Irreversible Oxidation of the Active-site Cysteine of Peroxiredoxin to Cysteine Sulfonic Acid for Enhanced Molecular Chaperone Activity*

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The thiol (–SH) of the active cysteine residue in peroxiredoxin (Prx) is known to be reversibly hyperoxidized to cysteine sulfenic acid (–SO2H), which can be reduced back to thiol by sulfiredoxin/sestrin. However, hyperoxidized Prx of an irreversible nature has not been reported yet. Using an antibody developed against the sulfonylated (–SO3H) yeast Prx (Tsa1p) active-site peptide (AFTFVCPTEI), we observed an increase in the immunoblot intensity in proportion to the H2O2 concentrations administered to the yeast cells. We identified two species of hyperoxidized Tsa1p: one can be reduced back (reversible) with sulfiredoxin, and the other cannot (irreversible). Irreversibly hyperoxidized Tsa1p was identified as containing the active-site cysteine sulfonic acid (Tsa1p-SO3H) by mass spectrometry. Tsa1p-SO3H and Tsa1p-SO2H was not an autoxidation product of Tsa1p-SO2H and was maintained in yeast cells even after two doubling cycles. Tsa1p-SO3H self-assembled into a ring-shaped multimeric form was shown by electron microscopy. Although the Tsa1p-SO2H multimer lost its peroxidase activity, it gained ~4-fold higher chaperone activity compared with Tsa1p-SH. In this study, we identify an irreversibly hyperoxidized Prx, Tsa1p-SO3H, with enhanced molecular chaperone activity and suggest that Tsa1p-SO3H is a marker of cumulative oxidative stress in cells.

A typical reaction mechanism of cysteine-based peroxide reductases, such as 2-Cys peroxiredoxins (Prxs), includes the oxidation of an active-site thiolate to sulfenic acid by accepting an oxygen atom from peroxide, resolution by a nearby cysteine residue to form a disulfide bond, and final reduction to thiolates sequentially (9–12). Hyperoxidation of the peroxidatic cysteine (CP) of Prx to cysteine sulfenic or cysteine sulfonic acid can occur both under normal culture conditions and under oxidative stress (13–16). During normal peroxide catalysis of human Prx I with a low and steady concentration of H2O2 (<1 μM), 0.072% of the CP thiolate hyperoxidizes to cysteine sulfenic acid per turnover (17). Hyperoxidation of a cysteine residue to cysteine sulfenic acid alters the structure and the catalytic activities of various proteins, such as Prxs, MMP-7 (matrix metalloproteinase-7) (18), glyceraldehyde-3-phosphate dehydrogenase (14), DJ-1 (20), carbonic anhydrase III (21), MKP3 (mitogen-activated protein kinase phosphatase-3) (22), and PTP-1B (protein-tyrosine phosphatase-1B) (23, 24). Cysteine sulfinylation in proteins has been recognized as a process of losing biological function, but it is now seen as a precise cellular regulation mechanism after the identification of the sulfiredoxin/sestrin-mediated retroreduction, which reverses cysteine sulfenic acid (Cys-SO2H) to sulfhydryl (Cys-SH), and the identification of the novel functions of sulfinated proteins (8, 25–28, 30–34). Sulfinyl hyperoxidation induces a functional switch of human Prx II, a 2-Cys Prx, from a cellular peroxidase to a molecular chaperone (32, 33). Temporal sulfinylation of Tsa1p, a yeast 2-Cys Prx, negatively regulates the Tsa1p-Pap1p redox relay (30). Reversible sulfinylation of human Prx II with an accompanying structural transition to large filamentous oligomeric forms regulates hydrogen peroxide-dependent cell cycle arrest and release (34).

In contrast to the reversibly hyperoxidized Cys-SO2H in proteins, irreversibly hyperoxidized Cys-SO3H is generally believed to be an impaired dead end product, which is destined to be removed (35, 36). The sulfonic acid state of hyperoxidized CP of mammalian 2-Cys Prx proteins, such as human Prxs II and III, was observed by in-gel tryptic digestion followed by

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‡‡1 The abbreviations used are: Prx, peroxiredoxin; CP, peroxidatic cysteine; SD, synthetic defined; Trx, thioredoxin; DTT, dithiothreitol; Smx1, sulfiredoxin 1; CHAPS, 3-[3-cholamidopropyl]dimethylammonio]-1-propanesulfonic acid.

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mass spectrometric analysis (15). Later, it was proven that sulfonylation of Cys occurs as a result of air oxidation from cysteine sulfinic acid during protease treatment for peptide fingerprinting under denaturing conditions (17). To date, there is no known systematic study examining the conditions of irreversible hyperoxidation or switching of enzyme activity by irreversibly hyperoxidized cysteine residues. Here, we report the hyperoxidation of the active-site Cys-SH to Cys-SO3H in Tsa1p, the reaction that induces activity switching from a peroxidase to an enhanced molecular chaperone. In addition, we also demonstrate the maintenance of Tsa1p-SO3H in yeast cells and the possibility of hyperoxidation systems that facilitate the irreversible hyperoxidation of Tsa1p in yeast cells.

EXPERIMENTAL PROCEDURES

Preparation of Antibodies Specific for Sulfonylated Tsa1p—A peptide (AFTVFCPTPEI) that corresponds to the active site of yeast Tsa1p was oxidized by performic acid, conjugated to key-peptide (AFTFVCPTEI) that corresponds to the active site of Tsa1p was verified with anti-Tsa1p-SO3H antiserum.

Hyperoxidation and Retroreduction of Tsa1p in Yeast Cells and in Vitro—The Saccharomyces cerevisiae wild-type strain S288C was used in this experiment. Yeast cells were grown overnight at 30 °C in YPD medium (1% yeast extract, 2% Bactopeptone, and 2% glucose) and then transferred to synthetic defined (SD) medium (0.67% yeast nitrogen base and 2% glucose). For hyperoxidation, log-phase cells (A600 = 0.5) were exposed to various concentrations of H2O2 for 10 min and harvested. For retroreduction of Tsa1p in yeast cells, after a 10-min exposure to H2O2, yeast cells were allowed to recover in fresh medium. Yeast cell lysates were prepared with glass beads using a minibead beater (BioSpec Products, Inc.) in lysis buffer (50 mM Tris-HCl (pH 7.5), 1 mM EDTA, and 1 mM phenylmethylsulfonyl fluoride), and the crude cell lysates were separated on SDS-polyacrylamide gel or a two-dimensional polyacrylamide gel. Tsa1p was visualized with anti-Tsa1p antiserum regardless of its oxidation state, and the presence of hyperoxidized Tsa1p was verified with anti-Tsa1p-SO3H antiserum. In vitro hyperoxidation of Tsa1p (0.31 μM) was performed in a thiolredoxin (Trx)-dependence oxidation mixture containing 50 mM HEPES-NaOH (pH 7.0), 5.95 μM Trx, 0.1 μM Trx reductase, and 0.25 mM NADPH with various H2O2 concentrations for 10 min at 30 °C. For retroreduction in vitro, hyperoxidized Tsa1p was incubated in retroreduction mixture containing 1.2 μM sulfiredoxin 1 (Srx1), 1 mM ATP, 1 mM MgCl2, and 2.5 mM dithiothreitol (DTT) for 1 h at 30 °C.

Two-dimensional SDS-PAGE—Yeast crude extracts (10 μg) were mixed with rehydration buffer (8 M urea, 2% CHAPS, 0.5% immobilized pH gradient buffer, 20 mM DTT, and 0.005% bromphenol blue) and loaded onto immobilized pH gradient strips (pH 3–10, 13 cm, Amersham Biosciences). Iselectric focusing was performed with an IPGphor isoelectric focusing unit (Amersham Biosciences) according to the instruction manual. After reduction and alkylation, second dimension electrophoresis was conducted on a 12% SDS-polyacrylamide gel using an SE600 vertical unit (Amersham Biosciences). Proteins from the SDS-polyacrylamide gel were electrochemically transferred onto nitrocellulose membrane. Tsa1p was probed on the membrane with anti-Tsa1p or anti-Tsa1p-SO3H antiserum, and the immune complexes were visualized with alkaline phosphatase-conjugated secondary antibodies using nitro blue tetrazolium and 5-bromo-4-chloro-3-indolyl phosphate.

Immunoprecipitation—Yeast cell lysates were prepared with buffer A (50 mM Tris-HCl (pH 8.0), 150 mM NaCl, and 1 mM phenylmethylsulfonyl fluoride), and anti-Tsa1p antibody-conjugated protein G-Sepharose (2 μL of antibody with 30 μL of resin) was equilibrated with buffer B (buffer A containing 0.1% Nonidet P-40). Cell lysates (500 μg/ml protein) were incubated with the resin for 1 h at 4 °C, and the supernatant was then removed by centrifugation. After washing with buffer B, immune complexes were subjected to immunoblot analysis with anti-Srx1 antibody.

Mass Analysis—Mass analysis was performed in parallel on a QSTAR pulsar quadrupole time-of-flight mass spectrometer (Applied Biosystems) equipped with nanoelectrospray ion sources (Protana, Odense, Denmark). The analytes were dissolved in acetonitrile/water (50:50, v/v) and loaded into pre-opened “medium” borosilicate spray capillaries for off-line nanoelectrospray (Proxeon Biosystems, Odense). A potential of 1000 V was applied to the loaded glass capillary tip. Electrospray ionization mass spectra were collected for 1 min, and the masses of proteins were calculated by deconvoluting multiple charged ions with m/z values and corresponding total charges using Analyst QS software (Version 1.0, Applied Biosystems).

Size Exclusion Chromatography—Size exclusion chromatography was performed using a TSK-GEL G3000SWXL column (Tosho). Intact Tsa1p-SH or Tsa1p-SO3H was injected into the column, which was equilibrated with 50 mM sodium phosphate buffer (pH 7.0) containing 100 mM NaCl at a flow rate of 0.5 ml/min. Determination of molecular masses was referenced against protein standards (Bio-Rad): thyroglobin (670 kDa), bovine γ-globulin (158 kDa), ovalbumin (44 kDa), equine myoglobin (17 kDa), and vitamin B12 (1.35 kDa).

Electron Microscopy—Transmission electron microscopy analyses were performed at the Korea Basic Science Institute. Proteins were negatively stained with 2% (w/v) uranyl acetate, and transmission electron microscope images were obtained with a Zeiss EM 912 Omega microscope operating at an acceleration voltage of 120 kV.

Trx-dependent Peroxidase Assay and Chaperone Assay—the rate of peroxide reduction was monitored as a decrease in absorbance at 340 nm by NADPH oxidation at 30 °C using a Jasco UV-visible spectrophotometer equipped with a thermostatted cell holder. Reactions were initiated by adding the designated concentrations of hydrogen peroxide in a Trx-dependent oxidation mixture containing 50 mM HEPES-NaOH (pH 7.0), 5.95 μM Trx, 0.1 μM Trx reductase, and 0.25 mM NADPH. Specific activities were calculated from the initial linear portion of the assay. Chaperone activity was measured by its capacity to inhibit thermal aggregation of firefly luciferase (Roche Applied Science). Luciferase (0.15 μM) was incubated with Tsa1p-SH or Tsa1p-SO3H in 25 mM HEPES-KOH buffer (pH 7.9) at 47 °C for 30 min, and the thermal aggregation of luciferase was monitored by measuring the light scattering at 320 nm.
RESULTS

Two Different Hyperoxidation States of Tsa1p—Hyperoxidation of C`p residues of Prx proteins to cysteine sulfenic acid (C`p-SO2H) and the corresponding spot migration on two-dimensional polyacrylamide gels have been studied in both mammalian and yeast cells (15–17). Although an antiserum against mammalian hyperoxidized Prx proteins was developed and characterized, its specificity for two hyperoxidized states, C`p-SO2H and C`p-SO3H, needs to be redetermined in detail (38). We also raised rabbit antiserum against the Tsa1p active-site peptide containing C`p-SO3H following the previous protocol for mammalian anti-Prx-SO3H antiserum (38). The reactivity of anti-Tsa1p-SO3H antiserum to yeast cell lysates increased in proportion to hydrogen peroxide concentrations administered to the cells (Fig. 1, A and B, upper panels). By comparing the mobilities of immunoreactive bands to anti-Tsa1p and anti-Tsa1p-SO3H antiserum (Fig. 1, A and B), we concluded that the hyperoxidized Tsa1p migrated slower than the reduced form on SDS-polyacrylamide gel. However, the extent of intensity increases in anti-Tsa1p-SO3H-immunoreactive bands upon hydrogen peroxide treatment did not correlate well with that of slow migrating anti-Tsa1p-immunoreactive bands. Although anti-Tsa1p-SO3H-immunoreactive bands showed linear increases up to 10 mM hydrogen peroxide (Fig. 1, A and B, upper panels), slow migrating anti-Tsa1p-immunoreactive bands showed saturation at 2 mM hydrogen peroxide, suggesting that the latter contains other species (probably Tsa1p-SO2H) as well as the species reactive to anti-Tsa1p-SO3H antiserum. Two-dimensional PAGE analysis showed results consistent with those from SDS-PAGE (Fig. 1C). With increasing concentrations of hydrogen peroxide administered to the cells, the hyperoxidized Tsa1p spot gained intensity, whereas the reduced Tsa1p spot lost intensity. Consistently, hyperoxidized Tsa1p migrated slower than the reduced form on two-dimensional polyacrylamide gel (Fig. 1C, left panel). The hyperoxidized spots detected by the two antisera showed different patterns of intensity increases, similar to the results in Fig. 1A, also suggesting the existence of different species at the same spot location.

Retrodeducibility of the Two Different Hyperoxidized Forms of Tsa1p—The retroreduction of Tsa1p-C`p-SO2H to Tsa1p-C`p-SH by the catalytic activity of Srx1/sestrin 2 in the presence of ATP and MgCl2 has been reported (25–28). We wondered if the two different hyperoxidized forms of Tsa1p revert to the reduced form (Tsa1p-C`p-SH). When cells were treated with 1 mM hydrogen peroxide, ~78% of the hyperoxidized Tsa1p was retroreduced within 3 h after the removal of peroxide. However, when cells were exposed to 2.5 and 5 mM hydrogen peroxide, 53 and 28% of the Tsa1p were retroreduced, respectively. Treatment with 10 mM hydrogen peroxide totally disabled retroreduction. Based on these results, high doses of hydrogen peroxide seem to make irreversibly oxidized products. Subtraction of reversibly oxidized products (fast migrating bands) (Fig. 2A, middle panel) from the total oxidized product (slow migrating bands) (upper panel) can yield irreversibly oxidized products (lower panel), presented in Fig. 2B. As expected, irreversibly hyperoxidized Tsa1p obtained by treatment with 10 mM hydrogen peroxide failed to associate with Srx1, whereas reversibly hyperoxidized Tsa1p obtained by treatment with 0.5 mM hydrogen peroxide associated with Srx1 (Fig. 2C). These results support the presence of irreversibly hyperoxidized Tsa1p, which is reactive to anti-Tsa1p-SO3H antiserum.

FIGURE 1. Different reactivity of hyperoxidized Tsa1p with anti-Tsa1p-SO3H antiserum. Log-phase yeast cells (A500 = 0.5) were exposed to the indicated concentrations of H2O2 for 10 min and disrupted in lysis buffer (50 mM Tris-HCl (pH 7.5), 1 mM EDTA, and 1 mM phenylmethylsulfonyl fluoride). Cell lysates were separated by SDS-PAGE (A and B) or two-dimensional PAGE (C) and probed with anti-Tsa1p or anti-Tsa1p-SO3H antiserum. The positions of hyperoxidized (Ox) and reduced (Re) forms of Tsa1p are indicated. WB, Western blot.

FIGURE 2. Reversibly and irreversibly hyperoxidized forms of Tsa1p. A, exponentially growing yeast cells (A500 = 0.5) were exposed to the indicated concentrations of H2O2 for 10 min. For the retroreduction of hyperoxidized Tsa1p, yeast cells were allowed to recover from H2O2 stress for 3 h in fresh SD medium in the presence of cycloheximide (50 μg/ml). Redox states of Tsa1p were examined by comparing immune reactivity with anti-Tsa1p and anti-Tsa1p-SO3H antisera. WB, Western blot. B, shown is a graphical demonstration of the relative proportions of reduced ( ), reversibly oxidized ( ), and irreversibly hyperoxidized ( ) Tsa1p. The relative intensities of the corresponding bands probed with anti-Tsa1p antibody were quantified by densitometry (Typhoon, GE Healthcare). C, immunoprecipitation (IP) with anti-Tsa1p antibody was performed using the crude yeast lysates exposed to 0.5 mM (for reversible hyperoxidation) or 10 mM (for irreversible hyperoxidation) hydrogen peroxide for 10 min. Immune complexes were separated on SDS-polyacrylamide gel and visualized with anti-Srx1 antiserum.
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Specificity of Anti-Tsa1p-SO3H Antiserum for Irreversibly Hyperoxidized Tsa1p—Purified Tsa1p-SH (0.3 μm) was incubated in a 150-μl reaction mixture containing 50 mM HEPES-NaOH (pH 7.0), 5.95 μm Trx, 0.1 μm Trx reductase, 0.25 mM NADPH, and 0.2 or 50 mM H2O2 for reversible or irreversible hyperoxidation of Tsa1p, respectively (A). The hyperoxidation reaction was performed for 10 min at 30 °C. The retroreducibility of hyperoxidized Tsa1p was determined by reaction with 1.2 μM Srx1, 1 mM ATP, 1 mM MgCl2, and 2.5 mM DTT for 1 h at 30 °C. The reaction specificity of anti-Tsa1p-SO3H antiserum was examined in the mixture of reversibly and irreversibly hyperoxidized Tsa1p in the indicated ratios to yield a total of 100 ng of Tsa1p (B) and in the same amount of the two oxidized forms (C). WB, Western blot.

Mass Analysis of Irreversibly Hyperoxidized Tsa1p—The mass of irreversibly hyperoxidized Tsa1p was determined by electrospray ionization mass spectrometry. Irreversibly hyperoxidized Tsa1p was purified from crude extracts of yeast cells treated with 10 mM hydrogen peroxide as described previously for the purification of Tsa1p-SH (39). Irreversibly hyperoxidized Tsa1p was identified by SDS-PAGE and two-dimensional PAGE analyses (Fig. 4, A and B). The mass spectra of reduced Tsa1p and irreversibly hyperoxidized Tsa1p with multiple charges are shown in Fig. 4C. The calculated monoisotopic mass values were 21,458.4 and 21,506.2 Da, respectively. The mass difference was ~48 Da, which suggests the presence of three oxygen atoms in irreversibly hyperoxidized Tsa1p.

Maintenance of Tsa1p-SO3H in Cells—Irreversibly oxidized proteins in cells are generally believed to be degraded by proteolytic systems employing principally the 20 S proteasome as well as the ubiquitin-26 S proteasome pathways (35, 36). To examine the life span of Tsa1p-SO3H in yeast cells, exponentially growing yeast cells were exposed to 10 mM hydrogen per-
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FIGURE 5. Tsa1p-SO3H is not an autoxidation product of Tsa1p-SO2H. A, for the preparation of Tsa1p-SO2H, yeast cells were exposed to 0.5 mM H2O2 at 30 °C for 10 min and then incubated for 3 h in fresh SD medium in the presence of 50 μM cycloheximide to confirm the reversibility of hyperoxidized Tsa1p. B, after exposure to 0.5 mM H2O2 for 10 min at A600nm = 0.5, the lysates were prepared in 50 mM HEPES-NaOH (pH 7.0) containing 1 mM EDTA and 1 mM phenylmethylsulfonyl fluoride. Crude lysates (~4 mg/ml) were diluted to 1 mg/ml with 50 mM HEPES-NaOH (pH 7.0) or 6 M guanidine HCl (final concentration) (pH 8.0) and then subjected to air exposure at 4 °C for 5 days. Samples were taken at the indicated time points, and autoxidation to Tsa1p-SO3H was analyzed by immunoblotting with anti-Tsa1p-SO2H antibody. C, Tsa1p-SO2H was prepared from purified Tsa1p-SH (0.3 mg/ml) by treatment with 0.2 mM H2O2 in the presence of a Trx-dependent oxidation mixture containing 50 mM HEPES-NAOH (pH 7.0), 5.95 μM Trx, 0.1 mM Trx reductase, and 2.5 mM DTT for 1 h at 30 °C. D, autoxidation of purified Tsa1p-SO2H to Tsa1p-SO3H was performed as described for B. WB, Western blot.

FIGURE 6. Maintenance of Tsa1p-SO2H in growing yeast cells. A, log-phase yeast cells (A600nm = 0.5) were exposed to 10 mM H2O2 and then incubated with fresh SD medium at 30 °C. At the indicated time points, yeast cells in 30 ml of culture were collected and disrupted. The same volume of each sample was separated on 12% SDS-polyacrylamide gel followed by immunoblotting with anti-Tsa1p and anti-Tsa1p-SO3H antibodies. WB, Western blot. B, growth of yeast cells in fresh SD medium after H2O2 exposure was monitored by absorbance at 600 nm (○). The relative amounts of Tsa1p-SO2H (○) and Tsa1p-SH (△) were estimated by densitometry (Gel Doc XR system, Bio-Rad). arb., arbitrary.

oxide for 10 min and incubated with fresh SD medium. After 4 h of lag phase, yeast cells achieved exponential growth with a doubling time of ~3 h (Fig. 6B). The Tsa1p-SO2H level did not change much until 6 h and then gradually decreased. From 6 to 12 h of the culture period, the cell number doubled twice, whereas total Tsa1p-SO2H decreased to half of the initial level. Because of the long lag period after the treatment with hydrogen peroxide, the determination of the half-life of Tsa1p-SO2H in yeast cells might be controversial. If we define the Tsa1p-SO2H half-life in exponentially growing yeast cells, it is reasonably estimated to be 6 h (Fig. 6). The negligible rate of irreversible hyperoxidation of Tsa1p under normal culture conditions and the maintenance of Tsa1p-SO2H over more than two doubling periods implicate a new role for Tsa1p-SO3H in yeast cells rather than overoxidized protein wastes.

Loss and Gain of Catalytic Activity by Hyperoxidation—Trx-dependent peroxidase activity was lost by hyperoxidation at Cp (Fig. 7A), but the considerable maintenance of Tsa1p-SO2H in the yeast cells led us to explore the endogenous catalytic activity of Tsa1p-SO2H. We measured the molecular chaperone activity of Tsa1p-SO2H because the oligomeric structure-dependent functional switching of Tsa1p from a peroxidase to a molecular chaperone and vice versa has been reported (31). Chaperone activities were determined by measuring the protection capacity of the heat-induced aggregation of luciferase. The molar ratios of Tsa1p-SH and Tsa1p-SO2H to luciferase requiring 70% protection were 2:1 and 0.5:1, respectively. Thus, Tsa1p-SO2H had ~4-fold higher chaperone activity compared with Tsa1p-SH (Fig. 7B).

Oligomeric Structure of Tsa1p-SO2H—The correlation between chaperone activities and their structures has been established in many small heat shock proteins (19, 29, 44). We explored whether this correlation is also applicable to Tsa1p-SO2H. A functional switching of Tsa1p from a peroxidase to a molecular chaperone accompanied by a shift to a higher molecular mass complex has been observed (31–33). Purified Tsa1p-SH from yeast cells was separated into two major peaks on a size exclusion chromatography column (TSK-GEL G3000SWXL) with calculated molecular masses of ~600 and 40 kDa, corresponding to two dodecameric forms and a dimeric form. Tsa1p-SO2H eluted with one major peak at ~530 kDa, which we suppose to be two dodecamers (Fig. 8A). To compare the oligomeric structures of Tsa1p-SH and Tsa1p-SO2H, we performed electron microscopy, and the proteins were visualized by negative staining with uranyl acetate. Tsa1p-SO2H showed a homogeneous ring-shaped structure 14 nm in diameter, whereas Tsa1p-SH showed an irregular configuration with...
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~10% of a sporadic ring-shaped structure (Fig. 8B). Sphereshaped particles were not found with either Tsa1p-SH or Tsa1p-SO3H. These results confirm that irreversible hyperoxidation induces oligomerization to higher molecular mass forms, which we suppose to be two dodecamers, and that these higher molecular mass structures are responsible for a molecular chaperone function.

Comparison of Sulfonyl Hyperoxidation in Yeast Cells and an in Vitro System—In vitro preparation of Tsa1p-SO2H using the Trx system was feasible, and the hyperoxidized product was mostly Tsa1p-SO2H (Figs. 3 and 5). In contrast, preparation of Tsa1p-SO2H using yeast cells was not successful because the hyperoxidized products were a mixture of Tsa1p-SO2H and Tsa1p-SO3H (data not shown). In addition, preparation of Tsa1p-SO3H was feasible, and the hyperoxidized product was Tsa1p-SO3H. These results confirm that irreversible hyperoxidation of Tsa1p in yeast cells is not successful because the hyperoxidized product is Tsa1p-SO3H.

DISCUSSION

In contrast to catalases, which directly catalyze the decomposition of peroxides to water and oxygen, peroxidases, which consume expensive cellular reducing equivalents to reduce peroxides, should be systematically regulated to conserve cellular reducing resources. After the finding of inactivation of Tsa1p during peroxidase catalysis in yeast (13), the hyperoxidation of active-site cysteine (Cp) from sulfhydryl to sulfinic or sulfonic acid in in vitro assays and in-cell analysis was reported (14–17). In addition, studies that followed reported the reversibility of sulfenic Cp to sulfhydryl in mammalian 2-Cys Prxs and enzymatic systems responsible for that retroligation in yeast and mammalian systems (25–28).

To quantify hyperoxidized Tsa1p in yeast cells using immunological methods, we developed rabbit antiserum against the peptide encompassing Tsa1p-Cp-SO3H using the same method as described previously for mammalian anti-Prx-SO3H antiserum (38). We identified two different forms of hyperoxidized Tsa1p (reversible Tsa1p-SO3H and irreversible Tsa1p-SO3H) (Figs. 1–4) based on the reactivity of the antiserum we developed. In contrast to mammalian anti-Prx-SO3H antisera, which equally recognizes both Prx-CP-SO2H and Prx-CP-SO3H (38), the reactivity of anti-Tsa1p-SO3H antiserum was specific to Tsa1p-Cp-SO3H (Fig. 3). Tsa1p-Cp-SO3H was detectable
above 20 ng in our alkaline phosphatase-based immune complex detection system, but 100 ng of Tsa1p-SO₂H was not visible using the same detection system (Fig. 3). When HeLa cell extracts treated with various concentrations of hydrogen peroxide were probed, anti-Tsa1p-SO₃H antiserum clearly showed distinguishable reactivity toward irreversibly hyperoxidized 2-Cys Prx spots, although the exact oxidation status of C₅₇ needs to be verified (data not shown). Considering the active-site amino acid sequence identity between yeast Tsa1p-(AFT-FVC₅₇PTEI-) and mammalian 2-Cys Prx (DFTFVC₅₇PTEI), anti-Tsa1p-SO₃H antiserum can also be used for the detection of mammalian irreversibly hyperoxidized 2-Cys Prx proteins. The points that Tsa1p-SO₃H has a half-life of more than two doubling times in yeast, Tsa1p-SO₂H is not an autodestruction product, and Tsa1p-SO₃H is an irreversibly hyperoxidized product together support the use of Tsa1p-SO₂H as a molecular marker of cumulative oxidative stress in cells. Anti-Tsa1p-SO₃H antiserum, which specifically recognizes irreversibly hyperoxidized Tsa1p as well as mammalian 2-Cys Prxs, can be an efficient tool for probing a cell's history of withstanding oxidative stresses.

Once impaired by oxidative stress, proteins are generally considered cellular waste and designated for degradation by the proteasome in cells. The longer half-life of irreversibly hyperoxidized Tsa1p-SO₂H (Fig. 6) compared with the average half-life of yeast proteins (37) suggests an intrinsic role for hyperoxidized Tsa1p-SO₂H in oxidatively stressed yeast cells. We identified the enhanced molecular chaperone activity and corresponding oligomeric structure of Tsa1p-SO₂H (Figs. 7 and 8). Tsa1p-SO₂H functioning as a molecular chaperone in cells is considered to be an advantage over Tsa1p-SH or Tsa1p-SO₃H in cells. Presumably, Tsa1p-SO₃H has a considerably shorter life span than Tsa1p-SO₂H because most of the Tsa1p-SO₂H is retroreduced to Tsa1p-SH when Srx1 is available. Because Srx1 is a noticeably inefficient enzyme with a very low turnover rate (kₑᵣᵣ = 0.1–0.8/min) (28) and because proper oligomeric structure is responsible for chaperone function, Tsa1p-SO₂H seems likely to act as a molecular chaperone only briefly until association with Srx1 for retroreduction. Tsa1p-SH has several different oligomeric forms (Fig. 8B) (31), which is the reason why it is a less effective molecular chaperone than Tsa1p-SO₂H. Therefore, Tsa1p-SO₂H functions as the most effective and longest lived molecular chaperone among the Tsa1p redox forms in yeast cells.

Comparing Tsa1p hyperoxidation in yeast cells with that in an in vitro oxidation system, we found that reversible sulfynil hyperoxidation was the favored reaction in the in vitro system but not in the yeast cells. In contrast, irreversible sulfynil hyperoxidation was preferred in yeast cells (Fig. 9). Our results show that atmospheric oxygen is not capable of oxidation of Tsa1p-SO₂H to Tsa1p-SO₃H (Fig. 5), which is consistent with previous observations (40–43). Hyperoxidation in vitro takes place only during peroxidatic turnover in the presence of a recycling system including Trx, Trx reductase, and NADPH (17). Taken together, the in vitro preference of sulfynil hyperoxidation and the in vivo preference of sulfynil hyperoxidation suggest the possible existence of an enzymatic system other than the Trx system that facilitates sulfynil hyperoxidation in yeast cells, and it remains to be elucidated. In summary, we have identified Tsa1p-SO₂H and its function as an effective and stable molecular chaperone and as a cellular marker that retains a cell's cumulative index of oxidative stress.

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