ExoMol molecular line lists V: the ro-vibrational spectra of NaCl and KCl

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ABSTRACT
Accurate rotation–vibration line lists for two molecules, NaCl and KCl, in their ground electronic states are presented. These line lists are suitable for temperatures relevant to exoplanetary atmospheres and cool stars (up to 3000 K). Isotopologues 21Na35Cl, 23Na37Cl, 39K35Cl, 39K37Cl, 41K35Cl and 41K37Cl are considered. Laboratory data were used to refine ab initio potential energy curves in order to compute accurate ro-vibrational energy levels. Einstein A coefficients are generated using newly determined ab initio dipole moment curves calculated using the CCSD(T) method. New Dunham Yj constants for KCl are generated by a re-analysis of a published Fourier transform infrared emission spectra. Partition functions plus full line lists of ro-vibration transitions are made available in an electronic form as supplementary data to this paper and at www.exomol.com.

Key words: molecular data – opacity – astronomical data bases: miscellaneous – planets and satellites: atmospheres – stars: low-mass.

1 INTRODUCTION
NaCl and KCl are important astrophysical species as they are simple, stable molecules containing atoms of relatively high cosmic abundance. Na, K and Cl are the 15th, 20th and 19th most abundant elements in the interstellar medium (Caris et al. 2004). In fact, NaCl could be as abundant as the widely observed SiO molecule (Cernicharo & Guelin 1987). Cernicharo & Guelin (1987) reported the first detection of metal halides, NaCl, KCl, AlCl and, more tentatively, AIF, in the circumstellar envelope of carbon star IRC+10216. These observations have been followed up recently by Agundez et al. (2012), who also observed CS, SiO, SiS and NaCN. NaCl has also been detected in the circumstellar envelopes of oxygen stars IK Tauri and VY Caris Majoris (Milam et al. 2007). Another environment in which these molecules have been found is the tenuous atmosphere of Jupiter’s moon Io. Submillimetre lines of NaCl, and more tentatively KCl, were detected by Lellouch et al. (2003) and Mouillet et al. (2013), respectively. NaCl has also been identified in the cryovolcanic plumes of Saturn’s moon Enceladus alongside its constituents Na and Cl (Postberg et al. 2011). K was also detected but the presence of KCl could not be confirmed. Furthermore, NaCl and KCl are expected to be present in super-Earth atmospheres (Schaefer, Lodders & Fegley 2012) and may form in the observable atmosphere of the known object GJ1214b (Kreidberg et al. 2014). It appears that the only theoretical transition line lists for these molecules are catalogued in the Cologne Database for Molecular Spectroscopy (CDMS; see Muller et al. 2005). They were constructed using data reported in Caris et al. (2002), Clouser & Gordy (1964), Uehara et al. (1989) and Leeuw et al. (1970) for NaCl, and Caris et al. (2004), Clouser & Gordy (1964) and Wachem & Dymanus (1967) for KCl. The lists are limited to v = 4, J = 159 and do not include a list for 41K37Cl. In this paper, we aim to compute more comprehensive line lists for the previously studied isotopologues and the first theoretical line list for 41K37Cl.

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IK Tauri and VY Caris Majoris (Milam et al. 2007) and may form in the observable atmosphere of GJ1214b (Kreidberg et al. 2014). Therefore, it is important to monitor their concentrations, which can be done spectroscopically provided the appropriate data are available.

The importance of NaCl and KCl spectra has motivated a number of laboratory studies, for example Rice & Klemperer (1957), Honig et al. (1954), Horiai et al. (1988), Uehara et al. (1989, 1990) and Clouser & Gordy (1964). The most recent and extensive research on KCl and NaCl spectra has been performed by Ram et al. (1997), who investigated infrared emission lines of Na35Cl, Na37Cl and 39K35Cl, Caris et al. (2004), who measured microwave and millimetre wave lines of 39K35Cl, 39K37Cl, 41K35Cl, 41K37Cl and 40K36Cl, and Caris, Lewen & Winnewisser (2002), who recorded microwave and millimetre wave lines of Na35Cl and Na37Cl.

Dipole moment measurements have been carried out by Leeuw, Wachem & Dymanus (1970) for Na35Cl and Na37Cl, Wachem & Dymanus (1967) for 39K35Cl and 39K37Cl, and Hebert et al. (1968) for 39K35Cl, Na35Cl and Na37Cl.

It should be noted that the only theoretical transition line lists for these molecules are catalogued in the Cologne Database for Molecular Spectroscopy (CDMS; see Muller et al. 2005). They were constructed using data reported in Caris et al. (2002), Clouser & Gordy (1964), Uehara et al. (1989) and Leeuw et al. (1970) for NaCl, and Caris et al. (2004), Clouser & Gordy (1964) and Wachem & Dymanus (1967) for KCl. The lists are limited to v = 4, J = 159 and do not include a list for 41K37Cl. In this paper, we aim to compute more comprehensive line lists for the previously studied isotopologues and the first theoretical line list for 41K37Cl.
The ExoMol project aims to provide line lists for all the molecular transitions of importance in the atmospheres of exoplanets. The aims, scope and methodology of the project have been summarized by Tennyson & Yurchenko (2012). Lines lists for \(X^1\Sigma^+\) \(\text{XH}\) molecules, \(\text{X} = \text{Be, Mg, Ca, and } \text{X}'\Sigma^+\) \(\text{SiO}\) have already been published (Yadin et al. 2012; Barton, Yurchenko & Ten- nyson 2013, respectively). In this paper, we present ro-vibrational transition lists and associated spectra for two \(\text{NaCl}\) and four \(\text{KCl}\) isotopologues.

2 METHOD

The nuclear motion Schrödinger equation allowing for Born–Oppenheimer breakdown (BOB) effects is solved for species \(\text{XCl}\) using the program \textsc{levell} (Le Roy 2007). To initiate these calculations, the program \textsc{dpotfit} (Le Roy 2006) was used to generate a refined potential energy curve (PEC) for each molecule by fitting analytic PEC functions derived from ab initio points to laboratory data.

2.1 Spectroscopic data

The most comprehensive and accurate sets of available laboratory measurements are the infrared ro-vibrational emission lines of Ram et al. (1997) and the microwave rotational lines of Caris et al. (2002) and Caris et al. (2004) all of which were recorded at temperatures in the region of 1000 C, see Table 1. No electronic transition data appear to be available. For \(\text{KCl}\) Fourier transform, infrared emission spectra measured by Ram et al. (1997) have been re-analysed and re-assigned as part of this work, see Section 2.2. The Dunham constants \(Y_i\) obtained from this new fit are provided in Table 2. Our new assignments for the ro-vibrational emission lines were used in place of those given by Ram et al. (1997).

2.2 Re-analysis of the KCl infrared spectrum

Ram et al. (1997) reported spectroscopic constants derived from an infrared emission spectrum of \(\text{KCl}\) recorded with a high-resolution Fourier transform spectrometer (FTS). By using the new constants derived from the millimetre wave spectrum by Caris et al. (2004) to simulate the infrared spectrum of \(39\text{K}^{35}\text{Cl}\) with \textsc{pgopher} (Western 2013), it was clear that Ram et al. (1997) had misassigned much of the complex spectrum. The Ram et al. (1997) spectrum was therefore re-analysed. As a first step, the millimetre wave line list from Caris et al. (2004) was refitted with the addition of two Dunham parameters, \(Y_23\) and \(Y_41\). These parameters were found to improve the quality of the fit. The Caris et al. (2004) constants plus \(Y_{23}\) and \(Y_{41}\) were then used to calculate band constants used as input for \textsc{pgopher}. Using \textsc{pgopher}, the infrared line positions were selected manually and then refitted along with the Caris data using our LSQ fit program. There were 253 \textit{R}-branch lines of \(39\text{K}^{35}\text{Cl}\) fit from the 6–5, 5–4, 4–3, 3–2, 2–1 and 1–0 bands, and the \(Y_{10}, Y_{20}\) and \(Y_{30}\) vibrational constants were added. The quality of the observed spectrum was insufficient to fit additional bands or \(P\)-branch lines. The final constants from our global fit are compared to the values reported by Caris et al. (2004) in Table 2. The \(Y_{10}\) and \(Y_{20}\) \((\omega_v\) and \(-\omega_v\Sigma_v^+\)) constants of Caris et al. (2004) were derived entirely from millimetre wave data using Dunham relationships and are in good agreement with the values we have determined directly from infrared observations.
2.3 Dipole moments

Experimental measurements of the permanent dipole as a function of the vibrational state have been performed by Leeuw et al. (1970), Wachem & Dymanus (1967) and Hebert et al. (1968), who considered NaCl, KCl and both molecules, respectively. Additionally, Pluta (2001) calculated dipole moments at equilibrium bond length as part of theoretical study comparing various levels of theory [SCF, MP2, CCSD and CCSD(T)]. Giese & York (2004) computed NaCl dipole moment curves (DMCs) using a multireference configuration interaction approach and extrapolated basis sets. However, there appear to be no published KCl DMCs, experimental or ab initio.

Figure 1. Ab initio DMCs for NaCl and KCl in their ground electronic states.

We determined new DMCs for both molecules using high-level ab initio calculations, shown in Fig. 1. These were performed using MOLPRO (Werner et al. 2010). The points defining the new dipole moment functions are given in Tables 3 and 4. The final NaCl DMC was computed using an aug-cc-pCVQZ-DK basis set and the CCSD(T) method, where both core-valence and relativistic effects were also taken into account. Inclusion of both effects is known to be important (Tennyson 2014). In the case of KCl, an effective core potential ECP10MDF (MCDHF + Breit) in conjunction with the corresponding basis set (Lim et al. 2005) was used for K and aug-cc-pV(Q+d)Z was used for Cl. In both cases, the electric dipole moments were obtained using the finite field method. The ab initio DMC grid points were used directly in LEVEL. Equilibrium bond length dipole moments are compared in Table 5. Our computed equilibrium dipole for KCl is about 1 per cent larger than the experimental value. For NaCl, this difference is closer to 2 per cent but our final CCSD(T) value is close to those calculated by Giese & York (2004).

Table 5. Comparison of Na$^{35}$Cl and 39K$^{35}$Cl dipole moments at equilibrium internuclear distance.

| Reference             | Method   | $\mu$(NaCl) | $\mu$(KCl) |
|-----------------------|----------|-------------|------------|
| Hebert et al. (1968)  | Experiment | 8.9721     | 10.2384    |
| Pluta (2001)          | SCF      | 9.2774     | 10.6626    |
| Pluta (2001)          | MP2      | 9.0740     | 10.4923    |
| Pluta (2001)          | CCSD     | 9.0715     | 10.4847    |
| Pluta (2001)          | CCSD(T)  | 9.0257     | 10.4542    |
| This work             | CCSD(T)  | 9.1430     | 10.3119    |

Table 3. Computed values of the NaCl dipole, $\mu$, as a function of bond length, $R$.

| $R$ (Å) | $\mu$ (D) | $R$ (Å) | $\mu$ (D) | $R$ (Å) | $\mu$ (D) | $R$ (Å) | $\mu$ (D) |
|-------|-----------|-------|-----------|-------|-----------|-------|-----------|
| 1     | 1.217 469 1604 | 1.9  | 7.400 231 757 | 2.5  | 9.734 259 68 | 3.4  | 13.999 391 946 |
| 1.1   | 1.861 172 7533 | 2    | 7.745 338 2459 | 2.6  | 10.173 5001 | 3.5  | 14.505 584 0304 |
| 1.2   | 3.038 629 5206 | 2.1  | 8.110 383 8383 | 2.7  | 10.623 436 162 | 3.6  | 15.015 891 0498 |
| 1.3   | 4.514 733 3677 | 2.2  | 8.493 459 0983 | 2.8  | 11.083 042 1271 | 3.7  | 15.529 730 5282 |
| 1.4   | 5.580 998 3601 | 2.3  | 8.892 826 3074 | 2.9  | 11.551 435 3763 | 3.8  | 16.046 526 8712 |
| 1.5   | 6.135 791 0101 | 2.35 | 9.098 124 5849 | 3    | 12.027 826 798 | 3.9  | 16.565 708 6553 |
| 1.6   | 6.477 882 0685 | 2.37 | 9.181 232 1663 | 3.1  | 12.511 486 8654 | 4    | 17.086 702 9683 |
| 1.7   | 6.773 525 7195 | 2.39 | 9.264 886 8554 | 3.2  | 13.001 729 995 | 4.1  | 17.608 925 658 |
| 1.8   | 7.076 663 2648 | 2.4  | 9.306 915 7558 | 3.3  | 13.497 905 996 | 4.2  | 18.131 767 3948 |

Table 4. Computed values of the KCl dipole, $\mu$, as a function of bond length, $R$.

| $R$ (Å) | $\mu$ (D) | $R$ (Å) | $\mu$ (D) | $R$ (Å) | $\mu$ (D) | $R$ (Å) | $\mu$ (D) |
|-------|-----------|-------|-----------|-------|-----------|-------|-----------|
| 1.5   | 1.830 345 58 | 2.4  | 8.996 620 28 | 2.76 | 10.772 468 | 3.3  | 13.490 859 15 |
| 1.6   | 3.739 305 15 | 2.42 | 9.095 698 99 | 2.77 | 10.821 950 94 | 3.4  | 14.005 846 02 |
| 1.7   | 4.869 647 22 | 2.45 | 9.244 043 43 | 2.78 | 10.871 458 14 | 3.5  | 14.524 113 63 |
| 1.8   | 5.674 082 82 | 2.5  | 9.490 746 12 | 2.79 | 10.920 990 47 | 3.6  | 15.045 513 41 |
| 1.9   | 6.334 728 37 | 2.55 | 9.737 0491 | 2.8  | 10.970 548 67 | 3.7  | 15.569 756 19 |
| 2     | 6.922 318 76 | 2.6  | 9.983 220 69 | 2.82 | 11.069 745 68 | 3.8  | 16.096 523 74 |
| 2.05  | 7.199 4271 | 2.62 | 10.081 704 13 | 2.85 | 11.218 753 05 | 3.9  | 16.625 479 73 |
| 2.1   | 7.468 781 85 | 2.65 | 10.229 485 29 | 2.88 | 11.368 029 41 | 3.95 | 16.890 670 47 |
| 2.15  | 7.732 1183 | 2.68 | 10.377 3678 | 2.9  | 11.467 036 66 | 4    | 17.156 2775 |
| 2.2   | 7.990 770 44 | 2.7  | 10.476 029 71 | 2.95 | 11.717 464 27 | 4.05 | 17.422 257 12 |
| 2.25  | 8.245 785 51 | 2.72 | 10.574 761 44 | 3    | 11.968 082 68 | 4.1  | 17.688 565 35 |
| 2.3   | 8.498 000 13 | 2.74 | 10.673 571 52 | 3.1  | 12.472 017 09 | 4.15 | 17.955 158 96 |
| 2.35  | 8.748 0928 | 2.75 | 10.723 008 51 | 3.2  | 12.979 644 97 | 4.2  | 18.221 994 22 |
2.4 Fitting the potentials

The PECs were refined by fitting to the spectroscopic data identified in Table 1. However, extending the temperature range of the spectra requires consideration of highly excited levels and extrapolation of the PECs beyond the region determined by experimental input values; hence, care needs to be taken to ensure that the curves maintain physical shapes outside the experimentally refined regions. In this context, we define a physical shape to be the shape of the ab initio curve. We tested multiple potential energy forms, namely the extended Morse oscillator (EMO), Morse long range (MLR) and Morse–Lennard Jones potentials (Le Roy 2011), to achieve an optimum fit to the experimental data whilst maintaining a physical curve shape. Data for multiple isotopologues were fitted simultaneously to ensure that the resulting curves are valid for all isotopologues. \( r_c \) and \( D_\beta \) were held constant in the fits, as the fits were found to be unstable otherwise.

For NaCl, BOB terms did not improve the quality of the fit and were not pursued. Of the 1370 lines used in the fit, 1060 were Na\(^{35}\)Cl and 310 were Na\(^{37}\)Cl. The final potential was expressed as an EMO:

\[
V_{\text{EMO}}(r) = D_e \left[ 1 - e^{-\beta(y-r_c)} \right]^2,
\]

where

\[
\beta(y) = \beta_{\text{EMO}}(y_p^{eq}(r)) = \sum_{i=0}^{N_p} \beta_i y_p^{eq}(r)^i,
\]

\[
\gamma_p^{eq}(r) = \frac{r_p - r_p^0}{r_p + r_p^0}
\]

and \( p \) was set to 3, \( N \) to 4, \( D_e \) to 34 120.0 cm\(^{-1} \) (Huber & Herzberg 1979) and \( r_c \) to 2.360 796 042 Å (Ram et al. 2014). Parameters resulting from the fit are given in Table 6.

For KCl potassium, centrifugal non-adiabatic BOB terms were included in the fit as they resulted in a reduction by up to 50 per cent, in the residuals (obs–calc) obtained for high \( J \)'s. Of the 549 lines used in the fit, 361 were \(^{35}\)K\(^{35}\)Cl, 82 were \(^{35}\)K\(^{37}\)Cl, 64 were \(^{41}\)K\(^{35}\)Cl and 40 were \(^{41}\)K\(^{37}\)Cl. The final potential was expressed as an MLR:

\[
V_{\text{MLR}}(r) = D_e \left[ 1 - \frac{u_{1\beta}(r)}{u_{\text{LR}}(r_c)} e^{-\beta(y-r_c)} \right]^2,
\]

where

\[
\beta(y) = \beta_{\text{MLR}}(y_p^{eq}(r)) = \gamma_p^{eq}(r) + \left[ 1 - \gamma_p^{eq}(r) \right] \sum_{i=0}^{N_p} \beta_i y_p^{eq}(r)^i,
\]

\[
y_p^{eq}(r) = \frac{r_p - r_p^0}{r_p + r_p^0}
\]

and parameters used in the KCl MLR potential, see equation (4). (Uncertainties are given in parentheses in units of the last digit.)

| \( N \) | \( \beta_i \) | \( t_j \) |
|---|---|---|
| 0 | 0.894 7078(17) | 0.0 |
| 1 | -0.287 528(48) | 0.0 |
| 2 | 0.005 81(11) | 0.0 |
| 3 | -0.0278(14) | 0.0 |
| 4 | -0.0290(37) | 0.0 |

Table 7. Fitting parameters used in the KCl MLR potential, see equation (4). (Uncertainties are given in parentheses in units of the last digit.)

The calculated energy levels, see Section 3, were summed in Excel to generate partition function values for a range of temperatures.
While sodium has only a single stable isotope, $^{23}\text{Na}$, both potassium and chlorine each have two: $^{39}\text{K}$ (whose natural terrestrial

\begin{table}
\centering
\caption{Comparison of theoretically predicted Na$^{35}\text{Cl}$ ro-vibrational wavenumbers, in cm$^{-1}$, with some of the laboratory measurements of Ram et al. (1997).}
\begin{tabular}{cccccc}
\hline $v'$ & $J'$ & $v''$ & $J''$ & Obs. & Calc. \\
\hline 1 & 99 & 0 & 98 & 387.0444 & 387.0446 \\
2 & 100 & 100 & 387.1219 & 387.1221 \\
3 & 101 & 3 & 387.1950 & 387.1957 \\
2 & 3 & 359.9248 & 359.9260 \\
4 & 2 & 359.3419 & 359.3444 \\
5 & 2 & 359.7587 & 359.7596 \\
6 & 2 & 380.2722 & 380.2746 \\
7 & 2 & 380.3014 & 380.3075 \\
8 & 2 & 380.3372 & 380.3365 \\
3 & 2 & 361.3425 & 361.3429 \\
4 & 2 & 361.6718 & 361.6730 \\
5 & 2 & 362.3244 & 362.3230 \\
6 & 2 & 362.7560 & 362.6104 \\
7 & 2 & 362.0519 & 362.0519 \\
8 & 2 & 363.4289 & 363.4254 \\
9 & 2 & 369.5373 & 369.5406 \\
10 & 2 & 369.5551 & 369.5562 \\
11 & 2 & 369.7581 & 369.7585 \\
12 & 2 & 362.1649 & 362.1606 \\
13 & 2 & 362.3244 & 362.3276 \\
14 & 2 & 362.4924 & 362.4910 \\
15 & 2 & 357.0519 & 357.0709 \\
16 & 2 & 357.9881 & 357.9876 \\
17 & 2 & 351.2687 & 351.2710 \\
\hline
\end{tabular}
\end{table}

\begin{table}
\centering
\caption{Comparison of theoretically predicted $^{39}\text{K}^{35}\text{Cl}$ ro-vibrational wavenumbers, in cm$^{-1}$, with some of the laboratory data of Ram et al. (1997), as re-assigned in this work.}
\begin{tabular}{cccccc}
\hline $v'$ & $J'$ & $v''$ & $J''$ & Obs. & Calc. \\
\hline 1 & 102 & 0 & 101 & 294.9349 & 294.9347 \\
2 & 103 & 0 & 102 & 295.0173 & 295.0154 \\
3 & 104 & 0 & 103 & 295.0955 & 295.0941 \\
4 & 105 & 0 & 104 & 295.1729 & 295.1710 \\
5 & 43 & 1 & 42 & 284.5787 & 284.5795 \\
6 & 49 & 1 & 48 & 285.6588 & 285.6551 \\
7 & 51 & 1 & 50 & 286.0004 & 285.6001 \\
8 & 52 & 1 & 51 & 286.1646 & 286.1700 \\
9 & 53 & 1 & 52 & 291.1544 & 291.1554 \\
10 & 54 & 1 & 53 & 291.2013 & 291.1992 \\
11 & 55 & 1 & 54 & 291.2433 & 291.2411 \\
12 & 56 & 1 & 55 & 291.3562 & 291.3555 \\
13 & 57 & 1 & 56 & 284.6069 & 284.6019 \\
14 & 58 & 1 & 57 & 284.7309 & 284.7300 \\
15 & 59 & 1 & 58 & 284.8579 & 284.8563 \\
16 & 60 & 1 & 59 & 285.1054 & 285.1036 \\
17 & 61 & 1 & 60 & 285.7192 & 285.7174 \\
18 & 62 & 1 & 61 & 285.8341 & 285.8375 \\
19 & 63 & 1 & 62 & 285.8912 & 285.8947 \\
20 & 64 & 1 & 63 & 286.0004 & 286.0037 \\
21 & 65 & 1 & 64 & 279.6821 & 279.6817 \\
22 & 66 & 1 & 65 & 279.9388 & 279.9355 \\
23 & 67 & 1 & 66 & 280.0552 & 280.0598 \\
24 & 68 & 1 & 67 & 280.6586 & 280.6552 \\
\hline
\end{tabular}
\end{table}

\begin{table}
\centering
\caption{Comparison of theoretically predicted $^{39}\text{K}^{35}\text{Cl}$ R-branch band heads, in cm$^{-1}$, with laboratory measurements from Ram et al. (1997) and this work.}
\begin{tabular}{cccc}
\hline Band & Observed & Calculated & O – C \\
\hline 1-0 & 296.702 & 296.703 & 0.001 \\
2-1 & 294.181 & 294.182 & 0.001 \\
3-2 & 291.680 & 291.682 & 0.002 \\
4-3 & 289.201 & 289.203 & 0.002 \\
5-4 & 286.742 & 286.745 & 0.003 \\
6-5 & 284.303 & 284.306 & 0.003 \\
7-6 & 281.884 & 281.887 & 0.003 \\
8-7 & 279.488 & 279.489 & 0.001 \\
9-8 & 277.110 & 277.110 & 0.0 \\
10-9 & 274.752 & 274.752 & 0.0 \\
11–10 & 272.414 & 272.411 & 0.003 \\
12–11 & 270.120 & 270.090 & 0.03 \\
\hline
\end{tabular}
\end{table}

We determined that our partition function is at least 95 per cent converged at 3000 K and much better than this at lower temperatures. Therefore, temperatures up to 3000 K were considered. Values for the parent isotopologues are compared to previous studies, namely Irwin (1981), Sauval & Tatum (1984) and CDMS, in Table 11.

For ease of use, we fitted our partition functions, $Q$, to a series expansion of the form used by Vidler & Tennyson (2000):

\begin{equation}
\log_{10} Q(T) = \sum_{n=0}^{6} a_n \left[ \log_{10} T \right]^n
\end{equation}

with the values given in Table 12.

### 2.6 Line-list calculations

While sodium has only a single stable isotope, $^{23}\text{Na}$, both potassium and chlorine each have two: $^{39}\text{K}$ (whose natural terrestrial
Table 11. Comparison of Na$^{35}$Cl and $^{39}$K$^{35}$Cl partition functions.

| T (K) | This work | CDMS | Irwin (1981) | Sauval & Tatum (1984) |
|-------|-----------|------|-------------|-----------------------|
| Na$^{35}$Cl |
| 9.375 | 30.3338  | 30.3307 | –           | –                     |
| 18.75 | 60.3352  | 60.3299 | –           | –                     |
| 37.5  | 120.3556 | 120.3455| –           | –                     |
| 75    | 240.6984 | 240.6770| –           | –                     |
| 150   | 496.6455 | 496.5538| –           | –                     |
| 225   | 802.3712 | 802.1167| –           | –                     |
| 300   | 1173.0397| 1172.5403| –         | –                     |
| 500   | 2506.9232| 2505.0340| –         | –                     |
| 1000  | 8161.702 | –      | 8204.6     | 8165.4               |
| 1500  | 17333.48 | –      | 17409.8    | 16960.3              |
| 2000  | 30294.77 | –      | 30370.1    | 29685.3              |
| 2500  | 47362.31 | –      | 47324.9    | 46807.2              |
| 3000  | 68909.60 | –      | 68530.1    | 68766.1              |
| 39$^{35}$K$^{35}$Cl |
| 9.375 | 51.1529  | 51.1495| –           | –                     |
| 18.75 | 101.9823 | 101.9724| –          | –                     |
| 37.5  | 203.6737 | 203.6504| –          | –                     |
| 75    | 409.1563 | 409.1053| –          | –                     |
| 150   | 876.2078 | 876.0902| –          | –                     |
| 225   | 1474.9611| 1474.7618| –         | –                     |
| 300   | 2225.1732| 2224.8905| –         | –                     |
| 500   | 49607.42 | 49600.33| –          | –                     |
| 1000  | 17102.33 | –      | 17277.73   | 17112.5               |
| 1500  | 37064.68 | –      | 37327.7    | 36147.1              |
| 2000  | 65580.22 | –      | 65747.4    | 64142.9              |
| 2500  | 103489.55| –      | 103058.6   | 102212.0             |
| 3000  | 151831.71| –      | 149837.7   | 151368.0             |

Table 12. Fitting parameters used to fit the partition functions, see equation 10. Fits are valid for temperatures between 500 and 3000 K.

| Na$^{35}$Cl | Na$^{37}$Cl | $^{39}$K$^{35}$Cl | $^{39}$K$^{37}$Cl | $^{41}$K$^{35}$Cl | $^{41}$K$^{37}$Cl |
|------------|-------------|-------------------|-------------------|-------------------|-------------------|
| $a_0$      | 35.528 812  | 39.941 335        | 71.922 595        | 72.206 029        | 74.926 932        |
| $a_1$      | $-65.142$ 353 | $-73.363$ 68       | $-138.068$ 267    | $-138.340$ 743    | $-143.689$ 312    |
| $a_2$      | 53.290 409  | 59.657 658        | 114.111 477       | 114.101 455       | 118.474 798       |
| $a_3$      | $-23.592$ 248 | $-26.212$ 185      | $-50.527$ 0109    | $-50.414$ 764     | $-52.319$ 838     |
| $a_4$      | 6.036 705   | 6.641 337         | 12.738 250        | 12.683 942        | 13.150 052       |
| $a_5$      | $-0.837$ 0958 | $-0.911$ 3382      | $-1.726$ 9110     | $-1.716$ 3406     | $-1.777$ 0609     |
| $a_6$      | 0.048 872 72 | 0.052 663 06       | 0.098 186 748     | 0.097 426 548     | 0.100 716 4997    |

Table 13. Summary of our line lists.

| Na$^{35}$Cl | Na$^{37}$Cl | $^{39}$K$^{35}$Cl | $^{39}$K$^{37}$Cl | $^{41}$K$^{35}$Cl | $^{41}$K$^{37}$Cl |
|------------|-------------|-------------------|-------------------|-------------------|-------------------|
| Maximum $v$ | 100         | 100               | 120               | 120               | 120               |
| Maximum $J$ | 557         | 563               | 500               | 500               | 500               |
| Number of lines | 4734 567 | 4763 324         | 7224 331          | 7224 331          | 7224 331          |

abundance is about 93.25 per cent) and $^{41}$K (6.73 per cent), and $^{35}$Cl (75.76 per cent) and $^{37}$Cl (24.24 per cent). Line lists were therefore calculated for two NaCl and four KCl isotopologues. 

3 RESULTS

The full line list computed for all isotopologue considered is summarized in Table 13. Each line list contains around 4 million

for two NaCl and four KCl isotopologues: Na$^{35}$Cl, Na$^{37}$Cl, $^{39}$K$^{35}$Cl, $^{39}$K$^{37}$Cl, $^{41}$K$^{35}$Cl and $^{41}$K$^{37}$Cl. The computed line lists are available in electronic form as supplementary information to this paper.
Table 14. Extract from start of states file for Na$^{35}$Cl.

| $I$ | $\tilde{E}$ | $g$ | $J$ | $v$ |
|-----|-------------|-----|-----|-----|
| 1   | 0.000 000   | 16  | 0   | 0   |
| 2   | 0.434 501   | 48  | 1   | 0   |
| 3   | 1.303 497   | 80  | 2   | 0   |
| 4   | 2.606 971   | 112 | 3   | 0   |
| 5   | 4.344 901   | 144 | 4   | 0   |
| 6   | 6.517 259   | 176 | 5   | 0   |

$I$: state counting number; $\tilde{E}$: state energy in cm$^{-1}$; $g$: state degeneracy; $J$: state rotational quantum number; $v$: state vibrational quantum number.

Table 15. Extract from start of states file for $^{39}$K$^{35}$Cl.

| $I$ | $\tilde{E}$ | $g$ | $J$ | $v$ |
|-----|-------------|-----|-----|-----|
| 1   | 0.0         | 16  | 0   | 0   |
| 2   | 0.256 466   | 48  | 1   | 0   |
| 3   | 0.769 393   | 80  | 2   | 0   |
| 4   | 1.538 778   | 112 | 3   | 0   |
| 5   | 2.564 613   | 144 | 4   | 0   |
| 6   | 3.846 887   | 176 | 5   | 0   |

$I$: state counting number; $\tilde{E}$: state energy in cm$^{-1}$; $g$: state degeneracy; $J$: state rotational quantum number; $v$: state vibrational quantum number.

Table 16. Extracts from the transitions file for Na$^{35}$Cl.

| $I$ | $F$ | $A_{IF}$ |
|-----|-----|----------|
| 1   | 2   | 1.89E−07 |
| 2   | 3   | 1.81E−06 |
| 3   | 4   | 6.55E−06 |
| 4   | 5   | 1.61E−05 |
| 5   | 6   | 3.21E−05 |
| 6   | 7   | 5.64E−05 |

$I$: upper state counting number; $F$: lower state counting number; $A_{IF}$: Einstein A coefficient in s$^{-1}$.

Table 17. Extracts from the transitions file for $^{39}$K$^{35}$Cl.

| $I$ | $F$ | $A_{IF}$ |
|-----|-----|----------|
| 2   | 3   | 7.21E−07 |
| 3   | 4   | 6.93E−06 |
| 4   | 5   | 2.50E−05 |
| 5   | 6   | 6.16E−05 |
| 6   | 7   | 1.23E−04 |
| 7   | 8   | 2.16E−04 |

$I$: upper state counting number; $F$: lower state counting number; $A_{IF}$: Einstein A coefficient in s$^{-1}$.

The CDMS data base contains 607 and 772 rotational lines for Na$^{35}$Cl and $^{39}$K$^{35}$Cl, respectively. Comparisons with the CDMS lines are presented in Fig. 4. As can be seen, the agreement is excellent for both frequency and intensity. In particular, predicted line intensities agree within 2 and 4 per cent for the KCl and NaCl isotopomers considered in CDMS, respectively.

Emission cross-sections for Na$^{35}$Cl and $^{39}$K$^{35}$Cl were simulated using Gaussian line-shape profiles with half-width = 0.01 cm$^{-1}$ as described by Hill, Yurchenko & Tennyson (2013). The resulting synthetic emission spectra are compared to the experimental ones in Figs 5 and 6. When making comparisons, one has to be aware of a number of experimental issues. The baseline in NaCl shows residual ‘channelling’: a sine-like baseline that often appears in FTS.
Figure 4. Absorption lines of Na\textsuperscript{35}Cl and Na\textsuperscript{39}K at 300 K: ExoMol versus CDMS.

Figure 5. Emission spectra of NaCl at 1273 K: left, Ram et al. (1997); right, ExoMol. [Reprinted from Ram et al. (1997). Copyright 1997, with permission from Elsevier.]

Figure 6. Emission spectra of KCl at 1273 K: left, Ram et al. (1997); right, ExoMol.
spectra due to interference from reflections from parallel optical surfaces in the beam. For KCl, the spectrum is very weak and the baseline, which has a large offset, was not properly adjusted to zero. Given these considerations, the comparisons must be regarded as satisfactory.

4 CONCLUSIONS

We present accurate but comprehensive line lists for the stable isotopologues of NaCl and KCl. Laboratory frequencies are reproduced to much more than sub-wavenumber accuracy. This accuracy should extend to all predicted transition frequencies up to at least \( v = 8 \) and 12 for NaCl and KCl, respectively. New ab initio dipole moments and Einstein A coefficients are computed. Comparisons with the semi-empirical CDMS data base suggest that the pure rotational intensities are accurate.

The results are line lists for the rotation–vibration transitions within the ground states of Na\(^{35}\)Cl, Na\(^{37}\)Cl, K\(^{35}\)Cl, K\(^{37}\)Cl, Na\(^{35}\)Cl, 3\(^{5}\)K\(^{35}\)Cl, 3\(^{7}\)K\(^{35}\)Cl, 4\(^{1}\)K\(^{35}\)Cl and 4\(^{3}\)K\(^{35}\)Cl, which should be accurate for a range of temperatures up to at least 3000 K. The line lists can be downloaded from the CDS or from www.exomol.com.

Finally, we note that HCl is likely to be the other main chlorine-bearing species in exoplanets. Comprehensive line lists for H\(^{35}\)Cl and H\(^{37}\)Cl have recently been provided by Li et al. (2013a, b).

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REFERENCES

Agundez M., Fonfria J. P., Cernicharo J., Kahane C., Daniel F., Guelin M., 2012, A&A, 543, A48
Barber R. J., Tennyson J., Harris G. J., Tolchenov R. N., 2006, MNRAS, 368, 1087
Barton E. J., Yurchenko S. N., Tennyson J., 2013, MNRAS, 434, 1469
Brewer L., Brackett E., 1961, Chem. Rev., 61, 425
Caris A., Lewen F., Winnewisser G., 2002, Z. Naturforsch. A, 57, 663
Caris A., Lewen F., Muller H. S. P., Winnewisser G., 2004, J. Mol. Struct., 695, 243
Cernicharo J., Guelin M., 1987, A&A, 183, L10
Clousier F. L., Gordy W., 1964, Phys. Rev. A, 134, 863
Giese T. J., York D. M., 2004, J. Chem. Phys., 120, 7393
Hebert A. J., Lovas F. J., Melendres C. A., Hollowell C. D., Story T. L., Jr, Street K., Jr, 1968, J. Chem. Phys., 48, 2824
Hill C., Yurchenko S. N., Tennyson J., 2013, Icarus, 226, 1673
Honig A., Mandel M., Stitch M. L., Townes C. H., 1954, Phys. Rev., 96, 629
Horai K., Fujimoto T., Nakagawa K., Uehara H., 1988, Chem. Phys. Lett., 147, 133
Huber K. P., Herzberg G., 1979, Molecular Spectra and Molecular Structure IV. Constants of Diatomic Molecules. Van Nostrand Reinhold Company, New York
Irwin A. W., 1981, ApJS, 45, 621
Kreidberg L. et al., 2014, Nature, 505, 66
Le Roy R. J., 2006, Chemical Physics Research Report CP-662R, DPotFit 1.1: A Computer Program for Fitting Diatomic Molecule Spectral Data to Potential Energy Functions. University of Waterloo, Waterloo, available at: http://leroy.uwaterloo.ca/programs/
Le Roy R. J., 2007, Chemical Physics Research Report CP-663, LEVEL 8.0: A Computer Program for Solving the Radial Schrödinger Equation for Bound and Quasibound Levels. University of Waterloo, Waterloo, available at: http://leroy.uwaterloo.ca/programs/
Le Roy R., 2011, Equilibrium Structures of Molecules. Taylor and Francis, London, p. 159
Leeuw F. H., Wachem R., Dymanus A., 1970, J. Chem. Phys., 53, 981
Lellouch E., Paulbert G., Moses J. I., Schneider N. M., Strobel D. F., 2003, Nature, 421, 45
Li G., Gordon I. E., Le Roy R. J., Hajigeorgiou P. G., Coxon J. A., Bernath P. F., Rothman L. S., 2013a, J. Quant. Spectrosc. Radiat. Transfer, 121, 78
Li G., Gordon I. E., Hajigeorgiou P. G., Coxon J. A., Rothman L. S., 2013b, J. Quant. Spectrosc. Radiat. Transfer, 130, 284
Linn I. S., Schwerdtfeger P., Metz B., Stoll H., 2005, J. Chem. Phys., 122, 104103
Milam S. N., Apponi A. J., Woolf N. J., Ziurys L. M., 2007, ApJ, 668, L131
Moullet A., Lellouch E., Moreno R., Gurwell M., Black J. H., Butler B., 2013, ApJ, 776, 32
Muller H. S. P., Schlodern F., Strutzki J., Winnewisser G., 2005, J. Mol. Struct., 742, 215
Pluta T., 2001, Mol. Phys., 99, 1535
Postberg F., Schmidt J., Hillier J., Kempf S., Srama R., 2011, Nature, 474, 620
Ram R. S., Dulick M., Guo B., Zhang K. Q., Bernath P. F., 1997, J. Mol. Spectrosc., 183, 360
Rice S. A., Klemperer W., 1957, J. Chem. Phys., 27, 573
Saaval A. I., Tatum J. B., 1984, ApJS, 56, 193
Schaefer F., Lodders K., Heglie B., 2012, ApJ, 755, 41
Tennyson J., 2014, J. Mol. Spectrosc., 296, 1
Tennyson J., Yurchenko S. N., 2012, MNRAS, 425, 21
Tennyson J., Hill C., Yurchenko S. N., 2013, in Gillasp J., Wiese W. L., Mosher J., eds, AIP Conf. Proc. Vol. 1545, Eighth International Conference on Atomic and Molecular Data and Their Applications ICAMDATA-2012. Am. Inst. Phys., New York, p. 186
Uehara H., Horai K., Nakagawa K., Fujimoto T., 1989, J. Mol. Spectrosc., 134, 98
Uehara H., Horai K., Konno T., Miura K., 1990, Chem. Phys. Lett., 169, 599
Vidler M., Tennyson J., 2000, J. Chem. Phys., 113, 9766
Wachem R., Dymanus A., 1967, J. Chem. Phys., 46, 3749
Werner H. J., Knowles P. J., Lindh R., Manby F. R., Schäfer M., 2010, MOLPRO: A Package of Ab Initio Programs. Available at: http://www.molpro.net/
Western C. M., 2013, PGOPHER 8.0: A Program for Simulating Rotational Structure. University of Bristol, Bristol, available at: http://pgopher.chm.bris.ac.uk
Yadin B., Vaness T., Conti P., Hill C., Yurchenko S. N., Tennyson J., 2012, MNRAS, 425, 34
Yang T., Kai X., Li R., Sun Y., He Y., 2014, Energy Sources A, 36, 15

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