A long-lifespan, flexible zinc-ion secondary battery using a paper-like cathode from single-atomic layer MnO2 nanosheets

Yanan Wang, Zeyi Wu, Le Jiang, Wenchao Tian, Chenchen Zhang, Cailing Cai and Linfeng Hu

Aqueous zinc ion secondary batteries (ZIBs) have recently attracted considerable attention and global interest due to their low cost, aqueous-based nature and great safety. Unfortunately, the intrinsic properties of poor cycle life, low energy density and uncontrolled dendrite growth during the charge/discharge cycling of ZIBs significantly hinder their practical application. In this work, we rationally designed two-dimensional (2D) 3d-MnO2 nanofluidic channels by the ordered restacking of exfoliated MnO2 single atomic layers, which exhibited a high zinc ion transport coefficient (1.93 × 10^{-14} cm^2 s^{-1}) owing to their appropriate d-spacing and the negative charge of the inner channel walls. More importantly, we found that Zn dendrite growth was prevented in the as-assembled ZIBs, resulting in superior stability compared with the bulk-MnO2 sample. Our design sheds light on developing high-performance ZIBs from two-dimensional nanofluidic channels, and this strategy might be applicable to the storage of other metal ions (Mg^{2+}, Ca^{2+}, Al^{3+}, etc.) in next-generation electrochemical energy storage devices.

1. Introduction

Owing to the high abundance of zinc, high theoretical capacity (820 mA h g^{-1}), low electrochemical potential (∼0.762 V versus the standard hydrogen electrode (SHE)), intrinsic safety and environmental friendliness, zinc ion rechargeable batteries (ZIBs) with mild aqueous electrolytes have attracted considerable attention and global interest.1-6 Nevertheless, it is still at an initial stage for ZIBs and their reported electrochemical performance has still remained far from practical requirements. The main challenges should be attributed to Zn dendrite formation during charge/discharge cycling. Similar to many other metals such as Li, Ni and Cu, zinc generally tends to deposit in the dendrite form especially at high current densities, which results in low coulombic efficiency, rapid degradation of capacity and poor stability. More seriously, the needle-like Zn dendrites may puncture the separator, leading to short circuits and safety concerns.7,8

Theoretically, the origin of dendrite growth can be ascribed to the metal-ion concentration gradient and tip effect during electroplating.9-12 In general, the depletion of metal ions in the electrolyte caused by the concentration gradient breaks the electrical neutrality at the plated electrode surface, leading to a local space charge and ramified Zn electro stripping/plating on the screw dislocations of the Zn surface.13 Extensive efforts have been devoted to solving the dendrite growth problem in ZIBs.14-17 Typical strategies, such as constructing Zn anodes with a hierarchical architecture and preparing a composite Zn anode coating as a thin inert layer to stabilize the Zn anode, have been developed to suppress dendritic growth in ZIBs.18 Most recently, Wang et al. achieved great progress on dendrite-free plating/stripping of Zn in a highly concentrated Zn(TFSI)_2 electrolyte with a supporting salt of LiTFSI (bis(trifluoromethanesulfonyl) imide).19 Nevertheless, either the complicated fabrication process of the inert layer or the high cost of the highly concentrated Zn-ion electrolyte hinders widespread application and industrial production. Thus, it is still highly desirable to design a new strategy to realize dendrite-free ZIBs with excellent reversibility and stability.

Two-dimensional (2D) nanostructures with a single-atomic layer thickness have attracted intensive attention for the development of energy-storage devices because of the emergence of novel physical and chemical properties which are distinct from those of or even absent in their bulk counterparts.18-21 In particular, 2D nanostructure units are rather soft and flexible and have been considered ideal building blocks of freestanding, flexible membrane samples as high-performance electrodes.22,23 Such freestanding electrodes are completely composed of 2D active materials without any current collectors and polymer binders, leading to a superior performance due to
the unique structure and significantly decreased interfacial resistance. More importantly, such a restacked structure in the freestanding, flexible membrane fabricated from the 2D building blocks could produce lamellar nano-channels and atomically smooth surfaces for nanofluids such as liquid or quasi-solid-state electrolytes. It has been verified that 2D nano-channels obtained by exfoliation and restacking of layered compounds are ideal frameworks for fast Li-ion storage because of their abundant interconnected (horizontal and vertical) channels, large active surface and short diffusion path.

Most recently, flexible ZIBs based on binder-free \( \alpha \)-MnO\(_2\) nanosheets/rGO on carbon cloth have been realized. However, the \( \alpha \)-MnO\(_2\) nanosheets in this work were synthesized by a hydrothermal route and their thickness is much larger than the single layer thickness. The resulting \( \alpha \)-MnO\(_2\) nanosheets/ rGO thin-film generally exhibits a rough surface and a loose, overlapping structure with the absence of 2D nanochannel characteristics. We note that the recent advances in layered-MnO\(_2\) demonstrate that it can be exfoliated into 2D negatively charged nanosheets with single-atomic layer thickness by a soft-chemical strategy. It is possible to construct unique 2D nanochannels with controllable interlayer spacing from these 2D building blocks? What would happen in these 2D nanofluidic channels for \( \text{Zn}^{2+} \) transport? It would be of great interest to study the performance of ZIBs constructed from such freestanding nanochannels and further understand the \( \text{Zn}^{2+} \) storage mechanism.

Inspired by this consideration, herein, we demonstrate a proof-of-concept strategy to construct 2D nanofluidic channels from liquid-exfoliated MnO\(_2\) single-layer nanosheets (Fig. 1). The as-constructed paper-like Zn ion battery from these nanofluidic channels exhibited excellent flexibility and a high energy density up to 365 W h kg\(^{-1}\). We found that Zn dendrite growth was prevented in the paper-like ZIBs, resulting in superior stability compared with batteries based on a bulk-MnO\(_2\) cathode. Our design sheds light on developing high-performance ZIBs from two-dimensional nanofluidic channels, and this strategy might be applicable to storage of other metal ions (Na\(^+\), Mg\(^{2+}\), Ca\(^{2+}\), Al\(^{3+}\), etc.) in next-generation electrochemical energy storage devices.

### 2. Experimental section

#### 2.1 Preparation of the H\(_{0.08}\)MnO\(_2\)·0.75H\(_2\)O precursor

Potassium permanganate (KMnO\(_4\), GR, 6 mmol, Sinopharm Chemical Reagent Co. Ltd (SCRC)) was dissolved in 137.5 mL Milli-Q water (18.2 M\( \Omega \) cm) to form a purple solution, and potassium hydroxide (KOH, AR, 3 mol, SCRC) was added to this solution in two batches under stirring at 400 rpm and then mixed with 20 mL of a manganese(II) chloride tetrahydrate (MnCl\(_2\)·4H\(_2\)O, AR, 0.6 M, SCRC) aqueous solution and stirred for 2 h to obtain a uniform precursor solution. This solution was hydrothermally treated at 175 °C for 3 days. The manganese oxide precipitate was washed with Milli-Q water 4 times to neutral pH, after which ammonium persulfate ([NH\(_4\)]\(_2\)S\(_2\)O\(_8\), AR, 0.25 mol, SCRC) and the neutral solution were added into a three-necked flask at 60 °C for 3 h (repeated 3 times) and then filtered and washed with Milli-Q water and finally dried in air at room temperature. The ideal reaction mechanism of the synthesis is:

\[
K_{0.33}\text{MnO}_2 \cdot 0.47\text{H}_2\text{O} + [\text{NH}_4]_2\text{S}_2\text{O}_8 \rightarrow H_{0.08}\text{MnO}_2 \cdot 0.75\text{H}_2\text{O}
\]

#### 2.2 Exfoliation

For the exfoliation process, the as-synthesized H\(_{0.08}\)MnO\(_2\)·0.75H\(_2\)O (0.36 g) and tetramethylammonium hydroxide (TMAOH, 10.4 mL, Aladdin) were dispersed into 80 mL Milli-Q water. The as-obtained suspension was shaken for 24 h at 100 rpm and washed with Milli-Q water to neutral pH by the centrifugation (6000 rpm for 10 min) method. Subsequently, a further peeling-off process was implemented, in which the as-obtained suspension and tetrabutylammonium hydroxide (TBAOH, 0.74 mL, Aladdin) were dispersed into 90 mL Milli-Q water and then shaken for 72 h at 80 rpm to obtain exfoliated MnO\(_2\) nanosheets.

#### 2.3 Preparation of the MnO\(_2\)/MWCNT membrane

The as-obtained exfoliated MnO\(_2\) nanosheet suspension (≈40 mmol L\(^{-1}\)) was mixed with carbon nanotubes (industrial multi-walled carbon nanotubes (MWCNTs-010-0), Tanfeng Technology Co., Ltd) and an N-methylpyrrolidone (NMP, AR, SCRC) solution (0.5 mg mL\(^{-1}\)) at a volume ratio of 1:2 by mechanical grinding for 10 minutes and ultrasonic dispersion for 30 minutes. The MnO\(_2\)/MWCNT membrane was prepared by vacuum filtering the obtained suspension, washing it with Milli-Q water 3 times and drying it under ambient conditions for 3 days. The loading mass of MWCNTs on the MnO\(_2\)/MWCNT membrane was around 20%. The mass density of MnO\(_2\)/MWCNT electrodes was \( \sim 0.96 \text{ g cm}^{-3} \).
2.4 Preparation of the $\delta$-MnO$_2$ bulk
The $\delta$-MnO$_2$ bulk sample was obtained by heating KMnO$_4$ powder in an air atmosphere at 350 °C for 5 h and cooling naturally. The product was washed with Milli-Q water 3 times and dried at 80 °C for 12 h and finally thoroughly ground. The ZIB cathode made of the $\delta$-MnO$_2$ bulk was prepared according to the following process: active powder : Ketjen black : poly(vinylidene fluoride) (PVDVF) (mass ratio: 7 : 2 : 1) was mixed homogeneously, and then the slurry mixture was coated onto 304 stainless steel foil and dried under vacuum at 120 °C for 12 h. Then the stainless steel foil after coating was punched into circular pieces ($d = 15$ mm, mass density: 1.1 g cm$^{-3}$) for the assembly of coin-type cells.

2.5 Characterization
The tensile strength of the MnO$_2$/MWCNT membrane was measured using an AUTO tensile tester (Labthink, XLW (PC)). A field emission scanning electron microscope (FE-SEM, FEI Nova Nano SEM 450) and an atomic force microscope (AFM, Bruker Dimension Icon) were used to characterize the sample’s morphology. X-ray photoelectron spectroscopy (XPS, PerkinElmer PHI 5000C ESCA, with Mg K$_\alpha$ radiation, standard signal: C 1s peak 284.6 eV) analysis was performed to investigate the elemental distribution and chemical components. X-ray powder diffraction (XRD) analysis was performed to investigate the crystal structure (Bruker D8-A25 diffractometer using Cu K$_\alpha$ radiation). The scan rate was 2° min$^{-1}$. Thermogravimetric (TG) analysis combined with differential scanning calorimetry (DSC, SDT Q600, the heating rate was 10 °C min$^{-1}$ and the air flow was 50 mL min$^{-1}$) was adopted to explore the sample’s thermal behavior. Transmission electron microscopy (TEM) and selected area electron diffraction (SAED) were performed using a Philips CM 200 FEG field emission microscope. A zeta-potential tester (Zetasizer Nano ZS90) was used to characterize the surface charge state of the nanosheets dispersed in solvents.

2.6 Electrochemical measurements
The electrochemical performances of the assembled coin-type MnO$_2$/MWCNTs/zinc and MnO$_2$ bulk/zinc unit cells, such as the discharge capacity, discharge C-rate capability, and cycling performance, were evaluated using a LAND battery cycle system (Wuhan LAND Electronics Co., LTD, China) in the voltage range of 1.0–1.9 V under ambient conditions. For the MnO$_2$/MWCNTs/zinc paper-like battery, the MnO$_2$/MWCNT cathode and the high purity zinc foil (thickness of 30 µm, Tengfeng Metal Materials LTD) were punched into circular pieces ($d = 15$ mm), Whatman glass fiber (GF, thickness of 350 µm) was used as the separator, and a mixture of zinc sulfate (ZnSO$_4$·7H$_2$O, 99.995%, 2 M, SCRC) and manganese sulfate (MnSO$_4$·H$_2$O, AR, 0.2 M, SCRC) aqueous solutions was used as the electrolyte. The charge and discharge cycling performance of the cells was measured at 0.1 A g$^{-1}$ for 600 cycles and the C-rate capability of the cells was examined from 0.1 A g$^{-1}$ to 2 A g$^{-1}$. Cyclic voltammetry (CV) and electrochemical impedance analysis of the Zn$^{2+}$/Zn unit cells were performed on an electrochemical workstation (CHI660e) over the frequency range of 10 mHz to 1 MHz. The performance (energy and power densities) was estimated based on the weight of the active MnO$_2$ only.

3. Results and discussion
Negatively charged MnO$_2$ single-atomic layer nanosheets were prepared by a liquid-exfoliation strategy from the H-birnessite bulk.$^{37}$ A dark-brown colloidal suspension was obtained with a high concentration (~40 mmol L$^{-1}$) of the MnO$_2$ nanosheets in the water phase (Fig. 2a). Clear Tyndall light scattering of the colloidal suspension was observed in a diluted nanosheet suspension (Fig. 2b). Their negative surface charge was proven by the negative zeta potential measurement (~32 mV, Fig. S1, ESI†). Atomic force microscopy (AFM) characterization revealed a typical 2D nanosheet morphology with average lateral dimensions of 1–1.5 µm (Fig. 2c) and a thickness of 1.0 nm (Fig. 2d and e). Note that the theoretical thickness of the MnO$_2$ single-layer nanosheets is estimated to be 0.52 nm. The measured thickness is slightly larger than the theoretical thickness by approximately 0.4 nm, which might be due to the distance between the substrate and MnO$_2$ as well as between MnO$_2$ and the probe. This phenomenon has been observed in a previous study.$^{38}$ The transmission electron microscope (TEM) image also shows ultrathin sheet-like objects with lateral dimensions similar to those observed in the AFM image (Fig. S2, ESI†).

The ultrathin nature, nanosheet-like morphology and inherent flexibility are quite promising for constructing free-standing membranes by a facile vacuum-assisted filtration process.
strategy. In order to improve the electronic conductivity of this membrane, we consider the incorporation of conductive multi-walled carbon nanotubes (MWCNTs) into this membrane sample. Mixing the two colloidal solutions of MnO₂ nanosheets and MWCNTs with sonication for an additional 30 min produces a homogeneous dispersion with no apparent precipitation. Then, a hybrid, black MnO₂/MWCNT membrane was easily fabricated by the vacuum-assisted filtration strategy with a lateral size of 4 cm (Fig. 2f). The as-obtained membrane is highly flexible (Fig. 2g), and it can be freely rolled up by hand and using glass rods (Fig. S3, ESI†). The optimal mass ratio of MWCNTs in the composite MnO₂/MWCNT membrane is ~20% because a high concentration of MWCNTs over this threshold can drastically decrease the flexibility of the hybrid membrane. The top-view SEM image shows a rather smooth and dense surface of our MnO₂/MWCNTs membrane (Fig. S5, ESI†).

The rapid decrease in contact angle may provide evidence of the facile permeation and transfer of the aqueous electrolyte in our membrane (Fig. 2i). X-ray photoelectron spectroscopy (XPS) analysis confirms the existence of Mn, O, and C elements in this hybrid membrane (Fig. S6, ESI†). The formation of well-defined 2D nanofluidic channels was further confirmed by X-ray diffraction (XRD) characterization (Fig. 2j), which exhibits a series of intense and sharp (00l) diffraction peaks with a d-spacing of 9.6 Å, 6.4 Å, 4.8 Å, and 3.2 Å, respectively. This result indicates a basal spacing of 1.6 Å with preferred orientation and ordered stacking along the [010] direction during the vacuum-assisted filtration process, and this is well consistent with that of the MnO₂/MXene hybrid membrane.† The preferred orientation of these restacked nanosheets results in ideal 2D nanofluidic channels with atomically smooth surfaces as highly expected.

The thermal stability of the MnO₂/MWCNT membrane was further evaluated by thermogravimetric analysis (TGA), which showed that our MnO₂/MWCNT membrane was thermally stable up to 225 °C (Fig. S7, ESI†). We further checked the wettability of our MnO₂/MWCNT membrane in the as-employed 2 M ZnSO₄/0.2 M MnSO₄ aqueous electrolyte for ZIB applications. This MnO₂/MWCNT membrane is hydrophilic with an initial contact angle of 18.7°, and this contact angle gradually decreased to 5.2° in a short time of ~15 seconds (Fig. S8, ESI†). The rapid decrease in contact angle may provide evidence of the facile permeation and transfer of the aqueous electrolyte in our MnO₂/MWCNT membrane.†

Based on the successful formation of 2D nanofluidic channels, the electrochemical properties of the MnO₂/MWCNT membrane cathode were subsequently measured in a typical coin-type cell. Fig. 3a shows the CV profile of the MnO₂/MWCNT membrane cathode at 0.05 mV s⁻¹, and two redox peaks at 1.42 V and 1.30 V were detected. Fig. 3b shows the cyclic voltamograms (CVs) of this membrane sample in an aqueous 2 M ZnSO₄/0.2 M MnSO₄ electrolyte within the coin-type cell. Our CV curves are mostly consistent with those of the δ-MnO₂ bulk electrode in the previous study.† Two cathodic peaks located at 1.35 eV and 1.17 eV can be observed, which are attributed to Zn²⁺ insertion into the restacked δ-MnO₂ nanosheets and the consequent reduction of Mn⁴⁺ to lower oxidation states, respectively. The appearance of an overlapped anodic peak at around 1.70 V may correspond to Zn-extraction which is accompanied by the recovery to the Mn⁰ state. We further estimate the pseudocapacitance-like contribution by clarifying the kinetics to separate the capacitive-controlled and diffusion-controlled capacities. In general, the capacitive contribution can be quantified using the modified power law:†

\[
I = I_{\text{capacitive}} + I_{\text{diffusion}} = k_1 v + k_2 v^{1/2}
\]

where \(I\) is the current; \(v\) is the scan rate; \(k_1\) and \(k_2\) correspond to the current contributions from the pseudocapacitive effect and diffusion-controlled insertion, respectively (Fig. S10, ESI†). The typical CV profile in Fig. 3b separates the capacitive current (red region) from the diffusion current (green region). The contribution of capacitive current is 14.4%, 18.2%, 23.9%, and 33.4% at scan rates of 0.2, 0.4, 0.6, 0.8, and 1.0 mV s⁻¹ (Fig. 3c), respectively. Fig. 3d depicts the characteristic galvanostatic charge/discharge profile of the MnO₂/MWCNT membrane electrode in the initial three cycles. Two discharge plateaus at 1.42 V and 1.30 V should be ascribed to H⁺ and Zn²⁺ insertion and extraction, respectively. In general, these two discharge plateaus were easily observed in manganese oxide-based ZIBs due to the coexistence of the reversible Zn ion insertion/extraction and proton reaction during the electrochemical process as follows:†

\[
\delta-MnO_2 + Zn^{2+} + e^- \rightarrow Mn^{3+} + Zn^{2+}
\]

\[
Mn^{3+} + 2H_2O + e^- \rightarrow Mn^{2+} + 2OH^- + H_2
\]
\[
\text{MnO}_2 + x\text{Zn}^{2+} + 2xe^- \leftrightarrow \text{Zn}_x\text{MnO}_2 \cdot x\text{H}_2\text{O}, \quad (0 < x < 1)
\]
\[
\text{(Zn ion insertion/extraction)}
\]
\[
\text{MnO}_2 + H^+ + e^- \leftrightarrow \text{MnOOH} \quad \text{(proton reaction)}
\]
\[
4\text{Zn}^{2+} + 6\text{OH}^- + \text{SO}_4^{2-} + 5\text{H}_2\text{O} \leftrightarrow 3\text{Zn(OH)}_2 \cdot \text{ZnSO}_4 \cdot 5\text{H}_2\text{O}
\]

To further understand the phase evolution more clearly, ex situ XRD characterization of this freestanding membrane was performed as shown in Fig. 3e and f. Notably, the XRD pattern at the initial discharging state (A) was drastically different from that of the fresh MnO2/MWCNT membrane just after vacuum filtration (Fig. 2n), which was also reported by T. Vaughey.59 During the following discharging process (A → E), the diffraction peaks significantly decreased in intensity, and some sharp peaks (including sharp peaks at 7.9°, 16.2°, and 24.2° and other weak peaks highlighted by grey cycles) emerged when it was discharged to 1.3 V, which can be indexed to 3Zn(OH)\(_2\) \cdot ZnSO\(_4\) \cdot 5H\(_2\)O (zinc hydroxide sulfate hydrate, JCPDS 78-0246). The formation of zinc hydroxide sulfate hydrate should be attributed to the increasing amount of OH\(^-\) groups in the electrolyte, further confirming the reversible proton reaction.41 During the subsequent charging process (E → J), the diffraction peaks from the intermediate phase of zinc hydroxide sulfate hydrate gradually disappeared, and the pattern finally well recovered to the initial state (A) when the charge voltage was 1.6 V (H). In general, the ex situ XRD characterization clearly confirms the excellent reversibility of the Zn\(^{2+}\) intercalation/deintercalation process in our MnO2/MWCNT membrane. The rate capability of the MnO2/MWCNT electrode in the range of 1–1.9 V is evaluated at various current densities (0.1–2 A g\(^{-1}\)) (Fig. 3g). The MnO2/MWCNT electrode shows high reversible capacities of 273.8, 265.1, 206.8, 135.5, 99, 71.4 and 270.2 mA h g\(^{-1}\) at current densities of 0.1, 0.2, 0.4, 0.6, 0.8, 1 and 2 A g\(^{-1}\), respectively (Fig. S11, ESI†). When the current density is decreased back to 0.1 A g\(^{-1}\), the capacity returns to 270.2 mA h g\(^{-1}\), demonstrating its outstanding capacity retention at both low and high rates.

It has been revealed that the electrolyte ions show drastically different behavior when confined in 2D nanochannels.22–24 Subsequently, the Zn\(^{2+}\) diffusion ability in our \(\delta\)-MnO2/MWCNT membrane was identified using electrochemical impedance spectroscopy (EIS). Combining the Nyquist spectra measured at the first (Point 1) and second (Point 2) discharge depth (Fig. 3h) and \(Z'\) vs. \(\omega^{-1/2}\) plots (Fig. 3i), the diffusion coefficient can be calculated based on the following equation:42,43

\[
\lambda_D = \frac{3.04 \times 10^{-10}}{z_i \sqrt{\epsilon_i}}
\]

where \(z_i\) and \(\epsilon_i\) are the valency and concentration of Zn\(^{2+}\), respectively. In our work, a 2 M ZnSO\(_4\)/0.2 M MnSO\(_4\) aqueous solution was employed as the electrolyte. Thus, the Debye length is estimated to be 0.11 nm. Note that the \(d\)-spacing (0.96 nm) of restacked \(\delta\)-MnO2 nanofluidic channels determined from the XRD pattern (Fig. 2n) is larger than the value of the Debye length, showing the same tendency as the Co\(_3\)O\(_4\) nanochannels in Li ion storage reported by Yu \textit{et al.}22 Accordingly, inside our 2D nanochannel, when the characteristic dimensions are comparable or less than the Zn\(^{2+}\) Debye length (0.11 nm), the negative MnO\(_2\) surface charges on the inner walls would repel counterions of the same charge (negative SO\(_4^{2-}\) groups), making Zn\(^{2+}\) the dominating charge carriers.55,43 This unipolar ionic transport is generally beneficial for high Zn\(^{2+}\) conductivity and a high diffusion coefficient.22,23

More attention has been paid to the long-term cycle performance of this MnO\(_2\)/MWCNT membrane. Our \(\delta\)-MnO\(_2\) membrane-like cathode delivers a capacity of 273.8 mA h g\(^{-1}\) for 600 cycles with a coulombic efficiency around 100%. In contrast, the \(\delta\)-MnO\(_2\) bulk counterpart displays rather poor cycling stability, the capacity of which drastically faded after the initial 10 cycles (Fig. 4a and S14, ESI†). In general, the formation of needle-like Zn dendrites on the Zn anode should be the main reason for the rapid degradation of capacity and poor stability.7 Thus, the reason for the difference in stability of our MnO\(_2\)/MWCNT membrane was further studied by SEM observation of the Zn foil anode after long-term cycling. Interestingly, one can clearly see that the Zn foil taken from the \(\delta\)-MnO\(_2\) bulk ZIBs shows a rough surface with dense Zn dendrites (Fig. 4b). However, the \(\delta\)-MnO\(_2\)/MWCNT membrane based ZIBs show a typical dendrite-free surface (Fig. 4e) with the absence of any ZnO by-products on the Zn foil after 600 charge/discharge cycles.
Although the origin of the dendrite-free characteristics in the membrane-based battery is not yet clear, we speculate that the smooth surface in our advanced 2D channel structure may be responsible. Recently, Liaw et al. found that the inhomogeneity of the distribution of the current density increases with the increase of the cathode particle size and surface roughness. In our case, taking advantage of the characteristic 2D planar morphology with a high aspect ratio, the nanosheets are generally tiled on the filter paper during the vacuum filtration process. As a result, there are almost no MWCNTs on the outermost surface of the MnO$_2$/MWCNT membrane (Fig. S5†). Such a smooth surface of the MnO$_2$/MWCNT membrane cathode could significantly reduce the concentration gradient in the electrolyte and induce homogeneous Zn$^{2+}$ stripping/plating on the zinc anode (Fig. 4d). In contrast, for the bulk MnO$_2$ cathode, the sharp corners and irregular morphology of the bulk MnO$_2$ gives rise to the tip-effect with a high concentration gradient during the Zn$^{2+}$ stripping/plating process (Fig. 4e). Finally, dense Zn dendrite growth occurred at the surface of the Zn foil, drastically decreasing the cycling stability of the battery system.\(^7\)–\(^9\)

We finally assembled a sandwiched, paper-like ZIB using the δ-MnO$_2$ membrane, a 2 mol L$^{-1}$ ZnSO$_4$ + 0.2 mol L$^{-1}$ MnSO$_4$ solution, PET non-woven fabric (thickness ~ 80 μm) and flexible Zn foil as the cathode, electrolyte, separator and anode, respectively (Fig. 5a). This paper-like battery shows a thickness of just ~132 μm (Fig. 5b) which is comparable with that of standard A4 printing paper (Fig. 5c). Two paper batteries connected in series were able to light up a red light-emitting diode (LED) indicator (Fig. 5d). The flexibility of this battery was tested by bending this device at various angles from 45° to 180°, while well maintaining the same energy storage characteristics as those in the initial state (Fig. 5d–i, ESI Video 1†). Compared to a single paper-like ZIB with a charge/discharge potential of 1.4 V, two ZIBs connected in series can achieve 2.8 V in a similar charge/discharge time (Fig. 5j). The flexible ZIB delivers a capacity of 181.6 mA h g$^{-1}$ for 630 cycles at 0.5 A g$^{-1}$ (Fig. S16†). The energy density and power density are 365 W h kg$^{-1}$ and 105 W kg$^{-1}$ (estimated based on the weight of the active MnO$_2$ only), respectively. Such a high energy density surpasses not only that of recently reported MnO$_2$-based ZIBs, but also that of other V-based and HCF-based counterparts (Fig. 5k).\(^{37,41,46–53}\)
4. Conclusions

In summary, we have successfully constructed 2D MnO₂ nanofluidic channels by the restacking of abundant single-layer δ-MnO₂ nanosheets. Compared to the δ-MnO₂ bulk, our 2D MnO₂ nanofluidic channels show fast Zn²⁺ transport and superior cycling stability. The high-performance of Zn²⁺ storage in our rationally designed nanofluidic channels should be associated with the following merits: (1) 2D nanochannels provide abundant interconnected ion transport pathways in vertical and horizontal directions to make the Zn²⁺ flow uniform, preventing the tip-effect and dendrite formation during Zn²⁺ deposition. (2) The implantable MWCNTs play a crucial role in the fast charge diffusion kinetics due to their highly disordered porous structure, and the channels in the vertical direction were cross-linked with the implantable MWCNTs for enhanced Zn ion migration. (3) The void space inside the nanofluidic channels may effectively alleviate structural stress and buffer the volume change during the Zn²⁺ insertion/extraction, avoiding the rapid decay of specific capacity of the ZIBs during cycling caused by irreversible structural collapse of the cathode. (4) Our free-standing δ-MnO₂/MWCNT membrane does not contain any polymer or insulating binder which will avoid sacrificing the active substance content in the ZIBs, thereby achieving enhanced specific capacity after normalization when compared to the conventional δ-MnO₂ bulk electrodes. Our design sheds light on developing high-performance ZIBs from two-dimensional nanofluidic channels, and this strategy might be applicable to storage of other metal ions (Mg²⁺, Ca²⁺, Al³⁺, etc.) in next-generation electrochemical energy storage devices.

Conflicts of interest

There are no conflicts to declare.

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