Effects of calcination temperatures on the structure–activity relationship of Ni–La/Al₂O₃ catalysts for syngas methanation†

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A series of Ni–La/Al₂O₃ catalysts for the syngas methanation reaction were prepared by a mechanochemical method and characterized by thermogravimetric analysis (TG-DTA), X-ray fluorescence (XRF), X-ray diffraction (XRD), N₂ adsorption–desorption, H₂ temperature-programmed reduction (H₂-TPR), and X-ray photoelectron spectroscopy (XPS). The calcination temperatures (350–700 °C) had significant impacts on the crystallite sizes and interactions between NiO and Al₂O₃. The catalyst calcined at 400 °C (cat-400) showed a 12.1% Ni dispersion degree and the maximum bound state of NiO (54%) through the Gaussian fitting of H₂-TPR. Cat-400 also achieved the highest CO conversion, CH₄ selectivity and yield. Cat-400 exhibited good stability and catalytic activity in a lifetime testing of 200 h. The deactivation of cat-400 was mainly caused by carbon deposition according to the data from XRD, TG-DTG and XPS.

Introduction

Catalytic methanation has attracted considerable attention since being reported by Sabatier and has been applied in various industrial processes such as the removal of oxo-compounds (COX) from the feed gas for the synthesis of ammonia, and in relation to the Fischer–Tropsch synthesis such as producing methane from synthesis gas. The preparation of natural gas from syngas has become an important factor determining the catalytic activity of catalysts.

Ni-based catalysts are widely used for CO methanation reactions due to their high activity and a competitive cost compared to noble metals. Nevertheless, despite exhibiting advantageous catalytic properties, there are many factors that affect the catalytic activity. Calcination temperature is an important factor determining the catalytic activity of catalysts. Echegoyen et al. investigated the effect of calcination temperature (450, 600, 800 and 1000 °C) on the performance of Ni–Al catalysts for the thermal catalytic decomposition of methane. At the calcination temperature of 600 °C, the Ni–Al catalyst exhibited Ni crystallites of the smallest size (17.9 nm) and the best catalytic activity with a yield of H₂ at 78%. Numpilai et al. prepared Fe–Co/K-Al₂O₃ catalysts for the hydrogenation of CO₂ to light olefins by a two-step incipient wetness impregnation method. The calcination temperature of Fe–Co/K-Al₂O₃ (400–800 °C) significantly impacted the size of metal oxide crystallites, the interaction between Fe₂O₃ and other metal oxides, and the transformation of potassium phases, which further affected CO₂ conversion, product selectivity and the yield of olefins. Dorner et al. investigated the effect of calcination temperature on Fe–Mn–K–Ce catalysts for CO₂ hydrogenation to olefins. Increasing the calcination temperature from 500 to 800 °C caused a drastic reduction in the specific surface area from 200 to 100 m² g⁻¹. Compared with the catalyst calcined at 400 °C, the catalyst calcined at 800 °C exhibited a 5% and 2.9% increase in olefin/paraffin (O/P) ratio and CO₂ conversion, respectively.

The materials prepared by the mechanochemical method can be regarded as homogeneous materials due to the highly dispersed components, the concentrated distribution of pore size and the specific surface area of over 100 m² g⁻¹. These materials exhibited excellent catalytic performances, such as high activity, good selectivity, and long lifetimes. Therefore,
the mechnochemical method was gradually applied in the preparation of catalytic materials. Using the mechnochemical method, Xing et al.\textsuperscript{21} prepared a visible light photocatalyst, N-doped nano-TiO\textsubscript{2}, which can degrade organic matter in aqueous solution. Kvet\u010an et al.\textsuperscript{22} prepared a Ni/Mo catalyst loaded on Al/Ce composite oxides and applied it in the hydro-desulfurization of 1-benzothiophene. Due to the promising applications of the mechnochemical method, it is necessary to study the effect of calcination temperature on the structure and catalytic performance of Ni-based catalysts for CO methanation in a slurry-bed reactor. In this work, a series of Ni–La/Al\textsubscript{2}O\textsubscript{3} catalysts were prepared by the mechnochemical method for CO methanation, and the influence of calcination temperature (350–700 °C) on the structure–activity relationship of the Ni–La/Al\textsubscript{2}O\textsubscript{3} catalyst for the syngas methanation reaction in a slurry-bed reactor was investigated in detail.

**Experimental**

**Material synthesis**

A ND7-2L planetary ball mill was used as the chemical reactor, and the grinding ball diameter was 6 mm with a ball/material quality ratio of 2 : 1. The reactants, 15 mmol Ni(NO\textsubscript{3})\textsubscript{2}·6H\textsubscript{2}O, 98 mmol Al(NO\textsubscript{3})\textsubscript{3}·9H\textsubscript{2}O, 160 mmol (NH\textsubscript{4})\textsubscript{2}CO\textsubscript{3}, and 1.3 mmol La(NO\textsubscript{3})\textsubscript{3}·6H\textsubscript{2}O were placed into the grinding jar and milled for 1 h with positive and negative alternating rotation. The mixture was dried at 100 °C for 10 h to obtain the precursor, which was further calcined in air for 4 h at 350, 400, 500, 600 and 700 °C, respectively. The calcined samples were screened to 80–120 mesh and placed in a fixed-bed reactor for reduction for 6 h at 850 °C with a flow of H\textsubscript{2} at 40 mL min\textsuperscript{−1} to acquire the reduced Ni–La/Al\textsubscript{2}O\textsubscript{3} catalysts. The reduced samples were screened to 60–80 mesh for evaluation. The chemical reactions that occurred in the process of the preparation of the catalyst are as follows:

\[
\begin{align*}
\text{(NH}_4\text{)}_2\text{CO}_3 + \text{Ni(NO}_3\text{)}_2·6\text{H}_2\text{O} & = \text{NiCO}_3 + 6\text{H}_2\text{O} + 2\text{NH}_4\text{NO}_3 \\
3\text{(NH}_4\text{)}_2\text{CO}_3 + 2\text{Al(NO}_3\text{)}_3·9\text{H}_2\text{O} & = 2\text{Al(OH)}_3 + 6\text{NH}_4\text{NO}_3 + 15\text{H}_2\text{O} + 3\text{CO}_2↑ \\
3\text{(NH}_4\text{)}_2\text{CO}_3 + 2\text{La(NO}_3\text{)}_3·6\text{H}_2\text{O} & = 2\text{La(OH)}_3 + 6\text{NH}_4\text{NO}_3 + 9\text{H}_2\text{O} + 3\text{CO}_2↑
\end{align*}
\]

**Material characterization**

Thermogravimetric analysis (TGA) was conducted on the sample precursor and the carbon deposited after the life test (Seiko Instruments EXSTAR TG/DTA 6300) in air with a flow rate of 100 mL min\textsuperscript{−1} and a heating rate of 5 °C min\textsuperscript{−1}. The Ni and La contents were determined by X-ray fluorescence (XRF) analysis performed on a 2424XRF instrument (Rh target, X-ray tube maximum power 2.4 kW), X-ray diffraction (XRD, Bruker D-8 Advance dillfractometer) was used to examine the phase and crystallinity of the samples over the 2θ range from 10 to 80° using a Cu-K\textsubscript{\alpha} radiation source (\(\lambda = 0.154060\) nm) with a step scan of 0.02°. The BET surface areas of the samples were determined by N\textsubscript{2} adsorption–desorption measurements using a Micromeritics ASAP 2020 instrument at 77 K. Prior to analysis, the samples were degassed under vacuum for 30 min at ambient temperature followed by fast-mode degassing at 300 °C for 12 h. The reducibility of the catalysts was examined by H\textsubscript{2} temperature-programmed reduction (H\textsubscript{2}-TPR) using a continuous-flow tube reactor. For each analysis, 100 mg of catalyst was packed in a quartz tube and a reducing gas containing 9.6% H\textsubscript{2} in argon flowed into the reactor at a rate of 30 mL min\textsuperscript{−1}. The adsorption properties (TPD) of the as-prepared catalysts were determined using a BELCAT-II-T-SP characterization system. CO chemical adsorption was used to determine the dispersion and surface area of metallic Ni in the sample. The chemical composition and elemental state of samples were determined by X-ray photoelectron spectroscopy (XPS, Thermo Scientific ESCALab 250Xi with 200 W monochromatic Al K\textsubscript{\alpha} radiation). The obtained binding energies were corrected with reference to C 1s (284.8 eV).

**Catalytic performances**

To determine the catalytic performance, 4.50 g of catalyst and 250 mL liquid paraffin (boiling range of 280–320 °C) were placed in a 1.0 L slurry-bed reactor with a rotating speed of 750 rpm. After the reaction, the mixture gas, including products, was passed through a condenser and were separated in a gas–liquid separator. Water was removed from the bottom of the separator and all of the H\textsubscript{2}, CH\textsubscript{4}, CO\textsubscript{2}, and CO\textsubscript{2} gases in the gas mixture from the top of the separator were quantitatively analyzed by an online gas chromatography system (3000A; Agilent Technologies; TCD) as shown in Fig. 1.

**Results and discussion**

**Selection of calcination temperature**

The calcination temperature range for the precursor was 350–700 °C. The influence of calcination temperature on the structure and catalytic performance of Ni-based catalysts for CO methanation was investigated in detail. This was done using TG/DTA analysis. Fig. 2 gives the TG/DTA profiles of the precursor, indicating that the weight loss process was divided into three phases. The low-temperature phase of 100–200 °C was ascribed to the evaporation of physically adsorbed water with a weight loss of 10.13%. The mid-temperature phase was from 200 to 400 °C with a weight loss of 73.31%. Al(OH)\textsubscript{3}, NiCO\textsubscript{3}, NH\textsubscript{4}NO\textsubscript{3} and La(OH)\textsubscript{3} decomposed
to Al2O3 and H2O, NiO and CO2, NH3 and NO2, and La2O3 and H2O, respectively, because the decomposition temperatures ranges for NiCO3/La(OH)3, Al(OH)3 and NH4NO3 were 200–300, about 300 °C and 230 °C, respectively. There was almost no weight loss in the high-temperature phase of 400–1000 °C.

Thus, 350, 400, 500, 600 and 700 °C were selected as the calcination temperatures, and the corresponding samples were named cat-350, cat-400, cat-500, cat-600, and cat-700, respectively.

**Evaluation of the catalytic performance**

The evaluation experiments for the five reduced samples were carried out on the device shown in Fig. 1. Experimental conditions were as follows: space velocity = 3000 mL (g h)−1, stirring rate = 750 rpm, temperature = 280 °C, pressure = 1.0 MPa, time = 10 hours, and H2/CO = 3.1 : 1 (molar). The main reactions of the methanation process were as follows:

**CO methanation reaction**

\[ \text{CO} + 3\text{H}_2 = \text{CH}_4 + 2\text{H}_2\text{O} \quad (\Delta H = -206 \text{ kJ mol}^{-1}) \]  \hspace{1cm} (1)

**Water–gas shift reaction**

\[ \text{CO} + \text{H}_2\text{O} = \text{CO}_2 + \text{H}_2 \quad (\Delta H = -41 \text{ kJ mol}^{-1}) \]  \hspace{1cm} (2)

**CO disproportionation reaction**

\[ 2\text{CO} = \text{C} + \text{CO}_2 \quad (\Delta H = 17 \text{ kJ mol}^{-1}) \]  \hspace{1cm} (3)

**CO2 methanation reaction**

\[ \text{CO}_2 + 4\text{H}_2 = \text{CH}_4 + 2\text{H}_2\text{O} \quad (\Delta H = -165 \text{ kJ mol}^{-1}) \]  \hspace{1cm} (4)

Fig. 3 shows the effects of calcination temperature on CO conversion (XCO), CH4 selectivity (SCH4), and CH4 yield (YCH4). As shown in Fig. 3(a)–(c), the curves of XCO, SCH4, and YCH4 for all the samples may be divided into two stages, namely, the initial phase (about 1 h) and stable phase (9 h, except for cat-350). XCO increased by 5 to 10% as shown in Fig. 3(a). For example, in the initial phase, XCO increase from 88.0% to 92.1, 94.0, 91.3 and 88.4% based on cat-350, cat-400, cat-500, and cat-600, respectively, and from 68.9% to 75.8% based on cat-700. In the stable phase, XCO, of cat-400, cat-500, cat-600 and cat-700 was kept at 94.3, 92.0, 89.1 and 73.3%, respectively, and only XCO of cat-350 decreased rapidly from 92.1% to about 80% before 6 hours.

In Fig. 3(b), SCH4 of cat-400 maintained the best value of 94.9% within 10 h, while that of the others was maintained between 84.7 and 89.6%. In Fig. 3(c), YCH4 of samples is ranked as follows: 88.4% for cat-400 > 81.6% for cat-500 > 77.6% for cat-600 > 68.3% for cat-350 > 65.80% for cat-700. Fig. 3(d) gives the average values of XCO, SCH4 and YCH4 during 10 h for the five samples, in which the curves for catalytic performance show an open downward parabolic tendency with the increase in the calcination temperature. XCO, SCH4 and YCH4 of cat-400 reached the maximum of 94.0, 94.9 and 88.4%, respectively. Therefore, 400 °C was chosen as the optimum calcination temperature.

Table 1 shows the theoretical and experimental contents and loadings of Ni and La in calcined samples. As shown, the Ni and La loadings were in the range of 93.80–95.13% and 91.0–92.5%, respectively, which almost reached the theoretical value.

![Fig. 2 TG/DTG profiles of the precursor.](image)

![Fig. 3 Influence of calcination temperature on CO methanation properties of the samples. (a) CO conversion, (b) CH4 selectively, (c) CH4 yield, (d) average of parameters.](image)

| Sample  | Theor. value | Exper. value | Load | Theor. value | Exper. value | Load |
|---------|--------------|--------------|------|--------------|--------------|------|
| Cat-350 | 15           | 14.32        | 2    | 1.85         | 92.5         | 92.5 |
| Cat-400 | 14.27        | 95.13        | 1.84 | 92.0         | 91.5         | 91.0 |
| Cat-500 | 14.16        | 94.27        | 1.83 | 91.0         | 91.0         | 91.0 |
| Cat-600 | 14.09        | 93.93        | 1.82 | 91.0         | 91.0         | 91.0 |
| Cat-700 | 14.07        | 93.80        | 1.82 | 91.0         | 91.0         | 91.0 |

*Theor. value: theoretical value, exper. value: experimental value.*
in the sequence of black, grayish-black, gray, gray-green and green, indicating that the calcination temperature deeply affected the types of NiO. According to the peak position of H$_2$-TPR profiles, the reducible NiO species could be divided into three types: \( \alpha \)-NiO (300–500 °C), \( \beta \)-NiO (500–800 °C), and \( \gamma \)-NiO (800–1000 °C). The \( \alpha \)-NiO reduction peak was attributed to the reduction of free NiO, which had almost no interaction with Al$_2$O$_3$. \( \beta \)-NiO reduction peak was generated from the reduction of bound NiO and there was a strong interaction between NiO and Al$_2$O$_3$. The \( \gamma \)-NiO reduction peak corresponds to the reduction of the nickel-aluminum spinel (NiAl$_2$O$_4$). NiO and Al$_2$O$_3$ were due to the interactions of NiAl$_2$O$_4$. Furthermore, there were \( \alpha \)-NiO, \( \beta \)-NiO and \( \gamma \)-NiO reduction peaks for both cat-350 and cat-400 samples, whereas the \( \alpha \)-NiO reduction peak disappeared, and the peak temperature of \( \beta \)-NiO and \( \gamma \)-NiO increased significantly for other samples. In Fig. 4(b), for the five samples, temperatures of the three reduction peaks increased, that is, the temperatures of the \( \alpha \)-peak, the \( \beta \)-peak and the \( \gamma \)-NiO increased from 402 to 411 °C, 647 to 766 °C, and 783 to 860 °C, respectively. Thus, the calcination temperature influenced the types of NiO present. Secondly, the relative content of \( \alpha \)-NiO decreased from 39.0% to 0 when the calcination temperature increased from 350 to 500 °C, whereas the \( \alpha \)-NiO content of cat-400 appeared as a parabolic curve with a downward opening, which reached the maximum at 54%. As such, 400 °C was the optimum calcination temperature. Thirdly, for cat-400, \( \alpha \)-NiO, \( \beta \)-NiO and \( \gamma \)-NiO were present and their relative contents were 18.8%, 54.0% and 27.2%, respectively.

In summary, calcination temperature affects the types of NiO present in the samples. The reason comes from the following two aspects. The first aspect is the three existing types of NiO. The influence degree order is bound NiO \( \gg \) free NiO \( \gg \) combined NiO, which is consistent with literature. The second is the relative content ratio of the three existing types and a reasonable ratio is 1 : 3 : 2.

Fig. 5 shows the XPS spectra of all samples. According to Fig. 5(a), O, La, Ni and Al species on the surface of the samples show similar O 1s, La 3d, Ni 2p and both Al 2s and Al 2p peaks, respectively. According to Fig. 5(b), the position of the Ni 2p$_{3/2}$ peak is in the range of 855.96–863.08 eV, indicating that there are two bands, where the first one located at 855.96–857.16 eV is associated with NiO exhibiting weak interactions between NiO and Al$_2$O$_3$. The second one is the bound NiO in NiAl$_2$O$_4$.
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the second one located at 861.69–863.08 eV is related to Ni$^{2+}$ in the formation of NiAl$_2$O$_4$ with strong interaction between NiO and Al$_2$O$_3$. The binding energy values of Ni 2p$_{3/2}$ for cat-350, cat-400, and cat-500 were at 855.96 and 861.69 eV, which were lower than those of 857.16 and 863.08 eV for cat-600 and cat-700. Interestingly, this observation suggests the existence of strong interactions between NiO and Al$_2$O$_3$ of NiAl$_2$O$_4$ while the calcination temperature changed from 350 $^\circ$C to 700 $^\circ$C. This result is in agreement with that of H$_2$-TPR.

Fig. 6 shows the XRD patterns of calcined (a) and reduced samples (b). As shown in Fig. 6(a), for the calcined samples, the characteristic diffraction peaks at 2$\theta$ = 37.5°, 45.5° and 66.4° are attributed to γ-Al$_2$O$_3$ (JPCDS 10-0425) and NiAl$_2$O$_4$ (JPCDS 10-0339). The diffraction peaks became stronger and sharper as the calcination temperature increased. This is consistent with the H$_2$-TPR results that when the calcination temperature increased from 350 $^\circ$C to 700 $^\circ$C, the relative content of NiAl$_2$O$_4$ was also enhanced from 24.1 to 67.8% because of the transition of NiO from the free-type to bound-type to combined type. However, the characteristic diffraction peaks of La$_2$O$_3$ and NiO (JPCDS 73-1519) at 2$\theta$ = 43.3°, 62.8° and 75.4° did not appear, which means that the NiO and La$_2$O$_3$ species were highly dispersed.

Furthermore, according to Fig. 6(b), there were new characteristic diffraction peaks attributed to Ni (JPCDS 04-0850) at 2$\theta$ = 44.5°, 51.8° and 76.4° for all samples reduced at 800 $^\circ$C for 3 h, indicating that NiAl$_2$O$_4$ was reduced to Ni, and the Ni crystallites were well dispersed on Al$_2$O$_3$.

Fig. 7 N$_2$ absorption–desorption isotherms (a) and pore size distributions (b) of samples.

Fig. 8 The influence of calcination temperature on the pore structure of the calcined samples.
Fig. S1† shows the effect of calcination temperature on the Ni grain size of the reduced samples according to Scherrer's formula (\(D = 0.89 \lambda/\beta \cos \theta\)) based on the peak of Ni at \(2\theta = 51.8^\circ\) in Fig. 6(b). Fig. S1† shows that the Ni grain size of cat-400 reached a minimum (7.08 nm), and those of cat-500 and cat-600 were less than 8.0 nm, whereas those of the remainder were more than 9.0 nm. Thus, combined with the results of H\(_2\)-TPR, it has been shown that the types of NiO can significantly affect the Ni grain size. This is why cat-400 showed the maximum values of \(X_{\text{CO}}\), \(S_{\text{CH}_4}\), and \(Y_{\text{CH}_4}\) in the 10 h evaluation experiments.\(^{25,27}\)

\(\text{N}_2\) adsorption–desorption isotherms and pore size distributions of the calcined samples are shown in Fig. 7(a) and (b). According to the IUPAC classification method, all samples exhibited type IV isotherms with hysteresis loops under relative pressure of \((p/p_0) \leq 0.4\)–0.45. As shown in Fig. 7(a), when the relative pressure of \(p/p_0\) was more than 0.4–0.45, the \(\text{N}_2\) adsorption amount increased rapidly due to capillary condensation, causing hysteresis loops to appear,\(^{33–35}\) indicating the presence of mesoporous structures in the five samples. However, there were significant differences in the hysteresis loop shapes of the five samples. Cat-350 and cat-400 showed type H2 isotherms, indicating that they were agglomerated with uniform size and narrow pore-size distribution, and the \(p/p_0\) range was 0.4–0.7. The maximum adsorption amount was due to saturation adsorption, and the pore size distributions were almost mesopores (2–9 nm). Cat-500, cat-600 and cat-700 all showed obvious H3 type hysteresis loops, which mean that the pore structures were slits formed by the stacking of tabular particles, and the \(p/p_0\) range was 0.4–1.0 without saturated adsorption capacity.

In Fig. 7(b), it can be seen that the pore-size distribution of the three samples was between 2 and 100 nm. Furthermore, the initial relative pressure of the hysteresis loops gradually increased from 0.4 to 0.45 with the increase in the calcination temperature. The relative pressure of capillary condensation occurring in the samples’ mesoporous holes gradually increased, indicating that the mesoporous diameters of the samples gradually increased.\(^{28,36}\) In summary, the calcination temperature has significant effects on the characteristics of the pore structures in the samples, and when the calcination temperature was 350–400 °C, the pore size was distributed within the mesopores.

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As shown in Fig. S2, the effects of calcination temperature on the dispersion, surface area and particle diameter of Ni were investigated through CO chemisorption testing of the reduced samples. The Ni dispersion and surface area of cat-400 were 11.3% and 10.73 m² g⁻¹, respectively. The dispersion and surface area of Ni of cat-700 were only 7.2% and 5.97 m² g⁻¹, and those of cat-350 were 8.7% and 8.46 m² g⁻¹, respectively, which coincided with the results shown in Fig. S1. Obviously, the larger dispersion and surface area of Ni can provide more active sites, and show better catalytic performance. The sizes of Ni crystallites in cat-350 were large but the dispersion degree was low. The size of Ni grains of cat-400 were the smallest (8.77 nm), while those of cat-700 and cat-350 were 12.54 nm and 11.24 nm, respectively. Compared to the size calculated by Scherrer’s formula, the difference between the two results was less than 15%. Thus, the analysis results in Fig. S1 and S2 are reliable. The pore size distribution is in the range of 3.9–7.3 nm as exhibited in Fig. 8, indicating the presence of mesoporous structures in all catalysts. The BET specific surface area of the catalysts decreases from 372.8 m² g⁻¹ to 243.6 m² g⁻¹ with the increased temperature as listed in Fig. 8.

Fig. 9 shows the pore structure parameters before and after the CO methanation reaction. The average pore diameter of the five samples increased by 2–3 nm after the reaction and the specific surface area reduced to 100–200 m² g⁻¹ after the reaction. This was mainly due to the collision between particles while the catalyst was scoured with the rapid liquid paraffin flow in the slurry-bed reactor. The formation of carbon deposits from the CO disproportionation, shown in eqn (3), may make these problems more serious. However, compared to the 10 h evaluation results, the catalytic activities of all samples did not exhibit a declining trend, except for cat-350. This means that in terms of the reaction, the average pore diameter of the catalysts changed within the nanoscale, and the change in the specific surface area was within a certain range, which had little effect on its activity, indicating that the pore structure parameters of the samples were not the sensitive factors affecting the activity.

The lifetime of a catalyst is a key consideration in the economical production of syngas methanation. Deposited carbon, an intermediate product during methanation, can lead to the deactivation of the catalyst. Fig. 10 exhibits the catalytic performance of cat-400 at 238 °C and 0.2 MPa within 200 h. The catalytic activity was stable within the first 25 h, that is, X_{CO}, Y_{CH₄}, and S_{CH₄} remained at 94.0, 89.1, and 94.9%, respectively. These values are consistent with the results of the evaluation experiment over 10 h. Afterwards, within 175 h, X_{CO} and Y_{CH₄} dropped to 76.8% and 72.2% with the decreasing rates of 0.1% and 0.11% per h, respectively, and S_{CH₄} was basically stable at 93%.

Fig. 11 shows the TG-DTG curves (a) and XRD patterns (b) of cat-400 after a lifetime test of 200 h. There are two peaks for the loss of weight in Fig. 11(a). The temperature of the first one was about 100 °C, corresponding to the weight loss of water, and the temperature of the second was 290 °C due to the oxidative weight loss of amorphous carbon. This indicates that carbon deposition occurs during the 200 h reaction because of the CO disproportionation during the 200 h reaction. The amorphous carbon types mainly formed because the reaction temperature was low, only about 280 °C. The weight loss at 200–1000 °C, i.e., the carbon deposition quantity, was 7.58%, and the carbon deposition rate was about 0.38 mg (h g_{cat})⁻¹. Fig. 11(b) shows the XRD patterns of cat-400 before and after the lifetime test. After the stability evaluation test, there was no change in

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Fig. 12 Comparison of survey XPS spectra before (a) and after (b) a lifetime test of 200 h and C 1s spectrum-wide scans of reduced cat-400 before (c) and after (d) the lifetime test.
the intensity of the Al2O3 diffraction peak in cat-400. However, the intensity of the Ni diffraction peak was weakened and diffused as compared with the fresh cat-400. Combining these results from the TG-DTG analysis, we conclude that the decrease in the catalytic activity was not due to Ni sintering but was caused by the amorphous carbon from the CO disproportionation reaction. There is no carbon diffraction peak in Fig. 11(b), which further confirms the surface coating by amorphous carbon.

Fig. 12 shows the XPS survey spectra before (a) and after (b) the lifetime test over 200 h. The C 1s spectrum-wide scans of cat-400 before (c) and after (d) the lifetime test for 200 h are also shown. On comparing Fig. 12(a) and (b), it can be seen that the peak location of each element near the surface of cat-400 was almost the same as the binding energy: 853.0 eV for Ni 2p, 835.6 eV for La 3d5, 531.2 eV for O 1s, 288.4 eV (C–O) peak, 284.8 eV (C–C) and 292.6 eV (C=O) for C 1s,

| Sample | Carbonous type | Binding energy (eV) | \( n_\text{c} \) (%) |
|--------|----------------|---------------------|---------------------|
| Fresh  | C 1s C–O       | 288.4               | 65.2                |
|        | C 1s C–C       | 284.8               | 23.9                |
|        | C 1s C=O       | 292.6               | 10.9                |
| Spent  | C 1s C–O       | 286.8               | 12.7                |
|        | C 1s C–C       | 284.8               | 76.3                |
|        | C 1s C=O       | 289.2               | 11.0                |

Table 2  
C 1s peak fitting XPS data for reduced cat-400 before and after the lifetime test of 200 h

There were three kinds of combined carbon species near the surface of the reduced cat-400 sample as shown in Fig. 12(c) and (d) after the stability evaluation experiment. The (285 ± 0.3) eV peak corresponds to the amorphous carbon sp3 hybridization, the (288 ± 0.3) eV peak was assigned to a C–O single bond, and peaks at energies higher than 289 eV were ascribed to the double bond of carbon–oxygen. Fitting (c) and (d) provided C 1s peak data for cat-400 before and after the lifetime test of 200 h, shown in Table 2.

Conclusions

The effects of calcination temperature on the structure–activity relationship of the Ni-La/Al2O3 catalyst prepared by the mechanochemical method for the syngas methanation reaction were comprehensively investigated. The interactions between NiO and Al2O3 support increased when the temperature increased from 350 °C to 700 °C because the relative content of free-state NiO decreased, whereas the relative contents of the bound and combined states of NiO increased. Cat-400 showed the highest catalytic activity, with the specific surface area of 372.8 m² g⁻¹ and the smallest (7.08 nm) Ni⁰ crystalline grain size, which reduced the internal diffusion resistance and offered more reactive sites. Cat-400 exhibited good stability and catalytic activity in the lifetime test of 200 h. From the TG-DTG, XRD and XPS analyses, the decrease in the catalytic activity of cat-400 resulted mainly from the deposition of amorphous carbon on the catalyst surface via CO disproportionation. This work not only shows the importance of calcination temperature, which can determine the physicochemical properties of catalyst and the catalytic performance on methanation of syngas, but it can also be used as a reference for material synthesis for other reactions.

Conflicts of interest

There is no conflicts to declare.

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