The effect of iron binding on uranyl(V) stability†

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Here we report the effect of UO$_2$$^{4+}$–Fe$^{2+}$ cation–cation interactions on the redox properties of uranyl(V) complexes and on their stability with respect to proton induced disproportionation. The tripodal heptadentate Schiff base trensal$^{3-}$ ligand allowed the synthesis and characterization of the uranyl(V) complexes [UO$_2$(trensal)K], [UO$_2$(trensal)Fe(py)$_3$], [UO$_2$(trensal)Fe(py)$_3$U(trensal)], [UO$_2$(trensal)Fe(py)$_3$U(trensal)]$,^3$, [UO$_2$(trensal)Fe(py)$_3$U(trensal)]$^2$, [UO$_2$(trensal)Fe(py)$_3$U(trensal)]$^2$, and [UO$_2$(trensal)Fe(py)$_3$U(trensal)]$^2$. The uranyl(V) complexes show similar stability in pyridine solution, but the presence of Fe$^{2+}$ bound to the uranyl(V) oxygen leads to increased stability with respect to proton induced disproportionation through the formation of a stable Fe$^{3+}$–UO$_2$$^{4+}$–U$^{4+}$ intermediate [UO$_2$(trensal)Fe(py)$_3$U(trensal)]$^-$, upon addition of 2 eq. of PyHCl to 6. The addition of 2 eq. of PyHCl to 3 results in the immediate formation of U(V) and UO$_2$$^{2+}$ compounds. The presence of an additional UO$_2$$^{4+}$ bound Fe$^{2+}$ in [UO$_2$(trensal)Fe(py)$_3$U(trensal)]$^3$, does not lead to increased stability. Redox reactivity and cyclic voltammetry studies also show an increased range of stability of the uranyl(V) species in the presence of Fe$^{2+}$ with respect both to oxidation and reduction reactions, while the presence of a proton in complex 2 results in a smaller stability range for the uranyl(V) species. Cyclic voltammetry studies also show that the presence of a Fe$^{2+}$ cation bound through one trensal$^{3-}$ arm in the trinuclear complex [UO$_2$(trensal)$_2$Fe], 5, does not lead to increased redox stability of the uranyl(V) showing the important role of UO$_2$$^{4+}$–Fe$^{2+}$ cation–cation interactions in increasing the stability of uranyl(V). These results provide an important insight into the role that iron binding may play in stabilizing uranyl(V) compounds in the environmental mineral-mediated reduction of uranium(VI).

Introduction

Uranyl(V)$^1$ has been proposed as an important transient intermediate in the biological or abiotic mineral-mediated transformation of soluble uranyl(vi) compounds into the insoluble uranium(iv) dioxide (UO$_2$). These processes provide a convenient strategy to sequester uranium in the environment and, as such, are very important for ground-water remediation. In particular, stable adsorbed or incorporated uranyl(v) species have been reported to form during the U(v) reduction by Fe(II)-bearing minerals such as mica$^{14}$ or magnetite ([$\text{Fe}^{2+}$]$_{Fe}$$^{3+}$)$_2$O$_4$)$^{2-}$ and the presence of iron as the second nearest neighbour has been identified.$^4$ Uranyl(V) is a low-stability compound in aqueous media and they quickly disproportionate to uranyl(V) and U(iv)$^5$, but the incorporation into iron minerals may prevent disproportionation or further reduction of U(v) to U(iv) and thus lead to long-term immobilization of U(v). However, the role of iron binding to uranyl(V) species and their stabilization remains ambiguous in spite of its importance for the correct speciation of uranium in the environment.

Dinuclear or polynuclear complexes of uranyl(v) built from the interaction of a uranyl(v) oxo group with the uranium centre from a UO$_2$$^{4+}$ moiety (UO$_2$$^{4+}$–UO$_2$$^{4+}$), also known as cation–cation interaction (CCI),$^6$ have been proposed as intermediates in the proton promoted disproportionation of uranyl(v) to afford UO$_2$$^{2+}$ and U(iv) species.$^5$ The subsequent addition of protons to these polynuclear uranyl(v) intermediates leads to complete electron transfer followed by dissociation of the resulting U(v)–U(iv) complex. In aprotic media stable polynuclear UO$_2$$^{4-}$–UO$_2$$^{4-}$ complexes have been isolated.$^6$ We showed that the addition of protons (PyHCl) to a pyridine solution of stable tetrameric UO$_2$$^{4-}$–UO$_2$$^{4-}$ complexes leads to the immediate disproportionation of the uranyl(v) species affording uranyl(vi) and U(iv) complexes and water.$^6b$ Disproportionation of polynuclear cation–cation complexes was also observed in the absence of protons upon addition of strong Lewis acids ($\text{Li}^+$ or U$^{4+}$)$^{6b,7}$ to stable uranyl(v) Schiff base complexes and was found to lead to complex mixtures of soluble mixed-valent U(v) U(vi) uranium oxo clusters. It was also reported that the binding of strong Lewis acids or Group 1 metals to the uranyl(vi) oxo group renders more favourable the reduction of U(vi) to U(v).$^6b$
Moreover, it has been demonstrated that the binding of strong Lewis acids such as B(C₆F₅)₃ to the uranyl(v) oxo groups renders more accessible the reduction of U(v) to U(iv). A few studies have been directed to investigate the effect of the interaction of uranyl(v) with 3d transition metals on the stability and redox reactivity of uranyl(v) species. Moreover, in spite of the fact that several uranyl(v) complexes stable in organic solution have been isolated in recent years, only a few examples of heteropolymetallic complexes presenting a UO₂⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻~-~-~

The addition of FeI₂ to an unstable putative uranyl(v) dipotassium complex of a macrocyclic Schiff base Pacman ligand was reported to result in a higher stability of the uranyl(v) Pacman complexes which was corroborated by the isolation of the corresponding heterobimetallic UO₂²⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻~-~-~

Here we report two new stable complexes of uranyl(v) supported by the tripodal Schiff base ligand 2,2’-bis(3-salicylideneiminomethylmethanium): the UO₂²⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻~-~-~

The reaction of K₅trensal with the nitrate salt of uranyl (vi) leads to the isolation of the uranyl(vi) complex [UO₂(trensal)K], in 59% yield. The broad ¹H NMR spectrum of 1 in pyridine suggests the presence of fluxional solution species. A higher resolution of the ¹H NMR spectrum is observed in deuterated THF and a well resolved ¹H NMR spectrum could be obtained in CD₃OD solution (Fig. S2†). X-ray quality crystals of 1 could not be obtained, but the addition of one equivalent of PyHCl to a pyridine solution of 1 led to the isolation of X-ray quality crystals of the neutral complex [UO₂(Htrensal)], in 60% yield. The ¹H NMR spectrum of 2 in pyridine shows the presence of 15 overlapping narrow signals in agreement with the presence of C₄ symmetric solution species (Fig. S3†). The X-ray crystal structure of this complex is presented in Fig. 1 and shows that the uranium atom is heptacoordinated, with a slightly distorted pentagonal bipyramidal coordination geometry, by two uranyl oxygen atoms in the axial position and five donor atoms of the trensal⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻~-~-~

Fig. 1  Ellipsoid plot at 50% probability of 2 (co-crystallised pyridine molecule and hydrogen atoms were omitted for clarity, C atoms are represented in grey, O in red, N in blue and U in green). Selected distances (Å) U(1)–O(2) 2.222(3), U(1)–O(2) 2.240(3), U(1)–O(3) 1.783(3), U(1)–O(4) 1.787(3), U(1)–N(1) 2.584(3), U(1)–N(2) 2.625(3), U(1)–N(3) 2.599(3), O(5)–N(4) 2.619(5) Å.

Results and discussion

Uranyl(vi) and uranyl(v) complexes of trensal⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻~-~-~

The addition of FeI₂ to an unstable putative uranyl(v) dipotassium complex of a macrocyclic Schiff base Pacman ligand was reported to result in a higher stability of the uranyl(v) Pacman complexes which was corroborated by the isolation of the corresponding heterobimetallic UO₂⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻⁻~-~-~

An X-ray structure determination shows that the uranium atom in this complex is heptacoordinated, with a slightly distorted pentagonal bipyramidal coordination geometry, by two uranyl oxygen atoms in the axial position and five donor atoms of the trensal⁻⁻⁻⁻⁻⁻⁻⁻~-~-~

Scheme 1  Synthesis of the complexes 3 and 4.

those found in other reported Schiff base complexes of uranyl(vi). In the attempt to reduce the uranyl(v) complex 2 we added 1 eq. of decamethyl cobalocene (Cp*₂Co) to pyridine solutions of 2. The ¹H NMR spectrum of the resulting reaction mixture immediately after addition shows the presence of a large number of signals in the −45 to 45 ppm range suggesting that a putative uranyl(v) intermediate complex undergoes rapid disproportionation (Fig. S4†). This suggests that the phenol arm protonates the more basic uranyl(v) (compared to uranyl(vi)) oxo group resulting in proton induced disproportionation.

In contrast, the uranyl(v) complex [UO₂(trensal)K], is conveniently prepared in 70% yield from the salt metathesis reaction between K₅trensal and [[UO₂(Py₃)](K₂Py₃)]ₙ in pyridine (Scheme 1).

The ¹H NMR spectrum of 3 in deuterated pyridine showed the presence of fluxional species with signals in the paramagnetic region (−11 to 15 ppm) characteristic of U(v). Cooling down or heating up the NMR sample did not lead to a better resolution of the spectrum (Fig. S5†).

The addition of stoichiometric amounts of 2.2.2 cryptand (4,7,13,16,21,24-hexaoxa-1,10-diazabicyclo[8.8.8]hexacosane) to pyridine solutions of 3 resulted in a well resolved ¹H NMR spectrum (Fig. S6†). This suggests that fluxional potassium binding to the uranyl oxygen is the cause of the broad features in the ¹H NMR spectrum of 3.

The complex [UO₂(trensal)] [K[2.2.2crypt]K[2.2.2crypt]], was obtained analytically pure as a green solid in 62% yield. The solid-state structure of 4 was determined by X-ray diffraction
studies and is presented in Fig. 2. The overall quality of the crystal structure of compound 4 is rather poor (very weakly diffracting sample) but its connectivity is well determined.

The coordination environment around the uranium centre is similar to that found in complex 2. In 4 the uranium atom is heptacoordinated in a pentagonal bipyramidal coordination geometry. Five donor atoms of the trensal ligand (two oxygen and three nitrogen atoms) occupy the equatorial plane of the uranium ion, while the third arm of the trensal ligand does not interact with any cation and one [K(2.2.2crypt)] cation is found as an isolated ion in the unit cell of 4. The bipyramidal axial positions in 4 are occupied by two oxo ligands with U-O distances (1.824(15) and 1.865(16) Å) significantly longer than those found in the uranyl(v) complex 2 (1.785(3) Å). These distances are in the range of those found in previously reported complexes of uranyl(v). \(^{13a,13f}\)

The \(^1\)H NMR spectrum of 4 in pyridine shows the presence of 12 narrow signals in agreement with the presence of \(C_2\) symmetric solution species (Fig. S6†). This indicates that the molecular anionic fragment \([\text{UO}_2(\text{trensal})]\) \([\text{K}(2.2.2\text{crypt})]^-\) found in the X-ray structure of 4 dissociates in pyridine solution and the \([\text{K}(2.2.2\text{crypt})]^-\) cation is not bound to the uranyl(v) oxo group in pyridine solution.

The solid-state X-band EPR spectra measured at 298 K and 10 K revealed that the complex 3 is EPR silent.

In contrast, the solid-state X-band (9.40 GHz) EPR spectrum of 4 shows an intense signal at 10 K with a fitted rhombic set of \(g\) values \((g_1 = 2.44; g_2 = 1.10; g_3 < 0.6)\), confirming the presence of uranium in the oxidation state +5 (Fig. S29†). Notably, encapsulation of potassium enables us to obtain an EPR signal from the otherwise EPR silent complex 3. The likely presence of two potassium binding both uranyl(v) oxo groups results in a different electronic structure of 3 compared to 4 (where only one potassium cation is bound) which results in the absence of the EPR signal. \(^{13e}\)

Complex 4 is stable up to one month in the solid state and in pyridine and THF solutions. In order to assess the stability of these uranyl(v) complexes with respect to proton induced disproportionation, we have investigated the reaction of 3 and 4 with protons. After addition of 1 eq. of PyHCl to complex 3, partial disproportionation of the uranyl(v) complex was observed by \(^1\)H NMR spectroscopy. The addition of 2 eq. of PyHCl resulted in the complete disproportionation of the uranyl(v) to afford the uranyl(v) complex 2 and unidentified \(U(v)\) products as indicated by \(^1\)H NMR spectroscopy (Fig. S7†) and single crystal X-ray diffraction of the isolated crystal of the complex 2. The \(U(v)\) compounds formed in the disproportionation were identified as the product of the hydrolysis of the \([\text{U}(\text{v})\text{Cl}]\) complex as confirmed by the \(^1\)H NMR spectrum of a 1 : 1 : 2 mixture of \([\text{UO}_2(\text{trensal})]\), \([\text{U}(\text{trensal})]\)X \((X = \text{I}, \text{Cl})\) and \(\text{H}_2\text{O}\) (Fig. S8†).

On the other hand, addition of PyHCl to 4 initially resulted in the formation of NMR silent species, but after 3 days the \(^1\)H NMR spectrum shows the formation of the same disproportionation products as those found in the reaction of 3 with 2 eq. of PyHCl (Fig. S9†).

Iron binding to uranyl(v) complexes

In view of the potential important role of iron binding in the abiotic reduction of uranyl(v) as well as in the stabilization of uranyl(v) at iron mineral surfaces we have investigated the reactivity of complexes 1 and 3 with iron salts.

The reaction of 1 with FeI\(_2\) affords the trinuclear complex \([\text{UO}_2(\text{trensal})_2\text{Fe}](\text{py})_2\), 5, in 93% yield according to Scheme 2.

The solid state structure of 5 (Fig. 3) shows the presence of a neutral trinuclear complex where two \([\text{UO}_2(\text{trensal})]\) moieties are held together by a \(\text{Fe}^{2+}\) cation bound by two trensal O, N donor atoms not involved in the coordination of the uranyl cation. Thus, the replacement of the potassium cation in 1 with a \(\text{Fe}^{2+}\) cation leads to formation of a trinuclear structure.

In order to prepare a trinuclear uranyl(v) analogue we allowed 5 to react with Cp*Co. The \(^1\)H NMR spectrum after addition of 1 eq. of Cp*Co to complex 5 revealed the formation of a complex reaction mixture. One of the products could be identified by X-ray diffraction studies, revealing the formation of the dinuclear heterobimetallic complex \([\text{UO}_2(\text{trensal})\text{Fe}(\text{py})_3]\), 6. Addition of 2.5 eq. of Cp*Co to a pyridine solution of 5 led to an intractable reaction mixture from which none of the components could be identified (Fig. S11c†).

Complex 6 can be conveniently prepared in 81% yield from the reaction of FeI\(_2\) with complex 3 in pyridine in a 1 : 1 ratio (Scheme 3).

The solid-state structure of 6, represented in Fig. 4, shows the presence of a neutral dinuclear complex where a \([\text{U}^{\text{v}}\text{O}_2(\text{trensal})]\) dianion binds a \(\text{Fe}^{2+}\) cation through a \(\text{UO}_2^{2+}\cdots\text{Fe}^{2+}\) CCl.

**Fig. 2** Ellipsoid plot at 50% probability of the anion \([\text{UO}_2(\text{trensal})]\) \([\text{K}(2.2.2\text{crypt})]^-\) in 4 (H and \([\text{K}(2.2.2\text{crypt})]^-\) were omitted for clarity. C atoms are represented in grey, O in red, N in blue, K in light blue and U in green). Selected distances (Å) \(\text{U}(1)\text{–O}(3)\) 1.865(16), \(\text{U}(1)\text{–O}(4)\) 1.824(15).
The $\text{Fe}^{2+}$ cation is hexacoordinated, with a slightly distorted octahedral ligand, one uranyl oxo group and three pyridine molecules. The uranyl cation is heptacoordinate with a slightly distorted pentagonal bipyramid geometry, by two uranyl oxygen atoms and five donor atoms of the trensal$^{3-}$ ligand in the equatorial plane. The mean U–O bond lengths lie in the range of values typically observed for uranyl($v$) complexes,$^{10,11,13,14}$ with the UO$_2^{2-}$···$\text{Fe}^{2+}$ interaction resulting in a slight shortening of the bond (U1–O4 = 1.930 (2) Å and U1–O3 = 1.837(3) Å). The value of the Fe–O bond length (2.018(3) Å) falls in the range of those found in the only two previously reported examples of uranyl(V) complexes bridged to Fe($\text{II}$) via a hydroxo group.$^{16}$

The $^1$H NMR spectrum of 6 in pyridine results in the partial disproportionation of the uranyl($v$) complex (Scheme 4) with a 2 : 1 ratio of 6 to the disproportionation product $[\text{UO}_2(\text{trensal})\text{Fe}(\text{py})_3]\text{U}(\text{trensal})$Cl $\text{7b}$ (Fig. S20†). The addition of 2 equivalents of PyHCl to 6 led to the complete disappearance of the signals of complex 6 in the $^1$H NMR spectrum (Fig. S13†) and to an increased intensity of the signals assigned to $\text{7b}$. The presence of the uranyl($v$) complex $[\text{UO}_2(\text{Htrensal})]$ as the second disproportionation product was also identified by $^1$H NMR spectroscopy.

However, in both cases the disproportionation was not complete. Notably, the trinuclear cation–cation complex $\text{7b}$ contains unreacted uranyl($v$) (Scheme 4).

The formation of the $[\text{UO}_2(\text{Htrensal})]$ by-product prevented the synthesis of $\text{7b}$ from the reaction of 6 with PyHCl (Scheme 4).

However, the iodide analogue $[\text{UO}_2(\text{trensal})\text{Fe}(\text{py})_3]\text{U}(\text{trensal})][\text{I}], \text{7}$ was prepared in 80% yield from the reaction of complex 6 with 1 eq. of the $[\text{U}(\text{trensal})]$ complex in pyridine (Scheme 4). This complex is a rare example of an actinide-functionalized uranyl complex and only the third example of a uranyl($v$) complex presenting a CCI between the uranyl($v$) oxo group and a U($v$) cation.$^{11,12}$

The structure of complex 7 (Fig. 5) shows the presence of a cationic trinuclear complex built via CCI between the U($v$) center from the $[\text{U}(\text{trensal})]^{+}$ complex and the oxo group of the uranyl($v$) $[\text{UO}_2(\text{trensal})\text{Fe}(\text{py})_3]$ fragment. The three metal ions adopt a close to linear arrangement with a Fe–O–U angle of 170.3(3)° and a U–O–U angle of 171.2(3)°. The U=O bond distance for the uranyl($v$) oxo group bound to the Fe($\text{II}$) remains unchanged at 1.922(6) Å compared to complex 6, but a significant shortening of the U=O bond is observed upon binding of the U($v$) cation in 7 (1.960(6) Å). The UO$_2^{2-}$···U($v$) distance

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**Scheme 4** Addition of PyHCl to the complex 6 and synthesis of 7.
The binding of U(IV) to the uranyl(V) oxo group was previously reported to have a UO$_2$\(^{2+}$–U(IV) CCI (2.198(13) and 2.245(3) \(\text{Å}\)).

The UO$_2$\(^{2+}$–Fe$^{2+}$ distance (2.144(6) \(\text{Å}\)) is slightly longer than in 6 but is in the range of those found in the two previously reported complexes presenting a UO$_2$\(^{2+}$–Fe$^{2+}$ interaction (1.946(4) \(\text{Å}\)–2.132(4) \(\text{Å}\)).

These results indicate that the presence of Fe$^{2+}$ increases the stability of uranyl(V) in 6 with respect to proton-induced disproportionation. Notably the addition of 2 eq. of PyHCl led to full disproportionation of the complexes 3 and 4 while it resulted only in the partial disproportionation of 6 and the formation of [UO$_2$(trensal)] and of the Fe–U(V)–U(IV) trimer. The addition of five equivalents of pyridinium chloride is required for the full disproportionation of 6 to occur affording the same uranyl(V) and U(IV) disproportionation products as observed after addition of acid to 3. This indicates that the iron bound uranyl(V) complex 6 displays an increased stability towards the proton-induced disproportionation compared to the potassium bound uranyl(V) complexes 3 and 4 (Fig. S14†).

The binding of U(IV) to the uranyl(V) oxo group was previously reported to promote partial disproportionation and formation of multimetallic U(IV)–U(IV) oxo-bridged complexes. The stabilizing effect of Fe$^{2+}$ compared to U$^{4+}$ can be explained in terms of the lower Lewis acidity of Fe$^{2+}$ compared to U$^{4+}$. This results in the formation of stable UO$_2$\(^{2+}$–Fe$^{2+}$ adducts where the uranyl oxo group becomes less accessible to protonation by the Bronsted acid H$^+$. In contrast, complex 7 is stable in pyridine solution over one month period. \(^1\)H NMR studies show that the addition of [U(trensal)]$^+$ to complex 7 also leads to the formation of a stable unidentified compound (Fig. S15†). The subsequent addition of Fe$_2$I$_2$ to this compound led to the formation of complex 7. These results suggest that stable U(IV)–U(IV) complexes also form in the absence of iron bound to uranyl(V) oxo group. However, the formation of these compounds is not observed during the addition of PyHCl to 3, which undergoes complete disproportionation after the addition of 2 eq. of PyHCl. Moreover, the addition of 2 eq. of PyHCl to the U(IV)–U(IV) adduct results in full disproportionation, as indicated by the \(^1\)H NMR spectrum, suggesting that the binding of U(IV) to the uranyl(V) oxo does not lead to increased stability (Fig. S15†).

This further confirms the stabilizing role of Fe(II) binding with respect to proton-induced disproportionation of uranyl(V).

The \(^1\)H NMR spectrum of 7 in pyridine shows the presence of 45 signals over a large range of chemical shifts (–35 to + 53 ppm) in agreement with the presence of the trinuclear complex 7 in solution (Fig. S16†). Additionally, the ESI/MS spectrum of 7 in pyridine solution \{[UO$_2$(trensal)Fe$_2$U(trensal)$_3$]$^+$\} \(m/z = 1474.42\) indicated that the complex 7 retains its trinuclear structure in the pyridine solution (Fig. S27†).

The addition of 1 eq. of pyridinium chloride to 7 results in partial disproportionation with a 3 : 1 ratio of complex 7 to the disproportionation products as shown by \(^1\)H NMR spectroscopy (Fig. S17†). The complete disproportionation of complex 7 requires the addition of 4 eq. of PyHCl. The coordination of U(IV) does not increase the stability of the uranyl(V) species in 7 with respect to 6.

In view of the increased stability of 7 and 6 compared to 3 towards proton-induced disproportionation, we set out to investigate how the coordination of a second Fe$^{2+}$ cation to complex 6 would affect the structure and reactivity of the U(IV) centre.

The \(^1\)H NMR of the reaction mixture resulting from the addition of 0.5 equivalents of iron(II) iodide to 6 in pyridine indicated the formation of a new species (Fig. S18†). X-ray quality crystals of [[UO$_2$(trensal)Fe(py)$_3$$_2$Fe(py)$_3$]I$_2$, 8, were obtained in 65% yield from this reaction (Scheme 5). The solid-state structure (Fig. 6) of 8 shows the presence of a pentametallic structure where a Fe(Py)$_3$ bridges two iron-bound uranyl(V) [UO$_2$(trensal)Fe(py)$_3$] moieties. Overall, this results in the presence of UO$_2$\(^{2+}$–Fe$^{2+}$ CCI$s$ for both uranyl(V) oxo-groups. The central Fe$^{2+}$ cation is penta-coordinated by one oxo atom from each of the two uranyl(V) groups and three pyridine molecules. The mean Fe(2)–O(oxo) bond lengths is 1.988 Å. The [UO$_2$(trensal)Fe(py)$_3$] moieties of the crystal structure possess the same geometry found in the mononuclear complex 6, but the additional uranyl–iron interaction results in a slight shortening of the UO$_2$\(^{2+}$–Fe$^{2+}$ bonds compared to 6 (2.061(4) Å vs. 2.018(3) Å). The arrangement of the 5 metal ions is not linear, with a Fe$_2$Fe$_2$Fe$_3$ angle of 136°.

The ESI/MS spectrum of 8 (Fig. S28†) did not show the presence of a pentanuclear architecture in pyridine solution but showed only the peaks corresponding to the trimetallic [Fe$_2$U–Fe] and dinuclear [Fe–U] complexes. The addition of 1 eq. of the complex [UO$_2$(trensal)Fe(py)$_3$]$_2$ to a solution of 8 in pyridine resulted only in a slight broadening of the \(^1\)H NMR signals of 8 suggesting the presence in solution of a fast exchange between the [UO$_2$(trensal)Fe(py)$_3$]$_2$ moiety and 8 (Fig. S19†).

Scheme 5 Synthesis of [[UO$_2$(trensal)Fe(py)$_3$]$_2$Fe(py)$_3$]I$_2$, 8.
The labile binding of the central Fe(Py)32+ cation in 8 does not lead to an increased stability of 8 towards proton induced disproportionation compared to 6. Notably, the 1H NMR indicated a 2 : 1 ratio between the starting complex 8 and the disproportionation products upon addition of 1 eq. of H+ per uranyl(v) which is identical to the ratio observed for the complex 6 (Fig. S20†).

Redox reactivity
Iron binding to the uranyl(v) oxo is anticipated to have an important effect on its redox reactivity. Moreover, it has been suggested that iron binding at mica surfaces leads to the stabilization of uranyl(v) intermediates but the effect of iron binding on the redox properties of isolated uranyl(v) complexes has not been investigated.

At first, we explored the chemical oxidation of uranyl(v) by Fe(III). The reaction of 3 with 1 eq. FeCl3 leads to the oxidation of the uranium center (Fig. S21†) and to the formation of the uranyl(v)–Fe(n) complex 2 [(UO2(trensala))2Fe] as identified by X-ray diffraction crystallography and 1H NMR spectroscopy. The oxidation of uranyl(v) complex to uranyl(vi) by Fe(n) is explained in terms of the respective redox potential (Fe(n)/Fe(n−1) = 0.0 V; UO22+/UO2 = −1.6 vs. V (Fe/Fc†)).

In order to probe the possibility of obtaining a uranyl(v)–Fe(n) complex we explored the reactivity of 3 and 6 with increasingly electron-rich FeLCln complexes (L = tpa and tdmba), tpa = (tris(pyridin-2-y1methyl)amine) and H2tdmba = (tris-(2-hydroxy-3,5-dimethylbenzyl)amine). The reaction of 6 with [Fe(tpa)Cl2] led to the oxidation of uranyl(v) to uranyl(vi) with concomitant formation of [Fe(tpa)Cl2] (as shown by X-ray diffraction studies and 1H NMR spectroscopy). The reaction of 6 with the neutral Fe(n) complex [Fe(tdmba)] did not result in any change observable in the 1H NMR spectrum of 6 (Fig. S23†) indicating that the Fe(n) cation in [Fe(tdmba)] does not form CCIs with the uranyl(v) oxo group but does not oxidize the uranyl(v) either. In contrast, when complex 3 is reacted with [Fe(tdmba)], 1H NMR spectroscopy indicated that a redox reaction occurs yielding uranyl(vi) and Fe(n) species (Fig. S24†). These results are in agreement with the reported influence of chelating agents on the reoxidation by Fe(n) of biogenic products of uranyl(vi) reduction.37

These results suggest that the presence of UO2+····Fe3+ CCI stabilizes the uranyl(v) oxidation state with respect to the oxidation. In order to further probe the effect of iron binding on the redox properties of uranyl(v) species we performed comparative cyclic voltammetry studies of complexes 1, 2, 5, 6 and 8 (Fig. 6 and SCV1–SCV4†). The voltammogram of 1 in pyridine (Fig SCV1†) shows an irreversible redox event at −1.75 V, but when the voltammogram of 1 is measured in the presence of the cryptand a reversible redox event is observed at E1/2 = −1.69 V vs. Fe/Fc (Fig. 7, green curve). The voltammogram of the protonated uranyl(v) complex 2 also shows the presence of a reversible redox event at E1/2 = −1.66 V vs. Fe/Fc† assigned to the U(vi)/U(v) couple. These values compare well with the values previously measured in pyridine for other uranyl(v) complexes of tetradeinate (E1/2 = −1.61 V or −1.67 V vs. Fe/Fc†)20,21 and pentadentate Schiff bases (E1/2 = −1.58 V vs. Fe/Fc). A second irreversible redox event is observed at E1/2 = −2.47 V vs. Fe/Fc for complex 2, but not for complex 1. This event is consistent with the reduction of the metal centre (values of redox potential ranging from −2.02 to −2.88 V vs. Fe/Fc were previously assigned to the U(vi)/U(v) couple). The possibility that this event could be related to the reduction of the Schiff base ligand is unlikely since this feature is absent from the voltammograms of the H2trensala, K2trensala ligands and of the complex 1. Moreover, the shift of the U(vi)/U(v) couple to a more positive potential in complex 2 could be explained by the presence of a proton on the complex.

Similar redox events are observed in the voltammogram of complex 6 in addition to the quasi-reversible wave at E1/2 = 0.0 V vs. Fe/Fc†, assigned to the Fe(n)/Fe(n−1) couple. However, the U(vi)/U(v) reduction process is found at −1.03 V vs. Fe/Fc in the voltammogram of 6 and the second reduction event occurs at E1/2 = −2.7 V vs. Fe/Fc demonstrating that the range of stability of the uranyl(v) species is significantly extended compared to complex 2 as a result of Fe(n) binding. Both reduction and oxidation of the uranyl(v) cation are more difficult in the

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**Fig. 6** Molecular structure of the dication [(UO2(trensala)Fe(py)3)]2+ in 8 (H atoms were omitted for clarity. C atoms are represented in grey, O in red, Fe in dark blue, N in blue, I in purple and U in green). Selected distances U1–O2 = 1.920 (4) Å, U1–O5 = 1.935 (4) Å, U2–O6 = 1.927 (4) Å and U2–O9 = 1.927 (5) Å.

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**Fig. 7** Room temperature cyclic voltammograms of 4 mM pyridine solutions of [UO2(trensala)K] 1 in the presence of 1 eq. of cryptand (green), of [UO2(Htrensala)] 2 (red) and of [UO2(trensala)Fe(n)py] 6 (blue) recorded in 0.1 M Bu4NPF6 at 100 mV s−1 scan rate. Cp3Fe/Cp3Fe† (Fc/Fc†) corrected.
presence of Fe(II). No additional redox stabilisation was observed upon addition of two or more equivalents of Fe(II) to complex 6 as indicated by the voltammogram of complex 8 (Fig. SCV47). This is probably due to the labile binding of the second Fe(II) cation to the uranyl(V) oxo group in pyridine.

Moreover, in the voltammogram of complex 5 (Fig. SCV2f) the redox event assigned to the U(V)/U(VI) couple is found at $E_{1/2} = -1.66 \text{ V vs. Fe}^{2+}/\text{Fe}^{3+}$ as in complexes 1 and 2 in spite of the presence of a Fe(II) ion bound through the Schiff base acting as a bridging ligand. These results indicate that cation–cation interaction between the uranyl(V) oxygen and the Fe$^{3+}$ is essential for the stabilization of U(V) while the presence of a Fe(II) bound through the ligand has no significant effect on the redox properties of uranyl(V).

Conclusions

In conclusion the tripodal heptadentate Schiff base trensal$^3$ ligand allowed the synthesis and characterization of uranyl(V) complexes presenting UO$_2$$^+$...K$^+$ or UO$_2$$^+$...Fe$^{2+}$ cation–cation interactions. The reported uranyl(V) complexes show similar stability in pyridine solution, but the presence of Fe$^{2+}$ bound to the uranyl(V) oxygen leads to increased stability with respect to proton induced disproportionation. A stable Fe$^{2+}$...UO$_2$$^+$...U$^+$ intermediate (7b) containing both UO$_2$$^+$...Fe$^{2+}$ and UO$_2$$^+$...U$^+$ cation–cation interactions formed upon addition of 2 eq. of PyHCl to the iron bound uranyl(V) complex (6). In contrast, the addition of 2 eq. of PyHCl to the potassium bound uranyl(V) complexes (3 and 4) resulted in the immediate formation of U(V) and UO$_2$$^{2+}$ complexes. The UO$_2$$^+$...Fe$^{2+}$ (6) complex reacts with an additional Fe$^{2+}$ cation leading to the formation of a pentanuclear Fe$^{2+}$...UO$_2$$^+$...Fe$^{2+}$...UO$_2$$^+$...Fe$^{2+}$ complex (8) but the additional Fe$^{2+}$...UO$_2$$^+$ cation–cation interactions do not lead to increased stability. Redox reactivity and cyclic voltammetry studies also show an increased range of stability of the uranyl(V) species in the presence of Fe$^{2+}$ with respect both to oxidation and reduction reactions, while the presence of a proton in the uranyl(V) complex (2) results in a smaller stability range for the uranyl(V) species. Cyclic voltammetry studies also show that the presence of a Fe$^{2+}$ cation bound only through one trensal$^3$ arm in the trinuclear complex ([UO$_2$(trensal)$_3$]Fe), 5 does not lead to increased redox stability of the uranyl(V) demonstrating the important role of UO$_2$$^+$...Fe$^{2+}$ cation–cation interactions in increasing the stability of uranyl(V). These results provide an important insight into the role that iron binding may play in stabilizing uranyl(V) species in the environmental mineral-mediated reduction of uranium(V).

Conflicts of interest

There are no conflicts to declare.

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