Optimization and prediction of the electron–nuclear dipolar and scalar interaction in $^1$H and $^{13}$C liquid state dynamic nuclear polarization$^+$

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During the last 10–15 years, dynamic nuclear polarization (DNP) has evolved as a powerful tool for hyperpolarization of NMR and MRI nuclides. However, it is not as well appreciated that solution-state dynamic nuclear polarization is a powerful approach to study intermolecular interactions in solution. For solutions and fluids, the $^1$H nuclide is usually dominated by an Overhauser dipolar enhancement and can be significantly increased by decreasing the correlation time ($r_c$) of the substrate/nitroxide interaction by utilizing supercritical fluids (SCF CO$_2$). For molecules containing the ubiquitous $^{13}$C nuclide, the Overhauser enhancement is usually a profile of both scalar and dipolar interactions. For carbon atoms without an attached hydrogen, a dipolar enhancement usually dominates as we illustrate for $sp^2$ hybridized carbons in the fullerenes, C$_{60}$ and C$_{70}$. However, the scalar interaction is dependent on a Fermi contact interaction which does not have the magnetic field dependence inherent in the dipolar interaction. For a comprehensive range of molecular systems we show that molecules that exhibit weakly acidic complexation interaction(s) with nitroxides provide corresponding large scalar enhancements. For the first time, we report that $sp^2$ hybridized (H–C) alkene systems, for example, the phenylacetylene–nitroxide system exhibit very large scalar dominated enhancements. Finally, we demonstrate for a wide range of molecular systems that the Fermi contact interaction can be computationally predicted via electron–nuclear hyperfine coupling and correlated with experimental $^{13}$C DNP enhancements.

Introduction

In dynamic nuclear polarization, the electron–nuclear dipolar and scalar couplings can be described by the Hamiltonian,

$$\hat{H} = \hbar \hat{S} \hat{D} \hat{I} + \hbar \hat{I} \hat{A} \hat{S}$$ (1)

where $\hbar$ is the Planck constant, $\hat{I}$ is the nuclear spin angular momentum operator, $\hat{S}$ is the electronic spin angular momentum operator, $\hat{D}$ is the dipolar coupling tensor, and $\hat{A}$ is the hyperfine coupling tensor (typically in units of MHz). One can compute the dipolar coupling tensor $\hat{D}$; however, knowledge of the electron–nuclear dipole–dipole correlation function is required to determine the enhancement factor.$^7$ Computation of this dipole–dipole correlation function can be achieved with a molecular dynamics simulation. However, such simulations are typically limited to select parameterized systems, and are very computationally intensive.$^{1,3}$

For liquid-state DNP, there has been a resurgence of interest in ameliorating liquid-state high magnetic field DNP enhancements.$^{6,10}$ Liquid state Overhauser effect DNP has been applied for signal enhancement of flow imaging$^{11-14}$ as well as the characterization of local water dynamics in biological systems.$^{15-21}$ More recently, $^1$H liquid state DNP measurements of water$^{22-26}$ and small organic molecules$^{27-29}$ at higher magnetic fields have been reported. However, liquid state DNP experiments also provide an excellent approach for understanding and predicting intermolecular solution interactions. The drawbacks for obtaining significant DNP enhancements ($A$) in the liquid state are well recognized, since the time-dependent electron–nuclear interaction dominates the Overhauser effect (Oe).$^{30-35}$ This can be derived from eqn (2)–(4), where $\gamma_S$ and $\gamma_I$ are the magnetogyric ratios for the electron and nuclear spins, respectively, and the coupling ($\rho$) and leakage ($f$) factors are
dipolar interaction dominates at low magnetic fields and/or short correlation times ($\tau_c$) in the extreme narrowing limit $\omega_S^2\tau_c^2 \ll 1$ ($\omega_S$ electron Larmor frequency). Unfortunately, at the high magnetic fields (3–12 T) commonly employed in NMR, the coupling factor approaches zero (shown in Fig. 2).

$$\rho = \frac{W_0^D - W_0^{Sc} - W_0^{10}}{W_0^D + W_0^{10} + W_0^{Sc} + 2W_0^{10}}$$

Scalar relaxation originates from the coupling of the magnetic moments of electrons and nuclei via the hyperfine coupling interaction. 34 Although $A$ in eqn (1) is the isotropic Fermi contact hyperfine coupling tensor, henceforth it is labelled as $a_{EC}$ in order to avoid confusion with the DNP enhancement, *A wide infra*.

For hydrogen containing molecules, $^1$H DNP enhancements are usually dominated by the dipolar relaxation mechanism. If one assumes a rotationally modulated mechanism dominates the electron–nuclear dipolar interaction, the corresponding rotational correlation time ($\tau_r$) can be estimated34,35,37 (see ESI† for detailed derivation). The corresponding DNP coupling factor modulated by exclusive rotational diffusion gives the magnetic

![Image](https://example.com/image.png)

**Fig. 2** Coupling factor ($\rho$) of the Overhauser effect enhancement depends on the electron Larmor frequency for rotational (solid line) and translational (dash line) motion with correlation time 10 ps (blue) and 42 ps (red); mixed with scalar enhancement of $\beta = 1$ (blue solid for $\tau_r = 10$ ps and red solid for $\tau_r = 42$ ps), and $\beta = 0.1$ (blue dash for $\tau_r = 10$ ps and red dash for $\tau_r = 42$ ps), where $\beta$ values define different contributions from dipolar relaxation competing with pure scalar relaxation. 34 Pure scalar enhancement (black solid line) is independent of magnetic field strength. Points A–E represent coupling factor values calculated from extrapolated experimental LLIT DNP enhancements polarized at 0.33 T: (A) $^1$H of 9% benzene/0.001 M TEMPO/SF CO$_2$; 33 (B) $^1$H of chloroform/0.0015 M TEMPO; 34 (C) $^1$H of 9% benzene/0.01 M TEMPO/C$_6$D$_6$; 33 (D) $^{13}$C-$\beta$ of phenylacetylene/0.01 M TEMPO; 35 and (E) $^{13}$C of chloroform/0.0052 M TEMPO. 34 The relative standard deviation for $\rho$ is about 15%. The two correlation times were estimated from $^1$H DNP coupling factors of benzene in SF CO$_2$ and normal solvent by a rotational diffusion modulated dipolar interaction mechanism (see ESI† for derivation).
field dependence shown in Fig. 2. More recently, it has been reported that translational diffusion predominantly modulates the dipolar interaction for water \(^4\text{,}^{5,6,29}\) and some small organic molecular systems \(^25,29\) with organic radicals. Using the same correlation times, the field dependence of the DNP coupling factor modulated by translational diffusion \(^11,35,41\) is also represented by colored dashed lines in Fig. 2. One notable exception for the \(^1\text{H}\) nuclide is a nitroxide–trifluoroacetic acid system reported by Bates where the scalar interaction dominates the \(^1\text{H}\) DNP enhancement.\(^{42}\)

On the other hand, for the \(^1\text{H}\) nuclide, the coupling interaction can be dominated by a strong scalar interaction \((W_{\text{SC}}^{90} \gg W_{\text{DP}}^{90}, W_{\text{DP}}^{90})\) or a modest dipolar interaction \((W_{\text{SC}}^{90} \ll W_{\text{DP}}^{90}, W_{\text{DP}}^{90})\), or by a mixture of scalar and dipolar interactions. As illustrated in Fig. 1, it is well recognized that molecules with a (C–H) group can exhibit weak hydrogen bonding with nitroxides and show modest to large scalar enhancements.\(^{25,44,46-49}\) Griffin and coworkers have pointed out the importance of the scalar interaction for certain other nuclides \((^{13}\text{F}, ^{15}\text{N}, ^{31}\text{P})\) besides \(^1\text{C}\) for high-field liquid-state DNP studies.\(^{50}\) Unfortunately, it is difficult to estimate the corresponding scalar correlation time \(\tau_{\text{SC}}\). For the case of acetone/TEMPO, we previously reported \(^1\text{C}\) DNP \(\tau_{\text{SC}}\) values 2–3 times longer than the dipolar correlation time \(\tau_{\text{DP}}\).\(^{49}\)

As first reported by our laboratory, the solid–liquid intermolecular transfer (SLIT) DNP experiment has certain advantages when the radical is not present in the solution.\(^{51,52}\) In the SLIT DNP experiment, \(^1\text{C}\) contact shifts and spectral line broadening are avoided in the high field NMR detection magnet. The SLIT approach also has the advantage of improved low to high magnetic field transfer efficiencies allowing the transfer of radical free hyperpolarized metabolites for biological uses.\(^{33,54}\) The use of immobilized radicals for DNP approaches has recently been reconsidered and utilized to generate hyperpolarized water for clinical MRI applications.\(^{11,33,55,56}\) The efficiency of interactions between the radical and nuclei in immobilized radical flow methods can be improved through better design of immobilizing methods,\(^{11,57}\) as well as the optimum synthesis of new radicals.\(^{58-60}\)

In the current study, we demonstrate that \(^1\text{H}\) LLIT (liquid–liquid intermolecular transfer) and \(^1\text{H}\) SLIT DNP with solutes dissolved in supercritical fluid (SF) \(^1\text{CO}_2\) result in enhanced electron–nuclear dipolar interactions and exhibit significant Oe DNP enhancements as a result of reduced correlation times. This is the first report of SF DNP with supercritical \(^1\text{CO}_2\) except for the earlier \(^1\text{H}\) DNP enhancement of supercritical ethylene by Wind and coworkers.\(^{41}\) It is well recognized that solutes dissolved in supercritical fluids exhibit significantly shorter molecular correlation times.\(^{42}\)

For the case of the \(^1\text{C}\) nuclide, we present dipolar dominated \(^1\text{C}\) DNP enhancements for the fullerene \(^{70}\) with a trend attributable to the difference in the intermolecular distance for electron–nuclear interaction. In a large number of molecules with acidic C–H groups, significant scalar interactions and corresponding large positive enhancements are observed in the nitroxide–substrate complex. For example, we show that sp hybridized (H–C) alkyne systems, such as, the phenylacetylene–nitroxide system exhibit very large scalar dominated enhancements. It is also possible to directly compare liquid–liquid intermolecular scalar (Fermi contact) hyperfine couplings \((a_{\text{FC}})\) with experimental scalar DNP data utilizing density functional theory (DFT) modeling.\(^{49}\) We will subsequently demonstrate that computational predictions of the scalar hyperfine coupling constants (HFCs) across a wide set of nitroxide–substrate systems correlate strongly with the DNP enhancements for systems dominated by the scalar interaction.

## Experimental

### Samples

The production method used to obtain the \(^{60}\) and \(^{70}\) fullerenes utilized an apparatus similar to the original Krätschmer–Huffman\(^{43}\) arc-burning method. Cored rods (6 mm diameter) packed with a mixture of graphite and amorphous carbon \((^{1}\text{C}\) labelled) were arc-burned in a fullerene generator under a dynamic He atmosphere. Typical electric-arc operating production parameters consist of 20–25 V, 95–115 A, and 180–220 Torr of helium. The fullerenes were extracted by dissolution in \(^2\text{CS}_2\) immediately after collection of the soot. Following this extraction protocol, the raw extract was filtered over a plug of glass wool. The \(^{60}\) and \(^{70}\) samples were separated chromatographically utilizing a pentabromobenzyl column, (PBB, 25 cm × 10 mm i.d., Phenomenex Co., Torrance, California) with \(^2\text{CS}_2\) as the mobile phase. A mass spectrometer for the sample confirmed that the sample was \(^1\text{C}\) labelled at a level of approximately 12%. All other reagents and solvents for the LLIT DNP experiments were obtained commercially. \(^{60}\) was purchased from the Isotic Inc, and other chemicals were purchased from Aldrich Chemical Co. and used without further purification. Sample solutions for the LLIT DNP experiments were prepared with TEMPO of varied concentrations from 0.001 to 0.14 M. The immobilized nitroxide materials used for the SLIT DNP experiments were synthesized in our lab by Rossi Gitti.\(^{44}\) The sample solutions were degassed by bubbling \(^2\text{N}_2\) gas through DNP experiments.

### Dynamic nuclear polarization instrument

The instrument used for the LLIT DNP experiments was described previously.\(^{44}\) The flow transfer DNP apparatus for normal liquids and supercritical fluid \(^1\text{CO}_2\) is shown in Fig. 4, where \(^1\text{CO}_2\) can be used as the flowing solvent. For flow \(^1\text{H}\) DNP performed under supercritical fluid conditions, \(^2\text{N}_2\) degassed neat liquid benzene containing free radical TEMPO \((0.1 \text{ M}–0.01 \text{ M})\) originates at the HPLC pump and flows at a rate of 0.1 mL min\(^{-1}\). The typical sample volumes inside the low field microwave cavity and high field NMR detector are 160 and 60 \(\mu\)L, respectively, and the volume of the transfer region is 80 \(\mu\)L. \(^1\text{CO}_2\) is pumped from a syringe pump of a Suprex SFC/200A Multipurpose Unit and equilibrated to a flow rate of 1.01 mL min\(^{-1}\) and 165 atm. The liquid sample is introduced into the \(^1\text{CO}_2\) flow, and at a total flow rate of 1.11 mL min\(^{-1}\) the sample solution enters the SFC oven where it is heated to 40 °C. The flow rate remains the liquid sample concentration ensuring...
the liquid solutes solubility and also keeps optimum condition based on transfer and residence times for the flow \( ^1H \) DNP apparatus.

The instruments for the SLIT DNP performed in normal liquids and SF \( CO_2 \) were reported in our previous studies.\textsuperscript{51–54} The immobilized nitroxide radical is contained in a flow cell and located in the 0.33 T low field cavity, where electron spins are polarized. The experimental procedure of SLIT \( ^1H \) DNP is similar to that of LLIT \( ^1H \) DNP. The rate of liquid flow containing the sample was adjusted accordingly to obtain preferred concentrations in normal liquids or SF \( CO_2 \) as the flowing solvent.

Determination of experimental parameters

The determination of the leakage factor \( f \) and the electron spin saturation factor \( s \) is described previously for the DNP flowing system.\textsuperscript{55} In ref. 44, we reported a theoretical model for determining the absolute enhancement for the low to high magnetic field transfer experiments, which is scaled according to the observed enhancement considering relaxation losses during the transfer. This model is restated in ESI.\textsuperscript{†} For most systems reported in this study, the DNP data were processed based on the method we applied previously.\textsuperscript{53,64,65} For benzaldehyde, nitrobenzene, toluene, diphenylmethane, triphenylmethane, benzonitrile, anisole, phenylamine and acetone systems, the absolute enhancements of \(^{13}C\) LLIT DNP were obtained utilizing a ratio method\textsuperscript{66–68} using the \(^{13}C\) enhancement of cyclohexane as the reference compound, with detailed procedures illustrated in the ESI.\textsuperscript{†} Enhancements of ethyl acetooacetate and diethyl malonate were determined by the ratio method using \(^{13}C\) of benzene as the reference. The DNP enhancements of \( C_70 \) were also determined using the ratio method using \(^{13}C_60 \) as the reference. Compared to the former model, the ratio method does not require extrapolation from observed enhancements at varied liquid flow rates and microwave powers. The applicability of the ratio method was carefully evaluated utilizing the adamantane/TEMPO system, where the DNP enhancement was validated by this method using different microwave power levels (Table S1†). A detailed discussion is included in the ESI.\textsuperscript{†} However, the ratio method does lead to low-to-high magnetic field transfer efficiency errors when the substrate molecules of interest contain quaternary carbons with significantly longer \( T_1 \) relaxation times. For the electron spin saturation parameter \( s \), we employed relatively high radical concentrations (0.1 M and higher) for DNP experiments in normal solvents in this work, in order to saturate a single EPR line and also have leakage factors approaching unity. For the benzene/SF \( CO_2 \) studies a low nitroxide radical concentration (0.001 M) was employed since fast Heisenberg electron–electron exchange led to a single broad EPR line for this system.

Theory: isotropic hyperfine coupling

The \(^{13}C\) hyperfine coupling constants were computed in the gas phase at the unrestricted B3LYP level of DFT employing the EPR-III basis\textsuperscript{75} for \( C, N, H, O \). We chose to use the B3LYP functional\textsuperscript{76} based on previously demonstrated good EPR performance for a nitroxide in aqueous solution.\textsuperscript{1} Hyperfine coupling constants were computed using the spin-orbit mean-field approximation SOMF(1X).\textsuperscript{77} For all substrate/TEMPO systems, Boltzmann weighted isotropic HFCs for nucleus \( j \), \( a_{iso} \), are determined from low-energy intermolecular configurations, \( c_j \), as computed from eqn (5). \( a_{iso} \), \( \Delta G^\circ \) is the isotropic HFC for nucleus \( j \) in geometrical configuration \( c_j \), \( \Delta G^\circ \) is the relative free
energy of geometrical configuration $c_i$, $R$ is the universal gas constant, and $T$ is the temperature (taken to be 298.15 K). Symmetry equivalent atoms are averaged together.

$$a_{\text{iso}} = \sum_i a_{\text{iso}}^{c_i} \left( \frac{e^{-\Delta G_i/RT}}{\sum_j e^{-\Delta G_j/RT}} \right)$$

(5)

Software

All optimization, thermochemistry, and solvation computations were accomplished using the Gaussian09 Rev D.01 suite of electronic structure programs. All EPR computations were accomplished in the gas phase using the ORCA 3.0.2 suite of electronic structure programs.

Results and discussion

LLIT $^1$H DNP experiments in supercritical fluid (SF) CO$_2$

As previously indicated, solutes dissolved in supercritical fluid (SF) CO$_2$ should exhibit significant Oe DNP enhancements as a result of reduced correlation times. The static and flow $^1$H NMR, and flow $^1$H DNP spectra for 9% benzene/0.001 M TEMPO dissolved in SF CO$_2$ are illustrated in Fig. 5 and the LLIT $^1$H DNP absolute enhancements of benzene dissolved in deuterated benzene and SF CO$_2$ are presented in Table 1.

Table 1

| System Absolute enhancement |
|-----------------------------|
| 9% benzene/C$_6$D$_6$/0.001 M TEMPO | $-120 \pm 18$ |
| 9% benzene/SF CO$_2$/0.001 M TEMPO | $-341 \pm 36$ |
| 9% benzene/C$_6$D$_6$/0.01 M TEMPO | $-146 \pm 20$ |
| 9% benzene/SF CO$_2$/0.01 M TEMPO | $-293 \pm 44$ |

The $^1$H DNP absolute enhancement of chloroform gives a coupling factor approaching the dipolar enhancement of SF CO$_2$ compared with liquid C$_6$D$_6$, which is in agreement with the decreased factor of molecular correlation times of small molecules in SF CO$_2$. Assuming a rotationally dominated mechanism, predicts correlation times of 10 and 42 ps, respectively for SF CO$_2$ and liquid C$_6$D$_6$. Furthermore, within experimental error the absolute enhancement of 9% benzene in SF CO$_2$ with 0.001 M TEMPO approaches the dipolar limiting value of $-330$ corresponding to a coupling factor $\rho = 0.5$, leading to nearly the highest possible dipolar-dominant enhancement in the $^1$H DNP experiment (point A in Fig. 2). This indicates that fast molecular motion of benzene in SF CO$_2$ satisfies the extreme narrowing limit $\omega_c^2 \tau_C^2 \ll 1$ shown as the flat region at the low magnetic field strength in Fig. 2, for the pure dipolar relaxation models at 0.33 T. Based on discussion in the introduction for cases where the enhancement is induced by mixed scalar and dipolar relaxation, it is clear that by conducting flow transfer $^1$H DNP using SF CO$_2$, both dipolar and scalar relaxation can result in substantial DNP Oe enhancements.

SLIT $^1$H DNP experiments in supercritical fluid (SF) CO$_2$

The solid–liquid intermolecular transfer (SLIT) $^1$H DNP experiment has the advantage of the radical not being present in the liquid or fluid. We have obtained SLIT $^1$H DNP results for benzene in supercritical CO$_2$ with two immobilized radicals that have been previously reported. Table 2 presents the SLIT $^1$H DNP absolute enhancements of benzene dissolved in deuterated benzene and SF CO$_2$ using the immobilized nitroxides (I)
and (II) as illustrated in Fig. 3. The absolute enhancements were obtained by estimating a transfer efficiency of ~80% (the efficiency with which the polarized sample bolus is transferred from the low to high field strength magnet). For I and II, a 3 to 4 fold increase is observed for the $^1$H SLIT DNP enhancement for SF CO$_2$, in comparison with liquid C$_6$D$_6$, which is due to an increase in molecular motion of the immobilized nitroxide/benzene complex in SF CO$_2$. The absolute DNP enhancement from nitroxide II is slightly higher ($A = -246$) in comparison with radical I ($A = -205$) presumably due to the increased mobility (shorter correlation time $\tau_c$) of the radical attached to a longer flexible chain off the surface of the silica gel for nitroxide II.

**LLIT $^{13}$C DNP experiments for molecular fullerene-nitroxide systems**

For the case of molecules containing the $^{13}$C nuclide, the Overhauser enhancement is usually a profile of both scalar and dipolar interactions. However, for molecules with quaternary sp$^3$ hybridized carbon atoms, a dipolar enhancement usually dominates. Our laboratory previously reported LLIT and SLIT $^{13}$C DNP results for the fullerene C$_{60}$/TEMPO system and observed dipolar dominated enhancements in both experiments. In contrast, the fullerene C$_{70}$ with ellipsoidal $D_{5h}$ symmetry is very different in comparison with C$_{60}$, with 5 different non-equivalent carbon sites on the fullerene cage surface with significant differences in surface curvature as illustrated in Fig. 6. Johnson and coworkers have previously unambiguously assigned the five different carbons by 2D $^{13}$C NMR techniques.

A surprising feature of the $^{13}$C DNP enhancements for the C$_{70}$/TEMPO system is the significantly higher dipolar enhancements ($A = -281$) observed at the polar cap carbons C$_1$ in comparison with the equatorial belt carbons C$_5$, $A = -160$. Moreover, there is a trend of decreasing enhancement values from the polar carbons (C$_1$) to the carbons at the equatorial position (C$_5$). The dipolar dominated $^{13}$C DNP enhancement value $A = -250$ for C$_{60}$ as well as the values for C$_{70}$ are remarkably large when compared with small aromatic molecules (e.g. benzene, $A = -220$, discussed in the next section). The relatively large dipolar $^{13}$C DNP enhancements can be partially attributed to the high $I_S$ and $D_{5h}$ symmetry of C$_{60}$ and C$_{70}$ with corresponding relatively short rotational correlation times. However, a model of mixed rotational and translational diffusion cannot be excluded for the electron-nuclear interaction of both systems. It is well recognized that $^{13}$C paramagnetic NMR shifts are a sensitive probe for the electron-nuclear hyperfine (scalar) interaction without the time-dependence noted above for dipolar interaction. The contact shifts of both fullerenes (Table S4†) indicate a very minor scalar contribution. Moreover, the $T_1$ relaxation times and corresponding similar $f$ factors for these carbon sites exclude the influence of leakage factor $f$ on the determination of absolute enhancements in the C$_{70}$ carbon sites (see Table S3† for relaxation data and leakage factor).

One possible origin of the enhancement differences of C$_{70}$ carbons is the variation of intermolecular distance between TEMPO and C$_{70}$ carbon from the polar to equatorial sites. Haddon has demonstrated the importance of expressing the curvature of a surface of sp$^2$ carbon atoms in terms of the $\pi$-orbital axis vector (POAV) to analyse local strain of fullerenes.

The local curvature of the C$_{70}$ carbons is assessed with the pyramidalization angle ($\theta_{\pi\pi} - 90$), where $\theta_{\pi\pi}$ is the angle between the carbon $\sigma$ orbital and the $\pi$ orbital as shown in Fig. 7. The curvature/POAV angle is related to the interaction distance between nitroxide radicals and the carbons of C$_{60}$ and C$_{70}$. The C$_{70}$ carbons with a larger pyramidalization angle are more accessible to the nitroxide radical in the weak intermolecular collisional complex, therefore resulting in a larger dipolar interaction. The correlation between the coupling factor

**Table 2**  SLIT $^1$H DNP enhancements polarized at 0.33 T for benzene/nitroxide in C$_6$D$_6$ and SF CO$_2$

| System                      | Absolute enhancement |
|-----------------------------|-----------------------|
| 9% benzene/C$_6$D$_6$/nitroxide (I) | $-67 \pm 15$          |
| 9% benzene/SF CO$_2$/nitroxide (I) | $-205 \pm 24$         |
| 9% benzene/C$_6$D$_6$/nitroxide (II) | $-61 \pm 17$          |
| 9% benzene/SF CO$_2$/nitroxide (II) | $-246 \pm 36$         |

**Fig. 6**  Structures of C$_{70}$, D$_{5h}$ and C$_{60}$, I$_h$, and LLIT $^{13}$C DNP enhancements polarized at 0.33 T for different carbons of C$_{70}$/C$_6$D$_6$/0.1 M TEMPO system and the carbon of C$_{60}$/C$_6$D$_6$/0.1 M TEMPO system. Carbon atoms of C$_{70}$ are numbered from the apical to the equatorial positions. The relative standard deviation for $A$ is about 15% (see Table S3†).

**Fig. 7**  The $\pi$-orbital axis vector (POAV) is the vector of the $\pi$ hybrid orbital with equal angles to the three $\sigma$ bonds at a conjugated carbon atom. The average angle is denoted $\theta_{\pi\pi}$. Figure reproduced from ref. 95.
of molecules with sp\(^3\) hybridized carbons (C–H) represented by red diamond; the pyramidalization angle of fullerene carbon(s) dependent on the hydrocarbon acidity. Actions ranging from modest dipolar to strong scalar interactions are observed in the interaction(s) of nitrooxide systems and density functional theory (DFT) computations: background

It is well recognized that molecules with C–H groups readily form weak complexes with nitrooxides, providing \(^{13}\)C DNP interactions ranging from modest dipolar to strong scalar interaction(s) dependent on the hydrocarbon acidity.\(^{44,46,50}\) For the case of molecules with sp\(^3\) hybridized carbons (C–H) that are not bound to electronegative elements, the \(^{13}\)C DNP presents modest dipolar enhancements as illustrated for the case of cyclohexane (A = −270) and adamantane (A = −254 and −209) (Table 6). It is well recognized that there is increasing acidity for the sp\(^3\) hybridized carbon (–CH\(^3\)) group in progressing from toluene (pK\(_a\) = 41.2) to diphenylmethane (pK\(_a\) = 33.4), and triphenylmethane (pK\(_a\) = 31.4). This trend results in a significant change in the DNP enhancement originating from a modest dipolar (A = −209) to scalar interaction (A = +129) for \(^{13}\)CH\(_X\) (X = 1–3) (Fig. 9). The aromatic ortho, meta and para carbons reflect nearly the same dipolar to scalar trend as the sp\(^3\) hybridized (–CH\(^3\)) group. This trend is especially relevant in view of the change in the molecular size (and corresponding correlation time) in progressing from toluene to triphenylmethane. Clearly, spin delocalization from the sp\(^3\) hybridized (–C–H\(_2\)) group to the aromatic pi system is an important factor dictating this trend. A second factor is the expected longer scalar correlation time for the nitroxide/triphenylmethane complex in comparison with the toluene complex, which results in a larger scalar component in the overall coupling factor. However, the sp\(^2\) aromatic \(ipso\) carbons exhibit dipolar enhancements (see Fig. S7†). The absolute enhancement trend in these cases could have a larger error because of significantly longer \(T_1\) relaxation times than the corresponding molecules with attached hydrogens (–C–H\(_2\) groups). On the other hand, for the case of sp\(^3\) hybridized carbons attached to electron withdrawing groups (X–C–H), these weakly acidic hydrocarbons lead to very significant scalar enhancements as represented by chloroform (HCCl\(_3\)) example with A = +2200, which is close to the scalar limit. In addition, there are numerous other examples of large scalar enhancements for electron withdrawing sp\(^3\) hybridized carbon (X–C–H\(_3\)) groups including nitromethane, diethyl malonate, ethylacetocetate, and acetonitrile as illustrated in Fig. 10 and Table 6. All of these examples have acidic (C–H) groups with pK\(_a\)s ranging from 10 to 25. Thus, sp\(^3\) hybridized carbon hydrocarbons exhibit a wide range of DNP enhancements ranging from those with pK\(_a\)s 40–50 (e.g., toluene, cyclohexane) that exhibit modest dipolar
interactions, to hydrocarbons with C–H groups exhibiting pK_a acidities between 10 and 30 with large scalar enhancements (A = +400 to +2200). Even molecules with weakly acidic (C–H) sites exhibit significant scalar enhancements, for example, the C_1 carbon in 1-chlorobutane (A = +460), when the remaining non-acidic (C–H) sites (C_2–C_4) exhibit dipolar enhancements (Table 6). The ability of nitroxide radicals to probe the weak acidity of C–H groups in more complex molecular systems is a unique application of solution state DNP.

Benzene/TEMPO

For the case of hydrocarbons with sp^2 hybridization, benzene has been an archetypal system for many previous solution state DNP studies. In addition, earlier NMR contact shift studies have clearly established a small scalar contribution to the dipolar dominated DNP enhancement for the benzene/TEMPO system, A \(\leq -220\). Electron withdrawing groups attached to sp^2 hybridized quaternary carbons dramatically increase this dipolar interaction. For example, the aromatic ipso (quaternary) sp^2 carbons \(^{13}\text{C-X}\) of monosubstituted benzenes such as benzonitrile, phenylamine, benzaldehyde, anisole, and nitrobenzene exhibit dipolar enhancements with both electron donating and electron withdrawing substituents (X) (see Fig. S5†). A similar trend is observed for sp^3 hybridized carbonyl carbon with electronegative oxygens directly attached (Fig. S5†). The highest dipolar enhancement in this trend is for the carbonyl carbon of acetone with a dipolar enhancement of A \(\leq -744\) representing a coupling factor of \(\approx -55\%\) of the dipolar limit at 0.33 T for this small molecule. For comparison, the significantly larger fullerene C_{60} molecule with sp^3 hybridized quaternary carbons exhibits a solution state enhancement of A \(\leq -250\) and does not have a significant scalar component vide supra.\(^{65}\) These results suggest that substrate molecules (100–700 Daltons) with sp^3 hybridized carbons that do not have a significant scalar contribution exhibit dipolar enhancements with TEMPO ranging from 25 to 50% of the dipolar limit at 0.33 T. As previously noted, at the lower end of this range, the benzene/TEMPO system represents a small molecule with both dipolar and scalar \(^{13}\text{C}\) interactions. We have previously reported based on DFT calculations that the formation of weak hydrogen bonding between C–H bond of substrate molecules and nitroxides results in unpaired spin density at the substrate carbon nuclei and results in non-trivial scalar contributions to the \(^{13}\text{C}\) DNP enhancements for acetonitrile/TEMPO and acetaldehyde/TEMPO.\(^{68}\) It is of considerable interest to establish the important orientation and interaction site(s) for the electron–nuclear interaction for the case of the benzene/TEMPO system. We have obtained the Boltzmann weighted \(^{13}\text{C}\) A_{FC} = 0.80 MHz via DFT computations using eqn (5) by considering the four lowest-energy molecular configurations (see S6–S9 and Table S5† for optimized orientations and their parameters). The two important conformations with the largest contributions to the hyperfine coupling are shown in Fig. 11 and 12. It is important to note that in both conformations the nitroxy N–O bond vector is either orthogonal to one C–H bond vector or orthogonal between (T shape) two C–H bond vectors with interaction distances of 0.230 and 0.239 nm, respectively. This orthogonal approach of the N–O and C–H bond is also consistent for the lowest energy conformations found below for other monosubstituted benzene/TEMPO complexes vide infra. The computational average A_{FC} value (A_{FC} = 0.80 MHz) indicates a non-trivial scalar contribution to the \(^{13}\text{C}\) DNP NMR signal of benzene. For the comparison of benzene and monosubstituted benzenes, their comparable molecular size and lack of significant conformational differences allow the assumption of relatively small changes in the correlation times.

Nitrobenzene/TEMPO

For the case of nitrobenzene/TEMPO (NMR and DNP spectra shown by Fig. 14), the introduction of the strong electron withdrawing nitro group on the aromatic ring leads to a lowest energy conformation with an orthogonal approach of the N–O and C–H bond vectors between the ortho and meta carbons (Fig. 13). The weighted Boltzmann averaging of HFC for the ortho carbon (A_{FC} = 1.424) is in strong agreement with the large scalar enhancement, A = +587 observed for this site (Table 3). The large \(^{13}\text{C} para\) and \(^{13}\text{C} meta\) contribution appears to originate from a slightly higher energy configuration (Fig. S23†).
where the para C–H bond vector is orthogonal to the nitroxide, N–O bond vector. However, delocalization of polarized spin from the dominant ortho conformation to the meta and para sites can not be excluded. It is interesting to note in this regard that the strong electron withdrawing nitro group leads to scalar enhancements, whereas, electron donating groups such as the amino group of phenylamine leads to dipolar enhancements at all ortho, meta, and para carbon positions (Table S10†).

### Benzaldehyde/TEMPO

In contrast with benzene and nitrobenzene, the benzaldehyde/ TEMPO system provides a system where two different weak sp² hybridized C–H bonds can complex with TEMPO (Fig. 15 and Table 4). The weighted Boltzmann averaging of hyperfine couplings continues to show strong agreement with the scalar enhancements measured by DNP (see Table S7† for details). In this case the ortho aromatic C–H bond and the aldehyde, O=C–H bonds both contribute as a “dual complexation site” with TEMPO for the lowest energy conformation. In addition, this conformation is similar to the orthogonal bond vector conformations found for benzene and nitrobenzene. However, the more acidic aldehyde O=C–H site exhibits a larger hyperfine coupling (a_{FC} = 1.155) and corresponding larger scalar enhancement (A = +373) in comparison with the modest hyperfine coupling and DNP enhancement found for ortho C–H site, a_{FC} = 0.442 and A = +131, respectively. It is also interesting

### Table 3

| System               | Nucleus   | Absolute enhancement | a_{FC} \(^a\) |
|----------------------|-----------|----------------------|---------------|
| Nitrobenzene         | ¹³C-ipsoc | −572 ± 86            | 0.399         |
|                      | ¹³C-orthoc | +587 ± 88            | 1.424         |
|                      | ¹³C-metac | +198 ± 30            | 0.329         |
|                      | ¹³C-parac | +366 ± 55            | 0.418         |

\(^{a}\) a_{FC} is the Fermi contact HFC in units of MHz computed as described.

### Table 4

| System               | Nucleus   | Absolute enhancement | a_{FC} \(^a\) |
|----------------------|-----------|----------------------|---------------|
| Benzaldehyde         | ¹³CHO–    | +373 ± 56            | 1.155         |
|                      | ¹³C-ipsoc | −456 ± 20            | 0.284         |
|                      | ¹³C-orthoc| +131 ± 20            | 0.442         |
|                      | ¹³C-metac | +93 ± 14             | 0.063         |
|                      | ¹³C-parac | +78 ± 12             | 0.099         |

\(^a\) a_{FC} is the Fermi contact HFC in units of MHz computed as described.

Fig. 13 The electronic spin density distribution for the lowest energy orientation between nitrobenzene and TEMPO where ortho and meta C–H moieties complex with the TEMPO N–O group. Dashed lines are given as guides for intermolecular distances, which are 2.45 Å for O–H (of ortho C–H) and 2.53 Å for N–H (of para C–H).

Fig. 14 ¹³C spectra of benzaldehyde/TEMPO in cyclohexane at 4.7 T (a) 0.1 M TEMPO without flow DNP, (b) 0.1 M TEMPO with flow DNP. (a) 3000 scans; (b) 735 scans.\(^{67}\)

Fig. 15 The electronic spin density distribution for the largest ¹³CHO– and ¹³C-ortho HFC contributing orientation between benzaldehyde and TEMPO where the aldehyde and ortho C–H moieties bite the N–O group. Dashed lines are given as guides for intermolecular distances, which are 2.39 Å for O–H (of ortho C–H) and 2.46 Å for N–H (of aldehyde C–H).
to note that the modest electron withdrawing aldehyde group leads to modest scalar enhancements at all ortho, meta, and para carbon positions (Table 4).

**Phenylacetylene/TEMPO**

For the case of sp hybridized carbons there is a paucity of data for the complexation of alkynes with nitroxides. For cyano sp hybridized carbons modest $^1$C dipolar DNP enhancements have been reported for benzonitrile and acetonitrile (Table 6). However, for the acidic sp hybridized C–H carbons there is a paucity of examples reported. In one notable early NMR study, a transient hydrogen bonding between phenylacetylene and di-tert-butyl nitroxide was proposed, since the former is a good proton donor with a $pK_a$ of 21 (Fig. 1). It should be noted that alkynes are key components in the widely employed “click” reaction extensively employed in organic, polymer, and biomedical functionalization studies. Phenylacetylene has an acetylenic proton of relatively high acidity ($pK_a$ of 21) compared with benzene as shown in Fig. 1.

**Table 5** LLI $^1$C DNP enhancements polarized at 0.33 T for phenylacetylene/0.001 M and 0.14 M TEMPO and $a_{FC}$ via DFT modeling

| System | Nucleus | Absolute enhancement | $a_{FC}$ |
|--------|---------|----------------------|----------|
| Phenylacetylene | $^{13}$C-$\beta$ | +1400 ± 210 | 3.281 |
| | $^{13}$C-α | +290 ± 44 | 0.871 |
| | $^{13}$C-ipso | −280 ± 42 | 0.277 |
| | $^{13}$C-ortho | −36 ± 5 | 0.579 |
| | $^{13}$C-meta | −32 ± 5 | 0.117 |
| | $^{13}$C-para | N/A | 0.181 |

$a_{FC}$ is the Fermi contact HFC in units of MHz computed as described. Inverse gated $^1$H decoupling was employed. C-$\beta$ results are based on measurements of 0.01 M TEMPO solution, while C-α and C-1–3 results are from 0.14 M TEMPO solution. C-4 signal is close to zero.

**Fig. 16** $^1$C spectra of phenylacetylene/TEMPO at 4.7 T (a) 0.01 M TEMPO without flow DNP, (b) 0.01 M TEMPO with flow DNP and (c) 0.14 M TEMPO. (a) 16 scans; (b) 16 scans; (c) 64 scans.

**Fig. 17** The electronic spin density distribution of the largest $^{13}$C-$\beta$ HFC contributing orientation between phenylacetylene and TEMPO where the alkyne points at N–O group. Dashed line is given as guide for intermolecular distance, which is 2.03 Å between oxygen atom of TEMPO and the hydrogen atom of acetylene group of phenylacetylene.
Clearly there is a large difference in the spin–lattice $T_1$ relaxation time of the scalar dominated β carbon relative to the α, μ, and ρ carbons of the aromatic ring. At low radical (TEMPO) concentrations (0.01 M), the β carbon polarization is notably enhanced, but the other carbon atoms are barely observable. For the ring carbons having dipolar-dominated interactions, whereas, at higher radical concentrations (0.14 M) dipolar enhancements become observable, but the β carbon is barely observed because of the very short nuclear relaxation times at this concentration which limits the transfer efficiency from the low (0.33 T) to the high observable magnetic field (4.7 T). However, the lower radical concentration used (0.01 M) may diminish the observed DNP enhancements for the β carbon and other carbons because of a lower leakage factor and three-spin effects\textsuperscript{100} (Fig. 16b). These results suggest that for molecules having specific strong scalar complexation sites, it is possible to independently detect enhancements contributed by the scalar and dipolar interactions by varying the radical concentration.

## Conclusions

In summary, we have found that for $^1$H DNP, supercritical fluids provide a convenient approach for decreasing the correlation time for solutes and provide an alternative approach for improving dipolar enhancements at high magnetic fields without the need for polarization at low temperatures (~4 K). For the system SF CO\textsubscript{2}/benzene/TEMPO, the DNP enhancement approaches the dipolar limiting value of $-330$ ($\rho = 0.5$) at 0.33 T.

### Table 6  LLIT $^{13}$C DNP enhancements polarized at 0.33 T for substrate/TEMPO systems and acid dissociation constants (pK\textsubscript{a})\textsuperscript{101,102} at 25 °C of $^{13}$C–H protons

| System (solute/solvent/TEMPO conc.) | Nucleus | Absolute enhancement | pK\textsubscript{a} | Type of carbon hybridization |
|-------------------------------------|---------|----------------------|-----------------|-----------------------------|
| Phenylacetylene/NA/0.01 M (ref. 45) | $^{13}$C-β | +1400 ± 210 | 21 | sp\textsuperscript{1} |
| Phenylacetylene/NA/0.01 M (ref. 45) | $^{13}$C-μ | +290 ± 44 | — | sp\textsuperscript{1} |
| Benzaldehyde/cyclohexane/0.1 M (ref. 67) | $^{13}$CHO– | +373 ± 56 | 14.9 | sp\textsuperscript{1} |
| | $^{13}$C-ortho | +131 ± 20 | — | — |
| Nitrobenzene/cyclohexane/0.1 M (ref. 66) | $^{13}$C-ortho | +587 ± 88 | — | — |
| | $^{13}$C-meta | +198 ± 30 | — | — |
| | $^{13}$C-para | +366 ± 55 | — | — |
| Chloroform/0.0052 M (ref. 44) | $^{13}$CHCl\textsubscript{3} | +2200 ± 330 | 15.5 | sp\textsuperscript{1} |
| Nitromethane/benzene/0.1 M (ref. 66) | $^{13}$CH\textsubscript{3}NO\textsubscript{2} | +996 ± 149 | 10 | — |
| Carbon tetrachloride/1-chlorobutane/0.08 M (ref. 44) | $^{13}$CCI\textsubscript{4} | +770 ± 115 | — | — |
| Acetonitrile/carbon tetrachloride/0.1 M (ref. 103) | $^{13}$CH\textsubscript{3}CN | +520 ± 100 | 25 | — |
| 1-Chlorobutane/carbon tetrachloride/0.08 M (ref. 44) | $^{13}$CH\textsubscript{2}CH\textsubscript{2}CH\textsubscript{2}Cl | +460 ± 70 | — | — |
| Ethyl acetoacetate/benzene/cyclohexane/0.1 M (ref. 66) | $^{13}$CH\textsubscript{2}CO\textsubscript{13}CH\textsubscript{2}COOCH\textsubscript{2}CH\textsubscript{3} | +440 ± 66 | 11 | — |
| Diethyl malonate/benzene/cyclohexane/0.1 M (ref. 66) | $^{13}$CH\textsubscript{2}[OOC$^{13}$CH\textsubscript{3}]\textsubscript{2} | +401 ± 60 | 13 | — |
| Triphenylmethane/cyclohexane/0.1 M (ref. 66) | $^{13}$CH(Ph)\textsubscript{3} | +129 ± 19 | 31.4 | — |
| Diphenylmethane/cyclohexane/0.1 M (ref. 66) | $^{13}$CH(Ph)\textsubscript{3} | +67 ± 10 | 33.4 | — |
| Benzene/0.1 M (ref. 66) | $^{13}$C\textsubscript{6}H\textsubscript{6} | $–$200 ± 30 | 43 | sp\textsuperscript{2} |
| C\textsubscript{6}O/C\textsubscript{2}D\textsubscript{6}/0.1 M (ref. 65) | $^{13}$C\textsubscript{6}O | $–$250 ± 20 | — | — |
| Benzonitrile/cyclohexane/0.1 M (ref. 67) | $^{13}$C-ipso | $–$281 ± 42 | — | — |
| Phenylamine/cyclohexane/0.1 M (ref. 67) | $^{13}$C-ipso | $–$318 ± 48 | — | — |
| Toluene/cyclohexane/0.1 M (ref. 66) | $^{13}$C-ipso | $–$423 ± 63 | — | — |
| Benzaldehyde/cyclohexane/0.1 M (ref. 67) | $^{13}$C-ipso | $–$456 ± 20 | — | — |
| Anisole/cyclohexane/0.1 M (ref. 67) | $^{13}$C-ipso | $–$457 ± 69 | — | — |
| Nitrobenzene/cyclohexane/0.1 M (ref. 66) | $^{13}$C-ipso | $–$572 ± 86 | — | — |
| Diethyl malonate/benzene/cyclohexane/0.1 M (ref. 66) | [CH\textsubscript{3}CH\textsubscript{2}OOC$^{13}$C\textsubscript{2}]\textsubscript{2} | $–$642 ± 96 | — | — |
| Ethyl acetoacetate/benzene/cyclohexane/0.1 M (ref. 66) | CH\textsubscript{2}COCH\textsubscript{2}$^{13}$COOCH\textsubscript{2}CH\textsubscript{3} | $–$711 ± 107 | — | — |
| | CH\textsubscript{2}$^{13}$COCH\textsubscript{2}COOCH\textsubscript{2}CH\textsubscript{3} | $–$720 ± 108 | — | — |
| Acetone/carbon tetrachloride/cyclohexane/0.1 M (ref. 66) | CH\textsubscript{3}$^{13}$CO | $–$744 ± 112 | — | — |
| Toluene/cyclohexane/0.1 M (ref. 66) | $^{13}$CH\textsubscript{Ph} | $–$209 ± 31 | 41.2 | — |
| Adamantane/benzene/0.1 M (ref. 66) | $^{13}$CH\textsubscript{12} | $–$209 ± 13 | — | — |
| | $^{13}$CH\textsubscript{13} | $–$254 ± 14 | — | — |
| Cyclohexane/NA/0.1 M (ref. 45) | $^{13}$C\textsubscript{6}H\textsubscript{12} | $–$270 ± 40 | 45 | — |

1-Chlorobutane/carbon tetrachloride/0.08 M (ref. 44) | CH\textsubscript{2}$^{13}$CH\textsubscript{2}CH\textsubscript{2}CH\textsubscript{2}Cl | $–$200 ± 30 | — | — |
T. For biomedical applications, there has been considerable interest in the DNP enhancement of water as reviewed by Günther. For example, Bennati and coworkers have experimentally reported an Oe dipolar DNP enhancement $A = -170$ for the water/TEMPO-D-15N system at 9.7 GHz (0.33 T).

For $^{13}$C DNP at 0.33 T, we have found that molecules (100–700 Daltons) dissolved in liquids containing carbon sites without a significant scalar contribution (e.g., carboxyls) exhibit dipolar enhancements ranging from, $A = -250$ to $-740$ with TEMPO *vide supra*. It is easy to predict that $^{13}$C DNP for these carbon sites could also exhibit $^{13}$C DNP enhancements 2 to 3 times greater if conducted in SF CO$_2$. For the biomedical metabolites, the enhancement in SF CO$_2$ followed by rapid dissolution in water or saline solutions and transfer to high magnetic fields could be a viable approach for NMR nuclides with long $T_1$ relaxation times for both NMR and MRI applications. Furthermore, other SF fluids, such as, supercritical water and nitrous oxide represent alternative fluids for decreasing correlation times at higher magnetic fields for DNP at high fields. For the case the C$_{60}$/TEMPO system, we have reported significantly higher dipolar enhancements at the polar cap carbons with greater sp$^3$ curvature in comparison with the equatorial belt carbons. For the case of $^{13}$C DNP where both dipolar and scalar interactions are important, the correlation between experimental scalar DNP enhancements ($A$) and the Boltzmann weighted hyperfine coupling constants ($\alpha_{\text{FC}}$) calculated using the DFT methods provides a predictive tool for any substrate systems exhibiting a non-trivial scalar interaction as illustrated in Fig. 18. A cautionary note of these predictions is the importance of the scalar and/or dipolar correlation time(s) that is not included in the current computational approach. For this work, we assume similar correlation times for the mono-substituted benzenes employed in this study. Finally, of critical importance, molecules with sp hybridized C–H groups, such as, the $\beta$ carbon of phenylacetylene deserve further study because of the large scalar enhancements observed for this system. For example, Bennati and coworkers have experimentally reported an Oe dipolar DNP enhancement $A = -170$ for the water/TEMPO-D-15N system at 9.7 GHz (0.33 T).

Fig. 18 Correlation between experimentally determined $^{13}$C LLIT DNP scalar dominated enhancements and their calculated AFC (Boltzmann weighted DFT computed Fermi contact hyperfine coupling constant) for the phenylacetylene/TEMPO, benzaldehyde/TEMPO, and nitrobenzene/TEMPO systems. After performing a linear regression with the line of best fit ($y = mx + b$) was determined having parameters $m = 415 \pm 23$, $b = 21 \pm 52$, and the goodness of the fit is computed as having an adjusted $R^2 = 0.97$.

in phenylacetylene$^{107}$ (C$_6$H$_5$C$^{13}$C$^{1}$H). In conclusion, solution state DNP provides a unique approach for studying intermolecular weak bonding interactions of solutes in normal liquids and SF fluids.

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