the ring potential at $-0.80$ V to electroreduce any Cu(I) and Cu(II) species, is consistent with the presence of soluble copper species at potentials preceding peaks $A_1$ and $A_{III}$ as well as in the pitting region [Fig. 3(d)]. Similar experiments were made by holding the ring at 0.60 V to detect only soluble Cu(I) species [Fig. 3(e)]. The latter is detected at potentials preceding those of peaks $A_1$ and $A_{III}$ and also during pitting. However, in the pitting region, the amount of soluble Cu(II) is greater than that of soluble Cu(I) species. The latter in the potential range of peak $A_1$ increases according to the NaCl concentration and to $\omega$. This fact indicates that Cu(I) in solution undergoes a reaction with Cl$^-$ ions yielding CuCl$_2$ in equilibrium with insoluble CuCl(s) (Table 2).

### Table 2. Amounts of soluble copper(I) species detected at current peak $A_1 - A_{III}$

| $c_{NaCl}$/M | $q$/μC cm$^{-2}$ | $\omega$/rpm |
|-------------|-----------------|-------------|
| 0           | 24.5            | 250         |
| 0.1         | 282.0           | 250         |
| 0.5         | 11222.0         | 250         |
| 0.5         | 19815.0         | 3000        |

3. Current transients in the borax buffer

Current transients in the borax buffer at different potential steps ($E_\alpha$) were run after following a two-step procedure, namely the electrode was firstly held at $-1.20$ V for 90 s to electroreduce the copper surface, and then at $-0.10$ V for another 90 s to form a thin prepassive Cu$_2$O layer of approx. 10 Å thickness [4]. Immediately afterwards the electrode was stepped to $E_\alpha$ and the corresponding current transient was recorded [Fig. 7(a)]. In the $0.20 \leq E_\alpha \leq 0.80$ V range the current decreases monotonously to reach the corresponding stationary passivity current value and this behaviour is independent of $\omega$. The charge resulting from the current transient integration for $E_\alpha > E_{AII}$ is associated with two main processes, namely the formation and growth of the passive layer ($j_\phi$) and the corrosion of copper through the passive layer ($j_\psi$). The latter implies the formation of soluble copper species as detected through the use of the rde. Practically all soluble species resulting during the potential step from the prepassive ($E = -0.10$ V) to the passive region correspond to soluble Cu(II) probably as Cu(OH)$_2$ (I) species[8] (Fig. 8).

![Fig. 7. (a) Current transients in borax buffer solution at different $E_\alpha$ values; (b) current transient recorded in the borax buffer + 0.5 M NaCl solution at $E_\alpha < E_5$ ($E_5 = 0.20$ V); 25°C.](image-url)
Kinetics of passivation and pitting corrosion of polycrystalline copper

4. Current transients in the borax buffer containing NaCl

Current transients recorded in borax buffer + 0.5 M NaCl at $E_a < E_b$ exhibit characteristics rather similar to those described for plain borax solution, although the instantaneous current becomes greater than that resulting in the absence of NaCl ([Fig. 7(b)]. This fact is accompanied by a greater production of soluble Cu(II).

In the $0.20 \leq E_a \leq E_b$ potential range only minor amounts of Cu(I) soluble species are caught at the ring electrode. On the other hand, the current transient recorded for $E_a > E_b$ ([Fig. 9(a)] initially decreases to a minimum ($I_m$) and after certain time ($t_1$) it rises to attain a maximum ($I_{max}$) at the time $t_{max}$. Finally, it decreases again but slowly. The increase in $E_a$ results in a decrease of $I_1$, an increase of $I_{max}$ and a decrease of $t_{max}$ ([Figs 10(a)-(c)]). At a constant $E_a$, the same effects are got by increasing $c_{NaCl}$ ([Fig. 11]). The apparent pitting current density ($j_1 - j_0$) vs time ($t$) plot fits the Engell-Stolica equation (Fig. 12):

$$j_1 - j_0 = k(t - t_1)^b$$

where $j_1$ and $j_0$ are the overall and the background apparent current densities, respectively; $k$ and $b$ are constants. For $t \to 0$ the values of $b$ are in the 1.3-2.0 range and when $t$ increases beyond a certain transition time ($t_1$) which depends on the applied potential, these values fall in the 0.5-1.0 range. When the potential step applied to the disc is greater than $E_a (E_a > E_b)$ and the potential at the ring is kept at $-0.80$ V, then the current transient at the ring due to the formation of soluble copper species at the disc electrode, follows that at the disc electrode ([Fig. 9(b)]. Notwithstanding, when the potential at the ring is 0.60 V, the corresponding current first decreases to a minimum and later increases to a maximum before the current at the disc reaches a maximum. Finally, the quasi-steady current is attained at the ring earlier than at the disc ([Fig. 9(c)].

5. SEM patterns and EDAX data

SEM observations of copper electrodes held at 0.5 V for 50 s in borax buffer + 0.3 M NaCl show an outer Cu(OH)$_2$ layer which spills away from the electrode to the cell. After removing this non-adherent layer an inner layer showing a white-orange colour is observed. SEM patterns reveal the presence of a thick layer made of cubic crystals. The EDAX signals of these layers result in copper dissolution mainly as Cu(II) soluble species, which further contributes to the formation of a CuCl$_2$ layer.

Fig. 8. (a) Current transient in the borax buffer solution at $E_a = 0.20$ V; (b) current density at the rotating ring electrode $E_a = -0.80$ V; (c) current density at the rotating ring electrode $E_a = 0.60$ V.

Fig. 9. (a) Current transient in borax buffer + 0.5 M NaCl solution at $E_a = 0.50$ V, $\omega = 250$ rpm; (b) current at rotating ring electrode $E_a = -0.80$ V; (c) current at the rotating ring electrode $E_a = 0.60$ V, 25°C.
Fig. 10. (a) Initial part of the current transients in the borax buffer + 0.5 M NaCl solution at different $E_1$ values; (b) detail of the initial part of current transients. ($\circ$, $+$, $\times$) experimental data; (- -) computed from Equations (17), (27) and (30); (c) current maximum of transients ($\circ$, $\times$, $\bullet$, $+$) experimental data; (- -) computed from Equation (33).

Fig. 11. Current transients recorded in borax buffer + x M NaCl solutions at $E_1 = 0.5$ V. ($\circ$, $\times$, $+$) experimental data; (- -) computed from Equations (17), (27) and (30); 25°C.

Fig. 12. Log $j$ vs log $(t-t_i)$ plots. Total current density at the rotating disc electrode ($\bullet$) $E_r = 0.50$ V, $\omega = 250$ rpm, borax buffer + 0.5 M NaCl solution; $j_1$ ($\circ$) $E_r = 0.60$ V; $j_1$ ($\bullet$) $E_r = 0.80$ V.
DISCUSSION

The anodic current–potential curve of polycrystalline copper in weak alkaline buffers containing NaCl can be divided into at least three potential regions. This division is useful for rationalizing the discussion of results. In the \( E_{\text{a}} < E < E_{\text{a}} \) potential range (region I) the formation of a hydrous anodic layer involving \( \text{Cu}_2\text{O} \), \( \text{Cu(OH)}_2 \) and a certain amount of \( \text{CuCl} \) takes place. This layer apparently involves an inner and outer layer complex structure[4]. At \( E > E_{\text{a}} \) a passive region extending up to \( E_p \) is observed (region II) and at \( E > E_p \) pitting corrosion by the aggressive chloride ions is accomplished (region III).

1. Interpretation of the electrochemical reactions

1.1. Analysis of region I. In the absence of NaCl the early stage of copper electrooxidation is interpreted as the electroformation of a \( \text{Cu(OH)}_2 \) monolayer on copper through the underpotential decomposition (UPD) of water yielding absorbed \( \text{OH}^- \) species[17]. This reaction occurring at \( E_c \) values lower than the equilibrium potential of the \( \text{Cu}/\text{Cu(I)} \) reversible electrode (Table 1) is followed as \( E_c \) becomes closer to \( E_{\text{a}} \) by the tridimensional growth of the hydrous \( \text{Cu}_2\text{O} \) layer. Both soluble \( \text{Cu(I)} \) and \( \text{Cu(II)} \) species are detected at this stage of copper anodization in alkaline electrolytes[5]. When \( E_c \) approaches \( E_{\text{a}} \) the electroformation of the hydrous \( \text{Cu(OH)}_2 \) layer takes place[4]. In this way, the complex anodic layer structure involving the inner \( \text{Cu}_2\text{O} \) and outer \( \text{Cu(OH)}_2 \) layers is built. In the presence of NaCl, the increase with \( c_{\text{NaCl}} \) of both the current density required for the tridimensional passive film growth and the amount of soluble \( \text{Cu(I)} \) suggests the prevalence of copper chlorocomplex electrodissolution over the \( \text{Cu}_2\text{O} \) electroformation. The formation of a certain amount of \( \text{CuCl} \) at \( E_{\text{a}} < E < E_{\text{a}} \) on the \( \text{Cu}_2\text{O} \) layer correlates with both the appearance of peak \( A_{\text{II}} \) and the thermodynamic data assembled in Table I. Then, the structure of the anodic layer can be represented by a poorly hydrous inner \( \text{Cu}_2\text{O} \) non-homogeneously covered by \( \text{Cu(OH)}_2 \) and \( \text{CuCl} \) precipitates. These results can be interpreted through a series of competitive adsorption reactions, involving \( \text{Cl}^- \) and \( \text{OH}^- \) ions at the copper surface, according to the following simple chemical formalism:

\[
\begin{align*}
\text{Cu} + \text{H}_2\text{O} & \rightarrow \text{Cu(OH)}_2 + \text{H}^+ + e^- \quad (6) \\
\text{Cu} + \text{OH}^- & \rightarrow \text{Cu(OH)}_2 + e^- \quad (7) \\
\text{Cu} + \text{Cl}^- & \rightarrow \text{CuCl}_2 + e^- \quad (8) \\
\text{CuCl}_2 + \text{OH}^- & \rightarrow \text{Cu(OH)}_2 + \text{Cl}^- \quad (9) \\
\text{Cu(OH)}_2 + \text{Cl}^- & \rightarrow \text{CuCl}_2 + \text{OH}^- \quad (10) \\
\text{Cu(OH)}_2 + \text{OH}^- & \rightarrow \text{Cu(OH)}_2^+ \quad (11) \\
\text{CuCl}_2 + \text{Cl}^- & \rightarrow \text{CuCl}_2^- \quad (12)
\end{align*}
\]

Reaction (6) is the UPD decomposition of water yielding adsorbed \( \text{OH}^- \). This reaction occurs as a reversible process at the monolayer level[17]. Reactions (7) and (8) correspond to anion adsorption processes. Reactions (9) and (10) indicate competitive adsorption stages and reactions (11) and (12) are related to copper electrodissolution through soluble \( \text{Cu(I)} \) species.

The second level of copper electrooxidation can follow, in principle, both pathways either through \( \text{Cu(OH)}_2 \) or through \( \text{CuCl}_2 \) species:

\[
\begin{align*}
\text{Cu(OH)}_2 & \rightarrow \text{oxide} \quad (i) \\
\text{CuCl}_2 & \rightarrow \text{CuCl} \quad (j)
\end{align*}
\]

The composition and structure of the tridimensional layer should change according to the relative contribution of reactions (13) and (14). The relative amount of adsorbed \( \text{OH}^- \) and \( \text{Cl}^- \) ions should depend on the local pH value which is mainly determined by reaction (6). Thus, when copper is in contact with concentrated NaCl acid solutions the anodic reaction yields \( \text{CuCl}_2^- \) or \( \text{CuCl}_2 \) (at \( c_{\text{NaCl}} < 0.7 \text{M} \) and \( c_{\text{NaCl}} > 0.7 \text{M} \), respectively) and a \( \text{CuCl} \) precipitate[18] as reaction products, although in the borax buffer (pH 9.00) \( \text{Cu}_2\text{O} \) electroformation is thermodynamically favoured. Therefore, in the latter case copper undergoes electrodissolution through a thin \( \text{Cu}_2\text{O} \) layer and \( \text{CuCl} \) deposition results on the \( \text{Cu}_2\text{O} \) film. In this case, the dependences of the current peaks on \( \omega \) suggest that \( \text{Cu}_2\text{O} \) and \( \text{CuCl} \) electroformations are controlled by diffusion[18].

1.2. Analysis of region II. In the borax buffer and in the borax buffer containing \( x \text{M NaCl} \) solutions, the passive current density \( j \) involves two main contributions, namely the current density associated with the passive layer \( j_p \) and that related to the copper electrodissolution through the passive layer \( j_c \)[19]. In this potential region under stationary conditions \( j_p \rightarrow 0 \) and \( j ightarrow j_c \). The contribution of \( j_c \) which is related to the formation of soluble \( \text{Cu(II)} \) increases with \( c_{\text{NaCl}} \), as expected from the fact that the formation of \( \text{CuCl}^- \) species is assisted by increasing \( c_{\text{NaCl}} \) in solution.

1.3. Analysis of region III. In NaCl-containing solutions with prepassivated copper electrodes, both the potentiostatic current transients and the voltamograms are typical of those expected for the nucleation and growth of a new phase on the electrode[20]. Both EDAX and SEM data indicate that \( \text{CuCl} \) appears at local points related to pitting corrosion of prepassivated copper. The increase of \( E_p \) with \( \omega \) can be regarded as the salt nuclei stabilization caused either by removal of chloro-complexes out of the reaction region or localized acidification[21].

The mechanism of pitting corrosion of prepassivated copper in NaCl solutions can be conceived as a sequence of stages involving the initial nucleation of \( \text{CuCl} \) at weak points and defects of the \( \text{Cu}_2\text{O} \) layer (stage 1) followed by salt nuclei penetration through the inner \( \text{Cu}_2\text{O} \) layer to reach the metal oxide interface (stage 2). The induction time \( t_i \) is related to stages 1 and 2[22]. The following stages are the direct attack of the base copper metal covered either by a discontinuous thin chloride phase (stage 3) or by the thick \( \text{CuCl} \) salt layer formed by precipitation (stage 4). The corresponding reactions imply the local formation of different ionic soluble complexes such as \( \text{CuCl}_2 \) and \( \text{CuCl}_2^- \) in equilibrium with \( \text{CuCl} \) (sat). The presence of soluble \( \text{Cu(II)} \) can be explained through the following reaction at the salt covered regions, involving \( \text{CuCl} \)[12].

\[
\text{CuCl} = \text{Cu}^{2+} + \text{Cl}^- + e^- \quad (15)
\]
2. A generalized model for the anodic behaviour of prepassivated copper

The kinetic characteristics of the various processes involved in the anodization of prepassivated copper in NaCl solutions can be estimated through the analysis of the potentiostatic current transients run in different potential ranges. Thus, the instantaneous average anodic current \( I_a \) is assigned to three main processes, namely the passive layer growth \( I_p \), the metal corrosion through the passive layer \( I_c \) and the random salt nuclei growth on the metal surface \( I_r \):

\[
I_a = I_p + I_c + I_r. \tag{16}
\]

Neglecting the contributions from the changes in the electrical double layer structure, in the absence of pitting, \( I_p \) and \( I_c \) become the most important contributions to \( I_a \). The growth of the passive layer on the prepassivated copper can be regarded as the formation of a new phase following an instantaneous nucleation and circular bidimensional growth mechanism under diffusion control. In this case, considering that the passive layer growth and the corrosion current are independent and separate processes occurring on the whole electrode surface, the rate equation in terms of the apparent current density is [23]:

\[
j_a = P_1 \exp (-P_2 t) \tag{17}
\]

where \( P_1 = \frac{q_{mon}}{zF}D_k N_0 \) and \( P_2 = K_c \pi N_0 D \); \( D \) is the diffusion coefficient of the reacting species; \( q_{mon} \) is the charge density required for the full coverage of the metal surface by the constant thickness passive layer; \( K_c \) is a proportionality constant and \( N_0 \) is the number of sites available for nucleation.

On the other hand, the rate of metal corrosion through the passive layer appears as the dissolution of tridimensional nuclei under diffusion control following a rate equation of the form [24]:

\[
j_c = \frac{P_3}{\sqrt{t}} \{1 - \exp (-P_4 t)\} \tag{18}
\]

where \( P_3 = zF \Delta c^{1/2} \pi / n^{1/2} \) and \( P_4 = K_c \pi N_0 D \). \( \Delta c \) is the concentration difference of the diffusing species and \( N_0 \) is the number of sites available for the dissolution process. In the absence of pitting Equations (17) and (18) with the set of parameters assembled in Table 3 reproduce satisfactorily the current transients [Figs 7(a)-(b)]. Obviously, the time scale associated with \( j_a \) and \( j_c \) are considerably different. The contribution of the latter principally manifests for \( t < 2s \), but as \( t \to \infty \), \( j_c \to j_r \). As no influence of solution stirring is observed, the second layer growth on the prepassivated copper specimen should involve a rate process in the solid phase such as the diffusion of either anion or cation or both through the CuO layer. In the range of \( 0.2 \leq E_a \leq 0.8 \) V the values of \( q_{mon} \) obtained from \( P_1/P_2 \) ratio are between 0.8 and 6 mC cm\(^{-2}\). Otherwise, the \( q_{mon} \) vs \( E_a \) plot coincides reasonably well with recently reported data resulting from potentiodynamic reduction of copper in borax solutions (Fig. 13) [14]. Moreover, the average thickness of the passive layer formed at \( E_a \) can be estimated from the equation:

\[
h = \frac{M}{zF} \frac{q_{mon}}{C_{CuO}} \tag{19}
\]

where \( M \) is the molecular weight and \( \rho \) is the density of the layer.

Table 3. Parameters used in simulation of potentiostatic current transients with Equations (17), (18) and with Equation (24)

| \( E_a / \text{V} \) | \( P_1 / \text{mA cm}^{-2} \) | \( P_2 / \text{min}^{-1} \) | \( P_3 / \text{mA min}^{1/2} \text{cm}^{-2} \) | \( P_4 / \text{min}^{-1} \) | \( E_a / \text{V} \) | \( P_3 / \text{mA min}^{1/2} \text{cm}^{-2} \) | \( P_4 / \text{min}^{-1} \) | \( P_5 / \text{mA cm}^{-2} \) |
|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|
| Borate boric acid buffer | 0.20 | 0.691 | 52.3 | 0.0855 | 0.793 | -0.80 | 0.125 | 1.16 | 0.0261 |
| | 0.40 | 1.32 | 51.0 | 0.118 | 1.55 | | | | |
| | 0.60 | 3.95 | 55.7 | 0.347 | 4.26 | | | | |
| | 0.80 | 4.98 | 49.8 | 0.398 | 6.00 | | | | |
| Borate-boric acid buffer + 0.5 M NaCl | 0.20 | 6.47 | 43.7 | 1.07 | 61.5 | -0.80 | 0.249 | 1.94 | 0.0809 |
Further details of the mechanism of prepassivated copper corrosion can be envisaged by considering that the increase in $E_c$ ($E_c > 0.40$ V) appears as a potential drop in the outer oxide layer. The electrical field strength ($\varepsilon$) in the outer oxide layer depends whether this layer is assigned to CuO or Cu(OH)$_2$. Thus, for CuO, $\varepsilon = 1.3 \times 10^5$ V cm$^{-1}$ and for Cu(OH)$_2$, $\varepsilon = 0.6 \times 10^6$ V cm$^{-1}$.

On the basis of these values the rate of metal corrosion through the passive layer ($j_c$) can be interpreted in terms of cation diffusion from the metal/film interface to the film/solution interface resulting in the dissolution at the metal/film interface[25]. From $P_s$, the difference in the cation concentration between the metal/film ($c_0$) and the film/solution ($c_f$) interfaces can be estimated. Thus, taking $D = 1 \times 10^{-14}$ cm$^2$ s$^{-1}$[26], $\Delta c$ values between 0.05 and 0.07 mol cm$^{-3}$ are obtained at the lower potentials. As $\Delta c$ is close to $10^5$ V cm$^{-1}$ for a film of 20 A and the potential drop ($\phi$) within the oxide film related to the ion transfer can be considered directly proportional to $\varepsilon$

$$\phi = \varepsilon h,$$

the $c_0/c_f$ concentration ratio can be estimated from $\phi$ according to the equation:

$$kT \ln \frac{c_0}{c_f} = \frac{\phi}{z e \varepsilon}$$

where $k$ is the Boltzmann constant and $e$ is the electron charge. For $\varepsilon = 10^6$ V cm$^{-1}$ and $h = 20$ A the $c_0/c_f$ ratio comes out close to $10^6$. This means that $\Delta c \simeq c_f$. Therefore, $K_s$ can be evaluated from the equation

$$K_s = \left( \frac{8 M c_0}{\rho} \right)^{1/2}$$

for $M = 63.54$ g mol$^{-1}$ and $\rho = 8.52$ g cm$^{-3}$, $K_s = 2.99$. Furthermore, from this figure and $P_s = 27$ min$^{-1}$, $N_f$ results equal $5 \times 10^12$ cm$^{-2}$. This value suggests that the metal surface involved in the corrosion process is in the order of a few percent of the total copper surface. The contribution of the copper electrode dissolution through the anodic layer to Cu(II) soluble species can be followed through the current transient of copper electrodeposition at the ring electrode ($E_r = 0.80$ V). This transient can also be fitted by an equation similar to Equation (18). Assuming that copper electrodeposition involves a progressive nucleation and threedimensional growth under diffusion control[20] and taking into account the background current at the ring gold electrode ($P_{s3}$), the corresponding rate equation becomes

$$j_{o3} = \frac{P_{s3}}{\sqrt{t}} \left[ 1 - \exp (-P_{s3} t^2) \right] + P_{s3}.$$  

Both contributions increase as the $q_{mon}$ increases. Equation (24) fits the experimental data (Fig 14) reasonably well with the set of parameters assembled in Table 3. These results confirm that the corrosion current through the passive film is considerably increased in the presence of Cl$^-$ ions.

In the borax buffer containing Cl$^-$ ions at $E_c < E_a$ ($E_a = 0.20$ V), the $P_1/P_2$ ratio gives a $q_{mon}$ value in the order of 8 mC cm$^{-2}$ suggesting that due to the presence of the aggressive anion either a thicker film probably containing Cl$^-$ ions as CuCl is formed or the corrosion area is enhanced. Microscopic observations reveal a certain increase in roughness of the electrode surface but this is insufficient to explain the increase in the $q_{mon}$ value. Furthermore, taking $c_f = 0.05$ mol cm$^{-3}$, as obtained in the free chloride solution, the estimated apparent value of $D'$ resulting in borax + 0.5 M NaCl at $E_r < E_a$ is close to $2 \times 10^{-12}$ cm$^2$ s$^{-1}$, a figure which is two orders of magnitude greater than that resulting in the plain buffer solution. This considerable difference suggests that a layer thicker and more defective is formed in NaCl-containing solutions. This may imply the inclusion of a CuCl layer into the oxide phase. This type of layer can be related to the current oscillations detected in the $E/j$ curves at $E_r < E_a$ (see Fig. 3) which are attributed to breakdown and repair events.

When $E_r > E_a$ localized corrosion of copper takes place when the direct contact between the CuCl salt layer and copper is established. This process involves an induction time which can be associated with the time required to complete the removal of the passive layer at defective points[22]. In this case, the contribution of $j_c$ can be expressed by the following equation

$$j_c = J_{o3} \left[ 1.0 - \exp (-P_{s3} t^2) \right] + P_{s3} \left[ 1.0 - \exp (-P_{s3} t^2) \right].$$  

The first term in Equation (25) is assigned to the corrosion of copper through the passive layer and the second one to the corrosion of copper at the passive film/growing salt nuclei interface. The latter process behaves as an instantaneous nucleation and conical growth under charge transfer control[27]. At $t \ll t < t$, it is reasonable to assume that most of the current is due to the second contribution[22], so that:

$$P_{s3} \left[ 1.0 - \exp (-P_{s3} t^2) \right] \gg \frac{P_{s3}}{\sqrt{t}} \left[ 1.0 - \exp (-P_{s3} t^2) \right]$$

and then, Equation (25) becomes

$$j_c = P_{s3} \left[ 1.0 - \exp (-P_{s3} t^2) \right].$$  

Fig. 14. Current at the rotating ring electrode by stepping the copper disc at $E_r$ (o = 250 rpm). (a) borax buffer solution, $E_r = 0.20$ V; (b) borax buffer + 0.5 M NaCl solution, $E_r = 0.20$ V; 25°C. (+) (x) experimental data, (——) computed from Equation (24).
Table 4. Parameters used in simulation of potentiostatic current transients with Equations (17), (27) and (30)

| $E_i$/V | $P_1$/mA cm$^{-2}$ | $P_2$/min$^{-1}$ | $P_3$/mA cm$^{-2}$ | $P_6$/min$^{-1}$ | $P_7$/mA cm$^{-2}$ | $P_8$/min$^{-1}$ | $t_i$/s |
|--------|-------------------|-----------------|-------------------|-----------------|-----------------|-----------------|--------|
| 0.30   | 5.56              | 24.8            | 1.98              | 124             | 4.42            | 3.68            | 12.8   |
| 0.40   | 13.4              | 81.4            | 6.99              | 572             | 35.4            | 16.8            | 2.10   |
| 0.50   | 21.9              | 74.7            | 15.9              | 808             | 72.2            | 18.6            | 0.530  |
| 0.55   | 42.5              | 111             | 13.4              | 1400            | 97.2            | 28.6            | 0.250  |

$c_{\text{NaCl}}$/M | $P_1$/mA cm$^{-2}$ | $P_2$/min$^{-1}$ | $P_3$/mA cm$^{-2}$ | $P_6$/min$^{-1}$ | $P_7$/mA cm$^{-2}$ | $P_8$/min$^{-1}$ | $t_i$/s |
|---------------------|-------------------|-----------------|-------------------|-----------------|-----------------|-----------------|--------|
| 0.25                | 8.50              | 19.9            | 4.25              | 168             | 9.68            | 11.8            | 3.32   |
| 0.30                | 11.7              | 53.9            | 11.1              | 453             | 24.7            | 24.9            | 2.52   |
| 0.50                | 21.9              | 74.7            | 15.9              | 807             | 72.2            | 18.6            | 0.530  |

where $P_3 = zFk_{3\parallel}$ and $P_6 = \pi M_e^2 k_{3\perp} N_{0s}/\rho_3^2$, $k_{3\parallel}$ and 
$k_{3\perp}$ are formal rate constants for the nuclei growth in 
the directions parallel and perpendicular to the metal 
plane, respectively, $\rho$ is the salt (CuCl) density whose 
molecular weight is $M_e$ and $N_{0s}$ is the number of sites 
available for salt nuclei formation. When $t \to \infty$, 
Equation (27) becomes:

$$j_* = P_3.$$  (28)

Equation (28) was previously used to describe the 
pitting behaviour of nickel and iron in the presence of 
different aggressive anions[15, 16]. In Equations (27) 
and (28) $j_*$ is an "apparent current density" related 
local points where salt nuclei grow through the 
 passive film. The passive layer can be totally removed 
at the sites of attack, which in turn become centres for 
another new phase formation, namely the nucleation 
and growth of salt nuclei on the base copper. The rate 
of the latter process can also be represented as an 
instantaneous nucleation and conical growth under 
charge transfer control, according to the equation

$$j_* = P_3 [1.0 - \exp (-P_6 t^2)].$$  (29)

Obviously, $j_*$ is also an "apparent current density" 
related to the copper surface regions where pitting 
occurs. The expressions of $P_3$ and $P_6$ are similar 
to those of $P_3$ and $P_6$, respectively. When salt nuclei 
formation implies an induction time ($t_{i3}$) Equation (29) 
becomes

$$j_* = P_3 (1.0 - \exp [-P_6 (t - t_{i3})^2]).$$  (30)

The current transients resulting for $E_i > E_s$ 
[Figs 10(a), (b), 11] are reproduced within 5% as 
maximum error, with Equations (17), (27), (30) using 
the parameters assembled in Table 4.

The potential dependence of $P_3$, $P_6$ and $P_7$, 
$P_8$ approaches linear semilogarithmic relationships 
with slopes $dE/d\ln P_3 = 0.22$ V, $dE/d\ln P_6 = 0.30$ V, $dE/d\ln P_7 = 0.18$ V, $dE/d\ln P_8 = 0.08$ V/decade, 
respectively. The slopes $dE/d\ln P_3$ and $dE/d\ln P_8$ are likely 
to be related to the activated electrodissolution of copper at 
the surface, where the chloride salt layer impedes the repassivation of copper. This conclusion agrees with previous kinetic studies on copper 
electrode dissipation in aqueous chloride solutions[6].

Equations (29) and (30) after series expansion yield the following relationships:

$$j_* = P_3 P_6 t^2$$  (31)

and

$$j_* = P_3 P_6 (t - t_{i3})^2$$  (32)

which correspond to particular cases of the Engell-Stolica equation where $b = 2.0$. This value of $b$ is 
approached at the initiation of pitting (Fig. 12).

The change in the kinetics of pit growth at $t \approx t_{i3}$, 
which is accompanied by a decrease in $b$ can be related to 
the precipitation and spread of a thick CuCl layer on 
the metal surface. Secondary breakdown of this layer 
results in copper electrode dissolution as Cu$_2^+$ leading to 
the precipitation of CuCl$_2$. Then, an outer Cu(OH)$_2$ 
layer is also formed acting as a barrier to separate the 
electrolyte in the pit from the bulk electrolyte. The formation 
and growth of this complex layer can be 
represented by an instantaneous nucleation and tridimensional growth under diffusion control. The corresponding rate equation is:

$$j_* = P_9 \sqrt{t} \{1.0 - \exp \left[ -P_10 (t - t_{i3}) \right]\}.$$  (33)

where $P_9 = zF \Delta C_+/\pi^{1/2}$ and $P_10 = \pi N_{0s} K_4 D_+ / \rho_3 M_e^2$; 
$N_{0s}$ denotes the number of sites available for salt 
precipitation inside the pits and $K_4$ is a proportionality 
constant.

When $P_10 (t - t_{i3}) \to \infty$, Equation (33) results:

$$j_* = P_9 \sqrt{t}.$$  (34)

The corresponding current transients [Fig. 10(c)] can be 
simulated with Equation (33) using the parameters 
assembled in Table 5. According to Equation (33) $t_{i3}$ 
depends inversely on $j_*$ as is found experimentally.

According to the proposed kinetic model the current 
transient decay beyond $I_m$ is a consequence of CuCl$_2$ 
formation and cation diffusion. Therefore, the value of 
$D_+$ in Equation (33) can be estimated using $\Delta C_+ = 3.31 \times 10^{-3}$ mol cm$^{-3}$ from CuCl$_2$ solubility[28], 
and the following equation:

$$D_+ = \frac{P_9 \pi^{1/2}}{zF \Delta C_+}.$$  (35)

Table 5. Parameters used in simulation of poten-
tiostatic current transients with Equation (33)

| $E_i$/V | $P_9$/mA cm$^{-2}$ | $P_{10}$/min$^{-1}$ | $t_{i3}$/s |
|--------|-------------------|-----------------|--------|
| 0.30   | 27.9              | 0.552           | 31.3   |
| 0.40   | 42.9              | 2.83            | 8.12   |
| 0.50   | 32.0              | 3.55            | 1.39   |
| 0.55   | 30.3              | 6.55            | 2.19   |


The \( D_z \) values are between \( 1.4 \times 10^{-6} \text{ cm}^2 \text{s}^{-1} \) and \( 3.4 \times 10^{-6} \text{ cm}^2 \text{s}^{-1} \) in reasonable agreement with values of the diffusion coefficient of ions in aqueous concentrated solutions.

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