Abstract: Chemical reactions that lead to a spontaneous symmetry breaking or amplification of the enantiomeric excess are of fundamental interest in explaining the formation of a homochiral world. An outstanding example is Soai’s asymmetric autocatalysis, in which small enantiomeric excesses of the added product alcohol are amplified in the reaction of diisopropylzinc and pyrimidine-5-carbaldehydes. The exact mechanism is still in dispute due to complex reaction equilibria and elusive intermediates. In situ high-resolution mass spectrometric measurements, detailed kinetic analyses and doping with in situ reacting reaction mixtures show the transient formation of hemiacetal complexes, which can establish an autocatalytic cycle. We propose a mechanism that explains the autocatalytic amplification involving these hemiacetal complexes. Comprehensive kinetic experiments and modelling of the hemiacetal formation and the Soai reaction allow the precise prediction of the reaction progress, the enantiomeric excess as well as the enantiomeric excess dependent time shift in the induction period. Experimental structural data give insights into the privileged properties of the pyrimidyl units and the formation of diastereomeric structures leading to an efficient amplification of even minimal enantiomeric excesses, respectively.

Introduction

Autocatalysis and in particular self-amplifying chemical processes are of great interest as they provide an explanation for the efficient replication of molecules with intrinsic error correction in general,[1] and the appearance of mirror image molecules with the same handedness, namely homochirality. Such processes are of fundamental importance in symmetry breaking[2] related to the emergence of life.[3] In an asymmetric reaction or catalysis, it is usually expected that the enantiomeric excess of the reagent or catalyst used will be transferred linearly to the product formation. Positive nonlinear effects,[4] which means that the use of only enantiomerically enriched reagents or catalysts lead to a significant increase in the enantiomeric excess in the product, are rarely observed. Mechanistic explanations for such reactions with positive nonlinear effects were discussed by Kagan[5] and Noyori,[6,7] considering reversible monomer-dimer associations. Frank[8] postulated a theoretical model leading to a spontaneous asymmetric synthesis. If dimers can be formed from their monomeric building blocks, for example, by intermolecular interactions, they are of the same configuration (homochiral) or of opposite configuration (heterochiral). Since these homochiral and heterochiral dimers are diastereomeric to each other, they have different intrinsic properties, which are reflected in their solubility, rate of formation and other physical properties. The formation of heterochiral dimers from an enantiomerically enriched mixture can increase the enantiomeric excess of the free monomeric major enantiomer.[9]

In 1995, Soai[10,11] reported an extremely remarkable reaction. When pyrimidine-5-carbaldehyde 4 reacts with diisopropylzinc (iPr2Zn) in the presence of a catalytic amount of pyrimidyl alcohol 1 with a low ee, asymmetric autocatalytic amplification of the enantiomeric excess gives the pyrimidine alcohol 1 with high ee as the final product (Scheme 1a). Autocatalysis and amplification are also observed in pyridyl-3-carbaldehydes,[12] however the 2-alkynyl substituted pyrimidine analogues are
Numerous findings and reports contributed to the mechanistic understanding of the Soai reaction.\cite{2} Still, open questions remain, especially regarding (a) the origin of enantioselectivity, (b) kinetic aspects of the reaction, that is, the reliable prediction of the chiral amplification, and (c) the privileged structure of the pyrimidyl-5-carbaldehydes and corresponding alcohols. While the last point (c) can be explained by experimental findings of similar reactions with the possible coordination of the nitrogen containing pyridyl or pyrimidyl rings and the associated activation of the alkyl zinc compounds as well as the formation of supramolecular structures,\cite{3} points a) and b) are not that obvious. The elucidation of the mechanism is highly challenging due to complex reaction equilibria and elusive intermediates.\cite{4,5,6,7} It is well established that isopropylzinc pyrimidyl alkoxides\cite{8} can form dimers, tetramers\cite{9,10} and oligomers.\cite{10} The dimers can be either homochiral ((R,R)-3 or (S,S)-3) or heterochiral ((R,S)-3) (Scheme 1 b).\cite{11,12} It is important to note, that the interconversion of the dimers 3 can be performed by direct exchange of the monomeric moieties without formation of the monomers.\cite{13} The implication of this equilibrium is, that the equilibrium constant for the heterochiral dimer formation $K_{	ext{hetero}}$ is twice the equilibrium constant for the homochiral dimer formation $K_{	ext{homo}}$.\cite{13} Thus, nonlinear effect can be well explained, because an imbalance of the enantiomers leads to amplification as soon as more stable heterochiral dimers (R,S)-3 have formed. Blackmond and Brown developed a model considering dimers 3 as catalytically active species based on reaction progress analysis by calorimetric measurements and NMR spectroscopy. These dimers, tetramers\cite{14,15} and oligomers\cite{16,17} were characterized by comprehensive NMR spectroscopic measurements and single-crystal X-ray diffraction analysis,\cite{18} and these findings are supported by quantum chemical computations.\cite{19} Kinetic studies corroborate these results with pronounced effects of the additive concentration\cite{20} and ee leading to an induction period and a sigmoidal kinetic profile typical for autocatalytic processes.\cite{21} Schiaffino performed quantum chemical calculations of the kinetic constants at the M05-2X/6-31G(d) level of theory and investigated the effect of the aza group in the pyrimidine moiety to activate the zinc reagent in the tetrameric complex.\cite{22}

In 2012 Brown, Blackmond and co-workers\cite{23} reported the identification of a transient hemiacetal intermediate in the Soai reaction of 2-(adamantylacetylene-1-yl)pyrimidine-5-carbaldehyde and 2-(adamantylacetylene-1-yl)pyrimidyl alcohol by $^1$H NMR spectroscopic kinetic studies. Gridnev and Vorobiev investigated by quantum chemical DFT calculations and kinetic analysis potential acetal intermediates. They concluded that the acetals are off-loop species because they are not precursor of the reaction product.\cite{24} In this context, another highly remarkable discovery was reported by Hawbaker and Blackmond, where hydroxy ethers interfere with the Soai reaction and even inhibit the reaction.\cite{25}

More recently, Denmark and co-workers\cite{26} performed investigations of the Soai reaction with focus on the role of the nitrogen atoms in the pyrimidine/pyridyl moiety, the structure of the Zn-alkoxides in solution by NMR spectroscopic studies and in situ IR kinetic studies using pyridyl-3-carbaldehyde as surrogant, ("Trojan-Horse" substrate). An alternative mechanism is proposed, considering a "cube escape" model. Such cubetype structures were also proposed by Noyori and co-workers\cite{27} in the enantioselective addition of dialkylzincs to aldehydes promoted by chiral β-amino alcohols, that is, (−)-3-exo-(dimethyl-amino)soborneol (DAIB) and has been discussed by Brown and co-workers in the context of the Soai reaction as potential tetrameric structure of the Zn-alkoxides.\cite{28}

Furthermore, the Soai reaction shows some peculiarities that are well documented, but still unexplained such as (a) an un-
usual inverse temperature dependence on the reaction kinetics, that is, the maximum reaction rate increases significantly as the reaction temperature is decreased,[41] and (b) a prolonged induction period, which are not yet rationalized by properties of potential catalyst structures or corresponding kinetic models. Herein, we seek to investigate the open mysteries of Soai’s asymmetric autocatalysis.

Results and Discussion

First, we performed reaction kinetic investigations using multiplexing HPLC[42] in the flow-injection mode[43] (Chiralpak IB, mobile phase n-hexane/THF 55:45, 1.2 mL min⁻¹), which provides temporal resolution of the stereoisomers formed (Figure 1 in the Supplementary Information). Reaction progress was observed injection of sample pulses with a time interval of 2.1 min from the reaction mixture onto the chiral separation column (Figure 1a and b). We systematically varied the concentrations of the reactants and additives of the Soai reaction (2-tert-butylacetylene-1-yl)pyrimidyl-5-carbaldehyde 4: 10.6–41 mmol, (R)-2-tert-butylacetylene-1-yl)pyrimidyl-5-(iso-butanol-1-ol) (R)-1 (ee > 99.9%): 0.266–4 mmol; Pr₃Zn: 30–130 mmol (Figure 1b and Figures S6–S23 in the Supporting Information).

Kinetic analysis of these data gives a reaction order of 1.9 in [4], an order of 1 in [1], and an order of 0 in [Pr₃Zn] (Figures 24–26 in the Supporting Information), confirming previous studies [Eq. (1)].[41]

\[
d[1]/dt = k[4]^{1.9}[1]^{1}[Pr₃Zn]^{0}
\]

Surprisingly, when we performed the HPLC separation of the reaction mixture using isopropanol instead of THF in the mobile phase we observed the enantiomers of another compound connected by plateau formation with 4. We identified these peaks as the isopropyl hemiacetals (R)-5ₚₚ, or (S)-5ₚₚ, (Figure 2a), that is, in the case of chiral alcohols diastereomeric hemiacetals are expected (vide infra).

Remarkably, these hemiacetals 5ₚₚ are subject to a dynamic interconversion, which we investigated by temperature-dependent enantioselective dynamic HPLC (DHPLC)[44,45] (Figure 2b and Figures S28–S35 in the Supporting Information).

The thermodynamic parameters of the formation of the hemiacetal 5ₚₚ were determined by linear regression of the thermodynamic Gibbs free energies ΔG(T), obtained from the equilibrium constants K, vs. the temperatures T (correlation coefficient r = 0.9949) to be ΔG = 3 kJ/mol, ΔH = −15.6 kJ/mol and ΔS = −62.5 J/(K mol) (Figure S36 in the Supporting Information).

The activation enthalpies ΔH* for the hemiacetal 5ₚₚ formation and decomposition were obtained via the slope and the activation entropies ΔS* via the intercept of the Eyring plots (ln(k/T) vs. 1/T) (Figure 2c). Deviations of the activation parameters ΔH* and ΔS* have been calculated by error band analysis of the linear regression with a level of confidence of 95%.

The activation parameters of the hemiacetal 5ₚₚ formation are ΔH* = 26.3 ± 0.2 kJ mol⁻¹ and ΔS* = −195 ± 34 J/(K mol) (r = 0.9990, residual deviation sₚₚ = 0.0306) and the hemiacetal 5ₚₚ decomposition are ΔH* = 47.7 ± 0.2 kJ mol⁻¹ and ΔS* = −112 ± 1 J/(K mol) (r = 0.9994, sₚₚ = 0.0630). The formation of hemiacetal 5ₚₚ from 4 is endergonic, however this is a highly dynamic process and interestingly, the kinetic parameters kₚₚ (293 K) = 4.1 × 10⁻³ (mol s⁻¹) and k₋ₚₚ (293 K) = 1.3 × 10⁻² s⁻¹. Hemiacetals, formed from the pyrimidine-5-carbaldehyde and its corresponding alcohol can function as a transient chiral ligand to activate the dialkylzinc reagent, very similar to the ß-dialkylaminoalcohols in Noyori’s DAIB catalysis[46] or Blackmond’s hydroxy ethers.[19] The in situ formation of a transient catalyst by reaction or interaction of molecules participating in the reaction is a fundamental mechanism leading immediately to autocatalysis and amplification. Similar mechanisms are well known in substrate activated enzyme catalysis to regulate biochemical reaction networks and in artificial systems.[46] Furthermore, the
thermodynamic data indicate that the formation of the hemiacetal is favored at lower temperature and if it is involved in the autocatalytic cycle, it correlates with the observation that the Soai reaction is accelerated at lower temperature. Moreover, the measured reaction kinetics of the formation of the hemiacetals agrees with the observed induction period of the Soai reaction.

1H NMR spectroscopic studies in CD$_3$OD reveal that the electron-deficient pyrimidine moiety of 4 and the unsubstituted pyrimidyl-5-carbaldehyde $4_H$ promote the formation of deuterated methyl hemiacetals (characteristic hemiacetal proton at $\delta = 5.7$ ppm; cf. Figures S39 and S42 in the Supporting Information) in 95% yield at room temperature, while for comparison benzaldehyde yields only 9% (Table S4 in the Supporting Information). Temperature dependent measurements confirm the trend that the hemiacetal formation is favored at lower temperature (Figures S40 and S41 in the Supporting Information). This formation of hemiacetals is also observed in toluene, the solvent used in the Soai reaction: Reaction of 4 and $4_H$ with rac-2-methyl-1-phenylpropan-1-ol in [D$_8$]toluene gives diastereomeric hemiacetals in 11% and 9% yield, respectively. More important, the diastereomeric ratio of 1:2.6 for the unsubstituted pyrimidine hemiacetal improves to 1:5.2 for the 2-(tert-butylacetylene-1-yl) substituted pyrimidine hemiacetal (HSQC spectra depicted in Figures S43 and S44 in the Supporting Information). In this context, it is important to note that the formation of stereolabile hemiacetals offers the conceptual mechanism of minor enantiomer recycling$^{[47]}$, leading to the amplification of the major enantiomer/diastereoisomer.

In the next step, we investigated the Soai reaction by in situ high-resolution mass spectrometric experiments. In these experiments there is no separation column involved to avoid quenching of the reaction intermediates. We monitored the course of the Soai reaction under inert conditions (anhydrous toluene, argon atmosphere) by feeding the reacting reaction mixture continuously (10 µL/ min) but pulsed (time interval of 30 s) via a 5 µL sample loop of a 6-port valve (anhydrous toluene as eluent, flow rate 200 µL/ min) into a high-resolution Orbitrap mass spectrometer (Figure S45 in the Supporting Information) using atmospheric pressure chemical ionization (APCI) under mild ionization conditions ($T = 150\, ^{\circ}\mathrm{C}$, N$_2$).

To identify all intermediates and transient intermediates during the Soai reaction, all pulsed injections were summed up over the complete reaction progress and the mass range between $m/z$: 180 and 800 in a first step (Figure 3a; Figures S46–49 in the Supporting Information). This mass spectrum shows the complexity of intermediates formed in the Soai reaction. We identified (cf. Figure 3b) the alcohol 1a and its fragment 1b, the monomeric Zn-alkoxides 2a and 2b, the dimeric Zn-alkoxides 3a and 3b (different charge states), pyrimidine-5-carbaldehyde 4, the Zn-complex of the hemiacetalate 5a, the hemiacetal fragments 5b and 5c, a hydroxylated structure 5d, which is probably formed by the ionization process, hemiacetal 5e with iPr$_2$Zn coordinated to it (this can be bridged or open), the hemiacetalate with coordinated toluene 5f. Structures 7a–7d are Zn–hemiacetalate complexes with another molecule or fragment of alcohol 1 coordinated, and structure 8 represents a dimeric Zn–hemiacetal complex.

The identification of the Zn complexes is facilitated by the characteristic isotope pattern of Zn (Figure 4a, right MS spectrum); for comparison, the high-resolution MS spectrum on the left side in Figure 4a shows the formation of hemiacetal 5 ($m/z$ 421.25822) by mixing 1 and 4 in toluene. High-resolution MS spectra of the identified structures are depicted in Fig-
are pre-mixed, the selectivity depends on the nature of the metal and the catalyst, which are not stable under the experimental conditions of the APCI.

Furthermore, these mixing experiments are in good agreement with the experiments by Amedjkouh using pyrimidine alcohol 1 as a chiral additive in the Soai reaction of pyridyl-3-carbaldehydes.12,48 The authors further characterized the formation of hemiacetals by partial rearrangement (tweezers) of the complexes, which activate the zinc reagent, and favor the formation of hemiacetals by spatial arrangement (tweezers) and electronic effects. Interestingly, no bridged pyrimidyl alcohols 1, pyrimidyl alcohols 2 or pyrimidine-5-carbaldehydes 4 are observed, which can be attributed to the noncovalent nature of such complexes, which are not stable under the conditions of the APCI.

To investigate the role of the dialkylzinc reagent and the alcohol additive in the Soai reaction, we performed mixing experiments. We systematically varied iPrZn and diethylzinc (Et₂Zn) as reagent in the reaction itself and to pre-form Zn-alkoxides 2.

In a second set of experiments, we mixed various concentrations of the additives (S)-1, (S)-1-H, and (R)-1-H to study the effect on the ee. While in the Soai reaction of 4 with Et₂Zn only low ee values are observed, the reaction with a 1:1 mixture of iPrZn and Et₂Zn/ (S)-1 gives ee values of 68% and 63% for (S)-1 and (S)-1-H, respectively (Scheme 2a).

If the reaction is started with pre-formed (S)-2 or (S)-2-EtH, higher ee values are observed for the isopropyl substituted alkoxide (S)-2 compared to the ethyl substituted alkoxide (S)-2-EtH (Scheme 2b). If iPrZn, Et₂Zn and (S)-1 are pre-mixed, the selectivity of the isopropyl substituted alkoxide (S)-2 dominates and gives higher ee values (Scheme 2c). This is also observed for the competitive reaction using (S)-2 and (S)-2-EtH simultaneously (Scheme 2d). These experiments show that the catalyst formed in the induction period strongly depends on the starting conditions and remains catalytically active and selective throughout the Soai reaction.

This becomes evident from the mixing experiments where the concentrations of the additives (S)-1, (S)-1-H, and (R)-1-H are varied. The 2-(tert-butylacetylene-1-yl) substituted alcohol (S)-1 dominates the selectivity, resulting in high ee values (Table 1). Entries 4 and 5 of Table 1 show that the catalyst formed from additive (S)-1 remains catalytically active and the selectivity is controlled by the catalyst, which is better stabilized in solution. The 2-(tert-butylacetylene-1-yl) substituent improves the residence time and with that the turnover number. This explains also the excellent selectivities of the 2-(adamantylacetylene-1-yl)- and 2-(trimethylsilylacetylene-1-yl)-substituted pyrimidyl alcohols.10c,37 Furthermore these mixing experiments are in good agreement with the experiments by Amedjkouh using pyrimidine alcohol 1 as a chiral additive in the Soai reaction of pyridyl-3-carbaldehydes.12d,48
A doping experiment corroborates that the Soai reaction is catalyzed by a transient catalyst formed in the course of the reaction. For this purpose, we transferred the reaction solution of a running Soai reaction to a just started Soai reaction.

For this, four Soai reactions (25 mM 2-(tert-butylacetylene-1-yl)pyrimidyl-5-carbaldehyde 4. 1.25 mM (R)-2-(tert-butylacetylene-1-yl)pyrimidyl-5-(iso-butanol-1-ol) (R)-1 (ee > 99.9%), 20 °C, all concentrations are final concentrations after addition of the iPrZn solution) were prepared from the same stock solutions and distributed in 4 vials under inert reaction conditions. 2 vials were used as reference vials, one started simultaneously (1st vial) with a 2nd vial, in which the transient catalyst is formed during the reaction (for experimental details a detailed timing Scheme is depicted in Figure S64 in the Supporting Information). The Soai reactions are started by addition of iPrZn (125 mM final concentration). After 210s the 3rd Soai reaction is started by addition of iPrZn (125 mM final concentration). 218 s after the start of the reference reaction and the Soai reaction in the 2nd vial 60 μL of the reaction solution are transferred from the 2nd vial into the 3rd vial. The 2nd reference (4th vial) was started after the completion of reactions 1 and 2. All reactions were monitored by multiplexing HPLC in the flow-injection mode. Analysis of the kinetic data (Figure 5) shows, that the induction period is reduced in the doped experiment. Quantitative kinetic analysis confirms that the inflection point (maximum reaction rate) of the "normal" Soai reaction (25.0 mM 4. 1.25 mM (R)-1 (ee > 99.9%) and 125 mM iPrZn in toluene at 20 °C) is at 25s with an initial reaction rate of 4.2 x 10⁻³ mol⁻¹ s⁻¹, while in the doped experiment the inflection point is at 200 s and the initial reaction rate is 1.4 x 10⁻² mol⁻¹ s⁻¹ (Figures S65–S70 in the Supporting Information). The induction period is shifted by 55 s in the doped experiment. It has to be pointed out, that adding 60 μL of a completed Soai reaction does not influence the induction period.

Considering all kinetic and thermodynamic data of the hemiacetal formation, the structural information obtained by the in situ high-resolution mass spectrometric reaction monitoring, the transient formation of the Zn–hemiacetolate 5 during the Soai reaction, the mixing experiments and the doping experiment, we propose a reaction mechanism that can explain the high amplification of the experimentally observed enantioselectivity (Scheme 3). This mechanism starts with the formation of the isopropylzinc pyrimidyl alkoxides 2 from (R)-1 and/or (S)-1. The isopropylzinc pyrimidyl alkoxides 2 are in equilibrium with the homochiral (R,R)-3/ (S,S)-3 and heterochiral (R,S)-3 dimers. Depending on this first selection process the dominating enantiomer directs if the autocatalytic cycle proceeds to the right autocatalytic R-cycle (green) or the left autocatalytic S-cycle (red).

A key step is the formation of the transient hemiacetal catalyst 5, which is in a dynamic equilibrium between the zinc aldehyde 2 and aldehyde 4 forming diastereomeric complexes (R,R)-5 or (S,S)-5, if (R)-1 dominates, or forming diastereomeric complexes (S,S)-5 or (R,R)-5, if (S)-1 dominates. These structures are very similar to the zinc β-dialkylaminoalcoholates structures in Noyori’s DAIB catalyst. DFT calculation at the PBE0-D3/LACVP** level of theory as implemented in the Jaguar 10.1 quantum chemistry program package[56–58] of optimized structures of (R)-5 and (S)-5 indicate that the diastereomer (R,R)-5 is favored by 6 kJ mol⁻¹ (cf. Supporting Information, Figures S118 and S119). In the following only the autocatalytic R-cycle on the right side will be discussed, which is mirror-symmetrical to the autocatalytic S-cycle. In the next step of this cycle pyrimidine-5-carbaldehyde 4 and iPrZn are coordinated to the hemiacetal (R,R)-5 forming adduct (R,R)-6. This spatial alignment of the carbaldehyde results in a transfer of the adjacent isopropyl group from the re side giving (R,R,R)-7. DFT calculations provide an energy barrier of 54 kJ mol⁻¹ (cf. Supporting Information Figures S120–S123). Insertion of another molecule of the pyrimidine-5-carbaldehyde 4 leads to the dimeric hemiacetal (R,R,R,R)-8, which splits into two monomeric hemiacetals (R,R)-5, which explains the rapid sigmoidal increase in the formation of catalytically highly active (R,R)-5 (Figure 4b) and is typical for an autocatalytic process. The dimeric hemiacetal (R,R,R,R)-8 represents a “super” diastereo-
meric complex and in combination with the dissociation into the monomeric hemiacetals ([R,R]-5), which can dynamically control the stereocenter of the hemiacetal group, gives a natural mechanism of autocorrection. It has to be noted that ([R,R]-7) can be also directly converted into ([R,R]-5) and ([R,R]-2) (and its corresponding dimer ([R,R]-3)).

For the evaluation of the kinetic data we developed three kinetic models (models I, II and III) with increasing complexity (Figures S71–S73 in the Supporting Information). The minimal model (I) (Figure S71 in the Supplementary Information, 7 reaction steps) considers only a single enantiomer, the extended model (II) (Figure S72 in the Supporting Information) considers both enantiomers, and the comprehensive model (III) (Figure S73 in the Supporting Information) takes the epimeric hemiacetals 5 into account. It has to be pointed out that all three models result in consistent intrinsic reaction rates. The comprehensive model III consists of 26 differential equations (see Supporting Information for details) based on the here presented mechanism (Scheme 3) was created and implemented in a software program (Soai 7). This program allows to calculate kinetic reaction profiles using an adaptive Runge–Kutta routine to solve the system of differential equations with the initial experimental parameters, that is, concentrations of the additives (R)-1 and (S)-1 (ee), the pyrimidin-5-carbaldehyde 4 and iPr₂Zn, the reaction time, reaction rate constants \( k_n \) (Scheme 3) and equilibrium constants \( K_n \). This program allows to define large data sets (2.25 million kinetic profiles each) with variable ranges for the reaction rate constants \( k_n \) and equilibrium constants \( K_n \). The calculated kinetic profiles are compared with the experimentally determined kinetic profiles of (R)-1, (S)-1 and 4 (Figure 1b and Figures S6–S23 in the Supporting Information) to refine the kinetic parameters. This method was applied iteratively to all kinetic data sets (in total 81 million kinetic profiles) and thus the rate constants for the respective partial steps were determined (Figures S79–S100 in the Supporting Information). The kinetic and thermodynamic parameters are summarized in Table 2.

The formation of the isopropylzinc pyrimidyl alkoxides 2 is a rapid process and agrees very well with kinetic data of the reaction of alkylzinc compounds with alcohols. The equilibrium between monomeric 2 and homochiral ([R,R]-3) ([S,S]-3) and heterochiral ([R,S]-3) is dynamic and not extremely shifted to the side of the dimers, as it is also observed in the mass spectra of the isopropylzinc pyrimidyl alkoxides.
(Figure 3a). More interesting is the equilibrium and the kinetic parameters of the hemiacetal 5 formation, which are in excellent agreement with the kinetic parameters determined by enantioselective DHPLC (Figure 2a–c) for the formation of $S_{\text{Pr}}$ $(k_{\text{f}}(293 \, \text{K}) = 4.1 \times 10^{-3} \, (\text{mol} \, \text{s}^{-1})^{-1}$) and $k_{\text{i}}(293 \, \text{K}) = 1.3 \times 10^{-2} \, \text{s}^{-1}$) and equilibria of the derivatives by $^1$H NMR spectroscopy. In the autocatalytic cycle the rate determining step is the transfer of the isopropyl group, while the other steps are energetically balanced.

The proposed mechanism and the kinetic model allow not only to predict kinetic reaction profiles of the conversion of the pyrimidine-5-carbaldehyde 4 into the reaction product 1 of the Soai reaction, but even more important the precise predictions of the nonlinear amplification of the ee and the induction period in dependence on the ee. When starting with an ee of 1% in 1 (2 mmol/L), 9.25% ee in the 1st step, 59.4% ee in the 2nd step, 94.6% ee in the 3rd step, 99.4% ee in the 4th step and 99.9% ee in the 5th step are obtained (Figures S101–S105 of the Supporting Information). A systematic variation of the initial en $e_{0}$ of 1 and the corresponding final product ee is plotted in Figure 6a (the corresponding ee simulations are plotted in Figures S106–S115 in the Supporting Information). Interestingly, if the reaction is performed under conditions, where the formed product with amplifying ee propagates through a reaction mixture, that is, by diffusion or starting with seeding on a chiral or enantiomorph surface, extraordinary enhancement of ee amplifications can be predicted, jumping immediately from $1 \times 10^{-5}$% to 55% and finally > 99.9% (Figure 6a; red line). Experimental investigations of reactions with variation of the starting ee of the alcohol additive and concentrations were compared with the prediction of ee values by simulation with the program Soai 7, giving an excellent correlation between experiment and simulation (see Supporting Information Table S5 and Figure S117).

Further, the simulations correctly predict the prolonged induction period, which is caused by the slow hemiacetal formation, and the time of the inflection point $t_{\text{ip}}$ in dependence on the initial ee of the pyrimidine alcohol 1 (Figure 6b).

**Conclusions**

In summary, the results of the high-resolution mass spectrometric measurements and the comprehensive kinetic analyses suggest the formation of a transient Zn–hemiacetalate complex, which is catalytically active in the Soai reaction. This intermediate can establish the here proposed autocatalytic cycle and the extraordinary amplification of the enantiomeric excess. This is supported by mass spectrometric profiling of the transient hemiacetal intermediate and by doping experiments, which demonstrate that the Soai reaction can be accelerated by adding the in situ formed catalyst. Kinetic and thermodynamic data of the highly dynamic formation and decomposition of the hemiacetal explain the unusual inverse temperature dependence on the reaction kinetics, the induction period and time shift of the inflection point. Furthermore, the results suggest that the formation of the transient diastereomeric Zn-hemiacetates amplify any initial imbalance of the formed product enantiomers, which is interestingly always given for an odd number of formed molecules. These results give a new guidance to structures envisioning potential processes leading to symmetry breaking.

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**Conflict of interest**

The authors declare no conflict of interest.

**Keywords:** autocatalysis · hemiacetal · kinetic analysis · mass spectrometry · Soai reaction

[1] a) J.-M. Lehn, *Angew. Chem. Int. Ed.* 2013, 52, 2836 –2850; Angew. Chem. 2013, 125, 2906 –2921; b) J.-M. Lehn, *Angew. Chem. Int. Ed.* 2015, 54, 3276 –3289; Angew. Chem. 2015, 127, 3326 –3340.
[2] D. G. Blackmond, *Chem. Rev.* 2020, 120, 4831 –4847.
[3] a) I. Weissbuch, L. Addadi, Z. Berkovich-Yellin, E. Gati, M. Lahav, L. Leiserowitz, *Nature* 1984, 310, 161 –164; b) D. K. Kondepudi, R. J. Kaufman, N. Singh, *Science* 1990, 250, 975 –976; c) B. L. Feringa, R. A. van Delden, *Angew. Chem. Int. Ed.* 1999, 38, 3418 –3438; Angew. Chem. 1999, 111, 3624 –3645; d) H. Zepik, E. Shavit, M. Tang, T. R. Jensen, K. Kjaer, G. Bolbach, L. Leiserowitz, I. Weissbuch, M. Lahav, *Science* 2002, 295, 1266; e) K. Mikami, M. Yamanaka, *Chem. Rev.* 2003, 103, 3369 –3400; f) D. G. Blackmond, *Proc. Natl. Acad. Sci. USA* 2004, 101, 5732 –5736; g) R. R. E. Steendam, M. C. T. Brouwer, E. M. E. Huijs, M. W. Kulka, H. Meekes, W. J. P. V. Enckevort, J. Raap, F. P. J. T. Rutjes, E. Vlieg, *Chem. Eur. J.* 2014, 20, 13527 –13530; h) S. Olsson, P. M. Björmark, T. Kokoli, J. Sundberg, A. Lennartson, C. J. Mckenzie, M. Håkansson, *Chem. Eur. J.* 2015, 21, 5211 –5219; i) J. M. Ribó, J. Crusats, Z. El-Hachemi, A. Moyano, D. Hochberg, *Chem. Sci.* 2017, 8, 763 –769.
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Angew. Chem. Int. Ed. 2003, 55, 126.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2012, 51, 3515 – 3515.

Angew. Chem. Int. Ed. 2015, 54, 2014 – 2014.

Angew. Chem. Int. Ed. 2012, 51, 15246 – 15249.

Angew. Chem. Int. Ed. 2015, 54, 21371 – 21371.

Angew. Chem. Int. Ed. 2006, 15, 15246 – 15249.

Angew. Chem. Int. Ed. 2007, 46, 5609 – 5613.

Angew. Chem. Int. Ed. 2007, 46, 5609 – 5613.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2014, 53, 11919 – 11922.

Angew. Chem. Int. Ed. 2004, 53, 15246 – 15249.

Angew. Chem. Int. Ed. 2015, 54, 15246 – 15249.

Angew. Chem. Int. Ed. 2007, 46, 5609 – 5613.

Angew. Chem. Int. Ed. 2015, 54, 15246 – 15249.

Angew. Chem. Int. Ed. 2015, 54, 15246 – 15249.

Angew. Chem. Int. Ed. 2007, 46, 5609 – 5613.

Angew. Chem. Int. Ed. 2015, 54, 15246 – 15249.

Angew. Chem. Int. Ed. 2006, 14, 1510 – 1512.

Angew. Chem. Int. Ed. 2009, 48, 1479 – 1482.

Angew. Chem. Int. Ed. 2009, 48, 1479 – 1482.

Angew. Chem. Int. Ed. 2014, 53, 11919 – 11922.

Angew. Chem. Int. Ed. 2004, 53, 15246 – 15249.

Angew. Chem. Int. Ed. 2007, 46, 5609 – 5613.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2012, 51, 15246 – 15249.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.

Angew. Chem. Int. Ed. 2009, 48, 1236 – 1236.
[45] a) O. Trapp, V. Schurig, J. Am. Chem. Soc. 2000, 122, 1424 – 1430; b) O. Trapp, Anal. Chem. 2006, 78, 189 – 198; c) O. Trapp, Electrophoresis 2006, 27, 534 – 541; d) O. Trapp, Electrophoresis 2006, 27, 2999 – 3006; e) F. Maier, O. Trapp, Angew. Chem. Int. Ed. 2012, 51, 2985 – 2988; Angew. Chem. 2012, 124, 3039 – 3043.

[46] a) H. Fanlo-Virgos, A.-N. R. Alba, S. Hamieh, M. Colomb-Delsuc, S. Otto, Angew. Chem. Int. Ed. 2014, 53, 11346 – 11350; Angew. Chem. 2014, 126, 11528 – 11532; b) C. Kremer, A. Lutzen, Chem. Eur. J. 2013, 19, 6162 – 6196.

[47] a) E. Wingstrand, A. Laurell, L. Fransson, K. Hult, C. Moberg, Chem. Eur. J. 2009, 15, 12107 – 12113; b) L. Fransson, C. Moberg, ChemCatChem 2010, 2, 1523 – 1532; c) L. Fransson, A. Laurell, K. Widyan, E. Wingstrand, K. Hult, C. Moberg, ChemCatChem 2010, 2, 683 – 693; d) A. Laurell, C. Moberg, Eur. J. Org. Chem. 2011, 3980 – 3984; e) Y.-Q. Wen, R. Hertzberg, I. Gonzalez, C. Moberg, Chem. Eur. J. 2014, 20, 3806 – 3812; f) C. Moberg, Acc. Chem. Res. 2016, 49, 2736 – 2745.

[48] M. Funes-Maldonado, B. Sieng, M. Amedjkouh, Eur. J. Org. Chem. 2015, 4081 – 4086.

[49] C. Adamo, V. Barone, Chem. Phys. Lett. 1998, 298, 113 – 119.

[50] a) J. P. Perdew, K. Burke, M. Ernzerhof, Phys. Rev. Lett. 1996, 77, 3865 – 3869; b) J. P. Perdew, K. Burke, M. Ernzerhof, Phys. Rev. Lett. 1997, 78, 1396.

[51] S. Grimme, E. Antony, T. Ehrlich, E. Krieg, J. Chem. Phys. 2010, 132, 154104.

[52] P. J. Hay, L. R. Wadt, J. Chem. Phys. 1985, 82, 299 – 310.

[53] P. C. Hariharan, J. A. Pople, Theor. Chim. Acta 1973, 28, 213 – 222.

[54] Jaguar, version 10.1, Schrodinger, Inc., New York, NY, 2018.

[55] A. D. Bochevarov, E. Harder, T. F. Hughes, J. R. Greenwood, D. A. Braden, D. M. Philipp, D. Rinaldo, M. D. Halls, J. Zhang, R. A. Friesner, Int. J. Quantum Chem. 2013, 113, 2110 – 2142.

[56] O. Trapp, Soai 7, compatible with Microsoft Windows 7, 8 and 10. The compiled executable program can be obtained from the author upon request.

[57] R. J. Herold, S. L. Aggarwal, V. Neff, Can. J. Chem. 1963, 41, 1368 – 1380.

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