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Frontiers and progress in cation-uptake and exchange chemistry of polyoxometalate-based compounds

Cation-uptake and exchange in polyoxometalates (POMs) and POM-based compounds are categorized and reviewed in three groups: POMs as inorganic crown ethers and cryptands, POM-based ionic solids as cation-exchangers, and reduction-induced cation-uptake in POM-based ionic solids, which is based on a feature of POMs that they are redox-active and multi-electron transfer occurs reversibly in multiple-steps. This method can be utilized to synthesize mixed-valence metal clusters in metal ion-exchanged POM-based ionic solids.

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Cation-uptake and exchange has been an important topic in both basic and applied chemistry relevant to life and materials science. For example, living cells contain appreciable amounts of Na⁺ and K⁺, and their concentrations are regulated by the sodium–potassium pump. Solid-state cation-exchangers such as clays and zeolites both natural and synthetic have been used widely in water softening and purification, separation of metal ions and biomolecules, etc. Polyoxometalates (POMs) are robust, discrete, and structurally well-defined metal-oxide cluster anions, and have stimulated research in broad fields of sciences. In this perspective, cation-uptake and exchange in POM and POM-based compounds are categorized and reviewed in three groups: (i) POMs as inorganic crown ethers and cryptands, (ii) POM-based solid ionic solids as cation-exchangers, and (iii) reduction-induced cation-uptake in POM-based ionic solids, which is based on a feature of POMs that they are redox-active and multi-electron transfer occurs reversibly in multiple steps. This method can be utilized to synthesize mixed-valence metal clusters in metal ion-exchanged POM-based ionic solids.

1. Introduction

Cation-uptake and exchange from aqueous solutions has been an important topic in both basic and applied chemistry relevant to life and materials science. For example, living cells contain appreciable amounts of Na⁺ and K⁺, and their concentrations are regulated by the sodium–potassium pump, which exchanges three Na⁺ with two K⁺. On the other hand, it is quite difficult to achieve high selectivity towards K⁺ artificially except for 18-crown-6 ether, which is a cyclic oligomer of ethylene oxide, and binds K⁺ by using all six oxygens as donor atoms. The denticity of the polyether influences the affinity toward various ions: 15-crown-5 and 12-crown-4 show high selectivity toward Na⁺ and Li⁺, respectively. Crown ethers have been widely used for cation recognition and separation, and as phase transfer catalysts.

Solid-state cation-exchangers play an especially important role in chemistry. A classic example is zeolites, which are microporous crystalline aluminosilicates with anionic frameworks due to the substitution of Si⁴⁺ by Al³⁺. Cations such as Na⁺, K⁺, Ca²⁺, Mg²⁺, etc. loosely interact with the anionic framework via Coulomb interaction, which can be exchanged by treating the zeolite in an aqueous solution containing excess foreign cations. It is well known that the pore sizes and adsorption properties of zeolites can be controlled by the types of cations: the effective pore size of Linde Type A (LTA) zeolite with K⁺ is 3 Å (Molecular Sieves 3Å), and the pore size is increased to 4 Å by the exchange of K⁺ with smaller Na⁺ (Molecular Sieves 4Å). Zeolites can adsorb gas and vapor (CH₄, H₂O, etc.) in the microporous structure, and the amounts of adsorption in alkaline earth metal ion-exchanged and alkali metal ion-exchanged faujasite (FAU) zeolites increase with the increase in the ionic potentials z/r (z and r are the charge and radius of the ion, respectively) of the counter cations.
Recently, because of relatively facile reaction conditions, cation-exchange has also been recognized as a strategy for post-synthesis and discovery of new solid materials.\textsuperscript{2} Metal–organic frameworks (MOFs), which can be recognized as “inorganic–organic zeolites” have emerged decades ago, but cation-exchange has been reported only recently. A landmark report in this research area is the exchange of guest Mn\textsuperscript{2+} in as-synthesized Mn\textsubscript{3}[(Mn\textsubscript{4}Cl\textsubscript{3})(BTT)\textsubscript{8}(CH\textsubscript{3}OH)\textsubscript{10}]\textsubscript{2} (BTT = 1,3,5-benzenetristetrazolate) with monovalent or divalent metal ions in methanol solution, which resulted in the formation of isostructural frameworks with a large variation in H\textsubscript{2} adsorption enthalpy.\textsuperscript{10,11} Cation-exchange has also been employed with nanocrystals and nanoparticles to fine-tune their structures and functions systematically.\textsuperscript{12} For example, in vivo cation-exchange of Ag\textsuperscript{+} with Hg\textsuperscript{2+} and Zn\textsuperscript{2+} in selenide/sulfide quantum dots enhanced the specificity of tumor imaging.\textsuperscript{13}

Polyoxometalates (POMs) are robust, discrete, and structurally well-defined oxide cluster anions that are mainly composed of high-valence transition metals (such as W\textsuperscript{6+}, Mo\textsuperscript{6+}, V\textsuperscript{5+}) and have stimulated research in broad fields of sciences.\textsuperscript{14–25} For example, α-Keggin-type silicododecatungstate, which is one of the most researched and popular POM, forms according to the following equation:

\[
\text{SiO}_{4}^{4-} + 12\text{WO}_{4}^{2-} + 24\text{H}^{+} \rightarrow [\alpha\text{-SiW}_{12}\text{O}_{40}]^{4-} + 12\text{H}_{2}\text{O}. \quad (1)
\]

The oxides of high-valence transition metals dissolve at high pH as an anion (e.g., WO\textsubscript{4}\textsuperscript{2–}), condensation proceeds via loss of

![Fig. 1](https://example.com/fig1.png)

**Fig. 1** Cation-uptake and exchange in POM and POM-based compounds categorized in three groups: (i) POMs as inorganic crown ethers, (ii) POM-based ionic solids as cation-exchangers, and (iii) reduction-induced cation-uptake in POM-based ionic solids.
water and formation of M–O–M linkages with acidification, and an anionic molecular framework of twelve octahedral tungsten oxoanions surrounding a central silicate is formed. One of the most noteworthy features of POMs are that they are redox-active, and multi-electron transfer occurs reversibly in multiple-steps:16

\[ [\text{z-SiW}_{12}\text{O}_{40}]^{3+} + e^- = [\text{z-SiW}_{11}\text{W}^\text{VI}\text{O}_{40}]^{5-} \]  
\[ (-0.22 \text{ V vs. SHE in 1 M HCl(aq.)}), (2) \]

\[ [\text{z-SiW}_{11}\text{W}^\text{V}\text{O}_{40}]^{5+} + e^- = [\text{z-SiW}_{10}\text{W}^\text{V}\text{O}_{40}]^{6-} \]  
\[ (-0.42 \text{ V vs. SHE in 1 M HCl(aq.)}), (3) \]

In this perspective, cation-uptake and exchange in POM and POM-based compounds are categorized and reviewed in three groups: (i) POMs as inorganic crown ethers and cryptands, (ii) POM-based ionic solids as cation-exchangers, and (iii) reduction-induced cation-uptake in POM-based ionic solids (Fig. 1). Unique functions related to these cation-exchanged POM-based compounds are introduced, and future works arising from these functions are also discussed. For past developments on polyoxometalates as cation-exchangers, the readers are directed to a legendary review article.26

2. Polyoxometalates as inorganic crown ethers and cryptands

Crown ethers4,5 and cryptands27,28 which are a family of synthetic cyclic and polycyclic multidentate organic ligands, have attracted great interest due to their structural topologies and applications especially in selective cation-uptake. Crown ethers can strongly bind alkali and alkaline earth metal ions size-selectively with the oxygen donors in gas, solution, or solid phases. Cryptands can bind these cations using both nitrogen and oxygen donors three-dimensionally, often showing higher selectivity and binding constants.

In contrast, POMs can serve as inorganic crown ethers and cryptands:29 an early example is a cyclic POM \([\text{As}_5\text{W}_{40}\text{O}_{140}]^{2-}\), which binds alkali and alkaline earth metal ions size-selectively with the oxygen donors in gas, solution, or solid phases. Cryptands can bind these cations using both nitrogen and oxygen donors three-dimensionally, often showing higher selectivity and binding constants.

A recent work by Kortz and co-workers on a wheel-shaped K⁺-templated POM \([\text{K}^+\{\text{HAs}^\text{III}\text{W}_8\text{O}_{30}\}\{\text{WO}_4\text{H}_2\text{O}\}]_3\) \(14^{-}\) exhibits high selectivity to Rb⁺, because the size of the central cavity is relatively large for K⁺.31 Preyssler–Pope–Jeannin-type POM with a general formula of \([\text{X}^\text{n+}\text{(H}_2\text{O})\text{P}_5\text{W}_{30}\text{O}_{110}]^{15-}\) is the smallest POM with an internal cavity allowing cation-exchange in aqueous solutions.32,33 Preyssler–Pope–Jeannin-type POM possesses a flexible \(\text{W}_2\text{O}_4\) cavity and can capture various cations from Na⁺, Ca²⁺, La³⁺ to tetravalent actinides (e.g., Th⁴⁺) (Fig. 2),34 so that they have been considered as a potentially useful material for separation of nuclear wastes. DFT calculations by López, Poblet, and co-workers have revealed that encapsulation of cations with larger charge is difficult (i.e., heating is needed) because energy cost for the cation encapsulation from aqueous solution is dependent on the dehydration enthalpy of the cation.34 A more recent report by Li, Su, Wang and co-workers shows that while Preyssler–Pope–Jeannin-type POMs with phosphorous \([	ext{P}_3\text{W}_{30}\text{O}_{110}]^{15-}\) exhibit high affinity to Na⁺, those with sulfur \([\text{S}_5\text{W}_{30}\text{O}_{110}]^{10-}\) exhibit high affinity to K⁺ because of the larger internal cavity.35

Müller and co-workers synthesized a series of nanoporous POM capsules with a general formula of \([\{(\text{Mo}^\text{VI})\text{Mo}^\text{VI}\text{O}_9\}\text{Mo}^\text{V}\text{O}_4\text{O}_5\text{(ligand)}\}_{30}\text{[Li}_3\text{W}_{30}\text{O}_{110}]^{15-}\) and these capsules allow systematic studies of uptake/release of cations in aqueous solutions.36,37 The capsules possess large negative charges, and the affinity and coordination environment for cations depend on the functional groups inside the capsules. For example, various cations can coordinate to \(\text{SO}_4^{2-}\) ligands via exchange with NH₄⁺: protonated urea molecules situate close to the pore openings while Ce³⁺ situates deeply inside the capsule due to the small ionic radius (Fig. 3).36,37 The protonated urea molecules can be removed and the pores open by cation-exchange with Ca²⁺ in water.36 This system can be recognized as an artificial cell since Ca²⁺ take an important role in life science.

Mizuno and co-workers synthesized a dimeric POM \([\text{H}_6\{\text{γ-SiW}_{10}\text{O}_{32}\}_2\{\mu-\text{O}_4\}_4]^{28-}\) by dehydrative condensation of

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**Fig. 2** Preyssler–Pope–Jeannin-type POM \([\text{X}^\text{n+}(\text{H}_2\text{O})\text{P}_5\text{W}_{30}\text{O}_{110}]^{15-}\) with flexible \(\text{W}_2\text{O}_4\) cavity for cation encapsulation.35 Light green and purple polyhedra show the \([\text{WO}_4\]) and \([\text{PO}_4\]) units, respectively.

**Fig. 3** Coordination environments of protonated urea (left, black: C and yellow: N/O) and Ce³⁺ (right, red) in the nanoporous POM capsule.36,37 Blue and green polyhedra show the \([\text{MO}_6\]) and \([\text{SO}_4\]) units, respectively.
3. Polyoxometalate-based ionic solids as cation-exchangers

Ammonium salt of α-Keggin-type phosphododecamolybdate (NH₃)₄[α-PMo₁₂O₄₀] has been long investigated as a cation-exchanger in aqueous solutions according to the following equation,

\[
(\text{NH₃})₄[\alpha-\text{PMo}_{12}\text{O}_{40}] + 3\text{A}^+ \rightleftharpoons \text{A}_3[\alpha-\text{PMo}_{12}\text{O}_{40}] + 3\text{NH}_3^+. \tag{4}
\]

The following affinity was derived Cs⁺ = Tl⁺ > Rb⁺ > Ag⁺ > K⁺ > H₂O⁺ > Na⁺ > Li⁺, which is in line with the trend in hydration radius or dehydration enthalpy of the cations.³⁴,⁴⁴ This trend means that it is more facile to remove the water of hydration from Cs⁺ than Li⁺ because of the large ionic radius (i.e., low ionic potential) of Cs⁺, so that Cs⁺ can more easily enter and diffuse through the solid state structure. Besides, [α-PMo₁₂O₄₀]³⁻ shows high affinity towards Tl⁺ or Ag⁺, and the bonds between Tl⁺ or Ag⁺ and O²⁻ of the POM are supposed to have a covalent character.

The selectivity, kinetics, and capacity of cation-exchange in POM-based ionic solids are determined both by the framework geometry and the behavior of extra-framework cations. For example, Nyman and co-workers reported that Keggin-type polyoxoniobates [XNb₁₂O₄₀]³⁻ ([X = Si, Ge, P] with [TiO₂]⁹⁻ or [NbO₂]⁷⁻) bridges form one-dimensional chains, and these chains have an overall negative charge of −10 or −12.⁴⁴ Single crystal X-ray diffraction, thermogravimetry, IR, and ²⁹H MAS-NMR combined with computational studies could distinguish the states of counter cations (Na⁺ and K⁺), and the mobile extra-framework cations can be exchanged with radionuclides (Sr²⁺, Np(NpO₂)₂⁺, and Pu⁴⁺) (Fig. 1).³⁸ Unlike with conventional POMs, polyoxoniobates are stable under basic conditions, and therefore should be less likely to decompose in the highly alkaline conditions of nuclear wastes. More recently, Ueda, Sadakane, and co-workers synthesized microporous solids with α-Keggin-type POMs and Zn²⁺ or Mn²⁺ as linkers (Fig. 4).⁴⁷ These solids possess 3D-cages and channels with an aperture of ca. 8 Å that can accommodate exchangeable cations (NH₄⁺ and Na⁺). These cations can exchange with K⁺, Rb⁺, and Cs⁺ in aqueous solutions, while exchange with H⁺ and Li⁺ is insufficient,⁴⁷ which is in line with the hydration radius and dehydration enthalpy of the cations. More recently, the same group has reported the synthesis and structure of a polyoxomolybdate with a one-dimensional molecular structure.⁴⁸ NH₄⁺ as counter cations surround the molecular wire, and NH₄⁺ is selectively exchanged with Cs⁺ among alkali metal ions in water, and large alkylammonium cations can also be incorporated due to the flexible solid-state structure.⁴⁹

Acidic salts of α-Keggin-type POM [H₃[α-PW₁₂O₄₀]·nH₂O, H₄[α-SiW₁₂O₄₀]·nH₂O] have been well known as excellent acid catalysts, and partial substitution of protons with Cs⁺ in aqueous solutions, stabilizes the solid-state structure and increases the surface area.²⁹–₃₂ The cesium hydrogen salt of silicododecatungstate Cs₄H₄⁻[α-SiW₁₂O₄₀] adopts a body-centered cubic cell in analogy to the cesium salt of phosphododecatungstate Cs₄[α-PW₁₂O₄₀].²⁹ Maybe we have shown that the use of [α-SiW₁₂O₄₀]³⁻ instead of [α-PW₁₂O₄₀]³⁻ leads to the formation of POM vacancies to compensate the excess negative charge, which give rise to channels exhibiting cation-exchange of Cs⁺ with other alkali metal ions in aqueous solutions (Fig. 5a and b).⁴⁴ Amounts of cation-exchange decreased in the order of Rb⁺ > K⁺ > Na⁺ > Li⁺, which is in line with the hydration radius and dehydration enthalpy of the alkali metal ions (Fig. 5c), and elemental mapping images confirmed the uniform distribution of the exchanged cations (Fig. 5d).²⁹ Recently, Sun and co-workers showed that the cation-exchange of Cs⁺ in the cesium hydrogen salt of silicododecatungstate with Bi (BiO⁺ and BiOH⁺) having stereoeactive 6s lone pair as a dopant, leads to near-infrared photoluminescence in the important biological and telecommunication optical windows, due to the asymmetric coordination geometry of the Bi species in the microporous framework.⁵³ This result offers a new strategy for the preparation of POM-based luminescent systems via cation-exchange.

We have reported a porous organic–inorganic ionic crystal K₂[Cr₃O(OC₃)₆(4-methylpyridine)₃][α-SiW₁₂O₄₀]·nH₂O composed of [α-SiW₁₂O₄₀]³⁻ with a molecular cation...
4. Reduction-induced cation-uptake in polyoxometalate-based ionic solids

Redox property of solids is a key for selective cation-uptake and exchange relevant to material science. For example, Yoshikawa, Awaga, and co-workers have reported that \( \text{z-Keggin-type phosphododecamolybdate} \) \( [\text{z-PMo}_{12}\text{O}_{40}]^{3-} \) exhibits reversible 24-electron redox during charging/discharging due to the twelve molybdenum atoms \( \text{Mo}^{\text{V/VI}} \) coupled with Li\(^+\) uptake/release, as a component of molecular cluster battery.\(^{29}\) Therefore, it can be suggested that one of the best way to engineer redox-active porous solids would be to incorporate redox-active components. A landmark example was reported by Cronin and co-workers: a porous solid composed of silicodectungstate \( \gamma\text{-SiW}_{10}\text{O}_{40}\)^{8-} and Mn\(^{3+}\) was synthesized, and the oxidation states of Mn\(^{II/III}\) can be switched by the addition of reducing/oxidizing reagents.\(^{30}\) They have later synthesized another redox-active porous solid with cyclic POM \( [\text{P}_2\text{W}_{12}\text{O}_{40}]^{90-} \) units and Mn\(^{III/II}\) (Fig. 7), and the cation-exchange rate and capacity can be controlled by the oxidation states of Mn.\(^{31}\)

We have reported a redox-active porous ionic crystal \( [\text{Cr}_{3}\text{O}(\text{OOCH})_6(4\text{-methylpyridine})_3]^+ \) \( n\text{H}_2\text{O} \) (A = alkali metal ions) possessing one-dimensional channels, and the treatment of the crystal with reducing (ascorbic acid) or oxidizing (chlorine water) reagents results in one-electron redox of the POM \( \text{Mo}^{\text{V/VI}} \) coupled with uptake/release of alkali metal ions, and the reaction rate depended on the type of alkali metal ions.\(^{32}\) The reaction rate increased in the order of \( \text{K}^+ < \text{Rb}^+ < \text{Cs}^+ \), which is in line with the order of hydration radius and dehydration enthalpy of the cations (Fig. 1).\(^{33,34}\) This work was extended by the utilization of \( [\text{z-SiMo}_{12}\text{O}_{40}]^{24-} \), which resulted in the formation of an ionic crystal with isolated pores instead of continuous one-dimensional ones (Fig. 8).\(^{35}\) The compound selectively adsorbed Cs\(^+\) among alkali and alkaline earth metal ions \( \text{via} \) reduction of the POM in the compound with ascorbic acid, showing potential applicability as an adsorbent for radioactive Cs\(^+\) removal from environmental water. Despite the high selectivity to Cs\(^+\), there were several tasks to solve: requirement of heating (343 K) and slow adsorption kinetics (12 h to reach equilibrium). In order to solve these tasks, large-molecular size and easily reducible Wells–Dawson-type POMs \( [\text{z-P}_{2}\text{M}_{18}\text{O}_{62}]^{3-} \) \( (\text{M} = \text{Mo, W}) \) were utilized to increase the pore volume and to facilitate the reduction-induced Cs\(^+\) uptake.\(^{36}\) As expected, Cs\(^+\)-uptake capacity and rate increased largely (only 1 h to reach equilibrium) at room temperature.

Metal clusters are a topic of great interest in materials science and have found numerous applications especially in catalysis and electro-optics.\(^{36}\) Microporous compounds offer versatile scaffolds for the formation and stabilization of metal clusters from metal ions. For example, small mixed-valence silver clusters have been synthesized in zeolites by calcination of Ag\(^+\)-exchanged zeolites at high temperature: Ag\(^+\)\(^{37}\) in MFI-
zeolite is active for the selective reduction of NO by propane with O₂ and H₂, and Ag⁺ in LTA-zeolite shows on–off switching of yellow-green photoluminescence (PL) by hydration–dehydration. A landmark report on formation of metal clusters in redox-active MOFs has been carried out by reducing Pd²⁺ via electron transfer from nitrilotrisbenzoate, which is a redox-active organic linker of the porous framework. However, redox of MOFs is mostly limited to the utilization of redox-active organic ligands because redox of the metal center ion induces large change in the coordination geometry causing to collapse the porous framework. Therefore, metal clusters in MOFs have been synthesized by adding reducing reagents such as H₂, NaBH₄, DMF, etc. to the MOF comprising metal ions, and homogeneous formation and distribution of metal clusters become a problem.

As explained above, POMs can store multiple electrons in the molecular framework and have been utilized as constituents of redox-active porous frameworks. Some compounds show cooperative migration of electrons with metal ions, so called cation-coupled electron-transfer (CCET) in relation to proton-coupled electron-transfer (PCET). Quite recently, we have utilized redox-active porous ionic crystals Cs₂[Cr₃O(OOCH)₆(4-methylpyridine)₃][z-PMo₁₁MoV₄O₄₀]nH₂O (Cs₂-red) and Cs₂[Cr₃O(OOCH)₆(4-methylpyridine)₃][z-PMo₁₁MoV₄O₄₀]nH₂O (Cs₂-red) (the abbreviations Cs-ox and Cs₂-red are based on the types and numbers of counter cations and the oxidation state of POM), to form and stabilize small mixed-valence luminescent silver clusters in the one-dimensional channel (Fig. 9). According to elemental analysis of cesium and silver in the compounds by atomic absorption spectrometry (AAS), we have found that reduction-induced ion-exchange of Cs⁺ in Cs₂-red with Ag⁺ from AgNO₃(aq.), and subsequent formation of a mixed-valence luminescent silver cluster Ag₄⁺ took <1 min (eqn (5)), while the simple ion-exchange with Cs⁺ in Cs-ox with Ag⁺ from AgNO₃(aq.) took >24 h (eqn (6)):

\[
\text{Cs₂}[\text{Cr₃O(OOCH)₆(4-methylpyridine)₃}][z-\text{PMo}_{11}\text{MoV}_{4}\text{O}_{4₀}] + 2\text{Ag}^+ \rightarrow \\
\text{Ag₄⁺Ag⁰}[	ext{Cr₃O(OOCH)₆(4-methylpyridine)₃}][z-\text{PMo}_{11}\text{MoV}_{4}\text{O}_{4₀}] + 2\text{Cs}^+.
\]

(5)

\[
\text{Cs}[\text{Cr₃O(OOCH)₆(4-methylpyridine)₃}][z-\text{PMo}_{11}\text{MoV}_{4}\text{O}_{4₀}] + \text{Ag}^+ \rightarrow \\
\text{Ag⁰}[\text{Cr₃O(OOCH)₆(4-methylpyridine)₃}][z-\text{PMo}_{11}\text{MoV}_{4}\text{O}_{4₀}] + \text{Cs}^+.
\]

(6)

Fig. 7 (Upper) Molecular structure of [P₈W₄₈O₁₈₄]₄₀⁻ comprising a nanometer-size cavity. (Lower) The molecular unit is linked by Mn²⁺ resulting in a 3D-POM framework.  

Fig. 8 (a) Crystal structure of (etpyH)₂[Cr₃O(OOCH)₆(etpy)₃][z-SiMo₂O₄₀]nH₂O (etpy = 4-ethylpyridine). Each void (in yellow-brown) has a size of ca. 6.5 Å × 12.5 Å. (b) Amounts of cations incorporated by the reduction-induced method. Note that there is a color change due to the reduction of POM upon Cs⁺ uptake.
POMs with sulfur ($S^{2-}$), selenium ($Se^{2-}$), etc.\textsuperscript{73} to tune the coordination environment. (ii) POM-based ionic solids as cation-exchangers: the next target would be to explore cooperative effects of the selectively adsorbed cations and POMs,\textsuperscript{55} especially as optical materials, magnetic materials, solid catalysts, etc. (iii) Reduction-induced cation-uptake in POM-based ionic solids: while it is difficult for conventional porous compounds such as zeolites and MOFs to support the geometry change in the framework that often accompany the redox processes, POM-based solids show great potential for the multiple and reversible uptake/release of cations with electrons.\textsuperscript{66–73} Such CCET reactions in solids can be applied not only to selective cation-uptake and sensing but also to the next-generation rechargeable batteries,\textsuperscript{78} solid catalysts for water splitting, chemical fixation of CO\textsubscript{2}, ammonia synthesis, etc.

Another challenge is anion-exchange in POM-based compounds. This notion includes substitution of O\textsuperscript{2-} in POMs with S\textsuperscript{2-}, Se\textsuperscript{2-}, N\textsuperscript{3-} or halide ions as well as incorporation of multiple types of anions in the ionic solid. A recent review on metal oxyfluorides and oxynitrides shows that incorporation of multiple anions in metal oxide-based compounds can finely modulate physicochemical properties such as catalysis, optics, conduction, magnetism, etc.\textsuperscript{77} Some MOFs with cationic frameworks show anion-exchange properties,\textsuperscript{78} according to the Hofmeister series\textsuperscript{79} (citrate (trivalent) > sulfate (divalent) > acetate (monovalent) > HCO\textsubscript{3}- > Cl\textsuperscript{-} > Br\textsuperscript{-} > I\textsuperscript{-} > NO\textsubscript{3}-, which is in line with the degree of hydration) or non-Hofmeister selectivity due to the utilization of Lewis acid and/or multidentate donors.\textsuperscript{80} We have recently reported the synthesis of cesium salts of $\alpha$-Keggin-type [$\alpha$-BW\textsubscript{12}O\textsubscript{40}]- (BW) and [$\alpha$-SiW\textsubscript{12}O\textsubscript{40}]\textsuperscript{4-} (SIW) blends, and the porosity is finely controlled by the BW/SIW ratio.\textsuperscript{81} The next aim would be to synthesize these mixed-POM compounds post-synthetically or by anion-exchange.

### Conflicts of interest

There are no conflicts to declare.

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