Adsorption Study of Rhodamine –B Dye on Plant (Citrus Leaves)

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Abstract:
The current research includes the adsorption of Rhodamine-B Dye on the surface of Citrus Leaves using the technique of UV. Vis spectrophotometer to determine data of quantitative adsorption at various contact time, ionic strength, PH and temperature conditions. As a function of temperatures 25,35,45,55 °C, the desorption phenomenon was examined, and the results showed that Rhodamine-B adsorption Citrus leaves rose with increasing temperatures on the surface (endothermic process). Using various NaCl solution concentrations, the effect of ionic strength on adsorption has also been studied. Increasing the importance of ionic strength has been shown to improve the amount of adsorption of Rhodamine-B on citrus leaves at constant temperatures. The quantity of Rh-B dye that was adsorbed on the citrus leaves was increased with increasing the PH of the solution in rang 3-7. Then the results were subjected to the practical results obtained with the use isotherms of the Langmuir, Freundelsh, and Temkin. The results can be shown that the isotherm suitable for adsorption applies to Ferndelsh and Temkin. The thermodynamic functions (the amount of change in enthalpy, the amount of change in entropy and the amount of change in the free energy of Gibbs) were also studied, and it was found that the thermodynamic results of the adsorption process of the Rhodamine-B dye on the powder of citrus leaves are endothermic and an automatic adsorption process.

Keywords: Rhodamine B, Adsorption, Citrus Leaves, Isotherm, Thermodynamic.

Introduction:
Today, all societies suffer from water pollution problems due to the drainage of industrial effluents into surface water contaminated with toxic dyes from a wide range of industries, such as textiles, leather, sheet, printing, food, beauty products, paint, rubber, plastics, pesticides, wood preserving chemicals and pharmaceutical-industries. Water pollution is defined as water contamination and the alteration of water's physical, chemical and biological characteristics. It entails the dumping of any sewage or commercial waste or any other fluid. In water, gaseous or solid compounds are likely to produce a nuisance by means of their physical presence, odor or taste. Such discharge can damage and harm Water for the public as far as domestic hygiene is concerned. They are concerned with commercial, manufacturing and agricultural life.

The discharge of dye effluents also imparts color from the sources for miles downstream to the receiving stream. This color is aesthetically unacceptable and therefore decreases the penetration of light into the water, thus limiting the efficacy of photosynthesis in aquatic organisms and thereby making an adverse impact on their growth.

Natural compounds like the toxins, surfactants, some synthetic dyes and their precursors have varying biochemical and morphological influences on plants and animals.

A dye is a colored compound that binds chemically to the substrate to which it is added. This separates dyes from pigments that do not bind to the substance they color chemically. The dye is normally added in an aqueous solution and will need a mordant to enhance the dye's intensity on the fiber. Dyes can cause skin inflammation, accelerated heartbeat, vomiting, diarrhea, tissue necrosis, and shock.

Dyes can be categorized in different ways: (e.g., according to the chemical constitution, application class, and end-use). The Principal categorizations of dyes are based on the natural
fibers to which they will be applied and the chemical of each dye character. In order to form chemical bonds, reactive dyes react with fiber molecules\(^6\). Dyes may be categorized as anionic: direct, acid, and reactive dyes; cationic: simple dyes; and non-ionic: disperse dyes\(^7\).

Rhodamine B (RH-B) is a synthesized organic colorant used as biological stain in paper, draw, food, leather, pharmacy, cosmetic, labs, etc. This colorant has a carcinogenic and toxic effect, hence the use in food and cosmetics is currently limited\(^8\).

For the removal of dyes, many methods involving biological, physical, chemical and related technology have been published. Adsorption is the treatment of choice among the many therapies and provides the best outcomes as it can be used to extract various forms of dyes simultaneously\(^9\).

In brief, by physical isolation, chemical methods or biological degradation, dye recovery from waste water can be accomplished. Adsorption, oxidation, biological therapy, electrochemical treatment, membrane filtration, and coagulation flocculation are several major techniques that are commonly used for the removal of dyes\(^10\).

This method of adsorption is used to remove chemical chemicals, hazardous substances, dyes, and other products that conventional techniques do not remove. Adsorption on certain textures, such as natural materials with porous properties such as activated charcoal, zeolite, silica gel, porous clays, some agricultural residues, and plant parts\(^11\), may largely eliminate these components\(^11\).

In contrast to other processes, adsorption of farm waste has a low cost and high performance benefit\(^12\)\(^-\)\(^15\).

In this work, the adsorption of RhB dye from aqueous solutions on Citurs leaves powder, which are considered environmentally friendly and of low-cost. The efficiency and behavior of adsorbents were studied at different concentration and temperatures.

**Material and Methods:**

**Chemicals and Apparatus**

Rhodamine B molecular dye formula \(\text{C}_{28}\text{H}_{31}\text{ClN}_{2}\text{O}_{3}\) and chemical composition Fig.1 is supplied by the HIMEDIA Company while citrus leaves were collected from the University of Baghdad gardens. The leaves of Citrus aurantium contain organic chemicals, including alkanes, alcohols, carboxylic acids, esters, soaps, etc.\(^16\)

Sodium hydroxide was supplied from BDH with purity 99.5%. Hydrochloric acid was provided from BDH with purity 37%, and Sodium chloride with purity 99.5%. Thermostated Shaker water bath model JEITECH (BS-11)-Germany was used to shake the aqueous solutions of dye with adsorbents, while Remi Centrifuge Model R&C. Bombay-India was utilized to separate the adsorbents from the aqueous solution of dye. UV-Vis spectrophotometer model PC1650 Doube beam – Shimadzu was used to record absorbance of Rhodamine B before and after adsorption experiments. pH-mater model Hanna-HI-8417-England was used to adjust the pH of the solutions.

![Figure 1. Structure of Rh-B](image1.png)

**Absorbed Substance**

By dissolving specifically weighted colour in refined water to the centralization of 8 ppm, the rohdamine B dye stock structure was set up, was used without aiding cleaning at \(\lambda_{\text{max}} 542\) nm, Fig.2 indicates the dye absorption spectrum. The arrangements were accomplished by weakening the structure of the color stock to different starting focuses in precise ranges from Fig.3 shows the calibration curve, which was used to convert the absorbance value of dye to concentration according to Beer-Lambert law.

![Figure 2. Absorption spectrum of Rhodamine B dye](image2.png)
Preparation of Citrus Leaves Powder
After they were picked, citrus leaves were washed and the leaves were cleaned with plain water from the soil clinging to them, then washed with purified water, dried in an oven (Hot Air Oven LTD, LDHAM, England at 90 °C for 30 min, and ground with an electric grinder after drying, sifted with a sieve 75μm and stored away from moisture.

Batch Adsorption Experiment
To prepare 100ppm from the dye, 0.1 gm of RhB was dissolved in 1L of distilled water. Different dye concentrations were prepared for RhB at concentrations of 2-14 ppm. 0.1M of HCl and 0.1M of NaOH solutions were used to adjust the pH according to the media in which the adsorption process was conducted takes place and to reduce the error resulting from changing the acidity of the media on the sample size and prepare NaCl solution 0.1, 0.01, 0.001 M. The adsorption was observed by evaluating the convergence of Rhodamine B in a double bar Visible spectrophotometer configuration, at λmax 542 nm. The percentage of dye expulsion and volume of rhodamine B adsorbed on adsorbent was determined using the following equation to determine the-quantity of dye adsorbed:

\[ Q_e = \frac{(C_0 - C_e)V}{W} \]  

Where the main and equanimity (mg/l) amounts of dye, respectively, are C₀ and Cₑ. At the moment of equanimity (mg/g), Qₑ is the sum of dye adsorbed on the adsorbent. V is the volume (l) of solution and W is the mass of the adsorbent (g).

Factors Effecting Adsorption System
Dosage of adsorbent. Original dosage of adsorbent 0.1, 0.2, 0.3, 0.4, and 0.5 g, 8 ppm concentration. Both decanters were put in a water bath fitted with a temperature 25°C shaking system 100 rpm and with a period 25°C, 30 min.

Contact Time of Adsorption System: Study adsorption process equilibrium at 25°C for an adsorbent sum of 0.4 g of citrus leaves powder in 25 ml. The sample was withdrawn from the shaker for regular time intervals of 10,20, 30, 40, 50, 60, and 70 min.

Temperature: At temperatures of 25, 35, 45, and 55 °C in an indoor regulator paired with a shaker, the adsorption process was repeated in the same way. The experiments were carried out by taking 25 ml of Rhodamine B arrangement with different color convergences varying from different beginning convergences.of 2 - 12 ppm.

Ionic strength: The adsorption experiments carried out at various concentrations of NaCl 0.1, 0.01 and 0.001M at 25 °C with 25 ml of dye concentration.

pH: Adsorption experiments were conducted at 25°C at pH 3, 7 and 11 with 25 ml of 8 ppm color focus. Hydrochloric acid 0.1M and sodium hydroxide 0.1M were used to change the acidic and basic pH of the solution.

Results and Discussion:
Adsorbent Dosage: -
The effect of the adsorbent dosage on the removal of RhB dye was explored in different quantities of powdered citrus leaves, while keeping the initial dye concentration 8ppm in 30 min at 25 °C. Fig.4 shows the effect of dosage of citrus leaves powder on the adsorption percentage removal of rhodamine B dye.

The surface mass that provided the lowest absorbency (highest adsorption value) was picked and was 0.4 gm from the surface of the Citrus aurantium leaf powder.

Although the rise in the weight of the powdery orange leaf powder surface contributes to an increase in the amount of active sites primed for dye adsorption on the adsorbent surface, this means an increase in dye adsorption/removal percentage to increase the surface effectiveness.
Determine the equilibrium time of adsorption systems (Contact Time)

To evaluate the time necessary for the adsorption of Rhodamine B on the surface of the powdered Citrus leaves to achieve the equilibrium state at temperature 298 k and concentration \( C_0 = 8 \) ppm for coloring and with a steady adsorbent weight 0.4 g . Fig.5 indicates that the dye (RB) is quickly absorbed when the dye is adsorbed into the powdered citrus leaf in the first 30 minutes. This arises because the active centers on the adsorbent surface are empty, but the adsorption slows down as these active centers are filled with dye particles.

Effect of Temperature

The effect of the adsorption temperature of the simple rhodamine B dye on the surface of the powder of Citrus aurantium leaves in the thermal range 25, 35, 45, and 55 °C was analyzed by fixing the adsorbent mass 0.4 g from the surface of the powder of Citrus aurantium leaves and at the concentrations 2,4,6,8,10,12 ppm of the dye Fig.6.

Effect of pH

PH has a major effect on the removal of dye from aqueous solutions, as it influences the adsorbent's superficial charge, the degree of dye ionization, the isolation of the groups added to the adsorbent's active site and the molecular structure of the dye. Fig.7

Effect of Ionic Strength

The effect of ionic strength of adsorption of rhodamine B dye Fig.8-10 on the surface of Citrus leaves powder 0.4 g was studied by adding 0.1, 0.01, and 0.001M salt in a temperature range 298-318 K contact time: 30 min. It can be clarified that the cationic sodium ions enhance the potential of the dye to interact on the surface of the adsorbents.
bearing in mind the overarching objective of how the additional set of sodium chloride courses of action triggers increased rivalry at first glance between rhodamine B particles and electrolyte particles over the expulsion aims.19.

Figure 8. The effect of Ionic strength on the Adsorption of Rhodamine B Dye at a Temperature (298 K)

Figure 9. The effect of Ionic strength on the Adsorption of Rhodamine B Dye at a Temperature (308K)

Adsorption Isotherm
The adsorption on the surface of Citrus aurantium leaf powder of rhodamine B dye from its aqueous solution was tested at varying amounts and at different temperatures 298, 308, 318, 328K. The findings demonstrate that Ferndelsh and Temkin are protected by an isotherm appropriate for adsorption.

- Langmuir Adsorption Isotherm
The Langmuir isotherm was calculated by the following Eq.2

\[ \frac{C_e}{Q_e} = \frac{1}{q_{\text{max}}} K_L + \left(\frac{1}{q_{\text{max}}}\right) C_e \] ………… 2

If \( C_e \) is the adsorbate equilibrium concentration (mg/L), \( Q_e \) is the sum of adsorbed per gram of equilibrium adsorbent (mg/g), \( q_{\text{max}} \) is the overall adsorption power (mg/g) and \( K_L \) is Langmuir is the adsorption energy constant Fig.11.
Figure 11. Langmuir linear relationship between Ce/Qe and Ce for Rhodamine B dye on Citrus leaf powder at a temperature range a) 298 b) 308 c) 318 d) 328 K.

**Freundlich Isotherm**

The isothermal equation was determined as follows

\[
\log Q_e = \log K_f + \frac{1}{n} \log C_e \quad \ldots \ldots \quad 3
\]

Where \( Q_e \) is adsorbed dye (mg/g) is the indicator of Rh B, \( C_e \) is the equilibrium grouping of in the structure, color (mg/l), \( K_f \) and \( 1/n \) are constants that integrate the elements influencing the adsorption cap and the force of adsorption, individually, separately Fig.12.

Figure 12. Freundlich Linear relationship between \( \log Q_e \) and \( \log C_e \) for rhodamine B dye on citrus leaves powder at a temperature range: a) 298 b) 308, c) 318 and d) 328 K.
**Temkin Isotherm**

Temkin discovered the impacts of certain circuitous sorbate/adsorbate relationships on adsorption isotherms and proposed that the warmth of adsorption of a large number of particles with scope decreases directly in the sheet in view of these relations. The Temkin isotherm was used after the Eq.4 Fig.13.

\[ Q_e = B_T \ln K_T + B_T \ln C_e \]  

Where

\[ B_T, K_T = \text{the equilibrium binding constant (1/g)} \]

![Figure 13. Temkin Isotherm relationship between Qe and lnCe for Rhodamine B dye on Citrus aurantium leaf powder at a temperature range: a) 298 b) 308, c) 318 and d) 328 K](image)

As can be seen from the calculated results in the Tab.1 the R coefficient figures were akin to 1, and for all isotherms, which means that a reasonable one is available in the relationship between Qe and ln C, the corresponding Ce at various temperatures that suggest the useful calculations of the constants. Adsorption of the Freundlich and Temkin have better explained isothermal for isotherm shows.

**Table 1. The factors of langmuir, Freundlich and Temkin condition for the adsorption of Rhodamine B dye on Citrus aurantium leaf powder**

| Tem. (K) | q_{max} | K_L | R_L | R^2 | n | K_F | R^2 | K_T | b | R^2 |
|----------|---------|-----|-----|-----|---|-----|-----|-----|---|-----|
| 298      | 0.4467  | 0.0991 | 0.558 | 0.5975 | 0.667 | 24.027 | 0.9719 | 1.12 | 9224.02 | 0.9705 |
| 308      | 0.3325  | 0.1263 | 0.497 | 0.5305 | 0.653 | 23.768 | 0.9622 | 1.124 | 9291.41 | 0.9672 |
| 318      | 0.746   | 0.08  | 0.61 | 0.4207 | 0.777 | 16.436 | 0.9765 | 1.076 | 10782.43 | 0.9851 |
| 328      | 0.2355  | 0.141 | 0.4699 | 0.7152 | 0.5884 | 33.136 | 0.9899 | 1.266 | 9595.33 | 0.9516 |

**Thermodynamic of Adsorption Process**

The effect of thermodynamic factors on the adsorption of rhodamine B on the surface of Citrus aurantium leaf powder at the temperature range 298-318K was studied. Where the study of the effect of temperature on the adsorption process enables us to specify the values of the thermodynamic functions (enthalpy, entropy, free
energy) for all adsorption forms, as shown by the Van't Hoff equation, the adsorption equilibrium constant ($K_{eq}$) was calculated as (ln $Q_e$/$C_e$) against the temperature as (1/T) by means of plotting\textsuperscript{20}. The outcomes are recorded in Tab.2 and Fig.14 for the importance of these functions in understanding the adsorption phase system.

$$\Delta G = -RT\ln K_{eq} \quad \text{...... 5}$$

$$\ln K_{eq} = -\Delta H/RT + \Delta S/R \quad \text{...... 6}$$

$$\Delta G = \Delta H - T\Delta S \quad \text{...... 7}$$

where $K_{eq}$ is adsorption equilibrium constant and $R$ is the gas constant.

![Figure 14. Van’t Hoff plot adsorption of Rh-B dye on citrus leaves powder](image)

| $C_o$ | $T$ (K) | $\Delta H$ | $\Delta G$ | $\Delta S$ |
|-------|---------|------------|------------|------------|
| 6ppm  | 298     | +8.703     | -10.877    | 65.7       |
|       | 308     |            | -11.088    | 64.26      |
|       | 318     |            | -11.605    | 63.87      |
|       | 298     |            | -10.877    | 89.07      |
| 8ppm  | 308     | +15.666    | -11.088    | 86.864     |
|       | 318     |            | -11.977    | 86.928     |

### Conclusion:

The adsorption of Rhodamine B dye by citrus leaves powder has been determined. The adsorption was measurement that the dose of adsorbents was 0.4g. The amount adsorption of Rhodamine B dye increasing in $PH = 3.7$ but decreased with increasing $PH$. The best contact time within of 30 min for citrus leaves powder. The ionic strength was increasing with an increasing concentration of Sodium chloride. The adsorption process supplied Freundlich and Temkin isotherm models. The thermodynamic factors: $\Delta G$ spontaneous, $\Delta H$ was endothermic in nature, in addition, the values of $\Delta S$ were positive (higher).

### Authors' declaration:

- **Conflicts of Interest:** None.
- We hereby confirm that all the Figures and Tables in the manuscript are mine ours. Besides, the Figures and images, which are not mine ours, have been given the permission for republication attached with the manuscript.
- **Ethical Clearance:** The project was approved by the local ethical committee in University of Baghdad.

### Authors' contributions:

A M.: Preparation, analysis, curation of data, Study, Writing- main draft.
A M. Far.: Conceptualization, analysis, Resources, Visualization, Editing Writing. The final paper was read and authorized by all writers.
H. I.: Resources, Visualization, Editing Writing.

### References:

1. Ebrahim A S, Tarek A F, Marwa N., Mohamed D. A Comparative Study for Adsorption of Methylene Blue Dye from Wastewater onto Three Different Types of Rice Ash. J. Pharm. Appl. Chem.[Internet]. Jan. 2018; 4 (1): 99-107.
2. Khattri, S D, Singh M K. Colour removal from dye wastewater using sugar cane dust as an adsorbent. Adsorption science & technology. April 1999; 17(4): 269-282.
3. Sughosh Madhav, Arif Ahamed, Amit K. Singh, Jyoti Kushwahawa, Joginder Singh Chauhan, Sanjeev Sharma, et al. Water pollutants: sources and impact on the environment and human health. Sensors in Water Pollutants Monitoring: Role of Material. Springer, Singapore, 2020;43-62.
4. Sivathasan J. Preparation of Clay-dye pigment and its dispersion in polymers. School of Civil, Environmental and Chemical Engineering RMIT University: September 2007; P141.
5. Tran H Tu, Phan Thi Ngoc Cam, Le Van Trong Huy, Mai Thanh Phong, Hoang Minh Nam, Nguyen Huu Hieu. Synthesis and application of graphene aerogel as an adsorbent for removal of dyes from water. Mater Lett. December 2018;238: 134-137.
6. Mohamed B, Rachid H, Mohammed K, Mohammed A, Omar C, Abderrahim B, et al. Textile finishing dyes and their impact on aquatic environs. Heliyon. November 2019;5(11): e02711.
7. Bhoi S K. Adsorption Characteristics of Congo Red Dye onto PAC and GAC based on S/N ratio:A Taguchi Approach. National Institute of Technology, Rourkela In partial fulfilment of the requirements of Bachelor of Technology (Chemical Engineering). India: May 2010.1-32.
8. Paul H A, Guy D F, Didier R, Kopoin A, Patrick D. Photocatalytic degradation of Rhodamine B dye with TiO2 Immobilized on SiC foam using full factorial design. Appl Water Sci [Internet]. August 2020; 10(207): 1-9.
9. D Li, Yan J, Liu Z. Adsorption kinetic studies for removal of methylene blue using activated carbon prepared from sugar beet pulp. Int. J. Environ. Sci. Technol. May 2016; 13(7): 1815–1822.
10. Akil A, Situ H, Sing C C, Asma K, Waseem A W, Rajeev Kumar, et al. Recent Advances in New Generation Dye Removal Technologies: Novel Search of Approaches to Reprocess Waste Water. RSC Advances. 2015; 5(39): 1-58.
11. ChunxiangLin, YuemingQiu, JingFan, MinWang, Lannei Ye, Yifan Liu, et al. Fabrication of photo-responsive cellulose based intelligent imprinted material and selective adsorption on typical pesticide residue. Eng Sci Chem. August 2020; 394: 124841.
12. Sandeep P, Sukanta M, Dilip K C, Chetan K D, Sanjay K S, Kumar A. Vinod. Effect of coal fly ash's particle size on U adsorption in water samples and thermodynamic study on adsorption. Environ Chem. Ecotoxicol. February 2020; 2: 32-38.
13. Khudhair. A. Al-Rudaini. Adsorption Removal of Rhodamine-B Dye from Aqueous Solution Using Rhamnus Stone as Low Cost Adsorbent. J N U S. March 2017; 20(1): 32-41.
14. H.Yang, M.Lu, D.Chen, R.Chen, L.Li, W.Han. Efficient and rapid removal of Pb2+ from water by magnetic Fe3O4@MnO2 core-shell nanoflower attached to carbon microtube: adsorption behavior and process study. J Colloid Interface Sci. December 2019; 563: 218-228.
15. Grégorio C, Eric L, Lee W, Nadia M-C. Conventional and non-conventional adsorbents for wastewater treatment. Environ Chem Lett. March 2019; 17(1): 195–213.
16. Haipeng Z, Mengjun C, Huan W, Zhenhua W, Jiajing C, Liu F, et al. Transcriptomic and metabolomic analyses provide insight into the volatile compounds of citrus leaves and flowers. BMC Plant Biol. January 2020; 20(7):1-14.
17. Aseel M. Aljeboree, Abbas N. Alshirifi, Ayad F. Arkaim. Kinetics and equilibrium study for the adsorption of textile dyes on coconut shell activated carbon. Arab J Chem. May 2017; 10(2): S3381-S3393.
18. Somayeh R, Abbas R, Muhammad N Z, Syed S S, Shahin A. Synthesis and characterization of MgO supported Fe–Co–Mn nanoparticles with exceptionally high adsorption capacity for Rhodamine B dye. J. Mater. Res. Technol. July 2019; 8(5): 3800-3810.
19. Deepak P, Shikha S, Pardeep S. Removal of methylene blue by adsorption onto activated carbon developed from Ficus carica bast. Arab. J. Chem. February 2017; 10(1): S1445–S1451.
20. Barsharani Priyadarshini, Tanasswini Patra, TapasRanjan Sahoo. An efficient and comparative adsorption of Congo red and Trypan blue dyes on MgO nanoparticles: Kinetics, thermodynamics and isotherm studies. J. Magnes Alloy. October 2020;9(2): 478-488.
دراسة الامتزاز لصبغة رودامين B على النبات (أوراق الحمضيات)

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الخلاصة:

تضمن الدراسة الحالية امتزاز صبغة الرودامين B على سطح أوراق النارنج باستخدام تقنية مقياس الأشعة فوق البنفسجية حيث تم تحديد البيانات الكيمياء للأمتزاز (زمن التلمس، القوة الأيونية، درجة الحموضة، ودرجة الحرارة). تمت دراسة تأثير درجة الحرارة على عملية الامتزاز عند النطاق (25، 35، 45، و55) درجة مئوية حيث أظهرت النتائج بأن عملية امتزاز الرودامين B تزداد بزيادة درجة الحرارة مما يدل على أن عملية الامتزاز عملية ماصة للحرارة، وباستخدام تراكيز مختلفة من محلول كلوير الصوديوم تم دراسة تأثير القوة الأيونية على عملية الامتزاز. فالأمتزاز يثبت أن زيادة القوة الأيونية تعمل على تحسين كمية امتزاز الرودامين B على أوراق النارنج عند درجات حرارة ثابتة. وأعطت النتائج أن كمية امتزاز صبغة الرودامين B تزداد بزيادة PH للحلول في المدى (3-7). ثم خضعت نتائج الامتزاز لنماذج متوازي الحرارة (Freundlich, Langmuir, Temkin), حيث تم تحديد معاملات كل من Langmuir, Freundlich, Langmuir, Temkin, و Temkin و Freundlich و Langmuir و Freundlich و Temkin و Freundlich معاملات الديناميكية الحرارية (مقدار التغير في المحتوى الحراري، ومقدار التغير في الانتربيا و مقدار التغير في الطاقة الحرارية لـ Gibbs) ووجد أن النتائج الديناميكية الحرارية لعملية امتزاز صبغة Rhodamine-B معاملات متوازي الحرارة وعملية الامتزاز العملية ثالثية.

الكلمات المفتاحية: الامتزاز، أوراق الحمضيات، الأيزوثرم ، رودامين B، الديناميكا الحرارية.