One-Pot Synthesis and Applications of N-Heteroaryl Iodonium Salts

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An efficient one-pot synthesis of N-heteroaryl iodonium triflates from the corresponding N-heteroaryl iodide and arene has been developed. The reaction conditions resemble our previous one-pot syntheses, with suitable modifications to allow N-heteroaryl groups. The reaction time is only 30 min, and no anion exchange is required. The obtained iodonium salts were isolated in a protonated form, these salts can either be employed directly in applications or be deprotonated prior to use. The aryl groups were chosen to induce chemoselective transfer of the heteroaryl moiety to various nucleophiles. The reactivity and chemoselectivity of these iodonium salts were demonstrated by selectively introducing a pyridyl moiety onto both oxygen and carbon nucleophiles in good yields.

Diaryliodonium salts have recently been recognized as efficient electrophilic arylation reagents with a wide range of nucleophiles under both metal-free and metal-catalyzed conditions.[1] The accessibility of these reagents has been greatly simplified by the recent development of efficient one-pot syntheses from the corresponding iodoarenes or iodine with arenes (Scheme 1 A,B).[2] A complimentary regiospecific route employing arylboronic acids gives access to diaryliodonium salts with all kinds of substitution patterns (Scheme 1 C).[3]

Despite the wide scope of the one-pot methods described above, they fail in the synthesis of N-heterocyclic iodonium salts.[4] N-heterocycles are common organometallic ligands and substructures in biologically active compounds,[5] and the introduction of such a moiety from a diaryliodonium salt would be of high utility in organic synthesis. Symmetric N-heterocyclic salts can be obtained in low yields by treatment of a highly unstable vinyliodonium dichloride[6] with aryl lithium reagents at −78 °C.[7] Unsymmetric salts can be obtained in a similar way.[8] A four-step synthesis to 3-pyridyl(aryl)iodonium salts via the corresponding 3-pyridyl-iodoiodochloride has also been reported.[9] The poor accessibility to N-heterocyclic diaryliodonium salts inspired us to develop a general one-pot synthesis of these compounds. The synthesis and chemoselective application of these salts to different nucleophiles are reported herein.

The reaction of 3-iodopyridine (1a) with benzene (2a) to give pyridyl iodonium salt 3a was chosen as a model system (Scheme 2). Initial attempts to form 3a under our standard conditions, 1.1 equiv meta-chloroperbenzoic acid (mCPBA) and 3 equiv trifluoromethanesulfonic acid (TfOH),[2b] resulted in mixtures due to competing N-oxidation of 1a.[10] This could be prevented by treating 1a with TfOH before addition of mCPBA, thus protecting the nitrogen from oxidation. The nitrogen remained protonated in the formed iodonium salt, and the isolated product was found to be phenyl(3-pyridinium)iodonium bistriflate (3a') rather than 3a.

The reaction conditions were further optimized, as detailed in Table 1. The amount of mCPBA was first investigated, and a slight excess proved better than equimolar amounts (Entries 1–3). Increasing the amount of TfOH from three to four equivalents further improved the yield, which can be explained by one equivalent of acid being consumed for nitrogen oxidation (Entry 4).[10] This could be prevented by treating 1a with TfOH before addition of mCPBA, thus protecting the nitrogen from oxidation. The nitrogen remained protonated in the formed iodonium salt, and the isolated product was found to be phenyl(3-pyridinium)iodonium bistriflate (3a') rather than 3a.

The reaction time and temperature were varied, and 30 min at 60 °C was as efficient as 3 h at 80 °C, while lower temperatures were insufficient (Entries 4–8). The high temperature required could be explained by slow I-oxidation of protonated 1a.
It is known that pyridyl(phenyl)iodonium salts provide insufficient chemoselectivity in reactions with nucleophiles under metal-free conditions.\cite{9,10} Therefore, further optimization was performed with anisole (2b), which is known to be a useful “dummy group” in reactions of diaryliodonium salts with various nucleophiles.\cite{12}

Syntheses of diaryliodonium salts with electron-rich arenes generally require special conditions, such as a milder acid or oxidant, to avoid side reactions. We have previously reported modified procedures in our one-pot syntheses, where the iodoarene is oxidized to the iodine(III) intermediate before addition of the electron-rich arene at low temperature.\cite{2b,3}

Because oxidation of protonated 1a is slow, it was deemed unsuitable to change the oxidation conditions. Thus, 1a was oxidized under the same conditions, followed by addition of 2 equiv water at 0°C to quench the triflic acid and create a milder acid. The electron-rich arene in dichloromethane was added dropwise, and additional stirring for 10 min afforded the desired product 3b in high yield (Scheme 3).

Pyridinium bistriflate 3′ might behave similarly to pyridyl salts 3 in applications, but it was desirable to also have access to deprotonated salts 3. A deprotonation procedure was therefore developed to remove the triflic acid from 3′ to obtain pyridyl salts 3. Several basic workup procedures were attempted to isolate salt 3b, but unwanted anion exchanges complicated such deprotonations.\cite{10} Deprotonation was best achieved using a basic Al₂O₃ column eluted with dichloromethane/methanol (20:1). The eluted material was concentrated in vacuo to give the pure deprotonated product 3b in 97% yield. Unfortunately, submitting the crude reaction mixture directly onto an Al₂O₃ column gave back the protonated material, and isolation of 3′ before deprotonation was necessary.

The optimized synthesis of 3′ and deprotonation to 3 was subsequently applied on a range of N-heterocyclic iodoarenes 1 and arenes 2. The arenes were selected to give good chemoselectivity in both metal-free and metal-mediated reactions (Scheme 4). As mentioned above, the anisyl moiety is a good dummy group in many reactions under metal-free conditions.

| Table 1. Optimization of reaction conditions.\textsuperscript{[a]} |
|---|---|---|---|---|---|
| Entry | mCPBA (equiv) | TFOH (equiv) | T (°C) | t (min) | Yield (%) |
| 1 | 1.1 | 3.0 | 80 | 180 | 54 |
| 2 | 1.5 | 3.0 | 80 | 180 | 60 |
| 3 | 2.0 | 3.0 | 80 | 180 | 52 |
| 4 | 1.5 | 4.0 | 80 | 180 | 68 |
| 5 | 1.5 | 4.0 | 80 | 10 | 48 |
| 6 | 1.5 | 4.0 | 80 | 30 | 60 |
| 7 | 1.5 | 4.0 | 60 | 30 | 69 |
| 8 | 1.5 | 4.0 | 40 | 30 | \textsuperscript{[b]} |

\textsuperscript{[a]} 1.1 equiv 2a was used. The product was isolated by concentration of the crude mixture in vacuo followed by precipitation by addition of Et₂O. \textsuperscript{[b]} Product isolation was difficult.

Scheme 4. General chemoselectivity trends.

Scheme 5. Modified synthesis with electron-rich arenes.

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n) should be chemoselective in metal-catalyzed reactions and other reactions sensitive to steric hindrance, for example, malonate arylations.

A reliable analytical tool was required to ensure that complete deprotonation had taken place for all salts. While HRMS delivered the same molecular weight for both $3'$ and $3$, the $^1$H NMR spectra in deuterated methanol differed. A concentration effect with slightly changing shifts is seen in samples of $3'$, due to the interchangeable proton, while shifts of $3$ are constant. Salt $3e'$ was synthesized to confirm the accuracy of this analytical method. Apart from the differences in $^1$H NMR spectra described above, $^{19}$F NMR analysis of $3e'$ gave a ratio of 6:1 between the two detected peaks, indicating the presence of two triflate moieties. The same analysis after deprotonation to $3e$ indeed gave a ratio of 3:1.

Application studies were next undertaken to demonstrate the utility of the prepared salts in chemoselective arylations. Furthermore, the reactivities of the protonated salts $3'$ and the deprotonated salts $3$ were compared. The electron-rich salts $3b'$ and $3b$ were applied in the arylation of phenols, using our previously reported protocol. Arylation of $4a$ with $3b$ gave pyridyl ether $5a$ in 88% yield (Scheme 6). The sterically hindered phenol $4b$ was also arylated to give $5b$, albeit in slightly lower yield. Reactions with salt $3b'$ were performed with 2.1 equiv base instead of 1.1 equiv to compensate for the TfOH present. The yields with this salt were lower than with $3b$; but considering the facile isolation of $3b'$, it might still be a useful alternative in some applications. The reactivity of N-heteroaryl salts $3c$ and $3c'$ was investigated by arylation of diethyl methylmalonate ($6$), which has previously been chemoselectively arylated using a mesityl dummy group.

Again, the deprotonated salt was more efficient than the protonated salt in formation of $7$, despite use of 2 equiv base (Scheme 7). Complete chemoselectivity was obtained in all reactions, illustrating the usefulness of inexpensive anisole and mesitylene as dummy groups in these arylations.

An efficient synthesis of N-heteroaryl iodonium triflates has been developed. The aryl groups were chosen to induce chemoselective transfer of the heteroaryl moiety to different nucleophiles.

This was demonstrated by selectively introducing a pyridyl moiety onto both oxygen and carbon nucleophiles. The protonated heteroaryl salts $3'$ give lower yields than the deprotonated salts $3$ in these reactions but might be as efficient in transformations that are not base mediated.

Scheme 5. Synthesized N-heteroaryl iodonium salts. Method A was used unless specified (see Table 1, Entry 7). Yield of $3$ is from isolated $3'$. [a] Method B was used (see Scheme 3).

Scheme 6. Chemoselective arylation of phenols.

Scheme 7. Chemoselective arylation of malonate $6$. This was demonstrated by selectively introducing a pyridyl moiety onto both oxygen and carbon nucleophiles. The protonated heteroaryl salts $3'$ give lower yields than the deprotonated salts $3$ in these reactions but might be as efficient in transformations that are not base mediated.
Experimental Section

General experimental procedures

Method A: Iodoarene 1 (0.24 mmol) was stirred with TFOH (4 equiv) in CH₂Cl₂ (1 mL) for 5 min at RT. Arene 2 (1.1 equiv) and mCPBA (1.5 equiv) were added, and the solution was stirred at 60 °C for 30 min.

Method B: Iodoarene 1 (1.51 mmol) was stirred with TFOH (4 equiv) in CH₂Cl₂ (10 mL) for 5 min at RT. mCPBA (1.75 equiv) was added, and the solution was stirred at 60 °C for 30 min. After H₂O (2 equiv) was subsequently added at 0 °C and then arene 2 (1.1 equiv), the solution was stirred at 0 °C for 10 min.

Isolation of 3: The crude mixture was concentrated in vacuo followed by precipitation by addition of Et₂O (1–3 mL). The solid was collected by filtration and washed with Et₂O.

Deprotonation to 3: Isolated 3' was added on a basic Al₂O₃ column and eluted with CH₂Cl₂/MeOH (20:1). The eluted mixture was concentrated in vacuo to give the pure deprotonated product 3.

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