Luminescence properties in relation to controllable morphologies of Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ and its thermal conversion product Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$

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Three types of morphologies of Eu$^{3+}$-doped Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O) phosphors were obtained via hydrothermal reactions by different kinds of raw materials. In addition, Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ phosphors were obtained by calcining the precursor Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$. The structure and morphology of the obtained samples were characterized by XRD, EDS, FT-IR, TG-DTA, SEM and HRTEM. Herein, the effects of the synthesis parameters, including the reaction temperature, boron sources and the rare earth doping dosage, on the photoluminescence (PL) properties of Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ were investigated in detail. The lifetime and absolute quantum yield (QY) of different morphologies of Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ were also measured. The PL properties of the Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ phosphor prepared by the precursor calcination method compared with those prepared by the high-temperature solid-state method are discussed.

Introduction

In the past, many borates were chosen as host lattices for phosphors because of their large band gap, high thermal stability, high luminescence efficiency and low cost.\textsuperscript{1-5} Borogermanates, as one kind of borates, have different chemical compositions and structures derived from the flexible coordination geometries for both B and Ge atoms. The use of borogermanates with alkali, alkaline-earth and transition metals has gained considerable interest due to their excellent Second Harmonic Generation (SHG) coefficients, high optical-damage thresholds, photocatalytic applications and excellent thermal stability.\textsuperscript{6-12} Moreover, borogermanates can be a promising luminescence host material. The luminescence properties of some rare earth borogermanates doped with different kinds of rare earth metals, such as LaBGeO$_{35}$, EuGeBO$_{5}$ and Ln$_2$GeB$_2$O$_3$, have been studied.\textsuperscript{13-17}

The properties and applications of luminescent materials are strongly dependent on their chemical composition, crystal structure, size and morphology.\textsuperscript{18} For instance, Zhu et al. have investigated the effects of trace alkali metal ion (Li$^+$, Na$^+$, K$^+$) doping on the morphology and photoluminescence properties, and they deemed that the morphology influences the luminescence properties by geometric effects. Scattering occurs more frequently because of the distinctive particle shapes and surface roughness, which results in less absorption and decreased luminous efficiency. After doping with the alkali metal, the particle surface becomes smooth and the particle size gets reduced, meanwhile decreasing the number of surface-related defects. It leads to weaker nonradiative transitions and longer decay time afterwards.\textsuperscript{19} Li et al. have developed a series of one-dimensional Zn$_2$GeO$_3$:Mn$^{2+}$ phosphors with H$_2$O/EG as a solvent; they thought that the smaller particle size generally presents larger surface area and more defects in the host matrix, which may result in a weaker luminescence. Therefore, the Zn$_2$GeO$_3$:Mn$^{2+}$ nanobelts exhibit broader diameters and larger particle size, which induces lower surface area and leads to fewer defects and stronger luminescence.\textsuperscript{20} Wang et al. have synthesized various Tb$^{3+}$-doped calcium borate micro/nanostructures, such as Ca$_4$B$_{10}$O$_{19}$·7H$_2$O micro/nanoflakes, Ca$_3$B$_2$O$_5$·H$_2$O micro/nanorods, hierarchical micro–nanostructures, monodispersed nanobelts and Ca$_5$B$_2$O$_8$ polyhedron-type particles, via a facile hydrothermal method. The influence of the morphology and crystalline phase on the photoluminescence of Tb$^{3+}$ ions in the borate micro-nanostructures was investigated. These results indicate that both the crystal lattice and the morphology have important influence on the emission of Tb$^{3+}$\textsuperscript{21,22} Zhu et al. have synthesized red phosphors of the Mn$^{4+}$-doped dodec-fluoride Li$_3$Na$_3$Ga$_2$F$_2$Mn$^{4+}$ (LNGF:Mn) with various morphologies. The micro-morphologies of LNGF:Mn change from flake to cube and to polyhedron shapes with an increase in the HF concentration from 10 to 40 wt%. The luminescence intensity of LNGF:Mn increases linearly with...
the increase in the HF concentration, which might be jointly caused by the formation of higher concentrations of luminescence centers and the polyhedrons has better crystallinity. Hence, controlling the size and morphology of the phosphors precisely enables us to manipulate their properties as desired. In addition, the properties are usually dependent on the preparation methods and the reaction parameters. To date, many researchers have prepared rare earth ion-doped inorganic phosphors of different morphologies only by changing the raw reaction materials. For instance, Ding et al. have prepared Eu$^{3+}$-doped InBO$_3$ phosphors with four kinds of morphologies (microsphere, flat microsphere with a hole, octahedron and oblique hexahedron) via a facile flux method only by adjusting the kinds of reaction boron source materials. Yue et al. have prepared Ln$^3+$ (Ln = Tb, Eu)-doped zinc phosphate with different morphologies (regular nanosheet, fractured nanosheet, irregular nanosheet, and self-assembly microflower) only by altering the phosphate sources. Zou et al. have synthesized YBO$_3$:Eu$^{3+}$ phosphors with different morphologies (flower-like and nano-scaled particles) by simply adjusting the B$^{3+}$ sources.

Ba$_3$(Ge$_2$B$_7$O$_{16}$(OH)$_2$)(OH)(H$_2$O):Eu$^{3+}$ and Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ are potential SHG materials. However, there are no reports on the use of these two barium borogermanates as matrix luminescence materials. In this paper, we first prepared Eu$^{3+}$-doped Ba$_3$Ge$_2$B$_7$O$_{16}$(OH)$_2$)(OH)(H$_2$O):Eu$^{3+}$ with rhombohedron-like morphology, thorn-ball-like morphology and hexagonal-flakes-like morphology by the hydrothermal method and we also prepared Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ phosphor by the precursor calcination method. The effects of the experimental parameters, including the rare earth doping dosage, boron sources, and reaction time, on the photoluminescence (PL) properties of Ba$_3$(Ge$_2$B$_7$O$_{16}$(OH)$_2$)(OH)(H$_2$O) and Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ were investigated in detail. In addition, the PL properties of Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ prepared by two methods are compared and discussed.

Experimental
Synthesis of samples
Preparation of Ba$_3$(Ge$_2$B$_7$O$_{16}$(OH)$_2$)(OH)(H$_2$O):Eu$^{3+}$ phosphors. All the reagents were of analytical grade and used directly without further purification. In a typical procedure, the raw materials BaCO$_3$ (0.7425 mmol, 0.1465 g), Eu$_2$O$_3$ (0.0038 mmol, 0.0013 g), GeO$_2$ (0.5 mmol, 0.052 g), H$_2$BO$_3$ (1.75 mmol, 0.108 g) and H$_2$O (2.0 mL) were sealed in an autoclave with a Teflon lining. The autoclave was then placed in an oven and maintained at 200 °C for 4 days. After cooling to room temperature, the products were obtained by filtering, washing with deionized water and ethanol for several times, and drying at 60 °C for 12 h.

For the synthesis of Ba$_3$(Ge$_2$B$_7$O$_{16}$(OH)$_2$)(OH)(H$_2$O):Eu$^{3+}$ phosphors of other morphologies, the procedures were the same as stated above except that the boron source of H$_2$BO$_3$ was replaced with an equal amount of NH$_4$B$_5$O$_8$·4H$_2$O (0.35 mmol, 0.0953 g) and Na$_2$B$_4$O$_7$·10H$_2$O (0.4375 mmol, 0.1669 g), respectively.

Preparation of Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ phosphors. Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ phosphors were prepared by calcining the Ba$_3$(Ge$_2$B$_7$O$_{16}$(OH)$_2$)(OH)(H$_2$O):Eu$^{3+}$ precursor at 900 °C for 5 h in air. For comparison, the sample Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ was also synthesized by a conventional high-temperature solid-state reaction method at 900 °C for 5 h according to the literature.

Characterization
The structures of the prepared samples were characterized by X-ray powder diffraction (XRD, Rigaku D/MAX-C, Cu K$_\alpha$ radiation at 40 kV and 30 mA), FT-IR spectroscopy (recorded over the 400 to 4000 cm$^{-1}$ region using a Nicolet NEXUS 670 spectrometer with KBr pellets at room temperature), TG-DTA (TA-SDT Q600, in a N$_2$ atmosphere at a heating rate of 10 K min$^{-1}$), and energy-dispersive X-ray spectrometer (EDS, Hitachi, JEOl-6700F). The morphology and size of the product were investigated by scanning electron microscopy (SEM, Philips-FEI, Quanta 200). The inter-planar spacing of the crystal stripes of the sample was characterized by field transmission electron microscopy (TEM) (Tecnai G2 F20, FEI). The photoluminescence (PL) spectra were recorded at room temperature using a molecular fluorescence spectrometer (FL-4600, Hitachi) equipped with a 150 W xenon lamp as an excitation source. The lifetime was recorded using an FLS920P Edinburgh Analytical Instrument apparatus equipped with a 450 W xenon lamp and a µF900H high-energy micro-second flash lamp as the excitation sources. The absolute photoluminescence quantum yield (QY) was measured by a Hamamatsu (C9920-02G) analytical instrument apparatus using an integrating sphere.

Results and discussion
Structure identification of the samples
Fig. 1A shows the powder XRD pattern of the as-synthesized Ba$_3$(Ge$_2$B$_7$O$_{16}$(OH)$_2$)(OH)(H$_2$O):Eu$^{3+}$ with three kinds of boron sources (a) H$_2$BO$_3$, (b) NH$_4$B$_5$O$_8$·4H$_2$O and (c) Na$_2$B$_4$O$_7$·10H$_2$O, and Ba$_3$(Ge$_2$B$_6$O$_{16}$(OH)$_2$)(OH)(H$_2$O):Eu$^{3+}$ single-crystal structure simulated pattern. Fig. 1B shows the powder XRD pattern of Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ samples with two different methods: (a) calcination precursor method and (b) high temperature solid-state method, and Ba$_3$Ge$_2$B$_6$O$_{16}$ single-crystal simulated pattern. It can be observed that all the diffraction peaks in Fig. 1A and B are in good agreement with the monoclinic phase Ba$_3$(Ge$_2$B$_7$O$_{16}$(OH)$_2$)(OH)(H$_2$O) and the triclinic phase Ba$_3$Ge$_2$B$_6$O$_{16}$ respectively. No impurity peaks were observed. A small amount of doped Eu$^{3+}$ ions did not cause any significant change in the crystal structure of the host Ba$_3$(Ge$_2$B$_7$O$_{16}$(OH)$_2$)(OH)(H$_2$O) and Ba$_3$Ge$_2$B$_6$O$_{16}$.

Fig. S1a and b† show the EDS pattern of Ba$_3$(Ge$_2$B$_7$O$_{16}$(OH)$_2$)(OH)(H$_2$O):Eu$^{3+}$ and Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$, respectively. The EDS patterns clearly indicate that the two obtained samples contain B, O, Ba, Ge and Eu elements, and there are no other impurity element peaks.

Fig. 2a and b show the FT-IR spectra of Ba$_3$(Ge$_2$B$_7$O$_{16}$(OH)$_2$)(OH)(H$_2$O):Eu$^{3+}$ and Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$, respectively. In Fig. 2a, the bands at 3427 cm$^{-1}$, 3167 cm$^{-1}$ and 1629 cm$^{-1}$ are
attributed to the stretching and bending vibrations of O–H. The absorption peak at 1383 cm$^{-1}$ can be assigned to the asymmetrical stretching of B(3)–O. The peak at 1260 cm$^{-1}$ is attributed to the in-plane bending vibration of B–O–H. The peaks at 1130 cm$^{-1}$ and 918 cm$^{-1}$ are attributed to the symmetric stretching of B(4)–O. The peaks at 870 cm$^{-1}$ and 836 cm$^{-1}$ are attributed to the asymmetric stretching of Ge(4)–O. The peak at 487 cm$^{-1}$ is attributed to the flexural vibration of Ge(4)–O. The peaks at 657 cm$^{-1}$ and 576 cm$^{-1}$ are attributed to the flexural vibration of B(3)–O and B(4)–O. In Fig. 2b, the absorption peaks at 1370 cm$^{-1}$ and 1260 cm$^{-1}$ can be assigned to the asymmetrical stretching of B(3)–O. The peaks at 1062 cm$^{-1}$ and 924 cm$^{-1}$ are attributed to the symmetric stretching of B(4)–O. The peaks at 871 cm$^{-1}$, 832 cm$^{-1}$ and 773 cm$^{-1}$ are assigned to the asymmetric stretching of Ge(4)–O. The peak at 459 cm$^{-1}$ is assigned to the flexural vibration of Ge(4)–O. The peaks at 610 cm$^{-1}$ and 583 cm$^{-1}$ are assigned to the flexural vibration of B(3)–O and B(4)–O. The corresponding absorption bands of Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ and Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ are similar to those reported in the literature, respectively.$^{26}$ Therefore, we conclude that we obtained the Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ and Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ phosphors.

Fig. S2† shows the simultaneous TG-DTA curves of Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$. TG studies indicate that there is no weight loss before 673 K, and it exhibits the initial step of weight loss in the range of 673–1073 K, which corresponds to the release of 2.5 water molecules per formula unit. The observed weight loss of 4.8% matches well with the theoretical calculation result (4.7%). This assignment is also in agreement with the endothermic peak at 883 K in the DTA diagram. This results indicated that Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ is thermally stable up to 673 K.

Fig. 3 shows the SEM images of the as-synthesized Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ samples with different boron sources as raw reaction materials. It is apparent that Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ synthesized with H$_3$BO$_3$ as the boron source displays rhombohedron-like morphology with an edge of about 2 μm (shown in Fig. 3a1 and a2). Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ synthesized with NH$_4$B$_5$O$_8$·4H$_2$O as the boron source displays uniformly sized thorn-ball-like morphology, which is constructed by lots of nano-particles, with diameters of about 10 μm (shown in Fig. 3b1 and b2). Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ synthesized with Na$_2$B$_4$O$_7$·10H$_2$O as the boron source displays hexagonal-flake morphology with an edge of about 1 μm and a thickness of about 100 nm (shown in Fig. 3c1 and c2). Different morphologies can be obtained by using different boron sources. When the different boron sources were introduced into the reaction system, the pH values and the existing cations were different; the acidity of H$_3$BO$_3$ is lower than those of NH$_4$B$_5$O$_8$·4H$_2$O and Na$_2$B$_4$O$_7$·10H$_2$O. The pH value can affect the kinds and qualities of the absorbed ions on the surface of the crystal.$^{23}$ Thus, the surface energy of the crystal is different. According to the Gibbs–Curie–Wulff theorem, the growth rates on different surface facets are dominated by the surface energy.$^{27}$ As a result, the different growth rates direct the different growth directions.

Fig. 2 FT-IR spectrum of samples: (a) Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ and (b) Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$. 

Fig. 1 (A) The powder XRD pattern of Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ with different synthetic methods: (a) H$_3$BO$_3$, (b) NH$_4$B$_5$O$_8$·4H$_2$O, and (c) Na$_2$B$_4$O$_7$·10H$_2$O. (B) The XRD pattern of Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ obtained by different methods: (a) calcination precursor method and (b) high temperature solid-state method.
and further direct the different final morphology of the obtained samples.

HREM images of the three different morphologies of Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ samples are displayed in Fig. 4. The inter-planar spacing of the rhombohedron-like morphology sample is 4.258 Å, which is consistent with the $d$ value (4.258 Å) of (020) faces in the Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O) crystal. The inter-planar spacing of the thorn-ball-like morphology sample is 3.268 Å, which is very close to the $d$ value (3.274 Å) of (221) faces in Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O) crystal. The inter-planar spacing of the hexagonal-flake morphology sample is 2.840 Å, which is close to the $d$ value (2.855 Å) of (222) faces in the Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O) crystal.

Fig. S3† shows the SEM images of Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ prepared by different reaction methods: (a) precursor method; and (b) solid-state method. It can be seen that the Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ sample prepared by the precursor method exhibits porous surface block structure. In addition, the Ba$_3$Ge$_2$B$_6$O$_{16}$:Eu$^{3+}$ sample prepared by the solid-state method exhibits irregular block structure.

Fig. 3  SEM images of Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ with different raw materials: (a$_1$ and a$_2$) H$_3$BO$_3$, (b$_1$ and b$_2$) NH$_4$B$_5$O$_8$·4H$_2$O and (c$_1$ and c$_2$) Na$_2$B$_4$O$_7$·10H$_2$O.

Fig. 4  High-resolution TEM images of Ba$_3$[Ge$_2$B$_7$O$_{16}$(OH)$_2$](OH)(H$_2$O):Eu$^{3+}$ with different morphologies: (a) rhombohedron-like morphology, (b) thorn-ball-like morphology and (c) hexagonal-flake morphology.
Photoluminescence properties

Effect of doped Eu$^{3+}$ concentration on the PL intensity of Ba$_3$[Ge$_2$B$_{16}$O$_{44}$]OH(H$_2$O)Eu$^{3+}$. The PL and PL spectra under UV excitation for the hexagonal-flake morphology of Ba$_3$[Ge$_2$B$_{16}$O$_{44}$]OH(H$_2$O)Eu$^{3+}$ ($x$ = 1%, 3%, 5%, and 7%, mole ratio) were recorded at room temperature. The excitation spectra of Ba$_3$[Ge$_2$B$_{16}$O$_{44}$]OH(H$_2$O)Eu$^{3+}$, recorded at an emission wavelength of 615 nm, are shown in Fig. 5 (inset). The excitation spectra exhibits a strong charge transfer band (CTB) from the 2p orbital of O$^2-$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transitions of Eu$^{3+}$.28 The excitation spectra exhibit a strong charge transfer band (CTB) from the 2p orbital of O$^2-$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transitions of Eu$^{3+}$.28 The excitation spectra exhibit a strong charge transfer band (CTB) from the 2p orbital of O$^2-$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transitions of Eu$^{3+}$.28 The excitation spectra exhibit a strong charge transfer band (CTB) from the 2p orbital of O$^2-$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transitions of Eu$^{3+}$.28 The excitation spectra exhibit a strong charge transfer band (CTB) from the 2p orbital of O$^2-$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transitions of Eu$^{3+}$.28 The excitation spectra exhibit a strong charge transfer band (CTB) from the 2p orbital of O$^2-$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transitions of Eu$^{3+}$.28

The emission spectra of the samples with different concentrations of doped Eu$^{3+}$ upon excitation at 242 nm wavelength. No remarkable differences in the spectroscopic characteristics were observed upon excitation at the same wavelength when different concentrations of Eu$^{3+}$ ions were doped. The peaks at the 589 nm and 600 nm are due to the split sublevels of the magnetic dipole transition $^5$D$_0$ → $^7$F$_1$, whose intensities are almost independent of the crystal field. The peaks at 615 nm originating from the $^5$D$_0$ → $^7$F$_2$ transition of Eu$^{3+}$ are the hypersensitive forced electric dipole transition, whose intensities are sensitive to the local environment around Eu$^{3+}$. The very weak $^5$D$_0$ → $^7$F$_1$ transitions can also be observed at 653 nm.29 The results show that the red emission $^5$D$_0$ → $^7$F$_2$ transition is dominant in intensity, and so, the Eu$^{3+}$ ions mainly occupied a non-inversion symmetry center in the crystal lattice. As the concentration of Eu$^{3+}$ increases, the emission associated with Eu$^{3+}$ continues to increase first, and above the 3% doping concentration, it decreases due to concentration quenching. The concentration quenching effect might be due to the enhanced excitation migration or due to the activator ion coupling or coagulation.30

Effect of the reaction time on the PL properties of Ba$_3$[Ge$_2$B$_{16}$O$_{44}$]OH(H$_2$O)Eu$^{3+}$. The excitation spectra of 3% Eu$^{3+}$-doped Ba$_3$[Ge$_2$B$_{16}$O$_{44}$]OH(H$_2$O)Eu$^{3+}$ with hexagonal-flake morphology, recorded at an emission wavelength of 615 nm, are shown in Fig. 7 (inset). The excitation spectra exhibit a strong charge transfer band (CTB) from the 2p orbital of O$^2-$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transitions of Eu$^{3+}$.28 The excitation spectra exhibit a strong charge transfer band (CTB) from the 2p orbital of O$^2-$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transitions of Eu$^{3+}$.28 The excitation spectra exhibit a strong charge transfer band (CTB) from the 2p orbital of O$^2-$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transitions of Eu$^{3+}$.28 The excitation spectra exhibit a strong charge transfer band (CTB) from the 2p orbital of O$^2-$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transitions of Eu$^{3+}$.28 The excitation spectra exhibit a strong charge transfer band (CTB) from the 2p orbital of O$^2-$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transitions of Eu$^{3+}$.28

The emission spectra of Ba$_3$[Ge$_2$B$_{16}$O$_{44}$]OH(H$_2$O)Eu$^{3+}$ prepared with different boron source raw materials under UV excitation are shown in Fig. 7. No remarkable differences in the spectroscopic characteristics of the samples with different morphologies were observed in the excitation spectra when synthesized using different boron source raw materials. The peaks at the 589 nm and 600 nm are originating from the $^5$D$_0$ → $^7$F$_1$ transition. The peak at 615 nm is originating from the $^5$D$_0$ → $^7$F$_2$ transition of Eu$^{3+}$.31 The very weak $^5$D$_0$ → $^7$F$_1$ transitions can also be observed at 653 nm. The sample prepared with Na$_2$B$_4$O$_7$.10H$_2$O as the boron source (hexagonal-flake morphology) exhibited the strongest intensity. It is widely accepted that the luminescence properties of the phosphors are strongly dependent on the morphology, size, crystallinity, the center of luminescence and the density of

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**Fig. 5** The emission spectral of Ba$_3$[Ge$_2$B$_{16}$O$_{44}$]OH(H$_2$O)Eu$^{3+}$ with different Eu$^{3+}$ doping concentrations ($\lambda_{ex}$ = 242 nm): (a) 1%, (b) 3%, (c) 5%, and (d) 7%.

**Fig. 6** The emission spectra of Ba$_3$[Ge$_2$B$_{16}$O$_{44}$]OH(H$_2$O)Eu$^{3+}$ prepared for different reaction times ($\lambda_{ex}$ = 242 nm): (a) 24 h, (b) 48 h, (c) 72 h, and (d) 96 h.
The quantum yield (QY) is defined as the ratio of the number of emitted photons to the number of absorbed photons. When comparing the QY values ($\lambda_{\text{ex}} = 394 \text{ nm}$) of the three morphologies of Ba$_3[\text{Ge}_2\text{B}_7\text{O}_{16}\text{(OH)}_2]\text{(OH)}\text{(H}_2\text{O)}\text{:Eu}^{3+}$ phosphors with those of some other Eu-doped compounds reported by different authors (given in Table 1), we can see that the obtained absolute QY of the hexagonal-flake morphology Ba$_3[\text{Ge}_2\text{B}_6\text{O}_{16}\text{(OH)}_2]\text{(OH)}\text{(H}_2\text{O)}\text{:Eu}^{3+}$ phosphor possesses the highest value (54.1%) than the rhombohedron-like morphology and thorn-ball-like morphology Ba$_3[\text{Ge}_2\text{B}_6\text{O}_{16}\text{(OH)}_2]\text{(OH)}\text{(H}_2\text{O)}\text{:Eu}^{3+}$ phosphors. Moreover, the QY of the hexagonal-flake morphology Ba$_3[\text{Ge}_2\text{B}_6\text{O}_{16}\text{(OH)}_2]\text{(OH)}\text{(H}_2\text{O)}\text{:Eu}^{3+}$ phosphor is equivalent to that of the Sr$_3\text{Y}_2\text{(BO}_3\text{)}_4\text{:Eu}^{3+}$ powder and SrO–La$_2\text{O}_3$–Al$_2\text{O}_3$–B$_2\text{O}_3$–SiO$_2$–Eu$^{3+}$ glass. Moreover, the QY of the hexagonal-flake morphology Ba$_3[\text{Ge}_2\text{B}_6\text{O}_{16}\text{(OH)}_2]\text{(OH)}\text{(H}_2\text{O)}\text{:Eu}^{3+}$ phosphor is much higher than that of the well-known commercial Y$_2\text{O}_3\text{:Eu}^{3+}$ phosphor. Therefore, it can be concluded that the hexagonal-flake morphology Ba$_3[\text{Ge}_2\text{B}_6\text{O}_{16}\text{(OH)}_2]\text{(OH)}\text{(H}_2\text{O)}\text{:Eu}^{3+}$ phosphor may be used as a new red phosphor for WLED.

**Comparing the PL properties of Ba$_3\text{Ge}_2\text{B}_6\text{O}_{16}\text{Eu}^{3+}$ phosphor prepared by different methods.** The excitation spectra of Ba$_3\text{Ge}_2\text{B}_6\text{O}_{16}\text{Eu}^{3+}$, recorded at an emission wavelength of 615 nm, is shown in Fig. 9A (inset). The excitation spectra exhibit a strong charge transfer band (CTB) from the 2p orbital of O$^{2-}$ to the 4f orbital of Eu$^{3+}$ with a maximum peak at about 242 nm. Other weak peaks are attributed to the f-f transition of Eu$^{3+}$. The emission spectra of Ba$_3\text{Ge}_2\text{B}_6\text{O}_{16}\text{Eu}^{3+}$ phosphor prepared by the precursor method and the solid-state method are shown Fig. 9A. It can be observed that the strongest emission peak for Ba$_3\text{Ge}_2\text{B}_6\text{O}_{16}\text{Eu}^{3+}$ obtained by the precursor method was located at 614 nm, which was ascribed to red emission, and Eu$^{3+}$ was mainly in the non-inversion center sites; meanwhile the strongest emission peak for the Ba$_3\text{Ge}_2\text{B}_6\text{O}_{16}\text{Eu}^{3+}$ obtained by high-temperature solid-state reaction was located at 591 nm, which was ascribed to orange emission, and Eu$^{3+}$ was mainly in the inversion center sites. As shown in Fig. 9B, there are two types of surrounding environments for the Ba atoms in Ba$_3\text{Ge}_2\text{B}_6\text{O}_{16}$, in which Ba1 is 6-coordinated by 6 O atoms, and it mainly occupies the superior symmetry site, whereas Ba2 is 7-coordinated by 7 O atoms, and it mainly occupies the inferior symmetry site. Therefore, we can conclude that the Eu$^{3+}$ occupies the Ba2 site for Ba$_3\text{Ge}_2\text{B}_6\text{O}_{16}\text{Eu}^{3+}$ obtained by the
precursor method, whereas Eu\(^{3+}\) occupies the Ba1 site for Ba\(_3\)Ge\(_2\)B\(_6\)O\(_{16}\):Eu\(^{3+}\) obtained by the high-temperature solid-state reaction. The corresponding variation in the CIE chromaticity coordinates under irradiation using a 242 nm UV lamp is demonstrated in Fig. 9C. The color tone was red (a) (0.524, 0.312) and orange (b) (0.434, 0.288) in the Ba\(_3\)Ge\(_2\)B\(_6\)O\(_{16}\):Eu\(^{3+}\) phosphors by two different preparation methods.

**Conclusions**

Three different morphologies of Ba\(_3\)[Ge\(_2\)B\(_7\)O\(_{16}\)(OH)\(_2\)](OH)(H\(_2\)O):Eu\(^{3+}\) have been obtained by a hydrothermal reaction by varying the type of boron sources, for the first time. The effects of the doping amount of Eu\(^{3+}\), synthetic raw materials and hydrothermal time on the luminescent properties were studied. The result indicated that the different reaction raw materials have a greater influence on the crystallinity and morphology of Ba\(_3\)[Ge\(_2\)B\(_7\)O\(_{16}\)(OH)\(_2\)](OH)(H\(_2\)O):Eu\(^{3+}\), and further influence the luminescence property. The lifetime of the Ba\(_3\)[Ge\(_2\)B\(_7\)O\(_{16}\)(OH)\(_2\)](OH)(H\(_2\)O):Eu\(^{3+}\) phosphor with hexagonal-flake morphology is 1.67 ms. The absolute QY of the obtained Ba\(_3\)[Ge\(_2\)B\(_7\)O\(_{16}\)(OH)\(_2\)](OH)(H\(_2\)O):Eu\(^{3+}\) with hexagonal-flakes-like morphology is 54.1, which is much higher than that of the commercial Y\(_2\)O\(_3\):Eu phosphor. When the Ba\(_3\)[Ge\(_2\)B\(_7\)O\(_{16}\)(OH)\(_2\)](OH)(H\(_2\)O):Eu\(^{3+}\) precursor was calcined at high temperatures, we can get Ba\(_3\)Ge\(_2\)B\(_6\)O\(_{16}\):Eu\(^{3+}\) phosphor. By comparing the luminescence properties of Ba\(_3\)Ge\(_2\)B\(_6\)O\(_{16}\):Eu\(^{3+}\) obtained by the precursor calcination method and that obtained by the high-temperature solid-state reaction, it was found that the strongest emission peak for Ba\(_3\)Ge\(_2\)B\(_6\)O\(_{16}\):Eu\(^{3+}\) obtained by the precursor calcination method appeared at 614 nm, which exhibits red emission (0.524, 0.312), meanwhile the strongest emission peak for Ba\(_3\)Ge\(_2\)B\(_6\)O\(_{16}\):Eu\(^{3+}\) obtained by the high-temperature solid-state reaction was located at 591 nm, which exhibits orange emission (0.434, 0.288).

**Conflicts of interest**

There are no conflicts to declare.

**Acknowledgements**

This work was supported by the Talent Introduction Research Project of Shaanxi Xueqian Normal University (2017DS02), the Science Foundation of Shaanxi Xueqian Normal University.
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