Combined pulsed laser deposition and non-contact atomic force microscopy system for studies of insulator metal oxide thin films

Daiki Katsube¹, Hayato Yamashita¹,², Satoshi Abo¹ and Masayuki Abe*¹

Abstract
We have designed and developed a combined system of pulsed laser deposition (PLD) and non-contact atomic force microscopy (NC-AFM) for observations of insulator metal oxide surfaces. With this system, the long-period iterations of sputtering and annealing used in conventional methods for preparing a metal oxide film surface are not required. The performance of the combined system is demonstrated for the preparation and high-resolution NC-AFM imaging of atomically flat thin films of anatase TiO₂(001) and LaAlO₃(100).

Introduction
The surfaces of metal oxides have attractive and useful properties, such as superconductivity, ferroelectricity and catalytic activity [1]. These properties are derived from the unique structure, the electron and spin interactions, the degree of freedom of the electron orbit and the lattice constant of the metal oxides. To understand and apply these characteristics, various studies have been conducted thus far, including investigations of physical properties by growth of ultra-high-quality metal oxide thin films [2,3], discovery of novel physical properties by fabricating heterointerfaces using different metal oxides [4,5], and expression of conductivity and magnetism by doping metal oxides with specific elements [6,7]. In studies of metal oxide surfaces at the atomic scale, developments of observation techniques have been advancing rapidly, especially in the field of transmission electron microscopy [5,8-13]. As atomic resolution methods, scanning probe microscopy including scanning tunneling microscopy (STM) [13-30] and non-contact atomic force microscopy (NC-AFM) [19,23,29,31-47] have played important roles. NC-AFM in particular can be used to elucidate the structures of surfaces at the atomic scale. An NC-AFM
measures the shift in cantilever resonance due to the interaction force between the tip and the sample, hence it is possible to use it to observe any surface regardless of the conductivity of the material [19,23,31-47]. STM on the other hand can only be used for observations of insulators with a thickness of few atomic layers on conductive substrates [20-22].

For observations of surface atoms of metal oxides using NC-AFM and STM, it is critical to prepare atomically flat and clean surfaces. A standard method for obtaining clean surfaces of metal oxides is performing iterations of Ar’ sputtering and annealing at high temperatures [23-26,37,41,42]. Different conditions of annealing and sputtering result in different types of surface reconstruction even under the same preparation conditions [26]. As an example of an insulator metal oxide, an Al₂O₃(0001) surface prepared using iterations of annealing and sputtering was imaged with atomic resolution in recent studies [35,37]. Difficulties with this method are a long preparation time and a low reproducibility. Recently, connecting an STM with a pulsed laser deposition (PLD) system [13,27,28] or with molecular beam epitaxy [18] has increased the types of measurements possible. Such combined systems allow for in situ observations, from sample preparation to measurements. In order to image surface atoms of insulator metal oxides with atomic resolution, in this study, we have developed a combined system consisting of an NC-AFM and PLD operated in ultra-high vacuum (UHV) at room temperature. In situ sample preparation and measurements have made it possible to image surfaces of insulator metal oxides under conditions where the sample does not become contaminated. Furthermore, the sample preparation time is significantly reduced. In the “Experimental” section, the configuration of the combined NC-AFM/PLD system and the heating method for annealing insulator samples are described. Then, in the section “Results and Discussion”, we introduce a method for finding the appropriate conditions for obtaining atomically flat surfaces, which is critical for obtaining atomic resolution images stably with NC-AFM. Finally, we demonstrate atomic resolution NC-AFM imaging of anatase TiO₂(001) and LaAlO₃(100). Both materials are important in the field of materials science, and it has been challenging to form and image atomically flat and clean surfaces of these two oxides thus far.

Experimental

Figure 1 shows a schematic diagram of the combined NC-AFM and PLD system that was developed. It includes UHV chambers for sample preparation, low-energy electron diffraction (LEED), and load locking. It is also designed to allow for adding surface analysis capabilities such as reflection high-energy electron diffraction (RHEED). Sample preparation and NC-AFM, STM and LEED measurements can be performed while keeping the sample under UHV. The PLD chamber was designed to be as small as possible for easy sample exchange and transfer. In addition, a small design would lead to smaller vibrations during AFM and/or STM measurements and, thus, better observations. The PLD chamber the side of which is fitted with an CF203 port includes a PLD target, a shutter, a sample stocker, and a RHEED apparatus. A small size commercial PLD target system (UNISOKU) is installed in the PLD chamber. Piezoelectric motors are used to rotate the PLD targets. The pulse laser (COHERENT COMPex Pro; KrF, λ = 248 nm). The laser enters the chamber diagonally from above, so the sample surface faces downward during PLD sample preparation. In order to prevent the sample from falling, both the sample and the holder are placed on the stocker using magnets, and the sample and holder do not fall even when the temperature is over 1000 °C.

During sample preparation with PLD, the sample is heated in an O₂ atmosphere. If the sample is a doped metal oxide, it can be heated by flowing a current through the sample as done in previous STM and NC-AFM studies. To anneal insulator samples, a tungsten thin film is sandwiched between the sample and a metal oxide substrate [35,37]. The sample is then heated by flowing current through the tungsten film. The configuration of the sample and the sample holder used are shown in Figure 2a [43]. The substrate is fixed with four screws and leaf springs on the ceramic part of the sample holder. Stainless steel electrodes under the ceramic are used as connections for flowing a current through the tungsten thin film.

The electrode for the flowing current is insulated from the base of the stocker by insulator materials. This prevents a short-circuit of the flowing current to the UHV chambers. The U-shaped base plate of the sample holder stocker is 1 mm thick to reduce heat diffusion during heating of the substrate. The chambers of the NC-AFM, the sample preparation and the load lock are the same as in our previous studies [29,37,41,44-48]. To perform stable atomic resolution imaging in this chamber, a mechanism to fix the unit by double spring vibration isolation and eddy current damping [49] is provided, to prevent vibration noise. Therefore, it is possible to obtain atomic resolution images with the same performance as in our previous study [48]. We are able to perform NC-AFM measurements without the influence of vibration noise caused by the linkage to the PLD chamber. An optical interferometer [50] was used for deflection detection of the probe used for the NC-AFM observations. To image the insulator metal oxide thin films with atomic resolution, NC-AFM with the frequency modulation mode is used [51]. A commercial cantilever (Budget Sensors, TAP190) is used to obtain NC-AFM topographic images. An Ar’ sputtering gun is installed in the preparation chamber to
Figure 1: (a) Top view and (b) side view of the combined PLD/AFM system. STM measurements can also be performed. LEED and RHEED apparatuses are also installed. The AFM (and also STM) is operated at room temperature in UHV.

Figure 2: (a) Sample holder and (b) holder stocker used in the PLD camber.
Results and Discussion
Preparation of atomically flat surfaces with PLD
For atomic scale observations of the metal oxide surface, it is necessary to prepare a clean and flat surface at the atomic level over a wide area. This is to avoid changes in the tip apex due to the presence of protrusions on the surface. Although there have been studies in which the growth of atomically flat metal oxides has been demonstrated [3,13,27,28,52-55], it is still difficult to find the best sample preparation conditions for PLD for atomic resolution imaging with NC-AFM and STM. Even if a clean diffraction pattern can be seen with RHEED performed during the PLD, it might be that the surface flatness necessary for the NC-AFM observation is not obtained. Parameters such as the sample temperature, oxygen partial pressure, and laser intensity during PLD need to be optimized for each sample. Since our UHV NC-AFM is optimized for atomic resolution measurements, scanning over a wide range cannot be performed. To find suitable conditions for efficient PLD deposition, a deposited sample was taken out into the ambient atmosphere and its topography was measured with a tapping AFM in air. Although it is difficult to obtain atomic resolution images in tapping mode, the larger scanning area of tapping AFM enabled us to check whether a step-and-terrace surface was formed over a wide range. Furthermore, the spatial resolution in the vertical direction of the tapping AFM is high, so images of a single-atomic step-and-terrace structure for confirming the flatness of the surface can be obtained. Evaluation using X-ray diffraction (XRD) in ambient air is also effective for checking the crystal structure, orientation plane, and crystallinity of a thin film. Once a single-atomic layer step is confirmed, a new sample is prepared under the same PLD conditions and in situ NC-AFM (and also STM) measurements can then be performed in UHV.

As an example, a process flow for determining the sample preparation conditions of anatase TiO$_2$(001) is shown in Figure 3 [56]. Referring to previous studies [52,53], we started the PLD with the following sample parameters: temperature $T_s = 700$ °C, oxygen partial pressure $P_O \approx 1 \times 10^{-3}$ Pa, laser density $I = 1.0$ J/cm$^2$, and laser repetition rate $f_p = 2.0$ Hz on a LaAlO$_3$(100) substrate. The sample was then taken out of vacuum, and its surface was measured with tapping mode AFM in air as shown in Figure 3a. It was found that an atomically flat area did not exist on the surface, instead grain structures appeared on the surface. A peak for anatase TiO$_2$(001) in the XRD pattern Figure 3c indicates that each small grain should contain anatase TiO$_2$ crystals. Since excessive oxygen could cause the grain structure observed, the sample was prepared after reducing the oxygen partial pressure to $P_O \approx 1 \times 10^{-4}$ Pa as shown in Figure 3b. In this case, many small islands including single- and multi-atomic steps were observed. In order to flatten the islands of the multi-atomic step structure, the sample temperature during the PLD was increased to $T_s = 800$ °C (Figure 3c). Increasing the annealing temperature then formed a flat area. Although anatase TiO$_2$(001) is said to undergo a structural change to rutile TiO$_2$(001) at around 750 °C, the XRD pattern showed that the anatase phase remained in this case. This is due to the lattice constant of the substrate. Considering that a phase transition from anatase to rutile may occur if the temperature is raised beyond 800 °C, we then changed the laser power density to a lower value ($I = 0.8$ J/cm$^2$) to reduce the deposition rate (Figure 3d). Since we obtained a step-and-

![Figure 3](image-url)

(a) $T_s \approx 2 \times 10^{-4}$ Pa, $I = 1.0$ J/cm$^2$
(b) $T_s \approx 2 \times 10^{-4}$ Pa, $I = 0.8$ J/cm$^2$
(c) $T_s \approx 1 \times 10^{-4}$ Pa, $I = 0.8$ J/cm$^2$
(d) $T_s \approx 1 \times 10^{-4}$ Pa, $I = 1.0$ J/cm$^2$

(e) XRD patterns for the samples shown in (a–d). The samples were prepared under different conditions of annealing temperature, oxygen partial pressure, and laser power density as shown in the AFM images. The pulsed laser frequency was $f_p = 2$ Hz. The samples were washed with acetone and ultrapure water before performing PLD. AFM images are processed using the WSxM software [57].
terrace surface, as shown in Figure 3d, we decided to use these conditions for the NC-AFM measurements in UHV. For other metal oxides, sample preparation at the atomic level is possible in the same way.

Atomic resolution imaging of insulator metal oxide films with NC-AFM

Atomic resolution imaging with NC-AFM of insulator thin films of anatase TiO$_2$(001) and LaAlO$_3$(100) prepared with PLD was carried out. Both materials are important metal oxides, and studies at the atomic level of both materials are still ongoing because of their high importance in materials science. Anatase TiO$_2$ is well known as a photocatalyst, but there have been few studies using NC-AFM due to the difficulty in preparing samples [58]. Thin films of LaAlO$_3$ epitaxially grown on substrate materials have been studied for correlated electron heterostructures and devices. One of the most important and common uses of epitaxial LaAlO$_3$ is its interface with SrTiO$_3$ for studies of electrical conductivity [4], superconductivity [59], photoconductivity [60], and magnetoresistance [61]. Figure 4 shows results of atomic-resolution measurements of anatase TiO$_2$(001) and LaAlO$_3$(100) using the NC-AFM. We determined the PLD conditions according to the procedure mentioned above, described in the caption of Figure 4. LEED measurements were also performed in situ after the NC-AFM measurements. The results showed that the both surfaces had structures with a period of four times the lattice constant (data not shown).

Figure 4a,b shows an atomic resolution image and a line profile of the anatase TiO$_2$(001), respectively. From the NC-AFM image and the line profile in Figure 4b, the periodic width of the column was 1.5 nm, which is approximately four times the lattice constant (0.3785 nm). This means that, under these preparation conditions, the anatase TiO$_2$(001) surface has been reconstructed to the $(1 \times 4)$ or $(4 \times 1)$ structure. This result is consistent with the results of LEED and the previous study [23]. Similarly, to the case of anatase TiO$_2$(001), the LaAlO$_3$(100) surface also has a periodic structure of 1.5 nm. Since the lattice constant of LaAlO$_3$(100) is 0.379 nm, the surface was also reconstructed with $(1 \times 4)$ or $(4 \times 1)$. Details of the NC-AFM image contrast of both surfaces will be described elsewhere.

As shown by the measurement results in Figure 4, by connecting PLD with NC-AFM, we succeeded in atomic-resolution observations of two insulator surfaces that have not been previously measured due to the difficulty of sample preparation. One of the critical points for obtaining atomic resolution images
is the sample preparation method described in the first part of section “Results and Discussion”. It has become possible to efficiently find the appropriate conditions for sample preparation with this method. The possibility of studies on surfaces of insulators has thus been greatly expanded by the development of this method. Since it is no longer necessary to repeat the sputtering and annealing that typically takes several days for the preparation of samples, the efficiency of experiments using metal oxides described in this work has dramatically increased. Also, reproducibility of the surface quality can be expected without the sputtering and annealing. Of course, one can perform experiments of the same surface without the combined system. Some insulator thin films such as TiO$_2$ made by PLD and left under ambient conditions can be cleaned by an iteration of sputtering and annealing for imaging with atomic resolution. We confirmed that this is the case for the anatase TiO$_2$(001) thin film, but not for LaAlO$_3$(100). However, even if one can obtain atomic-resolution images, the influence on the surface by, e.g., OH groups, cannot be completely eliminated. Furthermore, the contact potential difference (CPD) of the insulator surface synthesized with our PLD system is at most about 3 V. With our system, it is possible to measure without affecting the charging of the sample surface, which often occurs on insulator surfaces.

Conclusion

We have developed a combined system of NC-AFM and PLD for atomic-resolution measurements of insulator thin films. The system enabled preparing atomically flat surfaces with high efficiency without the usual iteration of sputtering and annealing. We performed in situ NC-AFM measurements on anatase TiO$_2$(001) and LaAlO$_3$(100) prepared by PLD and obtained images with atomic resolution. It is expected that atomic-scale studies of the surfaces of insulators that cannot be measured with STM can be carried out using the current system.

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References

1. Heinrich, V.; Cox, P. The Surface Science of Metal Oxides; Cambridge University Press, 1994.
2. Ohta, H.; Oriita, M.; Hirano, M.; Tanji, H.; Kawazoe, H.; Hosono, H. Appl. Phys. Lett. 2000, 76, 2740. doi:10.1063/1.126481
3. Ohta, H.; Oriita, M.; Hirano, M.; Hosono, H. J. Appl. Phys. 2002, 91, 3547. doi:10.1063/1.1448873
4. Ohtomo, A.; Hwang, H. Y. Nature 2004, 427, 423–426. doi:10.1038/nature02308
5. Gozar, A.; Logvenov, G.; Kourkoutis, L. F.; Bollinger, A. T.; Giannuzzi, L. A.; Muller, D. A.; Bozovic, I. Nature 2008, 455, 782–785. doi:10.1038/nature07293
6. Matsumoto, Y.; Murakami, M.; Shono, T.; Hasegawa, T.; Fukumura, T.; Kawasaki, M.; Ahmet, P.; Chikyow, T.; Koshihara, S.; Koinuma, H. Science 2001, 291, 854–856. doi:10.1126/science.1056186
7. Furubayashi, Y.; Hitosugi, T.; Yamamoto, Y.; Inaba, K.; Kinoda, G.; Hirose, Y.; Shimada, T.; Hasegawa, T. Appl. Phys. Lett. 2005, 86, 252101. doi:10.1063/1.1949728
8. Erdman, N.; Poeppelmeier, K. R.; Asta, M.; Warschkow, O.; Ellis, D. E.; Marks, L. D. Nature 2002, 419, 55–58. doi:10.1038/nature01010
9. Lanier, C. H.; Rondinelli, J. M.; Deng, B.; Kilaas, R.; Poeppelmeier, K. R.; Marks, L. D. Phys. Rev. Lett. 2007, 98, 086102. doi:10.1103/PhysRevLett.98.086102
10. Erdman, N.; Warschkow, O.; Asta, M.; Poeppelmeier, K. R.; Ellis, D. E.; Marks, L. D. J. Am. Chem. Soc. 2003, 125, 10050–10056. doi:10.1021/ja034933h
11. Yuan, W.; Wu, H.; Li, H.; Dai, Z.; Zhang, Z.; Sun, C.; Wang, Y. Chem. Mater. 2017, 29, 3189–3194. doi:10.1021/acs.chemmater.7b00284
12. Sugiyama, I.; Shibata, N.; Wang, Z.; Kobayashi, S.; Yamamoto, K.; Ikuhara, Y. Nat. Nanotechnol. 2013, 8, 266–270. doi:10.1038/nnano.2013.45
13. Ohsawa, T.; Saito, M.; Hamada, I.; Shimizu, R.; Iwaya, K.; Shiraki, S.; Wang, Z.; Ikuhara, Y.; Hitosugi, T. ACS Nano 2015, 9, 8766–8772. doi:10.1021/acsnano.5b02867
14. Koguchi, K.; Matsumoto, T.; Kawai, T. Science 1995, 267, 71–73. doi:10.1126/science.267.5194.71
15. Morales, E. H.; Diebold, U. Appl. Phys. Lett. 2009, 95, 253105. doi:10.1063/1.3275716
16. Tao, J.; Luttrell, T.; Batzill, M. Nat. Chem. 2011, 3, 296–300. doi:10.1038/nchem.1006
17. Gong, X.; Selloni, A.; Batzill, M.; Diebold, U. Nat. Mater. 2006, 5, 665–670. doi:10.1038/nmat1695
18. Ma, J. X.; Gillaspie, D. T.; Plummer, E. W.; Shen, J. Phys. Rev. Lett. 2005, 95, 237210. doi:10.1103/PhysRevLett.95.237210
19. Kubo, T.; Nozoye, H. Phys. Rev. Lett. 2001, 88, 1801. doi:10.1103/PhysRevLett.88.1801
20. Kresse, G.; Schmid, M.; Napechnog; E.; Shishkin, M.; Köhler, L.; Varga, P. Science 2005, 308, 1440–1442. doi:10.1126/science.1107753
21. Schinke, S.; Messerli, S.; Pivetta, M.; Patthey, F.; Liboule, L.; Stengel, M.; De Vita, A.; Schneider, W.-D. Phys. Rev. Lett. 2001, 87, 276801. doi:10.1103/PhysRevLett.87.276801
22. Glückler, K.; Sokolowski, M.; Soukopp, A.; Umbach, E. Phys. Rev. B 1996, 54, 7705. doi:10.1103/PhysRevB.54.7705
23. Tanner, R. E.; Sasahara, A.; Liang, Y.; Altman, E. I.; Onishi, H. J. Phys. Chem. B 2002, 106, 8211–8222. doi:10.1021/jp020617y
24. Cronnie, M. F.; Lutz, C. P.; Eigler, D. M. Nature 1993, 363, 524–527. doi:10.1038/363524a0
25. Hasegawa, Y.; Avouris, P. Phys. Rev. Lett. 1993, 71, 1071. doi:10.1103/PhysRevLett.71.1071
26. Gerhold, S.; Wang, Z.; Schmid, M.; Diebold, U. Surf. Sci. 2014, 621, L1–L4. doi:10.1016/j.susc.2013.10.015
27. Shimizu, R.; Iwaya, K.; Ohsawa, T.; Shiraki, S.; Hasegawa, T.; Hashizume, T.; Hitosugi, T. ACS Nano 2011, 5, 7967–7971. doi:10.1021/nn202477n
28. Ohsawa, T.; Shimizu, R.; Iwaya, K.; Hitosugi, T. ACS Nano 2014, 8, 2223–2229. doi:10.1021/nn405359u
29. Sawada, D.; Sugimoto, Y.; Morita, K.; Abe, M.; Morita, S. Appl. Phys. Lett. 2009, 94, 173117. doi:10.1063/1.3127503
30. Stetsovych, O.; Todorovic, M.; Shimizu, T. K.; Moreno, C.; Ryan, J. W.; Pérez León, C.; Sagisaka, K.; Palomares, E.; Matolin, V.; Fujita, D.; Perez, R.; Custance, O. Nat. Commun. 2015, 6, 7265. doi:10.1038/ncomms8265
31. Torbrügge, S.; Oselland, F.; Reichling, M. J. Phys. Chem. C 2009, 113, 4909–4914. doi:10.1021/jp804026v
32. Torbrügge, S.; Reichling, M.; Iahiyama, A.; Morita, S.; Custance, O. Phys. Rev. Lett. 2009, 99, 056101. doi:10.1103/PhysRevLett.99.056101
33. Fukui, K.; Onishi, H.; Iwasawa, Y. Phys. Rev. Lett. 1997, 79, 4202. doi:10.1103/PhysRevLett.79.4202
34. Canova, F. F.; Foster, A. S.; Rasmussen, M. K.; Meinander, K.; Besenbacher, F.; Lauritsen, J. V. Nanotechnology 2012, 23, 325703. doi:10.1088/0957-4484/23/32/325703
35. Barth, C.; Reichling, M. Nature 2001, 414, 54–57. doi:10.1038/35102031
36. Suzuki, S.; Ohminami, Y.; Tsutsumi, T.; Shoaib, M. M.; Ichikawa, M.; Asakura, K. Chem. Lett. 2003, 32, 1098–1099. doi:10.1246/cl.2003.1098
37. Katsube, D.; Takase, Y.; Yamashita, H.; Abe, S.; Wakaya, F.; Abe, M. Mater. Trans. 2015, 56, 1310–1313. doi:10.2320/matertrans.M2015133
38. Bettac, A.; Koeble, J.; Winkler, K.; Uder, B.; Maier, M.; Feltz, A. Nanotechnology 2009, 20, 264009. doi:10.1088/0957-4484/20/26/264009
39. Hoffmann, R.; Lantz, M. A.; Hug, H. J.; van Schendel, P. J. A.; Kappenberger, P.; Martin, S.; Barattof, A.; Güntherodt, H.-J. Appl. Surf. Sci. 2002, 188, 238–244. doi:10.1016/S0169-4332(01)00915-1
40. Heyde, M.; Sterrer, M.; Rust, H.-P.; Freund, H.-J. Appl. Phys. Lett. 2005, 87, 083104. doi:10.1063/1.212523
41. Abe, M.; Sugimoto, Y.; Morita, S. Nanotechnology 2005, 16, S68. doi:10.1088/0957-4484/16/3/013
42. Yurtsever, A.; Sugimoto, Y.; Abe, M.; Morita, S. Nanotechnology 2010, 21, 165702. doi:10.1088/0957-4484/21/16/165702
43. Yokoyama, K.; Ochi, T.; Uchihashi, T.; Ashino, M.; Sugawara, Y.; Suehira, N.; Morita, S. Rev. Sci. Instrum. 2006, 77, 128–132. doi:10.1063/1.1501747
44. Sugimoto, Y.; Abe, M.; Hirayama, S.; Oyabu, N.; Custance, O.; Morita, S. Nat. Mater. 2005, 4, 156–159. doi:10.1038/nmat1297
45. Sugimoto, Y.; Pou, P.; Abe, M.; Jenike, P.; Pérez, R.; Morita, S.; Custance, O. Nature 2007, 446, 64–67. doi:10.1038/nature05830
46. Abe, M.; Sugimoto, Y.; Namikawa, T.; Morita, K.; Oyabu, N.; Morita, S. Appl. Phys. Lett. 2007, 90, 203103. doi:10.1063/1.2739410
47. Sugimoto, Y.; Pou, P.; Custance, O.; Jenike, P.; Abe, M.; Perez, R.; Morita, S. Science 2008, 322, 413–417. doi:10.1126/science.1160601
48. Sugimoto, Y.; Abe, M.; Yoshimoto, K.; Custance, O.; Yi, I.; Morita, S. Appl. Surf. Sci. 2005, 241, 23–27. doi:10.1016/j.apsusc.2004.09.012
49. Okano, M.; Kajimura, K.; Wakiyama, S.; Sakai, F.; Mizutani, W.; Ono, M. J. Vac. Sci. Technol., A 1987, 5, 3313–3320. doi:10.1116/1.574189
50. Rugar, D.; Mamin, H. J.; Geuehner, P. Appl. Phys. Lett. 1989, 55, 2588–2590. doi:10.1063/1.101987
51. Albrecht, T. R.; Grütter, P.; Horne, D.; Ruger, D. J. Appl. Phys. 1991, 69, 668–673. doi:10.1063/1.347347
52. Kurita, D.; Ohta, S.; Sugiyura, K.; Ohta, H.; Koumoto, K. J. Appl. Phys. 2006, 100, 096105. doi:10.1063/1.2362990
53. Katayama, M.; Ikesaka, S.; Kuwano, J.; Koinuma, H.; Matsumoto, Y. Surf. Sci. 2008, 602, 132107. doi:10.1016/j.susc.2008.03.028
54. Chambers, S. A.; Wang, C. M.; Thevuthasan, S.; Droubay, T.; McCready, D. E.; Lea, A. S.; Shulthanandan, V.; Windisch, C. F., Jr. Thin Solid Films 2002, 418, 197–210. doi:10.1016/S0040-6090(02)00709-5
55. Shao, R.; Wang, C.; McCready, D. E.; Droubay, T. C.; Chambers, S. A. Surf. Sci. 2007, 601, 1582–1589. doi:10.1016/j.susc.2007.01.039
56. We have tried different parameters for sample temperature and oxygen partial pressure. The best conditions were found according to Figure 3.
57. Horcas, I.; Fernández, R.; Gómez-Rodríguez, J. M.; Colchero, J.; Gómez-Herrero, J.; Baro, A. M. Rev. Sci. Instrum. 2007, 78, 013705. doi:10.1063/1.2432410
58. Selvin, M.; Daniel, B.; Mansfeldova, V.; Kavan, L.; Scheiber, P.; Fidler, M.; Schmid, M.; Diebold, U. Surf. Sci. 2014, 626, 61–67. doi:10.1016/j.susc.2014.04.001
59. Gariglio, S.; Reyren, N.; Caviglia, A. D.; Triscone, J.-M. J. Phys.: Condens. Matter 2009, 21, 164213. doi:10.1088/0953-8984/21/16/164213
60. Tebano, A.; Fabbri, E.; Pergolesi, D.; Balestrino, G.; Traversa, E. ACS Nano 2012, 6, 1278–1283. doi:10.1021/nn203991q
61. Shalom, M. B.; Sachs, M.; Rakhmilevich, D.; Palevski, A.; Dagan, Y. Phys. Rev. Lett. 2010, 104, 126802. doi:10.1103/PhysRevLett.104.126802

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