The Direct Reduction of Iron Ore with Hydrogen

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Abstract: The steel industry represents about 7% of the world’s anthropogenic CO2 emissions due to the high use of fossil fuels. The CO2-lean direct reduction of iron ore with hydrogen is considered to offer a high potential to reduce CO2 emissions, and this direct reduction of Fe2O3 powder is investigated in this research. The H2 reduction reaction kinetics and fluidization characteristics of fine and cohesive Fe2O3 particles were examined in a vibrated fluidized bed reactor. A smooth bubbling fluidization was achieved. An increase in external force due to vibration slightly increased the pressure drop. The minimum fluidization velocity was nearly independent of the operating temperature. The yield of the direct H2-driven reduction was examined and found to exceed 90%, with a maximum of 98% under the vibration of ~47 Hz with an amplitude of 0.6 mm, and operating temperatures close to 500 °C. Towards the future of direct steel ore reduction, cheap and “green” hydrogen sources need to be developed. H2 can be formed through various techniques with the catalytic decomposition of NH3 (and CH4), methanol and ethanol offering an important potential towards production cost, yield and environmental CO2 emission reductions.

Keywords: hydrogen; iron ore; direct reduction; fluidized bed; solar energy

1. Introduction

1.1. Background Information

Most steel mills are currently using the blast furnace principle, fuelled by mostly coal, coke and coke oven gas. The successive processes of sinter belt, blast furnace and converter operate at extremely high temperatures, hence being energy-intensive and producing significant amounts of CO2. The steel industry represents about 7% of the world’s anthropogenic CO2 emissions. Within the objectives of COP21 (Paris, 2015), the target of steel mills is to reduce the CO2 emissions from the current (nearly) 3000 Mton/year to below 500 Mton/year. The use of renewable energy sources is, hence, a topic of major development, and ultimately the CO2-lean direct reduction of iron ore with hydrogen is considered to offer high potential. Two alternative direct H2 ore reductions (Figure 1, right-side) and the traditional fossil-fuel blast furnace operation (Figure 1, left-side) are illustrated in Figure 1 below.

Whereas the production of sponge iron from iron ore pellets is already commercially applied, the direct H2 reduction of crude iron ore powder has the advantages of omitting the sintering stage that is the major energy consumer and CO2 emission source of the iron-making process, together with the blast furnace. Sintering is indeed required to produce a sinter cake to be loaded into the blast furnace. It thermally agglomerates the mixture of iron ore fines, recycled iron-making products, fluxes, slag-forming agents and solid fuel (coke) at a temperature between 1300 and 1480 °C. Omitting the sintering operation will hence reduce the overall energy consumption and CO2 emissions by 35 to 40% [1].
The main reasons behind researching the reduction of CO\textsubscript{2} using hydrogen are based upon limiting both the required thermal energy and the associated CO\textsubscript{2} emissions. Major developments were reported since 2015, with research efforts illustrated in Figure 2. This figure was established from Web of Science statistics with appropriate keywords. Over 70% of the research dealt with the fundamentals of direct ore reduction using hydrogen and the reactors to be used (rotary furnace, plasma, fixed and fluidized beds). Additional publications tentatively focused upon process economics and environmental benefits. The fixed/moving or fluidized bed reactors represent only 7 out of 28 papers (see Table 1), stressing the need for additional research, as dealt with in the present study.

**Figure 1.** Different steel-making processes: traditional, using fossil fuels (*left*); H\textsubscript{2}-driven (*right*).

**Figure 2.** Review of publications since 2015, from Web of Science, with keywords iron ore, direct reduction, hydrogen, CO\textsubscript{2} reduction.

### Table 1. Important research contributions since 2017.

| References | Major Findings |
|------------|----------------|
| Sohn et al. [2] | Developed a novel iron-making technology based on fine iron ore concentrate in a flash reactor. Optimization towards using multiple burners, preheating the feed gas, or both is examined. |
| Abolpour et al. [3] | Studied the reaction kinetics of H\textsubscript{2} reduction of iron ore concentrate particles between 700 and 900 °C. |
| Murakami et al. [4] | Studied the effect of H\textsubscript{2} concentration in the reducing gas on the changes in mineral phases during reduction of iron ore sinter. The mineral composition of 10 types of sinter samples was analyzed. |
| Patisson & Mir- gauxb [5] | Analyzed a new route for making steel from iron ore based on the use of H\textsubscript{2}, with a 90% reduction in CO\textsubscript{2} emissions compared to those of the blast-furnace route. |
| Zhang et al. [6] | Described the reduction kinetics by a pellet-scale single-interface shrinking core model at temperatures below 1143 K. At higher temperatures, the rate-limiting step is the interfacial chemical reaction process. |
| Bai et al. [7] | Investigated the thermal reduction of pellets by H\textsubscript{2}. An un-reacted core and interfacial chemical reaction are reaction rate controlling steps in function of the operating temperature. |
| Wei et al. [8] | Conducted reduction experiments with a H\textsubscript{2}-N\textsubscript{2} mixture in a rotary drum reactor. A dense, increasingly thick layer of iron was formed during the reduction reaction. The reduction process was controlled by a first-order reaction model. |
| Bai et al. [9] | Studied the direct reduction behavior and dynamic characteristics of oxidized pellets under the 75% H\textsubscript{2}-25% N\textsubscript{2} atmosphere at 760, 900, 1000 °C. Holes and cracks were observed at higher temperature, hampering the reduction degradation behavior of the pellets. |
| Lyu et al. [10] | Investigated the effect of hydrogen addition on gas utilization in a blast furnace. The H\textsubscript{2} content should be controlled in the range of 5%–10%. |
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| Lyu et al. [10]    | Investigated the effect of hydrogen addition on gas utilization in a blast furnace. The H₂ content should be controlled in the range of 5%–10%. |
| Elzohiery et al. [11] | Demonstrated that the reduction is of a first-order dependence with the hydrogen partial pressure. |
| Du et al. [12]     | Proposed a new process that could simultaneously enrich CH₄ from coke oven gas and produce separated magnetite from low-grade hematite. |
| Sohn et al. [13]   | Reduced the concentrate by H₂ and CO gas mixtures formed from the partial oxidation of natural gas in a flash reactor. The rate equations were applied to experimental results from a laboratory flash reactor using CFD simulation. |
| Htet et al. [14]   | Investigated the gasification reactivity of carbon black compared to the carbonaceous materials used in the HiSarna process by TGA method at 1250 °C, 1350 °C and 1450 °C under atmospheric pressure. Carbon black is less reactive than thermal coal and charcoal. The random pore model best describes the gasification reaction of the selected samples. Activation energies were determined. |
| Bhaskar et al. [15] | Developed a mass and energy flow model based upon Python to explore the feasibility of using H₂-based direct reduction of iron ore (HDDR) coupled with an electric arc furnace (EAF) for carbon-free steel production. This technology could reduce specific emissions from steel production in the EU by more than 35%, at present grid emission levels (295 kgCO₂/MWh). |
| Vogl et al. [16]   | Hydrogen direct reduction (H-DR) steelmaking needs 3.48 MWh of electricity per ton of liquid steel, mainly for the electrolytic hydrogen production. H-DR emits only 2.8% of blast furnace CO₂. |
| He et al. [17]     | Investigated the micro-fluidized bed reduction behavior of Brazilian hematite in 20% H₂ and 80% Ar at 400–570 °C. The reaction rate was suggested to be determined by the phase boundary reaction. |
| Spreitzer & Schenk [18] | Investigated the fluidization behavior and reducibility of various iron ore grades. A stable fluidization of hematite- and limonite-based iron ores was observed, while magnetite-based iron ore could not be fluidized. |
| Knop and Kubiak [19] | Studies the comparison of a combined H₂-based direct reduction and electric steelmaking, and the ore-fines reduction in a novel fluidized bed process. |
| Hamadeh et al. [20] | Studied the moving bed principle in de direct reduction shaft furnace, providing evidence of 8 different zones according to the predominant thermal and reaction characteristics. |
| da Costa et al. [21] | In a counter-current direct-reduction moving bed reactor, CO₂ emissions were reduced by more than 80%. |
| Lazou et al. [22]  | Studied the CO₂ reduction in the conventional Pedersen Process by reducing iron or bauxite ore at temperatures below and above 560 °C. At temperatures above 860 °C, the formation of FeAl₂O₄ retards the reduction to metallic iron. |
| Moldenauer et al. [23] | A circulating fluidized bed reactor was applied in a chemical looping combustion. CO₂ reductions of 75 to 92% were achieved. |

Some selected and important publications are listed in Table 1 according to their main output. Research on fixed/moving or fluidized ore reduction was separately included in the Table. The solar energy potential was not previously reported upon.
Similar research efforts on reducing the CO\textsubscript{2} emissions from thermal processes were also presented for different fields, with a specific target of CO\textsubscript{2} modeling, such as in a large-scale coal-fired thermal power plant \cite{24}, in integrated processing of liquified natural gas \cite{25}, in a multigeneration cascade system \cite{26}, in the pyrolysis reaction of lignocellulosic waste \cite{27}, or through the integration of rigid porous high-temperature filters in thermal processes \cite{28}, among others.

1.2. Objectives, Structure and Contribution of the Reported Research

From the literature survey, it is clear that the direct H\textsubscript{2}-driven reduction of iron ore, and the considerably lower CO\textsubscript{2} emissions, are increasingly important. The previous research developed some fundamental principles about the reaction mechanisms, reduction kinetics and CO\textsubscript{2} reduction potential, as also dealt with in several studies about CO\textsubscript{2} modeling. The importance of using a zero CO\textsubscript{2} emission H\textsubscript{2} as a fuel toward CO\textsubscript{2} abatement in comparison with common fossil fuels is illustrated in Figure 3.

\begin{figure}[h]
\centering
\includegraphics[width=0.5\textwidth]{CO2_emissions.png}
\caption{CO\textsubscript{2} emissions when burning an equivalent energy unit of fuel.}
\end{figure}

The application of fluidized bed reactors in this direct reduction principle of iron ores was only scarcely investigated. The application of sustainable solar energy was not investigated. To complete and assess the previous findings, and develop the perspective of concentrated solar energy application in the fluidized bed to direct H\textsubscript{2} reductions of iron ore, this paper has multiple objectives.

Firstly, the properties of the iron ore powder that can be used in a fluidized bed H\textsubscript{2}-reduction reactor will be summarized.

Secondly, reaction fundamentals will be studied in both an isothermally operated macro-thermogravimetric reactor with electrical heating and in a vibrated fluidized bed reactor positioned in the cavity and focused on a concentrated solar furnace. Both the reduction yield and the fluidization reactor operation characteristics (velocities, pressure drop, temperature profiles within the reactor) will be determined. The results will confirm the potential of the direct H\textsubscript{2} reduction of iron ore in a solar fluidized bed reactor.

Thirdly, the production of “renewable” H\textsubscript{2} will be assessed. Although tests were performed on a lab-scale batch basis only, the economy of producing renewable H\textsubscript{2} will be demonstrated.

Finally, all experimental findings and future developments will be summarized in the conclusions.

These objectives are summarized in the research framework of the paper, as presented in Figure 4.
2. Materials and Methods

2.1. Materials

The physical and chemical properties of the fine Fe$_2$O$_3$ particles used to examine the fluidization and hydrogen reduction are listed in Table 2. These particles belong to the Geldart group C (cohesive) particles, based on particle size, particle density and gas density [29].

Table 2. Physical and chemical properties of fine Fe$_2$O$_3$ particles.

| Particle Property            | Value                                      |
|------------------------------|--------------------------------------------|
| Mean particle size (µm)      | 10–15                                      |
| Particle density (kg/m$^3$)  | 4800                                       |
| Shape                        | Non-spherical, sphericity ~0.85             |
| Chemical composition        | Iron 68%, Oxygen 30%                       |

2.2. Methods and Procedures

2.2.1. Experimental Set-Ups

Two experimental rigs were used in the study: an electrically heated furnace and a solar-heated reactor.

Isothermal conversions under the H$_2$ feed were determined by a macro-thermogravimetric analyzer (electrically-heated macro-TGA), where the weight loss under the H$_2$ feed was monitored, according to the method described by Fernandez et al. [30]. The set-up is illustrated in Figure 5. The total heating capacity is 8 kW, and the heating rate is controlled by the thermocouple installed in the Fe$_2$O$_3$ ore bed.
The iron ore (170 g) was loaded inside the reactor above the distributor. The static bed height was about 16 cm. The bottom section of the electrically heated reactor served as the preheating section of the gas feed. Nitrogen was used when starting the experiment. When the reactant bed temperature reached the set point of 773 K, the N\textsubscript{2} feed was stopped, and H\textsubscript{2} was fed to the reactor. The time of reaction was used as a process parameter. When the required time was reached, the feed of H\textsubscript{2} was stopped, N\textsubscript{2} was restarted, the electric heating was stopped, and the reactor was allowed to cool. Samples of feedstock and reaction products were also analyzed for their oxygen concentration.

Solar H\textsubscript{2} reductions were carried out in the solar thermal furnace, depicted in Figure 6. The maximum heat load in the cavity was ~20 kW. The vibrating fluidized bed was installed inside the 0.5 m × 0.5 m × 0.5 m cavity with an irradiation opening of 0.3 m diameter. Since we had to treat cohesive particles, a vibrating fluidized bed was used. The solar irradiation was collected by a heliostat and focused onto the cavity opening by a parabolic concentrator. The heat load could be controlled by partly defocusing the heliostat. A photograph of the experimental rig in operating conditions is shown in Figure 7.
Figure 7. Solar furnace in QinHuangDao (Hebei Province, China).

Table 3 lists the specifications of the solar test unit and test conditions. The internal diameter of the fluidized bed reactor was 50 mm. \( \text{H}_2 \) was preheated by a coil heat exchanger also installed in the cavity and visible at the right of Figure 6. About 500 g of fine \( \text{Fe}_2\text{O}_3 \) particles formed the bed. A multi-orifice distributor was used to ensure a uniform gas distribution in the fluidized bed reactor [31].

Table 3. Experimental conditions.

| Parameter                          | Value   |
|------------------------------------|---------|
| Fluidizing carrier gas             | \( \text{N}_2 \) |
| Reduction gas                      | \( \text{H}_2 \) |
| Superficial gas velocity (m/s)     | <0.02   |
| Bed height (m)                     | 0.25    |
| Vibration frequency [Hz]           | <50     |
| Vibration amplitude [mm]           | 0.6     |
| Vibration direction                | Vertical|
| Reaction temperature [K]           | 773–873 |

The vibration of the fluidized bed reactor was applied by adjusting the motor inverter frequency. The vibration frequency could be adjusted up to 47.7 Hz, and the amplitude was fixed to 0.6 mm. The pressure drop was checked in real-time using a differential pressure transmitter (Rosemount Inc., Shakopee, MN, USA) through pressure drop ports just above the distributor and at the top of the reactor. Hydrogen was used as the reduction gas, whereas nitrogen was used as fluidization gas at start-up (till the desired bed temperature was obtained) and at shut-down. Nitrogen and hydrogen gas was supplied to the fluidized bed using separate mass flow controllers. After the reduction was performed for a given reaction time, nitrogen was injected again to reduce the temperature of the reactor to ambient temperature, and the products were collected for analysis.

2.2.2. Data Treatment

For both experimental set-ups, the representative reduction reaction rate is shown in Equation (1) and evaluated on the basis of the weight loss during the reduction process and confirmed by the O-content of the \( \text{Fe}_2\text{O}_3 \), sampled at distinct times [32,33].

\[
\text{Reduction yield} = \frac{\text{Oxygen content in } \text{Fe}_2\text{O}_3 - \text{oxygen content in product}}{\text{Oxygen content in } \text{Fe}_2\text{O}_3}, \quad (1)
\]
3. Results and Discussion

3.1. The Fluidization Characteristics of the Fine Fe$_2$O$_3$ Particles

Due to the cohesiveness of the particles, fluidization with gas alone was impossible. Channels and rat holes formed in the cohesive particles. To increase the fluidity of the fine particles, fluidization must be performed with the help of an external force to decrease the interparticle Van der Waals force. The vibration causes the particles to remain non-agglomerated.

Vibrated fluidization provides a uniform distribution of the gas flow, prevents the formation of particle agglomerates, and weakens the interparticle force. When the combined forces of the vibration and the hydrodynamic shear force exerted on the particles by the gas flow (related in a fluidized bed to the bed weight) exceed the Van der Waals force, fluidization of the C-type particles is achieved. The vibration force depends upon the vibration frequency and amplitude. Equation (2) defines the effective acceleration, resulting from the effects of vibration frequency and amplitude [34,35]:

$$g_{ef} = g \left( 1 + \frac{A_v f^2}{g} \right),$$

with:

- $g$: gravitational acceleration ($9.81 \text{ m/s}^2$)
- $A_v$: amplitude of the vibration (m)
- $f$: frequency of the vibration (Hz; m/s)
- $g_{ef}$: effective acceleration factor (m/s$^2$)

Without vibration, the fluidization characteristics showed excessive channeling. A bubbling fluidized regime was achieved when vibration was applied. Figure 8 shows pressure drop values for the 0.25 m deep bed of fine Fe$_2$O$_3$ particles as a function of the temperature. At reaction temperatures from 273 to 993 K, the pressure drop values at 25 Hz were 4.8 kPa, irrespective of the applied temperature. When the vibration frequency was increased, the pressure drop slightly increased by a maximum of 5%. The minimum fluidization velocity, $U_{mf}$, hardly changed when the frequency and/or temperature increased, although a minimum vibration frequency of ~10 Hz was required (at the constant amplitude of 0.6 mm). The minimum fluidization velocity is around 1 mm/s.

![Figure 8. Pressure drop across the bed versus superficial gas velocity.](image)

Contrary to other Geldart particle types (A, B or D particles), fluidization has no distinct temperature-dependent tendency when vibration is applied to Geldart group C particles. At a vibration frequency of 10 Hz, the $U_{mf}$ value could not be determined because the pressure drop values for fine Fe$_2$O$_3$ particles did not exhibit a clear trend. As the external force increased, fluidization changed to the bubbling regime, with a constant pressure drop achieved in the bubbling regime.
3.2. The Fe$_2$O$_3$ Reduction Rate

Figure 9 shows the reduction yield for the vibrated fluidized bed operation. Under the specific conditions of the test, the reaction time was about 50 min for a maximum reduction yield of about 90%. Increasing the vibration frequency to 47 Hz improved the reaction, and yields close to 98% were achieved. The superficial H$_2$ velocity was 0.01 m/s.

![Figure 9. Reduction of Fe$_2$O$_3$ at 773 K with vibration at 20 (■) and 47 Hz (▲), respectively.](image)

Such high Fe$_2$O$_3$ reduction rates, achieved at the low temperature of 500 °C, demonstrate that fine iron ore particles can directly be converted into Fe without prior sintering by replacing the blast furnace with a vibro-fluidized bed and direct H$_2$ reduction. Moreover, the application of solar heat totally eliminates the use of fossil fuels, although the fossil fuel-fired operation is also possible and can be operated at a low temperature. The potential energy savings of the direct reduction processes are considerable since neither sintering nor high-temperature blast furnace reduction is required. Since these latter processes require significant applications of fossil fuels, the CO$_2$ savings are evident.

The remaining important issue is the production of H$_2$, preferably by a "green" production method, as discussed below. The current H$_2$ production methods rely mostly on steam reforming of CH$_4$, on the gasification of coal or biomass, and on the electrolysis of water to a minor extent. Producing H$_2$ by non-fossil fuel resources and by avoiding gasification should therefore be of primary importance.

4. Potential “Green” H$_2$ Production Methods

4.1. NH$_3$ Decomposition

Liquid NH$_3$ can be considered an H$_2$ storage medium and can be decomposed to N$_2$ and H$_2$. Although this decomposition is currently being tested, the literature provides sufficient background data to assess the reaction feasibility.

Electrolysis and thermo-chemical processes can be utilized for NH$_3$ decomposition. The thermo-chemical decomposition of NH$_3$ involves a catalyst inside a reactor using a reaction as follows:

$$\text{NH}_3 \rightleftharpoons 0.5 \text{N}_2 + 1.5\text{H}_2 \quad \Delta H^o = 46.2 \text{ kJ/mol}$$

The decomposition reaction is an endothermic thermochemical reaction that requires high temperatures, around 250–700 °C, and relatively high NH$_3$ partial pressures [36].

Theoretically, the thermodynamic equilibrium calculation produces the single-pass conversion rate at different decomposition temperatures, as presented in Table 4. The
decomposition can reach a complete reaction with a conversion rate of 99% at a temperature of at least 400 °C. Higher temperatures above 450 °C no longer strongly influence the reaction but rather its kinetics [36].

Table 4. Equilibrium NH$_3$ decomposition (at 1 bar) as a function of temperature, with original data from [36].

| Temperature (°C) | NH$_3$ Conversion (%) |
|-----------------|-----------------------|
| 300             | 95.7                  |
| 400             | 99.1                  |
| 500             | 99.7                  |
| 600             | 99.9                  |

Recent studies reported various single metal catalyst, in terms of activity, in the decreasing order Ru > Ni > Rh > Co > Ir > Fe > Pt > Cr > Pd > Cu >> Te, Se, Pb [37]. However, even if Ru shows the best conversion activity, this type of catalyst is still not sufficiently active for NH$_3$ decomposition at a temperature of 450 °C. More importantly, it will not be economically viable for large-scale applications. Multiple sources state that the Ru-based catalyst is active at a temperature of decomposition over 600 °C with a conversion rate of 97%.

Research and development of catalysts are done on precious metals, bimetallic materials, and metal oxides [38–42]. The bimetallic compound is proven to improve catalytic performance even though the manufacturing is complicated. It was also shown that metal oxides work by changing the electronic density of the active catalyst sites, and hence, are proposed as promoters and supporting materials for catalysts.

4.2. C$_2$H$_5$OH and CH$_3$OH Catalytic Decomposition

The catalytic decomposition of ethanol was studied in both the electric and solar furnaces in Figures 5–7. Results are shown in Figure 10 below for an ethanol/water mixture (1 to 6 mol-ratio), at a feed rate of 0.1 mL/min of ethanol/water and using two different catalysts, i.e., MnFe$_2$O$_4$ and a Co-loaded Al$_2$O$_3$.

The produced gas stream was continuously monitored by a gas chromatograph. The lag time of about 25 min is due to the piping between the reactor outlet and the GC. H$_2$, CO, CO$_2$ and CH$_4$ were the main products identified; acetone was also detected. The reaction mechanisms are currently under investigation.

A similar procedure was applied for the steam-methanol decomposition, with results illustrated in Figures 11 and 12. No coking of the catalyst was observed, and up to 5 repeat experiments for the same length of time and at successive days yield the same H$_2$ production rates.

The average results of three repeat experiments are used in the figures below. The deviation between the repeat results and the averages were below 10%, accepted as fair in view of the lag time in the reactor-to-GC piping.
Figure 10. Catalytic decomposition of ethanol-water (1:6) mixture using Co-Al$_2$O$_3$ or MnFe$_2$O$_4$ catalyst. Results obtained in the electric furnace at the given temperatures.

Figure 11. Catalytic decomposition of methanol-water (1:3) mixture using Co-Al$_2$O$_3$ or MnFe$_2$O$_4$ catalyst. Results obtained in the electric furnace at the given temperatures.
The results obtained demonstrate that excellent ethanol/methanol to H₂ conversions are achieved at low temperatures (300 to 500 °C is sufficient). In all the cases studied, the ratio of volumetric H₂ to CO₂ production was very high. Most of the feedstock carbon was detected as CO, CH₄ and acetone.

### 4.3. Tentative Techno-Economic Comparison of Different H₂ Production Technologies

From the experimental data and literature reports, the following comparison emerges (Table 5) for the levelized cost of hydrogen production (LCOH), for the GHG emission reduction potential and for the H₂ yield. Current production techniques of methane steam reforming, coal and biomass gasification and water electrolysis (using a renewable electricity source from wind turbines or photovoltaics) are compared with the novel methanol, methane and ethanol decomposition reactions.

**Table 5.** Techno-economic comparison (2017–2021).

|                        | Methane Steam Reforming | Coal Gasification | Membrane Electrolysis (Renewable Energy) | Biomass Gasification | Methanol Steam Reforming | Ethanol Steam Reforming | Methane Decomposition |
|------------------------|-------------------------|-------------------|-----------------------------------------|----------------------|--------------------------|------------------------|------------------------|
| **LCOH ($/kg H₂)**     | 1.25                    | 1.5               | 7.7                                     | 3                    | 1.5–2.0                  | 1.0–1.5                |
|                        |                         |                   | (for large scale)                        | Otherwise > 5        | (if acetone can be recovered and sold) | (if carbon black can be sold at 250–1000 $/ton) |
| **GHG emission**       | 8.1–11                  | 13–17             | 0                                       | +/−0                 | >7                       | 5–7                    | 2.5                    |
| (kg CO₂-eq/kg H₂)      |                         |                   |                                         |                      |                          |                        |                        |
| **H₂ yield (%)**       | 70–85                   | 50–60             | 70                                      | 20–40                | >90                      | 60–80                  | 75–80                  |

From this summary table, it is clear that (i) biomass and coal gasification should be avoided because of the low H₂ yield and fairly high LCOH; (ii) all other systems have a fairly similar H₂ yield, although the GHG emission levels disfavor methane steam reforming and coal gasification; methane, methanol and ethanol decomposition reactions can reduce the CHG emission levels, while still achieving a competitive LCOH.
5. Conclusions

Since steel mills need to reduce their carbon footprint, the traditional high-temperature fossil-fuel-fired (>1200 °C) processes need to be avoided. The solution of the direct reduction of iron ore with hydrogen has high potential: (i) the process has a high conversion efficiency (up to 98%) at low temperatures (500–600 °C) and both the use of H₂ and the significantly lower operating temperatures reduce the energy consumption and fossil-based CO₂ emissions by over 75%; (ii) H₂ can be generated by competitive and environmentally friendly processes, as examined in the research.

The limitations of the present research must, however, be considered.

(i) Fe₂O₃ ore particles are of small size and cannot readily be processed by a H₂ stream in a fixed or fluidized bed. A novel vibrated fluidized bed was developed and shown capable of fluidizing 1μm-particles at very low superficial fluidization velocities. The small-scale experiments need to be confirmed on a pilot-scale. Vibrated fluidization of the fine Fe₂O₃ particles caused good fluidization and yielded a nearly constant pressure drop. An increase in external force due to vibration slightly increased the pressure drop. Despite an increase in the reaction temperature, vibration produces a nearly constant minimum fluidization velocity.

(ii) Although very interesting reduction reaction yields for fine Fe₂O₃ particles were obtained on a small-scale, it is important to investigate the production of such fine Fe₂O₃ feedstock on an industrial scale.

(iii) The remaining issue relates to the “green” production of H₂, currently based primarily on fossil fuel feedstock with a high degree of GHG emissions. Although steam reforming of ethanol and methanol offers an important potential, the methane process is preferred since its GHG emission potential is the most attractive (except for electrolysis and biomass gasification) while producing H₂ at a very interesting cost. Further pilot-scale experiments are required to confirm these initial lab-scale results.

(iv) To confirm the solar production potential, additional experiments will be conducted in the solar furnace as soon as the Direct Normal Irradiance values are higher (from July onwards).

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