Research Article

Intermetallic Cu₅Zr Clusters Anchored on Hierarchical Nanoporous Copper as Efficient Catalysts for Hydrogen Evolution Reaction

Hang Shi 1, Yi-Tong Zhou 1, Rui-Qi Yao 1, Wu-Bin Wan 1, Qing-Hua Zhang 2, Lin Gu 2, Zi Wen 1, Xing-You Lang 1, and Qing Jiang 1

1Key Laboratory of Automobile Materials (Jilin University), Ministry of Education, and School of Materials Science and Engineering, Jilin University, Changchun 130022, China
2Beijing National Laboratory for Condensed Matter Physics, The Institute of Physics, Chinese Academy of Sciences, Beijing 100190, China

Correspondence should be addressed to Xing-You Lang; xylang@jlu.edu.cn and Qing Jiang; jiangq@jlu.edu.cn

Received 3 July 2019; Accepted 22 October 2019; Published 20 February 2020

Designing highly active and robust platinum-free electrocatalysts for hydrogen evolution reaction is vital for large-scale and efficient production of hydrogen through electrochemical water splitting. Here, we report nonprecious intermetallic Cu₅Zr clusters that are in situ anchored on hierarchical nanoporous copper (NP Cu/Cu₅Zr) for efficient hydrogen evolution in alkaline medium. By virtue of hydroxygenated zirconium atoms activating their nearby Cu-Cu bridge sites with appropriate hydrogen-binding energy, the Cu₅Zr clusters have a high electrocatalytic activity toward the hydrogen evolution reaction. Associated with unique architecture featured with steady and bicontinuous nanoporous copper skeleton that facilitates electron transfer and electrolyte accessibility, the self-supported monolithic NP Cu/Cu₅Zr electrodes boost violent hydrogen gas release, realizing ultrahigh current density of 500 mA cm⁻² at a low potential of -280 mV versus reversible hydrogen electrode, with exceptional stability in 1 M KOH solution. The electrochemical properties outperform those of state-of-the-art nonprecious metal electrocatalysts and make them promising candidates as electrodes in water splitting devices.

1. Introduction

Global climate change and environmental pollution caused by excessive CO₂ emission have raised tremendous demands for new energy sources with renewability and sustainability [1–3]. Hydrogen, as a high-density energy carrier in clean and sustainable energy framework of water-hydrogen interconversion, has long been considered as one of the most promising alternatives to traditional fossil fuels for meeting future energy needs [4, 5]. Electrochemical water splitting driven by renewable electricity from plentiful solar, wind, and hydro resources is an attractive and feasible approach to large-scale production of hydrogen [4, 6, 7] showing genuine potential in replacing commercial technologies (such as catalytic steam methane reforming, partial oxidation, and coal gasification) that are neither environmentally friendly nor cost-effective because of vast CO₂ emission and harsh high-temperature requirements. During the process of water electrolysis at room temperature, the hydrogen evolution reaction (HER) is a crucial step but suffers from a sluggish kinetics [6, 8], particularly in alkaline electrolytes [7, 9–14], which urgently calls for robust and efficient catalytic materials to minimize the HER overpotential at high current densities [6–16]. Highly conductive precious metals, including platinum (Pt) and ruthenium (Ru), are at present the best HER catalysts by virtue of their appropriate hydrogen-binding energy (HBE) values that are conducive to the adsorption of reactive hydrogen intermediates (H⁺) and their recombination into hydrogen molecules [6, 17, 18]. However, their scarcity, expensive cost, and poor durability essentially impede the practical use of Pt and Ru for scalable hydrogen production even in various nanostructured alloys or composites with high specific surface areas [19–23]. To address these issues, it is crucial to explore Earth-abundant...
nonprecious metal electrocatalysts with superior HER catalysis towards a successful hydrogen economy [7, 13, 15, 16, 24]. In view that monometallic transition metals (e.g., Fe, Co, Ni, Mo, and W) generally have too weak or too strong HBE values [6, 17, 25], considerable research efforts have been focused on developing their compounds (such as carbides [26, 27], nitrides [28, 29], phosphides [30, 31], sulfides [32, 33], oxides [34], or selenides [35]). Initial strides have been made in improving their HER catalysis by nanostructuring to increase the number of active sites or/and making use of C, N, P, S, or Se heteroatoms to enhance intrinsic activity via adjusting the HBE values to balance the H⁺ adsorption/desorption [6, 13, 15, 16]. Unfortunately, few of these materials operate at low overpotentials due to either insufficient catalytic activity or poor electron transfer caused by electron localization of electrovalent bonds. Even though nanostructuring is expected to ameliorate electron transferability by shortening the distance from electroactive to conductive materials, immobilizing these low-dimensional nanostructures on planar current collectors to construct macroscale electrodes inevitably leads to durability and efficiency issues [33, 35–37], unsatisfying practical requirements of large-scale hydrogen production, i.e., low working overpotential and long-term stability at high current density [26, 29, 30, 38–40]. Therefore, it is desirable to design and develop novel low-cost electrocatalytic materials with a rational electrode architecture that not only withstands violent gas evolution but has abundant high active sites and facilitates fast electron transfer and electrolyte transport to minimize the overpotential.

In this study, we report self-supported monolithic electrocatalytic materials, which are composed of high-density intermetallic Cu₅Zr clusters in situ integrated on bicontinuous nanoporous Cu current collector (NP Cu/Cu₅Zr), for highly efficient production of hydrogen in alkaline media. Therein, the constituent Cu₅Zr clusters serve as active sites to boost the sluggish reaction kinetics of the HER as a result of hydroxylation of surface Zr atoms activating their nearby Cu–Cu bridge sites with an appropriate HBE conducive to the H⁺ adsorption/desorption. The specific current density of Cu₅Zr clusters reaches 3.1 mA cm⁻², ~300-fold higher than that of bare Cu, at the potential of ~200 mV versus a reversible hydrogen electrode (RHE). Because bicontinuous nanoporous Cu skeleton is conductive and steady enough to facilitate fast electron transfer and withstand violent hydrogen gas evolution, the NP Cu/Cu₅Zr electrodes exhibit superior HER catalysis in 1 M KOH solution, with a low onset overpotential of ~30 mV, a small Tafel slope of ~68 mV dec⁻¹ and a long-term durability at high current densities. Furthermore, it takes the NP Cu/Cu₅Zr only a low potential ~280 mV versus RHE to achieve a current density of as high as 500 mA cm⁻², outperforming the state-of-the-art nonprecious metal-based catalysts. These impressive electrochemical properties make them promising alternatives to precious metal-based catalysts for efficient water electrolysis in basic solutions.

2. Results and Discussion

2.1. DFT Theoretical Calculations. Cu is a classic metal with high conductivity but usually exhibits poor catalytic activity. Here, we devise a highly efficient HER copper catalyst, which is modified by Zr via the formation of Cu–Zr intermetallic clusters (Figure 1(a)). Owing to the large electronegative difference between Cu and Zr, there forms structurally ordered Cu₅Zr intermetallic compound [41–44], one of the most thermodynamically stable phases with a face-centered cubic (fcc) structure in Cu–Zr system, in the Cu-rich component region [41–44]. On the basis of Cu₅Zr(111), first-principles density functional theory (DFT) calculations demonstrate that Zr atoms can effectively activate their surrounding negative-valence Cu atoms. As shown in partial density of states (PDOS) of Cu₅Zr(111) (Figure 1(b)), the center of d bands of the Cu₅Zr is shifted up toward the Fermi energy relative to bare Cu(111). Although surface Cu atoms tend to be oxidized into Cu(OH)₂ layer in the ambient surroundings, they are reduced to the metallic state in the HER potential range because of the Cu²⁺/Cu equilibrium potential of 0.61 V at pH = 14 [45]. This is demonstrated by the DFT simulation on the Cu₅Zr(111), on which except for the Zr atoms, the Cu atoms do not adsorb hydroxyl group (HO⁻) at 0 V vs. RHE (supplementary Figure S1). Distinguished from the hollow-site adsorption of Cu(111) (Figure 1(c)), the incorporation of Zr enables the Cu–Cu bridge sites of the Cu₅Zr(111) surface to adsorb hydrogen atoms, with a calculated HBE per H atom of ~0.44 eV (Figure 1(d), supplementary Figure S2b and Table S1). While it is thermodynamically favorable to adsorb HO⁻ on the surface Zr atoms (Cu₅Zr–OH) in the HER potential range (Figure 1(e) and supplementary Figure S1), the hydroxylation does not adjust the optimal adsorption sites except for slightly weakening the HBE value to ~0.36 eV (supplementary Figure S2d and Table S1). Such unique configuration is more favorable for the Cu₅Zr–OH(111) to accelerate the HER kinetics than the clean Cu₅Zr(111), suggesting a superior HER catalytic activity. This expectation is verified by the calculations of free-energy diagram for the HER reaction paths on the Cu₅Zr(111) facets with or without the hydroxylation of surface Zr atoms. As shown in Figure 1(f), the Cu₅Zr–OH(111) has an energy barrier of ~0.54 eV for the rate-limiting step of H⁺ adsorption/combination, much lower than the value of the Cu₅Zr (~0.67 eV).

2.2. Preparation and Structural Characterization. With an aim at developing highly efficient and robust HER catalytic materials that should comprise sufficient available surface with high catalytic activity and have a rational structure with fast electron transfer and reliable steadiness [36], we construct self-supported monolithic NP Cu/Cu₅Zr electrodes with a hierarchical nanoporous architecture by a facile and scalable dealloying strategy [46, 47]. Therein, both the nanoporous microstructure and the components are controlled by hereditary effect of precursor Cuₓ₀.₅ZₓAl₁ₓ alloys (x = 0, 1, 3, 5, and 20 at%) alloys during the dealloying processes [46]. Resembling the multiphase-structured Cuₓ₀.₅ZₓAl₁ₓ alloy (supplementary Figure S3a), the Cuₓ₀.₅ZₓAl₁ₓ (x = 1, 3, and 5 at%) alloys are primarily composed of several-hundred-nanometer-wide α-Al and CuAl₂ phases with a quasiperiodic distribution, as illustrated by a scanning
electron microscope (SEM) backscattered electron image for the Cu$_{17}$Zr$_3$Al$_80$ alloy (Figure 2(a)) and the corresponding energy-dispersive X-ray spectroscopy (EDS) elemental mappings (Figures 2(b)–2(d)). X-ray diffraction (XRD) characterization reveals that in addition to α-Al and CuAl$_2$ phases, there are intermetallic Cu$_5$Zr [48, 49], Al$_{9.83}$Zr$_{0.17}$ and Al$_3$Zr clusters that are embedded in Cu-rich and α-Al phases, respectively, with the incorporation of Zr components (supplementary Figures S3b–d and S4) [48, 49]. The multiphase structure enables the product of NP Cu/Cu$_5$Zr with a bimodal nanoporous structure, as that of bare NP Cu (supplementary Figure S5), in a two-step chemical dealloying process: the rapid dissolution of the α-Al phase embedded with Al$_{9.83}$Zr$_{0.17}$ and Al$_3$Zr to form large channels and the selective etching of the less-noble Al component from CuAl$_2$ alloy to generate small nanopores and expose Cu$_5$Zr clusters. As shown in a representative SEM image for the NP Cu/Cu$_5$Zr electrode dealloyed from the precursor Cu$_{17}$Zr$_3$Al$_80$ alloy (Figure 2(e)), it displays a uniform bimodal nanoporous structure consisting of large channels with ~300 nm in width and small nanopores with the diameter of ~18 nm (supplementary Figure S6). The small nanopore structure is further verified by N$_2$ adsorption/desorption isotherm (supplementary Figure S7a), which signifies a size distribution with maximum centered at ~18 nm (supplementary Figure S7b). Figure 2(f) shows a typical high-resolution transmission electron microscope (HRTEM) image of Cu/Cu$_5$Zr ligaments, where high-density intermetallic Cu$_5$Zr clusters with diameter of ~3-6 nm can be identified by their distinct contrasts and lattices. Atomic resolution high-angle annular-dark-field scanning transmission electron microscope

Figure 1: DFT calculation of Cu$_5$Zr intermetallic compound. (a) Schematic illustrating Cu matrix embedded with Cu$_5$Zr clusters. Cu atoms in the matrix and Cu$_5$Zr clusters in orange and golden, Zr atoms in cyan, O in red, and H in magenta, respectively; (b) Partial density of states (PDOS) of Cu atom in Cu(111), Zr atom in Zr(111), and Cu and Zr atoms in Cu$_5$Zr(111). (c–e) Optimized atomic structures of facets of Cu(111) (c), Cu$_5$Zr(111) (d), and Cu$_5$Zr-OH(111) (e) with hydrogen adsorption (H*). (f) Free-energy diagram of the Tafel route for the HER on Cu$_5$Zr and Cu$_5$Zr-OH surfaces.
(HAADF-STEM) image viewed along the [111] zone axis illustrates that these Cu₅Zr clusters are exposed and in situ anchored on the NP Cu skeleton with a seamless interfacial structure (Figure 2(g)), enhancing the stability of electroactive Cu₅Zr clusters and their electron transfer with conductive Cu skeleton. The hybrid structure of NP Cu/Cu₅Zr is further verified by the corresponding fast Fourier transform (FFT) patterns (inset of Figure 2(g)) and the XRD patterns (Figure 2(h)) with two sets of diffraction characteristics [48]. In the XRD patterns, the weak diffraction peaks at 2θ = 37.0, 43.7, and 45.7 are assigned to the (220), (311), and (222) planes of fcc Cu₅Zr clusters (JCPDS 40-1322) (inset of Figure 2(h)) [48], apart from the obvious ones attributed to monometallic Cu (JCPDS 04-0836). X-ray photoelectron spectroscopy (XPS) analysis reveals the presence of Zr and Cu with an atomic ratio of 92/5, in addition to a little residual Al according to inductively coupled plasma mass spectroscopy (ICP-MS) and EDS (supplementary Figure S8 and Table S2). The deviation of Zr/Cu ratio from the initial value in the
precursor alloy confirms the fact that those Al$_{9.83}$Zr$_{0.17}$ and Al$_2$Zr are removed with the dissolution of α-Al phase. Because of electron donation from Zr to Cu (inset of Figure 2(j)), the Zr 3d XPS spectrum shows the chemical state of Zr$^{2+}$ in addition to the Zr$^{4+}$ that is due to the oxidation of surface Zr atoms (Figure 2(i)) [50]. This is in contrast with the observation of solid Cu/Cu$_x$Zr bulk alloy with negligible Cu$_x$Zr clusters exposed at the surface (supplementary Figures S9a and b), of which the primary chemical state of Zr$^{2+}$ is attributed to the electron transfer from Zr to Cu (supplementary Figure S9c). As a result, Cu exhibits the chemical state of Cu$^{0}$ in addition to Cu$^{0}$ and Cu$^+$ that correspond to the Cu atoms locating at the internal and surface of Cu ligaments, respectively (Figure 2(j) and supplementary Figure S9d) [34, 51]. Similar dealloying processes also take place in the precursor Cu$_{20-x}$Zr$_x$Al$_{80}$ alloys with $x=1$ (Cu$_{19}$Zr$_1$Al$_{80}$) or 5 (Cu$_{15}$Zr$_5$Al$_{80}$), which gives rise to almost the same bimodal nanoporous structure (supplementary Figures S10a and S11a) except for their distinct Cu/Zr compositions of 97/1 (supplementary Figure S10b) and 80/16 (supplementary Figure S11b), respectively. In contrast, the precursor alloy with $x=20$ (i.e., Zr$_{20}$Al$_{80}$) produces NP Zr with a single-modal nanoporous structure during the chemical dealloying (supplementary Figure S12).

2.3. Electrochemical Characterizations. To assess the electrocatalytic HER activity of NP Cu/Cu$_x$Zr, all nanoporous catalytic materials are directly used as working electrodes for electrochemical measurements in N$_2$-saturated 1 M KOH aqueous electrolyte, based on a classic three-electrode configuration with a saturated calomel electrode (Hg/Hg$_2$Cl$_2$, SCE) and a graphite rod as the reference electrode and the counter electrode, respectively. All potentials are calibrated with respect to the RHE. Supplementary Figure S13 shows the first four polarization curves of NP Cu/Cu$_x$Zr electrode. In agreement with DFT simulation result, there takes place an irreversible reduction reaction of Cu hydroxide to metallic Cu at the potential of -0.17 V vs. RHE during the first HER measurement. This enables metallic surface Cu atoms to take part in the subsequent HER. In the plot of geometry area-normalized current density versus iR-corrected potential, Figure 3(a) shows the representative HER polarization curves of self-supported NP Cu/Cu$_x$Zr hybrid electrode and NP Cu and Zr monometallic electrodes at a scan rate of 1 mV s$^{-1}$, comparing with that of commercially available Pt/C, a benchmark HER catalyst, immobilized on NP Cu current collector. By virtue of the presence of Cu$_x$Zr clusters, the NP Cu/Cu$_x$Zr electrode exhibits remarkably enhanced HER catalytic activity, with an onset overpotential (~30 mV) much lower than the values of NP Cu (~140 mV) and Zr (~260 mV), which have either too low or too high HBE. The superior HER catalysis of NP Cu/Cu$_x$Zr is also demonstrated by the ultralow Tafel slope of ~68 mV dec$^{-1}$, compared with the values of bare NP Cu (~172 mV dec$^{-1}$) and NP Zr (~249 mV dec$^{-1}$) (Figure 3(b)). With the low onset overpotential and small Tafel slope due to high-density and high-activity Cu$_x$Zr clusters, the NP Cu/Cu$_x$Zr rapidly reaches the geometric area-normalized current density of 100 mA cm$^{-2}$ at the potential of -155 mV versus RHE, which is higher than that of NP Cu-supported Pt/C electrode (~199 mV) although the former exhibits slightly lower activity than the latter at low overpotential range from 0 to -86 mV (Figure 3(a)). The fast kinetics of NP Cu/Cu$_x$Zr electrode is further justified by electrochemical impedance spectroscopy (EIS) measurements in the frequency range of 100 kHz to 10 mHz. As shown in the Nyquist plots (Figure 3(c)), the EIS spectra of all nanoporous catalysts display characteristic semicircles in the high- and middle-frequency range, of which the diameters represent their charge transfer resistances ($R_{CT}$). Based on the equivalent circuit (inset of supplementary Figure S14), the NP Cu/Cu$_x$Zr is demonstrated to have the lowest $R_{CT}$ value (~12 $\Omega$) at the overpotential of 0.1 V among these catalysts with nanoporous architecture including the NP Cu-supported Pt/C (~41 $\Omega$) (inset of Figure 3(c)), in agreement with their observations in Tafel plots (Figure 3(b)). The impressive HER catalysis results from the Cu$_x$Zr clusters, on which the hydroxygenated Zr surface atoms not only adsorb strongly HO$^∗$ but adjust the H$^+$ adsorption at the Cu-Cu bridge sites, essentially boosting the HER kinetics as a result of the balance between the water dissociation and the H$^+$ adsorption/desorption [52]. The unique behavior differs from the NP Cu/ZrO$_2$ hybrid electrode, on which the constituent ZrO$_2$ nanoparticles serve as strong HO$^*$ adsorption sites but the Cu skeleton is too weak to adsorb the H$^+$, as well as the fully electrooxidized NP Cu/Cu$_x$Zr (NP Cu/Cu$_x$Zr-EO), on which there are only strong HO$^*$ adsorption sites. This leads to the distinct HER activity of the pristine NP Cu/Cu$_x$Zr, with the overpotential at 100 mA cm$^{-2}$ being ~100 and ~250 mV lower than those of the NP Cu/ZrO$_2$ and NP Cu/Cu$_x$Zr-EO (Figure 3(d)) in addition to the exceptional Tafel slope (Figure 3(e)).

To further attest that intermetallic Cu$_x$Zr clusters are indeed active sites for the HER, we investigate the NP Cu/Cu$_x$Zr electrode after removing the exposed Cu$_x$Zr clusters in H$_2$SO$_4$ solution. Supplementary Figure S15a shows a typical SEM image of the H$_2$SO$_4$-treated Cu/Cu$_x$Zr electrode, displaying almost the same nanoporous architecture as the pristine one. Nevertheless, the XRD characterization demonstrates that most of the Cu$_x$Zr are removed except for a few clusters imbedded in the Cu ligaments (supplementary Figure S15b), as confirmed by the Cu/Zr/Al component change of 98.9/0.4/0.7 (supplementary Figure S15c) as well as the Zr 3d and Cu 2p XPS spectra with little amount of Cu$^{0}$ and Zr$^{2+}$ (supplementary Figures S15d and e). Figure 3(f) compares the HER polarization curve of the H$_2$SO$_4$-treated NP Cu/Cu$_x$Zr with that of the pristine one. Owing to the absence of Cu$_x$Zr clusters at the surface, the H$_2$SO$_4$-treated NP Cu/Cu$_x$Zr exhibits remarkably degraded HER activity, with the overpotential at the current density of 100 mA cm$^{-2}$ increasing to ~350 mV, which implies the indispensable role of Cu$_x$Zr clusters in achieving superior HER performance of NP Cu/Cu$_x$Zr electrodes. This is also reflected by the linear dependence of the overpotentials at 100 mA cm$^{-2}$ on the surface areas, which are adjusted by controlling the dealloying time (supplementary Figures S16a–c) or precursor alloy
components (supplementary Figure S17). Although there is always residual Al component in the ligaments, it is not expected to significantly influence the HER catalytic activity of NP Cu/Cu/Zr. As shown in supplementary Figure S16d, the NP Cu/Cu/Zr electrodes with different residual amount of Al exhibit almost the same specific activities at the overpotential of 100 mV. To evaluate quantitatively the HER activity of Cu/Zr clusters, here we approximately estimate the specific current density ($J_{Cu/Zr}$) of the electroactive Cu/Zr clusters according to
the equation [53] $j_{scu5zr} = (\frac{j_{scu5zr}A_{scu5zr}}{A_{cu5zr}} - j_{scu}A_{scu})/A_{scu5zr}$, where $j_{scu5zr}$ and $j_{scu}$ are the specific current densities of NP Cu/Cu5Zr and the constituent Cu surface, $A_{scu5zr} = A_{scu} + A_{scu5zr}$ with $A_{scu}$ and $A_{scu5zr}$ being the specific surface areas of Cu and Cu5Zr components in the NP Cu/Cu5Zr, respectively. Considering that the NP Cu/Cu5Zr and bare NP Cu electrodes have almost the same nanoporous architectures with primary (111) surface (Figure 2(e) and supplementary Figure S5), the value of $j_{scu}$ refers to that of NP Cu while the $A_{scu}$ is calculated through the broad anodic peak between 0.5 and 0.71 V corresponding to the formation of CuO with a charge density of 360$\mu$C cm$^{-2}$ (supplementary Figure S18) [25]. With an assumption that the Zr$^{2+}$ corresponds to the surface Zr atom, $A_{scu5zr}$ is evaluated according to $A_{scu5zr} = cA_{scu}$ with $c$ being the surface Zr$^{2+}$/Cu component ratio. At the potential of $-200$ mV versus RHE, $j_{scu5zr}$ reaches $\sim3.1$ mA cm$^{-2}$, $300$-fold enhancement relative to Cu. This enlists NP Cu/Cu5Zr to outstand among the best surface, Aous solution (supplementary Figure S17a). It takes the NP densities of NP Cu/Cu5Zr and the constituent Cu lower than the values of NP Cu97Zr1Al2 and Cu80Zr16Al4.

To simulate the industrial product of H$_2$ at large current density, we perform a galvanostatic stability test at various current densities from 10 to 500 mA cm$^{-2}$ in 1 M KOH electrolyte (Figure 4(c)). When performing at each current density for 8 h, the NP Cu/Cu5Zr electrode exhibits stable potentials, indicating the outstanding durability. Even at the current density of as high as 500 mA cm$^{-2}$, the NP Cu/Cu5Zr electrode can withstand the violent H$_2$ gas evolution to maintain the original nanoporous structure in the long-term durability measurement (inset of Figure 4(c)). The impressive catalytic stability can be attributed to the electrochemical stability of Cu5Zr clusters and the robustness of nanoporous Cu structure for hydrogen bubble release. Moreover, the practical hydrogen production rate in Figure 4(d) agrees well with the theoretical value, revealing that the Faradaic efficiency reaches approximately 100%.

### 3. Conclusions

In summary, we have developed monolithic and hierarchical nanoporous Cu/Cu5Zr hybrid electrodes as cost-effective and robust catalytic materials towards the HER in alkaline environments. As a consequence of evident electron donation from Zr to Cu triggered by their large electronegative difference, the zirconium atoms essentially activate their nearby Cu/Cu5Zr bridge sites. While the hydroxylation of surface Zr atoms facilitates the HO$^-$ adsorption and adjusts an appropriate hydrogen-binding energy conducive to the adsorption/desorption of reactive hydrogen intermediates, the constituent Cu5Zr intermetallic clusters serve as stable HER catalysts with exceptionally high activity of $\sim3.1$ mA cm$^{-2}$ ($300$-fold enhancement relative to bare Cu) at the potential of $-200$ mV versus RHE. Associated with the unique architecture featured with steady hierarchical nanoporous Cu current collector that not only facilitates electron transfer to the constituent Cu5Zr clusters but also endures the violent gas evolution, the NP Cu/Cu5Zr electrode exhibits superior HER electrocatalytic activity and durability in 1 M KOH solution relative to their corresponding monometallic catalysts with similar architecture, such as the NP Cu and Zr electrodes. More impressively, the NP Cu/Cu5Zr only takes a low potential of $-280$ mV versus RHE to achieve the current density of as high as $300$ mA cm$^{-2}$, outperforming the state-of-the-art nonprecious HER catalysts. The outstanding electrochemical properties enable the NP Cu/Cu5Zr to show genuine potential as an attractive alternative to precious metal-based catalysts for large-scale and efficient hydrogen production through water electrolysis.

### 4. Methods

#### 4.1. Fabrication of Self-Supported Nanoporous Catalyst Electrodes

Catalyst electrodes with a nanoporous structure were fabricated by a facile alloying/dealloying procedure. Precursor Cu$_{20-x}$Zr$_x$Al$_{80}$ ($x = 0, 1, 3, 5, and 20$ at%) alloys were produced by arc melting pure Cu, Zr, and Al under an argon atmosphere. After cutting into thin sheets with thickness of $\sim400$ $\mu$m and further polishing by sandpapers (600, 1200, and 2000 Cw), nanoporous catalyst electrodes were prepared by chemical dealloying in a N$_2$-purged 6 M KOH aqueous electrolyte at 70°C until there is no gas to be produced. To identify the electroactive sites that enhanced the Zr atoms, the nanoporous Cu/Cu5Zr catalyst electrode was further treated in a N$_2$-purged 0.5 M H$_2$SO$_4$.
aqueous electrolyte at 70°C to remove surface Zr atoms. All these catalysts were rinsed in ultrapure water (18 MΩ) for multiple times to remove substance in nanoporous structure. Commercially available Pt/C catalyst with weight of 1.5 mg (20 wt%, Johnson Matthey) was mixed into an 0.05 wt%, Sigma-Aldrich) solution containing isopropanol (20%) and water (80%) to form Pt/C ink under rigorous sonication. 100 μL Pt/C ink was drop-cast onto NP Cu electrode (2 mm × 5 mm × 0.4 mm) to prepare the Pt/C catalyst electrode for electrochemical measurements.

4.2. Structural Characterization. Field-emission scanning electron microscope (JSM-6700F, JEOL, 15 keV) equipped with X-ray energy-dispersive spectroscopy (EDS) was employed to characterize microstructure and chemical composition of nanoporous catalysts. High-resolution transmission electron microscopy (HRTEM) and scanning transmission electron microscopy (STEM) characterizations were performed on a field-emission transmission electron microscope (JEM-2100F, JEOL, 200 kV) and a field-emission transition electron microscope (JEM-ARM200CF, JEOL) operated at 200 kV and equipped with double spherical-aberration correctors for both condenser and objective lens, respectively. X-ray diffraction patterns of all nanoporous catalyst electrodes were collected from a D/max2500pc diffractometer with a monochromated Cu Kα radiation. X-ray photoelectron spectroscopy (XPS) analysis was conducted on a Thermo ECSALAB 250 with an Al anode. Charging effects were compensated by shifting binding energies based on the adventitious C 1s peak (284.8 eV). Atomic ratio of elements was analyzed using ICP-MS (Thermo electron). N2 adsorption/desorption isotherms were collected at 77 K by a Micromeritics (ASAP 2020 Plus) system.

4.3. Electrochemical Measurements. All nanoporous catalyst electrodes were directly used as working electrodes for electrochemical measurements, which were performed in a classical three-electrode setup with a graphite rod as the counter electrode and a saturated calomel electrode (Hg/Hg2Cl2, SCE) as the reference electrode. The HER polarization curves of nanoporous catalyst electrodes were collected at a scan rate of 1 mV s⁻¹ in a N2-saturated 1 M KOH aqueous solution at room temperature. The reference electrode was calibrated to the reversible hydrogen potential (RHE). Electrochemical impedance spectroscopy (EIS) analysis was performed at various overpotentials with frequency 0.01 to 100,000 Hz with 5 mV amplitude. To evaluate the electrochemical surface area
Her stability tests of nanoporous Cu/Cu5Zr electrode were conducted at the current densities of 10, 100, 200, 300, 400, and 500 mA cm⁻² in 1 M KOH solution, respectively, for 8 h. In addition, a potential cycling in a window of -0.2 and 0.1 V was also performed for 30,000 cycles with a scan rate of 100 mV s⁻¹ in a N₂-saturated 1 M KOH solution. The hydrogen production via the HER experiments were further conducted at the current density of 20 mA cm⁻² for NP Cu/Cu5Zr electrode (0.35 cm⁻²) in a typical H-type cell separated by a proton exchange membrane, and the generated gas was analyzed by gas chromatography (GC-2014).

4.4. DFT Calculation. All spin-unrestricted DFT calculations were carried out using the Vienna Ab initio Simulation Package (VASP) with the Perdew, Burke, and Ernzerhof (PBE) functional. The projector augmented wave (PAW) method was adopted to describe the electron-ion interactions. According to our careful convergence tests, the cutoff of plane wave basis sets was set to 400 eV. A Monkhorst-Pack grid of 4 × 4 × 1 and 12 × 12 × 3 k-points were used for geometry optimization and electronic structure calculations, respectively. A first-order Methfessel-Paxon smearing of 0.2 eV was applied to accelerate electronic convergence for geometry optimization. The electronic structure and geometry optimization convergence tolerance were 10⁻⁵ eV, 0.005 eV Å⁻¹, respectively. The transition states were obtained using the Climbing Image Nudged Elastic Band (CI-NEB) method. The convergence tolerance of NEB image was 0.05 eV Å⁻¹. The Cu(111), Cu₅Zr(111), and single-Zr-atom-doped Cu(111) and Zr(111) were modeled by using five layers, the bottom two layers were fixed while others were fully relaxed. A vacuum layer of 15 Å was used to avoid the unwanted interaction between the slab and its period images. The optimized structures and the charge density difference plots were illustrated with VESTA software.

The reaction free energy (ΔG) was calculated according to \( \Delta G = \Delta E + \Delta E_{ZPE} - T\Delta S \), where \( \Delta G \) is the reaction energy, \( \Delta E_{ZPE} \) is the zero-point energy, \( T \) is the absolute temperature, and \( \Delta S \) is the change of entropy. \( \Delta E_{ZPE} \) was calculated by \( \frac{1}{2} \int F(\omega)\hbar\omega d\omega \), where \( h \) is the Planck constant, \( F(\omega) \) is the phonon density of states. \( \Delta S_H \) is considered as \(-1/2 S^0_{H2} \) due to the fact that the vibrational entropy in the adsorbed state is small meaning, where \( S^0_{H2} \) is the entropy of H₂ in the gas phase at standard conditions. In addition, the adsorption energy (\( E_{M-H} \) denoting the corresponding species) was determined by the equation \( E_{M-H} = E_M - E_H - E_{M-H} \), where \( E_T \) is the total energy of the catalyst with species, \( E_M \) and \( E_H \) are the energies of the isolated catalyst and the species, respectively.

Conflicts of Interest
The authors declare no competing financial interests.

Authors’ Contributions
X.Y.L. and H.S. conceived and designed the experiments. H.S., R.Q.Y., W.B.W., and Z.W. carried out the fabrication of materials and performed the electrochemical measurements. Q.H.Z. and L.G. contributed to microstructural characterizations. Y.T.Z. and Q.J. performed the DFT calculation. X.Y.L., Q.J., H.S., and Y.T.Z. wrote the paper, and all authors discussed the results and commented on the manuscript. H.S. and Y.T.Z. contributed equally to this work.

Acknowledgments
This work was supported by the National Natural Science Foundation of China (Nos. 51871107 and 51631004), the Top-Notch Young Talent Program of China (W02070051), the Chang Jiang Scholar Program of China (Q2016064), the Program for JLU Science and Technology Innovative Research Team (JLUSTIRT, 2017TD-09), the Fundamental Research Funds for the Central Universities, and the Program for Innovative Research Team (in Science and Technology) in University of Jilin Province.

Supplementary Materials
Figure S1: relative energies of HO⁺ adsorptions on Cu₅Zr at different potentials of 1.80, 1.31, 0.61, and 1.00 V versus RHE. Figure S2: (a, b) atomic schematic showing Cu₅Zr with hydrogen atom adsorption at five possible sites before (a) and after optimization (b). (c, d) Atomic schematic illustrating Cu₅Zr-OH with hydrogen atom adsorption at four possible sites before (c) and after optimization (d). Figure S3: XRD patterns of precursor Cu₂₀₋₅ZrAl₈₀ alloys with x = 0 (a), 1 (b), 3 (c), and 5 at% (d). The line patterns show reference cards 25-0012 for CuAl₂ (violet line), 65-2869 for α-Al (dark yellow line), 40-1322 for Cu₃Zr (pink line), 26-0041 for Al₉.₈₅Zr₀.₁₅ (orange line), and 48-1385 for Al₃Zr (olive line) according to JCPDS. Figure S4: typical SEM image of precursor Cu₁₇Zr₅Al₈₀ alloy and the corresponding EDS elemental mapping. Figure S5: typical SEM image of bare NP Cu that is fabricated by chemical dealloying Cu₂₀Al₈₀. It displays a bimodal and bicontinuous nanoporous architecture. Figure S6: distributions of small and large pore sizes for the NP Cu/Cu₅Zr electrodes dealloyed from Cu₁₇Zr₅Al₈₀ alloy. Figure S7: N₂ adsorption/desorption characterization for the NP Cu/Cu₅Zr electrode dealloyed from Cu₁₇Zr₅Al₈₀ alloy. (a) The N₂ adsorption/desorption isotherm from NP Cu/Cu₅Zr with BET specific area of 45.4 m²/g. (b) Small pore size distribution of NP Cu/Cu₅Zr. Figure S8: EDS spectrum of NP Cu/Cu₅Zr electrode that is fabricated by chemical dealloying Cu₁₇Zr₅Al₈₀ in 6 M KOH solution. Figure S9: (a, b) XRD patterns (a) and EDS spectrum (b) of solid Cu/Cu₅Zr bulk alloy with the Cu/Zr component of 94.85/5.15, the same as that of NP Cu/Cu₅Zr electrode. (c, d) High-resolution Zr 3d (c) and Cu 2p (d) XPS spectra for the solid Cu/Cu₅Zr bulk electrode. Figure S10: typical SEM image (a) and EDS spectrum (b) of NP Cu/Cu₅Zr electrode that is fabricated by chemical dealloying Cu₁₇Zr₅Al₈₀ in 6 M KOH solution. Figure S11: typical SEM image (a) and EDS spectrum (b) of NP Cu/Cu₅Zr electrode that is fabricated by chemical dealloying Cu₁₇Zr₅Al₈₀ in 6 M KOH solution. Figure S12: typical SEM image of NP Zr electrode that is fabricated by chemical dealloying Zr₂₀Al₈₀. It displays a single-modal and bicontinuous...
nanoporous architecture. Figure S13: the first four polarization curves of NP Cu/Cu5Zr electrode in 1 M KOH with scan rates of 5 mV s\(^{-1}\). The high-valence copper is reduced to metallic copper at -0.17 V. Figure S14: comparison of charge transfer resistances (RCT) and intrinsic resistance (Rs) for NP Cu/Cu5Zr, NP Cu-supported Pt/C, NP Cu, and NP Zr. Inset: equivalent circuit according to which their EIS spectra are analyzed. Figure S15: (a) typical SEM image of NP Cu/Cu5Zr electrode after H\(_2\)SO\(_4\) treatment. (b) Comparison of XRD patterns of NP Cu/Cu5Zr before and after H\(_2\)SO\(_4\) treatment. (c) EDS spectrum of NP Cu/Cu5Zr electrode after H\(_2\)SO\(_4\) treatment. (d) Zr 3d (d) and Cu 2p (e) XPS spectra of NP Cu/Cu5Zr after H\(_2\)SO\(_4\) treatment. Figure S16: (a, b) typical HER polarization curves (a) and CV curves (b) of NP Cu/Cu5Zr electrodes that are fabricated by chemical dealloying precursor Cu\(_{17}\)Zr\(_3\)Al\(_{80}\) alloy for 3 h, 5 h, and 7 h at room temperature in 6 M KOH solution, respectively. Therein, the Al component changes from 8.7% to 3.4%. (c) Zr 3d (c) and Cu 2p (d) XPS spectra of NP Cu/Cu5Zr electrodes that are prepared by chemical dealloying for 3, 5, and 7 h, respectively. Figure S17: (a) comparison of polarization curves of NP Cu/Cu5Zr electrodes with different Cu/Zr/Al compositions, which are fabricated from precursor Cu\(_{20-x}\)Zr\(_x\)Al\(_{80}\) alloys with \(x = 1, 3, 5\). Scan rate of 1 mV s\(^{-1}\). Electrolyte: 1 M KOH. (b) The overpotential at the current density of 500 mA cm\(^{-2}\) for NP Cu/Cu5Zr electrodes as a function of specific surface area. (d) Electrochemical surface area-normalized specific activity of the NP Cu/Cu5Zr electrodes that are prepared by chemical dealloying for 3, 5, and 7 h, respectively. Figure S18: typical CV curves of bare NP Cu (a) and NP Cu/Cu5Zr with various Cu/Zr/Al components (b–d) that are fabricated by chemical dealloying Cu\(_{20}\)Al\(_{80}\) and Cu\(_{20}\)Zr\(_3\)Al\(_{80}\) (\(x = 1, 3, 5\) at%), respectively. Scan rate: 20 mV s\(^{-1}\). Figure S19: (a, b) typical SEM image (a) and XRD patterns (b) of NP Cu/Cu5Zr electrode after potential cycling measurement (10,000 cycles). (c, d) Zr 3d (c) and Cu 2p (d) XPS spectra of NP Cu/Cu5Zr after 10,000 cycles. Table S1: the HBE values at the possible adsorption sites on the Cu5Zr and Cu2Zr-OH illustrated in Figure S2. Table S2: chemical compositions of NP Cu/Cu5Zr dealloyed from different precursors, which are determined by ICP-MS, EDS, and XPS analysis. Table S3: comparisons of the catalytic activity of NP Cu/Cu5Zr electrocatalytic materials in alkaline medium (pH = 14) with representative noble metal-free ones reported previously. (Supplementary Materials)

References

[1] S. Chu and A. Majumdar, “Opportunities and challenges for a sustainable energy future,” *Nature*, vol. 488, no. 7411, pp. 294–303, 2012.

[2] M. S. Dresselhaus and I. L. Thomas, “Alternative energy technologies,” *Nature*, vol. 414, no. 6861, pp. 332–337, 2001.

[3] N. S. Lewis and D. G. Nocera, “Powering the planet: chemical challenges in solar energy utilization,” *Proceedings of the National Academy of Sciences of the United States of America*, vol. 103, no. 43, pp. 15729–15735, 2006.

[4] J. A. Turner, “Sustainable hydrogen production,” *Science*, vol. 305, no. 5686, pp. 972–974, 2004.

[5] S. Chu, Y. Cui, and N. Liu, “The path towards sustainable energy,” *Nature Materials*, vol. 16, no. 1, pp. 16–22, 2017.

[6] I. Roger, M. A. Shippman, and M. D. Symes, “Earth-abundant catalysts for electrochemical and photoelectrochemical water splitting,” *Nature Reviews Chemistry*, vol. 1, no. 1, article 0003, 2017.

[7] V. R. Stamenkovic, D. Strmcnik, P. P. Lopes, and N. M. Markovic, “Energy and fuels from electrochemical interfaces,” *Nature Materials*, vol. 16, no. 1, pp. 57–69, 2017.

[8] Y. Jiao, Y. Zheng, M. Jaroniec, and S. Z. Qiao, “Design of electrocatalysts for oxygen- and hydrogen-involving energy conversion reactions,” *Chemical Society Reviews*, vol. 44, no. 8, pp. 2060–2086, 2015.

[9] R. Subbaraman, D. Tripkovic, D. Strmcnik et al., “Enhancing hydrogen evolution activity in water splitting by tailoring Li+-Ni(OH)\(_2\)-Pt interfaces,” *Science*, vol. 334, no. 6060, pp. 1256–1260, 2011.

[10] D. Strmcnik, M. Uchimura, C. Wang et al., “Improving the hydrogen oxidation reaction rate by promotion of hydroxyl adsorption,” *Nature Chemistry*, vol. 5, no. 4, pp. 300–306, 2013.

[11] J. Durst, A. Siebel, C. Simon, F. Hasché, J. Herranz, and H. A. Gasteiger, “New insights into the electrochemical hydrogen oxidation and evolution reaction mechanism,” *Energy & Environmental Science*, vol. 7, no. 7, pp. 2255–2260, 2014.

[12] Y. Zheng, Y. Jiao, A. Vasileff, and S. Z. Qiao, “The hydrogen evolution reaction in alkaline solution: from theory, single crystal models, to practical electrocatalysts,” *Angewandte Chemie International Edition*, vol. 57, no. 26, pp. 7568–7579, 2018.

[13] M. Mahmood, Y. Yao, J. W. Zhang, L. Pan, X. Zhang, and J. J. Zou, “Electrocatalysts for hydrogen evolution in alkaline electrolytes: mechanisms, Challenges, and prospective Solutions,” *Advanced Science*, vol. 5, no. 2, article 1700464, 2018.

[14] I. Ledezma-Yanez, W. D. Z. Wallace, P. Sebastián-Pascual, V. Climent, J. M. Feliu, and M. T. M. Koper, “Interfacial water reorganization as a pH-dependent descriptor of the hydrogen evolution rate on platinum electrodes,” *Nature Energy*, vol. 2, no. 4, article 17031, 2017.

[15] X. Zou and Y. Zhang, “Noble metal-free hydrogen evolution catalysts for water splitting,” *Chemical Society Reviews*, vol. 44, no. 15, pp. 5148–5180, 2015.

[16] J. R. McKone, S. C. Marinescu, B. S. Brunschwig, J. R. Winkler, and H. B. Gray, “Low-temperature using gas-phase atomic hydrogen: spectroscopic evidence for the carbonyl insertion mechanism on a transition metal surface,” *Journal of the American Chemical Society*, vol. 117, no. 9, pp. 2606–2617, 1995.

[17] P. Wang, X. Zhang, J. Zhang et al., “Precise tuning in platinum-nickel/nickel sulfide interface nanowires for synergistic hydrogen evolution catalysis,” *Nature Communications*, vol. 8, no. 1, article 14580, 2017.

[18] W. J. Mitchell, J. Xie, T. A. Jachimowski, and W. H. Weinberg, “Carbon monoxide hydrogenation on the Ru(001) surface at low temperature using gas-phase atomic hydrogen: spectroscopic evidence for the carbonyl insertion mechanism on a transition metal surface,” *Journal of the American Chemical Society*, vol. 117, no. 9, pp. 2606–2617, 1995.

[19] S. Trasatti, “Work function, electronegativity, and electrochemical behaviour of metals: III. Electrolytic hydrogen evolution in acid solutions,” *Journal of Electroanalytical Chemistry and Interfacial Electrochemistry*, vol. 39, no. 1, pp. 163–184, 1972.
evolution reaction,” *Nature Catalysis*, vol. 1, no. 12, article 195, pp. 985–992, 2018.

[21] J. Mahmood, F. Li, S. M. Jung et al., “An efficient and pH-universal ruthenium-based catalyst for the hydrogen evolution reaction,” *Nature Nanotechnology*, vol. 12, no. 5, pp. 441–446, 2017.

[22] J. Mao, C. T. He, J. Pei et al., “Accelerating water dissociation kinetics by isolating cobalt atoms into ruthenium lattice,” *Nature Communications*, vol. 9, no. 1, article 4958, 2018.

[23] F. Li, G. F. Han, H. J. Noh, I. Ahmad, I. Y. Jeon, and J. B. Baek, “Hydrogen Evolution Reaction: Mechanochemically assisted synthesis of a Ru catalyst for hydrogen evolution with performance superior to Pt in both acidic and alkaline media (Adv. Mater. 44/2018),” *Advanced Materials*, vol. 30, no. 44, article 1870330, 2018.

[24] B. You and Y. Sun, “Innovative strategies for electrocatalytic water splitting,” *Accounts of Chemical Research*, vol. 51, no. 7, pp. 1571–1580, 2018.

[25] W. Sheng, M. Myint, J. G. Chen, and Y. Yan, “Correlating the hydrogen evolution reaction activity in alkaline electrolytes with the hydrogen binding energy on monometallic surfaces,” *Energy & Environmental Science*, vol. 6, no. 5, pp. 1509–1512, 2013.

[26] N. Han, K. R. Yang, Z. Lu et al., “Nitrogen-doped tungsten carbide nanoarray as an efficient bifunctional electrocatalyst for water splitting in acid,” *Nature Communications*, vol. 9, no. 1, p. 924, 2018.

[27] E. H. Ang, K. N. Dinh, X. Sun et al., “Highly efficient and stable hydrogen production in all pH range by two-dimensional structured metal-doped tungsten semicarbazides,” *Research*, vol. 2019, article 4029516, 14 pages, 2019.

[28] J. Yin, Q. Fan, Y. Li et al., “Ni-C-N nanosheets as catalyst for hydrogen evolution reaction,” *Journal of the American Chemical Society*, vol. 138, no. 44, pp. 14546–14549, 2016.

[29] T. Wang, X. Wang, Y. Liu, J. Zheng, and X. Li, “A highly efficient and stable biphasic nanocrystalline Ni-Mo-N catalyst for hydrogen evolution in both acidic and alkaline electrolytes,” *Nanoscale*, vol. 22, pp. 111–119, 2016.

[30] C. Tang, R. Zhang, W. Lu et al., “Fe-doped CoP nanoarray: a monolithic multifunctional catalyst for highly efficient hydrogen generation,” *Advanced Materials*, vol. 29, no. 2, article 1602441, 2017.

[31] J. X. Feng, S. Y. Tong, Y. X. Tong, and G. R. Li, “Pt-like hydrogen evolution electrocatalysis on PANI/CoP hybrid nanowires by weakening the shackles of hydrogen ions on the surfaces of catalysts,” *Journal of the American Chemical Society*, vol. 140, no. 15, pp. 5118–5126, 2018.

[32] M. Cabán-Acevedo, M. L. Stone, J. R. Schmidt et al., “Efficient hydrogen evolution catalysis using ternary pyrite-type cobalt phosphosulphide,” *Nature Materials*, vol. 14, no. 12, pp. 1245–1251, 2015.

[33] X. Ge, L. Chen, L. Zhang, Y. Wen, A. Hirata, and M. Chen, “Nanoporous metal enhanced catalytic activities of amorphous molybdenum sulfide for high-efficiency hydrogen production,” *Advanced Materials*, vol. 26, no. 19, pp. 3100–3104, 2014.

[34] C. Panda, P. W. Menezes, M. Zheng, S. Ortmann, and M. Driess, “In situ formation of nanostructured core-shell Cu,N-CuO to promote alkaline water electrolysis,” *ACS Energy Letters*, vol. 4, no. 3, pp. 747–754, 2019.

[35] S. Deng, Y. Zhong, Y. Zeng et al., “Directional construction of vertical nitrogen-doped 1T-2H MoS2/graphene shell/core nanoflake arrays for efficient hydrogen evolution reaction,” *Advanced Materials*, vol. 29, no. 21, article 1700748, 2017.

[36] C. Tang, H. F. Wang, and Q. Zhang, “Multiscale principles to boost reactivity in gas-involving energy electrocatalysis,” *Accounts of Chemical Research*, vol. 51, no. 4, pp. 881–889, 2018.

[37] X. Y. Lang, H. Y. Fu, C. Hou et al., “Nanoporous gold supported cobalt oxide microelectrodes as high-performance electrochemical biosensors,” *Nature Communications*, vol. 4, no. 1, article 2169, 2013.

[38] P. C. K. Vesborg, B. Seger, and I. Chorkendorff, “Recent development in hydrogen evolution reaction catalysts and their practical implementation,” *Journal of Physical Chemistry Letters*, vol. 6, no. 6, pp. 951–957, 2015.

[39] Z. Ren, Y. Guo, and P. X. Gao, “Nano-array based monolithic catalysts: concept, rational materials design and tunable catalytic performance,” *Catalysis Today*, vol. 258, pp. 441–453, 2015.

[40] Q. Lu, G. S. Hutchings, W. Yu et al., “Highly porous non-precious bimetallic electrocatalysts for efficient hydrogen evolution,” *Nature Communications*, vol. 6, no. 1, article 6567, 2015.

[41] K. J. Zeng, M. Hämäläinen, and H. L. Lukas, “A new thermodynamic description of the Cu-Zr system,” *Journal of Phase Equilibria*, vol. 15, no. 6, pp. 577–586, 1994.

[42] Y. Liu, S. Liu, C. Zhang, Y. du, J. Wang, and Y. Li, “Experimental investigation and thermodynamic description of the Cu-Zr system,” *Journal of Phase Equilibria*, vol. 38, no. 2, pp. 121–134, 2017.

[43] G. Yuan, W. Luo, Y. Ouyang, and J. Liang, “The isothermal section of the Zr-Sn-Cu ternary system at 700 °C,” *Journal of Phase Equilibria and Diffusion*, vol. 39, no. 2, pp. 196–203, 2018.

[44] S. H. Zhou and R. E. Napolitano, “Phase stability for the Cu-Zr system: first-principles, experiments and solution-based modeling,” *Acta Materialia*, vol. 58, no. 6, pp. 2186–2196, 2010.

[45] W. H. Haynes, *Handbook of Chemistry and Physics*, CRC press, 95th edition, 2015.

[46] I. McCue, E. Benn, B. Gaskey, and J. Erlebacher, “Dealloying and dealloyed materials,” *Annual Review of Materials Research*, vol. 46, no. 1, pp. 263–286, 2016.

[47] C. Xu, L. Wang, R. Wang et al., “Nanotubular mesoporous bimetallic nanostructures with enhanced electrocatalytic performance,” *Advanced Materials*, vol. 21, no. 21, pp. 2165–2169, 2009.

[48] M. Y. W. Lou and N. J. Grant, “Identification of CuSn5Zr phase in Cu-Zr alloys,” *Metallurgical Transactions A*, vol. 15, no. 7, pp. 1491–1493, 1984.

[49] R. P. Singh, A. Lawley, S. Friedman, and Y. V. Murty, “Microstructure and properties of spray cast Cu-Zr alloys,” *Materials Science Engineering A*, vol. 145, no. 2, pp. 243–255, 1991.

[50] N. Köpfe, L. Mayr, D. Schmidmair et al., “A comparative discussion of the catalytic activity and CO₂-selectivity of Cu-Zr and Pd-Zr (intermetallic) compounds in methanol Steam reforming,” *Catalysts*, vol. 7, no. 12, p. 53, 2017.

[51] X. Zheng, Y. Ji, J. Tang et al., “Theory-guided Sn/Cu alloying for efficient CO₂ electroreduction at low overpotentials,” *Nature Catalysis*, vol. 2, no. 1, pp. 55–61, 2019.
[52] C. T. Dinh, A. Jain, F. P. G. de Arquer et al., “Multi-site electro-catalysts for hydrogen evolution in neutral media by destabilization of water molecules,” *Nature Energy*, vol. 4, no. 2, pp. 107–114, 2019.

[53] R. Q. Yao, Y. T. Zhou, H. Shi et al., “Nanoporous palladium-silver surface alloys as efficient and pH-universal catalysts for the hydrogen evolution reaction,” *ACS Energy Letters*, vol. 4, no. 6, pp. 1379–1386, 2019.