Design of oxophilic metalloporphyrins: an experimental and DFT study of methanol binding†

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By systematic measurements we have evaluated a series of tetraphenyl metalloporphyrins and halogenated tetraphenyl metalloporphyrin derivatives for binding to ligands with oxygen containing functional groups, using methanol, acetic acid and acetone as examples. Experimental binding constants identified three metalloporphyrins with good binding to all three ligands: MgTPFP, MgTPPB and ZnTPPB as well as a range of porphyrins binding to select ligands. Based on these results the optimal porphyrins can be selected for the desired binding interactions. We also show how to use DFT calculations to evaluate the potential binding between a metalloporphyrin and a ligand, which is deduced from free energies of binding \( \Delta G \), charge transfer \( \Delta Q \) and change of metal spin state. Computations on unsubstituted porphyrins in lieu of tetraphenyl porphyrin systems yield reliable predictions of binding interactions with good correlation to the corresponding experimental data. The calculations have also yielded interesting insights into the effect of halogenation in the \( \beta \)-position on the binding to ligands with oxygen containing functional groups.

Introduction

Metalloporphyrins are well known for their capability to bind ligands at the axial positions of the metal center. This results in a plethora of functionality, based on varying electronic properties of the metalloporphyrin, increased reactivity of the axial ligands, and the possible formation of supramolecular entities. Compounds with several metalloporphyrin units have attracted particular interest, since they allow for a stronger host-guest interaction due to cooperativity or reorganization. Amongst these, bisporphyrin molecular tweezers and clips with two metalloporphyrin units connected by a linker are particularly well known for their capacity of ditopic binding to guest molecules. This is the foundation of Exciton-Coupled Circular Dichroism (ECCD), aimed at the determination of the absolute stereochemistry of guest molecules. Recently, our group has extended the scope of such bisporphyrin tweezers to determine the relative stereochemistry of small flexible molecules using NMR spectroscopy via conformational restriction of the bound guest. In these studies, as in most ECCD investigations, zinc is the metal of choice. While zinc binds favorably to guests with nitrogen-containing functional groups, its affinity to oxygen-containing functional groups is much lower. Most molecules of interest for these analytical methods do not possess two nitrogen-containing functional groups, the presence of at least two oxygen-containing functional groups is more likely. Low affinity between host and guest may be compensated for in ECCD by a huge excess of the guest, since the signal originating from the bisporphyrin, and not the guest is detected. In our NMR-based method, instead we observe the signal originating from the guest molecule. Thus, a huge excess of guest is not tolerable, since it would result in NMR signals that are dominated by the contribution of unbound guest molecules.

Our aim therefore is to identify metalloporphyrins with increased affinity to oxygen-containing functional groups. On one hand it has been shown that this affinity may be increased by a derivatization of the porphyrin that reduces the electron density at the metal center. On the other hand, choice of suitable “oxophilic” metals may result in stronger binding. Also, it should be possible to combine both parameters. Although both approaches have been reported, we found that a systematic approach is lacking. A more complete knowledge of possible binding interactions between metalloporphyrins and common oxygen-containing functional groups also should make it possible to choose a metal based on the desired binding behavior. For example, differences in affinity might be exploited to design bismetalloporphyrins that selectively bind components of a mixture. To facilitate both ECCD and NMR studies our aim is to find porphyrins binding strongly enough so that interactions between host and guests are easily detectible without requiring guest excess.
To achieve this, we investigate here a series of metallated porphyrins, including, based on previous results with electron poor metal centers a number of halogenated metalloporphyrins. To facilitate NMR studies most of the investigated metals are diamagnetic. We have also performed DFT calculations in order to devise a method for accurate prediction of whether a metalloporphyrin will bind strongly enough to oxygen containing functional groups to be detectible even at low guest concentrations. For practical reasons, some of the calculations have been carried out for simplified systems, while others used the full molecular structures used in the experimental study.

Results and discussion

Metalloporphyrins

The porphyrins used in this study are meso-tetraphenylporphyrins (TPP) and derivatives thereof. This porphyrin is widely used in host guest chemistry, molecular electronics and molecular materials due to the relatively good solubility in organic solvents. The TPP derivatives chosen are the tetra(pentafluorophenyl)porphyrin (TPFPP) and the octabromo tetraphenylporphyrin (TPPBr8). Metals were chosen based on their prospective oxophilicity. For example, on the oxophilicity scale calculated by Kepp Mg2+ is more oxophilic than Zn2+. The insertion of metals into the porphyrins followed standard methods. All metalloporphyrins are shown in Scheme 1.

Ligands

The target of this study is to achieve reasonably strong binding of the common oxygen containing functional groups –OH, –COOH and –C=O to metalloporphyrins. To keep the studied systems as simple as possible the ligands chosen were methanol, acetic acid and acetone. The nitrogen containing ligand pyridine that is very well studied as a ligand to porphyrins was included as a reference.

Determination of binding constants

While the standard way to measure binding constants involving porphyrins is UV-vis spectroscopy due to simplicity and high sensitivity, we opted for NMR spectroscopy instead. This allows for direct observation of ligand : metalloporphyrin stoichiometry. Furthermore, possible interference of water, the presence of which is difficult to avoid completely, may be indicated in the spectra. Binding of water during titration of porphyrins may result in ligand exchange rather than binding. In UV-vis titrations, this situation would result in a possibly very small and difficult to detect red-shift. For the NMR titrations, it needs to be considered that the chemical shifts of ligand protons are affected by two counteracting effects. Binding to the Lewis-acidic metal is expected to result in an increase of chemical shift (Δδ > 0), whereas the anisotropy effect of the porphyrin ring system results in decreased chemical shift (Δδ < 0). In practice, normally the anisotropy effect dominates. The gradual change in chemical shift is recorded and the binding constants are then obtained by non-linear fitting (Fig. 1). To probe if the binding constant was directly related to the magnitude of the chemical shift change Δδ, this parameter also was recorded at the same [G]/[H] = 1 ratio for various metalloporphyrin complexes with G = ligand and H = metalloporphyrin (Table 2).

The stoichiometry of the formed complexes are mostly 1 : 1 [eqn (1)], but the zinc porphyrins form 1 : 2 complexes with pyridine [eqn (2)]. As mentioned above, the stoichiometry can be directly observed in NMR as the integrals of the peaks correlate to the number of protons with that chemical shift.

\[
K_1 = \frac{[HG]}{[H][G]} 
\]

(1)

\[
K_2 = \frac{[HG_2]}{[H][HG]} 
\]

(2)

Scheme 1 Metalloporphyrins and the naming used in this study.

Fig. 1 NMR titration curves for MgTPFPP with oxygen containing ligands.
The binding constants can then be calculated from the measured NMR data utilizing an iterative fitting program, using eqn (3), for 1:1 stoichiometry and eqn (4) for 1:2 stoichiometry.25

\[
[H\overline{G}] = \frac{1}{2} \left\{ \left( [\overline{G}]_0 + [H\overline{G}]_0 + \frac{1}{K} \right) - \sqrt{\left( [\overline{G}]_0 + [H\overline{G}]_0 + \frac{1}{K} \right)^2 - 4[H\overline{G}]_0[\overline{G}]_0} \right\}
\]

(3)

\[
[\overline{G}]^2(A) + [\overline{G}]^2(B) + [\overline{G}](C) - [\overline{G}]_0 = 0
\]

(4)

with: \(A = \{ K_1K_2 \}, B = \{ K_2[2K_1[H\overline{G}]_0 - K_2[G\overline{G}]_0 + 1] \}, C = \{ K_1[H\overline{G}]_0 - [G\overline{G}]_0 + 1 \} \).

The measured binding constants are presented in Table 1. In Fig. 2 the binding constants to different metallotetraphenylporphyrin (MTPP) complexes are visualized.

As can be clearly seen the metals binding to MeOH are Al(III), Co(III), Mg(II) and Zn(II). All these also bind to AcOH, with the exception of Mg(II) due to demetallation following porphyrin protonation.21 For Mg(II) we also observe binding to acetone. There is only one case where we observe binding to an oxygen containing functional group but not to methanol. This is in the case of Ti(IV) which binds readily to AcOH but to no other ligand. For Co(III)TPPCl the binding constant to pyridine is too high to be determined by this method. Here, pyridine binds to the metal without any exchange with free pyridine in the solution. This results in one set of signals for the bound pyridine and another set of signals for free pyridine at \([G]/[H] > 1\).

Somewhat surprisingly, Fe(II)TPP, Mn(III)TPPCl and Ni(II) TPP show no binding, as well as the Sn(OH)\(_2\)TPP complex with two anionic hydroxide ligands on the metal which apparently bind more strongly than uncharged methanol ligands.22,26,27

Introduction of halogen substituents, either in the \(\beta\)-positions (i.e. TPFPPs) or as tetrakis(pentafluorophenyl) substituents (i.e. TPFPP) results in enhanced binding to the oxygen containing ligands (Fig. 3–5).

When comparing MgTPFPP and MgTPPB\(_8\) to MgTPP the binding to MeOH is slightly increased (factor 2.6 and 2.8, respectively), but the binding to acetone more substantially so (factor 42 and 51, respectively). It is also interesting to note that the well-known demetallation of MgTPP under acidic conditions occurs substantially less strongly in MgTPFPP and MgTPPB\(_8\). Data from Table 1.

Table 1. Binding constants \(K_a\) for various metalloporphyrins from NMR titrations (1 : 1 stoichiometry, CDCl\(_3\) solutions). Standard errors are given in brackets.

| Porphyrin       | Methanol | Acetic acid | Acetone | Pyridine |
|-----------------|----------|-------------|---------|----------|
| Al(TPP)Cl       | 3.8 x 10^{-3} | (2.2 x 10^{-3})\(^a\) | 2.8 x 10\(^1\) | (2.6 x 10\(^{-3}\)) | nb         | —         | nb         | —         |
| Co(TPP)Cl       | 3.2 x 10^{-3} | (1.5 x 10^{-3}) | 1.7 x 10\(^1\) | (4.9 x 10\(^{-3}\)) | nb         | —         | cc         | —         |
| Co(TPFPP)Cl     | 5.4 x 10\(^0\) | (8.9 x 10\(^{-3}\)) | 1.8 x 10\(^1\) | (3.6 x 10\(^{-3}\)) | nb         | —         | cc         | —         |
| Fe(TPP)         | nb        | —           | nb      | —         | nb         | —         | —          | —         |
| Mg(TPP)         | 5.7 x 10\(^1\) | (3.0 x 10\(^{-2}\)) | D       | —         | 7.4        | (1.4 x 10\(^{-3}\)) | 1.7 x 10\(^1\) | (0.2)     |
| Mg(TPFPP)       | 1.5 x 10\(^2\) | (8.6 x 10\(^{-2}\)) | 2.3 x 10\(^3\) | (8.0 x 10\(^{-2}\)) | 3.1 x 10\(^2\) | (2.9 x 10\(^{-2}\)) | 1.3 x 10\(^1\) | (1.9 x 10\(^{-2}\)) |
| Mg(TPFPP)\(_B\) | 1.6 x 10\(^2\) | (4.6 x 10\(^{-2}\)) | 8.2 x 10\(^2\) | (1.0 x 10\(^{-3}\)) | 3.8 x 10\(^2\) | (2.2 x 10\(^{-2}\)) | 1.3         | (9.3 x 10\(^{-3}\)) |
| Mn(TPP)Cl       | nb        | —           | nb      | —         | nb         | —         | —          | —         |
| Ni(TPP)         | nb        | —           | nb      | —         | nb         | —         | —          | —         |
| Sn(TPP)OH\(_4\) | nb        | —           | nm      | —         | nm         | 1.3 x 10\(^4\) | (6.6 x 10\(^{-3}\)) |
| O=Ti(TPP)       | nb        | —           | 4.8 x 10\(^2\) | (8.1 x 10\(^{-1}\)) | nb         | —         | nb         | —         |
| Zn(TPP)         | 2.4 x 10\(^{-3}\) | (3.8 x 10\(^{-3}\))\(^a\) | 3.6 x 10\(^2\) | (1.0 x 10\(^{-3}\)) | nb         | —         | 6.8 x 10\(^1\) | (0.2)\(^b\) |
| Zn(TPFPP)       | 1.5 x 10\(^1\) | (1.0 x 10\(^{-2}\)) | 1.2 x 10\(^4\) | (1.1 x 10\(^{-3}\)) | 1.1 x 10\(^{-2}\) | (1.4 x 10\(^{-3}\)) | 2.7 x 10\(^1\) | (0.5)\(^b\) |
| Zn(TPFPP)\(_B\) | 2.0 x 10\(^{-3}\) | (1.1 x 10\(^{-3}\)) | 2.5 x 10\(^5\) | (3.1 x 10\(^{-3}\)) | 6.0 x 10\(^{-3}\) | (1.4 x 10\(^{-3}\)) | 3.2 x 10\(^2\) | (6.0 x 10\(^{-2}\))\(^b\) |

nb = no binding; nm = not measured; cc = cannot calculate, binding too strong; D = demetallation. \(^a\) Very broad signals. \(^b\) Stoichiometry 1 : 2.
For the zinc porphyrins the binding to methanol increases for the halogenated MTPFP and MTPFBr₈ (factors 625 and 8333, respectively) but the binding to acetic acid shows only slight changes (factors 3.3 and 0.7, respectively). The most interesting point is that ZnTPPB₈ binds to acetone, unlike the ZnTPP and ZnTPFP. This makes ZnTPPB₈ one of the most versatile metalloporphyrins investigated in this study as it binds to all four ligands with similar strength.

In the cobalt series the CoTPPB₈ was excluded as the binding to acetone for the Co(TPFPP)Cl porphyrin the binding to both methanol and acetic acid was increased as compared to CoTPPCCI. None of the cobalt porphyrins binds to acetone.

The ¹H chemical shift changes for the target ligands’ methyl groups are shown in Table 2. The largest changes for each ligand are indicated in bold face. The most apparent change in chemical shift is less impressive. The same can be said for the cobalt porphyrins. With the exception of the Co(TPFPP)Cl binding to methanol, the chemical shift changes are small whereas the binding constants are good (Table 1). Thus, Δδ seems more dependent on the metal of the metalloporphyrin than the binding strength. This might be a result of the two counteracting effects, i.e. an electronic effect from the metal and an anisotropy effect from the porphyrin ring (and possibly the phenyl rings).

### Computational studies

To rationalize the experimental results a DFT study was carried out. As almost all porphyrins binding to R–OH also bind to R–COOH the smaller hydroxyl group was chosen as functional group. Most of the calculations were carried out on simplified metalloporphyrin systems (MP) to reduce computational costs (vide infra). Calculations were also performed on β-halogenated porphyrins to investigate the effect of halogenation on the binding constant as well as the change in charge distribution. To investigate if calculations on the rudimentary metalloporphyrins (MP) gave results indicative of the full systems (MP) we performed a limited number of calculations on MTPP as well as halogenated MTPP derivatives (MTPFP and MTPFBR₈). All investigated systems are shown in Scheme 2.

Calculations on the metalloporphyrins were performed on the DFT level using the B3LYP-D3 method as described by Grimme. This method has previously been shown to predict the geometry, vibrations and electronic structure of porphyrin systems with reasonable accuracy. No solvent model was used in these calculations since their primary purpose was a comparative binding study of various metalloporphyrins independent of choice of solvent. Choice of solvent has a substantial impact on complexations and therefore without a solvent model the calculated binding constants will be considerably higher than the experimentally measured values, but should show the same trends. As it is well known

| Porphyrin    | Methanol | Acetic acid | Acetone |
|--------------|----------|-------------|---------|
| Zn(TPP)      | 3.37     | 1.64        | 0.48    |
| Zn(TPP)Br₈   | 2.79     | 2.04        | 2.13    |
| Zn(TPFPP)    | 2.33     | 2.09        | 2.02    |

*This metalloporphyrin could not be obtained free from water, thus the binding to oxygen containing functional groups here includes competition with water. nb = no binding.*
that metals can adopt several different spin states the three lowest spin states were investigated for each metalloporphyrin and its corresponding metalloporphyrin-methanol complex.\(^{35}\) For metals with an even number of electrons the states with total spin \(S = 0, 1, 2\) and multiplicity \(M = 1, 3, 5\) were calculated and for metals with an odd number of electrons the states with a total spin \(S = 1/2, 3/2, 5/2\) and multiplicity \(M = 2, 4, 6\) were calculated. The lowest states were used in the calculations of binding energies and theoretical binding constants. The energies of the chosen metalloporphyrins were calculated both with and without methanol as ligand. To investigate the change in charge distribution a natural population analysis (NPA) was carried out on the free methanol ligand and on its complexes.\(^ {36–38}\)

Geometries and symmetries

The simple MPs have planar structures whereas the substituted TPP, TPFPP and TPFPP\(_8\) porphyrins have more complex geometries. The phenyl rings are a source of steric repulsion and depending on their position will affect the symmetry and the energy of the system.\(^{39}\) Different rotamers with respect to the phenyl rings were used as starting geometries in order to find the most stable geometry. Our calculations show that for the free TPP metalloporphyrins (CoTPPCl, MgTPP, ZnTPP) the most stable positioning of the phenyl groups is with rings in opposing meso-positions in parallel orientation \((\|, \text{angle between ring planes } \approx 0^\circ)\), but when binding to methanol both MgTPP and ZnTPP have the opposite rings in an angular orientation \((X, \text{angle between ring planes } \approx 40^\circ)\) (Fig. 6).

For the TPFPP metalloporphyrins the \((X)\) configuration is favoured in both the free porphyrins and the complexes. The only exceptions are MgTPFPP where the \((\perp)\)-configuration is favoured for the free porphyrin, and ZnTPFPP where the \((\perp)\)-configuration is favoured in the methanol complex.

In the TPFPP\(_8\) the \textit{meso}-phenyl rings are restricted into adopting one position due to repulsion from the bromo substituents in the \(\beta\)-positions. These bromo substituents are also the cause of substantial distortions of the porphyrin ring planarity, resulting in a saddle shaped structure (Fig. 7).

We have also observed a distortion of the porphyrin ring in the methanol complexes of Co\([\mu]\)TPPCl, Co\([\mu]\)TPFPPCl and to
a lesser extent in the complexes of Zn(II)TPP and MgTPFPFPP, Fig. 8. This kind of distortion was previously observed for Co(II)TPP by de Melo et al.41

Free energies of binding
The lowest calculated energies of the free porphyrins and their methanol complexes were used to determine the free energies of binding $\Delta G$. From these energies the theoretical binding constants were calculated using (eqn (5)).

$$K_a = e^{-\frac{\Delta G}{RT}}$$

where $K_a$ is the binding constant, $\Delta G$ the free energy of binding, $R$ the universal gas constant and $T$ the absolute temperature. The calculated free energies of binding and the theoretical binding constants are shown in Table 3.

The results are presented graphically in Fig. 9. As can be seen, MgP shows the strongest binding interaction followed by Co(III)P. For the other metalloporphyrins $\Delta G$ are at the same level or weaker than for Zn(II)P. As it is known from the literature that Zn(II)P binds weakly to alcohols the porphyrins of most interest are the ones that show stronger interactions, i.e. (Mg(II)P and Co(III)P).15

The calculations on the $\beta$-halogenated porphyrins showed a stronger binding to methanol, Table 4 and Fig. 10. This effect seems independent of which halogen is used. As it has been observed that halogenation on the phenyl rings of TPP increases the binding constant this was an expected result.15,17 Worth noting is that for Co(III) and Mg(II) there is a clear trend with increased $\Delta G$ in the series F < Cl < Br, whereas for Zn(II) the brominated porphyrin shows the weakest increase in free energy of binding.

![Fig. 8 Distortions of the porphyrin ring planarity in the methanol complexes of CoTPPCl, CoTPFPFPP, ZnTPP and MgTPFPFPP.](image)

### Table 3 Calculated binding energies $\Delta G$ and theoretical binding constants $K_{a,\text{theo}}$ for the formation of MP-MeOH complexes

| Complex formation (spin state of metal indicated) | Change in spin | $\Delta G$ (kJ mol$^{-1}$) | $K_{a,\text{theo}}$ |
|--------------------------------------------------|----------------|--------------------------|-----------------|
| $^{1}\text{Al(III)}\text{PCl} + \text{MeOH} \rightarrow [^{1}\text{Al(III)}\text{PCl-MeOH}]$ | 1 → 1 | −20.4 | 3.70 × 10$^3$ |
| $^{3}\text{Co(II)}\text{P} + \text{MeOH} \rightarrow [^{3}\text{Co(II)}\text{P-MeOH}]$ | 2 → 2 | −18.9 | 2.06 × 10$^2$ |
| $^{5}\text{Ni(II)}\text{P} + \text{MeOH} \rightarrow [^{5}\text{Ni(II)}\text{P-MeOH}]$ | 1 → 1 | −21.8 | 6.72 × 10$^2$ |

* The singlet state is only 2.62 kJ mol$^{-1}$ higher in energy. * The quintet state is only 0.58 kJ mol$^{-1}$ higher in energy. * The triplet state is only 1.99 kJ mol$^{-1}$ higher in energy.

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Dalton Trans., 2018, 47, 11572–11585 | 11577

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The free energies of binding for the complexation between the MTPP and methanol are shown in Table 5. It can be seen that for Co(III)TPPCl and Zn(II)TPP the binding constants are similar to those for their MP congeners. For Mg(II)TPP the difference is almost one order of magnitude when compared to Mg(II)P. This supports the use of the simpler porphyrin derivatives (i.e. without meso-phenyl substituents) in most of the calculations.

The halogenated MTPP derivatives TPFPP and TPPBr8 show about as high or higher free energies of binding $\Delta G$ than the MTPP congener but, depending on the metal, different porphyrins present the best result (Table 6). With the distorted geometries of especially the MTPPBr8 in mind, a direct comparison is difficult. The structures of both the free porphyrin and the complex might be significantly stabilized by solvent. Further computational studies on those systems with different solvent models and ligands would be required to estimate this effect.

The spin state of the metal

The investigated MPs have different spin states as their most stable state. In Table 3 the change in spin state when forming the MP-MeOH complex is shown. As it is known that DFT has limitations regarding the calculation of spin states we also indicate when another spin state is very close in energy.35,42 Considering the free energy of binding $\Delta G$ it is indicated that the Fe(III) and Ti(IV) porphyrins do not form stable complexes, as their $\Delta G$ are positive. For the Al(III), Co(II), Mg(II), Mn(III) and Zn(II) porphyrins there is no change in spin state during complex formation. Co(III) porphyrin goes from a quintet to a singlet when forming the complex and also shows the second highest energy of binding. Both Fe(II) and Ni(II) porphyrin complexes change to a higher spin state. In comparison with the experimental results (Table 1) we might expect detectible binding for any metalloporphyrins that have similar or higher free energy of binding than Zn(II)P. However, of these Fe(II) and

### Table 4 Calculated free energies of binding $\Delta G$ and theoretical binding constants $K_{a,\text{theo}}$ of methanol with MP and MTPP

| Complex formation | $\Delta G$ (kJ mol$^{-1}$) | $K_{a,\text{theo}}$ |
|-------------------|-----------------------------|-------------------|
| Co(II)PF8Cl + MeOH $\rightarrow$ [Co(II)PF8Cl-MeOH] | −32.6 | 5.18 × 10$^5$ |
| Co(II)PCl8Cl + MeOH $\rightarrow$ [Co(II)PCl8Cl-MeOH] | −33.1 | 6.34 × 10$^4$ |
| Co(II)PBr8Cl + MeOH $\rightarrow$ [Co(II)PBr8Cl-MeOH] | −32.3 | 4.65 × 10$^3$ |
| MgPF8 + MeOH $\rightarrow$ [MgPF8-MeOH] | −55.3 | 4.96 × 10$^5$ |
| MgPCl8 + MeOH $\rightarrow$ [MgPCl8-MeOH] | −59.3 | 2.43 × 10$^6$ |
| MgPBr8 + MeOH $\rightarrow$ [MgPBr8-MeOH] | −62.7 | 9.72 × 10$^5$ |
| Zn(II)PF8 + MeOH $\rightarrow$ [Zn(II)PF8-MeOH] | −26.3 | 4.07 × 10$^4$ |
| Zn(II)PCl8 + MeOH $\rightarrow$ [Zn(II)PCl8-MeOH] | −30.7 | 2.43 × 10$^4$ |
| Zn(II)PBr8 + MeOH $\rightarrow$ [Zn(II)PBr8-MeOH] | −24.0 | 1.60 × 10$^4$ |
Ni(n) porphyrins did not show any binding in the experiments. Experimental binding is observed for those metalloporphyrins (Al(n), Co(n), Mg(n) and Zn(n)) that preserve or lower their spin state when forming the complex. The non-binding Fe(n) and Ni(n) porphyrins both would increase their spin state to form a stable complex according to the computations. Apparently, here an increase of spin state indicates non-binding although a negative ΔG indicates binding. It would be interesting to extend this study to probe whether this is a general trend in metalloporphyrins.

**Bond length**

We also investigated the ligand–metal bond length, Fig. 11. We can see a clear correlation between the binding energies and the metal–ligand bond length. A stronger binding corresponds to a shorter bond. For the linear fitting the period 3 metals (Mg and Al) where omitted as they are smaller and differ in orbital composition as compared to the period 4 transition and post-transition metals.

The calculated O–M bond lengths for the complexes of the full MTPP and halogenated MTPP derivative systems can be related to the change in chemical shifts observed on methanol during the experimental determination of the binding constants (Fig. 12). This effect is caused mainly by the anisotropy effect from the porphyrin ring current and is therefore directly related to the distance between the measured methyl group protons and the porphyrin ring.

**Charge distribution**

A natural population analysis (NPA) was performed on the optimized structures of the free metalloporphyrins and the methanol-metalloporphyrin complexes. To probe how the halogenation in β-position affected the charge distribution the change in natural charge for the metal and nitrogen atoms were investigated (Table 7). Interestingly the metals of Mg(n) PX₈ and of Zn(n)PX₈ become more electron depleted, as expected when introducing electron withdrawing groups, whereas the metal of Co(n)PX₈Cl gets more electron rich. The nitrogens of Co(n)PX₈Cl show mixed behaviour which suggest that looking on only the metal and nitrogens might be an over-simplification in the case of the Co(n)PX₈Cl porphyrins. Further study of this, including comparison of orbital overlap and more metals would be interesting. For all the studied porphyrins we also observe an increased free energy of binding upon β-halogenation (Fig. 10).

Investigating the change in charge distribution ΔQ in the MTPP and halogenated MTPP derivatives we can see a similar trend in the Mg(n) and Zn(n) metalloporphyrins (Table 8). For the Co(n) metalloporphyrins there is an increased charge on the nitrogen atoms in Co(n)TPFPPII, but very little change on the metal. In Co(n)TPFPBrCl however there is a large increase in negative charge on the metal and a decrease in negative charge on the nitrogen atoms.

The natural charges for free and complexed methanol were extracted (Table 9). A charge transfer can be observed in the porphyrin as well, but with the charge spread out over the system the change is more perspicuous for the methanol ligand. All of Al(n)PCl, Co(n)PCl, Mg(n)P and Zn(n)P show a negative change in charge when forming a complex with methanol. Those porphyrins also have calculated free energies of binding on the same level as Zn(n)P or higher. Co(n)P as

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**Table 6** Calculated free energies of binding ΔG and theoretical binding constants $K_{a,\text{theo}}$ of methanol with MTPP and halogenated MTPP derivatives

| Complex formation                                      | ΔG (kJ mol⁻¹) | $K_{a,\text{theo}}$  |
|--------------------------------------------------------|---------------|----------------------|
| Co(n)TPPPCl + MeOH $\rightarrow$ [Co(n)TPPMeOH]        | -26.7         | 4.74 x 10⁴           |
| Co(n)TPPFPICl + MeOH $\rightarrow$ [Co(n)TPPFPIClMeOH] | -43.8         | 4.80 x 10⁵           |
| Co(n)TPPBuCl + MeOH $\rightarrow$ [Co(n)TPPBuClMeOH]  | -26.2         | 3.89 x 10⁴           |
| Mg(n)TPP + MeOH $\rightarrow$ [Mg(n)TPPMeOH]          | -32.2         | 1.44 x 10⁵           |
| Mg(n)TPPFP + MeOH $\rightarrow$ [Mg(n)TPPFPMeOH]      | -44.8         | 7.07 x 10⁵           |
| Mg(n)TPPBu + MeOH $\rightarrow$ [Mg(n)TPPBuMeOH]      | -71.7         | 3.70 x 10⁻¹¹         |
| Zn(n)TPPP + MeOH $\rightarrow$ [Zn(n)TPPMeOH]         | -23.3         | 1.24 x 10⁵           |
| Zn(n)TPPFP + MeOH $\rightarrow$ [Zn(n)TPFPMeOH]       | -31.8         | 3.71 x 10⁵           |
| Zn(n)TPPBu + MeOH $\rightarrow$ [Zn(n)TPBuMeOH]       | -40.5         | 1.25 x 10⁷           |

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**Fig. 11** Calculated metal–ligand bond length vs. free energy of binding ΔG for MP-MeOH complexes. Mg and Al are included in the figure for comparison but not included in the data fitting. The reason for this is that they belong to a different period and have a substantially smaller ion radius, and thus are not directly comparable to the other metals in the study.
well as the non-binding Fe(II)P and Ni(II)P also have calculated free energies of binding on the same level as Zn(II)P but show a positive change in charge.

The charge transfer when methanol forms complexes with the MTPP, MTPFPP and MTPPBr$_8$ shows exactly the same pattern as observed for the simplified systems (MP) (Table 10).

In Table 11 the data calculated for the simplified MP systems are compared to the experimental results for the corresponding MTPPs. The calculations of free energies of binding $\Delta G$ result in a number of false positives, i.e. for (Fe(II), Mn(III) and Ni(II)), for which experimentally no binding to methanol was found. In contrast, the change of charge on the ligand correlates well to the experimental results, with negative values indicating binding and positive values indicating non-binding. From this we conclude that the change in charge on the ligand gives a good indication of whether binding can be expected. As these are relatively easy calculations we suggest this to be a useful tool for evaluation of different ligand–metal complexes before synthesising the compounds.

In Table 12 the calculated binding energies and charge transfer for the full systems (TPP, TPFPP and TPPBr$_8$) are compared to the experimentally determined binding constant $K_a$. They show similar results to those obtained for the simplified systems. The change in charge on the methanol ligand seems

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**Table 7** Change in natural charge $\Delta Q$ on the metal and nitrogen atoms of MP and MPX$_8$. The change in charge is calculated relative to MP and given in au

| Porphyrin | Charge of metal (au) | Average charge of N1–N4 (au) | $\Delta Q_{\text{metal}}$ (au) | $\Delta Q_{\text{N}}$ (au) |
|-----------|---------------------|-----------------------------|-------------------------------|-----------------------------|
| Co(μ)P    | 1.790               | −0.160                      | —                            | —                           |
| Co(μ)PF$_6$Cl | 1.782               | −0.162                      | −0.008                        | −0.002                      |
| Co(μ)PCl$_3$Cl | 1.781               | −0.155                      | −0.009                        | 0.004                       |
| Co(μ)PBr$_3$Cl | 1.780               | −0.154                      | −0.010                        | 0.005                       |
| Mg(μ)P    | 1.429               | −0.682                      | —                            | —                           |
| Mg(μ)PF$_6$ | 1.436               | −0.680                      | 0.007                         | 0.001                       |
| Mg(μ)PCl$_3$ | 1.446               | −0.669                      | 0.018                         | 0.012                       |
| Mg(μ)PBr$_3$ | 1.451               | −0.663                      | 0.022                         | 0.019                       |
| Zn(μ)IP   | 1.236               | −0.649                      | —                            | —                           |
| Zn(μ)PF$_6$ | 1.244               | −0.648                      | 0.009                         | 0.001                       |
| Zn(μ)PCl$_3$ | 1.254               | −0.637                      | 0.018                         | 0.013                       |
| Zn(μ)PBr$_3$ | 1.258               | −0.634                      | 0.023                         | 0.015                       |

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**Table 8** Change in natural charge $\Delta Q$ on the metal and nitrogen atoms of MTPP, MTPFPP and MTPPBr$_8$. The change in charge is calculated relative to MTPP and given in au

| Porphyrin | Charge of metal (au) | Average charge of N1–N4 (au) | $\Delta Q_{\text{metal}}$ (au) | $\Delta Q_{\text{N}}$ (au) |
|-----------|---------------------|-----------------------------|-------------------------------|-----------------------------|
| Co(μ)TPP  | 1.788               | −0.158                      | —                            | —                           |
| Co(μ)TPFPPCl | 1.789               | −0.162                      | 0.001                         | −0.004                      |
| Co(μ)TPPBr$_3$Cl | 1.742               | −0.151                      | −0.046                        | 0.007                       |
| Mg(μ)TPP  | 1.427               | −0.677                      | —                            | —                           |
| Mg(μ)TPFPP | 1.439               | −0.674                      | 0.012                         | 0.003                       |
| Mg(μ)TPPBr$_3$ | 1.453               | −0.663                      | 0.026                         | 0.015                       |
| Zn(μ)TPP  | 0.616               | −0.322                      | —                            | —                           |
| Zn(μ)TPFPP | 0.623               | −0.321                      | 0.006                         | 0.002                       |
| Zn(μ)TPPBr$_8$ | 0.628              | −0.317                      | 0.012                         | 0.005                       |

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$a$ The LanL2DZ basis set was used on all metals.

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Fig. 12 Calculated M–O bond lengths compared to the experimentally determined changes in chemical shift $\Delta \delta$ for the methyl group on methanol in $^1$H NMR. The CoTPPCl is not included in this figure as competition with water was observed during the experiments and this affects the binding interaction with the porphyrin (vide supra).
**Table 9** Natural charges for atoms in free and metalloporphyrin-bound methanol. The change in charge $\Delta Q$ upon complexation is shown in brackets and is given in au. Porphyrins not forming stable complexes in the calculations are omitted.

| Porphyrin | H ($\Delta Q$) | O ($\Delta Q$) | C ($\Delta Q$) | CH$_3$ ($\Delta Q$) | Total change |
|-----------|----------------|----------------|----------------|----------------------|--------------|
| None      | 0.24           | −0.37          | −0.16          | 0.10                 |              |
| Al(III)P  | 0.51 (0.28)    | −0.75 (−0.38)  | −0.31 (−0.16)  | 0.22 (0.12)          | −0.14        |
| Co(III)P  | 0.52 (0.28)    | −0.66 (−0.29)  | −0.31 (−0.15)  | 0.22 (0.12)          | −0.04        |
| Co(II)P   | 0.26 (0.02)    | −0.35 (0.03)   | −0.16 (0.001)  | 0.11 (0.01)          | 0.06         |
| Fe(III)P  | 0.24 (0.01)    | −0.35 (0.03)   | −0.15 (0.003)  | 0.11 (0.01)          | 0.05         |
| Mg(III)P  | 0.52 (0.28)    | −0.78 (−0.41)  | −0.31 (−0.16)  | 0.22 (0.12)          | −0.16        |
| Mn(III)P  | 0.25 (0.01)    | −0.34 (0.03)   | −0.15 (0.01)   | 0.10 (0.01)          | 0.05         |
| Ni(III)P  | 0.26 (0.02)    | −0.34 (0.03)   | −0.16 (0.001)  | 0.11 (0.01)          | 0.07         |
| V(III)P   | 0.25 (0.01)    | −0.36 (0.03)   | −0.15 (0.002)  | 0.10 (0.01)          | 0.04         |
| Zn(II)P   | 0.51 (0.27)    | −0.75 (−0.38)  | −0.31 (−0.15)  | 0.21 (0.11)          | −0.14        |

$^a$ The LanL2DZ basis set was used on all metals.

**Table 10** Natural charges for atoms in free and metalloporphyrin-bound methanol. The change in charge $\Delta Q$ upon complexation is shown in brackets and is given in au. Some MPs are included for reference.

| Porphyrin | H ($\Delta Q$) | O ($\Delta Q$) | C ($\Delta Q$) | CH$_3$ ($\Delta Q$) | Total change |
|-----------|----------------|----------------|----------------|----------------------|--------------|
| None      | 0.24           | −0.37          | −0.16          | 0.10                 |              |
| Co(III)P  | 0.52 (0.28)    | −0.66 (−0.29)  | −0.31 (−0.15)  | 0.22 (0.12)          | −0.04        |
| Co(III)PFPPCl | 0.52 (0.28) | −0.66 (−0.30)  | −0.31 (−0.15)  | 0.22 (0.12)          | −0.04        |
| Co(III)PFPPCl | 0.52 (0.28) | −0.66 (−0.30)  | −0.31 (−0.15)  | 0.22 (0.12)          | −0.04        |
| Co(III)PFPPBrCl | 0.52 (0.28) | −0.67 (−0.29)  | −0.31 (−0.15)  | 0.22 (0.12)          | −0.04        |
| Mg(III)P  | 0.52 (0.28)    | −0.78 (−0.41)  | −0.31 (−0.15)  | 0.22 (0.12)          | −0.04        |
| Mg(II)TPP | 0.52 (0.28)    | −0.78 (−0.41)  | −0.31 (−0.15)  | 0.22 (0.12)          | −0.04        |
| Mg(II)TPPP | 0.52 (0.28)   | −0.79 (−0.42)  | −0.31 (−0.15)  | 0.22 (0.12)          | −0.04        |
| Mg(II)TPPPBr$_a$ | 0.52 (0.28) | −0.80 (−0.43)  | −0.31 (−0.15)  | 0.22 (0.12)          | −0.04        |
| Zn(II)P   | 0.51 (0.27)    | −0.75 (−0.38)  | −0.31 (−0.15)  | 0.21 (0.11)          | −0.14        |
| Zn(II)TPPP | 0.51 (0.27)   | −0.75 (−0.38)  | −0.31 (−0.15)  | 0.21 (0.11)          | −0.15        |
| Zn(II)TPPPBr$_a$ | 0.51 (0.27) | −0.76 (−0.39)  | −0.31 (−0.15)  | 0.22 (0.12)          | −0.15        |
| Zn(II)TPPPBr$_a$ | 0.52 (0.28) | −0.77 (−0.40)  | −0.31 (−0.15)  | 0.22 (0.12)          | −0.14        |

**Table 11** Relation between free energies of binding $\Delta G$, change in spin states and change in charge on complexed MeOH to experimentally determined binding constants for the same metal centres.

| Metal$^a$ | DFT calculations (MP) | Experimental (MTPP) |
|-----------|------------------------|---------------------|
|           | $\Delta G$ (kJ mol$^{-1}$) | $\Delta Q$ on MeOH | Change in spin state | Binding constant $K_a$ (M$^{-1}$) |
| Al(III)   | −20.4                  | −0.14               | 1 $\rightarrow$ 1  | 3.8 $\times$ 10$^{-3}$         |
| Co(III)   | −28.4                  | −0.04               | 5 $\rightarrow$ 1  | 3.2 $\times$ 10$^{-3}$         |
| Fe(III)   | −18.9                  | 0.05                | 3 $\rightarrow$ 5  | nb                              |
| Mg(III)   | −10.6                  | 0.05                | 1 $\rightarrow$ 1  | 5.7 $\times$ 10$^1$            |
| Mn(III)   | −21.8                  | 0.07                | 1 $\rightarrow$ 3  | nb                              |
| Ni(III)   | −21.3                  | −0.14               | 1 $\rightarrow$ 1  | 2.4 $\times$ 10$^{-3}$         |

$^a$ Only including metals binding in DFT study and also present in experimental study. nb = no binding.

related to the metal centre rather than the porphyrin as all the Co(III), Mg(III) and Zn(II) porphyrins show almost identical results. When comparing the calculated free energies of binding to the experimentally determined binding constants we obtain a good linear correlation with a $R^2 = 0.9$ (Fig. 13).

When performing calculations on the full systems instead of MP we also gain some additional insights into the relative strengths of the binding constants and the bond lengths (see ESIf), which can be related to chemical shift changes observed in the $^1$H NMR spectra. For the purpose of predicting if a
Another interesting porphyrin is the O
useful for ligand recognition and separation of mixtures. This could potentially be very
strong. Also worth noting is that unlike most of the other
porphyrins the MgTPPBr₈ binds stronger to the oxygen con-
ing and change in
charge for complexed MeOH to experimentally determined binding
constants $K_a$

| Porphyrin                  | Δ$G$ (kJ mol⁻¹) | Δ$Q$ on MeOH | Experimental Binding constant $K_a$ (M⁻¹) |
|----------------------------|----------------|--------------|----------------------------------------|
| Co[η]TPPBr₈                | −26.7          | −0.04        | $3.20 \times 10^3$                      |
| Co[η]TPPBr₈Cl              | −43.8          | −0.04        | $5.36 \times 10^3$                      |
| Mg[η]TPP                  | −52.2          | −0.16        | $5.65 \times 10^3$                      |
| Mg[η]TPPBr₈               | −44.8          | −0.16        | $1.54 \times 10^2$                      |
| Mg[η]TPPBr₈Cl             | −71.7          | −0.16        | $1.63 \times 10^3$                      |
| Zn[η]TPP                 | −23.3          | −0.15        | $2.40 \times 10^3$                      |
| Zn[η]TPPBr₈               | −31.8          | −0.15        | $1.53 \times 10^3$                      |
| Zn[η]TPPBr₈Cl             | −40.5          | −0.14        | $6.61 \times 10^1$                      |

Fig. 13 Calculated free energies of binding Δ$G$ plotted against the
experimentally determined binding constant $K_a$ with MgTPFPP treated
as an outlier. The free energy of binding for MgTPFPP is clearly underes-
timated in the DFT calculations. The $R^2$-value with MgTPFPP included is 0.7.

certain metalloporphyrin will bind to methanol calculations
on a simplified MP system are sufficient.

Conclusion

For the purpose of binding to oxygen containing ligands
Mg[η]-porphyrins and to a lesser extent Co[η]- and Zn[η]-por-
phyrins are the prime choice. In particular, the porphyrin
derivatives TPFPP and TPPBr₈ support binding to various
ligands, i.e. MeOH, AcOH, and acetone. The MgTPFPP is
the most versatile porphyrin derivative in this study. A weaker
binding is observed for ZnTPPBr₈. Both the CoTPPBr₈ and the
ZnTPFPP bind to all ligands except acetone with reasonable
strength. Also worth noting is that unlike most of the other
porphyrins the MgTPPBr₈ binds stronger to the oxygen con-
taining ligands than to pyridine. This could potentially be very
useful for ligand recognition and separation of mixtures. Another interesting porphyrin is the O=TiTPP that only binds
to the carboxylic acid but binds with a high binding constant.

This could also be used to discriminate between ligands or
functional groups on the same ligand.

We also investigated the change in chemical shift of bound
ligands caused by the porphyrin ring current anisotropy effect. With respect to maximization of this effect the TPFPP por-
phyrin stands out. The TPPBr₈ also gives better results than the
TPP. The best choice to get both binding to all tested ligands and a large change in chemical shift is the MgTPFPP metalloporphyrin.

The DFT calculations using B3LYP-D3 show that the metalloporphyrins with both a high free energy of binding and
a negative change in charge on the methanol ligand are those
that experimentally have been found to bind methanol. Thus we
propose to use such calculations as a relatively fast and
easy method to predict binding between metalloporphyrins and
prospective ligands. The calculations on the full MTTP systems showed correlations to the experimental binding
constants and the observed changes in chemical shifts, indicating
that they are a good representation of the studied systems.
Furthermore, the calculations also indicated that formation of
a stable complex cannot be expected if a change to a higher
spin state of the metal is required. It would be interesting to
confirm this preliminary observation by more extended com-
putations involving further metalloporphyrins.

Experimental

Commercially available compounds were used without purifi-
cation. meso-Tetraphenylporphyrin (TPP) was purchased from
Porphyrin Systems GbR, Germany. Octabromo meso-tetra-
phenylporphyrin (TPPB₉) was purchased from Frontier
Scientific. meso-(Tetraperfluoro phenyl)porphyrin (TPFPP) was purchased from PorphyChem. The free porphyrins were
metallated using known methods.²¹,⁴³–⁵⁰ Mn[η] meso-tetrape-
nylporphyrin chloride (Mn(TPP)Cl), Zn[η] meso-(tetrape-
fluoro phenyl)porphyrin (ZnTPPBr₈) and Sn[η] meso-tetrape-
nylporphyrin dichloride (Sn(TPP)Cl₂) were purchased from
Porphyrin Systems GbR, Germany. Ni[η] meso-tetraphenylpor-
phyrin (Ni(TPP)) was purchased from Frontier Scientific.
Analytical TLC was performed using Merck precoated silica gel
60 F₂₅₄ plates and for column chromatography Matrex silica
gel (60 Å, 35–70 μm) or Merck basic aluminum oxide
(60–200 mm) was used.

Computational details

The calculations were performed with the B3LYP functional as
implemented in the Gaussian 09 program package.²⁸,³⁰
Geometries were optimized using the 6-31G(d,p) basis set for
the H-, C-, O-, F-, Cl-atoms and SDD basis set and pseudo
potential for all metals.³⁰,³¹ Frequency calculations were per-
formed at the same level to confirm that a minimum had been
reached and to extract free energy corrections which were evalu-
ated at 298.15 K. The energies were also corrected for dis-
perion effects according to the B3LYP-D3 method by
Grimme.³² When optimizing the geometries of the MTTP and

![Table 12 Comparison of free energies of binding Δ$G$ and change in charge for complexed MeOH to experimentally determined binding constants $K_a$.](image-url)
MTPP structures different rotations of the phenyl rings where used as starting points in order to find the most stable geometry.

As metals can adopt several different spin states the three lowest spin states were investigated for each species. For species with an even number of electrons the states with total spin

\[ S = 0, 1, 2 \quad (M = 1, 3, 5) \]

were calculated and species with an odd number of electrons the states with total spin

\[ S = 1/2, 3/2 \quad (M = 2, 4, 6) \]

were calculated. A stability analysis was performed to ensure that a stable wave-function was attained for all species. As the goal of the computational study was to compare the most stable free metalloporphyrin and methylmetalloporphyrin complex, the states which did not converge after repeated tries were deemed unstable and not relevant to this study.

Natural charges were calculated with Natural Population Analysis (NPA) using Natural Bond Orbital analysis (NBO) version 3 with the B3LYP functional and the 6-31G(d,p) basis set for the H-, C-, O-, F-, Cl-atoms and SDD basis set and pseudo potential for all metals. The MgPBr8 and MgTPPBr8 the LanL2DZ basis set was used for all metals.

UV-Vis spectra were recorded on a Varian Cary 3 Bio spectrophotometer using 5 mm quartz cuvettes.

\(^1\)H and \(^13\)C NMR spectra were recorded on Varian Mercury Plus (\(^1\)H at 300.03 MHz, \(^13\)C at 75.45 MHz), Agilent 400-MR DD2 (\(^1\)H at 399.98 MHz, \(^13\)C at 100.58 MHz), or Varian Unity Inova (\(^1\)H at 499.94 MHz, \(^13\)C at 125.7 MHz) spectrometers at 25 °C unless noted otherwise. Chemical shifts are reported referenced to tetramethylsilane

\([\text{CDCl}_3, \text{H} = 7.26 \text{ and } \text{C} = 77 \text{ ppm}]\)

Mass spectra were recorded on a Voyager-DE PRO MALDI-TOF spectrometer in positive or negative ion-mode without matrix.

For NMR titrations, aliquots of freshly prepared ligand solutions (\([\text{CDCl}_3, \text{AlO}x\text{-filtered}, \text{dried over 3 Å molecular sieves}]\) were added to a solution of the porphyrin in an NMR tube. All NMR spectra were recorded on a 500 MHz spectrometer with CDCl3, \(^1\)H at 7.26 and \(^13\)C at 77 ppm.

The binding constants (Kb) for the complexes were determined by NMR titration utilizing the iterative fitting program published by P. Thordarsson in Matlab R2012b. Both 1:1 and 1:2 complexation models were tested. Kb is calculated by (eqn (3) and (4)). The standard error (SEy) is estimated by (eqn (6)).

\[
SE_y = \sqrt{\frac{\sum (y_{\text{data}} - y_{\text{calc}})^2}{n - k}} \quad (6)
\]

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

We thank Professor Helena Grennberg for valuable insights into porphyrin synthesis and purification. Financial support by the Swedish Research Council (grant nr. 621-2012-3379) and by the Carl Tryggers Foundation (CTS 16:156) is gratefully acknowledged. The computations were performed on resources provided by the Swedish National Infrastructure for Computing (SNIC) at National Supercomputer Centre (NSC), Linköping University.

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