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Photocatalytic Deoxygenation of N–O bonds with Rhenium Complexes: from the Reduction of Nitrous Oxide to Pyridine N-Oxides

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Abstract

The accumulation of nitrogen oxides in the environment calls for new pathways to interconvert the various oxidation states of nitrogen, and especially their reduction. The large spectrum of reduction potentials covered by nitrogen oxides makes it however difficult to find general systems capable of efficiently reducing various N-oxides. Here photocatalysis unlocks high energy species able to both circumvent the inherent low reactivity of the greenhouse gas and oxidant N₂O (E°(N₂O/N₂) = +1.77 V vs. SHE), and reduce pyridine N-oxides (E°(pyridine N-oxide/pyridine) = −1.04 V vs. SHE). The rhenium complex [Re(4,4'-tBu-bpy)(CO)₂Cl] proved to be efficient to perform both reactions under ambient conditions, enabling the deoxygenation of N₂O as well as synthetically relevant and functionalized pyridine N-oxides.

Introduction

The modern development of intensive soil exploitation and – corollary – the massive production of nitrogen-based fertilizers from atmospheric N₂ has enabled to feed an ever-growing population.¹ Collateral damages however include increased production and accumulation of noxious nitrogen oxides, such as nitrates NO₃⁻, nitrates NO₂⁻ or nitrous oxide N₂O, their reduction back to N₂ being mostly ensured by the natural nitrogen cycle.²³ As defined by environmental scientists, planetary boundaries for the biogeochemical nitrogen cycle have indeed already been crossed, the current nitrogen anthropogenic fixation representing more than twice the estimated boundary (150 Tg N/year vs. 62 Tg N/year).⁴⁶ Since breaking the highly stable N₂ bond via the Haber-Bosch process consumes yearly 2% of the energy worldwide, leakage of nitrogen oxides in the environment represents as well a loss of energy.⁷⁸ In the search for a more sustainable economy, there is hence a need for new pathways able to perform efficient interconversion between oxidation states of nitrogen, and especially reduction of nitrogen oxides to less oxidized molecules.⁹ Chemically speaking, this raises the issue of deoxygenating the N–O bond of both inorganic and organic N-oxides. With the aim to explore new pathways, we decided to harness the potential of homogeneous photoredox catalysis, which has already proven its utility in the deoxygenation of C–O bonds of C₅-molecules, such as CO₂.¹⁰¹¹ In contrast to their C–O carbon counterparts,¹² nitrogen oxides cover a larger spectrum of reduction potentials (over 2.8 V vs. 0.4 V for CO₂) thus raising the level of difficulty to find general photocatalysts (Figure 1a).¹³ A simple model that can be used to showcase this deoxygenation is the ozone-depleting nitrous oxide N₂O, since it leads directly to N₂. Although N₂O is a strong oxidant (E°(N₂O/N₂) = +1.77 V vs. SHE),¹³ homogeneous deoxygenation reactions are scarcely reported, for it is highly inert and a poorly coordinating molecule.¹⁴¹⁵ Among the few thermocatalytic deoxygenating methods, the group of Milstein disclosed the hydrogenation of N₂O to N₂ and H₂O catalyzed by a PNP pincer rhenium complex at 65 °C, under a total pressure of 7 bar (Figure 1b).¹⁶ More recently, our laboratory proposed the deoxygenation of N₂O using disilanes and a catalytic amount of fluoride source under ambient conditions (Figure 1b).¹⁷ To the best of our knowledge, only heterogeneous photodeoxygenation of N₂O has been so far reported.¹⁸ At the opposite end of the redox scale, pyridine N-oxides have, in contrast, considerably lower reduction potentials (E°(pyridine N-oxide/pyridine (1a)/pyridine (2a)) = −1.04 V vs. SHE, Figure 1b) and they represent a much greater challenge.¹⁹ These compounds are useful intermediates in the synthesis of N-heteroaromatic-containing pharmaceutical or agrochemical products, where N-oxides are used for instance as directing groups for C–H bond functionalization.²⁰ In that context, there is a need for mild methods to post-reduce the N-oxide group and afford the N-heterocycle. The main reported methods involve sacrificial oxophilic reagents or high energy reductant based on phosphines,²¹ silane²² or borane²³ derivatives, with thermochemical or (photo)electrochemical activation.²⁴²⁷ Nevertheless, all these methods imply the use of either heating systems or high energy reductants (with E° < 0 V), leading to a loss of energy. There is moreover no report of systems capable of addressing the wide range of potentials (> 2.5 V) required to reduce both nitrous oxide and pyridine N-oxides. Nevertheless, all these methods imply the use of either heating systems or high energy reductants (with E° < 0 V), leading to a loss of energy. There is moreover no report of systems capable of addressing the wide range of potentials (> 2.5 V) required to reduce both nitrous oxide and pyridine N-oxides.

Herein, we disclose that a versatile photocatalytic system based on [Re(bpy)(CO)₂Cl]-type complexes²⁸ is able to efficiently deoxygenate both N₂O and pyridine N-oxides under ambient conditions (20 °C, 1 bar, Figure 1c).
Results and discussion

Design of the system for the photodeoxygenation of N₂O.

The potential of [Re(bpy)(CO)₃Cl] (Re-1) to photocatalyze the conversion of N₂O to N₂ was tested, using Re-1 (5 mol%) as catalyst in acetonitrile in the presence of triethanolamine (TEOA) and 1 bar of N₂O. After 2 h of irradiation with white LEDs at 20 °C, N₂ was formed to our delight in 66% yield according to GC analysis of the gaseous fraction. This result represents the first homogeneous photocatalytic deoxygenation of N₂O to N₂ (Table 1, entry 1).

In the absence of Re-1, no conversion of N₂O was observed: cleavage of the N–O bond does not occur spontaneously under irradiation (Table 1, entry 2). Other blank experiments confirmed that no N₂ is formed in the absence of light, N₂O, or electron donor (Table 1, entries 3–5).

Interestingly, the reaction performed without N₂O afforded H₂ as sole product, as a result from TEOA decomposition in the presence of the catalyst under irradiation (Table 1, entry 4). The detection of traces of H₂ at the end of the reaction (1.5% of the gaseous phase, Table 1, entry 1) was thus ascribed to the use of TEOA. Sacrificial amines in photocatalysis are indeed known to supply the catalytic cycle not only with electrons but also protons, which can in turn be reduced by the photocatalyst to H₂. This prompted us to investigate the influence of the sacrificial electron donor on the outcome of N₂O photocatalytic reduction.

Another tertiary amine, triethylamine (TEA), was used, inducing a drop of the N₂O conversion, the gaseous fraction of N₂ after 2 h decreasing from 66% down to 36% without suppressing the generation of H₂ (Table 1, entry 6). Since H₂ was likely produced through reduction of the protons located α to the nitrogen atom, choice was made to use a bulkier tertiary amine, diisopropylethylamine (DIPEA, Hüning’s base). When DIPEA was used as electron donor, H₂ could not be detected in GC after 2 hours of irradiation, and the N₂ yield increased up to 76% (Table 1, entry 5).
Catalyst optimization. When the reaction was scaled up from 0.1 to 1.2 mmol \( \text{N}_2\text{O} \), the yield with Re-1 as catalyst dropped to 50%, affording a TON of 10 and a TOF of 1.9 \( \text{h}^{-1} \) (Table 2, entry 1). GC-monitoring showed no evolution after 22 h of reaction, indicating potential deactivation of the catalyst. We hence explored the effects of substitution on the bipyridine ligand. We synthesized a series of \([\text{Re}(|\text{bpy}|\text{CO})\text{Cl}]\)-type catalysts with bipyridines substituted either on the 4,4’ or the 6,6’-position, and we studied their ability to photocatalyze the deoxygenation of \( \text{N}_2\text{O} \). Complexes Re-2 and Re-3, possessing bipyridines substituted in the 4,4’-position with methyl or t-butyl groups, respectively, were also active in the deoxygenation of \( \text{N}_2\text{O} \), and led to higher yields in \( \text{N}_2 \) than unsubstituted Re-1, 70% and 86% respectively, corresponding to TONs of 14 and 17. They also showed higher activity than Re-1: TOF reached 3.7 \( \text{h}^{-1} \) and 4.3 \( \text{h}^{-1} \) for Re-2 and Re-3 respectively vs. 0.7 \( \text{h}^{-1} \) for Re-1 (Table 2, entries 1-3). On the other hand, Re-4 and the novel Re-5 complex, featuring methyl and bulky methyldisubstituted in the 6 and 6’ positions of bipyridines, were less efficient on a 24 h timescale, leading to moderate yields in \( \text{N}_2 \) of 55% and 61% respectively (Table 2, entries 4 and 5). This behavior is also present in their initial activities: lower TOF0 (2.7 and 1.1 \( \text{h}^{-1} \) for Re-4 and Re-5 respectively) were measured compared to Re-2 and Re-3.

A more precise monitoring of kinetic profiles for each catalyst was also performed (Figure 2b). The evolution of the \( \text{N}_2 \) yield over 60 h showed that Re-1 deactivated after 22 h, sooner than all the other substituted catalysts, which accounts for the low yields obtained with Re-1. In contrast, Re-2 and Re-3, presenting para-substituted bipyridines, showed both high catalytic activity and increased stability. This phenomenon has been studied in the case of CO2 electroreduction by Kubiak et al.: they showed that substituting the bipyridine ligand in the 4,4'-positions provided sufficient steric hindrance to inhibit the dimerization of a ReO reactive intermediate to an inactive bimetallic complex. The stability increased with the steric bulk of the substituent. The same tendency is observed here in the case of the photoreduction of \( \text{N}_2\text{O} \), and hints towards a similar mechanistic pathway. Complexes Re-4 and Re-5, substituted in the ortho position, demonstrated an increased stability compared to Re-1 but, interestingly enough, the increase in stability brought by substitution on the bipyridine did not come along with an increase in activity. These catalysts indeed required respectively 100 h and 150 h to yield 55% and 61% of \( \text{N}_2 \). A structural explanation could arise from the steric strain between the substituents on the bipyridine and the equatorial carbonyls in Re-4 and Re-5.

Rhenium photocatalysts of type \([\text{Re}(|\text{bpy}|\text{CO})\text{Cl}]\) are thus able to catalyze the deoxygenation of \( \text{N}_2\text{O} \), suggesting that other N=O bonds could be deoxygenated under similar reaction conditions.

### Table 2: Performances of the photocatalysts in the deoxygenation of \( \text{N}_2\text{O} \)

| Entry | Catalyst | TOF0 (h\(^{-1}\)) | \% \( \text{N}_2 \) (TON) | Time (h) |
|-------|----------|-------------------|--------------------------|----------|
| 1     | Re-1     | 1.9               | 50 (10)                  | 22       |
| 2     | Re-2     | 3.7               | 70 (14)                  | 50       |
| 3     | Re-3     | 4.3               | 86 (17)                  | 115      |
| 4     | Re-4     | 2.7               | 55 (11)                  | 100      |
| 5     | Re-5     | 1.1               | 61 (12)                  | 150      |

Reaction conditions: \( \text{N}_2\text{O} \) (1.0 bar, 1.2 mmol), catalyst (50 \( \mu \text{mol} \), 5 \( \text{mol}\% \)), \( \text{CH}_3\text{CN} \) (7.2 mL), DIPEA (1.4 mL), 20 °C.

1, entry 7). We therefore selected DIPEA as the sacrificial electron donor for the rest of our study.
Photodeoxygenation of pyridine N-oxides. Encouraged by the success of N₂O deoxygenation, we considered the possibility of applying this methodology to the deoxygenation of more challenging N=O bonds across the scale of redox potentials, namely of pyridine N-oxides (E₁/₂(pyridine N-oxide (1a)/pyridine (2a)) = -1.04 V vs. SHE). Although they present lower reduction potentials compared to that of N₂O (E₁/₂(N₂O/N₂) = +1.77 V vs. SHE), we envisaged that rhenium-based catalysts should be able to deoxygenate as well their N=O bond (E₁/₂(Re-3/Re-3⁻) = -1.21 V vs. SHE). The reaction was monitored by NMR using rhenium-based Re-1–5 photocatalysts. To our delight, a first attempt using the tBu-substituted catalyst Re-3 under the reaction conditions developed for N₂O enabled to deoxygenate pyridine N-oxide (1a) to pyridine (2a) in 82% yield after 34 h (Table 3, entry 1). The blank experiments confirmed that all components (light, photocatalyst, electron donor) were necessary to perform the reduction (see Supplementary Table S1). Remarkably, the other catalysts also afforded full conversion of compound 1a, albeit with longer reaction times (44-87 h), leading to pyridine (2a) in 85-99% yield (Table 3, entries 2-5). This stands in contrast with our observations with N₂O: here the yields showed steady increase until full conversion with all catalysts. Two explanations may arise for this phenomenon: either the pyridine produced during the reaction delays catalyst deactivation through the stabilization of reactive intermediates, or the higher oxidative character of N₂O has a detrimental effect on the catalyst stability. This effect was particularly beneficial in the case of the 6,6'-dimethylbipyridine rhenium complex Re-4, which led only to 55% yield of N₂ after 100 h and gave full conversion of pyridine N-oxide (1a) within 55 h. Again, 4,4'-substituted catalysts Re-2 and Re-3 displayed higher catalytic activity (TOF₀ = 11 and 16 h⁻¹) than unsubstituted Re-1 and 6,6'-substituted Re-4 and Re-5 (TOF₀ = 5.5, 7.5 and 2.6 h⁻¹, respectively), suggesting similar mechanistic patterns for the deoxygenation of both N₂O and pyridine N-oxide (1a) (Table 3). To study the effect of irradiation, a light on/light off experiment was also performed (Figure 3a). The substrate was only converted when the system was exposed to light, and the reaction stopped once the system was in the dark. Continuous irradiation is thus required to perform the reaction.

From a kinetic standpoint, the possible improvement of catalytic performances as pyridine was released was particularly noticeable in the case of 6,6'-substituted catalysts Re-4 and Re-5. If their initial catalytic activities were low (TOF₀ = 7.5 and 2.6 h⁻¹), they showed steady catalytic activity past 2 hours of irradiation, while the conversion rates for the other catalysts were gently decreasing (Figure 3b). This resulted in Re-4 and Re-5 reaching full conversion before Re-1 and Re-2.

Table 3: Photocatalytic deoxygenation of pyridine N-oxide with different photocatalysts

| Entry | Cat. | TOF₀ (h⁻¹) | Yield (%) | Time to full conversion (h) |
|-------|------|------------|-----------|-----------------------------|
| 1     | Re-3 | 16         | 82        | 34                          |
| 2     | Re-1 | 5.5        | 99        | 63                          |
| 3     | Re-2 | 11         | 85        | 87                          |
| 4     | Re-4 | 7.5        | 96        | 55                          |
| 5     | Re-5 | 2.6        | 87        | 44                          |

Reaction conditions: pyridine N-oxide (100 µmol), catalyst (5 µmol, 5 mol%), CD₃CN (0.5 mL), DIPEA (0.1 mL), 20 °C
Figure 3. Photoreduction of pyridine N-oxide. a. Light and dark experiment with Re-3. Yellow: light on; white: light off. b. Kinetic profiles of pyridine N-oxide photoreduction with different catalysts. Reaction conditions: pyridine N-oxide (100 µmol), catalyst (5 µmol, 5 mol%), CD3CN (0.5 mL), DIPEA (0.1 mL), 20 °C. [a] Determined by 1H NMR analysis of the crude reaction mixture.

Figure 4. Photocatalytic reduction of para-substituted pyridine N-oxides. a. Scope of substrates (yield, time of irradiation until maximum yield). Corresponding Hammett constants of the substituents. Reaction conditions: pyridine N-oxide (100 µmol), Re-3 (5 µmol, 5 mol%), CD3CN (0.5 mL), DIPEA (0.1 mL), 20 °C. Yields were determined by 1H NMR analysis of the crude reaction mixture. b. Yields of the photocatalytic deoxygenation of para-substituted pyridine N-oxides after 30 minutes of irradiation.

Several para-substituted pyridine N-oxides 1b-f could be deoxygenated in 77-99% yield using photocatalyst Re-3 (Figure 4). The electronic parameters of the substituents affected the initial reaction rate: the yields after 30 min of irradiation evolved in line with the electron-withdrawing ability of the substituents as given by their Hammett sigma constants (Figure 4b). Indeed, compounds bearing electron-donating substituents were only poorly converted after 30 min: 1b-d bearing dimethylamino, methoxy or methyl groups (σH = -0.83, -0.27 and -0.17 respectively) reached only 3 to 25% yield, while 1a, 1e and 1f, bearing –H, chloro or cyano groups (σH = 0, 0.23 and 0.66 respectively) yielded after 30 min 26, 30 and 43% of the corresponding pyridines, respectively. This was also reflected in the time to reach full conversion: pyridine N-oxides 1b-d bearing electron-donating substituents (–NMe2, –OMe, –Me) in the para position required longer reaction times (14-126 h) than 1e-f with electron-withdrawing substituents (8 and 0.5 h respectively for the chloro and cyano derivatives). In particular the conversion of the nitrile derivative 1f was extremely fast, maximal (43%) after only 30 min of irradiation. Further irradiation however led to over-reduction of 2f to pyridine 2a (for details see Supplementary Fig. S1).

This is the first photocatalytic system able to perform such a deoxygenation reaction without additional energy sources. In contrast, strong reductants based on Hantzch esters or hydrazine were used by the groups of Konev and Wangelin and Lee.

In order to demonstrate the selectivity of this reaction and its synthetic utility, we explored the deoxygenation of a more complex substrate, that is a 2,6-substituted pyridine N-oxide. Pyridine N-oxide indeed feature an enhanced reactivity compared to the reduced pyridine, enabling the functionalization of the 2,6-positions by C-H activation and, as an example, Hiyama et al. recently reported the synthesis of 2,6-substituted pyridine 3, based on the oxidation of o-picoline to form the 2-methylpyridine N-oxide 4. A Ni-catalyzed coupling reaction leads to 2,6 substituted pyridine N-oxide 5, whose deoxygenation presents multiple challenges that are chemoselectivity and steric hindrance. In their publication, the group of Hiyama used PCl3 as a reducing agent to perform such a deoxygenation. We were very pleased to observe that using our system, Re-3 catalyzed the...
deoxygenation with 82% yield after 8 h of irradiation (Scheme 1). This demonstrates the validity of such an approach for mild deprotection strategies, which may be extended to other substrates.

![Scheme 1. 2-Methylpyridine functionalization via the N-oxide strategy](image)

**Conclusion**

In summary, a new photochemical method has been developed to deoxygenate N–O bonds with radically different reduction potentials, over a 2.8 V window. [Re(4,4'-Bu-bpy)(CO)Cl] (Re-3) is indeed capable of reducing both N₂O under ambient conditions, and pyridine N-oxides 1 and 4 in good to excellent yields. Those results open new perspectives concerning the photocatalytic deoxygenation of nitrogen oxide-containing compounds. Further mechanistic studies on catalyst Re-3 are underway in our laboratory. We believe that these will give new clues to understanding N–O bonds deoxygenation chemistry.

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**Data availability**

Data relating to the characterization data of materials and products, general methods, optimization studies, experimental procedures, gas chromatography and NMR spectra are available in the Supplementary Information. Crystallographic data for compounds Re-4 and Re-5 are available free of charge from the Cambridge Crystallographic Data Centre (CCDC) under reference numbers 2056048 and 2056049, respectively.

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Supporting Information for
Photocatalytic Deoxygenation of N–O bonds with Rhenium Complexes: from the Reduction of Nitrous Oxide to Pyridine N-Oxides

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1 Supplementary Data

1.1 Control experiments for the photocatalytic reduction of pyridine N-oxide 1a

Table S1: Photocatalytic reduction of pyridine N-oxide 6a and control experiments

| Entry | Deviation from standard conditions[a] | Conversion of 1a (%)[b] | Yield of 2a (%)[b] |
|-------|--------------------------------------|------------------------|-------------------|
| 1     | None                                 | 100                    | 78                |
| 2     | No light                             | 0                      | 0                 |
| 3     | No photocatalyst                     | 0                      | 0                 |
| 4     | No pyridine N-oxide                  | 0                      | 0                 |
| 5     | No DIPEA                             | 0                      | 0                 |

[a] Standard reaction conditions: 6a [100 µmol], Re-3 [5 µmol, 5 mol%), CD3CN (0.5 mL), DIPEA (0.1 mL), 20 °C, 24 h of irradiation.
[b] Conversions and yields measured by 1H NMR analysis of the crude reaction mixture (internal standard: mesitylene).

1.2 Photocatalytic reduction of 4-cyanopyridine N-oxide 1f

In the case of the photocatalytic reduction of 1f, the conversion to 2f was maximal (77%) after 30 minutes of irradiation. Further irradiation of the same NMR tube afforded pyridine 2a in 100% conversion after 8 hours (Figure S1). The decyanation of 2f to 2a by irradiation of \([\text{Re(bpy})(CO)_3(2f)]^+\) in DMF:TEOA was reported by the group of Ishitani, affording \([\text{Re(bpy})(CO)_3(CN)]\) and free pyridine 2a.¹ Thus, the conversion of 2f to 2a observed here is proposed to occur via the formation of intermediate \([\text{Re(bpy})(CO)_3(2f)]^+\).
Figure S1: Photocatalytic reduction of 4-cyanopyridine N-oxide 1f with Re-3: evolution of the reaction mixture (1H NMR, aromatic region).
2 Experimental Details

2.1 Material and methods

2.1.1 General

All reactions and manipulations were performed at 20 °C in a recirculating mBraun LabMaster DP inert atmosphere (Ar) drybox and/or using Schlenk lines. Glassware was dried overnight at 120 °C. Unless otherwise stated, all the reagents were purchased from commercial suppliers (Aldrich, Acros, Alfa Aesar, TCI). Nitrous oxide was purchased from Messer (purity ≥ 99.998%). Non deuterated solvents were thoroughly dried and distilled by standard methods prior to use. Deuterated solvents were dried on molecular sieves (4 Å; Aldrich) before use. 1H and 13C NMR spectra were recorded on a Bruker Advance Neo spectrometer operating at 400 MHz for 1H. Chemical shifts for 1H and 13C{1H} NMR spectra were referenced to solvent impurities. Coupling constants J are given in Hz. The following abbreviations are used: s, singlet; d, doublet; t, triplet; m, multiplet. Gas chromatography data were collected on a Shimadzu GC-2010 Plus. UV-visible spectra were recorded on a PerkinElmer Lambda 900 spectrometer.

2.1.2 Experimental setup for irradiation

A custom-made irradiation setup was used for all photocatalytic experiments. It is constituted of a hexagonal aluminum chassis with 6 regularly disposed holes, one on each side. A Light Emitting Diode (cool white LED, GU5.3 / MR16 12V 8W SMD 80°, 6000-8000 K) is built into each hole. On top, the apparatus displays 6 equivalent positions that can host NMR tubes for reproducible irradiations, and one hole in the middle used for scale-up experiments (Figure S2, right). A cooling fan under the reactor enables to evacuate the heat produced by the absorbing solution and/or by exothermic reactions. The whole apparatus is supported by a magnetic stirrer; stirring enables to maximize the absorption of photons by the solution. In our experiments, stirring was performed at 300 rpm.
2.2 Synthetic procedures for catalysts

Complexes Re-1 to Re-3 were prepared according to reported procedures. The NMR data are in agreement with those reported in the literature. Complexes Re-4 and Re-5 were synthesized using the same methodology.

2.2.1 Synthesis of [Re(6,6′-dmb)(CO)₃Cl] Re-4

\[
\text{Re(CO)₅Cl} + \text{6,6′-dimethylbipyridine} \xrightarrow{\text{toluene, 110 °C, 2 h}} \text{[Re(6,6′-dmb)(CO)₃Cl]} \]

Scheme S1: Synthetic pathway towards [Re(6,6′-dmb)(CO)₃Cl] Re-4.

In a glovebox, a 50 mL round-bottomed flask equipped with a stirring bar and a J-Young valve was charged with 6,6′-dimethylbipyridine (88 mg, 478 µmol, 1 equiv.), [Re(CO)₅Cl] (173 mg, 478 µmol, 1 equiv.), and 15 mL of anhydrous toluene. The flask was sealed, brought out of the glovebox, and immersed in an oil bath at 110 °C. After 2 hours, the resulting yellow solution was cooled down to room temperature and the solvent was removed in vacuo to afford a crude yellow solid. The latter was suspended in pentane and centrifuged three times for 10 min. The solid was dried under vacuum overnight to give [Re(6,6′-dimethylbipyridine)(CO)₃Cl] Re-4 (170 mg, 73%) as a yellow powder. Single crystals suitable for X-ray diffraction were obtained by slow diffusion of pentane into a THF solution of Re-4 at room temperature.
The NMR, IR and UV-visible data for Re-4 are in agreement with that reported in the literature.⁴

\[ ^1H \text{NMR (400 MHz, CD}_2\text{Cl}_2) \delta 8.01 (d, J = 8.0 \text{ Hz}, 2\text{H}), 7.92 (t, J = 7.9 \text{ Hz}, 2\text{H}), 7.47 (d, J = 7.4 \text{ Hz}, 2\text{H}), 3.10 (s, 6\text{H}), \text{ ppm} \]

\[ ^{13}C \text{NMR (400 MHz, CD}_2\text{Cl}_2) \delta 163.0, 158.3, 139.4, 126.9, 121.3, 30.4 \text{ ppm} \]

IR \( \nu(\text{CO}): 2013, 1894, 1881 \text{ cm}^{-1} \)

UV-Vis (acetonitrile) \( \lambda_{\text{max}} (\varepsilon) = 315 (6270), 329 (5330), 355 \text{ nm (1760 M}^{-1} \text{.cm}^{-1}) \)

Figure S3: View of Re-4 with displacement ellipsoids drawn at the 50% probability level and hydrogen atoms omitted.

2.2.2 Synthesis of \([\text{Re}(6,6'\text{-mesbpy})(\text{CO})_3\text{Cl}] \text{ Re-5}\)

Synthesis of 6,6'-dimesityl-2,2'-bipyridine (mesbpy) was performed as previously reported by the Suzuki coupling of 6,6'-dibromo-2,2'-bipyridine with mesitylboronic acid.⁵

\([\text{Re}(6,6'\text{-mesbpy})(\text{CO})_3\text{Cl}] \text{ Re-5}\) was prepared as follows: in a glovebox, a 50 mL round-bottomed flask equipped with a stirring bar and a J. Young valve was charged with 6,6'-dimesityl-2,2'-bipyridine (140 mg, 357 mmol, 1 equiv.), \([\text{Re(CO})_5\text{Cl}] (130 \text{ mg, 359 mmol, 1 equiv.)}, \) and 15 mL of anhydrous toluene. The flask was sealed, brought out of the glovebox, and immersed in an oil bath at 110 °C. After 2 hours, the resulting yellow solution was cooled down to room temperature, and the solvent was removed \textit{in vacuo}. The resulting crude yellow solid was purified by column chromatography on silica
gel (DCM:MeOH 99:1). The solid obtained after evaporation of the solvent was suspended in pentane and centrifuged three times for 10 min. The solid was dried under vacuum overnight to give [Re(6,6’-dimesitylbipyridine)(CO)₃Cl] Re-5 (170 mg, 68%) as a yellow powder. Single crystals suitable for X-ray diffraction were obtained by slow diffusion of pentane into a THF solution of Re-5 at room temperature.

¹H NMR (400 MHz, CD₂Cl₂) δ 8.29 (d, J = 8.3 Hz, 2H), 8.10 (t, J = 7.9 Hz, 2H), 7.40 (d, J = 7.7 Hz, 2H), 6.98 (s, 4H), 2.34 (s, 6H), 2.15 (s, 6H), 2.00 (s, 6H) ppm

¹³C NMR (400 MHz, CD₂Cl₂) δ 192.3, 190.1, 164.3, 158.0, 139.6, 139.0, 138.7, 136.4, 135.4, 128.9, 128.3, 122.4, 20.9, 20.8, 20.1 ppm

IR ν(CO): 2016, 1918, 1897 cm⁻¹

UV-Vis (acetonitrile) λmax (ε) = 315 (7150), 329 (5940), 370 nm (1400 M⁻¹.cm⁻¹)

Elemental analysis calcd (%) for C₃₁H₂₈ClN₂O₃Re (698.13): C 53.33, H 4.04, N 4.01; found: C 53.60, H 3.88, N 3.97.
2.3 Photocatalytic procedures

2.3.1 Photocatalytic reduction of N₂O on NMR scale

The procedure is detailed with photocatalyst \([\text{Re(bpy})(\text{CO})_3\text{Cl}] \text{ Re-1}\). In a glovebox, a 3 mL NMR tube equipped with a \(J. \text{ Young} \) valve was charged with \([\text{Re(bpy})(\text{CO})_3\text{Cl}] \text{ Re-1}\) (2.3 mg, 5 \(\mu\)mol, 5 mol%), followed by CD₃CN (0.5 mL) and DIPEA (0.1 mL). The tube was sealed and brought out of the glovebox. The reaction mixture was degassed by 3 freeze-pump-thaw cycles and exposed to a N₂O atmosphere (1 bar, ca. 100 \(\mu\)mol, ca. 2.4 mL, 1 equiv.). The tube was then left 2 hours under irradiation using the setup described in Figure S2. The gaseous fraction was sampled using the setup described in Figure S10 and analyzed by GC to determine the N₂ yield.

2.3.2 Photocatalytic reduction of N₂O on a 1.2 mmol scale

The procedure is detailed with photocatalyst \([\text{Re(bpy})(\text{CO})_3\text{Cl}] \text{ Re-1}\). In a glovebox, a Schlenk flask (Figure S5) equipped with two \(J. \text{ Young} \) valves, a GC septum (Supelco, thermogreen TM LB-2) and a stirring bar was charged with \([\text{Re(bpy})(\text{CO})_3\text{Cl}] \text{ Re-1}\) (22.4 mg, 60 \(\mu\)mol, 5 mol%), followed by 6 mL acetonitrile and 1.2 mL DIPEA. The two valves were sealed, and the flask was brought out of the glovebox. The resulting yellow solution was degassed by three freeze-pump-thaw cycles, then exposed to a N₂O atmosphere (1.0 bar, ca. 1.2 mmol, ca. 28.8 mL, 1 equiv.). The flask was left to irradiate in the center of the setup described in Figure S2. At regular intervals valve (*) was opened and 150 \(\mu\)L of the headspace were sampled and analyzed by GC; then, valve (*) was closed.

![Figure S5: Sideview of the scale-up reaction setup.](image)
TONs for $N_2$ were calculated as follows:

$$\text{TON} = \frac{n(N_2)}{n(\text{catalyst})}$$

where $n(N_2)$ is the final quantity of $N_2$ in the reaction setup.

The different values of $n(N_2)$ for each catalyst and corresponding TONs are given in Table S2.

### Table S2: Calculation of TON $N_2$ for each catalyst on 1.2 mmol scale. Reaction conditions: $N_2O$ (1.0 bar, 1.2 mmol), photocatalyst (60 µmol, 5 mol%), $CH_3CN$ (6 mL), DIPEA (1.2 mL).

| Entry | Catalyst | TOF$_0$ (h$^{-1}$) | Time of irradiation (h) | n($N_2$) (mmol) | % $N_2$ (TON) |
|-------|----------|-------------------|-------------------------|-----------------|--------------|
| 1     | Re-1     | 0.7               | 22                      | 0.3             | 22 (5)       |
| 2     | Re-2     | 3.7               | 50                      | 0.84            | 70 (14)      |
| 3     | Re-3     | 4.3               | 115                     | 1.02            | 86 (17)      |
| 4     | Re-4     | 2.7               | 100                     | 0.66            | 55 (11)      |
| 5     | Re-5     | 1.1               | 150                     | 0.73            | 61 (12)      |

### 2.3.3 Photocatalytic reduction of pyridine $N$-oxides

The procedure is detailed in the case of photocatalyst [Re(tbbpy)(CO)$_3$Cl] Re-3:

In a glovebox, a 3 mL NMR tube equipped with a J. Young valve was charged with Re-3 (2.9 mg, 5 mol%), followed by pyridine $N$-oxide (9.5 mg, 100 µmol, 1 equiv.), $CD_3CN$ (0.5 mL), DIPEA (0.1 mL), and mesitylene as an internal standard (10 µL). The tube was sealed, brought out of the glovebox, then left under irradiation using the setup described in Figure S2. The conversion rate and pyridine yield were determined by $^1$H NMR integration (protons ortho to the nitrogen atom) vs. mesitylene as an internal standard.

Representative NMR spectra for the photocatalytic deoxygenation of pyridine $N$-oxide 1a with Re-3 are given in Figure S6. The NMR data for all substituted pyridines are in agreement with those reported in the literature.$^6$
2.3.3.1 *Synthesis of 6-methyl-2-(4-octen-4-yl)pyridine from 2-methylpyridine N-oxide* 4

2-methylpyridine was oxidized to 2-methylpyridine N-oxide according to a reported literature procedure, then was engaged in a nickel-catalyzed hydroheteroarylation with 4-octyne following the procedure reported by Hiyama *et al.* The product obtained, 6-methyl-2-(4-octen-yl)pyridine N-oxide 5 (E/Z: 93/7), was deoxygenated using the following procedure: in a glovebox, a 3 mL NMR tube equipped with a J. Young valve was charged with Re-3 (2.9 mg, 5 µmol, 5 mol%), followed by 6-methyl-2-(4-octen-yl)pyridine N-oxide (21.9 mg, 100 µmol, 1 equiv.), CD$_3$CN (0.5 mL), DIPEA (0.1 mL), and mesitylene as an internal standard (10 µL). The tube was sealed, brought out of the glovebox, then left under irradiation using the setup described in Figure S2. The conversion rate and pyridine yield were determined by $^1$H NMR integration (alkene proton) vs. mesitylene as an internal standard. The reaction afforded 6-methyl-2-(4-octen-4-yl)pyridine 3 in 82% yield after 8 h of irradiation. The NMR data are in agreement with those reported in the literature.

![Figure S6: Photocatalytic reduction of pyridine N-oxide 1a with Re-3: evolution of the reaction mixture ($^1$H NMR, aromatic region).](image-url)
3 GC analyses

3.1 GC conditions

Column: Carboxen 1010 Plot fused silica capillary column (30 m × 0.53 mm × 30 μm); injection temperature: 230 °C; column temperature: 150 °C; flow: 5 mL/min; purge: 2 mL/min; split ratio: 5.0; carrier gas: argon; detector: TCD 230 °C, 30 mA.

3.2 GC calibration

![Figure S7: GC Calibration curve for N₂.](image)

![Figure S8: GC Calibration curve for N₂O.](image)
### Figure S9: GC Calibration curve for H$_2$.

### Table S3: Retention times and response coefficients of analyzed species.

| Species | Retention time (min) | Response factor (µV.min.µL$^{-1}$) |
|---------|----------------------|-------------------------------------|
| H$_2$   | 1.96                 | 177650                              |
| N$_2$   | 2.08                 | 15910                               |
| N$_2$O  | 4.45                 | 15639                               |

### 3.3 Analysis of the gaseous fraction from a J-Young NMR tube

The J. Young NMR tube was connected to a small glass chamber, itself connected to the Schlenk line. The chamber was sealed with a GC septum (Supelco, thermogreenTM LB-2, Figure S10) and placed under high vacuum ($V = 1$ mL, $P = 10^{-2}$ mbar) to limit air contamination. The chamber was closed, and the J. Young NMR tube was open, allowing the gaseous phase to expand. The gas phase was sampled through the septum using a Hamilton® SampleLock syringe and injected right away into the GC apparatus.
Figure S10: Side view of the apparatus used to analyze the gas phase from a J. Young NMR tube.

Table S4: Results of the control experiments.

| Entry | Deviation from standard conditions[a] | H$_2$ (%) [b] | N$_2$ (%) [b,c] | N$_2$O (%) [b] |
|-------|--------------------------------------|--------------|----------------|--------------|
| 1     | None                                 | 1.5          | 66             | 32           |
| 2     | No light                             | 0            | 4              | 96           |
| 3     | No photocatalyst                     | 0            | 2              | 98           |
| 4     | Argon instead of N$_2$O              | 91           | 9              | 0            |
| 5     | No TEOA                              | 0            | 8              | 92           |
| 6     | DIPEA instead of TEOA                | 0            | 76             | 24           |

[a] Standard reaction conditions: N$_2$O (1 bar, 100 µmol), Re-3 (5 µmol, 5 mol%), CD$_3$CN (0.5 mL), DIPEA (0.1 mL), 20 °C, 12 h of irradiation. [b] Determined by GC analysis of the gaseous phase. [c] It was impossible to obtain N$_2$ levels lower than 2 µL in the blank experiments; that residual amount was attributed to contamination by ambient air between sampling and injection for analysis.
4 Analysis data

4.1 Luminescence quenching experiments

For luminescence quenching experiments, a 1 mM solution of the photocatalyst in dry acetonitrile was prepared in a glovebox filled with argon. Out of the glovebox, the photocatalyst was irradiated: at 370 nm for Re-1, at 410 nm for Re-3. The change of the luminescence upon addition of DIPEA at different concentrations was recorded. These demonstrated that the luminescence was efficiently quenched by DIPEA for both complexes.

![Figure S11: Changes in the luminescence spectra of Re-1 upon addition of DIPEA at different concentrations in acetonitrile.](image1)

![Figure S12: Stern-Volmer quenching of Re-1 luminescence with DIPEA.](image2)
Figure S13: Changes in the luminescence spectra of Re-3 upon addition of DIPEA at different concentrations in acetonitrile.

Figure S14: Stern-Volmer quenching of Re-3 luminescence with DIPEA.

4.2 Crystallography

The data for compound Re-4•0.5THF were collected on a Bruker D8 Quest diffractometer equipped with an Incoatec Microfocus Source (IμS 3.0 Mo) and a PHOTON III area detector, and operated through the APEX3 software, while those for compound Re-5•0.5THF were collected on a Nonius Kappa-CCD area detector diffractometer using graphite-monochromated Mo Kα radiation (λ = 0.71073 Å). The crystals were mounted either into a glass capillary or on a Mitegen micromount with a
protective coating of Paratone-N oil (Hampton Research). The data were processed with SAINT (Re-4) or HKL2000 (Re-5).

Absorption effects were corrected empirically with the programs SADABS (Re-4) or SCALEPACK (Re-5). All structures were solved by intrinsic phasing with SHELXT, expanded by subsequent difference Fourier synthesis and refined by full-matrix least-squares on $F^2$ with SHELXL-2014, using the ShelXle interface. All non-hydrogen atoms were refined with anisotropic displacement parameters. The hydrogen atoms were introduced at calculated positions and they were treated as riding atoms with an isotropic displacement parameter equal to 1.2 times that of the parent atom (1.5 for $\text{CH}_3$, with optimized geometry). The solvent THF molecule has twofold rotation symmetry in Re-4·0.5THF while that in Re-5·0.5THF is disordered around an inversion center and has been given an occupancy parameter of 0.5 accordingly. The molecular plots were drawn with ORTEP-3.

**Crystal data for compound Re-4·0.5THF.** $\text{C}_{17}\text{H}_{16}\text{ClN}_{2}\text{O}_{3.5}\text{Re}$, $M = 525.97$, orthorhombic, space group $Pbcn$, $a = 13.4935(10)$, $b = 22.1756(17)$, $c = 11.2515(8) \text{ Å}$, $V = 3366.8(4) \text{ Å}^3$, $Z = 8$, $D_c = 2.075 \text{ g cm}^{-3}$, $\mu = 7.399 \text{ mm}^{-1}$, $F(000) = 2016$. Refinement of 224 parameters on 3201 independent reflections out of 41305 measured reflections ($R_{int} = 0.051$) led to $R_1 = 0.052$, $wR_2 = 0.114$, $S = 1.506$, $\Delta \rho_{\text{max}} = 1.91 \text{ e Å}^{-3}$.

**Crystal data for compound Re-5·0.5THF.** $\text{C}_{33}\text{H}_{32}\text{ClN}_{2}\text{O}_{3.5}\text{Re}$, $M = 734.25$, monoclinic, space group $C2/c$, $a = 35.2615(16)$, $b = 8.1575(3)$, $c = 22.7433(10) \text{ Å}$, $\beta = 116.166(3)^\circ$, $V = 5871.6(5) \text{ Å}^3$, $Z = 8$, $D_c = 1.661 \text{ g cm}^{-3}$, $\mu = 4.269 \text{ mm}^{-1}$, $F(000) = 2912$. Refinement of 394 parameters on 5550 independent reflections out of 155149 measured reflections ($R_{int} = 0.065$) led to $R_1 = 0.030$, $wR_2 = 0.061$, $S = 1.101$, $\Delta \rho_{\text{min}} = -1.01$, $\Delta \rho_{\text{max}} = 1.66 \text{ e Å}^{-3}$.
4.3 Copies of NMR spectra

4.3.1 Re-4

Figure S15: $^1$H NMR spectrum of Re-4 in CD$_2$Cl$_2$. 
4.3.2 Re-5

Figure S16: $^{13}$C NMR spectrum of Re-4 in CD$_2$Cl$_2$.

Figure S17: $^1$H NMR spectrum of Re-5 in CD$_2$Cl$_2$. 
Figure S18: $^{13}$C NMR spectrum of Re-5 in CD$_2$Cl$_2$.

4.4 Copies of GC traces

4.4.1 Representative GC traces for the photocatalyzed reduction of N$_2$O by Re-3

Figure S19: GC analysis of the gaseous phase after 24 h of irradiation, on NMR scale.
Figure S20: GC analysis of the gaseous phase after 5 h of irradiation, on a 1.2 mmol scale.

4.4.2 GC trace for the reduction of N₂O catalyzed by Re-1 in presence of TEOA

Figure S21: GC analysis of the gaseous phase after 2 h of irradiation.
4.4.3 GC traces for the control experiments for the photoreduction of N$_2$O with Re-1 (see Table 1)

![Figure S22: Control experiment without light.](image1)

![Figure S23: Control experiment without Re-1.](image2)
Figure S24: Control experiment with Ar instead of \( \text{N}_2\text{O} \).

Figure S25: Control experiment without TEOA.
4.5 Copies of UV-Vis spectra

4.5.1 Absorption spectrum of \textbf{Re-4}

\begin{figure}[h]
\centering
\includegraphics[width=\textwidth]{figure_s27.png}
\caption{Absorption spectrum of \textit{Re-4} (175 \textmu M in acetonitrile) at 293 K.}
\end{figure}
4.5.2 Absorption spectrum of Re-5

Figure S28: Absorption spectrum of Re-5 (175 µM in acetonitrile) at 293 K.
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