The influence of substituent field and resonance effects on the ease of N-heterocyclic carbene formation from imidazolium rings†

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Using a set of twelve selected substituents, the influence of substituent properties on the ease of deprotonation of imidazolium cations and mesoionic imidazolium-4-olates measured by the CREF index has been investigated. Significant correlations between CREF values and the Swain and Lupton field (F) and resonance (R) substituent constants have been found. In all cases the field effect has the greatest influence but resonance effects are also significant.

Introduction

We have demonstrated that the deprotonation energy (ZPE corrected) of N-heterocyclic carbene (NHC) precursors, formulated as an index, provides a convenient overview of the relative ease of formation of a wide range of normal (nNHC), abnormal (aNHC) and remote (rNHC) neutral and anionic NHCs.† This index, which we describe as the CREF (Carbene Relative Energy of Formation) index, is easily calculated using modern DFT methods. The results direct attention to properties of novel or unexplored molecules of potential interest, including their potential σ-bonding strength as ligands.

In addition, we have observed, in accord with the Hammond postulate,† that there is good correlation between CREF indexes and experimental rate constants for deprotonation or H–D exchange for a range of heterocycles for which kinetic data is available.

Our previous studies have been largely restricted to the influence of ring heteroatoms (N, O and S) on NHC formation. We now report a study of the influence of ring substituents.

Results and discussion

To explore the influence of ring substituents on NHC formation, we selected a set of eleven common substituents that represents a broad spectrum of sigma- and pi-electron-donor and electron-acceptor characteristics. The unsubstituted molecules were also included for reference. These substituents are shown in Table 1 together with their Swain and Lupton field (F) and resonance (R) constants.† Each parameter shows good variation (F 0.0 to +0.65 and R –0.73 to +0.42), and there is low correlation between the two (r 0.56).

Initially we studied the influence of 4-substituents on the formation of the NHCs by deprotonation of the imidazolium cations. The CREF values for these transformations are shown in Table 2 and multiple regression analysis shows a statistically significant relationship (eqn (1), Table 2A) between CREF values and the Swain and Lupton constants. As might be expected, since deprotonation primarily involves heterolytic cleavage of a σ bond, the larger effect is the field effect, but the resonance effect is also significant. Neither parameter alone (F nor R) gives a significant correlation. Using eqn (1) the CREF values of most common organic substituents can be estimated using published F and R values.

The observed substituent effects depend on the relative effects of substituents on reactant and product. Electron-
CREF = 0.416(±0.002) – 0.038(±0.004)F – 0.021(±0.003)R (1),

n = 12, r = 0.98, s = 0.003, F = 134.0, p < 0.0001

CREF = 0.418(±0.004) – 0.069(±0.009)F – 0.028(±0.006)R (2),

n = 12, r = 0.97, s = 0.006, F = 69.1, p < 0.0001

Table 2 CREF index values and correlations for formation of normal NHCs

| Subst. R | Precursor 1 E + ZPEa | Product 2 E + ZPEa | CREF |
|----------|----------------------|-------------------|------|
| a NHMe   | −399.8209            | −399.3944         | 0.427|
| b OMe    | −419.6950            | −419.2753         | 0.420|
| c F      | −404.4263            | −404.0218         | 0.405|
| d Cl     | −764.7876            | −764.3815         | 0.406|
| e Me     | −344.4758            | −344.0567         | 0.419|
| f cycHex | −539.7656            | −539.3424         | 0.423|
| g OCF3   | −717.5314            | −717.1286         | 0.403|
| h H      | −305.1679            | −304.7548         | 0.413|
| i NO2    | −509.6969            | −509.3096         | 0.387|
| j CN     | −397.4147            | −397.0229         | 0.392|
| k CF3    | −642.2949            | −641.8977         | 0.397|
| l NO     | −434.4790            | −434.0888         | 0.390|

| Subst. R | Precursor 3 E + ZPEa | Product 4 E + ZPEa | CREF |
|----------|----------------------|-------------------|------|
| a NHMe   | −494.4616            | −494.0342         | 0.427|
| b OMe    | −534.2104            | −533.7884         | 0.422|
| c F      | −503.6798            | −503.2840         | 0.396|
| d Cl     | −1224.4070           | −1224.0071        | 0.400|
| e Me     | −383.7819            | −383.3579         | 0.424|
| f cycHex | −774.3580            | −773.9273         | 0.431|
| g OCF3   | −1129.8927           | −1129.4988        | 0.394|
| h H      | −305.1679            | −304.7548         | 0.413|
| i NO2    | −714.2134            | −713.8449         | 0.369|
| j CN     | −489.6589            | −489.2844         | 0.374|
| k CF3    | −979.4124            | −979.0276         | 0.385|
| l NO     | −563.7831            | −563.4090         | 0.374|

*a Hartrees.

withdrawing substituents reduce density around neighbouring atomic nuclei, exposing greater nuclear charge, and thus lower σ-orbital energies in both reactant and product.

Since deprotonation effectively involves the transfer of one electron from a hydrogen 1s environment to a carbon σ environment, the lowering of σ energy levels might be expected to not only polarise the C–H bond in the precursor but also stabilise the product, which gains an extra sp2 σ electron. This is consistent with electron-withdrawing substituents, e.g., NO2, CN, CF3, lowering CREF values. Other effects may contribute; substituents that are π donors, e.g., NHMe, OMe, may stabilise the cationic precursors 1 relative to the products 2 contributing to higher CREF values. The origins of the electronic effects of substituents may be complex but eqn (1) conveniently summarises the net effect and there are both σ and π contributions.

Fig. 1 illustrates the effect of three structural changes (cations 5–7) on the CREF index of the imidazolium cation 1h (CREF 0.413). A 4-trifluoromethyl substituent lowers the value by 0.016 units, which is comparable to the effect of 4-aza substitution. Cations 5 and 6 both have acidity comparable to the 1,3-thiazolium cation 7 (Fig. 1). The effect of a 4-nitro substituent (1i, Table 2) is even larger (∆CREF 0.026 units). We have previously estimated that a reduction of the index by 0.01 units leads to a reduction of aqueous pKa of 2.4. On this basis, the estimated pKa of 16.8 for the 4-nitroimidazolium ring is comparable to that of the 1,3-oxazolium cation (pKa 16.9). The associated weakening of the C(2)–H bond in 4-nitroimidazolium rings is consistent with their formation of hydrogen bonds with chloride ions.

The effect of a second substituent is cumulative. An equally significant correlation (eqn (2), Table 2B) was obtained for deprotonation of the disubstituted derivatives 3 giving the NHCs 4. Relative to eqn (1), the coefficients for F and R in eqn (2) are increased by 1.82 and 1.40, respectively. A number of 4,5-disubstituted imidazolium salts with electron-withdrawing substituents have been investigated as potential energetic ionic liquids. Based on the calculated CREF values, the dicyano derivative 3j (CREF 0.374) can be predicted to have pKa ~ 13.6, and to have C–H acidity comparable to 1,4-dialkyltetrazolium salts (CREF 0.372).

Similar statistically significant correlations between CREF index values and the Swain and Lupton constants were found for the formation the substituted NHCs 9 (Table 3A) and 11 (Table 3B). For the deprotonation 8 → 9 the contributions of F

Fig. 1 Three structural changes to the imidazolium cation that result in similar CREF values.
Table 3  CREF index values and correlations for formation of abnormal NHCs

\[ \text{CREF} = 0.445(\pm 0.002) - 0.042(\pm 0.006)F - 0.022(\pm 0.004)R \] \( (3), \) 
\( n = 12, \quad r = 0.97, \quad s = 0.004, \quad F = 72.8, \quad p < 0.0001 \)

| Subst. | Precursor 8 E + ZPE | Product 9 E + ZPE | CREF |
|--------|---------------------|------------------|------|
| a NHMe | -399.8211           | -399.3675        | 0.454|
| b OMe  | -419.6920           | -419.2447        | 0.448|
| c F    | -404.4274           | -403.9925        | 0.435|
| d Cl   | -764.7879           | -764.3522        | 0.436|
| e Me   | -344.4771           | -344.0280        | 0.449|
| f cycHex | -539.7674      | -539.3132        | 0.454|
| g OCF3 | -717.5305           | -717.0970        | 0.434|
| h H    | -305.1679           | -304.7258        | 0.442|
| i NO2  | -509.6897           | -509.2772        | 0.413|
| j CN   | -397.4117           | -396.9924        | 0.419|
| k CF3  | -642.2870           | -641.8610        | 0.426|
| l NO   | -434.4773           | -434.0640        | 0.413|

* Hartrees.

Table 4  CREF index values and correlations for formation of anionic NHCs

\[ \text{CREF} = 0.573(\pm 0.003) - 0.040(\pm 0.008)F - 0.023(\pm 0.005)R \] \( (5), \) 
\( n = 12, \quad r = 0.95, \quad s = 0.005, \quad F = 44.0, \quad p < 0.0001 \)

| Subst. | Precursor 12 E + ZPE | Product 13 E + ZPE | CREF |
|--------|---------------------|------------------|------|
| a NHMe | -474.6740           | -474.0966        | 0.577|
| b OMe  | -494.5473           | -493.9711        | 0.576|
| c F    | -479.2917           | -478.7231        | 0.569|
| d Cl   | -839.6561           | -839.0907        | 0.565|
| e Me   | -419.3327           | -418.7548        | 0.578|
| f cycHex | -614.6214       | -614.0457        | 0.576|
| g OCF3 | -792.4036           | -791.8447        | 0.559|
| h H    | -380.0312           | -379.4554        | 0.576|
| i NO2  | -584.5940           | -584.0559        | 0.538|
| j CN   | -472.3062           | -471.7592        | 0.547|
| k CF3  | -717.1766           | -716.6242        | 0.552|
| l NO   | -509.3800           | -508.8396        | 0.540|

* Hartrees.
and $R$ are similar to those for the formation of the nNHCs 2
(Table 2A). This is consistent with the fact that the relationships
between substituent location and carbene centre are similar.
For the deprotonation $10 \rightarrow 11$ the contribution of $F$ is greater
and that of $R$ is slightly smaller; in this case the substituents are
adjacent to the carbene centre.

We have also investigated the effect of substituents on
anionic NHC formation (Table 4). The CREF index values for
anionic NHCs are higher, not least because the deprotonation
process involves charge separation. Again, significant correla-
tions are observed. For the formation of the nNHCs ($12 \rightarrow 13$
(Table 4A), the correlation between CREF values and the elec-
tronic parameters $F$ and $R$ (eqn (5)) is very similar to that for the
deprotonations $1 \rightarrow 2$ (Table 2A) and $8 \rightarrow 9$ (Table 3A). For
formation of the anionic aNHCs ($14 \rightarrow 15$) (Table 4B), the
contribution of the field effect ($F$) (eqn (6)) is appreciably larger.
The reason for this enhanced field effect is not clear, but it may
be due to the association of position 4 with particularly high
electron density.

**Computational details**

Calculations were performed using the Gaussian 09 program,\(^{11}\)
and the hybrid B3LYP functional,\(^{12,13}\) accompanied by the 6-
311+G(d,p) basis set,\(^{14}\) was used. All geometry optimizations
were followed by frequency calculations to ensure that the
stationary points obtained were true minima on the potential
energy surface and to calculate the zero-point vibrational
corrections (ZPE) to energy. The CREF index was calculated as
the (ZPE-corrected) energy difference between the protonated
precursor and the carbene (obtained via C–H bond heterolytic
cleavage).

The multiple regression relationships and associated statistics
based on the ordinary least squares method were deter-
mined using online software.\(^{15}\) The meaning of the statistical
parameters are the following: $n$ = sample size, $r$ = multiple
regression correlation coefficient, $s$ = residual standard deviation,
$F$ = Fisher test value, $p = p$-value.

**Conclusions**

Statistically significant correlations have been found between
the CREF index, which is a measure of C–H acidity, and elec-
tronic substituent constants for imidazolium cations and
mesoionic imidazolium-4-olates. Both $\sigma$ and $\pi$ contributions to
the acidity are observed. Derivatives disubstituted with electron-
withdrawing groups (e.g., CF$_3$, CN, NO$_2$) are predicted to have
acidity comparable to 1,3-dialkytriazolium cations,\(^8\) but
without the instability associated with rapid loss of nitrogen.\(^{16}\)

The CREF index is a measure of the strength of C–H $\sigma$ bonds
in heteroaromatic rings, and, by analogy, the strength of C–C $\sigma$
bonds (relevant to decarboxylation) and C-metal $\sigma$ bonds, for
example when NHCs act as metal ligands. In the latter case,
there is no measure of C-metal $\pi$ bonding (or steric effects).
In the context of ligand binding, the effectiveness of metal
complexes of the imidazole-ylidenes 4 as catalysts has been
found to vary with ring substituents.\(^{17}\) This has been attributed
to possible variation in $\pi$ bonding between ligand and metal.
This may be the case, but the correlations described in Table 2
show that $\pi$ effects also influence $\sigma$ bond strength and thus
catalytic efficiency.

In 1966 Olofson and Landesberg concluded, justifiably, that
the use of LCAO-MO methods to predict H/D exchange in
heterocyclic cations was futile.\(^{18}\) The DFT approach that we have
explored previously\(^2\) and extended in this study seems more
promising. Experimentally determined physical properties of
heterocycles are limited and there are gaps in the literature.\(^{19}\) It
is hoped that the DFT-based approach described above will be
helpful in understanding variations in physical properties of
heteroaromatic rings and augment experimental data.

**Conflicts of interest**

There are no conflicts to declare.

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