Ionic liquids and deep eutectic solvents for the recovery of phenolic compounds: effect of ionic liquids structure and process parameters

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Water pollution is a severe and challenging issue threatening the sustainable development of human civilization. Besides other pollutants, waste fluid streams contain phenolic compounds. These have an adverse effect on the human health and marine ecosystem due to their toxic, mutagenic, and carcinogenic nature. Therefore, it is necessary to remove such phenolic pollutants from waste stream fluids prior to discharging to the environment. Different methods have been proposed to remove phenolic compounds from wastewater, including extraction using ionic liquids (ILs) and deep eutectic solvent (DES), a class of organic salts having melting point below 100 °C and tunable physicochemical properties. The purpose of this review is to present the progress in utilizing ILs and DES for phenolic compound extraction from waste fluid streams. The effects of IL structural characteristics, such as anion type, cation type, alkyl chain length, and functional groups will be discussed. In addition, the impact of key process parameters such as pH, phenol concentration, phase ratio, and temperature will be also described. More importantly, several ideas for addressing the limitations of the treatment process and improving its efficiency and industrial viability will be presented. These ideas may form the basis for future studies on developing more effective IL-based processes for treating wastewaters contaminated with phenolic pollutants, to address a growing worldwide environmental problem.

1. Background of study

The rapid development of population, industries, and burgeoning urbanization are contributing enormously to global environmental issues. Water pollution, in particular, is a problem that continues to worsen.1 Water is a very essential component of life, environment, and agriculture, and the quality of water influences both soil fertility and grain yield. According to the World Health Organization (WHO), in 2030, about 62% of the population of the world will suffer from water shortage. In developing countries, about 80% of wastewater is discharged into water bodies without any proper treatment.1 According to health experts, water-born diseases are increasing at a considerable rate and millions of people die annually on account of poor quality of water available for drinking. Approximately 1.2 billion people have no access to safe drinking water while 2.6 billion people have very little access.2 Water pollution is a severe problem, which is one of the leading causes of deaths and diseases worldwide.3 There are a large number of sources of water pollution, including domestic waste, industrial effluents, and the use of fertilizer, herbicides, and pesticides by the agriculture sector. These hazardous substances such as dyes, organic and inorganic pollutants, which have adverse effects on living organisms and biota, are the major cause of pollution in rivers, lakes, and oceans.3,4 There are many organic pollutants present in wastewater such as various organic solvents, phenolic compounds, dibenzofurans, dioxins, pesticides, chlorophenols, polychlorinated biphenyls, dyes, and new emergent organic pollutants.4

Among these, phenolic compounds are considered to be some of the most serious contributors to water pollution due to their high toxicity and carcinogenicity. These compounds are largely produced from different industrial processes and are often discharged into the environment without proper treatment.5 Due to the extensive usage of phenols in modern industries such as polymeric resin, paint, petroleum, and petrochemical industries, they have become prevalent in the environment causing severe water pollution.6 It is estimated that more than 10 million tons of phenol are discharged to the environment per year.7 The presence of phenol in water and air, even in very low concentrations, is a serious environmental
and safety concern. The permissible level of phenol discharge from industrial effluents should not be greater than 5 ppm. Furthermore, even discharge to inland or water bodies with concentrations as low as 1 ppm is considered undesirable and toxic. The environmental standard for phenol in portable drinking water should not exceed 0.001 mg L⁻¹ according to the World Health Organization (WHO).

The environmental standard for phenol discharge into the environment has been set at 1 ppm by governmental organizations including the European Union and Malaysian Environmental Protection Agency. Thus, it is a matter of great importance to remove these phenolic compounds from wastewater to meet the above stringent standards.

### 2. Phenol properties

Phenolic compounds exist in nature in different forms. The parent compound is phenol (C₆H₅OH) which has a specific pungent smell that is a disagreeably sweet, medicinal, or tar-like odor. It is a flammable compound and causes disintegration of some coating materials, rubber, and some plastics on contact. It exists naturally in different types of food and animal and human waste. Table 1 lists some major physical and chemical properties of phenol. The second naturally occurring form of phenol is phenolic acids, also called phenic or carbolic acids, which were isolated in 1834 from coal tar for the synthesis of carbolic soap due to its antiseptic properties and has been sold under the name of carbolic or phenol. Phenolic compounds are used widely in many industries including resin production. About 35% of phenol is used to produce low-cost phenol-formaldehyde resin. Phenol-formaldehyde has a large number of industrial applications and can be used in construction, adhesive material, and appliance industries and automotive. Similarly, about 28% of phenol is used to produce epoxy resin. About 16% of phenol consumption goes towards the production of a mixture of cyclohexanone and cyclohexanol, which is further converted into caprolactam, the monomer of nylon. Adipic acid, another important monomer of nylon, is also derived from phenol. Adipic acid is produced by oxidation of the mixture of cyclohexanone and cyclohexanol by nitric acid. Phenol may be converted into xylenols, alkylphenols, chlorophenols, aniline, and other secondary intermediates in the production of surfactants, fertilizers, explosives, paints, and paint removers, textiles, rubber, and plastic plasticizers and antioxidants, and curing agents and so on. The typical phenol concentrations in waste released by various industries are given in Table 2.

**Table 1** Physical and chemical properties of phenol

| Property                  | Value                        |
|---------------------------|------------------------------|
| Chemical formula          | C₆H₅OH                      |
| Molecular weight          | 94.11 g mol⁻¹                |
| Boiling point             | 181.75 °C                    |
| Melting point             | 40.9 °C                      |
| Heat of Fusion            | 122.2 J g⁻¹                  |
| Density at 40 °C          | 1.0545 g mL⁻¹                |
| Density at 60 °C          | 1.0413 g mL⁻¹                |
| Solubility in water       | 9.3 g L⁻¹                    |
| Vapor pressure at 25 °C   | 0.35 mm Hg                   |
| Color                     | White crystalline            |
| pKₐ                       | 9.89                         |
| The wavelength and maximum absorbance (λmax) | 270 nm |
| Henry’s law constant      | 0.034 pa m³ mol⁻¹            |

**Table 2** Phenol sources and their typical concentrations

| Industrial source         | Concentration range (mg L⁻¹) | Reference |
|---------------------------|------------------------------|-----------|
| Coal conversion           | 1700–7000                    | 24        |
| Coke oven                 | 600–3900                     | 25        |
| Phenolic resin            | 1200–1600                    | 26        |
| Petrochemical             | 200–1220                     | 27        |
| Textile                   | 100–150                      | 28        |
| Fiberglass industry       | 40–2564                      | 29        |
| Leather                   | 4.4–5.5                      | 30        |
| Pulp and paper industry   | 20–80                        | 31        |
| Paint industry            | 1.1                          | 32        |

4. Industrial applications of phenols

Phenolic compounds are used widely in many industries including resin production. About 35% of phenol is used to produce low-cost phenol-formaldehyde resin. Phenol-formaldehyde has a large number of industrial applications and can be used in construction, adhesive material, and appliance industries and automotive. Similarly, about 28% of phenol is used to produce epoxy resin. About 16% of phenol consumption goes towards the production of a mixture of cyclohexanone and cyclohexanol, which is further converted into caprolactam, the monomer of nylon. Adipic acid, another important monomer of nylon, is also derived from phenol. Adipic acid is produced by oxidation of the mixture of cyclohexanone and cyclohexanol by nitric acid. Phenol may be converted into xylenols, alkylphenols, chlorophenols, aniline, and other secondary intermediates in the production of surfactants, fertilizers, explosives, paints, and paint removers, textiles, rubber, and plastic plasticizers and antioxidants, and curing agents and so on. The typical phenol concentrations in waste released by various industries are given in Table 2.
| Methods                     | Advantages                                                                 | Disadvantages                                                                 | References |
|-----------------------------|----------------------------------------------------------------------------|------------------------------------------------------------------------------|------------|
| **Biological treatments**   |                                                                            |                                                                              |            |
| Biological degradation      | • Phenol is consumed by microorganisms such as bacteria, algae, yeast, and fungi and convert them to harmless compound | • Can lead to toxic by-products                                               | 67–69      |
|                             |                                                                            | • Growth control problem                                                     |            |
|                             |                                                                            | • Not suitable for high concentrations of phenol                             |            |
|                             |                                                                            | • Sludge production                                                          |            |
|                             |                                                                            | • Requires use of co-solvent when the concentration of phenol is low         |            |
|                             |                                                                            | • Growth control problem                                                     |            |
|                             |                                                                            | • Not suitable for high concentrations of phenol                             |            |
|                             |                                                                            | • Sludge production                                                          |            |
| Enzyme degradation          | • Enzymatic reactions are specific in nature and happened under moderate pH and temperature | • Non-reusability of the enzymes                                             | 69–71      |
|                             | • Higher catalytic efficiency and lower cost than the traditional chemical methods | • Enzyme instability in the harsh environment of the wastewater              |            |
| **Chemical treatments**     |                                                                            |                                                                              |            |
| Oxidative process           | • In gaseous oxidation, there is no increase in the volume of wastewater and sludge | • Use of expensive chemicals                                                |            |
|                             | • Simplicity of application                                                | • Incomplete oxidation of phenol                                              | 72 and 73  |
|                             |                                                                            | • Safety issue due to the use of hazardous chemicals                         |            |
|                             |                                                                            | • H₂O₂ needs to be activated by some other means                            |            |
|                             |                                                                            | • Wet oxidation of phenol is not economical due to the need for high pressure and temperature |            |
| Electrochemical destruction | • No need for expensive chemicals                                         | • Requires expensive equipment                                               | 74         |
|                             | • Sludge is not produced                                                   | • High energy consumption                                                    |            |
|                             |                                                                            | • Safety issues in handling toxic chemicals                                   |            |
| Photochemical               | • Phenols are greatly degraded, and sludge is not produced                 | • By-products are formed                                                     | 75 and 76  |
| Fenton reagents (H₂O₂ + Fe(II) salts) | • Fenton reagents are environmentally safe and therefore can be easily handled | • Expensive equipment is needed                                              | 77 and 78  |
| Irradiation                 | • Effective oxidation at lab scale                                         | • The requirement of a high amount of dissolved O₂                           |            |
| **Physical treatments**     |                                                                            |                                                                              |            |
| Membrane filtration         | • Removes all types of dyes                                                | • Concentrated sludge is produced                                            | 79         |
| Distillation                | • Economically feasible                                                    | • Production of large amounts of sludge                                       | 80         |
|                             | • Phenols are separated from aqueous media based on relative volatility    | • Needs high energy consumption                                              | 37         |
|                             |                                                                            | • Used for low concentration of phenol removal                                |            |
| Adsorption                  | • Good for removal of phenol                                               | • Regeneration is difficult                                                  | 4, 81–83   |
|                             | • Need mild temperature and pressure                                       | • Regeneration need calcination or the use of solvent                        |            |
|                             | • Economical                                                              | • Many adsorbents have low adsorption efficiency                             |            |
|                             | • Easy to operate and no expensive equipment’s required                     | • Not suitable for adsorption low level of phenol                            |            |
|                             |                                                                            | • Sometimes chemicals are used for adsorbent modification which are expensive and toxic |            |
| Liquid–liquid extraction    | • Easy to operate                                                         | • Sometimes need a high amount of adsorbent required                          | 66, 83–85  |
|                             | • Performed at mild conditions                                            | • Sometimes has low selectivity                                              |            |
|                             | • The extract can be recycled as a raw material                            | • Use of toxic, flammable, and volatile solvents in the extraction process   |            |
| Ion exchange                | • Regeneration of adsorbent                                                | • Regeneration of solvent might be expensive and challenging                 | 66         |
|                             |                                                                            | • Not effective for all dyes                                                  |            |
safely used in surgery since 1867. Various types of medicine, such as sore throat spray, as an oral analgesic for the relief of pain in or around the mouth, are prepared using phenol. In veterinary medicine, it is utilized as an internal and gastric anesthesia. Additionally, it was also utilized in the form of phenol injection for muscle tightening to treat a condition known as muscle spasticity and employed as a preservative for vaccines used against pneumonia, typhoid, smallpox, and polio. The most important pharmaceutical application of phenol is in the synthesis of aspirin. Furthermore, phenolic compounds derived from plants are known as natural antioxidants and can stop free radicals, thus protecting DNA from breakage and hence helping to prevent cancer.

5. Toxicity of phenol

It has been observed that the release of phenolic substituted compounds from industrial effluents into water bodies causes various human health problems. High exposures to phenol may be fatal to human beings, with infants being more susceptible. Direct human exposure to phenol can cause excretion of dark urine, sore throat, decreased vision and can irritate eyes and skin. Furthermore, phenol exposure can cause some chronic side effects such as anorexia, weight loss, diarrhea, vertigo, growth retardation, and inflammation in the gastrointestinal, liver, kidney, central nervous system, and cardiovascular tissues. Phenolic contaminated wastewater is discharged frequently to the aquatic system and, as such, adversely affects the aquatic ecosystem, vertebrates, invertebrates, algae, and protozoa. This may lead to growth inhibition and a decrease in the survival of their offspring. Thus, wastewater containing phenolic compounds must be treated properly before discharge into the environment.

6. Phenol removal from contaminated wastewater

As explained previously, the treatment of water contaminated with phenol is very important to reduce the impact on environmental and human health problems. The technology available for removal of phenol can be generally classified into three categories: chemical, physical, and biological treatment technologies, with each having its advantages and drawbacks. However, the wastewater coming from different sources can be complex and no single method can be used to treat wastewater effluent to the desired level. Practically, a combination of methodologies must be applied for wastewater treatment to achieve the desired output quality.

Chemical methods include chemical oxidation, photocatalysis, electrochemical, ion exchange, precipitation, irradiation, electroflotation, and coagulation/flocculation. The effectiveness of these methods depends on the pollutants in the wastewater and their interactions with the reactants. The chemicals used in these methods can either assist in the separation process or neutralize some of the toxic effects caused by the pollutants. However, in these methods, various expensive chemicals are used and, therefore, they are considered not economically feasible to be employed on the industrial scale. Furthermore, these methods generate large amounts of sludge and secondary pollutants after treatment resulting in sludge disposal problems. Powerful oxidizing agents such as hydroxyl radicals are produced during oxidative processes which are very effective in pollutants degradation.

Fig. 1 Number of publications per year on phenol extraction using ILs.
However, this approach is considered to be chemically and energy-intensive. In some chemical degradation methods, chlorine is used as an oxidizing agent which produces highly toxic compounds such as organochlorine. Hence, chemical methods are not preferred due to the associated high treatment and sludge disposal costs, which make such processes unattractive for use on an industrial scale.

Various physical methods such as electrodialysis, adsorption techniques, membrane filtration processes, nanofiltration, and reverse osmosis are employed for phenol removal from contaminated wastewater. The main shortcoming of the membrane separation method is the short lifetime of the membrane and, thus, it becomes economically unattractive due to the need for periodic replacement. Adsorption is extensively used in wastewater treatment due to its effectiveness in removing pollutants and decolorizing the water effluent. A properly designed wastewater treatment plant utilizing an adsorption system produces a high-quality treated effluent. Adsorption offers an attractive alternative treatment method for contaminated waters, particularly in cases where the adsorbent is inexpensive, locally available, abundant, and requires only simple pre-treatment before it is used.

Biological degradation of phenol is economical compared to the chemical and physical treatment process, but the processes involved are considered complex in nature. Various types of microorganisms such as algae, fungi, yeast, and bacteria have the potential to break down various types of pollutants including phenol. In this method, different types of microorganisms such as sphingomonas, pseudomonas strains, white-rot fungi, microbial cultures, under aerobic, anaerobic, or mixed conditions, are employed in phenol removal. However, biological degradation methods are time-consuming and hence, are considered impractical for the treatment of large quantities of industrial wastewater contaminated with phenol. The advantages and disadvantages of the available phenol removal technologies are listed in Table 3.

The liquid–liquid extraction technique is currently considered the most attractive method for the removal of phenol and other pollutants from contaminated wastewater. However, this method is generally applied where the concentration of phenol is very high. Liquid–liquid extraction is based on the
differential solubility of the phenols between the water phase and the water-immiscible solvent phase employed as the extractant. After contacting the phases and extracting the phenols, the two layers are separated from one another. This method is easy to operate under mild conditions and it does not cause any change to the extraction solvent or phenol. The phenol after extraction can, therefore, be recovered and reused as a raw material. Different types of organic solvents, such as hydrocarbons and oxygenated compounds, are used to extract phenol from wastewater.\textsuperscript{38,62} Most frequently, phenol is extracted from water using volatile aromatic and aliphatic organic solvents such as ketones, acetates, ethers, and alcohol. However, typical organic solvents have high vapor pressures, resulting in atmospheric contamination as volatile organic compounds (VOCs). Furthermore, they are usually highly flammable and are often toxic.\textsuperscript{63,64} Therefore, considering these environmental and safety issues, researchers are trying to design novel solvents, such as room temperature-hydrophobic ionic liquids (IL), which can efficiently extract phenol from wastewater in an environmentally benign and safe manner.

7. Overview of ILs

Though the exact definition of ILs is debatable, in general, ILs are organic salts that contain more than 95% ionic moieties (cations + anions) and have melting points below 100 °C.\textsuperscript{86-88} Therefore, various terms have been used in the literature, such as liquid electrolytes, ionic liquids, ionic fluids, ionic glasses, fused salts, and liquid salts.\textsuperscript{36-45} Due to their ionic nature, ILs have superior properties over other common organic solvents for certain applications and hence are considered to be attractive solvents to both academic and industrial fields.\textsuperscript{89-93} The most important properties of ILs are their negligible volatility, excellent solvation properties, and their thermal, electrochemical and chemical stability.\textsuperscript{92,93} The properties of ILs are effected by the chemical nature and structure of the anions and cations, as well as the alkyl chain length. Investigation of the recent literature reveals that interest in ILs as hydrophobic solvents for liquid–liquid extraction of phenol has increased considerably in the last few years. This is reflected by the significant increase in the number of publications, as illustrated in Fig. 1. This trend demonstrates the attractive properties of ILs as substitutes for organic solvents.

7.1 Quantitative structure–properties relationships for IL tunability

Interest in ILs for industrial use is increasing due to their physicochemical tunability that results from changing the combinations of anions and cations. Therefore, ILs can be designed according to the desired need by appropriate selection of anions and cations, while fulfilling the criterion that the designed salt has a melting point below 100 °C.\textsuperscript{86-94} To fulfill this very basic and fundamental requirement, the ions of ionic liquids should be asymmetric and univalent so that a close packing lattice structure is not possible, thus lowering the melting points to the desired level.\textsuperscript{86-94} Due to their tunability and excellent physicochemical properties, ILs have found several applications in the field of biomass processing, carbon dioxide and sulfur dioxide capturing, electrical energy storage, organic and inorganic synthesis, extraction of compounds from water, polymerization, drug delivery, extraction of collagen, bone filler, and lubricants, to name a few.\textsuperscript{95-98}

In general, ILs are designed by combining bulky organic cations such as imidazolium, pyridinium, piperidinium, pyrazolium, phosphonium, and ammonium with various organic or inorganic anions. The selection of ion pairs plays a major role in the thermo-physical properties as well as the solubility and miscibility of ILs in various organic and inorganic solvents.\textsuperscript{99,101-103} The physical properties such as density, surface tension, thermal properties, refractive index, viscosity, and acidity are key factors in the design and performance of ILs based industrial processes.\textsuperscript{102,103} These physicochemical properties can be tuned to a greater extent either by changing the cation or the anion.\textsuperscript{103} Hence, a wide range of experience in the selection of smart anion and cation is required to design ILs having desired physicochemical properties for the desired application. Literature review reveals that acyclic cations such as phosphonium, ammonium, sulphonium, cholinium, and cyclic cations such as imidazolium, pyridinium, pyrrolidinium, and piperidinium, are employed in the design of ILs (Fig. 2)\textsuperscript{87,89,104-106} with imidazolium cations being the most extensively studied.\textsuperscript{104} On the other hand, more flexibility is available for the selection of anions since these could be inorganic, such as halides, cyanate, phosphate, sulfate, nitrate, borate, as well as organic, such as phenolate, benzoate, malonate, and amines (Fig. 3). It is worth mentioning that the properties of ILs could be further altered by the functionalization of both the selected cation and anion.\textsuperscript{105-107} This functionalization could involve the incorporation of long-chain hydrocarbons, ether, hydroxyl, phenyl, nitrile, amide and halide groups into ILs.

7.2 Technical and ecological aspects of ILs

7.2.1 Thermal stability and melting point. Although they are often considered expensive, ILs can be economically feasible for industrial use when are thermally stable, have a wide liquid range, and can be easily regenerated for further use.\textsuperscript{108} The liquid range of an IL is the region between the melting point and the decomposition temperature. ILs with a high decomposition temperature and low melting point possess a wide liquid range and, therefore, can be used over a wider range of operating conditions. High thermal stability makes ILs suitable for thermal storage liquids applications in electric power, for example. IL thermal stability can be managed by careful selection of anions and cations.\textsuperscript{109} In general, IL melting point is a function of both the symmetry and charge of both cations and anions. Those ILs having symmetrical cations have a higher melting point than those having asymmetrical cations,\textsuperscript{110} and the melting point increases considerably with the elongation of alkyl chain length attached to the cation. For example, Shimizu et al. found that [C<sub>18</sub>mim][NTf<sub>2</sub>] melts at a higher temperature than ILs with shorter alkyl chain lengths. This increase in the melting point with an increase in alkyl chain length is due to the
increase in van der Waals interaction.\textsuperscript{111} Hence, it is vital to determine the thermal properties of ILs before their selections for the desired application. ILs containing anions derived from Bronsted acids and imidazolium cations tend to be thermally stable and are considered suitable to use at high temperatures.\textsuperscript{112–114} High thermal stability is also important to avoid the risk of an explosion when ILs are used in industrial applications.\textsuperscript{115}

### 7.2.2 Low vapor pressure

The low vapor pressure of ILs is one of the most vital and important properties of ILs, which makes them distinctive solvents for industrial applications.\textsuperscript{116,117} In contrast to the typically used volatile organic solvents, ILs have negligible vapor pressure and, therefore, do not escape into the atmosphere. ILs have very large latent heat of vaporization ($\Delta H_{\text{vap}} = 120–200$ kJ mol\textsuperscript{-1}), which is considerably higher than common organic solvents, thus enabling them to be used for the extraction of pollutants from water and oils without contamination of the atmosphere.\textsuperscript{114}

#### 7.2.3 Broad range of solubility and miscibility

One of the most important properties of ILs is their wide range of solubility and strong solvation properties in contrast to conventional organic solvents.\textsuperscript{64,104,114} It is known that ILs have the ability to dissolve a very wide range of organic and inorganic materials. For example, ILs can dissolve CO$_2$, H$_2$S, SO$_2$, cellulose, biomass,
inorganic salts, transition metals, antibiotics, coal, collagen, asphaltene, and enzymes. The broad range ability of ILs to solubilize different materials can be attributed to the large freedom in the selection of anions and cation for a specific application. For example, it has been shown that, even though the polarities of ILs and short-chain alcohols are quite similar, their solvation properties are quite different. The solvation properties of ILs can be easily managed for the dissolution of polar and non-polar substances by the selection of appropriate cations and anions. Efficient and realistic use of ILs as solvents requires the knowledge of their physical and chemical properties such as density, viscosity, surface tension, and thermal stability. Just as important, yet less studied, is knowledge of IL polarity, as this property reflects the solvation capability of these fluids for a given solute as well as reaction rates, reaction mechanism, product yield, and enzymatic activity among others. In this perspective, it is crucial to obtain information regarding the polarity of ILs before their utilization for a specific application.

7.2.4 Wettability and regeneration of ILs. One of the most important properties of ILs is their wettability of hydrophilic and hydrophobic surfaces. The wettability can be adjusted, like other physicochemical properties, by the careful selection of the cation–anion combination. Hydrophilic and hydrophobic properties of ILs are controlled by both anions and cations. ILs which have longer alkyl chain length cation are more hydrophobic. Similarly, the ILs containing bis-(trifluoromethanesulfonyl)amide ([NTf2]), hexafluorophosphate ([PF6]) and tetrafluoroborate ([BF4]) anions are hydrophobic in nature. This hydrophobic character and higher efficiency of ILs are the main features of ILs which make them distinct to be used in phase separation. Furthermore, the regeneration and recycling of ILs are considered the most important and vital property in ensuring the environmental and economic feasibility of their industrial application. As a result, this aspect has attracted the attention of many researchers in this field.

8. Limitations of ILs

Although ILs are considered a unique class of solvents having a large number of applications in various fields, they still suffer from certain drawbacks that limit their usage on commercial scales. These limitations include toxicity, high viscosity, high cost, and biodegradability. A short discussion on each is given below.

8.1 Toxicity and biodegradability of ILs

Though ILs can be considered greener than volatile organic solvents due to their low vapour pressure, non-flammability, thermal stability, and ease of regeneration, which minimizes their loss to the environment, recent research shows that many ILs can pose some risks to both terrestrial and aquatic biota. It has been documented in the literature that the toxicity of ILs depends on their cationic and anionic nature. It has been found the alkyl chain length and the nature of the functional groups are important factors in determining the toxicity of ILs. An increase in alkyl chain length on cations causes an increase in the toxicity of ILs. It has been observed that the toxicity of guanidinium based ILs towards Vibrio fischeri increases as alkyl chain length increases from C7 to C12 (ref. 135) However, it has also been reported in the literature that toxicity of some ILs increases with alkyl chain length up to a certain limit and, after that, no increase in toxicity is observed, which is called the “cut-off” effect. The toxicity of phosphonium based ILs against Vibrio fischeri also increases with the increase in alkyl chain length from [P4,4,4,1] to [P6,6,6,14]. However, incorporating certain functional groups onto the alkyl chain can reduce toxicity and enhance their biodegradability. Wang studied the effect of functionalization of cation on the toxicity of ILs against Clostridium sp. and found that 1-methoxethyl imidazolium tetrafluoroborate [Moemmi][BF4] which contains the methoxy functional group (CH3O) is comparatively less toxic than [C4mim][BF4]. Therefore, it is necessary to choose such cations and anions or functional group carefully before designing task–specific ILs to reduce their hazardous potential. The fluorinated anions such PF6 offer potential risk due to their tendency to be hydrolyzed. It has also been found that ILs based on choline chloride are less toxic than imidazolium, pyridinium, piperidinium, and pyrrolidinium-based ILs.

In order to minimize the loss of ILs to the environment, and improve their regenerability, attempts were made to immobilize ILs on solid supports. Furthermore, new types of IL containing cations derived from natural sources such as amino acids, choline, and fatty acids were developed with higher biodegradability and, thus, facilitating greener and more sustainable processes. The toxicity of ILs can be also mitigated by the proper selection of anions, with NO3 being shown to be environmentally friendly. Therefore, it is crucial to study the biodegradability of ILs as a function of ion pairs, functional groups, and alkyl chain length. Thus, well documented toxicological data of ILs is needed before they can be used extensively for industrial applications.

8.2 Viscosity of ILs

Viscosity is an important thermophysical property of ILs from an engineering point of view and it is required for many calculations such as mass transfer, modelling, fluid flow, and equipment design. However, the viscosities of ILs are higher than those for common organic solvents. This high viscosity is considered a great drawback and poses a serious limitation in industrial applications. On the other hand, high viscosity in certain applications, such as lubrication, is favorable. IL viscosity is considered an important property in the extraction process, however, there is no report available on the effect of viscosity on phenol extraction. It is expected that increasing the viscosity will lead to a decrease in the rate of molecular transport and hence reduce the efficiency of phenol extraction. Low viscosity ILs are considered better in most applications due to several factors such as handling, recovery, increased mass transfer rate, and having better extraction and catalytic properties. The viscosity of ILs can be altered through the selection of the ion pairs. Viscosity of ILs is also effected by...
addition of salts,

147,148 temperature and addition of molecular solvents.149,150 It is noted that a large number of possible combinations of anions and cations are possible so that the desired viscosity can be achieved.

8.3 Cost of ILs

The cost of ILs is considered one of the major factors limiting the use of ILs on an industrial scale. Compared to common organic solvents, ILs are considerably more expensive. The high cost of ILs can be attributed to the high price of precursors as well as the complex synthesis and purification processes.151,152 However, it is expected that the cost of ILs will decrease with economies of scale as the technology advances and the demand increases.153,154

8.4 Purity of ILs

The purity of ILs can considerably impact their physicochemical properties. For example, the presence of a small amount of impurities or water in ILs significantly affect properties such as viscosity.155 This factor is hard to control due to the purity of the precursors during synthesis and the limited purification techniques to obtain the ILs in pure form. For example, during the synthesis of [C₄mim][C₈H₁₇OSO₃] excess reactants and the solubility of hydrocarbon anion limit isolation of the product in the pure form.156

9. Effect of ILs structure on phenol extraction

9.1 Effect of anion

Hydrophobic ILs such as [C₄mim][BF₄], [C₄mim][PF₆], [C₄mim][BF₄], and [C₄mim][PF₆] were used to study the effect of both anion and alkyl chain length on the extraction of phenol, p-hydroxybenzoic acid and tyrosol.157,158 It was reported that ILs containing BF₄ anion were more efficient in the extraction of all the above compounds compared to ILs contain PF₆ anion.157 Furthermore, various phenolic compounds such as phenol, chlorophenol, 1-naphthol, 2-naphthol, proline, sarcosine, and lysine showed high efficiency of phenol extraction (>97%) with small differences between them.141 In this case, the nature of the cation predominates in determining phenol removal efficiency. In another study with an imidazolium-based IL containing lactate anion, similar observations were reported with efficiency towards phenol removal >99.9%.144 This observation further supports the previous conclusion that the nature of the cation could predominate over the anionic nature in determining the extraction efficiency. To test this hypothesis, two types of ILs with different combinations of anions and cations were used for phenol extraction.145 For addressing the effect of anions, two ILs contain the same cation [C₄C₁Py] with bis(fluorosulfonyl)imide (NF₂) and NTf₂ were tested for the extraction of phenolic compounds and showed that NTf₂ is more efficient than that of NF₂.162 In addressing the effect of the nature of cations, two ILs containing C₄C₁Py and C₄Py cation and NTf₂ and NF₂ anions were tested for the removal of phenol, with the former being the most efficient. This observation was attributed to the aromatic character of cations, which supposedly enhances the extraction efficiency. In a further study,146 ILs having [C₄mim] as cation with BF₄ and PF₆ as anion were tested for phenol removal and the results showed that the removal efficiency of phenol by BF₄ based ILs was higher than that with PF₆. This was due to differences in the nature of anions and their H-bonding interaction with phenolic compounds. Theoretical calculation by quantum mechanics supported the experimental result by confirming that the effective charge of BF₄ anion is stronger than that of PF₆.146 Hence, the H-bonding interaction of BF₄ with –OH of phenol is predicted to be stronger than PF₆ and will effectively remove higher amount of phenols.

9.2 Effect of cation

IL is constituents by anions and cations, the nature of cation also plays a major role in the extraction of phenolic compounds. Five ILs were selected to test with the aim to study the effect of the nature of cation on the removal efficiency of phenol.143 These include the cations [C₄Py], [C₄C₁Py], [C₄C₁Pip], [C₄C₁Py], and [C₄C₁Py] with NTf₂ as the anion. The results indicate that, in this case, the nature of the cation plays an important role in the extraction efficiency of phenol, with [C₄C₁Py] being the most efficient among those studied, followed by [C₄Py]. This observation was attributed to the aromatic character of the cation which might enhance the extraction efficiency of ILs. On the other hand, another study on the extraction efficiency of various phenolic compounds such as phenol, o-cresol, and resorcinol by ILs containing non-aromatic ([C₄mim][NTf₂]) and aromatic ([C₄mim][NTf₂]) cations,163 showed that ILs containing non-

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aromatic pyrrolidinium cations have higher efficiency than ILs containing aromatic ([C₆mim]) cations. Hence, the above conclusion regarding the aromaticity of cations cannot be generalized. Therefore, it can be concluded that H-bonding is the main contributor to the phenol extraction by these ILs. Egorov synthesized hydrophobic ILs containing ammonium cations, such as tetrahexylammonium dihexylsulfosuccinate (THADHSS) and trioctylmethyammonium salicylate (TOMAS) for phenol extraction. It was found that TOMAS was better in the extraction of phenol than THADHSS. The extraction properties of these two ILs were better than imidazolium-based ILs using for phenol extraction. The author claim that this higher extraction efficiency of both ILs might be due to the dispersive interaction of phenol with the cation of both ILs.

9.3 Effect of alkyl chain length

Pei et al. used [C₆mim] and [C₈mim] based ILs containing hydrophobic BF₄ anion instead of organic solvents such as phenol and phenyl amines. This study suggested that the extraction efficiency of phenolic compounds increases with the increase in alkyl chain length from C₆ to C₈. These observations indicate that the alkyl chain length leads to higher distribution efficiency of phenol in IL increases with increasing alkyl chain length on the imidazolium cation from C₄ to C₈. These observations indicate that the alkyl chain length leads to higher distribution of phenol in ILs by virtue of the enhanced hydrophobic interaction between the ILs and phenolic compounds. Fan et al. studied the impact of ILs structure on phenolic compounds extraction from water at pH = 7. To make the comparison easy, the ILs were classified into eight groups as shown in Fig. 4. These are: (I) [C₄mim][NTf₂], [C₄C₇im][NTf₂], [C₄C₉im][NTf₂], and [C₄C₁₂im][NTf₂]; (II) [C₈mim][PF₆], [C₄C₇im][PF₆], [C₄C₉im][PF₆] and [C₄C₁₂im][PF₆]; (III) [C₈mim][BF₄], [C₄C₇im][BF₄], [C₄C₉im][BF₄] and [C₄C₁₂im][BF₄]; (IV)

Fig. 4 Structure of ILs cations used by Fan et al.
[C4C6OHim][NTf2], [C4C8OHim][NTf2] and [C4C11OHim][NTf2]; (V) [C4C6OHi][PF6], [C4C8OHi][PF6] and [C4C11OHi][PF6]; (VI) [C4C6OHim][BF4] and [C4C11OHi][BF4]; (VII) [C4Beim][BF4] and (VIII) [C4Beim][NTf2]. As far as the alkyl chain length on cation is concerned, it has been observed that increasing the hydrophobicity of ILs by increasing alkyl chain length has no considerable effect on the phenolic compounds extraction. Khan et al. found that phenol extraction efficiency decreases with increasing alkyl chain length of cation of ILs. This finding further supports the above conclusion (Section 9.2) that hydrogen bonding plays a vital role in phenolic compound extraction in these cases.

9.4 Effect of functional groups on anion and cation

Few studies in the literature are available on the effect of incorporating functional groups to the ILs on phenol extraction. It is expected that functionalizing the cations and anions of ILs will considerably affect the extraction of phenol from aqueous media or crude oil. Fan et al. studied the effect of functional groups on the extraction efficiency of phenol from aqueous media using imidazolium-based ILs containing various anions (BF4, PF6, and NTf2) and alkyl chain length. In order to study the functional group effect on phenol extraction, benzyl, hydroxyl, and dialkyl functional group were incorporated into the IL cation. It has been shown that the incorporation of the hydroxyl group into the IL structure significantly increases the phenol extraction efficiency for ILs having NTf2 and BF4 as the anion. This increase in extraction efficiency is due to the increase in hydrogen bonding with the incorporated functional group. Furthermore, it has been confirmed by studying the thermodynamic parameters, that H-bonding plays a vital role in phenolic compound extraction. The negative values of the thermodynamic parameters of extraction such as ΔS and ΔH suggest that H-bonding is mainly responsible for phenolic compound extraction.

Yao et al. synthesized and employed four different dual functionalized ILs containing various cations for phenol extraction from oil. These are 1,2-(diethylamino)ethyl)-3-methyl imidazolium chloride ([Et2NEmim][Cl]), 1,2-(diethylamino)ethyl)-3-methyl morpholinium chloride ([Et2NEmnor][Cl]), 1,2-(diethylamino)ethyl)-3-methyl pyridinium chloride ([Et2NEmpyr][Cl]) and 1,2-(diethylamino)ethyl)-3-methyl pyridinium chloride ([Et2NEmpic][Cl]). The extraction efficiency of these ILs was compared with the already reported ILs such as choline chloride and [C4min][Cl]. It was found that all these newly prepared ILs are very efficient in phenol extraction, with a very low mole ratio of 0.3, which is almost half the amount of that used in conventional extraction. This high efficiency was attributed to the interaction of the Cl ion with the H atom of the hydroxyl group on phenol.

Recently, phenolic compounds have been extracted using IL-based aqueous biphasic systems (IL-based ABS). Aqueous solutions of hydrophobic ILs can converted into IL-based ABS by adding inorganic salts to the ILs. IL-based ABS is an alternative and attractive approach applied for separation processes instead of liquid-liquid extraction. [C4min][PF6] and [C2min][Cl]/(K2CO3, K2HPO4 or K2PO4) ABS were used for extraction of phenol, and 4-nitrophenol from aqueous phase. Other research group have used ABSs containing [C3-8mim][BF4], 6-(hydroxymethyl)oxane-2,3,4,5-tetrol and water to extract phenol. Various process parameters effect such initial phenol concentration, temperature, phase forming components, and alkyl chain length attached to imidazolium cation were investigated. Increases in the phase forming components, concentration, particularly the concentration of glucose and alkyl chain length attached to the imidazolium ring resulted in a considerable increase in the phenol extraction to IL-rich phase. The value of D for phenol was about 78, which is comparable with literature.

10. Effect of process parameters on phenol extraction

10.1 Effect of pH

It is well known that the pH of a sample solution could significantly influence extraction efficiency, particularly when acidic or basic solutes are extracted. Phenol speciation is known to be a function of pH, which leads to different forms due to the ionization of –OH group. At pH lower than 9.23 (<pK_a), the phenol exists in the molecular form (C6H5OH), whereas at pH > 9.23 (>pK_a), the phenolate ion (C6H5O^-) is the predominant form. NTf2-based ILs such as [C4min][NTf2], [C2min][NTf2] and [C4min][NTf2] were investigated for phenol extraction in pH range from 2 to 10. No change in phenol extraction was observed in the pH range from 2 to 6, however, it was found to sharply decrease when pH reached 10. Sualiman et al. also studied the effect of pH on the phenolic compounds using NTf2-based ILs and found that maximum removal is obtained at pH < 7. Deng et al. used trihexyltetradecylphosphonium tetra-chloroferrate ([3C6PC14][FeCl4]) for the phenolic compound extraction. A higher distribution ratio was found in acidic conditions for all phenolic compounds. Furthermore, an investigation by Egorov confirmed that the maximum distribution of phenol occurs at acidic pH levels and a drastic decrease in distribution occurs at pH > 12. In another study, Vidal et al. found that the extraction of phenol was higher at acidic pH levels using [C4min][BF4], [C4min][BF4], [C4min][PF6] and [C4min][PF6]. This higher percent removal of phenol at low pH could be attributed to the higher affinity of the IL toward the molecular form of phenol. This might be due to the hydrogen bonding between the –OH of phenol with ILs. However, at higher pH (above pK_a), the OH group of phenol ionized leading to the formation of phenolate ion (C6H5O^-), which results in a reduction of the strength of the H-bonding interaction between phenol and ILs.

10.2 Effect of initial concentration

Sas et al. extracted chlorophenol and resorcinol from water using the bis(trifluoromethylsulfonyl) imide [NTf2]-based ionic liquids (ILs). These include 1-methyl-3-propylimidazolium bis(trifluoromethylsulfonyl) imide ([C4min][NTf2]), 1-butyl-3-methylimidazolium bis(trifluoromethylsulfonyl) imide ([C4min][NTf2]), and 1-hexyl-3-methylimidazolium bis(trifluoromethylsulfonyl) imide ([C6min][NTf2]), 1-butyl-3-methylimidazolium bis(trifluoromethylsulfonyl) imide ([C4min][NTf2]) and 78, which is comparable with literature.
bis(trifluoromethylsulfonylimide ([C6mim][NTf2]). The starting concentrations of the phenolic compounds in water were 3, 5, 10, 100, 300, 500, 2000, 5000, 10 000, and 15 000 mg L\(^{-1}\) at pH lower than the pK\(_a\) values of the phenolic compounds in all cases. It was concluded that an increase in the initial concentration of the phenolic compounds causes an increase in the extraction efficiency by the selected ILs. Hence, it can be concluded that saturation has not been reached and the solubility of these phenolic compounds in the ILs phase is much higher than the concentrations used in these studies. Furthermore, NTf\(_2\) based ILs were used to study the effect of the initial concentration of phenol on its extraction efficiency. It was reported that the phenol extraction efficiency of all the selected ILs increases with the increasing initial concentration of phenol. Specifically, the extraction efficiency of the phenolic compounds from aqueous solution increases from 50% to more than 90% when the initial concentration of phenol is higher than 5 g L\(^{-1}\). On the other hand, a study by Gonzalez revealed that a decrease in phenolic compound extraction is observed with the increase in the initial concentration of phenol using NTf\(_2\) based ILs. This observation was supported by a study performed by Brinda et al. who found that the extraction efficiency of [Bmim][BF\(_4\)] decreases with an increase in the initial concentration of phenol. In addition, a study by Ji et al. used dicaticonic imidazolium-based ILs such as [1,2-bis(N,N'-methylimidazolium)]-ethane dibromide, 1,3-bis[N-(N'-methyl-imidazolium)]propane dibromide and 1,4-bis[N-(N'-methyl-imidazolium)]butane dibromide for phenol extraction from oil. In this case, a reduction in the extraction of phenol from oil was found as the initial concentration increases from 50 to 200 g L\(^{-1}\). However, Fan et al. reported no effect of the initial concentration of phenol on its extraction efficiency using 1-butyl-3-(11-hydroxyundecyl)imidazolium tetrafluoroborate.

### 10.3 Effect of phase ratio

Phase ratio is a crucial parameter in determining the efficiency of the extraction of the phenolic compounds by ILs. This parameter allows the optimum selection of the IL flow rate in any process which leads to improved economics on an industrial scale. Fan et al. used various imidazolium-based ILs contain PF\(_6\) and BF\(_4\) anions and various ranged having an alkyl chain length from butyl ([C4mim]) to octyl ([C8mim]) for phenol extraction. About 1 mL of IL solution was mixed with 5 mL of water solution (W), which gives a phase ratio of 1 : 5 (IL/W). In another study, Deng et al. used [3C6PC14][FeCl4] for phenolic compound extraction using phase ratios from 1 : 40 to 1 : 280 (IL/W). It was found that the optimum phase ratio for the maximum extraction of phenol is obtained at 1 : 120 (IL/W). The extraction efficiency decreases after 0.05 mL : 6 mL and therefore 0.05 : 6 mL was selected as the optimum. Fan et al. used hydroxyl, dialkyl, and benzyl ILs in various phase ratios for phenol extraction. By increasing the phase ratio, a decrease in extraction efficiency was observed. Maximum phenol extraction was obtained using a phase ratio 1 : 10 (IL/W). Furthermore, imidazolium and pyrrolidinium based ILs containing NTf\(_2\) as anion were used to study the effect of phase ratio on phenol extraction and reported an optimum value of 2 : 3 (IL/W) due to performance and economic values. [C6mim][NTf2] and ([C8mim][NTf2]) ILs was also used by Brinda et al. to study the effect of phase ratio in the range of 1 : 1, 1 : 2 and 1 : 5 (IL/W) on phenolic extraction from different concentrations of phenolic solutions. The extraction efficiency of phenol was found to decrease as the phase ratio increases from 1 : 3 to 1 : 5. However, no considerable difference was observed for phenol extraction when the phase ratio is between 1 : 1 and 1 : 2, indicating that the ratio of 1 : 2 is the optimum.

The ratio of the moles of IL to phenol in the aqueous phase was also used to optimize the required amount of IL for certain applications. Hou et al. used [C4mim][Cl] IL for the extraction of phenol from the aqueous phase. The mole ratio was varied from 0.1 to 3.5. It was found that at mole ratio of 1, 99.0% phenol was extracted, indicating that the optimum mole ratio is obtained at 1. Yao et al. used dual functionalized ILs ([Et2NEMPPr][Cl], ([Et2NEmim][Cl] and ([Et2NEMPic][Cl]) for phenol extraction using model oil. It has been observed that equilibrium for all the initial concentrations of 50, 100, and 200 mg L\(^{-1}\) of the phenolic solution is achieved when the IL to phenol mole ratio is 0.3 or more.

### 10.4 Effect of temperature

Hou et al. studied the effect of temperature in the range of 10 °C to 40 °C on phenol extraction from hexane using [C6mim][Cl] and [C4mim][PF\(_6\)] ILs. The results show that the phenol extraction efficiency decreases with increasing temperature. Specifically, the extraction efficiency decreases from 99.3% to 98.7% for [C4mim][Cl], and from 85.2% to 67.3% for [C4mim][PF\(_6\)] when temperature increased from 10 °C to 40 °C. Inspection of the above result reveals that the extraction of phenol using [C4mim][PF\(_6\)] is more sensitive to temperature than that using [C4mim][Cl]. This was attributed to the strong interactions between phenol and [C4mim][Cl] compared to [C4mim][PF\(_6\)]. In another study, the effect of temperature on phenol extraction was reported using ILs containing imidazolium cation [C8m-1mim] and PF\(_4\) and PF\(_6\) as anions in the temperature range of 15 °C to 45 °C. It was reported that, in contradiction to a previous study by Hou, no effect of temperature on phenol extraction efficiency was observed. Additionally, this observation was supported by a report from Vidal et al. who found no effect of temperature on phenol extraction using [C4mim][PF\(_6\)] and [C8mim][BF\(_4\)]. On the other hand, a report by Sas et al. using [C4mim][NTf2] and [C8mim][NTf2] ILs for the extraction of phenol at a high and low concentration from aqueous solution indicated that the temperature effect is concentration-dependent. Specifically, no considerable effect of temperature was observed for the high concentration range, whereas for low concentration range, a significant decrease in phenol extraction was observed with increasing temperature. Furthermore, Yao et al. reported that, upon using dual functionalized ILs such as [Et2NPPr][Cl], [Et2NEMPPr][Cl], and [Et2NEMPic][Cl] for extraction of phenol from model oil, the extraction efficiency decreases by 4% upon increasing temperature from 25 °C to 65 °C. This slight decrease
in phenol extraction with increasing temperature might be due to the negative enthalpy of extraction which renders the process exothermic.

11. Thermodynamics of phenol extraction

The thermodynamic study of phenol extraction is very important because it reveals the position of equilibrium as well as the spontaneity of the extraction process. Transfer of phenol from aqueous solution or from model oils to ILs phase depends on various types of intermolecular forces such as van der Waals, H-bonding, electrostatic, π−π interaction, and hydrophobic interaction.

The distribution of phenol between the aqueous phase and IL phase can be presented by the equilibrium given in eqn (1).

\[ \text{Phenol}_{\text{aq}} \leftrightarrow \text{Phenol}_{\text{IL}} \]  

(1)

The equilibrium constant in this case is represented by the distribution coefficient \( D \) and is given by eqn (2).\(^{176,177}\)

\[ D = \frac{[\text{Phenol}]_{\text{IL}}}{[\text{Phenol}]_{\text{aq}}} \]  

(2)

The distribution coefficient \( D \) is calculated using eqn (3).\(^{176,177}\)

\[ D = \frac{[C_i - C_{(w)}]}{C_{(w)}} \times \frac{V_{\text{aq}}}{V_{\text{IL}}} \]  

(3)

Thermodynamic functions for the extraction process, such as Gibbs free energy change \( (\Delta G) \), enthalpy change \( (\Delta H) \), and entropy change \( (\Delta S) \), determine the driving force involved in the transfer of phenol from aqueous media to IL media. These parameters can be determined using the eqn (4)–(6).

\[ \Delta G = -RT \ln D \]  

(4)

since: \( \Delta G = \Delta H - T\Delta S \)  

(5)

Hence

\[ \ln D = -\frac{\Delta H}{RT} + \frac{\Delta S}{R} \]  

(6)

where \( R \) is known as a universal gas constant in J mol\(^{-1}\) K\(^{-1}\). Eqn (6) implies that a plot of \( \ln D \) vs. \( 1/T \) will yield a straight line with slope \( = -\Delta H/R \) and intercept \( = \Delta S/R \).

The value of \( \Delta G \) is calculated at different temperatures using eqn (4). If the value of \( \Delta G \) is negative, then the removal of phenol from aqueous media to IL media is spontaneous and feasible under a given set of experimental conditions. It should be noted that eqn (6) was derived based on the assumption that \( \Delta H \) is temperature independent. However, in practice, \( \Delta H \) is highly dependent on temperature and it may be negative or positive. Hence, the van’t Hoff equation (eqn (7)) should be applied in this case to evaluate its value as a function of temperature.

\[ \left( \frac{\delta \ln D}{\delta T} \right)_p = \frac{\Delta H}{RT^2} \]  

(7)

If \( \Delta H \) is positive, indicating an endothermic process, it means that the extraction of phenol from aqueous to IL media increases with increasing temperature and vice versa.\(^{178}\) This dependence was also found to affect the entropy of the distribution. At around room temperature, its value is positive, indicating that a less ordered structure is presumably obtained in the IL phase. At higher temperatures, \( \Delta S \) becomes negative, indicating that phenol is trapped within the structured IL, which is most likely leading to the observed decrease in the entropy of the distribution. The extraction of phenol is favored at low temperatures which is economically advantageous. Jiang \( et \ al. \) also observed that the removal of cresol from aqueous media using ILs was exothermic in nature and the distribution coefficient decreasing with increasing temperature.\(^ {179} \) In another study, Chen \( et \ al. \) also found that the removal of methylene blue dye from the aqueous phase is exothermic and decreases with an increase in temperature from 20 °C to 45 °C and is accompanied by simultaneous decrease in the distribution coefficient.\(^ {182,183} \)

12. Phenol extraction from model oil using ILs

Besides the extraction of phenolic compounds from aqueous media, ILs were also used for the extraction of phenolic compounds from model oil. Mostly hydrophilic ILs are used to extract phenolic compounds from model oils. Imidazolium based, 1-ethyl-3-methyl imidazolium lactate ([Emim][Lac]) was used for the extraction of phenol from model oil.\(^ {180} \) This IL has the potential to extract phenol from model oil as investigated by COSMO-SAC and demonstrated experimentally. The extraction mechanism shows that the IL extracts phenol via the formation of H-bonds. The extraction efficiency of [Emim][Lac] for phenol reached 99.9% in 30 min at room temperature using [Emim] [Lac]/model oil mass ratio of 1 : 1. Amine based ILs, such as propylamine formate ([PA][FA]) and propylamine acetate ([PA][Ac]), were synthesized and characterized for the extraction of phenolic compounds from coal tar model oil.\(^ {180} \) P-Cresol was extracted by [PA][FA] and [PA][Ac] under different experimental conditions. Among the tested ILs, [PA][FA] has higher efficacy of 97.8% and a distribution coefficient of 27.59 using IL/model oil in 0.2 ratio at room temperature. ILs having dual basic sites were designed by a simple neutralization reaction of 1,1,3,3-tetramethylguanidine acid (TMG) with dihydrogen phosphate, tetramethylguanidine tetrafluoroborate ([TMG][BF₄]) showed higher extraction efficiency and removed 98.2% phenol in 35 min at room 30 °C. Sidek \( et \ al. \) synthesized IL containing benzyl imidazolium cation with different substituents such as vinyl, ally, and benzyl for extraction of phenolic compounds from hexane as the model oil.\(^ {182} \) The effects of process parameters such as phase ratio, contact time, and temperature, were studied and optimized to achieve
higher removal efficiency of phenolic compounds. The ILs having allylic substituents were best among the tested ILs and demonstrated 95% efficiency for removal of phenolic compound under optimized conditions. Besides the extraction of phenol from hexane, various other model oils such as ether, heptane, cyclohexane, and petroleum were used. No considerable change in mass and efficiency of ILs were noted after six cycles. Different types of tetraethylammonium amino acid (TAAA) based ILs, having no corrosive halide, were synthesized for the extraction of phenol from the oil mixture.183 The effect of process parameters such as extraction time, phenol concentration, ILs type, types of phenol, water contents in phenol on the extraction of phenol using amino acid-based ILs were investigated. The extraction efficiency of 99.0% was achieved using IL: phenol mole ratio of 0.60. After regeneration and reuse, no decrease in efficiency of ILs for the extraction of phenol was found. Zhuang et al.184 used imidazolium-based ILs such as [C2mim][BF4], [C2mim][PF6], [C4mim][PF6] and [Emim][NTf2] for phenolic compounds extraction from model oil. The effect of the phase volume ratio of ILs to phenol, extraction time (0–80 min), and temperature on extraction efficiency was studied. It was found that all these ILs efficiently removed phenolic compounds from oil. Gai et al.185 reported on the application of ILs containing tetramethylguanidinium cation and anions derived from l-proline, acetic acid, and tetrafluoroboric acid for phenolic compounds extraction from model oil. The results revealed that the order of the efficiency of phenol extraction follow the following order: [BF4]– > [Ac]– > [Pro]–. This difference in extraction efficiency was attributed to differences in anions structure and electronegativity. The lower efficiency of [Pro] anion based ILs was due to its bulky size, which creates steric hindrance towards the interact with phenol.

13. Phenol adsorption on solid supported ILs

Liquid-liquid phase extraction of phenolic compounds from aqueous solutions using hydrophobic ILs has been extensively studied. However, extraction using ILs has serious limitations when applied in practical industrial applications for wastewater treatment. To improve its performance in practical applications, the support of ILs on solid substrates provides a potential practical route that is gaining more focus in separation fields in recent years.185 More specifically, solid-supported ILs for phenolic compounds extraction is also an area of interest. Currently, various researchers have reported the use of supported ILs on different solid supports for the extraction of phenolic compounds. Zhang et al. synthesized novel polymeric ILs using imidazolium monomers.186 The synthesized ILs have a porous structure and Lewis basic active sites which efficiently adsorbed both small and large phenolic molecules from water. The adsorption capacity of tannic acid, 4-nitrophenol, and 4-chlorophenol on the synthesized solid polymeric ILs were 911, 460, and 433 mg g⁻¹, respectively. This higher adsorption capacity of polymeric ILs is attributed to the Lewis basic sites which are due to the N and O atoms. Besides, higher adsorption efficiency for phenolic compounds, the polymeric ILs can be easily regenerated and reused. Another research group has synthesized polystyrene-based resin supported ILs as an adsorbent for the removal of p-nitrophenol from aqueous media.187 The adsorbent was used in batch and continuous flow systems and adsorption efficiency of 1269.8 mg g⁻¹ was achieved within 30 min. Besides promising efficiency for nitrophenol removal, the adsorbent was recycled and reused 10 times without any loss in the adsorption potential for nitrophenol.

Zhu et al. synthesized an IL functionalized polymer by grafting 1-butyl-3-vinylimidazolium bromide, which is used as a monomer, on the surface of silica.188 The silica-supported polymeric material has a rough surface with an area of 205.49 m² g⁻¹. Grafting ILs onto the silica surface significantly improved the adsorption efficiency for phenolic compounds (2,4-dichlophenol, 4-dinitrophenol, and bisphenol). The kinetic study revealed that the adsorption of phenolic compounds on the solid-supported ILs followed a pseudo-second order kinetic model. The adsorption capacity of this synthesized solid-supported ILs for 2,4-dichlophenol, 2,4-dinitrophenol, bisphenol, and nitrophenol were 239.7, 64.28, 56.86, and 68.39 mg g⁻¹, respectively. Graphene oxide (GO) nanocomposite with 1-amino-3-methylimidazole chloride was synthesized and used for the removal of phenol from aqueous media.189 The GO-IL composite attained a surface area of about 110.44 m² g⁻¹ and a total pore volume of 0.2839 cm³ g⁻¹. The experimental results showed that the adsorption efficiency of GO-IL nanocomposite for phenol was 95.3%. Layered double hydroxide (LDH) has good potential to adsorb phenolic compounds.190 IL functionalize Zn₃Al-LDH was synthesized for the removal of phenol. Zn₃Al-LDH was functionalized with Aliquat 336 using a co-precipitation ultrasonication method. Among the various synthesized adsorbents, IL-Zn₃Al showed a higher adsorption efficiency (64.7 mg g⁻¹) for phenol adsorption from aqueous media. Polymeric ILs modified with graphene oxide-grafted silica (GO-SiO₂) were synthesized for the extraction of phenolic compounds.191 The surface of silica modified with polymeric ILs has higher positive potential and therefore attained strong electrostatic interaction for acidic compounds than the native GO-SiO₂. Marwani et al. synthesized192 ILs based solid-supported composite using sol–gel method for phenol adsorption from aqueous media. A new composite material (SiO₂-CIPrNTf₂) was prepared from silica and chloropromazine bis-(trifluoromethanesulfonimide) for the adsorption of 4-chlorophenol from water. The IL supported solid material has 626.25 mg g⁻¹ efficiency for 4-chlorophenol from water at pH 1. The performance of this IL-based composite was confirmed by applying it to a real sample with satisfactory separation results. Zhu et al.193 synthesized N-butylimidazolium functionalized chloromethylated macroporous styrene-divinylbenzene copolymer with the aim of adsorbing phenol from aqueous solutions. The synthesized IL functionalized polymer can remove phenol from both ion acidic and basic media. The maximum adsorption capacity was 92.9 mg g⁻¹ at pH 11. The adsorption mechanism shows that, in an acidic medium, the adsorption on the surface is mainly molecular, while that in alkaline medium is through anion exchange. Balasubramanian et al.194 prepared emulsified liquid membrane (ELM), by dissolving [Bmim][PF₆]...
in tetrabutyl phosphate for the extraction of phenolic compounds such as phenol, p-chlorophenol, 2,4-dichlorophenol, 2,4,6-trichlorophenol and pentachlorophenol from synthetic an aqueous solution. The IL-based ELM has a higher removal efficiency for the extraction of phenolic compounds which follows the order: phenol (99.5%) > p-chlorophenol (95.84%) > 2,4- dichlorophenol (93.12%) > 2,4,6-trichlorophenol (91.07%) > pentachlorophenol (90.53%).

Garavand et al. used a microemulsion liquid membrane. A microemulsion liquid membrane (MLM) extractor was constructed for the separation and the concentration of phenolic compounds from pistachio peeling effluent water streams. The extraction efficiency was 64% for the MLM compared to 46% for the corresponding emulsion liquid membrane (ELM).

14. Deep eutectic solvents for phenol extraction

Deep eutectic solvents (DESs) have emerged as green solvents to be used as an alternative to the conventional ILs. By definition, a DES is a solvent formed by a combination of proper hydrogen bond donor (HBD) and a hydrogen bond acceptor (HBA), mainly hydrogen bond interaction is the main type of complexation which results in a new solvent having a melting point lower than two components. DES having properties analogous to conventional ILs, with the advantage of being easily synthesized using low-cost precursors with high purity. The interest in using DES has increased due to their greener and environmentally friendly nature, cost-effectiveness, and favorable physicochemical properties. A DES is chemically tunable, so it can be synthesized easily for targeted applications by the proper selection of HBD and HBA. However, most of the DES reported in the literature have a hydrophilic nature and, therefore, cannot be used in aqueous media. This is due to the fact that the strong hydrogen bonding caused by water molecules results in breaking the HBD–HBA complexation and dissolution of the DES components. Consequently, the functionality of the DES is inhibited and this makes it difficult to be separated. However, in the last few years, DES having hydrophobic characteristics have emerged as new potential solvents to be used for extracting non-polar and inorganic compounds from aqueous media. It is expected that, in the near future, the use of hydrophobic DES will replace the use of hydrophilic ILs used for the extraction of various pollutants from aqueous solutions.

Florindo synthesized hydrophobic DESs based on a group of fatty acids that may act as HBD and HBA for the removal of bisphenol from aqueous solutions. This new DES was prepared by using various types of fatty acids such as octanoic, nonanoic, decanoic, and dodecanoic acid. These fatty acids-based DES are hydrophobic and stable in water. The extraction efficiency of binary and ternary DES was around 92%. Sas et al. synthesized various fatty acid-based (dodecanoic acid, decanoic acid, octanoic acid) hydrophobic DES by mixing with menthol or thymol for the extraction of phenolic compounds such as phenol, chlorophenol, and o-cresol from aqueous solution. These hydrophobic DES have a high separation efficiency (about 85%) for all phenolic compounds from the aqueous phase. The extraction efficiency for phenolic compounds followed the order: 2-chlorophenol > o-cresol > phenol.

Adeyemi et al. prepared and characterized seven different types of hydrophobic DES for the extraction of phenolic compounds, such as 3-chlorophenol, 2-chlorophenol, and 2,4-dichlorophenol, from the aqueous phase. Various DES based on menthol with hexanoic, octanoic, and decanoic acid were prepared by mixing in 1 : 2 ratio, and menthol–thymol in 1 : 1, 1 : 2, 1 : 3, 1 : 4 ratio. All these DES were found to be efficient for the extraction of chlorophenol and about 94% phenol was removed from the aqueous media under optimized conditions. The extraction efficiency of these DES followed the order 3-chlorophenol, 2-chlorophenol, and 2,4-dichlorophenol. In another study, Yang et al. prepared binary and ternary DES by mixing two and three carboxylic acids having different chain lengths ranging from C₈ to C₁₂. Various types of DES were synthesized by using different types of carboxylic acids and by adjusting their various molar ratios. The ternary DESs containing C₈ : C₁₀ : C₁₂ mixed in 3 : 2 : 1 molar ratio had a high phenolic compounds removal efficiency of more than 91%. It was found that this DES is very effective, and its efficiency is not influenced by the volume of the sample (>1000 mL). The synthesized DES was very efficient in the removal of the phenolic compound from a large volume of water which indicates their suitability from an industrial point of view.

Very recently, hydrophobic DESs were synthesized using methyltrierylammonium chloride (N₃₈₈₁Cl), tetrabutylammonium chloride (C₄₄₄₄Cl), and menthol as HBA, and octanoic, decanoic, and dodecanoic as HBD. These HDESs were used for the extraction of phenolic compounds from synthetic winery wastewater. Among the tested DES, ammonium-based DESs such as N₃₈₈₁Cl-menthol and N₃₈₈₁Cl-octanoic acid have the highest efficiency for extraction of the phenolic compound from winery wastewater.

Lawal et al. used multiwall carbon nanotubes (CNT) modified with DES for the extraction of phenol from water. The DES was prepared by the combination of methyltriphenyolphosphonium bromide and glycerol. The synthesized DES was used for the surface modification of CNT. The DES modified CNT was characterized using FTIR, SEM, and XRD techniques. The phenol adsorption capacity was 298 mg g⁻¹, which is much higher than that of pristine CNT, 128.6 mg g⁻¹.

Gu et al. used liquid-phase microextraction for phenolic compounds extraction from model oil method using DES in the presence of ultrasonic waves to shorten the extraction time. DESs were synthesized from choline chloride and γ-naphthaleneacetic acid with urea and ethylene glycol. All the DES demonstrated good potential for the extraction of phenol and o-cresol from model oil. Gu et al. synthesized DES from various types of ammonium salts for extraction of the phenolic compound from model oil. These DES were synthesized from different types of quaternary ammonium salts such as tetramethylammonium chloride, tetraethylammonium bromide, tetraethylammonium chloride, tetraethylammonium bromide, tetraethylammonium bromide, tetrapropylammonium bromide, tetrapropylammonium bromide.
chloride, tetrabutylammonium chloride, methyl-tetraethylammonium chloride, choline chloride, choline bromide, ammonium chloride. The DESs based on the ammonium salt have higher efficiency for phenol which reached as high as 99.9%. The phenol removal efficiency was not affected by temperature. These DES are easily recycled and reused without a reduction in separation efficiency. Pang et al. synthesized DES from ammonium salts, especially choline chloride, were found to effectively remove phenolic compounds such as phenol and cresol from model oil. Efficient removal of phenolic compounds was achieved in a short time and the removal efficiency of phenolic compounds was not sensitive to temperature. The choline chloride can be easily recycled and reused without weight loss. Yi et al. used a DES for the extraction of phenolic compounds from coal-based liquids oils. The DES was synthesized from choline chloride and glycerol and their interaction with model oil was investigated. The effect of the composition of DES, temperature, and DES amount were investigated on the extraction efficiency of phenol. The separation of phenolic compounds was found to take place via hydrogen bonding with DES. The extraction efficiency of phenol from oil was 98.3% under optimized experimental conditions. Yao et al. used quaternary ammonium-based zwitterions (L-carnitine and betaine) to form DES for removing phenol from model oil. The effect of process parameters such as temperature, extraction time mole ratio of zwitterions to phenol, and initial phenol concentration of phenol on removing efficiency of phenol from model oils were studied and optimized to achieve maximum phenol removal of 94.6%. The DES was also found to be insoluble in the model oil. Very recently, the ammonium salt-containing hydroxyl group (triallyl-2,3-dihydroxypropylammonium chloride) has been used as a liquid eutectic forming salt to extract phenol from toluene. Incorporation of alkyl groups and the dihydroxypropyl group to the ammonium cation was found to enhance the extraction efficiency of phenol and also inhibit the miscibility of DESs and toluene. The extraction efficiency of these functionalized DES were greater than choline chloride based DESs for phenol. Yi et al. designed DES based on choline chloride and glycerol (1 : 1) for dephenolization of model oil. Choline based DES have been shown to have high efficiency for extraction of phenolic compounds from coal-based liquids phenol. The higher efficiency of choline based DES is due to the presence of the hydroxyl group, the short alkyl chain, and the small central cation atom. Three imidazolium-based diatomic ILs were synthesized and used for the formation of DES for the separation of phenolic compound from an oil mixture. 1,2-Bis[N-(N-methylimidazolium)]ethane dibromide (DIL1), 1,3-bis[N-(N-methylimidazolium)]propane dibromide (DIL2), 1,4-bis[N-(N-methylimidazolium)]butane dibromide (DIL3) were the ILs synthesized for phenolic compound extraction from oil mixtures. All these ILs have high efficiency (96.6%) for extraction of phenolic compounds and the order of phenol extraction was found to be DIL1 < DIL2 < DIL3. These dicaticon ILs are thermally stable and can be used over a broad temperature range. These ILs were recycled and reused without any change in their structure after four cycle. DIL3 was also used for extraction of phenolic compounds from coal tar with a removal efficiency of 93.1%.

15. Conclusions

In comparison to traditional solvents, ILs are alternative solvents known as greener solvents because of their negligible vapor pressure and reduced impact on the environment and human health. The structural characteristics of ILs such as anion type, cation type, alkyl chain length, and functional groups play a key role in their performance as phenol extractants. ILs having hydrophobic anions and aromatic cations are more efficient for phenol extraction. As far as the alkyl chain length on cation is concerned, it has been observed that increasing the hydrophobicity of ILs by increasing alkyl chain length has no considerable effect on the extraction of phenolic compounds. A wide range of ILs with desirable properties can be synthesized by selecting proper anions and cations which may offer great potential for the extraction of phenolic compounds. Moreover, the experimental parameters such as initial pH of the phenol solution, phase ratio (Vn : Vo), phenol concentration, and temperature each play a vital role in phenol extraction; therefore, their optimization is necessary to achieve maximum extraction of phenol. Phenol extraction efficiency increasing with increase in pH of phenol solution from acidic to basic, contact time and initial phenol concentration. By increasing the phase ratio, a decrease in phenol extraction efficiency was observed. Solid supported ILs are consider better for phenol extraction from aqueous and model oils due to their ease recyclability and reusability. As a result, ILs technology provides an opportunity to develop novel and improved methods for phenol removal from wastewater. In this context, task specific ILs will play a crucial role in wastewater processing technology. DESs in their pristine form, or as surface modifiers for nano-materials, were also explored for their applications in the treatment of fluidic waste either. These solvents were used successfully for the treatment of aqueous and non-aqueous waste fluids contaminated with phenolic compounds. These solvents have great industrial potential as alternatives for ILs.

Finally, it is hoped that this paper may provide valuable and useful data for researchers and industrialists working to develop novel cleaner processes using ionic liquids. It is anticipated also that these novel cleaner processes for wastewater treatment containing phenolic compounds will enhance both the sustainability and innovation in the relevant industries.

Abbreviations

| Abbreviation | Description |
|--------------|-------------|
| EPA | Environmental Protection Agency |
| ppb | Part per billion |
| ppm | Part per million |
| ILs | Ionic liquids |
| ΔH_{vap} | Heat of vaporization |
| CO_{2} | Carbon dioxide |
| H_{2}S | Hydrogen sulfide |
| SO_{2} | Sulphur dioxide |
| Ion | Chemical Formula | Description |
|-----|-----------------|-------------|
| BF<sup>−</sup> | Tetrafluoroborate | |
| PF<sub>6</sub> | Hexafluorophosphate | |
| Cl<sup>−</sup> | Chloride | |
| Br<sup>−</sup> | Bromide | |
| I<sup>−</sup> | Iodide | |
| HCOO<sup>−</sup> | Formate | |
| CF<sub>3</sub>COO<sup>−</sup> | Trifluoroacetate | |
| CH<sub>3</sub>COO<sup>−</sup> | Acetate | |
| HSO<sub>4</sub> | Hydrogen sulphate | |
| CF<sub>3</sub>SO<sub>3</sub> | Trifluoromethanesulfonate | |
| CH<sub>3</sub>SO<sub>3</sub> | Methanesulfonate | |
| NT<sub>2</sub>F<sub>−</sub> | Bis(trifluoromethanesulfonyl)amide | |
| NO<sub>3</sub> | Nitrate | |
| AlCl<sub>4</sub> | Tetraethyl ammonium | |
| SCN<sup>−</sup> | Thiocyanate | |
| H<sub>2</sub>PO<sub>4</sub> | Dihydrogen phosphate | |
| CF<sub>3</sub>SO<sub>3</sub> | Trifluoromethanesulfonate | |
| H<sub>2</sub>SO<sub>4</sub> | Sulfuric acid | |
| Ac| | Acetate |
| Pro| | Propionate |
| [C<sub>4</sub>mim]<sup>+</sup>| 1-Methyl-3-propylimidazolium | |
| [NT<sub>2</sub>]<sup>−</sup>| bis(trifluoromethanesulfonyl)imide | |
| [C<sub>6</sub>mPyr]<sup>+</sup>| 1-Methyl-3-hexylpyridinium | |
| [NT<sub>2</sub>]<sup>−</sup>| methylimidazoliambis(trifluoromethanesulfonyl)imide | |
| [C<sub>4</sub>mim]<sup>+</sup>| 1-Hexyl-3-methylimidazolium | |
| [NT<sub>2</sub>]<sup>−</sup>| hexafluorophosphosphate | |
| [C<sub>4</sub>mim]<sup>+</sup>[BF<sub>4</sub>]<sup>−</sup> | 1-Methyl-3-hexylimidazolium tetrafluoroborate | |
| [C<sub>4</sub>mim]<sup>+</sup>[PF<sub>6</sub>]<sup>−</sup> | 1-Methyl-3-hexylimidazolium hexafluorophosphate | |
| [C<sub>4</sub>mim]<sup>+</sup>[BF<sub>4</sub>]<sup>−</sup> | 1-Methyl-3-octylimidazolium tetrafluoroborate | |
| [C<sub>4</sub>mim]<sup>+</sup>[PF<sub>6</sub>]<sup>−</sup> | 1-Methyl-3-octylimidazolium hexafluorophosphate | |
| [N<sub>2222</sub>]<sup>+</sup>| Tetraethyl ammonium | |
| [N<sub>2222</sub>]<sup>+</sup>-[1-Pro]<sup>−</sup> | Tetraethyl ammonium prolinate | |
| [N<sub>2222</sub>]<sup>+</sup>-[Ser]<sup>−</sup> | Tetraethyl ammonium serine | |
| [N<sub>2222</sub>]<sup>+</sup>-[Gly]<sup>−</sup> | Tetraethyl ammonium glycine | |
| [N<sub>2222</sub>]<sup>+</sup>-[Ala]<sup>−</sup> | Tetraethyl ammonium alanine | |
| [N<sub>2222</sub>]<sup>+</sup>-[Lys]<sup>−</sup> | Tetraethyl ammonium lysine | |
| [C<sub>4</sub>C<sub>1</sub>Py]<sup>−</sup> | 3-Methyl-1-butyldiethylinidium | |
| MEA | Monoethanol amines | |
| DEA | Diethanol amines | |
| TEA | Triethanol amines | |
| C<sub>4</sub>C<sub>1</sub>Py<sup>+</sup> | Butyldiethylinidium | |
| C<sub>6</sub>C<sub>1</sub>Py<sup>+</sup> | 1-Methyl-3-butyldiethylinidium | |
| C<sub>6</sub>C<sub>1</sub>Pip<sup>+</sup> | 1-Methyl-3-butyldiethylpyrrolidinium | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-hexylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>4</sub>mim]<sup>+</sup> | 1-Methyl-3-hexylimidazolium bis(trifluoromethanesulfonyl)amide | |
| THADHSS | Tetrahexylammonium dihexylsulfosucinate | |
| TOMAS | Trioctylmethylammonium salicylate | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-octylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| | 1-Methyl-3-dodecylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-octylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
| [C<sub>6</sub>mPyr]<sup>+</sup> | 1-Methyl-3-heptylimidazolium bis(trifluoromethanesulfonyl)amide | |
ΔG  Change in Gibbs free energy
ΔH  Change in enthalpy
ΔS  Change in entropy
R  Universal gas constant
[Emim][Lac]  1-Ethyl-3-methyl imidazolium lactate
[PA][FA]  Propylamine formate
[PA][Ac]  Propylamine acetate
MG  1,1,3,3-Tetramethylguanidine acid
[TMG][BF₄]  1,1,3,3-Tetramethylguanidine tetrafluoroborate
TAAA  Tetraethylammonium amino acid
GO  Graphene oxide
LDH  Layered double hydroxide
ELM  Emulsified liquid membrane
DESS  Deep eutectic solvents
HBD  Hydrogen bond donor
HBA  Hydrogen bond acceptor
CNT  Carbon nanotubes

Conflicts of interest
We wish to confirm that no conflict of interest associated with this publication.

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