Supporting Information

Hydroxyl Porous Aromatic Frameworks for Efficient Adsorption of Organic Micropollutants in Water

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Table S1. Fitting parameters of pseudo-second-order kinetics and pollutant uptake rates by PAF materials.

| sorbent | BPA  | 2-NO | PCMX |
|---------|------|------|------|
|         | $K_{obs}$ ($g\,mg^{-1}\,min^{-1}$) | $R^2$ | $t_{eq}$ (min) | $K_{obs}$ ($g\,mg^{-1}\,min^{-1}$) | $R^2$ | $t_{eq}$ (min) | $K_{obs}$ ($g\,mg^{-1}\,min^{-1}$) | $R^2$ | $t_{eq}$ (min) |
| PAF-80  | 0.0069 | 0.999 | >45 | 0.0035 | 0.98 | >60 | 0.0113 | 0.99 | >45 |
| PAF-81  | 0.0099 | 0.98 | >60 | 0.0097 | 0.98 | >60 | 0.0036 | 0.97 | >60 |
| PAF-82  | 0.3203 | 0.998 | 15 | 0.1192 | 0.99 | 30 | 0.1949 | 0.999 | 10 |
Fig. S8 (a-c) Langmuir adsorption isotherms of BPA onto PAF-80, PAF-81, and PAF-82. (d-f) Langmuir adsorption isotherms of 2-NO onto PAF-80, PAF-81, and PAF-82. (g-i) Langmuir adsorption isotherms of PCMX onto PAF-80, PAF-81, and PAF-82.
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Fig. S10 Molecule sizes of (a) Bisphenol A, (b) 2-Naphthol, and (c) p-Chloroxylenol.
Fig. S11  TGA of PAF-82 before adsorption (red curves) and after adsorption (purple curves) to different micropollutants of (a) BPA, (b) 2-NO and (c) PCMX.

References

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