Role of Plasma Catalysis in the Microwave Plasma-Assisted Conversion of CO$_2$

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**Abstract**

Climate change and global warming caused by the increasing emissions of greenhouse gases (such as CO$_2$) recently attract attention of the scientific community. The combination of plasma and catalysis is of great interest for turning plasma chemistry in applications related to pollution and energy issues. In this chapter, our recent research efforts related to optimization of the conversion of CO$_2$ and CO$_2$/H$_2$O mixtures in a pulsed surface-wave sustained microwave discharge are presented. The effects of different plasma operating conditions and catalyst preparation methods on the CO$_2$ conversion and its energy efficiency are discussed. It is demonstrated that, compared to the plasma-only case, the CO$_2$ conversion and energy efficiency can be enhanced by a factor of $\sim 2.1$ by selecting the appropriate conditions. The catalyst characterization shows that Ar plasma treatment results in a higher density of oxygen vacancies and a comparatively uniform distribution of NiO on the TiO$_2$ surface, which strongly influence CO$_2$ conversion and energy efficiencies of this process. The dissociative electron attachment of CO$_2$ at the catalyst surface enhanced by the oxygen vacancies and plasma electrons may explain the increase of conversion and energy efficiencies in this case. A mechanism of plasma-catalytic conversion of CO$_2$ at the catalyst surface in CO$_2$ and CO$_2$/H$_2$O mixtures is proposed.

**Keywords:** green energy, CO$_2$ conversion, plasma catalysis, microwave discharge, NiO/TiO$_2$ catalyst

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1. Introduction

One of the biggest concerns of the twenty-first century is the effect of global warming and its destructive consequences on the global ecosystem. Evidence indicated that the main cause of increased amounts of greenhouse gas emissions (such as CO$_2$) are the human activities related to the industrialization and burning of the fossil fuels. The process of conversion of CO$_2$ along with the other greenhouse gases into value-added chemicals or new fuels has attracted considerable attention because of the current background of fossil resources depletion and increase of CO$_2$ emissions. If a renewable (from sunlight or wind) or nuclear energy source drives the dissociation of CO$_2$ or other greenhouse gases, the renewable or nuclear energy can be stored as chemical energy in fuels. In this context, CO$_2$ reutilization to synthesize syngas, valuable fuels, or chemical compounds as well as pure CO$_2$ dissociation into CO and O$_2$ is of a special interest. Large efforts have been made to develop energy-efficient technologies [1–7].

Different techniques for conversion of CO$_2$ into value-added chemical compounds or fuels activated by heat or electricity have been developed and studied, namely thermolysis, photocatalysis, and electrochemical methods [1–5]. The CO$_2$ thermal splitting is thermodynamically and energetically favorable only at very high temperatures (1400–1800°C) and the conversion efficiency is very low. A prototype of CO$_2$ thermolysis chamber using only solar energy was developed by Traynor and Jensen [2]. The measured net conversion of CO$_2$ to CO was close to 6% under optimal conditions. The maximum conversion of solar to chemical energy was 5%. Recently, photocatalytic conversion of CO$_2$ using solar energy has attracted more and more attention. However, one of the greatest drawbacks is the low conversion efficiency at present. Moreover, its application is limited due to the numerous disadvantages [3, 4]: (1) inefficient reactor system, (2) catalyst activation needing long periods of irradiation, and (3) inefficient exploitation of visible light. High-temperature electrolysis is another promising method since it can effectively convert the electrical energy into chemical energy (78–87%) [5]. Electrolysis of CO$_2$ in the liquid phase has been investigated extensively while only few studies have been reported for electrolysis of CO$_2$ in the gas phase. Much effort should be made to overcome the drawbacks such as short-term stability of electrolytes, using expensive or scarce materials, and whole system heating.

Plasma technology represents an alternative approach to the above-mentioned methods. The nonthermal plasma technology is considered to be an attractive alternative to other (classical) technologies for converting inert carbon emissions (such as CO$_2$) into valuable fuels and chemicals, due to its nonequilibrium characteristics, environmental-safety, scalability, and low power requirements [6]. The fundamental properties and usefulness of the nonthermal plasmas for conversion of carbon dioxide as well as for related plasma technologies are extensively reviewed by Fridman [6], Zou and Liu [7], and by the other numerous authors. Nonthermal plasmas including dielectric barrier discharges (DBDs) [8–13], microwave discharges (MWs) [14–21], gliding arc plasmatrons (GAPs) [22–24], glow discharges [25], radio frequency (RF) discharges [26], and corona discharge [27] have been investigated for CO$_2$ conversion so far. However, a known trade-off between the energy efficiency and conversion efficiency is observed in plasma-assisted dissociation of pure CO$_2$. In order to overcome this trade-off, a plasma-assisted catalytic process with plasma activation of the catalysts to
increase the conversion efficiency while maintaining high energy efficiency has been proposed. The combination of heterogeneous catalysis and plasma activation, known as plasma catalysis, has attracted increasing interest [28–31]. Plasma-catalytic conversion of CO₂ is a complex and challenging process involving a large number of physical and chemical reactions. The efficiency of this process can be controlled by means of the plasma parameters themselves, as well as by the catalyst properties. This suggests that more systematic studies on both the plasma effects and the chemical effects of the catalyst are necessary.

It is well known that the vibrational excitation is considered to be the most efficient elementary process in CO₂ plasmas to stimulate the endothermic CO₂ decomposition via the so-called vibrational ladder climbing process [6]. Due to its high degree of nonequilibrium, microwave plasma excites highly vibrational states of CO₂ molecules, which are energy-efficient for CO₂ decomposition. In this regard, combining catalysis with microwave plasmas could be especially promising, as the catalyst surface can reach a somewhat elevated temperature and possibly induce a synergy. Up to now, there are only a limited number of attempts for combining catalyst with microwave plasma [15, 17, 19–21].

The present Chapter critically summarizes the plasma-catalyst phenomena for optimization of the microwave plasma-assisted CO₂ conversion focusing on NiO/TiO₂ catalysts. It is focused on the investigation of several plasma-catalytic effects and their influences including the catalyst preparation methods, plasma activation and NiO content, aiming at enhancing our understanding of the plasma-catalytic CO₂ conversion process.

2. Brief theoretical background

The excitation of both CO₂ and water vapor is required for effective dissociation. The high energy plasma electrons have the ability to efficiently excite and dissociate CO₂ and H₂O via collisional processes. The dissociation of a CO₂ molecule is represented by the following global reaction [6]:

\[
\text{CO}_2 \rightarrow \text{CO} + \frac{1}{2}\text{O}_2, \Delta H = 2.9 \text{ eV/molecule} \tag{1}
\]

The process given by Eq.1 can be represented by an electron impact CO₂ dissociation [6]:

\[
\text{CO}_2 \rightarrow \text{CO} + \text{O}, \Delta H = 5.5 \text{ eV/molecule} \tag{1a}
\]

following either by O recombination into O₂ [Eq. (1b)] or by the reaction with a ground state (but vibrationally excited) CO₂ molecule [Eq. (1c)]:

\[
\text{M} + \text{O} + \text{O} \rightarrow \text{O}_2 + \text{M} (\text{M is a particle}) \tag{1b}
\]

\[
\text{O} + \text{CO}_2 \rightarrow \text{CO} + \text{O}_2, \Delta H = 0.3 \text{ eV/molecule}. \tag{1c}
\]

Besides this, CO can be easily converted into CO₂ with the production of H₂ in the presence of water, which is known as the water-gas shift reaction (WGSR):

\[
\text{CO} + \text{H}_2\text{O} \rightleftharpoons \text{CO}_2 + \text{H}_2, \Delta H = -0.4 \text{ eV/molecule (forward direction)}. \tag{2}
\]
Rather than disassociating the CO\(_2\) directly, H\(_2\) can reduce CO\(_2\) to CO via the reverse direction depending on the reaction conditions. The analysis of the CO\(_2\) and H\(_2\)O dissociation thermodynamics shows that the WGSR is favored in the forward direction if the temperature is less than about 827°C. Above that value the WGSR is favored in the reverse direction [5].

The decomposition of a H\(_2\)O molecule is given by the following reaction [6]:

\[
\text{H}_2\text{O} \rightarrow \text{H}_2 + \frac{1}{2}\text{O}_2, \Delta H = 2.6 \text{ eV/molecule}
\]  

(3)

Traditionally, in order to characterize the process efficiency, two main parameters reflecting the energy efficiency and conversion efficiency are used: conversion efficiency and energy efficiency. By definition [4], the energy efficiency \(\eta\) of dissociation process is given by

\[
\eta = \frac{\Delta H}{SEI}
\]

(4)

where \(\Delta H\) is the dissociation enthalpy of the global reaction [Eq. (1)], that is, 2.9 eV/molecule in the case of CO\(_2\), \(\chi\) is the CO\(_2\) conversion efficiency [Eq. (5)], and SEI is the specific energy input per molecule (in eV).

The conversion efficiency of CO\(_2\) is defined based on the following ratio:

\[
\chi = \frac{\text{moles of CO}_2 \text{ converted}}{\text{moles of CO}_2 \text{ supplied}}
\]

(5)

In our case, \(\chi\) is calculated by comparing the gas chromatography (GC) CO\(_2\) peak area without and with the presence of plasma discharge, as described in Ref. [15]. The specific energy input per molecule, SEI, is given by the ratio of the discharge power \((P)\) to the gas flow rate \((F)\) through the discharge volume and it can be expressed in eV per molecule.

\[
\text{SEI} = \frac{P}{F}
\]

(6)

### 3. Experimental part

#### 3.1. The microwave discharge reactor

Figure 1 shows a schematic diagram of the surface-wave microwave set-up used in this work. The discharge was sustained in a quartz tube (14 mm in diameter and 31 cm in length) by an electromagnetic wave with a filling frequency of 915 MHz. The quartz tube was additionally cooled by 5°C oil. In order to minimize the reflected microwave power, the impedance of the waveguide was automatically adjusted by a three-stub tuning system. The reflected power measured for each condition was always below the detection limit of the measuring probe and thus considered negligible. The CO\(_2\) or the other gases (Ar, O\(_2\), H\(_2\), or H\(_2\)O) used in our study are injected from the top of the system. The gas flow rate was regulated by electronic mass flow controllers ranging between 0 and 6 standard liter per minute (slm). The water vapor was generated in an Omicron technologies vaporization system producing a vapor flow rate between 0 and 3 slm. The obtained discharge pressure was ranging from 30 to 60 Torr. A reactor containing the catalyst was placed at the end of the plasma quartz tube, about 2 cm below the end of the tube. The sine microwave waveform (at 915 MHz) is modulated in square pulses using a dedicated power supply. Accordingly, the kHz range frequency is related to the
modulation frequency, on top of the base frequency. In the present study, a 50% duty cycle was used: a pulse duration of 300 µs and a period of 600 µs (pulse frequency = 1.67 kHz). A more detailed description of the microwave set-up can be found elsewhere [15,19,32].

3.2. Characterization of plasma and dissociation products

The composition of the postdischarge was analyzed by a Bruker-450 gas chromatograph (GC) system equipped with a carbon molecular sieve column and a molecular sieve 5A column in series and connected to a thermal conductivity detector (TCD). Argon was used as a carrier gas in order to facilitate detection of H$_2$, CO$_2$, CO, H$_2$, and O$_2$ were successfully detected by GC. No hydrocarbons were detected by this current system. The microwave discharge was also characterized by optical emission spectroscopy (OES), which is a useful technique to study the large variety of chemical processes in plasma. The light emission from diatomic molecules is generally represented by a set of ro-vibrational emission bands in the studied spectral range (≈300–800 nm) [33]. The vibrational populations, gas temperature, densities of the various particles and radicals, vibrational temperature, and molecular dissociation degree can be obtained through the analysis of these emission bands (as well as the atomic emission lines). The corresponding experimental data is given in the previous Chapter and also can be found elsewhere [17, 18]. The present Chapter is mainly dedicated to the effects of plasma catalysis.

Figure 1. Schematic representation of surface-wave microwave set-up with the discharge diagnostic facilities used in this work.
3.3. Catalyst preparation

In the present work, we consider two different types of the catalyst treatment: conventional calcination and plasma treatment. Pellets (6 mm diameter and 1 mm thick) made from pure anatase powder (Sigma-Aldrich) were used as a catalyst support. We use plasma calcination as a replacement of the conventional calcination in the plasma-treatment process. Three types of catalyst were prepared by plasma processing and are denoted in this work as: NiO/TiO\textsubscript{2}(O\textsubscript{2}), NiO/TiO\textsubscript{2}(Ar), and NiO/TiO\textsubscript{2}(CO\textsubscript{2}) which corresponds to the catalysts pretreated with pure O\textsubscript{2} plasma, pure Ar plasma, and pure CO\textsubscript{2} plasma, respectively. The catalysts prepared by conventional calcination are referred to as, NiO/TiO\textsubscript{2}(Air-C), NiO/TiO\textsubscript{2}(Ar-C), for calcination in air and argon, respectively. A catalyst (NiO/TiO\textsubscript{2}(Ar-P)) was also studied with an Ar plasma for a longer time (similar to conventional calcination). The summary of the preparation conditions and Brunauer-Emmett-Teller (BET) surface areas is given in Table 1. More details about these two catalysts and the preparation can be found in the corresponding literature studies [15, 21].

3.4. Catalyst characterization techniques

The surface area of the catalyst was determined by nitrogen adsorption at a temperature of 77 K using BET analysis. At the same time, the analysis of the catalyst crystalline structure was conducted using X-ray diffraction (XRD) on the catalyst powder (Bruker advanced X-ray diffractometer (40 kV, 40 mA) with a Cu K\textalpha (0.154 nm) at a scanning rate of 2°/min from 20 to 90°). TEM (transmission electron microscopy) (Philips CM 200) images were used to characterize the morphology and crystalline size. The Raman spectra were collected using a Bruker spectrometer equipped with a 532-nm argon ion laser as the excitation source (the laser power was 2 mW). The spectrum range was from 40 to 1000 cm\textsuperscript{-1} and the spectral resolution was 0.5 cm\textsuperscript{-1}. Raman spectrometry was also used to detect the presence of defects in the crystal lattice. Ultra violet -to- visible (UV-VIS) diffuse reflectance spectra were obtained by a Cary spectrophotometer using BaSO\textsubscript{4} as a background at a resolution of 1 cm\textsuperscript{-1}. It was used

| Catalysts         | Description                                      | BET surface area (m\textsuperscript{2}/g) |
|-------------------|--------------------------------------------------|------------------------------------------|
| TiO\textsubscript{2} | Catalyst support (pure anatase)                  | 12                                       |
| TiO\textsubscript{2}(Ar) | Ar plasma treated for 40 minutes                | 12                                       |
| NiO/TiO\textsubscript{2}(O\textsubscript{2}) | O\textsubscript{2} plasma treated for 40 minutes | 15                                       |
| NiO/TiO\textsubscript{2}(CO\textsubscript{2}) | CO\textsubscript{2} plasma treated for 40 minutes | 19                                       |
| NiO/TiO\textsubscript{2}(Ar) | Ar plasma treated for 40 minutes                | 19                                       |
| NiO/TiO\textsubscript{2}(Air-C) | Conventional calcination in air for 2 hours | 18                                       |
| NiO/TiO\textsubscript{2}(Ar-C) | Conventional calcination in Ar for 2 hours | 17                                       |
| NiO/TiO\textsubscript{2}(Ar-P) | Ar plasma treated for 2 hours                  | 19                                       |

Table 1. Summary of various catalysts and pure TiO\textsubscript{2} applied for plasma-assisted catalytic conversion of CO\textsubscript{2}. 
4. Plasma-assisted catalytic conversion of CO₂

4.1. The influence of catalyst preparation method

In the field of catalysis, the nature of the catalyst in terms of its physical structure and chemical properties is a key factor for its effectiveness in a particular process. The activity of the adsorption sites provided by the catalyst determines the mechanisms that provide the lower energy reactive pathways [28]. The formation of these adsorption sites could be sensitively affected by catalyst preparation conditions, and the understanding of the chemical basis for this is essential. The nonthermal plasma, having highly energetic electrons and chemically reactive species (e.g., free radicals, excited atoms, ions, and molecules), may modify surface properties of the catalyst, enhance the dispersion of the supported metals, and even create disorder in their crystallite structure, depending on the nature of the gas phase, and the experimental treatments conditions. In this work, in order to study the relationship between the catalyst preparation method (its structure, morphology, and defect) and its efficiency in the CO₂ conversion, two different catalyst preparation methods, namely the conventional calcination and plasma treatment, were considered. Comparison of CO₂ conversion and energy efficiencies in the microwave discharge with and without catalysts as well as the effect of different preparation methods are investigated and presented in Figure 2. Thermally calcinated

![Figure 2](http://dx.doi.org/10.5772/67874)

**Figure 2.** CO₂ conversion and energy efficiencies are shown for the NiO/TiO₂ catalysts prepared by different methods (conventional calcination vs Ar plasma treatment), (a) Plasma only, (b) NiO/TiO₂ (Air-C), (c) NiO/TiO₂ (Ar-C), (d) NiO/TiO₂ (Ar-P). Flow rate = 2 slm; pressure = 30 Torr; SEI = 6.95 eV/molecule. pulse frequency = 1.67 kHz.
catalysts (NiO/TiO$_2$ (Air-C) and NiO/TiO$_2$ (Ar-C)) affect the CO$_2$ conversion rather insignificantly (in comparison with the plasma-only-assisted CO$_2$ dissociation). On the other hand, it is clear that the presence of plasma-activated catalyst (NiO/TiO$_2$(Ar-P)) significantly enhances the CO$_2$ conversion and energy efficiencies compare to the plasma-only-assisted decomposition. The CO$_2$ conversion and energy efficiencies increase from 23 to 43% and 10 to 18%, respectively, compared to the plasma only case. For better understanding of the influence of the different preparation methods on chemical and physical properties of the catalysts, the chemical composition and morphology of the catalysts were studied by XRD, BET analysis, TEM, and Raman spectroscopy.

The catalyst characterization based on XRD and TEM results shows that an Ar plasma treatment results in a very uniform distribution of nickel oxide on the TiO$_2$ and the possibility of a higher Ni$^{2+}$ concentration into the TiO$_2$ lattice. Raman spectroscopy was used in order to access the phase constitution and lattice defects of the catalyst support (TiO$_2$). All observed Raman peaks correspond to the characteristic Raman modes of the anatase TiO$_2$ structure [34, 35] (Figure 3). The peak shift in Raman spectra indicates that the vibrations of the TiO$_2$ lattice are considerably affected by the presence of Ni ions. The largest shift and broadening occur for the Ar plasma-treated catalyst (NiO/TiO$_2$(Ar-P)), as shown in Figure 3. The $E_g$ peaks are associated with the symmetric stretching vibration mode of O-Ti-O in TiO$_2$. This mode is very sensitive to local oxygen coordination surrounding the metal ions. According to a previous detailed analysis of Raman spectra of TiO$_2$ and doped-TiO$_2$ samples [15], the broadening and shifting of the main bands can be attributed to the presence of oxygen vacancies. These results indicate that plasma treatment can induce higher density of oxygen vacancies compared to the conventional calcination. To get a valuable insight into the catalyst activation process, a two-dimensional model of Ar plasma was developed in the modeling framework PLASIMO [36–38]. The results from the simulation imply the effective catalyst activation by the Ar metastable species (4s 1s$^3$ and 4s 1s$^3$ states) at the catalyst chamber region [21].

![Figure 3](image-url)

Figure 3. (a) Raman spectra of conventional calcined and plasma-treated catalysts and pure TiO$_2$; (b) full-width at half-maximum (FWHM) in $E_g$ modes in the different catalysts and TiO$_2$ support.
in plasma could lead to the unusual catalyst structure, formation of smaller particles, better dispersion, and stronger interaction between the catalyst and the support.

4.2. The influence of plasma gas for catalyst activation

As mentioned above, plasma plays an important role in tailoring the properties of a catalyst by the selective doping of a material with heteroatoms and its subsequent modification of the bandgap of a material which can affect both activity and selectivity [28]. However, modifications of the catalyst depend on the nature of plasma gas and experimental treatment conditions. The influence of preactivation of the catalyst using three different gases, namely O$_2$, Ar, or CO$_2$, on the CO$_2$ conversion and energy efficiencies as well as on the chemical and physical properties is studied in this section. The CO$_2$ conversion and energy efficiencies are not altered significantly in the presence of TiO$_2$ (Ar) support (without doping agent) compared to the plasma-only CO$_2$ decomposition case, as shown in Figure 4. The presence of NiO/TiO$_2$ (Ar) significantly increases the CO$_2$ conversion efficiency and energy efficiency almost twice from 23 to 42% and 9.6 to 17.5%, respectively, comparing with the plasma-only experiment results. However, the CO$_2$ conversion and energy efficiencies were not significantly affected when NiO/TiO$_2$ (O$_2$) and NiO/TiO$_2$ (CO$_2$) catalysts are used. To understand this effect, the analysis of different catalysts and supports was undertaken by XRD, BET, Raman spectroscopy, and UV-VIS spectroscopy. Similar crystallite sizes (around 46 nm) were observed for pure TiO$_2$ support, TiO$_2$ (Ar) and NiO/TiO$_2$ catalysts prepared by plasma [15]. On the other hand, the lattice constant c increases slightly with the addition of NiO and the largest increase

![Figure 4](http://dx.doi.org/10.5772/67874)

Figure 4. CO$_2$ conversion and energy efficiencies are shown for the NiO/TiO$_2$ catalysts prepared by plasma treatment with different gases (O$_2$, Ar, CO$_2$). Flow rate = 2 slm; pressure = 30 Torr; SEI = 6.95 eV/molecule. Pulse frequency = 1.67 kHz. Reprinted with permission from Ref. [15]. Copyright 2016 Elsevier.
is observed for NiO/TiO$_2$ (Ar) catalyst based on the comparison of the lattice values for the NiO/TiO$_2$ catalysts with pure TiO$_2$. The observed TiO$_2$ lattice expansion might imply that a few Ni$^{2+}$ ions are incorporated into the TiO$_2$ lattice (the ionic radius of Ni$^{2+}$ (0.83 Å) is larger than that of Ti$^{4+}$ (0.75 Å) [39]). As we have discussed in Section 4.1, the oxygen vacancies are probably mainly responsible for the blue and red shifts as well as for widening of the Raman peaks corresponding to $E_g$ modes. Together with the XRD results, we conclude the substitution of Ti$^{4+}$ ions by Ni$^{2+}$ distorts the TiO$_2$ lattice and generates oxygen vacancies to maintain charge neutrality. As it is noted from Figure 5 that the largest shift and widening occur for the Ar plasma-treated NiO/TiO$_2$ catalyst. This indicates that the Ar plasma treatment leads to a higher concentration of oxygen vacancies in the presence of NiO. In order to study the doping effect on the catalyst optical properties, the bandgap energy $E_g$ was calculated based on the UV-VIS spectra [15]. The obtained bandgap energies are 2.85, 2.79, 2.60, 2.49, and 2.17 eV for pure TiO$_2$, TiO$_2$ (Ar), NiO/TiO$_2$ (O$_2$), NiO/TiO$_2$ (CO$_2$), and NiO/TiO$_2$ (Ar), respectively. These results are consistent with the previous findings that the increasing of oxygen vacancies results in a narrowing bandgap energy for TiO$_2$ [40–43]. The pure TiO$_2$ support and the TiO$_2$ support pretreated with Ar plasma were also investigated. No influence of Ar plasma pretreatment on the TiO$_2$ support surface area and its crystalline structure as well as on the defect formation has been found. Therefore, it proves that an important condition for the formation of oxygen vacancies by means of Ar plasma pretreatment is the presence of Ni$^{2+}$ which leads to incorporation of the Ni$^{2+}$ ions in the TiO$_2$ lattice.

**Figure 5.** Raman spectra of TiO$_2$-supported NiO catalysts, TiO$_2$ (Ar), and pure TiO$_2$ support. The enlarged view of the intense $E_g$ peak is shown in the inset. Reprinted with permission from Ref. [15]. Copyright 2016 Elsevier.
Figure 6. Evolution of the conversion efficiency and energy efficiency of CO$_2$ as a function of the discharge pressure at different SEI.

Figure 7. Dependence of the CO$_2$ conversion and energy efficiencies on the NiO content in the NiO/TiO$_2$ catalysts treated by Ar plasma. Flow rate = 5 slm; pressure = 60 Torr; SEI = 2.9 eV/molecule. Pulse frequency = 1.67 kHz.
4.3. The influence of pressure and NiO content

First, the total pressure effect on the CO\textsubscript{2} decomposition was studied in the absence of catalyst in the range of 20–60 Torr at different SEI (Figure 6). As the pressure increases, our experimental data clearly show an improvement in CO\textsubscript{2} conversion, which is in good agreement with published results obtained in microwave plasma at low pressure [6]. These pressure-dependent results can be understood based on the increase of electron-neutral collision frequency in the plasma phase which favors the vibrational excitation of the asymmetric vibrational mode of CO\textsubscript{2} and leads to higher CO\textsubscript{2} conversion efficiency [6]. In addition, $\eta$ drops when SEI increases whereas at the same time $\chi$ increases. This trade-off between $\chi$ and $\eta$ has been observed in other CO\textsubscript{2} experiments as well [6, 10]. From this result it is clear that $\chi$ does not increase to the same extent as SEI does. The effect of nickel contents on the CO\textsubscript{2} conversion and energy efficiencies was investigated at 60 Torr. Increasing the nickel content from 2.5 to 10 wt% results in an increase of the CO\textsubscript{2} conversion and energy efficiencies up to 42% (Figure 7). However, at higher nickel content, there is no significant change in the conversion of CO\textsubscript{2}. Figure 8 shows the X-ray powder diffraction patterns of NiO/TiO\textsubscript{2} catalysts for different NiO contents. For nickel contents ranging from 2.5 to 10 wt%, very weak diffraction peaks were observed at $2\theta = 43.5^\circ$, corresponding to the (200) crystal planes of face-centered cubic NiO lattice. As a result of further increase of the NiO content, NiO diffraction peaks become clearly visible. This increase in the NiO content does not increase the population of
Ti⁴⁺ substituted by Ni²⁺ in the TiO₂ lattice. Further investigations are necessary to obtain a better understanding of the role of NiO.

5. Discussion

Among several factors giving significant improvements in both χ and η values, the following two groups are of special interest: the nature of plasma activation gas and NiO content on the TiO₂ surface. Compared to the plasma-only case, the CO₂ conversion and energy efficiencies were multiplied by a factor of 2.1 by selecting the appropriate catalyst preparation methods (plasma) and the NiO content. In agreement with previous study [19] on the CO₂ conversion efficiency and the corresponding relative Raman shifts of the three E_g modes observed in NiO/TiO₂ (O₂, Ar, CO₂), NiO/TiO₂ (Air-C), NiO/TiO₂ (Ar-C), NiO/TiO₂⁹ (Ar-P), and in TiO₂ (Ar) (shown in Figures 3 and 5), the present results indicates that the higher activity observed on Ar plasma-treated NiO/TiO₂ catalysts could be linked to the formation of surface oxygen vacancies.

It is well known that photocatalysts can be activated through the formation of electron-hole (e⁻-h⁺) by photoillumination of the catalyst surface. TiO₂-based catalysts are extensively studied for CO₂ photoreduction because of the adequate bandgap of TiO₂ (3.2 eV for anatase phase). Liu et al. [44] have shown that the mechanism of activation and subsequent reduction of CO₂ involves the participation of protons and electron transfer. However, with a pure TiO₂ (Ar) support, no difference in the CO₂ conversion efficiency compared to plasma only case was found, as shown in Figure 8. Hence, the observed improvements in our work could not be explained by the possible photoreduction of CO₂ induced by the plasma radiation.

Nowadays, the fabrication of oxygen-deficient surfaces has increasing interest, which is an important strategy for improving the photocatalytic CO₂ dissociation [44–46]. Oxygen vacancy (Vo) has been considered as the active site for the adsorption and activation of CO₂. The electron attachment to a CO₂ molecule adsorbed by an oxygen vacancy leads to the formation of a transient negative ion CO₂⁻ that decomposes to give CO and fill the Vo with an oxygen atom between a Ni²⁺ and Ti⁴⁺ ions (the oxygen atom is O_{br}) [15]. The threshold energy for this one-electron process is found to be 1.4 eV [44, 47]. The dissociative electron attachment (DEA) of CO₂ observed on the catalytic surface can occur also in the gas phase [6]. However, the maximum value of the DEA cross-section is low (∼10⁻²² m²) and the threshold energy of 5–10 eV [6] is high compared to the threshold energy of 1.4 eV [47] for the DEA at the catalyst surface. So, the DEA process for CO₂ occurs more easily for adsorbed CO₂ on the TiO₂ defective surface than for CO₂ in the gas phase. As the activity of our catalyst remains constant over time, a regeneration of Vo is necessary [15].

A study of CO₂ conversion efficiency in different gas mixtures in the presence of NiO/TiO₂ (Ar-P) has also been investigated [19, 21]. It was found that the CO₂ conversion efficiency was also enhanced independently of the mixture when NiO/TiO₂ (Ar-P) is used. A synergy between the catalyst and plasma is clearly demonstrated. The CO₂ conversion and energy efficiencies in 90% CO₂–10% H₂O increase from 30 to 52% and 12.5 to 22%, respectively,
when the activated catalyst is combined with the plasma. We found previously that oxygen vacancies enhanced the CO₂ conversion in the plasma-catalytic dissociation of pure CO₂ [15]. The electrons supplied by the plasma enhance the dissociative attachment of CO₂ on the oxygen-deficient catalyst surface. We suggest that the following processes can be distinguished on the catalyst surface in the plasma-catalytic CO₂/H₂O conversion (Figure 9). CO₂ could be reduced to CO via reaction with H radical or direct dissociation by healing the oxygen vacancy sites [1, 48, 49]. The active species OH and adsorbed oxygen atom react with each other to form atomic hydrogen and O₂, resulting in the regeneration of oxygen vacancy. The oxygen vacancy can also be regenerated via the recombination on the surface of a bridging oxygen atom with a gaseous oxygen atom [15, 50, 51]. The presence of atomic

Figure 9. A tentative mechanism of plasma-catalytic CO₂ decomposition on the catalyst surface.
O peaks and water dissociation product (OH and H) peaks in the emission spectrum of CO\textsubscript{2}/H\textsubscript{2}O microwave discharge was detected and reported in our previous study [17]. The plasma supplies energy to the catalyst surface and thus additionally enhances the recombination process. The regeneration of oxygen vacancies can be confirmed by our experimental results [19, 21]. The CO\textsubscript{2} conversion and energy efficiencies remain more or less constant over time, at least for 150 minutes of continuous operation. Hence, a synergistic effect between the plasma and the NiO/TiO\textsubscript{2}(Ar-P) catalyst leads to significantly improved CO\textsubscript{2} conversion and energy efficiencies, which can be attributed to the dominant surface reactions driven by the plasma.

A two-dimensional $\chi$-$\eta$ chart compares the energy efficiency for CO\textsubscript{2} conversion using different types of non-thermal plasmas (Figure 10). The influence of the NiO content, as well as of the other effects studied (such as H\textsubscript{2}O effect) in this and previous work [19, 21] is clearly visible. Selected results on CO\textsubscript{2} conversion taken from the literature are given for comparison. The data points corresponding to our study look very promising. The highest energy efficiency of about 90% was reported by Asisov et al. in microwave plasma in supersonic flow [14]. The corresponding CO\textsubscript{2} conversion was only about 10%, which is significantly lower than that (42%) obtained in this work for pure CO\textsubscript{2} decomposition with a catalyst. The highest reported conversion efficiency was about 80% while the energy efficiency was less than 6% [18]. A balance between CO\textsubscript{2} conversion and energy efficiencies in the plasma processing of CO\textsubscript{2} is especially important for the development of an efficient and cost-effective process. The combination
of the microwave plasma and the Ar plasma-treated NiO/TiO$_2$ catalyst leads to significant enhancement in the CO$_2$ conversion and energy efficiencies of the plasma process, as well as the balance between them.

6. Conclusions

This work critically summarizes the effects of various plasma operating conditions and catalyst preparation methods on the CO$_2$ conversion and its energy efficiency in a surface-wave sustained microwave discharge. As a result, one can conclude that the catalyst preparation method has a significant impact on the chemical and physical properties of the catalysts, which in turn strongly influence CO$_2$ conversion and energy efficiencies. Ar plasma treatment is supposed to result in a higher density of oxygen vacancies and a very favorable distribution of nickel oxide on the TiO$_2$ surface comparing with the classical calcination in air or in argon.

The oxygen vacancies are the key factor explaining the catalytic activities in CO$_2$ decomposition. These vacancies are stabilized by the presence of Ni$^{2+}$ ions in the anatase lattice. The plasma pretreatment allows to both create oxygen vacancies, and to incorporate the nickel into the TiO$_2$ lattice to stabilize these vacancies. The dissociative electron attachment of CO$_2$ at the catalyst surface enhanced by the oxygen vacancies and plasma electrons can explain the observed increase of CO$_2$ conversion as well as the energy efficiency. A mechanism is proposed that explains the observed plasma-catalyst synergy, which has led to improved CO$_2$ conversion and energy efficiencies.

At the same time, there is still a room for further improvement of the CO$_2$ conversion and energy efficiencies through the optimization of the plasma parameters (e.g., high pressure and high flow rate) and the modification of catalysts (e.g., loading different metal nanoparticles on TiO$_2$). In addition, it is very promising to explore the possibility of H$_2$O-CO$_2$ as an alternative system to conversion of carbon dioxide to value-added compounds, such as syngas, methanol, and formic acid. We expect that the results presented in this study will provide useful insights into the plasma-assisted CO$_2$ conversion in the presence or absence of catalysts, which may be used for greenhouse gas conversion in industry.

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