Ring opening and ring expansion of 8-cyano-tetrazolo[1,5-a]pyridine with secondary amines. Reactions of azides, tetrazoles and nitrenes with nucleophiles, Part 2

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Dedicated to Professor Henk van der Plas on the occasion of his 80th birthday

Abstract
8-Cyanotetrazolo[1,5-a]pyridine 6T undergoes photochemical ring expansion to afford 1,3-diazepine 7 with diisopropylamine, but with stronger nucleophiles such as dimethylamine a rapid, quantitative ring opening reaction affords dienyltetrazoles 8 and 9 in the dark (an Addition of Nucleophile – Ring Opening reaction).

Keywords: Pyridine, tetrazole, nucleophilic ring opening

Introduction

Heterocyclic compounds undergo numerous and sometimes unexpected transformations.2 The formation of substituted 2-pyridylformamidines PyNCHNR2 and 2-pyridylhydrazines PyNHNHNR2 in dark reactions of valence tautomeric tetrazolo[1,5-a]pyridines/2-azidopyridines and amines was described in Part 1.1 Tetrazolo[1,5-a]pyridines/2-azidopyridines 1T/1A are also convenient sources of 1H- and 5H-1,3-diazepines (e.g. 1H-diazepines 4) and diazepinones 5 via ring expansion of photochemically (and sometimes thermally) generated 2-pyridylnitrenes 2 to 1,3-diazacyclohepta-1,2,4,6-tetraenes 3 followed by trapping with nucleophiles, most commonly amines, alcohols, and water. A large number of 1,3-diazepines 4 and diazepinones 5 have been prepared in this way as illustrated in Scheme 1.3-5
Here we report that 8-cyanotetrazolo[1,5-a]pyridine/2-azido-3-cyanopyridine 6T/6A undergoes photochemical ring expansion to afford 1H-1,3-diazepine 7 in the presence of diisopropylamine. Other diazepines and a diazepinone analogous to 5 were obtained with alcohols and water. However, no diazepine product was isolable from the reaction with the more nucleophilic dimethylamine. Instead, a rapid ring opening reaction took place in the dark.

Results and Discussion

8-Cyanotetrazolo[1,5-a]pyridine 6T exists as the tetrazole in the solid state. The azide valence tautomer 2-azido-3-cyanopyridine 6A was not detectable in the IR (KBr) or NMR (DMSO) spectra. Photolysis of a dilute solution of 6T in acetonitrile containing diisopropylamine afforded the 1,3-diazepine 7 as orange crystals in 47 % yield (Scheme 2). The structure was ascertained by NMR spectroscopy, drawing on our previous investigations of diazepines. In the 1H NMR spectrum a doublet at 5.78 ppm is assigned to H-5, and a broader doublet at 6.20 ppm to H-1. Assignment of two triplets at 5.06 and 5.70 ppm to H-6 and H-7, respectively, was confirmed by homonuclear decoupling experiments. In the IR spectrum, the NH and CN stretching absorptions appeared at 3375 and 2221 cm⁻¹, respectively.
The same type of diazepine synthesis could not be achieved with dimethylamine and other
strong nucleophiles (diethylamine and pyrrolidine) because 6T reacts rapidly with these amines
in the dark at room temperature by nucleophilic attack on C-5 of the pyridine ring to form the
ring-opened butadienyltetrazoles 8, 9 as shown in Scheme 3.

Scheme 3

The acetonitrile solution of 6T turned yellow immediately upon admission of dimethylamine
gas. After evaporation and recrystallization, the dimethylamine adduct 9 of the dienyltetrazole 8
was obtained as yellow needles. The $^1$H NMR spectrum of 9 showed two singlets at 2.60 and
2.95 ppm, each integrating for six protons and assigned to the methyl groups. The latter signal
was quite broad, probably due to hindered rotation of a dimethylamino moiety. A triplet at 6.50
ppm and two doublets at 6.97 and 7.12 ppm were assigned to the vinylic H-8, H-7 and H-9,
respectively. No other geometric isomers were observed. The infrared spectrum showed a
medium nitrile stretching vibration at 2201 cm$^{-1}$. The extraneous H-bonded dimethylamine was
removed from 9 by dissolution in water and addition of HCl, thus affording the free base 8. In
the $^1$H NMR spectrum of 8 there were two broad methyl group signals at 2.97 and 3.18 ppm
due to the slow rotation of the dimethylamino group. The H-8 proton signal (5.46 ppm) was at ca.
1 ppm higher field than the corresponding signal in 9, while those for H-7 and H-9 were shifted
downfield by 0.5 – 0.7 ppm. This can be explained by an allylic resonance whereby the NMe$_2$
lone pair is delocalized onto C8, whereas C9 becomes an iminium ion carbon. This gives the N-
C9 bond some double bond character, as is clearly indicated in the X-ray structure described
below. The $^{13}$C NMR spectrum of 8 showed two methyl group signals at 37.2 and 44.8 ppm, the
vinylic carbons C6 – C9 at 74.7, 152.2, 95.5, and 157.5 ppm, respectively, the tetrazole carbon
C5 at 153.9, and the nitrile carbon at 117.7 ppm. The CN stretching absorption appeared at 2206
cm$^{-1}$ in the infrared spectrum.

The crystal structure of compound 8, crystallized as its hemi-methanol solvate (8·½MeOH),
was determined. The structure comprises four independent tetrazoles in the asymmetric unit in
addition to a methanol molecule refined at half occupancy and disordered over four different
sites. A view of molecule 'A' is shown in Figure 1. The four tetrazoles (molecules A, B, C and D)
have the same Z-E configuration, where the NH group in the ring is anti to the cyano group. In
each molecule, the terminal -NMe$_2$ group is conjugated with the diene backbone as evident from
the C9n-N10n (n = A, B, C and D) bonds (1.312(4) - 1.319(4) Å) which approach those of
typical C=N double bonds. The four independent tetrazole molecules are essentially planar, and
all atoms lie within 0.2 Å of the least squares plane calculated from all non-H atoms. Each NH group (N1n) donates an H-bond to N3n on an adjacent molecule.

![Figure 1. ORTEP view of compound 8 (molecule A, 30% probability ellipsoids).](image)

Both 8 and 9 eliminate dimethylamine when passed through a heated gas chromatography column. The GC trace and mass spectra demonstrated that tetrazolopyridine 6T was regenerated. This also happened when an acetonitrile solution of 9 was heated to 50°C. We assume that E-Z isomerization in 8 can take place easily on heating. Cyclization can then take place via attack of the formally negatively charged tetrazole N4 on the iminium-type carbon atom C9 to regenerate the adduct 10 (a 'covalent amination' product), which reverts to 6T by elimination of dimethylamine, as illustrated in Scheme 4 below.

![Scheme 4](image)

Other examples of this type of nucleophilic ring opening reaction have been reported; the 2-(p-chlorophenyl)tetrazolo[1,5-a]pyridinium tetrafluoroborate 11 affords the ring-opened
dienyltetrazoles 12 with dimethylamine.\(^6\) Tetrazole 6T reacts with sodium hydroxide to afford a mixture of Z,E- and E,E-dienyltetrazoles, 13 (Scheme 5).\(^7\)

\[
\text{Ar} = p\text{-chlorophenyl}
\]

Scheme 5

The ring opening of 6 to 8 and 9 described in Scheme 3 also takes place with other secondary amines, including diisopropylamine, diethylamine and pyrrolidine. Only the reaction with diisopropylamine is slow enough to allow the diazepine 7 to be prepared successfully by rapid room temperature photolysis. It should be possible to prepare diazepines analogous to 7 (Scheme 2) from the other, more reactive amines by using low temperature photolysis.

Conclusions

Photolysis of 8-cyanotetrazolo[1,5-a]pyridine/2-azidopyridine-3-carbonitrile 6T/6A in the presence of diisopropylamine permits the isolation of the 1,3-diazepine derivative 7. However, less hindered/more nucleophilic amines attack the tetrazole 6T in position 5 in a fast, dark reaction, which results in ring opening to cyanodienyltetrazole 8 and its dimethylamine adduct 9. The reaction is reversible: heating of compound 8 causes elimination of dimethylamine and regeneration of 6T.

Experimental Section

8-Cyanotetrazolo[1,5-a]pyridine (6T). This compound was prepared according to the literature;\(^7\) mp 185-188 °C (dec) (lit.\(^7\) 183 °C).
4-Cyano-2-diisopropylamino-1H-1,3-diazepine (7). A solution of \(6T\) (0.10 g, 0.69 mmol) in 90 mL of acetonitrile was placed in a quartz vessel and purged with \(N_2\) for 45 min. Diisopropylamine (1 mL) was added, the \(N_2\) purge was discontinued, and irradiation commenced immediately at room temperature using a 1000 W high pressure Hg-Xe lamp. After 2.5 h, the orange solution was evaporated to give a brown solid, which was chromatographed on basic alumina with 25 % hexane/ether. The orange band was collected and evaporated to give a yellow–orange solid, which was recrystallized by dissolving in ether and then adding hexane dropwise until precipitation commenced. Yield 0.07g (47 %). Mp 120-121 °C. MS: 218 (M\(^+\)), 203 (-Me), 175 (-propyl radical), 161, 132. IR (KBr) cm\(^{-1}\): 3375 m, 2997 w, 2970 m, 2932 m, 2221 m, 1642 m, 1593 s, 1450 w, 1419 w, 1366 m, 1356 m, 1332 m, 1305 w, 1230 w, 1208 m, 1137 m, 1119 m, 1042 w, 933 w, 895 w, 840 w, 764 m, 689 m. \(^1\)H NMR (DMSO-d\(_6\), 400 MHz): \(\delta\) 1.16 (d, \(3J\) 6.72 Hz, 12H, 4 x Me); 3.76 (quintet, \(3J\) 6.68 Hz, 2H, iPr); 5.06 (dd, \(J\) 6.52 Hz, 1H, H6); 5.70 (dd, J 6.66 Hz, 1H, H7); 5.78 (d, \(3J\) 5.56 Hz, 1H, H5); 6.20 (d, \(3J\) 5.32 Hz, 1H, H1; septet of \(^3\)Pr not fully resolved). \(^13\)C NMR (DMSO-d\(_6\), 100 MHz): \(\delta\) 20.9 Me; 46.9 iPr; 113.7 C6; 120.1 C4 or CN; 126.0 C5; 127.3 CN or C4; 136.7 C7; 157.1 C2. Anal. Calc. for C\(_{12}\)H\(_{18}\)N\(_4\): C, 66.02; H, 8.31; N, 25.67. Found: C, 66.02; H, 8.48; N, 25.63 %.

Z,E-5-Dimethylamino-1-(1H-tetrazolyl)-1,3-butadiene carbonitrile dimethylamine adduct (9). Compound \(6T\) (0.20 g, 1.38 mmol) was dissolved in 7 mL of acetonitrile, and dimethylamine gas was bubbled into the solution at room temperature (ca. 25 °C) for a few seconds. The color changed immediately to yellow, and then rapidly to light orange. TLC (Al\(_2\)O\(_3\), CHCl\(_3\)) revealed that \(6T\) was no longer present. Evaporation of the solution afforded 9 (0.31 g; 96 %). Crystallization in a refrigerator afforded yellow crystals, mp 148-150 °C. Heating an acetonitrile solution of 9 to 50 °C caused partial decomposition with evolution of dimethylamine and regeneration of tetrazolopyridine \(6T\). Data for 9: IR (KBr) cm\(^{-1}\): 3432 m, 3014 m, 2986 m, 2950 m, 2907 m, 2683 m, 2454 m, 2263 w, 2201 s, 2013 w, 1887 w, 1621 s, 1567 s, 1485 s, 1436 m, 1415 s, 1400 s, 1335 m, 1304 m, 1286s, 1263 m, 1224 s, 1203 s, 1191 s, 1134 s, 1110 s, 1066 m, 1038 w, 1024 m, 1016 m, 1009 s, 946 w, 916 m, 888 w, 865 m, 824 m, 768 w, 742 w, 727 w, 691 m, 607 w, 577 w, 504 w, 476 w, 439 w. \(^1\)H NMR (DMSO-d\(_6\), 400 MHz): \(\delta\) 2.60 (s, 6H, (C\(_3\)H\(_3\))\(_2\)NH); 2.95 (br s, 6H, (C\(_3\)H\(_3\))\(_2\)N-R); 6.50 (t, \(3J\) 12.28 Hz, 1H, H8); 6.97 (d, \(3J\) 11.92 Hz, 1H, H7 or H9); 7.12 (d, \(3J\) 12.60 Hz, 1H, H9 or H7) (NH not observed). \(^13\)C NMR (DMSO-d\(_6\), 100 MHz): \(\delta\) 34.9 (CH3)\(_2\)NH; 84.5 C6; 97.3 C8; 123.7 CN; 147.0 C7; 154.1 C9; 158.3 C5. Anal. Calcd for C\(_{10}\)H\(_7\)N\(_7\) (9): C, 51.05; H, 7.28; N, 41.67. Found: C, 50.86; H, 7.33; N, 41.94 %.

Removal of dimethylamine from 9 to give 8.

Compound 9 (0.05 g, 0.21 mmol) was dissolved in water (0.5 mL), and 2 drops of 32 % HCl were added. A yellow precipitate of 8 formed. Water (2 mL) was added, and the mixture was transferred to a Buchner funnel, where it was washed with water then dried under suction. Recrystallization was achieved with acetonitrile or methanol. Yield 0.035 g (87 %). Mp 128-
129 °C (dec). IR (KBr) cm⁻¹: 3549 w, 3423 w, 3125 w, 3027 w, 2917 w, 2638 w, 2498 w, 2206 m, 1636 s, 1581 s, 1477 m, 1420 s, 1410 s, 1340 w, 1311 w, 1280 m, 1240 m, 1202 w, 1119 m, 1090 m, 1014 w, 980 m, 917 w, 866 w, 812 w, 756 w, 740 w, 701 w, 597 w, 432 w. ¹H NMR (DMSO-d₆, 400 MHz): δ 2.97 (s, 3H, Me); 3.18 (s, 3H, Me); 5.46 (t, ³J 12.16 Hz, 1H, H8); 7.56 (d, ³J 11.96 Hz, 1H, H7 or H9); 7.81 (d, ³J 12.32 Hz, 1H, H9 or H7) (NH not observed); ¹³C NMR (DMSO-d₆, 100 MHz): δ 37.2 Me; 44.8 Me; 74.7 C6; 95.5 C8; 117.7 CN; 152.2 C7; 153.9 C5; 157.5 C9. Anal. Calc. for C₈H₁₀N₆: C, 50.52; H, 5.30; N, 44.18. Found: C, 50.58; H, 5.39; N, 43.98 %.

Crystallography
Cell constants were determined by least-squares fits to the setting parameters of 25 independent reflections measured on an Oxford Diffraction Gemini CCD diffractometer employing graphite-monochromated Mo Kα radiation (0.71073 Å). Data reduction and absorption correction was performed with the Crysalis package. The structure was solved by direct methods with SHELXS and refined by full-matrix least-squares analysis with SHELXL-97. All non-H atoms were refined with anisotropic thermal parameters. H-atoms were included in estimated positions using a riding model. Drawings were produced with ORTEP.

Crystal data. Compound 8·½MeOH: C₈.₅H₁₂N₆O₀.₅, M = 206.24, triclinic, space group P ₁̅₁, a = 8.0828(4), b = 15.6043(8), c = 17.3420(9) Å, α = 87.289(4), β = 77.839(4), γ = 87.795(4)°, U = 2134.9(2) Å³, Z = 8, Dc = 1.283 g cm⁻³, µ = 0.090 mm⁻¹, 16168 reflections measured, 7365 unique (Rint = 0.0475), R₁ = 0.0581 (for 2794 observed data, I > 2σ(I), wR₂ = 0.1118 (all data).

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Supplementary information

Crystallographic data for compound 8·½MeOH have been deposited with the Cambridge crystallographic Data Centre (CCDC 698469). This information can be obtained free of charge from The Director, CCDC, 12 Union Road, Cambridge, CB2 1EZ, UK (fax: +44-1223-336033; e-mail: deposit@ccdc.cam.ac.uk or http://www.ccdc.cam.ac.uk).
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