Pd-Catalyzed Aerobic Oxidative Coupling of Arenes: Evidence for Transmetalation between Two Pd(II)-Aryl Intermediates

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Supporting Information

ABSTRACT: Pd-catalyzed aerobic oxidative coupling of arenes provides efficient access to biaryl compounds. The biaryl product forms via C–H activation of two arenes to afford a PdArAr intermediate, which then undergoes C–C reductive elimination. The key PdArAr intermediate could form via a “monometallic” pathway involving sequential C–H activation at a single PdII center, or via a “bimetallic” pathway involving parallel C–H activation at separate PdII centers, followed by a transmetalation step between two PdII-aryl intermediates. Here, we investigate the oxidative coupling of 3,4,3′-xylene catalyzed by a PdX2/fluoropyridine catalyst (X = trifluoroacetate, acetate). Kinetic studies, H/D exchange experiments, and kinetic isotope effects provide clear support for a bimetallic/transmetalation mechanism.

The oxidative coupling of hydrocarbons (eq 1) is an important contemporary topic with significant commercial implications. Prominent targets include the conversion of methane to higher hydrocarbons and the coupling of arenes to biaryls. The latter application is relevant to the industrial production of heat transfer fluids (e.g., biphenyl) and monomers for high-performance polymers (e.g., Upilex). Moreover, oxidative cross-coupling of arenes represents an efficient route to unsymmetrical biaryl reagents relevant to the pharmaceutical, agrochemical, and fine chemical industries. PdII catalysts show significant promise in oxidative coupling reactions as they are emissarily compatible with O2 as the stoichiometric oxidant and therefore could be accomplished with near-ideal atom economy. Mechanistic understanding of key steps could facilitate the development of new and/or improved reactions of this type. Here, we address unresolved mechanistic questions related to C–H activation and C–C bond formation in the Pd-catalyzed aerobic oxidative homocoupling of o-xylene, a reaction relevant to Upilex production.

R–H + R′–H + [O] \xrightarrow{cat, (R, R′ = alkyl, aryl)} R–R′ + [O]H2

The first example of PdII-mediated oxidative biaryl coupling, reported by van Helden and Verberg in 1965, featured the oxidative coupling of benzene with stoichiometric PdCl2 and NaOAc. Within a few years, several groups reported PdII-catalyzed biaryl coupling methods that used O2 as the oxidant. Related catalytic methods were later developed for the commercial synthesis of 3,4,3′-A′-tetramethyl biphenyltetracarboxylate, a monomeric precursor to Upilex. In subsequent decades, numerous additional examples of Pd-catalyzed methods for oxidative homocoupling of arenes have been reported. The vast majority of work in this area has focused on empirical development of new catalyst systems, but important mechanistic questions remain unanswered.

A reasonable mechanism for Pd-catalyzed oxidative coupling of arenes consists of a PdII/Pd0 catalytic cycle with three general steps (Scheme 1): (i) PdII-mediated activation of two aryl C–H bonds to produce a PdIIIArAr′ intermediate, (ii) reductive elimination of Ar–Ar′ from PdIIIArAr′, and aerobic oxidation of PdII to PdIII. While the mechanism of arene C–H activation by PdII has been the focus of considerable investigation, the pathway for formation of PdIIArAr′ is not clear. At least two different mechanisms are possible. A “monometallic” mechanism (Scheme 1A) involves sequential C–H activation at a single PdII center, while a “bimetallic” mechanism (Scheme 1B) involves parallel C–H activation at two separate PdII centers followed by transmetalation between the two PdII-aryl species.

The “monometallic” mechanism has been commonly invoked in recent reports, especially in the context of Pd-catalyzed oxidative cross-coupling reactions, and C–H activation of arenes by PdIIArX species has been studied experimentally.
The oxidative coupling of \( \text{o-xylene} \) biaryl mechanism in Scheme 1. biaryl coupling, and it provides preliminary support for the observation of a second-order \([\text{Pd}]\) dependence in oxidative dependence data at low \([\text{Pd}]\), shown in Figure 1. H/D exchange the latter observation complements the second-order \([\text{Pd}]\) - biaryl product formation is second-order in \([\text{Pd}]\) (Figure 3B). This reaction proceeds with high selectivity to the symmetrical biaryl product shown in Figure 1. Kinetic studies were carried out to establish the dependence of the reaction rate on \([\text{Pd}]\) (2–50 mM, 0.04–1.0 mol % with respect to \( \text{o-xylene} \)) (Figure 1). Initial rates of the reactions, monitored by gas chromatography, reveal a second-order dependence at low \([\text{Pd}]\) (≤5 mM; inset, Figure 1) and a first-order dependence at high \([\text{Pd}]\) (≥15 mM; Figure 1). To our knowledge, this is the first observation of a second-order \([\text{Pd}]\) dependence in oxidative biaryl coupling, and it provides preliminary support for the bimetallic mechanism in Scheme 1.

Deuterium kinetic isotope effects (KIEs) were evaluated to gain insight into the contribution of C–H activation to the overall turnover rate. The oxidative coupling of \( \text{o-xylene} \) and \( \text{o-xylene-}\text{dimethyl} \) monitored independently with 0.5 mol % \([\text{Pd}]\) in HOAc (Figure 2), reveal a very large KIE: \( k_{2H}/k_{2D} \approx 2–5 \). The reaction was also performed in DOAc to probe solvent isotope effects, and a significant inverse KIE was observed: \( k_{\text{HOAc}}/k_{\text{DOAc}} = 0.31 \pm 0.02 \) (Figure 2).

Further insights into \( \text{Pd}^{0} \)-mediated C–H activation and the origin of the substrate and solvent isotope effects were obtained from H/D exchange experiments. The reaction of \( \text{o-xylene-}\text{dimethyl} \) in HOAc (0.5 mol % \([\text{Pd}]\)) shows that H/D exchange occurs ∼30-fold faster than biaryl product formation (Figure 3A); however, this ratio changes at different \([\text{Pd}]\). For example, H/D exchange is favored over biaryl formation by a ratio of ∼120:1 at 0.1 mol % \([\text{Pd}]\). Plots of H/D exchange and biaryl product yields show that H/D exchange exhibits a first-order dependence on \([\text{Pd}]\) while biaryl product formation is second-order in \([\text{Pd}]\) (Figure 3B). The latter observation complements the second-order \([\text{Pd}]\)-dependence data at low \([\text{Pd}]\), shown in Figure 1. H/D exchange occurs exclusively at aromatic positions and favors exchange at \( \text{meta} \) positions. The 2:1 \( \text{meta/ortho} \) regioselectivity for H/D exchange is substantially lower than the regioselectivity of biaryl product formation, which exhibits ∼10:1 selectivity for \( \text{meta} \text{−}\text{meta} \) over \( \text{ortho} \text{−}\text{ortho} \) coupling (\( \text{ortho} \text{−}\text{ortho} \) coupling is not observed).

The above data led us to undertake a more systematic assessment of kinetic isotope effects as a function of \([\text{Pd}]\). Substrate and solvent KIEs were determined at four different \([\text{Pd}]\) (2.5–67.5 mM, 0.06–1.5 mol % of substrate) (Figure 4; for Supporting Information for individual kinetic plots), and the data reveal that the substrate KIE can vary from 7.9 to 25, depending on the \([\text{Pd}]\) and solvent (HOAc vs DOAc) (Figure 4A). The reciprocal of the solvent KIE (i.e., \( k_{\text{DOAc}}/k_{\text{HOAc}} \)) varies similarly from 3.1 to 23 (Figure 4B). At least three general trends are worth noting. The substrate KIEs are significantly larger in HOAc (KIE = 18–25) than in DOAc (KIE = 7.9–12). The reciprocal solvent KIEs increase significantly as the \([\text{Pd}]\) is lowered and also are substantially larger with \( \text{o-xylene-}\text{dimethyl} \) than with \( \text{o-xylene} \).

Collectively, the kinetic data, KIEs, and H/D-exchange results clearly distinguish between the two possible mechanisms in Scheme 1. Rate laws derived for both mechanisms (eqs 2–4); see...
SI for derivations) show that the monometallic mechanism always exhibits a first-order dependence on [Pd], while the bimetallic mechanism could have a first-order or second-order dependence on [Pd], depending on the identity of the rate-limiting step. The switch from second-order to first-order kinetic behavior as [Pd] is increased (cf. Figure 1) provides support for the bimetallic mechanism. At low [Pd], a lower steady-state concentration of the [L,PdArX] intermediate leads to rate-limiting transmetalation (eq 4). At higher [Pd], the bimetallic transmetalation rate will increase faster than the unimolecular limit. The switch from second-order to first-order dependence on [Pd], depending on the identity of the rate-limiting step (i.e., protonolysis of \( L \mathrm{PdArX} \) is the same for both substrate isotopologs (Scheme 1B and eq 4). Previously reported KIEs for PdII-mediated arene C–H activation vary from 3 to 5, which corresponds to a net KIE of 9–25 when this value is squared. Therefore, the KIEs of 18–25 in Figure 4A are not usual and fit in the range expected for a bimetallic mechanism when transmetalation is rate-limiting.23 Under conditions that lead to rate-limiting C–H activation, the KIE will approach the intrinsic isotope effect associated with C–H activation (i.e., 3–5). We never fully reach this limit under the conditions of our experiments, and the lowest KIE that we observe is 7.9 (Figure 4A).

Inverse solvent isotope effects are evident in all of the reactions described here, and they arise from faster protonolysis of the L,PdArX intermediate by HOAc relative to DOAc. The slower protonolysis with DOAc will lead to a higher steady-state concentration of the L,PdArX intermediate and result in a higher rate of product formation via transmetalation from this species. The inverse solvent KIE will be most substantial when transmetalation is rate-limiting. Under these conditions, C–H activation is reversible, and the isotope effect associated with protonolysis of L,PdArX (\( k_1' \)) is squared in the denominator of eq 4 (because two L,PdArX species are needed for the transmetalation step). This scenario is favored when [Pd] is low and \( o \)-xylene-\( d_{10} \) is the substrate, as evident in Figure 4B. Both conditions reduce the steady-state concentration of L,PdArX.24

This mechanistic study of aerobic oxidative coupling of \( o \)-xylene with the PdX₂ (X = OAc, TFA)/2Fpy catalyst system provides clear evidence for a bimetallic/transmetalation mechanism. Evidence for this pathway includes a bimolecular kinetic dependence on [Pd] (Figure 1) and unusually large KIEs (Figures 2 and 4). The bimetallic mechanism, but not the monometallic mechanism, readily accommodates these observations. The insights obtained from the relatively simple catalytic reaction described here provide an important foundation for future studies of more complex Pd-catalyzed oxidative biaryl coupling reactions, such as those that employ redox-active cocatalysts (e.g., CuII and polyoxometalates), as well as cross-coupling reactions between two different arenes.

### ASSOCIATED CONTENT

* Supporting Information

Experimental procedures, reaction time courses, additional mechanistic experiments and discussion. This material is available free of charge via the Internet at http://pubs.acs.org.
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(23) Transmetallation will be comparatively slow when [Pd] is low and/or protonolysis of the $\text{L}^\text{PdArX}$ intermediate is fast (e.g., in H$\text{OAc}$ vs D$\text{OAc}$; cf. Figure 4A).

(24) See Supporting Information for additional discussion on the KIE result and its consistency with the bimetallic mechanism.