QTAIM study of the bonding in triosmium trihydride cluster
[Os$_3$(µ-H)$_3$(µ$^3$-$η^2$-CC$_7$H$_3$(2-CH$_3$)NS)(CO)$_8$]

Nadia Ezzat Al-kirbasee$^1$ Shatha Raheem Helal Alhimidi$^2$
Muhsen Abood Muhsen Al-Ibadi$^3$

$^1$ Department of Chemistry, College of Education for Women, University of Kufa, Iraq
$^2$ Department of Basic Medical Science, College of Dentistry, Al-Muthanna University, Iraq
$^3$ Department of Chemistry, College of Science, University of Kufa, Iraq

Received 4/3/2020, Accepted 7/9/2020, Published Online First 30/4/2021

Abstract:
The Atoms in Molecules (AIM) analysis for triosmium cluster, which contains trihydride, carbon, carbonyl and 2-methylbenzothiazolide ligands, [Os$_3$(µ-H)$_3$(µ$^3$-$η^2$-CC$_7$H$_3$(2-CH$_3$)NS)(CO)$_8$] is reported. Bonding features in this cluster has been analyzed based on QTAIM ("Quantum Theory of Atoms in Molecules") in this work. The topological indices derived from electron density of relevant interactions in triosmium compound have been studied. The major interesting point of the AIM analyses is that the core of part (Os$_3$H$_3$) reveals the absence of any critical points and bond paths connecting any pairs of Os metal atoms. However, bond critical points with their bond paths occurring between all Os-H interactions were observed. For the bridged core part (Os$_3$H$_3$), a 6c–6e multicenter interaction is proposed. The topological parameters computed for the bridging 2-methylbenzothiazolide ligand interactions indicate that all of these interactions are typical related to the covalent bonds with a contribution of some double-bond character.

Keywords: DFT calculations, Osmium cluster, QTAIM bonding analysis.

Introduction:
The reactivity of benzoheterocyclic ligands with osmium carbonyl clusters giving the electron-deficient triosmium compounds, plays a central role in organometallic chemistry. In recent years, the electron-deficient triosmium clusters have extensively been studied to model the heterogeneously catalyzed hydrodinitrification of these heterocycles (1,2). Also, these clusters have been widely used to functionalize heterocyclic substrates resulting in the synthesis of new useful compounds which are not easy to gain by using classical organic methods (3,4). The analysis of the electron density distribution, using QTAIM theory (5-9), has been considered as a good tool to analyze chemical bonds and to compare the properties of atoms and molecules (10,11). This model is very important to distinguish the existence of chemical bond interactions between any two neighboring atoms within the expression of bond critical points (BCP) along the nuclear attractors in the neighborhood atoms (bond path, Bp), (12). Based on the QTAIM analysis, the electronic energy density $\rho(b)$, electron density $\rho(b)$ and Laplacian $\nabla^2\rho(b)$ calculated at critical points of the bond, offer important information about the properties of the bond. This methodology is able to a deep study of the characteristics of a metal–metal and metal–ligand interactions in cluster compounds (13). In this work, we present AIM study of triosmium cluster [Os$_3$(µ-H)$_3$(µ$^3$-$η^2$-CC$_7$H$_3$(2-CH$_3$)NS)(CO)$_8$] (14), shown in Fig. 1, focusing on the bonding behaviour between various atomic interactions, such as Os–Os, Os–Hbridged, Os–Cbridged and Os–Ccarbonyl interactions. The complex under this study has been chosen due to the lack of any previous AIM study on triosmium compound contains tribridging hydride ligands and bridging C spans the Os-Os edges. In addition, this work also discusses and compares between different topological properties of a variety interactions in M–M and M–L.
Theoretical Methods:
The optimized geometry of the triosmium cluster was obtained from the PBE1PBE level (15) using Gaussian 09 (16) program. The SDD (17) basis set augmented with a set of f-polarization functions (18), was used for Os atoms, whereas 6-31G (d,p) (19) basis set was used for other atoms of (H, C, N, O and S). The cluster was also characterized by analytical calculation of frequencies and has positive eigenvalues thus confirming that the optimised geometry is minima on the potential energy surface. The QTAIM analysis was calculated using the AIM2000 package (20) at the PBE1PBE/WTBS for Os atoms (21) in conjunction with the 6-31G(d,p) basis set for other atoms.

Results and Discussion:
According to QTAIM method, the structure of a compound is totally specified by the set of critical points at which the gradient of electron density $\nabla_2^2 \rho(b)$ is zero. The existence of a bcp and bp connecting between two neighbored atoms is taken as an indicator for the chemical bonding (22). The molecular graph for triosmium compound is presented in Fig. 2. The bcps and their bps correspond to the Os-H, Os-C, Os-N, C-C, C-H, C-N, and C-S bonds were clearly found. Remarkably, the results showed the lack of any bond critical point or bond path between any two of the osmium atoms. Six rcp (ring critical point) were found for the "N(1)-C(8)-S(1)-C(6)-C(7), C(2)-C(3)-C(4)-C(5)-C(6)-C(7), Os(1)-C(1)-C(2)-C(7)-N(1), C(1)-Os(1)-H(2)-Os(3), C(1)-Os(1)-H(1)-Os(2) and C(1)-Os(3)-H(3)-Os(2) rings".

Figure 1. (i) Schematic structure and (ii) optimized geometry of triosmium cluster.

Figure 2. Molecular graph for triosmium cluster, small red balls indicate bond critical points, yellow balls indicate ring critical points and solid lines indicate bond paths.
A gradient trajectory map for 2-methylbenzothiazole ligand, presented in Fig. 4, clearly shows the bond and ring critical points linked by bond paths associated with this ligand. For the interactions of Os(1) and C(1) with the 2-methyl benzo thiazole, the bps and bcps can also be observed.

**The Nature of the Interactions in the Core:**

Within the QTAIM paradigm, at the bond critical point, calculated topological properties such as the electron density ($\rho_b$), Laplacian ($\nabla^2 \rho_b$) and electronic energy density ($H$) award important insights into characteristic of the atomic interactions. The atomic interactions with the large $\rho_b$ values (the contraction of electronic charge in the interatomic surface), and the values of $\nabla^2 \rho(b) < 0$ and $H(b) < 0$ indicate the open-shell interactions (covalent and polar bonds). Whereas the atomic interactions with the small values of $\rho_b$ (the electronic charge is depleted in the interatomic surface), and values of $\nabla^2 \rho(b)$ and $H > 0$ belong to closed-shell interactions (ionic bond) (23). From the literature (24), the electronic energy density $H(b)$ is considered to be a suitable index to characterize an interaction. This term is calculated based on the equation $H(b) = G(b) + V(b)$, where $G(b)$ and $V(b)$ are the kinetic and potential energy densities, respectively. The energy density term is computed from the equation $1/4 \nabla^2 \rho(b) = 2G(b) + V(b)$. The results of topological indexes for the Os cluster are given in Table 1.

**Table 1. The topological indexes calculated at bcps of the studied cluster.**

\[
\begin{array}{cccccccc}
\text{Bond} & \rho_b[\text{eÅ}^{-3}] & \nabla^2 \rho_b[\text{eÅ}^{-5}] & G_b[\text{eÅ}^{-1}] & H_b[\text{eÅ}^{-1}] & V_b[\text{eÅ}^{-1}] & \varepsilon_b \\
\hline
\text{Os(1)-H(1)} & 0.109 & 0.165 & 0.086 & -0.045 & -0.131 & 0.027 \\
\text{Os(1)-H(2)} & 0.071 & 0.194 & 0.065 & -0.016 & -0.081 & 0.058 \\
\text{Os(2)-H(1)} & 0.045 & 0.113 & 0.033 & -0.005 & -0.038 & 0.619 \\
\text{Os(2)-H(3)} & 0.074 & 0.184 & 0.064 & -0.018 & -0.083 & 0.124 \\
\text{Os(3)-H(2)} & 0.085 & 0.188 & 0.072 & -0.025 & -0.097 & 0.048 \\
\text{Os(3)-H(3)} & 0.080 & 0.190 & 0.070 & -0.022 & -0.092 & 0.048 \\
\text{Os(1)-C(1)} & 0.116 & 0.192 & 0.0936 & -0.046 & -0.139 & 0.024 \\
\text{Os(2)-C(1)} & 0.131 & 0.230 & 0.113 & -0.056 & -0.169 & 0.040 \\
\text{Os(3)-C(1)} & 0.108 & 0.184 & 0.086 & -0.040 & -0.126 & 0.059 \\
\text{Os-C(CO)*} & 0.154 & 0.558 & 0.206 & -0.066 & -0.272 & 0.081 \\
\text{Os(1)-N(1)} & 0.082 & 0.373 & 0.106 & -0.0130 & -0.119 & 0.197 \\
\text{N(1)-C(7)} & 0.301 & -0.806 & 0.210 & -0.411 & -0.621 & 0.115 \\
\text{N(1)-C(8)} & 0.356 & -0.863 & 0.327 & -0.543 & -0.870 & 0.238 \\
\text{S(1)-C(6)} & 0.204 & -0.379 & 0.062 & -0.156 & -0.218 & 0.190 \\
\text{S(1)-C(8)} & 0.207 & -0.387 & 0.066 & -0.162 & -0.228 & 0.268 \\
\text{C(8)-C(9)} & 0.258 & -0.629 & 0.068 & -0.225 & -0.293 & 0.049 \\
\text{C=O*} & 0.460 & 0.511 & 0.926 & -0.798 & -1.724 & 0.001 \\
\end{array}
\]

*Average values.
The main point of interest, as mentioned above, is the full absence of any bond critical point between Os atoms of the core which are bridged by both H and C atoms. This finding is significant since it is attributed to the absence of any trace of localized electron density between Os atoms. Therefore, it can be concluded that there are no features of chemical bonding between any pair of Os atoms in the core (22). The strong ligand bridging interaction (H and C groups) destroys the topological Os-Os bond critical points with their bond paths, have been observed between supported and unsupported metal-metal atoms interaction (25-26). For Os-H bonds, the topological properties were calculated as presented in Table 1. The calculated electron density values are significantly greater than zero (0.045 - 0.109 eÅ⁻³), the Laplacian values are positive (0.113- 0.194 eÅ⁻⁵) and the negative values for H(b) (-0.005- -0.045 he⁻¹). According to these data and as argued in the Literature (27), Os-H interactions are classified as an open-shell model. Furthermore, the computed AIM ellipticities for Os-H bonds are greater than zero (of positive values), pointing out that these interactions are approximately a straight bonding. The plot of Laplacian map for the cluster core (Fig. 5-i) shows the polarization of the valence shell charge concentration (VSCC) of bridging hydrides toward the midpoint of Os, Os interactions. Whereas, the VSCC of the carbonyl C atoms are polarized toward the Os (3) atom (Fig. 5-ii) (28).

As mentioned above, the absence of a bcps between any pair of the Os metal atoms limits the interpretation of the Os–Os interaction. Instead, the delocalization index, \( \delta(A, B) \) (29), is a measure of the electrons shared between two atoms. This index can be considered as an appropriate AIM indicator to characterize the chemical bonding between atoms which do not depend on the existence of a bond critical point. The values of delocalization index \( \delta(A-B) \) computed for Os cluster are presented in Table 2. The magnitude of \( \delta(Os, Os) \) is within range 0.158-0.172 and this, in fact, is very similar to that computed by Cabeza group (28) for the Os...Os interaction (within the range 0.171-0.177) in the bridged form of [Os₃(l-H)(l-Cl)(CO)₁₀] and [Os₃(l-H)(l-OH)(CO)₁₀], which possesses no direct Os...Os bond. Furthermore, the \( \delta(Os, Os) \) computed data are similar to, or even greater than, many other published bridged M–M compounds (13,27,30). On the other hand, those data are not comparable to those values computed for unbridged Os-Os bonds in [Os₃(l-H)(l-OH)(CO)₁₀] and [Os₃(l-H)(l-Cl)(CO)₁₀] (28) 0.461 and 0.462, respectively. This finding leads us to conclude that the delocalization index is very useful to study ligand bridged M-M interactions. Then, by summarizing the data (summation of the delocalization indexes values of three Os...Os nonbonding and six Os-H bonding interactions), it can be concluded that the total \( \delta(A-B) \) for the core is 3.06 electron pairs in the core part (Os₃H₃). Therefore, the interaction in this core can be described as a 6c–6e interaction.
The computed topological values for Os-C(1) and Os-C(CO) bonds are listed in Table 1 and they show a clear variation. These bonds have obviously an open-shell bonding features, which point toward low positive values of $\nabla^2 \rho(b)$ (between 0.108 to 0.116 eÅ$^{-5}$ for Os-C(1) and 0.150 to 0.158 eÅ$^{-5}$ for Os-C(CO)), and a negative values for H(b) (between 0.040 to 0.056 he$^{-1}$ for Os-C(1) and -0.062 to -0.070 he$^{-1}$ for Os-C(CO)). Rather interestingly, the Os-C(CO) have topological parameter values significantly greater than Os-C(1). Moreover, the topological values of the bridging C ligand are very similar to those calculated for bridging H and comparable to the results of bridging CO ligand of [FeCo(CO)$_5$] (31), bridging H on "[Cr$_2$(μ-H)(CO)$_{10}$]" (32) and bridging CH on "[Ru$_2$(μ-H)$_2$(μ$_3$-MeImCH)(CO)$_9$]" (13). Concerning to the calculated data of Os(1)-N(1) bond, the property of this bond is typical for the open-shell interaction with a small positive value of $\nabla^2 \rho(b)$ (0.373 eÅ$^{-5}$), a low ρ(b) value (0.082 eÅ$^{-3}$), and a negative of H(b) value close to zero (-0.013 he$^{-1}$).

C-N, C-S and C-C Interactions in the 2-Methylbenzothiazolide Ligand.

The topological properties of the atom-atom interactions in bridging 2-methylbenzothiazolide ligand are presented in Tables 1. In this ligand, the nature of the bonding may also be appreciated in the graphically representations of the gradient map (shown in Fig. 4) and the Laplacian map (depicted in Fig. 6). Those two maps clearly show all the bond critical point, bond paths and atomic basins which existed in the bridging 2-methylbenzothiazolide ligand. Also, the polarization of the VSCCs of the bridging C(1) is shifted toward the interatomic surface of Osmium atoms (Fig. 6), whereas the N(1) polarised to the Os(1).

The topological parameters computed for the bonds within bridging 2-methylbenzothiazolide, presented in Table 1, are typical for covalent interactions, with values for ρ(b) higher than zero and a significantly negative $\nabla^2 \rho(b)$ and H(b) values, with some degree of delocalization (22,33-35).

| Atom pairs (A, B) | δ(A, B) | Atom pairs (A, B) | δ(A, B) ** |
|-------------------|---------|-------------------|-----------|
| Os(1)--Os(2)      | 0.172   | Os(3)--H(1)       | 0.008     |
| Os(1)--Os(3)      | 0.158   | Os(3)--H(2)       | 0.498     |
| Os(2)--Os(3)      | 0.172   | Os(3)--H(3)       | 0.395     |
| Os(1)--H(1)       | 0.558   | Os(1)--C(1)       | 0.772     |
| Os(1)--H(2)       | 0.405   | Os(2)--C(1)       | 0.804     |
| Os(1)--H(3)       | 0.009   | Os(3)--C(1)       | 0.772     |
| Os(2)--H(1)       | 0.254   | Os(1)--N(1)       | 0.422     |
| Os(2)--H(2)       | 0.007   | Os(2)--N(1)       | 0.008     |
| Os(2)--H(3)       | 0.455   | Os(3)--N(1)       | 0.009     |

Average values.

** (δAB) delocalisation indices

**Os-C(1), Os-C(CO) and Os-N(1) Interactions.

Figure 6. Laplacian distribution $\nabla^2 \rho(b)$ plot of the triosmium cluster though the Os(1)-Os(2)-C(1)-N(1) plane

Regarding the N(1)-C(7) and N(1)-C(8) interactions, it is very important to compare their topological parameters with the C(6)-S(1) and C(8)-S(1) bonds. The computed electron density ρ(b) and H(b) for N-C bonds slightly higher than S-C bonds and this finding indicates that the former has a slightly double bond character.

Conclusion:

Topological properties of electron density for triosmium cluster "[Os$_3$(μ-H)$_2$(μ$_3$-η$_3$-CC$_7$H$_5$(2-CH$_3$-NS)(CO)$_9$]" have been calculated and interpreted based on AIM theory. Also, the bonding features in this cluster have been analyzed. The AIM analyses of the core part, Os$_3$H$_8$, reveals the existence of a bond critical point and its bond path within Os–H bond. Conversely, no bond critical points linked by bond paths between any pair of Os metal atoms are observed. However, a delocalization index has been computed. The interaction in the Os$_3$H$_8$ core can be described as a multiple 6c–6e interaction. The numerical values of the topological indices of the bridging 2-methylbenzothiazolide ligand interactions indicate that all the bonds are typical for covalent bonding mode with a slight degree of double-bond character.
Authors’ declaration:
- Conflicts of Interest: None.
- We hereby confirm that all the Figures and Tables in the manuscript are mine ours. Besides, the Figures and images, which are not mine ours, have been given the permission for republication attached with the manuscript.
- Ethical Clearance: The project was approved by the local ethical committee in University of Kufa.

References:
1. Ghosh S, Uddin MN, Begum N, Hossain GMG, Azam KA, Kabir SE. Synthesis, Structure and Reactivity of Electron Deficient Triosmium Cluster Bearing 2, 6-Dimethylbenzoazolylidene Ligand. J Chem Crystallogr. 2010;40(7): 572–8.
2. Rosenberg E, Abedin MJ, Rokhsana D, Viale A, Gobetto R, Milone L and Hardcastle K., Ligand dependent structural changes in the acid–base chemistry of electron deficient benzoheterocycle triosmium clusters. Inorganica Chim Acta. 2002;334:343–54.
3. Rosenberg E, Kumar R. New methods for functionalizing biologically important molecules using triosmium metal clusters. Dalt Trans. 2012;41(3):714–22.
4. Kabir SE, Kolwaite DS, Rosenberg E, Hardcastle K, Cresswell W, Grindstaff J. Synthesis, Structure, and Reactivity of Electron-Deficient Complexes of Quinolines with Triosmium Clusters. Organometallics. 1995;14(8):3611–3.
5. Ruiz J, Sol D, Garcia L, Mateo MA, Vivanco M, and Maelen JV. Generation and Tunable Cyclization of Formamidinate Ligands in Carbonyl Complexes of Mn(I): An Experimental and Theoretical Study. 2019;38(4):916–925.
6. Tezcan H, Šenöz H, Tokay N. Experimental and quantum chemical studies of the structural and spectral properties of novel diazeyl formazans. J Mol Struct. 2019;1190:171–83.
7. Kumar PSV, Raghavendra V, Subramanian V. Bader’s Theory of Atoms in Molecules (AIM) and its Applications. 2016;128(10):1527–36.
8. Van der Maelen JV, Cabeza JA. A topological analysis of the electron density in the N-heterocyclic carbene trithium cluster [Ru3(μ3-H)2(MeImCH)(CO)6] (MeIm = l,3-dimethylimidazol-2-ylidene). Organometallics. 2009;28(13):3666–72.
9. Kabir SE, Malik KMA, Mandal HS, Mottalib MA, Abedin MJ, Rosenberg E. Reaction of Os3(CO)(9)(eta3)-eta3-(2-C3H3NS)(mu-H) with diazomethane. The first example of a trimetallic cluster containing a mu-methylidene and a sigma-methyl. Organometallics. 2002;21(13):2593–5.
10. Adamo C., and Barone V., Toward reliable density functional methods without adjustable parameters: The PBE0 model. J. Chem. Phys., 1999; 110, 6158-69. DOI: 10.1063/1.478522.
11. Frisch MJ, Trucks GW, Schlegel HB, Scuseria GE, Robb MA, Cheeseman JR, et al. gaussian 09, Revision d. 01, Gaussian. Inc., Wallingford CT. 2009:201.
12. Fuentealba P, Szentpály L, Stoll H, Fraschino FX, Preuss H. Pseudopotential calculations including core-valence correlation: Alkali compounds. J Mol Struct THEOCHEM. 1983;93:213–9.
13. Hollwarth, A., Boehme, M., Dapprich, S., Ehlers, A. W., Gobbi, A., Jonas, V., Koehler K., Stegmann R., Veldkamp, A., and Frenking, G. A set of d-polarization functions for pseudo-potential basis sets of the main group elements AlBi and F-type polarization functions for Zn, Cd, Hg. Chem. Phys. Lett. 1993; 208, 237–240.
14. Petersson G. A., Bennett, Tensfeldt A. T. G., Al-Laham M. A., Shirley W. A., and Mantzaris J., A complete basis set model chemistry. I. The total energies of closed-shell atoms and hydrides of the first-row atoms, J. Chem. Phys., 1988; 89: 2193-218. DOI: 10.1063/1.455064.
15. Piegler-König F, Schönbohm J. Update of the AIM2000-program for atoms in molecules. J. comput. Chem. 2002; 23(15):1489.
16. Huzinaga, S. and Klobukowski, M. Well-tempered Gaussian basis sets for the calculation of matrix Hartree-Fock wavefunctions. Chem. Phys. Lett. 1993; 212, 260–264.
17. Bader, R. F. W., A bond path: a universal indicator of bonded interactions. J Phys Chem A. 1998;102(37):7314–23.
18. Kutzenigg W. Atoms in Molecules. A Quantum Theory.(Reihe: International Series of Monographs on Chemistry, Vol. 22.) Von RFW Bader. Clarendon
Analysis of the bonds in incomplete molecules. J Am Chem, 1975;97(26):7391-7399.

Bo C, Sarasa J. P., Poblet J. M. Laplacian of charge density for binuclear complexes: Terminal vs bridging carbynyls. J Phys Chem. 1993;97(24): 6362 – 6.

Macchi P, Garlaschelli L, Sironi A. Electron density of semi-bridging carbynyls. Metamorphosis of CO ligands observed via experimental and theoretical investigations on \([\text{FeCo(CO)}_3]\). J Am Chem Soc. 2002;124(47):14173 – 84.

Macchi P, Donghi D, Sironi A. The electron density of bridging hydrides observed via experimental and theoretical investigations on \([\text{Cr}_2(\mu^2-\text{H})(\text{Co})_{3.5}]\). J Am Chem Soc. 2005;127(47):16494–504.

Ahlhomi SRH, Al-Ibadi MAM, Alkurbasy NE, Alhimidi SRH. The topological classification of the bond in \([\text{CP}^\prime \text{Ru}](\mu^2-\text{N})(\mu-\text{H})_2\mu^2-\text{H})_2\text{Haminon-N}, \text{Ni}(\text{CO})_3\]). In: Journal of Physics: Conference Series. IOP Publishing; 2018, 12068.

Van der Maelen JF, García-Granda S, Cabeza JA. Theoretical topological analysis of the electron density in a series of triosmium carbonyl clusters: \([\text{Os}_3(\mu-\text{H})_3(\mu^1-\eta^2-\text{CC}_7\text{H}_5(2-\text{CH}_3)\text{NS})(\text{CO})_8],\text{Os}_3(\mu-\text{H})_3(\mu^1-\eta^2-\text{C}_8\text{H}_7)(\text{CO})_8],\text{Os}_3(\mu-\text{H})_3(\mu^1-\eta^2-\text{C}_8\text{H}_7)(\text{CO})_8]\). Comput Theor Chem [Internet]. 2011:968(1-3):55–63. Available from: http://dx.doi.org/10.1016/j.comptc.2011.05.003

Bader R, Stephens M. Spatial localization of the electronic pair and number distributions in H-bonded complexes. Theor Chem. 2002;26(7):844–50.

Macchi P, Sironi A. Chemical bonding in transition metal carbonyl clusters: complementary analysis of theoretical and experimental electron densities. Coord Chem Rev. 2003;238:383–412.

Feliz M, Llusar R, Andrés J, Berski S, Silvi B. Topological analysis of the bonds in incomplete cuboidal \([\text{Mo}3\text{S}4]\) clusters. New J Chem. 2002;26:7(8):14–40.

Van der Maelen JF, Garcia-Granda S, Cabeza JA. Theoretical topological analysis of the electron density in a series of triosmium carbonyl clusters: \([\text{Os}_3(\mu-\text{H})_3(\mu^1-\eta^2-\text{CC}_7\text{H}_5(2-\text{CH}_3)\text{NS})(\text{CO})_8],\text{Os}_3(\mu-\text{H})_3(\mu^1-\eta^2-\text{C}_8\text{H}_7)(\text{CO})_8],\text{Os}_3(\mu-\text{H})_3(\mu^1-\eta^2-\text{C}_8\text{H}_7)(\text{CO})_8]\). Comput Theor Chem [Internet]. 2011:968(1-3):55–63. Available from: http://dx.doi.org/10.1016/j.comptc.2011.05.003

Bader R, Stephens M. Spatial localization of the electronic pair and number distributions in H-bonded complexes. Theor Chem. 2002;26(7):844–50.