Pyridyl-Thioethers as Capping Ligands for the Design of Heteroleptic Fe(II) Complexes with Spin-Crossover Behavior

Ökten Üngör 1, Dilyara Igimbayeva 1,2, Alina Dragulescu-Andrasi 3, Sandugash Yergeshbayeva 1, Teresa Delgado 4, Samuel M. Greer 1,5, Gabrielle Donalson 1,6, Minyoung Jo 1, Rakhmetulla Erkasov 2 and Michael Shatruk 1,5,*

1 Department of Chemistry and Biochemistry, Florida State University, 95 Chieftan Way, Tallahassee, FL 32306, USA; ou15@my.fsu.edu (Ö.U.); dilya92-22@mail.ru (D.I.); sy18t@my.fsu.edu (S.Y.); sgreer@lanl.gov (S.M.G.); gabby_donalson@yahoo.com (G.D.); mj13h@my.fsu.edu (M.J.)
2 Department of Chemistry, L.N. Gumilyov Eurasian National University, 5 Munaitpasov Str, Astana 010008, Kazakhstan; erkass@mail.ru
3 Department of Chemistry, University of Idaho, Moscow, ID 83844, USA; aandras@uidaho.edu
4 Département de Chimie Physique, Université de Genève, Quai Ernest Ansermet, 30, 1211 Genève, Switzerland; MariaTeresa.DelgadoPerez@unige.ch
5 National High Magnetic Field Laboratory, 1800 E Paul Dirac Dr, Tallahassee, FL 32310, USA
6 Department of Physical & Environmental Sciences, Colorado Mesa University, 1100 North Ave, Grand Junction, CO 81501, USA
* Correspondence: shatruk@chem.fsu.edu

Abstract: Mononuclear heteroleptic complexes [Fe(tpma)(bimz)](ClO4)2 (1a), [Fe(tpma)(bimz)](BF4)2 (1b), [Fe(bpte)(bimz)](ClO4)2 (2a), and [Fe(bpte)(bimz)](BF4)2 (2b) (tpma = tris(2-pyridylmethyl)amine, bpte = S,S′-bis(2-pyridylmethyl)-1,2-thioethane, bimz = 2,2′-biimidazoline) were prepared by reacting the corresponding Fe(II) salts with stoichiometric amounts of the ligands. All complexes exhibit temperature-induced spin crossover (SCO), but the SCO temperature is substantially lower for complexes 1a and 1b as compared to 2a and 2b, indicating the stronger ligand field afforded by the N4S2-coordinating tpma relative to the N4-coordinating tpma. Our findings suggest that ligands with mixed N/S coordination can be employed to discover new SCO complexes and to tune the transition temperature of known SCO compounds by substituting for purely N-coordinating ligands.

Keywords: spin crossover; thioether ligands; heteroleptic complexes; Mössbauer spectroscopy

1. Introduction

The design of coordination environments for achieving reversible spin-state switching (spin crossover, SCO) in transition metal complexes is an interesting endeavor that crosses the boundaries between inorganic, organic, and physical chemistry. It is directly related to the discovery of magnetically bistable materials, as well as to the synthesis of new ligands and evaluation of relative stabilities of the low-spin (LS) and high-spin (HS) electronic configurations. While there are many examples of SCO among complexes of 3d metal ions, the majority of them (>90%) are represented by Fe(II) complexes [1]. In turn, the majority of the SCO Fe(II) complexes feature the metal ion in the [N6] coordination environment [2]. In order to expand the possibilities in the design of SCO complexes, it is important to find alternative coordination environments that can provide the appropriate ligand field strength for the occurrence of SCO.

In recent years, our and other groups have reported SCO in several Fe(II) complexes with the [N6S2] coordination, in which the 5-donor atoms are provided by thioether functionalities [3–11]. In general, one thinks of thioethers as rather weak ligands, and initially we were quite surprised by the observation of SCO in such coordination environment. Nevertheless, the use of Mössbauer spectroscopy unequivocally confirmed the conversion...
between the LS and HS states of the Fe(II) ion in these complexes [5]. Another interesting aspect of such complexes is the large change in the Fe–S bond length (~0.3 Å) that takes place upon conversion between the LS and HS structures, due to population of the antibonding \( \epsilon_g \)-orbitals in the HS state. In comparison, the typical change in the Fe–N bond lengths is ~0.2 Å [12,13]. Thus, the SCO in the complexes with the [N4S2] coordination may also result in larger structural changes propagated through the crystalline lattice. These findings and considerations suggest new possibilities in the synthesis of SCO materials and their spin-state switching behavior.

To obtain further insight into the potency of thioether-containing ligands, we performed a comparative study of Fe(II) complexes containing ligands with and without S-donor atoms (Scheme 1). To that end, we used tetradentate ligands \( S,S' \)-bis(2-pyridylmethyl)-1,2-thioethane (bpte) and tris(2-pyridylmethyl)amine (tpma), which had been shown to afford heteroleptic SCO complexes with the Fe(II) ion [3–5,14–17]. Our study of heteroleptic complexes [Fe(bpte)(bim)]\( X_2 \) (bim = 2,2′-biimidazole; \( X = \text{ClO}_4^-, \text{BF}_4^-, \text{OTf}^- \)) [7] revealed that they exhibit SCO at much higher temperatures (>300 K) as compared to the SCO in the analogous complexes with tpma (~200 K) [15]. The homoleptic complex [Fe(bim)]\( ^2+ \) is known to exhibit only HS behavior [18], but combining the bidentate bim with the tetradentate bpte or tpma consistently leads to SCO in the heteroleptic complexes mentioned above. On the other hand, a hydrogenated derivative of bim, 2,2′-biimidazoline (bimz), affords a homoleptic Fe(II) complex that shows SCO around 110 K [19]. The [Fe(bimz)]\( ^2+ \) cation, however, is quite air-sensitive in solution, as well as in the solid state. All these considerations have prompted us to synthesize and investigate heteroleptic complexes [Fe(tpma)(bimz)]\( X_2 \) (1a: \( X = \text{ClO}_4^- \); 1b: \( X = \text{BF}_4^- \)) and [Fe(bpte)(bimz)]\( X_2 \) (2a: \( X = \text{ClO}_4^- \); 2b: \( X = \text{BF}_4^- \)). Herein, we demonstrate that the former complexes exhibit SCO below 300 K, while the latter show SCO only 300 K. Thus, we confirm once again that bpte is a stronger ligand than tpma, and the mixed N,S-coordinating ligands with thioether functionalities provide a promising platform for the discovery of new SCO materials.

![Scheme 1. Ligands used in this work.](image)

2. Results

2.1. Synthesis

Complexes 1a, 1b, 2a, and 2b were synthesized by reacting Fe(II) perchlorate or tetrafluoroborate salts first with the bidentate ligand (bimz) and then with the tetradentate ligand (tpma or bpte) in a MeOH-EtOH solvent mixture. We found such sequential addition to be the most effective, resulting in good yields of the target heteroleptic complexes. Despite the low solubility of bimz and the resulting complexes in the reaction solvent, this method yields rather pure precipitates of 1a-2b, which can be conveniently recrystallized from MeCN-Et2O. The complexes are sufficiently air-stable in the solid state, but should
be kept in a desiccator for prolonged storage. The only exception was observed for complex 1a, in which case the sample tended to develop a small amount of an Fe(III) impurity upon storage, as will be shown by Mössbauer spectroscopy below.

2.2. Crystal Structures

Single-crystal X-ray diffraction confirmed that compounds 1a-2b crystallize as mononuclear heteroleptic complexes. The FeII ion resides in a distorted octahedral coordination environment formed by one tpma and one bimz ligand in 1a and 1b and by one bpte and one bimz ligand in 2a and 2b (Figure 1). All crystal structures are solvent-free, except for that of 2a, which contains an interstitial acetonitrile molecule.

The TGA analysis of a freshly filtered sample of 2a revealed ~2.5% mass loss when the sample was heated to 340 K (67 °C). This mass loss corresponds to ~0.4 molecule of MeCN per formula unit, which is lower than the solvent content determined by the crystal structure determination, 2a·MeCN. This observation indicates that the interstitial MeCN molecules are lost immediately from the sample of 2a if it is not covered with mother liquid. Indeed, the elemental analysis performed on the solids stored for several days after filtration indicated the lack of interstitial solvent molecules in sample 2a, as well as in the other three samples (see the “Materials and Methods” section).

Examination of metric parameters for the structures determined at different temperatures (Table 1) revealed that the Fe–N bonds in structure 1b undergo substantial elongation when the temperature is increased from 100 to 230 K. The bond lengths at 100 K are characteristic of the LS FeII ion while the ones at 230 K approach typical values for the HS FeII ion [20]. The octahedral distortion parameter $\Sigma_9$, defined as the sum of absolute deviations of 12 cis-N–Fe–N angles from the ideal octahedral value of 90°, also exhibits a large increase at higher temperature. These changes indicate that complex 1b undergoes thermally driven SCO.
Complex 1a exhibits much smaller changes in the Fe–N bond lengths. This structure contains two unique Fe sites in the asymmetric unit, and for both of them the Fe –N distances observed at 230 K are typical of the HS FeII ion. The Fe–N distances observed at 90 K are only slightly smaller, while the \( \Sigma_{90} \) values remain nearly the same at both temperatures.

In contrast to 1b, the crystal structure parameters of 2a and 2b remain almost unchanged between 100 and 230 K, with only a negligible increase in the bond lengths associated with thermal expansion of the lattice. The Fe–N and Fe–S distances are consistent with the LS FeII ions, based on the comparison to the bond lengths in the \( \{\text{N}_4\text{S}_2\} \)-coordinated complexes that exhibit SCO [5].

It is also interesting to compare the octahedral distortion parameter \( \Sigma_{90} \) observed for the \( [\text{Fe}(\text{tpma})(\text{bimz})]^2+ \) cations described in this work to the \( \Sigma_{90} \) values in the previously reported \( [\text{Fe}(\text{tpma})(\text{bim})]^2+ \) cations. The SCO complex \( [\text{Fe}(\text{tpma})(\text{bim})](\text{ClO}_4)_2 \) showed \( \Sigma_{90} = 59.9^\circ \) in the LS state and \( 95.1^\circ \) in the HS state [15], while the corresponding values found for \( [\text{Fe}(\text{tpma})(\text{bimz})](\text{BF}_4)_2 \) (1b) are substantially larger, 79.5(5)\(^\circ\) at 100 K and 114.4(2)\(^\circ\) at 230 K. Likewise, in \( [\text{Fe}(\text{bpte})(\text{bim})]X_2 \) complexes, both of which are LS at 230 K, the values of \( \Sigma_{90} \) are 39.2\(^\circ\) for \( X = \text{ClO}_4^- \) and 43.0\(^\circ\) for \( X = \text{BF}_4^- \) [7], while in \( [\text{Fe}(\text{bpte})(\text{bimz})]X_2 \) these values increase to 45.2(2)\(^\circ\) in 2a and 54.4(2)\(^\circ\) in 2b, respectively. These comparisons clearly demonstrate that the coordination environment of the FeII ion in the bimz-containing complexes becomes more distorted due to the non-aromatic, non-planar structure of the bimz ligand. At the same time, the distortion from the ideal octahedral coordination is smaller in the complexes with bpte as compared to those with tpma, which might in part contribute to the greater stability of the LS state in the former (see below).

In the crystal structure of 1a, the \( [\text{Fe}(\text{tpma})(\text{bimz})]^2+ \) cations form layers separated by the \( \text{ClO}_4^- \) anions (Figure 2a). The interlayer interactions between the cations are weak. The layers are parallel to the (110) planes of the lattice. Within each layer, one of the imidazoline rings of the bimz ligand engages in a \( \pi-\pi \) interaction with a pyridyl ring of the tpma ligand from an adjacent complex, and such interactions create a chain pattern (Figure 2b). The other imidazoline ring of bimz is involved in weaker \( \sigma-\pi \) interactions with a tpma pyridyl ring from the neighbor chain.
Figure 2. Crystal packing of 1a: a side view of four layers of [Fe(tpma)(bimz)]^{2+} cations (a) and a top view of one layer (b). The π-π and σ-π contacts are indicated with black and gray arrows, respectively. Color scheme: Fe = cyan, Cl = yellow, O = red, N = blue, C = gray. Hydrogen atoms are omitted for clarity.

The crystal packing of 1b can be described as chains of [Fe(tpma)(bimz)]^{2+} cations arranged parallel to the c axis, with the BF_{4}^{-} anions separating the chains (Figure 3a). The structural arrangement is quite different from the one observed for 1a. Within each chain, the adjacent cations exhibit π-π interactions between an imidazoline ring of bimz and a pyridyl ring of tpma (Figure 3b). In addition, less pronounced σ-π interactions are observed between C–H groups of the other imidazoline ring of bimz and one of the remaining pyridyl rings of tpma.

Figure 3. Crystal packing of 1b: a top view of chains of [Fe(tpma)(bimz)]^{2+} cations (a) and a side view of one chain (b). The π-π and σ-π contacts within the chain are indicated with black and gray arrows, respectively. Color scheme: Fe = cyan, F = green, O = red, N = blue, C = gray, B = coral. Hydrogen atoms are omitted for clarity.

The crystal packing of 2a-CH\textsubscript{3}CN reveals columns of [Fe(bp\textsubscript{te})(bimz)]^{2+} cations arranged parallel to the c axis, with the ClO\textsubscript{4}^{-} anions and CH\textsubscript{3}CN molecules separating the columns (Figure 4a). Each [Fe(bp\textsubscript{te})(bimz)]^{2+} cation in the column interacts with two neighbor cations via π-π contacts between imidazoline rings, with the interplanar spacing of ~3.49 Å (Figure 4b). Interactions between the columns occur via σ-π contacts between ethylene fragments and pyridyl rings of neighbor bp\textsubscript{te} ligands.

In contrast to structures 1a, 1b, and 2a-CH\textsubscript{3}CN, in which the assembly of cationic complexes into columns takes place via pronounced π-π and σ-π interactions, the structure of 2b reveals the lack of any pronounced crystal packing features. In principle, one could...
distinguish columns of $[[\text{Fe(bpte)(bimz)}]^2]^+$ cations propagating along the $c$ axis (Figure 5a), but only weak van-der-Waals interactions are observed between the cations, with the BF$_4^-$ anions separating the chains. Perhaps the most interesting observation in the crystal packing of 2b is the existence of van-der-Waals S···S contacts (3.58 Å) between the columns (Figure 5b).

![Figure 4](image1.png)

**Figure 4.** (a) The crystal packing of 2a·CH$_3$CN viewed down the $c$ axis, showing the top view of the columns of $[\text{Fe(bpte)(bim)}]^2^+$ cations. (b) A side view of the column of cations, emphasizing the $\pi-\pi$ interactions between imidazoline rings. Color scheme: Fe = cyan, Cl = green, S = yellow, O = red, N = blue, C = gray. Hydrogen atoms are omitted for clarity.

![Figure 5](image2.png)

**Figure 5.** (a) The crystal packing of 2b viewed down the $c$ axis, showing the top view of the columns of $[\text{Fe(bpte)(bim)}]^2^+$ cations. (b) A side view of two columns of cations. The van-der-Waals S···S contacts between the columns are shown with red dotted lines. Color scheme: Fe = cyan, S = yellow, F = teal, N = blue, C = gray. Hydrogen atoms are omitted for clarity.

2.3. Magnetic Properties

Variable-temperature magnetic susceptibility measurements were carried on polycrystalline samples of all four complexes. The temperature was varied at a rate of 1 K/min. For complex 1b, a complete and gradual SCO is observed (Figure 6a). The $\chi T$ value decreases from 3.06 emu·K/mol at 300 K to nearly zero below 130 K. The midpoint of SCO, at which the HS and LS fractions are equal ($f_{\text{HS}} = f_{\text{LS}} = 0.5$), is observed at $T_{1/2} = 175$ K. Such
magnetic behavior is in agreement with the substantial change in the Fe–N bond lengths, which are indicative of the LS configuration at 100 K and the nearly complete HS configuration at 230 K (Table 1). A more gradual SCO was observed for complex 1a. The room-temperature $\chi T$ value of 3.43 emu·K/mol (Figure 6) is consistent with the presence of one HS Fe$^{II}$ ion per formula unit ($S = 2$). The onset of SCO takes place below 180 K, but the decrease in the $\chi T$ value is very gradual, and a plateau is observed below 70 K. The residual HS fraction of $f_{\text{res}} \sim 0.4$ was estimated from this plateau, indicating an incomplete SCO. This result can be explained by the kinetic effect, which “freezes” the SCO process due to the decreased energy of vibronic coupling between the SCO centers below ~60–70 K [5,21]. The decrease in the $\chi T$ curve below 20 K is due to zero-field splitting. The slow decrease of the $\chi T$ value below 180 K explains the rather small changes in the Fe–N bond lengths observed for the crystal structure of 1a between 230 and 90 K (Table 1). The more abrupt SCO for complex 1b can be attributed to the more efficient interactions between the cationic complexes along the columns observed in the crystal structures (cf. Figures 2 and 3).

In contrast to the complexes prepared with tpma, complexes 2a and 2b are characterized by much higher stability of the LS state. Upon warming up, an onset of SCO is observed only at 320 K for 2a and at 350 K for 2b (Figure 6). Thus, it appears that the increase in the ligand field strength due to the replacement of tpma with bpte shifts the SCO to substantially higher temperatures. Interestingly, complex 2a exhibits an abrupt and complete spin transition with $T_{1/2} = 330$ K, while complex 2b also exhibits an abrupt increase in $\chi T$ at 360 K, but the change appears to correspond to ~25% of the Fe centers changing their spin state from LS to HS, hinting at the formation of a mixed LS/HS state. Such mixed states are not uncommon for SCO complexes, and their formation might be related to crystal structure transformations during the spin transition [12,22].

![Figure 6. Temperature dependence of $\chi T$ for complexes 1a and 1b (a) and 2a and 2b (b).](image)

2.4. Mössbauer Spectroscopy

Mössbauer spectra of 1a and 1b were collected at 80 and 250 K in zero applied magnetic field. The isomer shift ($\delta$), quadrupole splitting ($\Delta E_Q$), and linewidth ($\Gamma_{L/R}$) parameters obtained by the spectral fitting procedure, as well as relative areas corresponding to various Fe species detected in the samples, are summarized in Table 2. The spectrum of 1a obtained at 250 K (Figure 7a) exhibits a single quadrupole doublet with $\delta = 1.00(4)$ mm/s and $\Delta E_Q = 2.84(6)$ mm/s. These values are typical of the HS-Fe$^{II}$ ion [23]. In contrast, two quadrupole doublets were observed in the spectrum recorded at 80 K. The first one is again typical of the HS-Fe$^{II}$ species, with $\delta = 1.07(2)$ mm/s and $\Delta E_Q = 3.24(2)$ mm/s, while the second doublet, with $\delta = 0.52(1)$ mm/s and $\Delta E_Q = 0.51(2)$ mm/s, is indicative of the LS-Fe$^{II}$ ion [23]. The ratio between the LS and HS fractions is 0.53(3):0.47(3) $\approx$ 1:1, which is in agreement with the ratio evaluated from magnetic susceptibility data at this temperature (Figure 6).
In contrast, the Mössbauer spectrum of sample 1b recorded at 80 K (Figure 7b) revealed the presence of Fe\textsuperscript{II} ions only in the LS state, with $\delta = 0.509(5)$ mm/s and $\Delta E_Q = 0.598(5)$ mm/s being similar to those observed for the LS-Fe\textsuperscript{II} species of 1a. The Mössbauer spectrum of 1b, however, also revealed the presence of 7% of an impurity phase, with the parameters characteristic of a dinuclear oxo-bridged Fe\textsuperscript{III} complex. We carried out multiple experiments to eliminate this impurity, but it was persistently present in the bulk samples of 1b, indicating that this complex, when prepared in a larger amount for Mössbauer experiments, is susceptible to slight oxidation. The fraction of the ferric impurity remains unchanged in the spectrum of 1b recorded at 250 K, while the dominant Fe\textsuperscript{II} species exist as a mixture of HS-Fe\textsuperscript{II} and LS-Fe\textsuperscript{II} in $\sim$4:1 ratio (Table 2). These observations are consistent with the magnetic behavior of 1b, which reveals the onset of SCO from the HS to LS state upon cooling below $\sim$275 K and the complete conversion to the LS state below 120 K (Figure 6a).

Mössbauer studies on complexes 2a and 2b were not pursued, due to the limitation of our instrument that did not allow probing the range of SCO temperatures (above 300 K) observed for these compounds.

| Complex | $T$ (K) | Site       | $\delta$ (mm/s) | $\Delta E_Q$ (mm/s) | $\Gamma_{LR}$ (mm/s) | Area (%) |
|---------|---------|------------|----------------|---------------------|----------------------|---------|
| 1a      | 80      | LS-Fe\textsuperscript{II} | 0.52(1)       | 0.51(2)             | 0.33(1)              | 53(3)   |
|         |         | HS-Fe\textsuperscript{II} | 1.07(2)       | 3.24(2)             | 0.31                 | 47(3)   |
|         | 250     | HS-Fe\textsuperscript{II} | 1.00(4)       | 2.84(6)             | 0.41                 | 100     |
| 1b      | 80      | LS-Fe\textsuperscript{III} | 0.509(5)      | 0.598(5)            | 0.27                 | 93(2)   |
|         |         | HS-Fe\textsuperscript{III} | 0.46(2)       | 1.47(2)             | 0.27                 | 7(2)    |
|         | 250     | LS-Fe\textsuperscript{II} | 0.51(8)       | 0.56(8)             | 0.37                 | 18(5)   |
|         |         | HS-Fe\textsuperscript{III} | 1.07(2)       | 2.82(2)             | 0.32/0.40            | 75(5)   |
|         |         | HS-Fe\textsuperscript{III} | 0.46(8)       | 1.47(8)             | 0.27                 | 7(5)    |

Figure 7. Zero-field $^{57}$Fe Mössbauer spectra of 1a (a) and 1b (b) recorded at 250 and 80 K in zero applied magnetic field. The spectral components corresponding to the HS and LS states are shown with solid red and blue curves, respectively. In the case of 1b, the ferric impurity is indicated with a solid green curve. The total spectral simulation is traced with a solid gray line.
2.5. LIESST Effect in Complex 1b

SCO complexes that exhibit complete thermally driven spins-state conversion below 300 K frequently show light-induced conversion from the ground LS state to the metastable HS state at low temperatures, a phenomenon known as light-induced excited spin state trapping (LIESST) [24]. While the high temperatures of SCO observed for complexes 2a and 2b are not conducive to the LIESST behavior, and complex 1a shows a rather gradual and incomplete SCO, the slightly more abrupt and complete SCO with \( T_{1/2} = 175 \) K, observed in the magnetic behavior of 1b, prompted us to study the possibility of LIESST in this complex. The optical absorption spectra were recorded on a single crystal of 1b of ~100 \( \mu \)m thickness in the range from 475 to 800 nm. At room temperature, when the sample is in the HS state, a relatively weak absorption band centered at 600 nm was observed, with an extinction coefficient of 30 M\(^{-1}\) cm\(^{-1}\) according to the optical density (the black line in Figure 8a). As the temperature was decreased, the intensity of this band increased and maximized at 6 K in the LS state. Hence, this band corresponds to the \(^1\text{A}_1 \rightarrow ^1\text{T}_1\) d-d transition of the LS state. A tail of the much more intense metal-to-ligand-charge-transfer (MLCT) band was observed at higher energies in the HS state whereas this band saturated at wavelengths below 550 nm in the LS state.

Once the spectra of both the HS and LS states were obtained, the LIESST effect was probed by irradiating the crystal at 6 K. First, the crystal was irradiated with a 532 nm laser operating at 2 mW mm\(^{-2}\). After 5 min of irradiation, the LS \( \rightarrow \) HS conversion achieved only 3.3%, based on the change in the optical density at 600 nm (the green line). Neither longer irradiation times nor a higher laser power improved the conversion. However, by employing a longer wavelength laser, with 650 nm and 2 mW mm\(^{-2}\), a higher conversion was achieved after 5 min (the brown line) due to the improved penetration depth. The conversion further increased under longer irradiation times, reaching a maximum value of 32% after 150 min (the red line).

The stability of the LIESST state was studied by heating the sample at 0.3 K min\(^{-1}\) after irradiation. A gradient HS \( \rightarrow \) LS relaxation was observed above 30 K, with the complete recovery of the LS ground state at 50 K (Figure 8b). Thus, while complex 1b exhibits the LIESST behavior, the stability of the photoinduced HS state is rather low, as compared, for example, to the related complex [Fe(tpma)(xbim)](ClO\(_4\))\(_2\) (xbim = 1,1'-\( \alpha,\alpha'-o\)-xylyl)-2,2'-bimidazole) [15]. This might be explained by the lower cooperativity of intermolecular interactions and less rigid structure of the bidentate ligand in the structure of 1b.

Figure 8. (a) Optical absorption spectra recorded on a single crystal of 1b at room temperature and 6 K before and after photo-excitation. The variation in the d-d absorption band of the LS state under different excitation wavelengths and times is shown. (b) The change in the fraction of the HS state as a function of temperature upon heating from 6 K at 0.3 K min\(^{-1}\).
3. Conclusions

The comparative study of SCO behavior in a series of heteroleptic Fe(II) complexes with the bidentate ligand, bimz, and a tetradentate ligand, either tpma or bpte, demonstrates the stronger ligand field of the N$_2$S$_2$-coordinating bpte as compared to the N$_4$-coordinating tpma. The SCO in the tpma-containing complexes occurs below 300 K, while in the bpte-containing complexes the SCO is observed above 300 K. These findings agree with our earlier observation of substantially higher SCO temperatures in heteroleptic Fe(II) complexes that combine biimidazole-type ligands and bpte as compared to analogous complexes with tpma. Besides, partial LS$\rightarrow$HS photoconversion based on the LIESST effect has been proven for one of the tpma-containing complexes.

This work clearly demonstrates that ligands combining pyridyl and thioether functionalities create a rather strong ligand field at the metal center. It should be pointed out, however, that tpma and bpte do not bear exact topological equivalence, since tpma has a tripodal metal-coordinating skeleton while bpte has a linear skeleton. A more accurate comparison would be offered by investigating complexes of bpte relative to complexes of $N,N'$-bis(2-pyridylmethyl)-1,2-ethanedianime [25]. Such studies, as well as studies of other N/S-coordinating ligands for the synthesis of SCO complexes, are currently under way in our labs, and their results will be reported in due course.

4. Materials and Methods

**Synthesis.** All reactions were performed under an inert N$_2$ atmosphere using standard Schlenk techniques. All reagents and anhydrous solvents were purchased from Millipore Sigma and used as received. Ligands tpma [26], bpte [27], and bimz [28] were synthesized according to the published procedures. Anhydrous commercial solvents were additionally purified by passing through a double-stage drying/purification system (Pure Process Technology, Nashua, NH, USA). Elemental analyses were performed by Atlantic Microlab, Inc. (Norcross, GA, USA).

**Caution:** Complexes between metal ions and organic ligands with perchlorate anions are potentially explosive. The compounds should be prepared in small amounts and handled with great care!

**[Fe(tpma)(bimz)](ClO$_4$)$_2$ (1a).** A yellow solution of Fe(ClO$_4$)$_2$·6H$_2$O (72.6 mg, 0.200 mmol) in 2 mL of methanol (MeOH) was added to a suspension of bimz (27.6 mg, 0.200 mmol) in 8 mL of ethanol (EtOH). The resulting orange suspension was stirred for ca. 30 s, followed by the addition of a solution of tpma (58.0 mg, 0.200 mmol) in 2 mL of MeOH. The mixture was stirred vigorously for 30 min. The resulting dark orange-brown solution was filtered under inert atmosphere, and the filtrate was layered with 8 mL of diethyl ether and left undisturbed. Orange crystals that formed after several days were recovered by filtration, washed successively with cold MeOH (1 mL) and diethyl ether (Et$_2$O, 1 mL), and dried by suction. Yield = 51.2 mg (37.5%). **Elem. analysis:** calcd. (found) for FeCl$_2$O$_8$N$_8$C$_{24}$H$_{28}$ (1a): C, 42.19 (41.74); H, 4.13 (4.45); N, 16.40 (16.43).

**[Fe(tpma)(bimz)](BF$_4$)$_2$ (1b).** A yellow solution of Fe(BF$_4$)$_2$·6H$_2$O (67.6 mg, 0.200 mmol) in 2 mL of MeOH was added to a suspension of bimz (27.6 mg, 0.200 mmol) in 8 mL of EtOH. The resulting orange suspension was stirred for ca. 30 s, followed by the addition of a solution of tpma (58.0 mg, 0.200 mmol) in 2 mL of MeOH. The mixture was stirred vigorously for 30 min. An orange precipitate that formed was recovered by filtration and dissolved in 6 mL of acetonitrile (MeCN). The resulting solution was layered with 8 mL of EtO. Orange crystals that formed after several days were recovered by filtration, washed successively with cold MeOH (1 mL) and diethyl ether (Et$_2$O, 1 mL), and dried by suction. Yield = 81.7 mg (62.1%). **Elem. analysis:** calcd. (found) for FeB$_2$F$_8$N$_8$C$_{24}$H$_{28}$ (1b): C, 43.81 (44.06); H, 4.29 (4.22); N, 17.03 (17.01).

**[Fe(bpte)(bimz)](ClO$_4$)$_2$ (2a).** Complex 2a was prepared in the same manner as complex 1b, starting with 72.6 mg (0.200 mmol) of Fe(ClO$_4$)$_2$·6H$_2$O, 27.6 mg (0.200 mmol) of
bimz, and 55.3 mg (0.200 mmol) of bpte. Yield = 87.6 mg (61.7%). Elem. analysis: calcd.
(found) for FeCl₂O₈N₆C₂₀H₂₆S₂ (2a): C, 35.89 (35.87); H, 3.92 (4.02); N, 12.56 (12.59).

[Fe(bpte)(bimz)](BF₄)₂ (2b). Complex 2b was prepared in the same manner as complex
1b, starting with 67.6 mg (0.200 mmol) of Fe(BF₄)₂·6H₂O, 27.6 mg (0.200 mmol) of
bimz, and 55.3 mg (0.200 mmol) of bpte. Yield = 85.0 mg (66.0%). Elem. analysis: calcd.
(found) for FeB₂F₈N₆C₂₀H₂₆S₂ (2b): C, 37.30 (37.36); H, 4.07 (3.98); N, 13.05 (12.97).

Magnetic Measurements. Magnetic susceptibility measurements were carried out on
polycrystalline samples, using a Magnetic Property Measurement System MPMS-XL
(Quantum Design MPMS-XL) equipped with a superconducting quantum interference
device (SQUID). Magnetic susceptibility was measured in a direct-current applied magnetic
field of 1000 Oe in the 5–400 K temperature range, at cooling and heating rates of 1
K/min. The data were corrected for the diamagnetic contribution from the sample holder
and for the intrinsic diamagnetism using tabulated constants [29].

Mössbauer Spectroscopy. ⁵⁷Fe Mössbauer spectra were collected on a constant acceler-
ation instrument at 80 and 250 K, under zero applied magnetic field. A polycrystalline
sample was placed directly in a custom-made polyethylene cup and constrained with a
tightly fitted stopper. Isomer shifts are quoted relative to an α-Fe metal foil at room tem-
perature. The spectra were simulated with the MossA software [30].

Optical Spectroscopy. Optical measurements were performed on a single crystal of
1b (4.5 × 3.0 × 0.1 mm³) mounted on a copper plate to completely cover a small aperture
drilled in the plate. The assembly was inserted into a cryostat (Janis-Sumitomo) capable
of reaching 4 K. Full optical spectra were recorded in the absorption mode in the 475–800
nm range on a double beam spectrometer (Cary 5000). Investigation of light-induced spin-
state conversion was performed by irradiating the sample at 6 K with a HeNe laser (2
mW/mm²). Two irradiation wavelengths were used, 532 nm and 650 nm, and the latter
was found to be optimal for the penetration depth and efficiency of the light-induced
transformation. The fraction of the HS state (fHS) at any time t was evaluated from the
Vegard’s law: fHS(t) = (ODLS − ODt)/(ODLS − ODHS), where ODLS and ODHS are optical den-
sities at 600 nm measured for the pure LS and HS states, respectively, and ODt is the op-
tical density measured at the specific time at the same wavelength. The ODLS and ODHS
values were obtained from the absorption spectra recorded for non-irradiated samples at
6 K and 300 K, respectively. The OD values were corrected from an eventual baseline jump
or shift by taking the difference between the OD at 600 nm and the OD at 800 nm, where
no noticeable absorption bands occurred in the two states.

X-Ray Crystallography. Single-crystal X-ray diffraction was performed on a Bruker
APEX-II diffractometer equipped with a CCD detector or on a Rigaku-Oxford Diffraction
Synergy-S diffractometer equipped with a HyPix detector. A graphite-monochromated
Mo-Kα radiation source (λ = 0.71073 Å) was used in each case. In a typical experiment, a
single crystal was suspended in Paratone-N oil (Hampton Research) and mounted on a
cryoloop, which was cooled to the desired temperature in an N₂ cold stream. The data set
was recorded as ω-scans at 0.3–0.5° step width and integrated with the Bruker SAINT [31]
or CrysAlis [32] software package. A multi-scan adsorption correction was applied based
on multiple equivalent measurements (SADABS) [33]. The space group was determined
with XPREP [34], and the crystal structure solution and refinement were carried out using
the SHELX software [35]. The final refinement was performed with anisotropic atomic
placement parameters for all non-hydrogen atoms. All H atoms were placed in calcul-
ated positions and refined in the riding model, except for the N-bound H atoms of bimz,
which were located on the difference Fourier electron density map and refined in the rid-
ing model. Full details of the crystal structure refinement and the final structural param-
eters were deposited with the Cambridge Crystallographic Data Centre (CCDC). The
CCDC registry numbers and a brief summary of data collection and refinement are pro-
vided in Table A1 (Appendix A).
Author Contributions: Conceptualization, Ö.Ü., R.E. and M.S.; methodology, Ö.Ü., D.I., A.D.-A., S.Y., T.D., S.M.G. and M.S.; validation, Ö.Ü., D.I., A.D.-A., S.Y., T.D. and S.M.G.; formal analysis, Ö.Ü., A.D.-A., S.Y., T.D., S.M.G. and M.S.; investigation, Ö.Ü., D.I., A.D.-A., S.Y., T.D., S.M.G., G.D., M.J. and M.S.; resources, R.E. and M.S.; writing—original draft preparation, Ö.Ü., T.D. and M.S.; writing—review and editing, Ö.Ü., A.D.-A., S.M.G. and M.S.; visualization, Ö.Ü., A.D.-A. and T.D.; supervision, R.E. and M.S.; funding acquisition, R.E. and M.S. All authors have read and agreed to the published version of the manuscript.

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Appendix A

Table A1. Data collection and structure refinement parameters for complexes 1a, 1b, 2a and 2b.

| Formula | FeClO\textsubscript{3}NiC\textsubscript{6}H\textsubscript{12} (1a) | FeFe\textsubscript{2}NiC\textsubscript{6}B\textsubscript{2}H\textsubscript{6} (1b) | FeClO\textsubscript{3}N\textsubscript{2}C\textsubscript{6}H\textsubscript{8}S\textsubscript{2} (2a\textsubscript{CH\textsubscript{3}}CN) | FeSF\textsubscript{2}Fe\textsubscript{2}NiC\textsubscript{6}H\textsubscript{8}S\textsubscript{2} (2b) |
|---------|-----------------|-----------------|-----------------|-----------------|
| T, K    | 90(1)           | 230(1)          | 100(1)          | 230(1)          |
| CCDC number | 1965486        | 1965490         | 1965488         | 1965484         |
| Space group | P\(\overline{T}\) | P\(\overline{T}\) | P\(\overline{2}\)/c | P\(\overline{2}\)/c |
| \(a, \AA\) | 12.2975(3)      | 12.4579(7)      | 12.836(4)       | 13.348(3)       |
| \(b, \AA\) | 13.6406(4)      | 13.8651(8)      | 17.474(6)       | 17.562(4)       |
| \(c, \AA\) | 17.3808(5)      | 17.512(1)       | 15.887(4)       | 16.243(4)       |
| \(\alpha, ^\circ\) | 81.282(2)       | 82.266(5)       | 12.4579(8)      | 12.2975(3)      |
| \(\beta, ^\circ\) | 77.321(5)       | 77.320(5)       | 129.56(2)       | 130.85(1)       |
| \(\gamma, ^\circ\) | 84.175(2)       | 84.248(2)       | 108.122(1)      | 108.093(1)      |
| \(V, \AA^\text{3}\) | 2809.3(1)       | 2916.4(3)       | 2747(2)         | 2880(1)         |
| \(Z\) | 4               | 4               | 4               | 4               |
| Crystal color | light yellow | light yellow | dark orange | orange |
| Crystal size, mm\textsuperscript{3} | 0.50 × 0.30 × 0.48 × 0.32 | 0.13 × 0.13 × 0.19 × 0.18 | 0.53 × 0.47 | 0.29 × 0.12 × 0.13 × 0.13 × 0.17 × 0.13 |
| \(d_{calc}, \text{g/cm}^\text{3}\) | 1.616           | 1.556           | 1.591           | 1.518           |
| \(\mu, \text{mm}^{-1}\) | 0.791           | 0.762           | 0.636           | 0.606           |
| \(\lambda, \AA\) | 0.71073         | 0.71073         | 0.71073         | 0.71073         |
| \(2\beta, ^\circ\) | 27.48           | 27.48           | 26.37           | 26.37           |
| Total refls. (\(R_{int}\)) | 39132 (0.052)   | 33276 (0.047)   | 26980 (0.148)   | 33221 (0.036)   |
| Unique refls. | 12690           | 13208           | 5568            | 6920            |
| Parameters/Restrains | 787/0           | 787/2           | 456/32          | 454/10          |
| \(R_1, wR_2 \) \([I > 2\sigma(I)]\) | 0.069, 0.174    | 0.076, 0.206    | 0.072, 0.200    | 0.042, 0.120    |
| \(R_1, wR_2 \) (all data) | 0.086, 0.186    | 0.108, 0.228    | 0.077, 0.204    | 0.064, 0.129    |
| Goodness of fit \(^\circ\) | 1.031           | 1.023           | 1.095           | 1.041           |
| Diff. peak/hole, \(e/\AA^\text{3}\) | 1.60, -0.87     | 1.58, -0.50     | 1.59, -1.10     | 0.58, -0.49     |

\(R_1 = \Sigma |F_o| - |F_c|/\Sigma |F_o|; wR_2 = [\Sigma w(F_o^2 - F_c^2)]/[\Sigma w(F_o^2)]; \) \(Goodness-of-fit = [\Sigma \left| w(F_o^2 - F_c^2) \right|/(N_{obs} - N_{params})]^{1/2}\), based on all data.

* \(R_1 = \Sigma |F_o| - |F_c|/\Sigma |F_o|; wR_2 = [\Sigma w(F_o^2 - F_c^2)]/[\Sigma w(F_o^2)]; \) \(Goodness-of-fit = [\Sigma \left| w(F_o^2 - F_c^2) \right|/(N_{obs} - N_{params})]^{1/2}\), based on all data.
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