Review Article
Cardiolipin, Perhydroxyl Radicals, and Lipid Peroxidation in Mitochondrial Dysfunctions and Aging

Alexander V. Panov¹ and Sergey I. Dikalov²

¹Federal Scientific Center for Family Health and Human Reproduction Problems, 16 Timiryasev str., Irkutsk, 664003, Russian Federation, Russia
²Division of Clinical Pharmacology, Vanderbilt University Medical Center, Nashville, TN 37232, USA

Correspondence should be addressed to Alexander V. Panov; alexander.panov55@gmail.com

Received 16 April 2019; Accepted 19 February 2020; Published 9 September 2020

Guest Editor: Armen Y. Mulkidjanian

Copyright © 2020 Alexander V. Panov and Sergey I. Dikalov. This is an open access article distributed under the Creative Commons Attribution License, which permits unrestricted use, distribution, and reproduction in any medium, provided the original work is properly cited.

Mitochondrial dysfunctions caused by oxidative stress are currently regarded as the main cause of aging. Accumulation of mutations and deletions of mtDNA is a hallmark of aging. So far, however, there is no evidence that most studied oxygen radicals are directly responsible for mutations of mtDNA. Oxidative damages to cardiolipin (CL) and phosphatidylethanolamine (PEA) are also hallmarks of oxidative stress, but the mechanisms of their damage remain obscure. CL is the only phospholipid present almost exclusively in the inner mitochondrial membrane (IMM) where it is responsible, together with PEA, for the maintenance of the superstructures of oxidative phosphorylation enzymes. CL has negative charges at the headgroups and due to specific localization at the negative curves of the IMM, it creates areas with the strong negative charge where local pH may be several units lower than in the surrounding bulk phases. At these sites with the higher acidity, the chance of protonation of the superoxide radical (O₂•⁻), generated by the respiratory chain, is much higher with the formation of the highly reactive hydrophobic perhydroxyl radical (HO₂•). HO₂• specifically reacts with the double bonds of polyunsaturated fatty acids (PUFA) initiating the isoprostane pathway of lipid peroxidation. Because HO₂• is formed close to CL aggregates and PEA, it causes peroxidation of the linoleic acid in CL and also damages PEA. This causes disruption of the structural and functional integrity of the respirosomes and ATP synthase. We provide evidence that in elderly individuals with metabolic syndrome (MetS), fatty acids become the major substrates for production of ATP and this may increase several-fold generation of O₂•⁻ and thus HO₂•. We conclude that MetS accelerates aging and the mitochondrial dysfunctions are caused by the HO₂•-induced direct oxidation of CL and the isoprostane pathway of lipid peroxidation (IPLP). The toxic products of IPLP damage not only PEA, but also mtDNA and OXPHOS proteins. This results in gradual disruption of the structural and functional integrity of mitochondria and cells.

1. Introduction

For the last five decades, the free radical theory of aging, first proposed by Harman [1–3], was regarded as one of the most important and investigated among other hypotheses of aging. Accumulation of mtDNA point mutations [4] and an exponential increase with age of mtDNA deletions [5] were considered as important hallmarks of ageing and the age-related diseases. Accumulation of mtDNA mutations with increasing age was demonstrated mainly in heart, skeletal muscles, and brain, which undergo a high range of workloads and associated metabolic activities. In the liver, which has relatively stable high rate of metabolism and high regeneration capacity, there was shown no significant accumulation of mtDNA mutations with age [6]. For decades, it was generally agreed that ROS-associated somatic mutations of mtDNA contribute to human aging and to the decline of energetic capabilities at higher age [5, 7].

In recent years, however, new experimental data have discredited to some extent the original theory of direct participation of free radicals in mtDNA mutations [8–10]. It has also been found that most common ROS species either are
not enough chemically active (O₂•*NO) to initiate mutations or excessively active (•OH) and have a too short life span in order to reach mtDNA [11, 12]. Several decades ago, researchers from the Vanderbilt University have discovered the isoprostane pathway of lipid peroxidation (IPLP), which results in the formation of prostaglandin-like compounds with enormous variations in molecular positional isomerism and stereoisomerism [13–16]. Considering the chemistry of IPLP, we have proposed the mechanism of initiation of IPLP involving formation in the inner mitochondrial membrane of perhydroxyl radical (HO₂•), which is the protonated form of superoxide radical (O₂•*) [12, 17]. We suggest that the mechanisms of action of perhydroxyl radical (HO₂•) can explain the mechanism of initiation of the autoxidation of polyunsaturated fatty acids that are still part of the membrane’s phospholipid. In addition, HO₂• can be regarded as a “carrier” of the highly toxic hydroxyl radical (•OH) inside the hydrophobic core of the membrane and thus can oxidize cardiolipin. This can happen in spite the fact that normally CL contains four linoleic acids with only two double bonds, which are not substrates for IPLP. In this review, we will focus on the roles of cardiolipin in oxidative stress due to the ability of this phospholipid to create, close to the surfaces of the inner mitochondrial membrane, the narrow layer of structured water with low pH and high conductance for protons [18–20]. This facilitates conversion of O₂•*, produced by the respiratory chain, into HO₂• that initiate IPLP [12, 17]. We consider that formation of toxic products, such as isolevuglandins (aka isoketals), produced by IPLP, as the major event in oxidative stress that is responsible for the slow but inevitable process of aging. The gradual accumulation of oxidative damages to mitochondrial enzymes and the membrane phospholipids, first of all phosphatidylethanolamine and cardiolipin, causes disruption of the structural and functional integrity of the respirosomes and ATP synthase.

2. Properties of Cardiolipin

Cardiolipin (CL) is the only phospholipid that is present almost exclusively in the inner mitochondrial membrane of all animals and in the plasma membrane of aerobic bacteria. That is in the membranes, where ATP is formed by the F₀F₁ ATP synthase [21]. CL is absent in the thylakoids of plants, where it is substituted by the plant sulfolipid (Sulfoquinovosyl dipalmitoylglycerol), probably because chloroplasts have another isoform of ATP-synthase (CF₀CF₁) [22], and, most importantly, the reversed, as compared to mitochondria, orientations of the energy-transforming enzymes, pH gradients, and electrical charges. The tight connection between the F₀F₁-ATP-synthase and cardiolipin was proved by experiments, in which facultative anaerobes were placed into the oxygen-depleted atmosphere, and after a while, they lost not only F₀F₁-ATP-synthase, but also cardiolipin [21].

The structure of the cardiolipin molecule was thoroughly described and discussed in [21, 23, 24]. The significance of CL specifically for the energy-transforming membranes comes from several intrinsic properties of the phospholipid, which originate from its unique structure (Figure 1). Due to the conical form, CL, together with also conical phosphatidylethanolamine (PEA), supports the superstructural organization of the respiratory chain and ATP synthase and allows accommodation of the multienzyme complexes into the

![Cardiolipin structure](image-url)
sharp curves of the inner membrane. Due to the conical form and strong negative charges at the headgroups, CL has the propensity to form inverted hexagonal (HII) structures, which are essential for activation of the respiratory chain enzymes and ATP synthesis by promoting binding and conductance of protons through the hydrophobic milieu of the inner membrane [25–28] (Figure 1).

Figure 1 does not reflect fully the fact that the shape of CL resembles that of a pyramid or cone with the broad base and a very small head, which has the length of a glycerol molecule with three C atoms. In most cardiolipins, the fatty acids have usually two double bonds and therefore are strongly bent. For this reason, CL, similar to PEA, cannot form "normal" bilayer membranes [25, 27]. The two phosphate groups give the CL molecule the ability to trap protons [29] and thereby to facilitate proton translocation along the membrane surface [30].

In general, it has been found that composition of the four fatty acids in CL has small variations in the number of double bonds for the given organism or even tissue [31]. The selection of fatty acids in CL is specific for every species and organ, but unsaturated fatty acids are most common [32]. In many animals, the “mature” CL, that is after remodeling, has four linoleic acids with two double bonds (C18:2). In the fast respiring mitochondria of animals, CL may comprise up to 20% of all lipids, and up to 80% of CL is located in the inner leaflet of the inner mitochondrial membrane [33]. In the liver mitochondria, which have relatively low respiratory activity, CL is distributed more or less evenly between the two sides of the inner membrane. The unique structure of cardiolipin results in a number of important properties [23, 34] briefly listed as follows:

1. As the two small molecules of phosphatidic acid are located at the small "head" of the phospholipid, which is much smaller than the charged heads of other phospholipids, namely, phosphatidylcholine (PC), phosphatidylserine (PS), or phosphatidylglycerol (PG), the four fatty acids of CL are closer to each other with stronger attachments, and thus the temperature of phase transitions is higher than in the surrounding membrane. In other words, the lipids of CL are solid in comparison with the fluid lipids of surrounding phospholipids, most of which contain polyunsaturated fatty acids (PUFA). This allows protein complexes “float” being imbedded into the rafts, made of cardiolipin, and thus rearrange into functional supercomplexes, such as respirosomes

2. Small charged heads of CL allow the phospholipid to fit proteins into negative curvatures of the inner leaflet (the matrix side) of the inner membrane

3. Under physiological conditions, the presence of one or two negative charges at the phosphate groups allows CL to create the negatively charged regions that electrostatically interact with proteins and peptides. The excessive negative charges on the surface of rafts, comprised of many CL molecules, create the so-called antennae, when the Coulomb radii of several cardiolipins overlap and amplify a local negative charge

4. The mobility of the head group of CL is restricted because of binding of the glycerol to two phosphatidates. Therefore, the self-shielding of phosphate groups in CL is much lower than in other phospholipids in the membrane. As a result, the phosphates of cardiolipin on the surface of the inner membrane are easily accessible for interactions with proteins, peptides, and ions, as compared with other membrane phospholipids [23]. The interactions of CL with some proteins are so strong that CL was present in the crystal structures of the isolated mitochondrial proteins, for example, complex 3 and ANT [23]. Moreover, CL is known to specifically coordinate divalent cations with high affinity [23], and the ion-lipid stoichiometry is dependent on the CL headgroup formal charge [35].

It was suggested that distribution of the headgroup charge determines the types of nonbonding interactions that can occur among lipids, including electrostatic and hydrogen-bonding interactions [36, 37]. Clearly, for each of the above phenomena, the ionization state of CL headgroup is of fundamental importance; whether this lipid exists in the mono- or dianionic state, the charge strongly influences behavior of CL in the bilayer [27].

3. Formation of Cardiolipin Clusters with the Negatively Charged Antennae

The ability of CL to accommodate the sharp curvatures of the mitochondrial cristae explains the fact that most of CL molecules are specifically located in the inner leaflet of the inner mitochondrial membrane (IMM) and non-covalently linked to various proteins. The list of proteins linked to CL is very long and includes, first of all, the electron transport complexes of the respiratory chain, ATP-synthase, and a number of transmembrane carriers, such as ATP/ADP carrier (ANT), inorganic phosphate carrier, and uncoupling protein [26, 38]. A relatively small amount of CL, which is present in the outer leaflet of the inner mitochondrial membrane, was found to be bound to cytochrome c, and at the contact sites of the inner and outer membranes, CL interacts with creatine kinase and ANT [reviewed in 38].

From studies of the proteins crystals, it was deduced that CL construct bridges between proteins and thus promote existence of many proteins as dimers. It is known that all respiratory chain complexes and ATP-synthase are present in the membrane as homodimers, which, in their turn, are assembled into supercomplexes [39]. One of the consequences of the unique structure of CL is that it tends to form rafts by clustering at the sites with negative curvature of the IMM. At these sites, CL segregates laterally from other phospholipids and stabilizes the membrane, simultaneously with holding together the respirosomes [38, 40]. Although CL is the main anionic phospholipid comprising up to 20% of all phospholipids of the IMM [33], since it is located...
predominantly in the inner leaflet of the IMM, particularly at the negative curves of the cristae, at some locations, concentration of CL may be high. The regions with overlapping Coulomb radii of negative charges, called antennae, are rather common all over the matrix side of the IMM [33, 41].

The phosphoester groups of most phospholipids behave as strong acids with pKₐ between 1 and 3, which coincide with the pKₐ₁ of the orthophosphoric acid [27]. Therefore, in the absence of the intramolecular and intermolecular interactions, which could shield the phosphates at the head of CL, it could be expected that both phosphates of CL behave similarly, and that at the neutral pH of a cell CL behaves as a dianion [27]. Indeed, some researchers observed CL as a dianion with pKₐ₁ ≈ 2 for both phosphates [27], whereas others observed large difference between pKₐ₁ ≈ 2 and pKₐ₂ ≈ 8 [21] and CL behavior as a monoanion. According to Haines [21], the reason for the discrepancies was caused by different properties of CL in the artificial membranes [27] and in the biological membranes [21].

Under normal conditions, the ΔpH between the cytosolic and the matrix sides of the inner mitochondrial membrane is close to 1.0, which is equivalent to 60 mV [42, 43]. It was suggested that one of the consequences of different dissociation of the CL phosphates may be as follows: on the cytosolic side (the outer surface of the IMM) CL traps protons, and thus converts ΔpH into ΔΨ, by enhancing the “+” charge on the outer surface of the inner membrane. On the matrix side, CL behaves as a dianion, and thus enhances the “−” charge of the ΔΨ at the inner leaflet of the IMM [28].

CL has an important function of buffering protons, particularly on the matrix side of the inner membrane [44]. In most active mitochondria of the heart, brain, skeletal muscles, and kidney, it is hardly possible to talk about pH of the matrix because mitochondrial proteins, together with water, exist as a quasicrystalline phase [45]. Antonenko et al. [46] suggested the existence of a kinetic barrier for proton transfer from the surface of bilayer phospholipid membrane to bulk water. Thus, there is no “normal” diffusion of protons and substrates in the bulk volume of the matrix. However, close to the surface of the inner membrane bearing negative charge, there are few layers of structured water, which serve as a conductor for protons by the Grotthuss mechanism, and thus the local concentration of protons may be several orders higher than in the bulk volume of the matrix [21].

To explain the mechanisms of oxidative damage of cardiolipin and other mitochondrial phospholipids, we have suggested that low pH at the surfaces of the inner mitochondrial membrane, that is at the border between the lipid phase of the membrane and the water phases of the matrix and intermembrane compartments, is of paramount significance for the conversion of the relatively chemically inactive superoxide radical (O₂⁻) to the highly active perhydroxyl radical (HO₂⁺) [12, 17].

4. Properties of the Perhydroxyl Radical (HO₂⁺)

Superoxide radical, when formed in the mitochondrial membrane, is rapidly removed from the lipid phase into the matrix or the intermembrane space because it is anion [49]. For some time, O₂⁻ was considered as the main radical responsible for oxidative stress and aging, but soon, it was discovered that O₂⁻ very poorly interacts with polyunsaturated fatty acids (PUFA) and amino acids [11, 18]. Therefore, the suspects became oxidants, which may originate from superoxide radicals, such as peroxynitrite (ONOO⁻) and hydroxyl radical (•OH). Peroxynitrite, as well as peroxynitrate (O₂NOO⁻), react relatively slowly with most, but not all, biological molecules, making them rather selective oxidants. Both ONOO⁻ and O₂NOO⁻ modify tyrosine in proteins, leaving a footprint detectable in vivo [50–52]. The most active oxidant, the •OH, indeed is so active that it reacts within 1 to 5 molecular diameters of their site of formation [53]. According to Pryor [53], •OH reacts with free linolate with rate constants that are nearly diffusion-controlled. The lifetime of •OH radicals have been estimated 10⁻⁶ sec. It reacts indiscriminately with organic molecules/groups and, therefore, cannot reach the inner layer of lipid membrane containing the unsaturated double bonds to initiate lipid peroxidation. All of the above mentioned radicals are hydrophilic. Perhydroxyl radical is a protonated form of the superoxide radical and has molecular formula HO₂⁺, and it is hydrophobic and much more active chemically than O₂⁻. Unlike the superoxide (O₂⁻), HO₂⁺ is a powerful oxidant [53]. The above considerations leave us with the only plausible candidate for initiation of the autoxidation of polyunsaturated fatty acids (PUFA) and CL, the perhydroxyl radical (HO₂⁺) [12, 17] (Figure 2).

Perhydroxyl radical is always present in the cell due to reversible reaction O₂⁻ + H⁺ ↔ HO₂⁺ with pKa = 4.88 [54]. Because of the low pKa, it was widely accepted that at pH 7.2 in the cytoplasm, much less than 1% of [O₂⁻] is present as HO₂⁺ [55]. Perhaps for this reason, many researchers presumed that HO₂⁺ has little or no role in initiation of lipid peroxidation [56]. However, several authors pointed out that pH values in the microvolumes in the vicinity of charged membranes may be several units lower than in the bulk volume of a cell [18, 21, 57]. This may occur around the negatively charged heads of cardiolipin, and other phospholipids, such as phosphatidyleserine and phosphatidylinositol which may retain protons. Thus, not the bulk pH, but local pH at the interfaces of the IMM in matrix and cytosol is critical for formation of HO₂⁺, and because it has no charge, HO₂⁺ can easily go back into the lipid core of the membrane and cross it [58]. It should be also kept in mind that on both sides of the IMM, protons have much higher mobility by the Grotthuss mechanism in the few layers of structured water molecules close to the charged surfaces of the membrane, particularly around the rafts with antennae made of cardiolipin [21, 59, 60]. It is important that interaction of O₂⁻ with H⁺ occurs at the interfaces of the matrix and cytoplasmic sides of the IMM where CL clusters hold together respirosomes and thus where superoxide radicals are produced [47]. For this reasons, CL and PEA may be the first phospholipids with unsaturated bonds affected by HO₂⁺.
In comparison with other oxidants, HO$_2^*$ shows high specificity in reaction with PUFA, linoleic (C18:2), and linolenic (C18:3) acids [11]. Thus, when HO$_2^*$ encounters PUFA attached to a phospholipid, or two double bonds of linoleic acids (18:2) of CL, it reacts with them, with high probability and fast. This might explain why CL is one of the most sensitive phospholipids for oxidative damage [46, 61, 62]. Indeed, it has been shown that increased production of mitochondrial HO$_2^*$ in cardiovascular conditions is associated with an increased cardiolipin oxidation, which represent a specific biomarker of mitochondrial oxidative stress [63]. The outcome of HO$_2^*$ reaction with a different number of double bonds will be very distinct: IPLP, which requires PUFA with at least three double bonds present, produces racemic mixtures of hundreds of final products, whereas reaction with the most common fatty acid in CL, linoleic acids with two double bonds, yields “regular” lipid hydroperoxide [64].

5. The Isoprostane Pathway of Lipid Peroxidation

More than 50 years ago, it was observed that during autoxidation of linolenic acid or prolonged storage of the human blood plasma at -20°C, there were formed products similar to prostaglandins H and F$_2$ [65]. Roberts and Morrow at
the Vanderbilt University have shown the nonspecific auto-
oxidative formation of prostaglandins in vivo [15, 16, 66–
68] and discovered the nonenzymatic pathway of PUFA lipid
peroxidation with racemic mixture of products possessing
enormous variations in molecular positional isomerism and
stereoisomerism in structure and biological activities.

A large number of products of this type of PUFA autox-
diation possess high reactivities with lipids and proteins, and
the resulting products may be considered as the most reliable
and sensitive early markers of oxidative damages of lipids
and proteins during aging and aging-associated pathologies
[68–71]. The major features of the IPLP type of lipid peroxi-
dation have been described in a number of publications [16,
66, 72]. The isoprostanes (IsoPs) that contain F-type pro-
stane rings analogous to PGF2 were the first class of oxidation
products discovered during IPLP in abundance in vitro and
in vivo [71].

IsoPs and the cyclooxygenase-derived prostaglandins
(PGs) have a number of distinctions in their origin and prop-
erties, which have been discussed in a number of publications
[12, 66, 71, 73–76]. Here we briefly list the most important
distinctions: (1) The side chains of PGs are almost always ori-
ented trans to the prostanate ring whereas the products of IPLP
have mostly the side chains with cis orientation [66, 71]. (2)
The IsoPs are formed in situ from PUFA, which are esterified
to phospholipids, while PGs are generated exclusively from
the free AA and DHA [74]. (3) The products of IPLP are the
racemic mixture of products with a very large number of
diffuse stereoisomers and positional isomers, whereas
the products of the enzymatically produced prostaglandins
have only one optical isomer each [66, 74].

Some of the products of PUFA autoxidation, such as γ-
ketoaldehydes, are highly reactive molecules, which form
adducts with primary amines of the lysine-containing pro-
teins and phosphatidylethanolamine [69]. The most active
among γ-ketoaldehydes formed from AA via the IPLP are
isolevuglandins (IsoLGs). IsoLGs are so reactive, that were
revealed only as adducts with proteins or ethanolamine of
PEA. A number of IsoPs possess potent biological activity and
thus can function as mediators of the oxidant injury, or
convey abnormal cellular signaling [69].

In addition to arachidonic acid (AA), which is the most
common among phospholipids, other PUFA such as eicsa-
pentaenoic acid (EPA) and docosahexaenoic acid (DHA)
have been found as substrates for the IPLP [70]. Because
DHA is present in a larger quantity in neurons, the products
of IPLP were correspondingly named neuroprostanes and
neuroketals (Figure 3).

The differences between IPLP and the “classical” lipid
peroxidation have been discussed in [17]. However, the
mechanisms, which are responsible for such enormous diver-
sity of stereoisomers and positional isomers of the final pro-
duct of PUFA during the IPLP, remained unknown, as well as
the radical responsible for initiation of IPLP.

Figure 2(a) illustrates one of the common presentations
of PUFA autoxidation. Initially, the radical that initiates
autoxidation of AA was denoted as *OH [76], Morrow and
Roberts usually showed no radical at all [66, 74, 75]. The
problem with hydroxyl radical is that it cannot by itself
appear in the middle of the hydrophobic core of the mem-
brane. HO2⁻ is the only radical that can spontaneously come
in contact with a PUFA, which is part of a phospholipid.

Recently, a hypothesis on the mechanism of initiation of
IPLP by perhydroxyl radical (HO2⁻) was presented [12, 17]. If
we accept that it is the perhydroxyl radical that abstracts
the first H atom from AA, then the product of the reaction will be
hydrogen peroxide (H2O2). It has been shown [77, 78] that in
the hydrophobic milieu, H2O2 undergoes homolytic splitting
with formation of two hydroxyl radicals: H2O2 → 2*OH,
which instantly will abstract another two H atoms, more
likely from the same PUFA. This will delocalize all electrons
in the double bonds with unpredictable structure of interme-
diate forms. Therefore, the intermediate structures of AA
shown in Figure 2(a) above “endoperoxides” have no sense
because nobody knows what structures will be acquired by
intermediate metabolites. Therefore, in Figure 2(b), we show
only the sequence of chemical intermediates for the perhy-
droxyl radical and show only the starting molecule AA and
the possible structures of the final products. The proposed
mechanism of IPLP initiation was first presented in [12, 17].

Although the abstractions of the three H atoms are
shown schematically as separate and consecutive events, see
Figures 2(a) and 2(b) and [77], in reality, abstraction of all
hydrogen atoms occurs extremely fast, as one chain reaction,
between two O2 molecules join the remaining skeleton of AA.
Rapid abstractions of three H atoms occur randomly at any
double bond, this makes the molecule of PUFA highly unsta-
able. For this reason, the O2 molecules bind randomly with

\[
\text{Arachidonic acid} \rightarrow \text{IPLP} \rightarrow \text{Isoketals} \text{ Neuroprostanes} \text{ Neuroketals}
\]

\[
\text{Docosahexaenoic acid} \rightarrow \text{IPLP} \rightarrow \text{Neuroprostanes} \text{ Neuroketals}
\]

\[
\text{HO} \rightarrow \text{HO}_2 \rightarrow \text{HO}_2^- \rightarrow \text{HO}_2^- \rightarrow \text{hydroxyl radical}
\]
formation of variants of regioisomers in accordance with the number of double bonds in the parent PUFA and surrounding conditions. The subsequent intramolecular rearrangements also occur randomly with the formation of different variations of the final product. Figure 3 shows a comparison of parent molecules of AA and DHA with the structures of the resulting possible products. For AA, these may be F2-isoprostanes and highly toxic isoketals, and for DHA, correspondingly, neuroprostanes and neuroketales (Figure 3). The products of IPLP have cis configuration relatively to the cyclopentane ring.

According to Antonenko et al. (2008) in experiments in vitro, cardiolipin is highly sensitive to peroxidation in the presence of "OH [44]. The principal distinction between the in vitro conditions described for "OH and HO₂⁺ radicals [18, 79], and the interactions of HO₂⁺ with AA, or "OH with CL, is that in vivo reactions proceed in the completely hydrophobic environment. According to Gebicki and Bielski [18], in the water-ethanol mixture, the reactions proceed in consent with the reaction sequence of the "classical" lipid peroxidation (LP), and abstraction of the first hydrogen atom from linoleic acid results in formation of hydrogen peroxide: LH + "OH → L* + H₂O₂, with the final formation of the stable end product linoleic hydroperoxide: LOO⁻ + LH⁺ → LOOH + L* [18]. Abstraction of hydrogen atom from linoleic acid by "OH results in the conversion of the "OH radical to water: LH + "OH → L* + H₂O. Bielski et al. [11] have concluded that reaction of HO₂⁺ radical with a double allylic H atom of a PUFA is proportional to the number of double allylic H atoms, which makes PUFAs highly specific targets with high affinity for perhydroxyl radicals [11, 18]. No selectivity was observed under similar conditions with hydroxyl radicals, which abstract H atoms randomly. According to Roberts and Morrows group [66, 74, 75], the reaction of HO₂⁺ with linoleic, linolenic, and arachidonic acids in water-ethanol solution, the formation of a stable products, which were the corresponding hydroperoxide [11], was relatively slow. On the contrary, studies on IPLP suggest that in the fully hydrophobic environment, reactions of HO₂⁺ with PUFA and linoleic acid are extremely fast, probably because of formation of hydroxyl radicals inside the hydrophobic milieu. For this reason, we can consider "OH as a "carrier" of 2 "OH inside the lipid core of the membrane.

As a result, the selective peroxidation of CL during aging, which we presume is driven to a large degree by formation of HO₂⁺ from the superoxide radical (O₂⁻), abstraction of the first H atom from CL results in formation of hydrogen peroxide. In the hydrophobic milieu, H₂O₂ undergoes homolytic splitting into two molecules of "OH, which instantly interact with another two H atoms of the same or adjacent linoleic acid followed by interaction of O₂ with formation of the stable linoleic hydroperoxide, which is still attached to CL. Molecular dynamic simulation studies have shown that upon oxidation, the molecule of linoleic acid changes its conformation (Figure 4) [64], which may affect the ability of CL to maintain the superstructure of the mitochondrial enzyme complexes. Thus oxidatively damaged CL is a specific hallmark for the oxidative stress in mitochondria.

6. IPLP as the Mechanism of Aging

Numerous studies have demonstrated that IsoPs are the most early and reliable among available markers of lipid peroxidation in vivo, and recent studies provided valuable information about participation of IPLP in pathogenesis of numerous human diseases [15, 16, 68, 69]. According to our model of IPLP initiation by HO₂⁺, the perhydroxyl radical upon encounter with a PUFA produces one of many variants of isoPG, iso-γ-ketoacids, or iso-levuglandins. When reacting with fatty acids with two unsaturated bonds, such as linoleic acids of cardiolipin, HO₂⁺ produces corresponding hydroperoxides.

A large number of data proves that oxidative damages to cardiolipin [63, 80, 81] and phosphatidylethanolamine
(PEA) [82, 83] may serve as markers for the mitochondrial autoxidation caused by aging. Above, we have presented our arguments that the major mechanism of oxidative damages of fatty acids esterified with phospholipids is associated with formation of HO$_2^*$ during “normal” production of superoxide radicals. The properties of perhydroxyl radical make it clear that regardless how small may be production of HO$_2^*$, it will cause some direct damages to mitochondria and other cellular membranes, or via the formation of adducts with PEA and lysine-containing proteins [82, 83], and peroxidation of cardiolipin [46, 63, 80, 81]. Even though the level of HO$_2^*$ production may be very low, it is, probably, the major mechanism of small, but persistent, accumulation of damages and regulatory signals caused by iso-prostaglandins, which, probably, often are wrong signals. Taking into consideration the crucial roles of CL and PEA in maintaining the structures of respiratory complexes and other polyenymatic complexes, we can state that during aging the primary causes of mitochondrial dysfunctions are not limited by mutations of mtDNA, but more likely are caused by the functional and structural changes of phospholipids and proteins of mitochondria themselves.

7. The Importance of Fatty Acids Oxidation for Increased Rate of ROS Production

Mammalian mitochondria generate superoxide and hydrogen peroxide (ROS) from at least 11 different sites associated with substrates catabolism and the electron transport chain [48]. All mitochondrial sites of ROS production have very distinct properties [48, 84]. They can be divided into two groups: six sites operate at the redox potential of the NADH/NAD$^+$ isopotential pool, about -280 mV, and five sites operate at the redox potential of the ubiquinol/ubiquinone (QH$_2$/Q) isopotential pool, about +20 mV [48, 84].

The increased mitochondrial respiration can result in either increased production of ROS, if the accelerated respiration was caused by increased substrate supply, or in a decrease of ROS production, if the accelerated respiration was the result of increased utilization of ATP [48]. Heart, skeletal, muscles, and brain may work at very different workloads. In order to produce more ATP during high workloads, the mitochondria must receive the correspondingly increased supply of electrons into the respiratory chain. The commonly used substrates, often regarded as complex I substrates, such as glutamate and pyruvate, have relatively low rates of respiration because the NADH dehydrogenase of complex I is the rate limiting step [85]. The accelerated rates of respiration may be achieved by activation of glutamate and pyruvate transamination, or by the use of substrates mixtures, such as pyruvate + glutamate + malate (for the brain). Under these conditions, activation of transaminase reactions in mitochondria produce, in addition to NADH, 2-oxoglutarate, which is then converted to succinate in the TCA cycle [85–87]. Brand has stressed the importance of fatty acids oxidation for the bioenergetics of the skeletal muscle and heart mitochondria [48, 89]. During β-oxidation of fatty acids by the multienzyme complexes, the electron-transferring flavoprotein-ETF-Q oxidoreductase system reduces the membrane pool of ubiquinone. In the presence of palmitoyl carnitine and other substrates (glutamate, pyruvate or succinate), there is no inhibition of SDH (complex II) [88], and thus electrons enter the respiratory chain both through the NADH/NAD$^+$ and QH$_2$/Q pathways allowing at high workloads fast production of ATP [48, 89, 90]. If consumption of ATP is limited, the excessive electrons activate at high membrane potential the reverse electron transport, which results in increased production of superoxide and H$_2$O$_2$ [48, 89].

We have suggested that in people with metabolic syndrome, which is characterized by increased utilization of fatty acids for production of energy, aging of the heart and brain may be accelerated [88, 90]. Indeed, recent studies support the cross-talk between the oxidative stress and metabolic conditions in mitochondria [91].

8. Oxidative Stress and the Mitochondrial Membrane Integrity

Mitochondria, particularly the inner membrane, contain very large amount of proteins, many of them in hundreds and thousands copies. Between proteins are located phospholipids, which comprise only around 20–25% of the total mass. Cardiolipin is located almost exclusively in the inner membrane of mitochondria (IMM). However, only phosphatidylcholine, together with few other phospholipids, form a biological membrane. Interactions of phosphatidylethanolamine (PEA) and cardiolipin (CL) with proteins allow integration of large proteins and multiprotein complexes into numerous curves of the inner membrane. PEA has a conical form because at C2, this phospholipid has a polyunsaturated fatty acid, usually arachidonic acid (C20:4, ω-6) or docosahexaenoic acid (C22:6, ω3), which have a curved shape. During oxidative stress, due to the high affinity of HO$_2^*$ to polyunsaturated fatty acids, PEA also undergoes peroxidation as well as the linoleic acid of cardiolipins. These two phospholipids are located close to each other because they share the function of maintaining the structural integrity of the IMM and multienzyme complexes. Therefore, oxidative damages to CL and PEA result in the malfunctions of the lipid-protein interactions and dysfunction of proteins. Accumulation of oxidized PEA and CL and their depletion are mitochondrial hallmarks of aging [80, 92, 93]. Recently, it has been shown that during aging and diabetes, the fatty acid composition of CL may undergo remodeling when C18:2 fatty acids can be replaced by arachidonic acid (C20:4, ω-6) or docosahexaenoic acid (C22:6, ω3) [94, 96], which are substrates for IPlP [13, 14]. Because CL is directly involved in formation of HO$_2^*$, it may explain why CL and PEA are particularly sensitive to oxidative damages [62, 92, 93].

It is conceivable that oxidative damages of even small amounts of CL and PEA may cause significant dysfunctions in energy production by mitochondria because both phospholipids are responsible for the superstructural organization of proteins involved in oxidative phosphorylation [96, 97]. In addition, CL is tightly bound with the transmembrane carriers of respiratory substrates, cytochrome c, and ANT. In the absence of CL, the electron transport between the respiratory complexes becomes disrupted, the membrane
potential drops significantly, and the synthesis of ATP becomes inhibited [98, 99].

In comparison with CL, the effects of PEA on mitochondrial functions are less studied. However, it was established that in the absence of PEA, the transport and assembly of proteins are also abnormal, only in this case, with the presence of only CL, the proteins tend to form unusually large superstructures, which result in the loss of membrane potential and ATP synthesis [98, 99]. Normal structural organization and functioning of respiratory chain and ATP-synthase requires coordinated interactions between CL and PEA.

9. Aging, Oxidative Stress, and Metabolic Syndrome

Oxidative stress is regarded as the main cause of aging, mitochondrial dysfunctions, and thus is one of the main pathogenic mechanisms of many diseases [100–105]. This is particularly true for those diseases, which are associated with the metabolic syndrome that develops at the certain stage of aging in humans [100–102]. It is clear, however, that the metabolic syndrome is not the result of only accumulated errors and dysfunctions in the course of life for whatever reason. From the point of view of the normal ontogenesis of an individual after birth, development of the metabolic syndrome represents a normal metabolic transition from the reproductive to post productive state of an individual.

Recently, we have shown that the full-scale oxidation of fatty acids by synaptic brain and heart mitochondria at all metabolic states occurs synergistically in the presence of other mitochondrial substrates, such as pyruvate, glutamate, or succinate [88]. We have also shown that oxidation of fatty acids may cause a several-fold increase in production of ROS [88, 90]. Figure 5 illustrates how simultaneous oxidation of palmitoyl-carnitine and other substrates affects production of ROS.

One of the important features of the metabolic syndrome is a dramatic increase in utilization of fatty acids for production of energy by mitochondria, particularly in women after menopause [106, 107]. From Figure 5, we can deduce that in aged individuals with metabolic syndrome, increased utilization of fatty acids may cause increased oxidative stress and thus accelerate the rate of aging.

10. The Rate of Aging is Proportional to the Rate of Superoxide Radical Production

Formation of HO$_2^*$ is proportional to the level of O$_2^*$ present at the interfaces of the inner membrane with the matrix and intermembrane spaces, where concentration of O$_2^*$ at any moment is determined by the rates of its production and elimination [48]. Taking into consideration that HO$_2^*$ is extremely reactive and dangerous, it is understandable that removal of superoxide radicals is very important for protecting cells from the deleterious effects of its protonated form as HO$_2^-$. The activities of SOD2 and SOD1 are of most importance for the heart and the central nerve system, where the alternative antioxidant systems are relatively weak, whereas the contents of AA and DHA are the highest [33]. Because HO$_2^-*$ interacts with PUFA and CL inside the membranes, any type of antioxidants, including superoxide dismutases, will have no effect on the aging caused by HO$_2^-*$ but will have effects on aging processes caused by other radicals. This means that the rate of HO$_2^-*$ formation is strictly dependent on the rate of superoxide radical formation. As we have discussed above, at old age, when fatty acids become the predominant substrates for energy provision, the rate of aging may also increase. Thus, to delay aging, first of all, it is necessary to delay production of superoxide radical, and secondly, to diminish conversion of superoxide into HO$_2^-*$. Local cellular conditions also affect the amount of O$_2^*$ converted to HO$_2^-*$, for example, accumulation of lactic acid during high physical loads or mild hypoxia may increase oxidative damages to skeletal muscle cells or cardiomyocytes due to the acidification-induced higher levels of HO$_2^-*$ production [17]. From this point of view, of particular interest for our discussions are recent publications of the Skulachev’s school of researchers on the effects of the mitochondria-targeted antioxidants with plastoquinone as an electron acceptor [46, 98–111].

Summarizing the data of several publications of the Skulachev’s research team, it has been concluded [98] that the plastoquinone-based mitochondria-targeted antioxidants operate in two quite different ways: (i) by preventing peroxidation of cardiolipin [46, 99] and (ii) by mild uncoupling of mitochondria resulting from fatty acid cycling that inhibits the formation of reactive oxygen species in mitochondria [111].

Shabalina et al. (2017) used the mtDNA mutator mice, which exhibit marked features of premature aging. The authors have suggested that accelerated aging was caused by the increased mitochondrial ROS that interact with polyunsaturated fatty acids in cardiolipin, releasing malondialdehyde and 4-hydroxynonenal that form protein adducts and thus diminish mitochondrial functions [109]. Treatment of animals with accelerated aging with SkQ1 counteracted these changes as it scavenges mitochondrial ROS. As the results, the normal mitochondrial ultrastructure was preserved in liver and heart; the phosphorylation capacity of skeletal muscle mitochondria as well as the thermogenic capacity of brown adipose tissue also improved. The SkQ1-treated mice lived significantly longer (335 versus 290 days) [109].

11. Conclusions

In this review, we examined the unique structure of cardiolipin and how this structure is reflected in several unique functions of this phospholipid. Cardiolipin (CL), together with phosphatidylethanolamine (PEA), maintain the superstructural organization of the enzymes that perform the most important function of mitochondria–oxidative phosphorylation. Both CL and PEA are not capable to build up a flat biological membrane, but they accommodate the functional complexes of proteins into the sharp curves of the inner mitochondrial membrane. At some regions of the IMM, cardiolipin creates regions with a strong negative charge, where local pH may be several units lower than in the
bulk phase of a compartment. We suggest that at these regions, the probability of protonation of the superoxide radical is relatively high. The resulting molecule of perhydroxyl radical is hydrophobic and specifically reacts with the double bonds of PUFA and linoleic acid of cardiolipin, producing correspondingly racemic mixture of isoprostanes and isoketals from PUFA and linoleate hydroperoxide. This explains why CL and PEA are the most common markers of the specific oxidative damages in mitochondria. Damages to CL and PEA cause gradual disarrangement of the structural organization of the functional complexes of oxidative phosphorylation as well as other enzymes of the

Figure 5: Production of superoxide radicals by rat heart mitochondria-oxidizing palmitoyl-carnitine. Designations: 1. supporting substrate only; 2. palmitoyl-carnitine only, and 3. palmitoyl-carnitine + supporting substrate. Substrates: Figure 5(a) (pyruvate 2.5 mM + malate 2 mM), Figure 5(b) (glutamate 5 mM + malate 2 mM), and Figure 5(c) (succinate 5 mM). Experimental conditions are described in [93]. The incubation medium contained Amplex Red 2 μM, horse radish peroxidase 2 units, substrates as indicated above, and volume 1 ml. The reaction was initiated by addition of 50 μg of mitochondria. Initial rates were measured for 3 minutes. Numbers at the traces are the rates of H_2O_2 production in picomol H_2O_2/min/mg protein RHM. The rates were corrected for the time control rate with RHM incubated without added substrates. The figure was taken from [95].
mitochondria. For this reasons, we consider that oxidative damages to CL and PEA are the primary cause of aging and the age-associated diseases. Therefore, developments of the mitochondria-targeted antioxidant drugs, similar to SkQ, are very promising in finding and making a reality the existence of the mythical “Elixir of Youth.”

Abbreviations

| Abbreviation | Description |
|--------------|-------------|
| AA           | Arachidonic acid (C20:4 ω6) |
| ANT          | Adenine nucleotide translocase |
| CL           | Cardiolipin |
| COX          | Cyclooxygenase |
| DHA          | Docosahexaenoic acid (C22:6 ω3) |
| HO2·         | Perhydroxyl radical |
| IMM          | Inner mitochondrial membrane |
| IsoPs        | Isoprostanes |
| IPLP         | Isoprostane pathway of lipid peroxidation |
| LOX          | Lipoxigenases |
| LP           | Lipid peroxidation |
| MetS         | Metabolic syndrome |
| mtDNA        | Mitochondrial deoxyribonucleic acid |
| NO           | Nitric oxide radical |
| O2·          | Superoxide radical |
| OH           | Hydroxyl radical |
| OMM          | Outer mitochondrial membrane |
| ONOO·        | Peroxynitrite |
| O2,NO2·      | Peroxynitrate |
| OXPHOS       | Oxidative phosphorylation |
| PEA          | Phosphatidylethanolamine |
| PGF2α        | Isoprostanes containing F-type prostanate rings |
| PGs          | Prostaglandins |
| PUFA         | Polyunsaturated fatty acids |
| RET          | Reverse electron transport |
| ROS          | Reactive oxygen species |
| SOD          | Superoxide dismutase |
| SOD1         | Cytoplasmic isoform of Cu,Zn-SOD |
| SOD2         | Mitochondrial isoform of Mn-SOD |

Conflicts of Interest

The authors declare that there is no conflict of interest regarding the publication of this paper.

Acknowledgments

This work was supported by R01HL124116 and R01HL144943 National Institutes of Health grants.

References

[1] D. Harman, “The aging process,” *Proceedings of the National Academy of Sciences of the United States of America*, vol. 78, no. 11, pp. 7124–7128, 1981.
[2] D. Harman, “Free radical theory of aging: consequences of mitochondrial aging,” *Age*, vol. 6, no. 3, pp. 86–94, 1983.
[3] D. Harman, “Free radical theory of aging: an update: increasing the functional life span,” *Annals of the New York Academy of Sciences*, vol. 1067, pp. 10–21, 2006.
[4] A. Chomyn and G. Attardi, “MtDNA mutations in aging and apoptosis,” *Biochemical and Biophysical Research Communications*, vol. 304, no. 3, pp. 519–529, 2003.
[5] B. Kadenbach, C. Münchner, V. Frank, J. Müller-Höcker, and J. Napiwotzki, “Human aging is associated with stochastic somatic mutations of mitochondrial DNA,” *Mutation Research/DNAging*, vol. 338, no. 1-6, pp. 161–172, 1995.
[6] R. M. Anson, E. Hudson, and V. A. Bohr, “Mitochondrial endogenous oxidative damage has been overestimated,” *The FASEB Journal*, vol. 14, no. 2, pp. 355–360, 2000.
[7] D. C. Wallace, “Mitochondrial genetics: a paradigm for aging and degenerative diseases?,” *Science*, vol. 256, no. 5057, pp. 628–632, 1992.
[8] M. Pinto and C. T. Moraes, “Mechanisms linking mtDNA damage and aging,” *Free Radical Biology and Medicine*, vol. 85, pp. 250–258, 2015.
[9] K. Szczepanowska and A. Trifunovic, “Origins of mtDNA mutations in aging,” *Essays in Biochemistry*, vol. 61, no. 3, pp. 325–337, 2017.
[10] K. L. DeBalsi, K. E. Hoff, and W. C. Copeland, “Effect of the mitochondrial DNA replication machinery in mitochondrial DNA mutagenesis, aging and age-related diseases,” *Aging Research Reviews*, vol. 33, pp. 89–104, 2017.
[11] B. H. Biesalski, R. L. Arudi, and M. W. Sutherland, “A study of the reactivity of HO2/O2- with unsaturated fatty acids,” *The Journal of Biological Chemistry*, vol. 258, no. 8, pp. 4759–4761, 1983.
[12] A. Panov, “Mitochondrial production of perhydroxyl radical (HO2•) as inducer of aging and related pathologies,” *Journal of Biochemical and Biophysical*, vol. 1, 2017.
[13] L. Jackson Roberts II, T. J. Montine, W. R. Markesbery et al., “Formation of isoprostane-like compounds (neuroprostanes) in vivo from docosahexaenoic acid,” *The Journal of Biological Chemistry*, vol. 273, no. 22, pp. 13605–13612, 1998.
[14] J. D. Morrow, J. A. Awad, H. J. Boss, I. A. Blair, and L. J. Roberts 2nd, “Non-cyclooxygenase-derived prostanoids (F2-iso prostanes) are formed in situ on phospholipids,” *Proceedings of the National Academy of Sciences of the United States of America*, vol. 89, no. 22, pp. 10721–10725, 1992.
[15] P. Montuschi, P. J. Barnes, and L. Jackson Roberts II, “Iso prostanes: markers and mediators of oxidative stress,” *The FASEB Journal*, vol. 18, no. 15, pp. 1791–1800, 2004.
[16] S. S. Davies and L. Jackson Roberts II, “Fj,-isoprostanes as an indicator and risk factor for coronary heart disease,” *Free Radical Biology and Medicine*, vol. 50, no. 5, pp. 559–566, 2011.
[17] A. Panov, “Perhydroxyl radical (HO2•) as inducer of the isoprostane lipid peroxidation in mitochondria,” *Molecular Biology*, vol. 52, no. 3, pp. 295–305, 2018.
[18] J. M. Gebicki and B. H. J. Biesalski, “Comparison of the capacities of the perhydroxyl and the superoxide radicals to initiate chain oxidation of linoleic acid,” *Journal of the American Chemical Society*, vol. 103, no. 23, pp. 7020–7022, 1981.
[19] D. A. Cherepanov, B. A. Fenikou, W. Junge, and A. Y. Mulkidjanian, “Low dielectric permittivity of water at the membrane interface: effect on the energy coupling mechanism in biological membranes,” *Biophysical Journal*, vol. 85, no. 2, pp. 1307–1316, 2003.
[20] A. D. N. J. de Grey, “HO2•: the forgotten radical,” *DNA and Cell Biology*, vol. 21, no. 4, pp. 251–257, 2002.
[21] T. H. Haines, “A new look at Cardiolipin,” Biochimica et Biophysica Acta, vol. 1788, no. 10, pp. 1997–2002, 2009.
[22] A. A. Benson, H. Daniel, and R. Wiser, “A sulfolipid in plants,” Proceedings of the National Academy of Sciences of the United States of America, vol. 45, no. 11, pp. 1582–1587, 1959.
[23] R. N. A. H. Lewis and R. N. McElhaney, “Cardiolipin, the heart of mitochondrial metabolism,” Cellular and Molecular Life Sciences, vol. 65, no. 16, pp. 2493–2506, 2008.
[24] J. Jouhet, “The physicochemical properties of cardiolipin bilayers and cardiolipin-containing lipid membranes,” Biochimica et Biophysica Acta, vol. 1788, no. 10, pp. 2069–2079, 2009.
[25] R. H. Houtkooper and F. M. Vaz, “Cardiolipin, the heart of mitochondrial metabolism,” Cellular and Molecular Life Sciences, vol. 138, no. 1-2, pp. 38–48, 2009.
[26] C. Arnarez, S. J. Marrink, and X. Periole, “Identification of cardiolipin binding sites on cytochrome c oxidase at the entrance of proton channels,” Scientific Reports, vol. 3, no. 1, 2013.
[27] M. Sathappa and N. N. Alder, “The ionization properties of cardiolipin and its variants in model bilayers,” Biochimica et Biophysica Acta, vol. 1858, no. 6, pp. 1362–1372, 2016.
[28] T. H. Haines and N. A. Dencher, “Cardiolipin: a proton trap for oxidative phosphorylation,” FEBS Letters, vol. 528, no. 1-3, pp. 35–39, 2002.
[29] M. Yates, J. Y. Syz, D. Gossler, and T. H. Haines, “pH-dissociation characteristics of cardiolipin and its Z'-deoxy analogue,” Lipids, vol. 28, no. 10, pp. 877–882, 1993.
[30] T. H. Haines, “Anionic lipid headgroups as a proton-conducting pathway along the surface of membranes: a hypothesis,” Proceedings of the National Academy of Sciences of the United States of America, vol. 80, no. 1, pp. 160–164, 1983.
[31] M. Schlame, M. Ren, Y. Xu, M. L. Greenberg, and I. Haller, “Molecular symmetry in mitochondrial cardiolipins,” Chemistry and Physics of Lipids, vol. 138, no. 1-2, pp. 38–49, 2005.
[32] C. Ye, Z. Shen, and M. L. Greenberg, “Cardiolipin remodeling: a regulatory hub for modulating cardiolipin metabolism and function,” Journal of Bioenergetics and Biomembranes, vol. 48, no. 2, pp. 113–123, 2016.
[33] S. E. Horvath and G. Daum, “Lipids of mitochondria,” Progress in Lipid Research, vol. 52, no. 4, pp. 590–614, 2013.
[34] J. M. Seddon, “Structure of the inverted hexagonal (H₁) phase, and non-lamellar phase transitions of lipids,” Biochimica et Biophysica Acta, vol. 1031, no. 1, pp. 1–69, 1990.
[35] J. A. Cohen and M. Cohen, “Adsortion of monovalent and divalent cations by phospholipid membranes. The monomerdimer problem,” Biophysical Journal, vol. 36, no. 3, pp. 623–651, 1981.
[36] T. J. Pinheiro, A. A. Duraslki, and A. Watts, “Phospholipid headgroup–headgroup electrostatic interactions in mixed bilayers of cardiolipin with phosphatidylycholines studied by 2H NMR,” Biochemistry, vol. 33, no. 16, pp. 4896–4902, 1994.
[37] J. M. Boggs, “Lipid intermolecular hydrogen bonding: influence on structural organization and membrane function,” Biochimica et Biophysica Acta, vol. 906, no. 3, pp. 333–404, 1987.
[38] M. Schlame, D. Rua, and M. L. Greenberg, “The biosynthesis and functional role of cardiolipin,” Progress in Lipid Research, vol. 39, no. 3, pp. 257–288, 2000.
[39] H. Schagger, “Respiratory chain supercomplexes,” IUBMB Life, vol. 52, no. 3-5, pp. 119–128, 2001.
[40] F. L. Hoch, “Cardiolipins and mitochondrial proton-selective leakage,” Journal of Bioenergetics and Biomembranes, vol. 30, no. 6, pp. 511–532, 1998.
[41] P. Adelroth and P. Brzezinski, “Surface-mediated proton-transfer reactions in membrane-bound proteins,” Biochimica et Biophysica Acta, vol. 1655, no. 1-3, pp. 102–115, 2004.
[42] V. P. Skulachev, “Chemiosmotic systems and the basic principles of cell energetics,” in Molecular Mechanisms in Bioenergetics, L. Ernster, Ed., pp. 37–73, Elsevier, Amsterdam, London, New York, Tokyo, 1992.
[43] V. P. Skulachev, “Chemiosmotic concept of the membrane bioenergetics: what is already clear and what is still waiting for elucidation?,” Journal of Bioenergetics and Biomembranes, vol. 26, no. 6, pp. 589–598, 1994.
[44] M. Schlame and M. Ren, “The role of cardiolipin in the structural organization of mitochondrial membranes,” Biochimica et Biophysica Acta, vol. 1788, no. 10, pp. 2080–2083, 2009.
[45] P. A. Srere, “Protein crystals as a model for mitochondrial matrix proteins,” Trends in Biochemical Sciences, vol. 6, pp. 4–7, 1981.
[46] Y. N. Antonenko, A. V. Avestisyan, L. E. Bakeeva et al., “Mitochondria-targeted plastoquinone derivatives as tools to interrupt execution of the aging program. 1. Cationic plastoquinone derivatives: synthesis and in vitro studies,” Biochemistry, vol. 73, no. 12, pp. 1273–1287, 2008.
[47] M. D. Brand, “The sites and topology of mitochondrial superoxide production,” Experimental Gerontology, vol. 45, no. 7-8, pp. 466–472, 2010.
[48] M. D. Brand, “Mitochondrial generation of superoxide and hydrogen peroxide as the source of mitochondrial redox signaling,” Free Radical Biology and Medicine, vol. 100, pp. 14–31, 2016.
[49] M. P. Murphy, “How mitochondria produce reactive oxygen species,” The Biochemical Journal, vol. 417, no. 1, pp. 1–13, 2009.
[50] J. S. Beckman and W. H. Koppenol, “Nitric oxide, superoxide, and peroxynitrite: the good, the bad, and ugly,” The American Journal of Physiology, vol. 271, pp. C1424–C1437, 1996.
[51] W. A. Pryor and G. L. Squadrato, “The chemistry of peroxynitrite: a product from the reaction of nitric oxide with superoxide,” The American Journal of Physiology, vol. 268, pp. L699–L722, 1995.
[52] W. A. Pryor, K. N. Houk, C. S. Foote et al., “Free radical biology and medicine: it’s a gas, man!,” American Journal of Physiology-Regulatory, Integrative and Comparative Physiology, vol. 291, no. 3, pp. R491–R511, 2006.
[53] W. A. Pryor, “Oxy-radicals and related species: their formation, lifetimes and reactions,” Annual Review of Physiology, vol. 48, pp. 657–667, 1986.
[54] B. H. J. Bielski, “Reevaluation of the spectral and kinetic properties of HO2AND O2–FREE radicals,” Photochemistry and Photobiology, vol. 28, no. 4-5, pp. 645–649, 1978.
[55] J. M. McCord and I. Fridovich, “Superoxide dismutase. An enzymic function for erythrocuprein (hemocuprein),” The Journal of Biological Chemistry, vol. 244, pp. 6049–6055, 1978.
in nonsynaptic brain mitochondria,” The Journal of Biological Chemistry, vol. 284, no. 21, pp. 14448–14456, 2009.

[87] A. Panov, Z. Orynbayeva, V. Vavilin, and V. Lyakhovich, “Fatty Acids in Energy Metabolism of the Central Nervous System,” BioMed Research International, vol. 2014, Article ID 472459, 22 pages, 2014.

[88] I. V. Perevoschikova, C. L. Quinlan, A. L. Orr, A. A. Gerencser, and M. D. Brand, “Sites of superoxide and hydrogen peroxide production during fatty acid oxidation in rat skeletal muscle mitochondria,” Free Radical Biology & Medicine, vol. 61, pp. 298–309, 2013.

[89] A. V. Panov, “Synergistic oxidation of fatty acids, glucose and amino acids metabolites by isolated rat heart mitochondria EC,” Cardiology, vol. 5, pp. 198–208, 2018.

[90] A. V. Panov and S. I. Dikalov, “Mitochondrial metabolism and the age-associated cardiovascular diseases,” EC Cardiology, vol. 5, pp. 750–769, 2018.

[91] S. I. Dikalov and A. E. Dikalova, “Crosstalk between mitochondrial hyperacetylation and oxidative stress in vascular dysfunction and hypertension,” Antioxidants & Redox Signaling, vol. 31, no. 10, pp. 710–721, 2019.

[92] T. Sen, N. Sen, S. Jana, F. H. Khan, U. Chatterjee, and S. Chakrabarti, “Depolarization and cardiolipin depletion in aged rat brain mitochondria: relationship with oxidative stress and electron transport chain activity,” Neurochemistry International, vol. 50, no. 5, pp. 719–725, 2007.

[93] G. Petrosillo, N. Moro, V. Paradies, F. M. Ruggiero, and G. Paradies, “Increased susceptibility to Ca(2+)-induced permeability transition and to cytochrome c release in rat heart mitochondria with aging: effect of melatonin,” Journal of Pineal Research, vol. 48, no. 4, pp. 340–346, 2010.

[94] H. J. Lee, J. Mayette, S. I. Rapoport, and R. P. Bazinet, “Selective remodeling of cardiolipin fatty acids in the aged rat heart,” Lipids in Health and Disease, vol. 5, no. 1, p. 2, 2006.

[95] X. Han, J. Yang, K. Yang, Z. Zhao, D. R. Abendschein, and R. W. Gross, “Alterations in myocardial cardiolipin content and composition occur at the very earliest stages of diabetes: a shotgun lipidomics study,” Biochemistry, vol. 46, no. 21, pp. 6417–6428, 2007.

[96] K. Beyer and M. Klingenberg, “ADP/ATP carrier protein from beef heart mitochondria has high amounts of tightly bound cardiolipin, as revealed by 31P nuclear magnetic resonance,” Biochemistry, vol. 24, no. 15, pp. 3821–3826, 1985.

[97] M. Schlame, K. Beyer, M. Hayer-Hartl, and M. Klingenberg, “Molecular species of cardiolipin in relation to other mitochondrial phospholipids: is there an acyl specificity of the interaction between cardiolipin and the ADP/ATP carrier?,” European Journal of Biochemistry, vol. 199, no. 2, pp. 459–466, 1991.

[98] J. B. Bultema, H.-P. Braun, E. J. Boekema, and R. Kouril, “Megacomplex organization of the oxidative phosphorylation system by structural analysis of respiratory supercomplexes from potato,” Biochimica et Biophysica Acta, vol. 1787, no. 1, pp. 60–67, 2009.

[99] I. Wittig and H. Schägger, “Supramolecular organization of ATP synthase and respiratory chain in mitochondrial membranes,” Biochimica et Biophysica Acta (BBA) - Bioenergetics, vol. 1787, no. 6, pp. 672–680, 2009.