Ring-opening polymerization of rac-lactide catalyzed by crown ether complexes of sodium and potassium iminophenoxides†

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A series of iminophenoxide ligand precursors [2-(RN=CH)C₆H₅OH] (HL): R = C₆H₅; HL²: R = 2,6-Pr₂C₆H₃ and [2-(RN=CH)-4,6-Bu₂C₆H₃OH] (HL³: R = C₆H₅; HL⁴: R = 2,6-Pr₂C₆H₃) were synthesized. These compounds reacted with NaN(SiMe₃)₂/15-crown-5 or KN(SiMe₃)₂/18-crown-6 to afford corresponding crown ether complexes of sodium and potassium iminophenoxides (1: (15-C-5)NaL¹, 2: (15-C-5)NaL², 3: (15-C-5)NaL⁴; 4: (18-C-6)KL¹, 5: (18-C-6)KL², 6: (18-C-6)KL³, 7: (18-C-6)KL⁴). Catalysts of the complexes toward the ring-opening polymerization of rac-lactide was studied. Each of the complexes exhibited high catalytic activity at room temperature. Complexes 2, 3, 6 and 7 showed poor isotactic selectivity and relatively broad molecular weight distributions. Complexes 1 and 5 resulted in more stereoregular polymers with Pm values of 0.58 and 0.66, respectively. Complex 4 led to the best selectivity for isotacticity (Pm = 0.75) when the polymerization was performed in toluene at 0 °C.

Introduction

Poly(lactide) (PLA), an important biodegradable and biocompatible polymer, has been used in a wide variety of applications such as drug delivery systems, resorbable sutures, and medical implants.1-11 The ring-opening polymerization (ROP) of lactides catalyzed by metal complexes is a major method for the synthesis of PLA with controlled molecular weights, low polydispersity indices (PDIs), and specific stereo-microstructures.12-14 A lot of metal complexes such as Zn,15-24 Mg,25-37 Al,38-45 Sn,46-57 and rare earth metal complexes58-71 have been proven to be effective for the ROP of lactides. In recent years alkali metal complexes also attracted great interest owing to their high catalytic activity and cheapness, wide availability and nontoxic nature of alkali metals.75-101 Several alkali metal complexes supported by iminophenoxide ligands have been reported to efficiently initiate the ROP of lactides. In 2012 Chen et al. reported the synthesis and catalysis of lithium and sodium iminophenoxide complexes toward the ROP of l-lactide.92 In 2015 Cano et al. synthesized a series of lithium, sodium, and potassium iminophenoxide complexes, and found that the lithium complexes resulted in heterorich-PLA (Pᵢ = 0.75),93 while sodium and potassium complexes presented poor stereo-selectivity in the polymerization of rac-lactide.92 Chakraborty et al. reported similar complexes which catalyzed solvent free ROP of rac-lactide with high catalytic activity and led to heterorich-PLA.93 Recently Wu and co-workers reported that crown ether complexes of sodium and potassium mono-phenoxides exhibited high catalytic activity and high iso-selectivity for the ROP of rac-lactide.94-101 However, the phenoxylic ligands are required to have very bulky ortho-substituents on the aromatic rings to achieve good selectivity.94-101 It is interesting to design simpler ligands to get high catalytic selectivity of complexes. A strategy is to use chelate ligands to restrict rotation of M-OAr bonds of alkali metal complexes. In an earlier work we demonstrated that crown ether complex of potassium quinolin-8-olate did initiate iso-selective ROP of rac-lactide. The Pm value reached 0.75 when the polymerization of 100 equiv. of rac-lactide was performed in toluene at 0 °C.104 We intended to examine more chelate ligands to improve catalytic selectivity of the chelate alkali metal complexes. Herein we report synthesis and catalytic study of crown ether complexes of sodium and potassium iminophenoxides with different substituents on the aromatic rings.

Results and discussion

Inimophenols 2-(PhN=CH)C₆H₄OH (HL¹), 2-(2,6-Pr₂C₆H₃N=CH)-C₆H₄OH (HL²), 2-(PhN=CH)-3,5-Bu₂C₆H₃OH (HL³) and 2-(2,6-Pr₂C₆H₃N=CH)-3,5-Bu₂C₆H₃OH (HL⁴) were synthesized according to reported procedures.92 Treatment of the iminophenols with NaN(SiMe₃)₂/15-crown-5 or KN(SiMe₃)₂/18-crown-6 in toluene at room temperature afforded corresponding...
crown ether complexes of sodium and potassium iminophenoxides 1–7 respectively (Scheme 1). However, treatment of HL with NaN(SiMe3)2 in toluene or THF in the presence of 15-crown-5 could not generate expected crown ether sodium complex.

Complexes 1–7 are yellow solids, soluble in DMSO, THF, and slightly soluble in toluene. They were characterized by elemental analyses, 1H and 13C{1H} NMR spectroscopy. Complexes 3 and 4 were further characterized by single crystal X-ray diffraction techniques. The ORTEP drawing of complex 3 is presented in Fig. 1, along with selected bond lengths and angles. The structure shows that the phenoxy ligand coordinates to the central sodium cation in a monodentate mode, rather than a O,N-chelate mode. This might be due to steric repulsion between the N-arylimino group and the crown ether ligand which combines with Na+ through coordination of the five oxygen atoms. The C(13)–O(6)–Na(1) angle of 159.67(17)° reveals approximately linear arrangement of C(13), O(6) and Na(1) atoms. The O6–Na1 distance of 2.1393(18) Å is slightly shorter than corresponding one in sodium 2-(anthracen-9-yl)-4,6-di-tert-butylphenolate crown ether complex (2.1478(18) Å).

The ORTEP drawing of complex 4 is presented in Fig. 2, along with selected bond lengths and angles. In the structure the phenoxy ligand also coordinates to the central metal in a monodentate mode. However, the C(15)–O(7)–K(1) angle of 130.8(3)° shows that the connectivity of C15, O7 and K1 atoms is not linear. The C15–O7–K1 plane is almost perpendicular to the adjacent aromatic ring plane, the torsion angle of K1–O7–C15–C14 being 96°. The K1–C7 distance of 2.623(3) Å is within a normal range compared with the reported crown ether potassium phenoxide complexes. The O(7)–C(15) distance of 1.289(5) Å is comparable with those in crown ether potassium phenoxide complexes.

In the presence of benzyl alcohol, catalysis of complexes 1–7 in the ROP of rac-lactide was tested, and the results are summarized in Table 1. Each of the complexes showed high catalytic activity in toluene at room temperature (Table 1, entries 1–7). When using a molar ratio of 100 : 1 : 1 for [LA]0/[M]0/[BnOH]0, complex 1 led to faster polymerization of rac-lactide than complexes 2 and 3 for the three similar sodium complexes (Table 1, entries 1–3). Potassium complexes 4–7 exhibited similar activity order. Thus, the complexes without ortho-tert-butyl group on the phenoxy rings led to faster ROP of rac-lactide than the other ones with ortho-tert-butyl groups on the phenoxy rings. It seems that the stereo hindrance of ortho-tert-butyl groups on the phenoxy rings decreases the ROP rate (Table 1, entries 4–7). But we cannot rule out influence of electron effect of the ortho- and para-position tert-butyl groups.
The ortho-position substituents on the N-aryl rings exhibited less effect on the catalytic activity of the complexes, which might be because the aromatic rings are far from the metal centers. In addition, the potassium complexes are less active than corresponding sodium complexes (5 vs. 1, 6 vs. 2, and 7 vs. 3), which is different from the previous report results.\textsuperscript{95-98} The molecular weights of the final polymers generated by those catalysts are lower than expected. A plausible reason could be the presence of transesterification side reaction in the process of polymerization.\textsuperscript{94-101} The relatively wide molecular weight distributions (1.18 to 1.39) and MALDI-TOF MS analysis (Fig. 3) of the polymers support this explanation. For example, the MS analysis of the polymer showed two main series of peaks. A series of peaks at 144 m + 108 + 23, which can be assigned to m(C6H8O4)+, whereas another series of peaks at 72 m + 108 + 23, which can be ascribed to m(C5H4O2)+ + BnOH + Na+. The series of peaks separated by m/z = 72 Da is indicative of transesterification occurring in the process of polymerization.

The sodium and potassium complexes 1 and 4 were selected for a more detailed study on their catalysis. At first, higher monomer loadings were tested using 1/BnOH and 4/BnOH, respectively (Table 1, entries 8–11 and 14–17). 1/BnOH showed higher activity than 4/BnOH. For example, 1/BnOH drove 500 equiv. of rac-lactide to PLA in 15 min in 99% conversion, whereas 4/BnOH drove 89% conversion of 500 equiv. of rac-lactide to spend 30 min. Each of 1- and 4-catalyzed ROP gave a linear relationship of the molecular weights of polymers to the ratio of monomer to catalyst (Fig. 4 and 5), implying that the polymerizations are controlled. Complexes 1 and 4 were demonstrated to catalyze the immortal ROP of 500 equiv. of rac-LA in the presence of 10 equiv. of BnOH, resulting in polymers with controlled molecular weights (Table 1, entries 12 and 18). We also examined the catalytic reaction by 4/BnOH in CH2Cl2 and THF, respectively (Table 1, entries 19 and 20). Both CH2Cl2 and THF were less effective than toluene. In these solvents the catalyst exhibited lower activity than in toluene and the molecular weight of the polymers were markedly lower than the theoretical values.

Table 1  ROP of rac-LA catalyzed by complexes 1–7

| Entry | Cat. | [LA]0/[Cat.0]/[BnOH]0 | Time (min) | Conv. (%) | M_{n,calc} \textsuperscript{c} (g mol \textsuperscript{-1}) | M_{n,GPC} \textsuperscript{d} (g mol \textsuperscript{-1}) | PDI | P_m \textsuperscript{e} |
|-------|------|----------------------|-----------|--------|------------------|------------------|------|--------|
| 1     | 1    | 100 : 1 : 1          | 1         | 99     | 14 400           | 10 100           | 1.32 | 0.58   |
| 2     | 2    | 100 : 1 : 1          | 5         | 97     | 14 100           | 9200             | 1.25 | 0.52   |
| 3     | 3    | 100 : 1 : 1          | 5         | 97     | 14 100           | 8600             | 1.30 | 0.52   |
| 4     | 4    | 100 : 1 : 1          | 3         | 98     | 14 200           | 7400             | 1.18 | 0.69   |
| 5     | 5    | 100 : 1 : 1          | 3         | 98     | 14 200           | 10 400           | 1.39 | 0.66   |
| 6     | 6    | 100 : 1 : 1          | 7         | 98     | 14 200           | 8900             | 1.27 | 0.53   |
| 7     | 7    | 100 : 1 : 1          | 10        | 97     | 14 100           | 7800             | 1.28 | 0.52   |
| 8     | 1    | 200 : 1 : 1          | 3         | 99     | 28 600           | 25 400           | 1.36 | 0.60   |
| 9     | 9    | 300 : 1 : 1          | 5         | 99     | 42 900           | 30 800           | 1.39 | 0.59   |
| 10    | 1    | 400 : 1 : 1          | 10        | 99     | 57 200           | 43 000           | 1.36 | 0.62   |
| 11    | 11   | 500 : 1 : 1          | 15        | 99     | 7200             | 56 400           | 1.37 | 0.63   |
| 12    | 12   | 500 : 1 : 10         | 2         | 99     | 7200             | 8500             | 1.34 | 0.61   |
| 13\f | 13\f | 100 : 1 : 1          | 15        | 99     | 14 400           | 9500             | 1.16 | 0.71   |
| 14    | 14\f | 200 : 1 : 1          | 5         | 97     | 28 100           | 12 500           | 1.17 | 0.70   |
| 15    | 15\f | 300 : 1 : 1          | 10        | 99     | 42 900           | 21 000           | 1.28 | 0.69   |
| 16    | 16\f | 400 : 1 : 1          | 15        | 96     | 55 500           | 25 800           | 1.25 | 0.69   |
| 17    | 17\f | 500 : 1 : 1          | 30        | 89     | 64 200           | 29 400           | 1.18 | 0.66   |
| 18    | 18\f | 500 : 1 : 10         | 3         | 99     | 7200             | 8100             | 1.37 | 0.70   |
| 19\f | 19\f | 100 : 1 : 1          | 30        | 19     | 2800             | 400              | 1.21 | —      |
| 20\f | 20\f | 100 : 1 : 1          | 30        | 15     | 2300             | 400              | 1.19 | —      |
| 21\f | 21\f | 100 : 1 : 1          | 20        | 97     | 14 100           | 5600             | 1.06 | 0.75   |

\textsuperscript{a} [Cat.0] = 5 mmol L\textsuperscript{-1}, reactions were conducted in 4 cm\textsuperscript{3} solvent under 25 °C. \textsuperscript{b} rac-LA conversion was determined by \textsuperscript{1}H NMR spectra. \textsuperscript{c} M_{n,calc} = 144.13 \times [LA]_0/[BnOH]_0 \times conv. (%) + 108.14. \textsuperscript{d} Obtained from GPC analysis and calibrated against the polystyrene standard, multiplied by 0.58.\textsuperscript{165} \textsuperscript{e} Determined by analysis of all of the tetrads in the methine region of the homonuclear-decoupled \textsuperscript{1}H NMR spectra. \textsuperscript{f} Polymerization reaction was run at 0 °C. \textsuperscript{g} Polymerization reaction was carried out in CH2Cl2. \textsuperscript{h} Polymerization reaction was carried out in THF.
experimental result, we speculate that the phenoxide ligands might coordinate to alkali metal center in a O,N-chelate mode during the process of catalysis, which improved the catalytic selectivity of the complexes through restricting rotation of the M–Oaryl single bond. A less sterically hindered phenoxo group and N-aryl group would benefit formation of O,N-chelate complexes.

We also noticed that the $P_m$ values of the polymers catalyzed by 1/BnOH slightly increased with the increase of the ratio of rac-LA to catalyst (Table 1, entries 1 and 8–11). However, the $P_m$ values of the polymers catalyzed by 4/BnOH were approximately unchanged when the ratio of monomer to catalyst varied from 100 to 400 and slightly decreased when the ratio of monomer to catalyst was 500 (Table 1, entries 4 and 14–17). Reduction of polymerization temperature led to increase of selectivity whether the catalyst was 1 or 4. The $P_m$ value of the PLA was 0.71 when 100 equiv. of rac-LA was catalytically polymerized by 1/BnOH at 0°C and the $P_m$ value of the PLA reached 0.75 when 100 equiv. of rac-LA was catalytically polymerized by 4/BnOH at 0°C (Table 1, entries 13 and 21). The end group analysis of the PLA proved that the polymer chain was capped with one benzyl ester end and one hydroxyl end (Fig. 7), which are consistent with an insertion mechanism of a benzyloxy group into the lactide as supposed for most alkali metal phenoxides.94–100
Synthesis of complex 1
To a stirred solution of HL$^2$ (0.28 g, 1.00 mmol) and 15-crown-5 (0.22 g, 1.00 mmol) in toluene (10 cm$^3$) was slowly added Na[N(SiMe)$_3$]$_2$ (0.55 cm$^3$, 2 M solution in THF, 1.1 mmol) at room temperature. The mixture was stirred for 10 h to generate a yellow solution. The resulting solution was concentrated in vacuo. n-Hexane (20 cm$^3$) was added to the solution to afford complex 1 as a yellow solid (0.30 g, 57%). $^1$H NMR (400 MHz, DMSO-d$_6$): δ 8.44 (s, 1H), 7.59 (d, J = 7.7, 2.0 Hz, 1H), 7.04 (d, J = 7.6 Hz, 2H), 6.98–6.91 (m, 1H), 6.89–6.81 (m, 1H), 6.17 (d, J = 8.4 Hz, 1H), 5.93 (t, J = 7.2 Hz, 1H), 3.56 (s, 20H), 3.02–2.91 (m, 2H), 1.09 (d, J = 6.9 Hz, 12H). $^{13}$C($^1$H) NMR (101 MHz, DMSO-d$_6$): δ 174.06, 161.93, 152.06, 137.66, 132.25, 126.78, 123.05, 122.87, 122.28, 122.18, 106.98, 69.34, 27.16, 23.39. Anal. calcd for C$_{37}$H$_{58}$NNaO$_6$: C 66.52, H 8.08, N 2.67; found: C 66.23, H 8.22, N 2.70.

Synthesis of complex 2
To a stirred solution of HL$^3$ (0.30 g, 1.00 mmol) and 15-crown-5 (0.22 g, 1.00 mmol) in toluene (10 cm$^3$) was slowly added Na[N(SiMe)$_3$]$_2$ (0.55 cm$^3$, 2 M solution in THF, 1.1 mmol) at room temperature. The mixture was stirred for 10 h and yellow precipitates were formed. The precipitates were collected by filtration, washed with toluene, and dried in vacuo to give yellow solids (0.25 g, 45%). $^1$H NMR (400 MHz, DMSO-d$_6$): δ 8.79 (s, 1H), 7.38 (d, J = 2.9 Hz, 1H), 7.31–7.24 (m, 2H), 7.07–6.97 (m, 3H), 6.95 (d, J = 2.9 Hz, 1H), 3.55 (s, 20H), 1.33 (s, 9H), 1.19 (s, 9H). $^{13}$C($^1$H) NMR (101 MHz, DMSO-d$_6$): δ 173.23, 162.37, 155.59, 139.89, 128.71, 126.24, 126.04, 125.82, 120.82, 120.28, 69.46, 34.90, 33.37, 31.79, 29.52. Anal. calcd for C$_{29}$H$_{42}$NNaO$_6$: C 66.19, H 8.20, N 2.54. Found: C 66.23, H 8.22, N 2.70.

Synthesis of complex 3
To a stirred solution of HL$^4$ (0.39 g, 1.00 mmol) and 15-crown-5 (0.22 g, 1.00 mmol) in toluene (10 cm$^3$) was slowly added Na[N(SiMe)$_3$]$_2$ (0.55 cm$^3$, 2 M solution in THF, 1.1 mmol) at room temperature. The mixture was stirred for 10 h and yellow precipitates were formed. The precipitates were collected by filtration, washed with toluene, and dried in vacuo to give yellow solids (0.32 g, 50%). $^1$H NMR (400 MHz, DMSO-d$_6$): δ 8.46 (s, 1H), 7.50 (d, J = 2.7 Hz, 1H), 7.02 (d, J = 7.6 Hz, 2H), 6.96 (d, J = 2.8 Hz, 1H), 6.90 (t, J = 7.6 Hz, 1H), 3.54 (s, 20H), 3.05–2.93 (m, 2H), 1.33 (s, 9H), 1.21 (s, 9H), 1.09 (d, J = 6.9 Hz, 12H). $^{13}$C($^1$H) NMR (101 MHz, DMSO-d$_6$): δ 172.41, 163.01, 152.59, 139.42, 137.74, 125.98, 125.67, 122.13, 121.69, 120.92, 119.87, 69.20, 34.87, 33.33, 31.86, 29.64, 27.09, 23.40. Anal. calcd for C$_{37}$H$_{58}$NNaO$_6$: C 67.49, H 8.40, N 2.54. Found: C 67.02, H 8.57, N 2.66.

Conclusions
In summary, we have synthesized and characterized a series of sodium and potassium iminophenoxides bearing 15-crown-5 or 18-crown-6 as the auxiliary ligand. In the presence of benzyl alcohol, all the complexes can efficiently catalyze the ROP of rac-lactide. The catalysts with bulky substituents on the aromatic rings of phenoxy and N-aryl groups are unfavored for the selectivity of the polymers. The potassium complex without any substituents except imino group on the aromatic rings, [(18-crown-6)KOC$_6$H$_4$[2-(PhN=CH)]] (4), exhibited the best selectivity of isotacticity ($P_{\text{isotactic}} = 0.75$) when the ROP was carried out at 0 °C. During the catalytic process, the iminophenoxide ligands might coordinate to the alkali metal center in a Q$_3$N-chelate mode which improved the catalytic selectivity of the complexes through restricting rotation of the M–O$_2$aryl single bond.

Experimental
All air or moisture sensitive manipulations were performed under dry N$_2$ using standard Schlenk techniques. Solvents were distilled under nitrogen over sodium (toluene), sodium/benzophenone (n-hexane and tetrahydrofuran), or CaH$_2$ (dichloromethane) and degassed prior to use. NaN(SiMe$_3$)$_2$, KN(SiMe$_3$)$_2$ and other reagents were purchased from local chemical companies and used as received. BnOH was distilled from CaO. rac-Lactide was purchased from Daigang Biomaterial Co. and recrystallized three times from toluene prior to use. DMSO-d$_6$, purchased from EMD Millipore Corporation, was degassed, and stored over 4 Å molecular sieves. CDCl$_3$ was purchased from Cambridge Isotope Laboratories and used as received. NMR spectra were recorded on a Bruker AVANCE III 400 spectrometer at ambient temperature. The chemical shifts of $^1$H and $^{13}$C($^1$H) NMR spectra were referenced to internal solvent resonances. Elemental analyses were performed on an Elementar Vario EL III analyzer. Gel permeation chromatography (GPC) was recorded on a Waters 150C instrument equipped with UltraStyragel columns (103, 104, and 105 Å) and a 410 refractive index detector, using monodispersed polystyrene as the calibration standard. THF (HPLC grade) was used as an eluent at a flow rate of 1 cm$^3$ min$^{-1}$.

Fig. 7 The $^1$H NMR spectrum of PLA initiated by 4/BnOH (Table 1, entry 4) (NMR: 400 MHz, solvent: CDCl$_3$, reference: TMS).
Synthesis of complex 4
To a stirred solution of HL^2 (0.19 g, 1.00 mmol) and 18-crown-6 (0.27 g, 1.00 mmol) in toluene (10 cm^3) was slowly added KN(SiMe_3)_2 (1.1 cm^3, 1 M solution in THF, 1.1 mmol) at room temperature. The resultant mixture was stirred at room temperature for 10 h and yellow precipitates were formed. The precipitates were collected by filtration, washed with toluene, and dried in vacuo to give yellow solids (0.21 g, 42%). \(^1^H\) NMR (400 MHz, DMSO-d_6): \(\delta\) 8.48 (s, 1H), 7.51 (dd, \(J = 7.8, 2.1\) Hz, 1H), 7.29 (t, \(J = 7.8\) Hz, 2H), 7.09–6.96 (m, 3H), 6.83–6.75 (m, 1H), 6.69 (dd, \(J = 8.6, 0.7\) Hz, 1H), 5.82 (t, \(J = 7.2\) Hz, 1H), 3.53 (s, 24H). \(^{13}\text{C}[^1\text{H}]\) NMR (101 MHz, DMSO-d_6): \(\delta\) 174.88, 161.91, 152.26, 137.68, 119.63, 69.41, 34.87, 33.30, 31.86, 29.57, 27.09, 23.37. Anal. calcd for C_{39}H_{62}KNO_7: C 67.30, H 8.98, N 2.01; found: C 67.30, H 8.87, N 2.06.

Synthesis of complex 5
To a stirred solution of HL^2 (0.28 g, 1.00 mmol) and 18-crown-6 (0.27 g, 1.00 mmol) in toluene (10 cm^3) was slowly added KN(SiMe_3)_2 (1.1 cm^3, 1 M solution in THF, 1.1 mmol) at room temperature. The resultant mixture was stirred at room temperature for 10 h to generate a yellow solution. The solution was concentrated in vacuo. Then n-hexane (20 cm^3) was added to the concentrated solution to afford complex 5 as a yellow solid (0.28 g, 48%). \(^1^H\) NMR (400 MHz, DMSO-d_6): \(\delta\) 8.40 (s, 1H), 7.56 (dd, \(J = 7.6, 1.9\) Hz, 1H), 7.02 (d, \(J = 7.6\) Hz, 2H), 6.96–6.87 (m, 1H), 6.83–6.75 (m, 1H), 6.05 (dd, \(J = 8.6\) Hz, 1H), 5.82 (t, \(J = 7.1\) Hz, 1H), 3.54 (s, 24H), 3.04–2.89 (m, 2H), 1.07 (d, \(J = 6.9\) Hz, 12H). \(^{13}\text{C}[^1\text{H}]\) NMR (101 MHz, DMSO-d_6): \(\delta\) 173.51, 162.32, 155.71, 139.94, 128.69, 126.00, 125.93, 122.37, 120.76, 119.88, 69.42, 34.89, 33.36, 31.79, 29.48. Anal. calcd for C_{33}H_{50}KNO_7: C 61.15, H 6.68, N 2.96. RT. The resultant mixture was stirred at room temperature for 10 h and yellow precipitates were formed. The precipitates were collected by filtration, washed with toluene, and dried in vacuo to give yellow solids (0.32 g, 46%). \(^1^H\) NMR (400 MHz, DMSO-d_6): \(\delta\) 8.48 (s, 1H), 7.51 (s, 1H), 7.02 (d, \(J = 7.6\) Hz, 2H), 6.95 (s, 1H), 6.90 (t, \(J = 7.6\) Hz, 1H), 3.52 (s, 24H), 3.07–2.94 (m, 2H), 1.34 (s, 9H), 1.21 (s, 9H), 1.10 (d, \(J = 6.8\) Hz, 12H). \(^{13}\text{C}[^1\text{H}]\) NMR (101 MHz, DMSO-d_6): \(\delta\) 172.81, 162.99, 152.70, 139.49, 137.72, 125.59, 125.46, 122.10, 121.59, 120.77, 119.63, 69.41, 34.87, 33.30, 31.86, 29.57, 27.09, 23.37. Anal. calcd for C_{39}H_{62}KNO_7: C 67.30, H 8.87, N 2.06.

X-ray crystallography
Single crystals of complexes 3 and 4 were respectively mounted in Lindemann capillaries under nitrogen. Diffraction data were collected at 298(2) K on a Bruker Smart CCD area detector with graphite-monochromated Mo K\(_\alpha\) radiation (\(\lambda = 0.7073\) Å). The structures were solved by direct methods using SHELXS-97 (ref. 106) and refined against F^2 by full-matrix least squares using SHELXL-97.\(^{107}\) Hydrogen atoms were placed in calculated positions. Crystal data and experimental details of the structure determination are listed in Table 2.

Polymerization of rac-LA
A typical polymerization procedure was exemplified using 1/BnOH as the catalyst. Complex 1 (0.104 mg, 0.02 mmol) and dry toluene (10 cm^3) were respectively mounted in Lindemann capillaries under nitrogen. Diffraction data were collected at 298(2) K on a Bruker Smart CCD area detector with graphite-monochromated Mo K\(_\alpha\) radiation (\(\lambda = 0.7073\) Å). The structures were solved by direct methods using SHELXS-97 (ref. 106) and refined against F^2 by full-matrix least squares using SHELXL-97.\(^{107}\) Hydrogen atoms were placed in calculated positions. Crystal data and experimental details of the structure determination are listed in Table 2.

**Table 2** Details of the X-ray structure determinations of complexes 3 and 4

| Complex | 3 | 4 |
|---------|---|---|
| Empirical formula | C_{33}H_{50}NaO_6 | C_{33}H_{50}KNO_7 |
| \(f_0\) | 635.83 | 499.63 |
| Crystal system | Monoclinic | Monoclinic |
| Space group | P2\(1\)/n | C2/c |
| \(a\) (Å) | 15.9289(13) | 24.8172(2) |
| \(b\) (Å) | 11.4741(9) | 17.2980(15) |
| \(c\) (Å) | 21.5944(18) | 17.5045(16) |
| \(\beta\) (deg) | 102.398(2) | 133.796(3) |
| \(V\) (Å\(^3\)) | 3854.83(5) | 5424.0(6) |
| \(Z\) | 4 | 8 |
| \(D_{calc}\) (g cm\(^{-3}\)) | 1.096 | 1.224 |
| \(F(000)\) | 1384 | 2128 |
| \(\mu\) (mm\(^{-1}\)) | 0.082 | 0.237 |
| \(\theta\) Range for data collection (deg) | 2.21 to 25.02 | 2.33 to 25.02 |
| No. of reflections collected | 19317 | 13632 |
| No. of independent reflections (\(R_{int}\)) | 6798 (0.0635) | 4778 (0.1302) |
| Restraints/params | 0/416 | 0/307 |
| Goodness of fit on \(F^2\) | 1.052 | 1.011 |
| Final R indices\(^a\) \([I > 2\sigma(I)]\) | R\(_{1}\) = 0.0492; | R\(_{1}\) = 0.0715; |
| | wR\(_{2}\) = 0.0526 | wR\(_{2}\) = 0.1323 |
| R indices (all data) | R\(_{1}\) = 0.1527; | R\(_{1}\) = 0.1586; |
| | wR\(_{2}\) = 0.0597 | wR\(_{2}\) = 0.1517 |
| Largest diff peak and hole [e Å\(^{-3}\)] | 0.189 and −0.187 | 0.618 and −0.225 |

\(^a\) \(R_1 = \sum||F_o|-|F_c||/\sum|F_o|, wR_2 = (\sum w(F_o^2 - F_c^2)^2/\sum w(F_o^2))/1/2.\)
BnOH (0.20 cm³, 0.1 M in toluene, 0.02 mmol) were added in sequence to toluene (1.8 cm³). The resultant mixture was stirred at room temperature for 10 min and then added to a stirred mixture of rac-LA (0.2883 g, 2.00 mmol) and toluene (2.0 cm³) at the same temperature. The polymerization reaction was terminated after 1 min by adding a drop of water. White precipitates were filtered under reduced pressure and washed with hexane. Drying the wet cake under vacuum gave the polymer as white solid. For GPC analysis, the sample was dissolved in THF, precipitated under reduced pressure and washed with hexane, passed through a short neutral aluminum oxide column, precipitated in methanol, and dried under vacuum.

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103 $P_m$ is the probability of meso linkages between monomer units. $[\text{mmm}] = P_m^2 + (1 - P_m)P_m/2$, $[\text{mmr}] = (1 - P_m)P_m/2$, $[\text{rmr}] = (1 - P_m)^2/2$, $[\text{mrm}] = (1 - P_m)^2 + P_m(1 - P_m)/2$.

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