The promotion effect of manganese on Cu/SAPO for selective catalytic reduction of NO$_x$ with NH$_3$

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The activity and hydrothermal stability of Cu/SAPO and xMn–2Cu/SAPO for low-temperature selective catalytic reduction of NO$_x$ with ammonia were investigated. An ion-exchanged method was employed to synthesize xMn–2Cu/SAPO, which was characterized by N$_2$ adsorption, ICP-AES, X-ray diffraction (XRD), NH$_3$-temperature programmed desorption (NH$_3$-TPD), NO oxidation, X-ray photoelectron spectrum (XPS), UV–vis, H$_2$-temperature programmed reduction (H$_2$-TPR) and diffuse reflectance infrared Fourier transform spectra (DRIFTS). 2Mn–2Cu/SAPO and 4Mn–2Cu/SAPO showed the best SCR activity, in that at 150 °C NO conversion reached 76% and N$_2$ selectivity was above 95% for the samples. NO oxidation results showed that the 2Mn–2Cu/SAPO had the best NO oxidation activity and the BET surface area decreased as manganese loading increased. XRD results showed that the metal species was well dispersed. NH$_3$-TPD showed that the acid sites have no significant influence on the SCR activity of xMn–2Cu/SAPO. H$_2$–TPR patterns showed good redox capacity for xMn–2Cu/SAPO. UV–vis and H$_2$–TPR showed that the ratio of Mn$^{4+}$ to Mn$^{3+}$ increased as manganese loading increased. XPS spectra showed a significant amount of Mn$^{5+}$ and Mn$^{4+}$ species on the surface and addition of manganese increased the ratio of Cu$^{4+}$. The promotion effect of manganese to 2Cu/SAPO comes from the generation of Mn$^{5+}$ and Mn$^{4+}$ species. Deducem the DRIFTS spectra, the Etal–Rideal mechanism was effective on 4Mn–2Cu/SAPO.

1 Introduction

NO$_x$ consists of NO, NO$_2$, N$_2$O and their related derivatives, which can lead to numerous environmental and health hazards. By forming nitric acid, NO$_x$ can lead to the generation of acid rain; by forming ammonium nitrate, it can accelerate the generation of PM2.5; by reacting with hydrocarbons under sunlight, it can cause the generation of photo chemical smog and ozone; by infiltrating into lungs, it can result in respiratory morbidity, like impaired host defense and lung inflammation. Thus, legislations on NO$_x$ emission are now becoming more and more stringent. In China, according to the emission standard of air pollutants for thermal power plants (GB 13223-2011), the NO$_x$ emission limit is 100 mg m$^{-3}$ for new natural gas-fired industrial boilers and all oil and coal-fired industrial boilers. For compression ignition and gas fueled positive ignition engines of vehicles, the NO$_x$ emission limit in Chinese National 3 standards is 57% of that in Chinese National 4 standards and 40% of that in Chinese National 3 standards.

The treatment of environmentally harmful NO$_x$ compounds emitted from mobile or stationary sources remains a challenging task, especially for industrial applications such as cement plants or iron and steel plants, because the suitable process temperature for NO$_x$ abatement is usually around 150 °C, under which commercial SCR catalysts V$_2$O$_5$–MoO$_3$ (WO$_3$)/TiO$_2$ cannot work well.

Manganese oxides were reported by Smirniotis et al. as potential catalysts for low-temperature NH$_3$-SCR reactions. They studied the promoted manganese oxides supported on TiO$_2$ systematically and found that the surface Mn$^{4+}$ species was reported to be highly active for the SCR of NO reaction with ammonia at low temperatures.

Recently, copper modified zeolites with a CHA structure attracted much attention owing to their excellent activity, N$_2$ selectivity and hydrothermal stability for low temperature NH$_3$-SCR reaction. However for the Cu/SAPO-34 catalyst, the oxygen activation of transiently formed Cu pairs [Cu$(\text{NH}_3)_2$] to [(NH$_3$)$_2$Cu$^{II}$O$_2$–Cu$^{II}$][NH$_3$] was rate-limiting in the catalytic cycle. Feng Gao also thought that the oxidation reaction of Cu$^{+}$ to Cu$^{2+}$ was the rate-determine step, so tuning the redox properties of the active site by introducing a second cation might be a promising approach. Li et al. and Chen et al. modified the Cu/SAPO catalyst with CeO$_2$ for SCR reaction; they found that the introduction of CeO$_2$ increased the SCR activity at high temperatures and cerium helped increase the amount of isolated Cu$^{2+}$ ions.

Leistner et al. found that Cu in Cu/SAPO-34 was more easily reduced compared to Cu/SSZ-13, which could facilitate the
redox processes and increase the SCR activity. Ma et al. also found that Cu-SAPO-34 showed higher DeNO\textsubscript{x} catalytic activity than Cu-SSZ-13. The usage of SAPO-34 in low-temperature SCR might be favored in low temperature SCR reactions. When manganese is introduced into the Cu/SAPO system, it not only promotes the oxidation reaction of Cu\textsuperscript{+}, but also acts as an active site for low temperature SCR reaction. Thus the manganese modified Cu/SAPO-34 catalyst would probably show better SCR activity.

In this work, the treatment of exhaust from a gas-fired boiler which contains a high concentration of H\textsubscript{2}O and a very low concentration of SO\textsubscript{2} was focused upon. xMn–Cu/SAPO catalysts with different manganese loadings were prepared and their performance in SCR reaction was discussed. Various characterization methods (XRD, XPS, H\textsubscript{2}-TPR, UV-vis) were used to probe the promotion effect of manganese on Cu/SAPO for low temperature SCR reaction.

2 Experimental

2.1 Catalyst preparation

xMn–2Cu/SAPO catalysts were prepared by an ion-exchange method. The alkali metals in commercial Na/SAPO-34 powder (Jiangsu XFNANO) would decrease SCR activities, so it was transformed to H\textsuperscript{+}/SAPO as follows. Na/SAPO-34 powder was ion exchanged using 11% wt NH\textsubscript{4}NO\textsubscript{3} (Aladdin, >98.5%) solution whose pH value was adjusted to 3.0–4.0 by 2 M ammonium hydroxide (Aladdin, 25–28%) solution at 80 °C for 4 h. It was then filtered and washed with distilled water three times. Finally, it was dried at 110 °C for 16 h and calcined at 550 °C for 3 h.

xMn–2Cu/SAPO (x = 0, 1, 2, 4, 8) was prepared in two steps. Firstly, H\textsuperscript{+}/SAPO was mixed with Cu(NO\textsubscript{3})\textsubscript{2} (across, >95%) solution at 80 °C for 6 h under vigorous stirring. Then it was dried and calcined at 550 °C for 3 h. In the second step, it was mixed with Mn(NO\textsubscript{3})\textsubscript{2} (across, >95%) solution at 80 °C for 6 h under vigorous stirring, and at last it was dried and calcined at 550 °C for 3 h.

2.2 SCR activity measurements

Catalytic activity evaluation was carried out using a flow-through powder reactor system equipped with a Fourier transform infrared (FT-IR) spectrometer (THERMO SCIENTIFIC IGS). In order to prevent condensation along upstream tubing, all the gas lines were heated and maintained at 120 °C. The gas mixture was 500 ppm NO + 500 ppm NH\textsubscript{3} + 3% vol O\textsubscript{2} balanced with N\textsubscript{2}. The gas hourly space velocity (GHSV) was 65 000 h\textsuperscript{−1} containing 500 ppm NO and 3% vol O\textsubscript{2} balanced with N\textsubscript{2}. The typical time to achieve the steady state was 2.5 h. NO conversion, promotion effect, and (E\textsubscript{p}) and N\textsubscript{2} selectivity were calculated at steady state using the equation below:

\[
X_{\text{NO}} = \frac{C_{\text{NO}_{\text{in}}}}{C_{\text{NO}_{\text{out}}}} \times 100\%
\]

where \(C_{\text{NO}_{\text{in}}}\) and \(C_{\text{NO}_{\text{out}}}\) represent the inlet and outlet NO\textsubscript{x} concentration. \(X_{\text{NO}_{\text{in}}}-\text{2Cu/SAPO}, X_{\text{Cu/SAPO}}\) and \(X_{\text{Mn/SAPO}}\) respectively represent the NO\textsubscript{x} conversion of xMn–2Cu/SAPO, 2Cu/SAPO and xMn/SAPO at 210 °C, and \([N_2O]_{\text{outlet}}\) represents the outlet N\textsubscript{2}O concentration.

2.3 Characterization

X-ray diffraction patterns were collected on a Bruker D8 Advance X-ray diffractometer with a Ni-Filtered Cu K\alpha with a step size of 0.02 in the 2θ range of 5° to 40°.

Ammonia temperature programmed desorption (NH\textsubscript{3}-TPD) experiments were carried out as follows. 150 mg sample was first pretreated in 21% vol O\textsubscript{2}/N\textsubscript{2} at 500 °C for 30 min and then cooled to 100 °C in N\textsubscript{2} and a total flow of 100 ml min\textsuperscript{−1} containing 2500 ppm NH\textsubscript{3} in N\textsubscript{2} was injected into the reactor for 2 hours to achieve a steady state. Once the catalyst was saturated, NH\textsubscript{3} was switched off and the catalyst was swept by N\textsubscript{2} overnight. Finally, the catalyst was heated in N\textsubscript{2} at a temperature ramp to 700 °C with a heating rate of 10 °C min\textsuperscript{−1}.

NO oxidation experiments were carried out as follows. 150 mg sample was first pretreated in 21% vol O\textsubscript{2}/N\textsubscript{2} at 500 °C for 30 min and then cooled to 210 °C in N\textsubscript{2}. Then a total flow of 100 ml min\textsuperscript{−1} containing 500 ppm NO and 3% vol O\textsubscript{2} was injected into the reactor for 30 min to achieve the steady state. The contents of the elements were determined by ion coupled plasma (ICP) optical emission spectroscopy (Thermo IRIS Intrepid II) after microwave digestion.

The BET surface area, pore volume and pore size of the catalyst samples were measured by N\textsubscript{2} adsorption using the MICROMERITICS ASAP 2020 surface area and porosity analyzer.

Diffuse reflectance UV-vis spectra were recorded in the range of 200–800 nm against a BaSO\textsubscript{4} as a reference standard on a HitachiU-3900 UV-vis spectrophotometer equipped with an integration sphere.

H\textsubscript{2}-TPR experiments were performed using 20 mg sample as follows. The catalysts were first pretreated at 500 °C for 30 min in a highly pure O\textsubscript{2} (40 ml min\textsuperscript{−1}) stream. Then the furnace temperature was decreased to room temperature, and feed containing 5% vol H\textsubscript{2} in N\textsubscript{2} was fed at a flow rate of 40 ml min\textsuperscript{−1}.

X-ray photoelectron spectroscopy (XPS) analyses were performed using PHI quantera SXMM Scanning ESCA Microprobe (Physical Electronics) with a hemispherical detector operating at a constant pass energy (PE = 55 eV). An X-ray source at 210 W (I = 15 mA, U = 14 kV) and Al K\alpha radiation (1486.6 eV) were used. Intensities were estimated from the integration of each peak, after smoothening, subtracting the L-shaped background,
and fitting the experimental curve to a combination of Lorentzian and Gaussian lines of variable proportions. All binding energies were referenced to the C 1s line at 284.8 eV.

Diffuse reflectance infrared Fourier transform spectra (DRIFTS) were measured on an FT-IR spectrometer (Thermo Nicolet NEXUS870) with an MCT detector and high temperature reaction chamber (Harrick Scientific Praying Mantis) with ZnSe Windows, which was connected to a gas-dosing system. The powder sample was placed in a sample cup and heated by a cartridge heater underneath the sample. The temperature was adjusted using a K-type thermocouple connected to a Harrick temperature controller. Before each measurement, oxidation pretreatments were executed at 500 °C for 1 h. Background spectra were collected before adsorption for 32 scans with a resolution of 4 cm⁻¹ in N₂. DRIFTS spectra were recorded in the range of 4000–650 cm⁻¹ for 32 scans with a resolution of 4 cm⁻¹.

3 Results and discussion

3.1 Low temperature NH₃-SCR activity and N₂ selectivity

Fig. 1a shows the NO conversion of xMn–2Cu/SAPO. There was not much difference in the catalytic activity between 2Cu/SAPO and 1Mn–2Cu/SAPO, 2Mn–2Cu/SAPO and 4Mn–2Cu/SAPO. The activity of 8Mn–2Cu/SAPO was slightly lower than the activity of 2Mn–2Cu/SAPO. As manganese loading increased from 1% to 2%, the activity increased obviously and the SCR activity of 2Mn–2Cu/SAPO, 4Mn–2Cu/SAPO and 8Mn–2Cu/SAPO was much higher than that of 2Cu/SAPO and 1Mn–2Cu/SAPO.

For 2Mn–2Cu/SAPO and 4Mn–2Cu/SAPO, the conversion of NO at 120 °C was about 60%; as the temperature increased to 210 °C, the conversion reached 95%.

Fig. 1b shows the N₂ selectivity of xMn–2Cu/SAPO. Good N₂ selectivity was observed, being above 95% for all catalysts. Other researchers also found good N₂ selectivity for the supported SAPO-34 catalysts.

3.2 Water resistance test

Fig. 2 shows the water resistance test of xMn–2Cu/SAPO. The poison effect of water might come from the reversible adsorption of H₂O or irreversible dealumination process of zeolite.

In Fig. 2, all catalysts showed strong resistance to water at high temperature. It could be caused by the lower adsorption capability of water at high temperature, since SAPO-34 was highly hydrothermally stable and the poison effect might originate from the competitive adsorption of water with ammonia.

**Fig. 1** (a) Catalytic activity and (b) N₂ selectivity of xMn–2Cu/SAPO. Conditions: 500 ppm NO + 500 ppm NH₃ + 3% vol O₂ + N₂ balance, GHSV = 65 000 h⁻¹.

**Fig. 2** Water resistance test of xMn–2Cu/SAPO for low-temperature SCR. Conditions: 500 ppm NO + 500 ppm NH₃ + 16.3% H₂O + 3% vol O₂ + N₂ balance, GHSV = 65 000 h⁻¹.

Compared with xMn/SAPO (x = 1, 2, 4, 8), the previously prepared xMn–2Cu/SAPO showed better SCR activities, especially for 2Mn–2Cu/SAPO, indicating that the interactions between copper and manganese might increase its SCR activity.
3.3 NO oxidation activity and promotion effect

Fig. 3 showed the NO oxidation activity of xMn–2Cu/SAPO and the promotion effect. 2Mn–2Cu/SAPO had the best NO oxidation activity, indicating the interaction between copper and manganese elements, which accelerated the NO oxidation process. Good oxidation activity would certainly help the oxidation of Cu\(^+\) to Cu\(^{2+}\) in SCR reaction. The \(E_p\) profile was almost the same as the NO conversion profile, indicating that the increase in NO oxidation activity would promote the low-temperature SCR activity. Compared with the xMn/SAPO \((x = 1, 2, 4, 8)\) previously prepared, the NO oxidation activities of xMn–2Cu/SAPO were much higher, which could be one reason for the increased SCR activity.

3.4 Physicochemical properties

As showed in Table 1, the BET surface area and micro pore volume decreased gradually as manganese loading increased; the surface areas of 2Mn–2Cu/SAPO, 4Mn–2Cu/SAPO and 8Mn–2Cu/SAPO were similar. xMn/SAPO \((x = 1, 2, 4, 8)\) previously prepared showed the same trend: the surface area and pore volume decreased as manganese loading increased.

The manganese loading, copper loading and Mn/Cu molar ratio from ICP result were almost the same as the set values. Meanwhile the Mn/Cu molar ratio from the XPS results was quite different from that of ICP results. Since XPS and ICP reflect the elemental information of the surface phase and the bulk phase respectively, the difference might come from the enrichment of manganese on surface.

3.5 XRD

XRD patterns of xMn–2Cu/SAPO are depicted in Fig. 4. The diffraction peaks of chabazite phase with the space group of \(R3\text{m}\) were identified, indicating that the crystalline structure of SAPO-34 remained unchanged after catalyst preparation. When manganese loading increased from 0 to 4%, only the chabazite phase was detectable, indicating that the manganese species were well dispersed. As manganese loading reached 8%, the diffraction peaks of the Mn\(_2\)O\(_3\) phase appeared, indicating the aggregation of manganese species and decreased manganese dispersity. On xMn/SAPO \((x = 1, 2, 4, 8)\), which was previously prepared, there too were no peaks attributed to manganese oxides until manganese loading increased to 8%.

3.6 NH\(_3\)-TPD

Fig. 5 showed the effluent NH\(_3\) profiles during the NH\(_3\)-TPD process. The amount of NH\(_3\) desorption was closely related to the amount of acid sites. As expected, when manganese loading increased, the amount of NH\(_3\) desorption decreased, which was caused by the replacement of H\(^+\) in the framework of SAPO by Mn\(^{2+}\) in the ion exchange process.

In Fig. 5 there are three NH\(_3\) desorption peaks, \(i.e.\) a low temperature peak at 250 \(^\circ\)C, a middle temperature peak at 400 \(^\circ\)C and a high temperature peak at 500 \(^\circ\)C, which could be assigned to the weak acid site, the middle strong acid site and the strong acid site, respectively.\(^{17}\)

The shapes of 1Mn–2Cu/SAPO, 2Mn–2Cu/SAPO and 4Mn–2Cu/SAPO were similar while the activity differed a lot, indicating that the related acid sites might not have significant influence on the SCR activity for xMn–2Cu/SAPO. For xMn/SAPO \((x = 1, 2, 4, 8)\), the amount of ammonia desorption also decreased as manganese loading increased.

3.7 H\(_2\)-TPR

The H\(_2\)-TPR profiles of the four catalysts are depicted in Fig. 6. For 8Mn–2Cu/SAPO, there were two sharp peaks at 278 \(^\circ\)C and 327 \(^\circ\)C, respectively. The former indicated the reduction of MnO\(_2\) to Mn\(_2\)O\(_3\), while the latter indicated the reduction of Mn\(_2\)O\(_3\) to Mn\(_3\)O\(_4\)\(^{18}\). Compared with Mn/TiO\(_2\),\(^{18,19}\) the reduction peaks moved to a lower temperature, indicating its good redox capacity.

Table 1  Physico-chemical properties

| Sample     | BET surface area \((\text{m}^2 \text{ g}^{-1})\) | Micro pore volume \((\text{cm}^3 \text{ g}^{-1})\) | Micro pore width (nm) | Mn loading\(^a\) (wt%) | Cu loading\(^a\) (wt%) | Mn/Cu molar ratio\(^a\) | Mn/Cu molar ratio\(^b\) |
|------------|-----------------------------------------------|-----------------------------------------------|-----------------------|------------------------|------------------------|-------------------------|-------------------------|
| 2Cu/SAPO   | 467                                           | 0.231                                         | 2.36                  | 0.051                  | 1.951                  | 1.051                   | 1.848                   |
| 1Mn–2Cu/SAPO | 459                                         | 0.230                                         | 1.90                  | 1.054                  | 2.016                  | 0.51                    | 1.587                   |
| 2Mn–2Cu/SAPO | 430                                         | 0.228                                         | 2.53                  | 2.108                  | 1.998                  | 1.053                   | 2.594                   |
| 4Mn–2Cu/SAPO | 426                                         | 0.206                                         | 1.96                  | 3.846                  | 2.048                  | 1.878                   | 3.424                   |
| 8Mn–2Cu/SAPO | 424                                         | 0.196                                         | 2.54                  | 8.306                  | 1.848                  | 4.495                   | 9.104                   |

\(^a\) Determined by ICP. \(^b\) Determined by XPS.
In the TPR patterns, the reduction process of copper in SAPO-34 zeolite could be separated into two steps, i.e. the reduction of Cu$^{2+}$ to Cu$^{+}$ at about 300 °C and the reduction of Cu$^{+}$ to Cu$^{0}$ at about 440 °C, while in Fig. 6 the peak strength of 1Mn–2Cu/SAPO was relatively low, indicating that the peak strength of 2Cu/SAPO was even lower, making the assignment of copper species difficult and unreasonable, so the assignment of copper was not discussed. The hydrogen consumption amount of copper reduction was relatively low, so the ignorance of copper species would not influence the discussion of manganese species.

As manganese loading increased, the peaks at 278 °C and 327 °C increased, indicating that the proportion of Mn$^{3+}$ and Mn$^{4+}$ species increased. For 1Mn–2Cu/SAPO catalysts, only one broad peak between 250 and 400 °C was detectable, indicating that most of the manganese species were Mn$^{2+}$ since Mn$^{2+}$ cannot be reduced below 700 °C. Since Mn$^{2+}$ species were inactive for SCR reaction, addition of manganese did not help increase the activity for 1Mn–2Cu/SAPO, which was consistent with $E_p$ trends and NO oxidation results. For 2Mn–2Cu/SAPO the peaks belonging to Mn$^{3+}$ and Mn$^{4+}$ species started to generate, indicating its good redox capacity. For 4Mn–2Cu/SAPO and 8Mn–2Cu/SAPO, the peaks at 278 °C and 327 °C increased greatly, indicating that the ratio of Mn$^{3+}$ and Mn$^{4+}$ increased. When manganese loading was above 2%, Mn$^{3+}$ and Mn$^{4+}$ species began to generate, which is consistent with the $E_p$ trends and NO oxidation results.

Compared with $x$Mn/SAPO ($x = 1, 2, 4, 8$) previously prepared, for which the temperature peak of H$_2$-TPR was about 353 °C and 442 °C, the peak temperature moved to a lower value, indicating that the interactions between copper and manganese species might accelerate oxidation reaction and increase the catalytic oxidation activity, which would promote low temperature SCR reaction.

Table 2 represents the hydrogen consumption of $x$Mn–2Cu/SAPO during TPR processes. As manganese loading increased, hydrogen consumption increased gradually, while the actual hydrogen consumption was much lower than the theoretical value. The ratio of experimental hydrogen consumption to...
theoretical hydrogen consumption of 8Mn–2Cu/SAPO and 4Mn–2Cu/SAPO was similar, as well as that of 1Mn–2Cu/SAPO and 2Mn–2Cu/SAPO. Since the ratio of experimental hydrogen consumption to theoretical hydrogen consumption was highly related to the proportion of Mn$^{3+}$ and Mn$^{4+}$, the proportion of Mn$^{3+}$ and Mn$^{4+}$ for 1Mn–2Cu/SAPO and 2Mn–2Cu/SAPO were similar, as well as that of 8Mn–2Cu/SAPO and 4Mn–2Cu/SAPO.

### Table 2: Hydrogen consumption

| Hydrogen consumption$^a$ (10$^{-2}$ mmol) | 8Mn–2Cu/SAPO | 4Mn–2Cu/SAPO | 2Mn–2Cu/SAPO | 1Mn–2Cu/SAPO |
|------------------------------------------|--------------|--------------|--------------|--------------|
| Hydrogen consumption$^b$ (10$^{-2}$ mmol) | 2.04         | 1.11         | 0.468        | 0.327        |
| Hydrogen consumption$^a$/hydrogen consumption$^b$ (%) | 57.8         | 53.4         | 35.7         | 33.1         |

$^a$ The actual hydrogen consumption calculated from TPR profiles. $^b$ The theoretical hydrogen consumption, assuming all of the Cu species are Cu$^{2+}$ and reduced to Cu$^0$, and all of the Mn species are Mn$^{4+}$ and reduced to Mn$^{2+}$ in TPR process.

### 3.8 DR UV-vis

The UV-vis spectra of the samples are given in Fig. 7. In Fig. 7, the absorption bands at 240 nm were related to the charge transfer processes between the framework aluminum and oxygen atoms of aluminophosphate. The band at 220 nm was attributed to the oxygen-to-metal charge-transfer of the isolated Cu$^+/Cu^{2+}$ bound on the framework of zeolites. The bands at 280–300 nm were assigned to copper oxide clusters. The bands at 700–800 nm were assigned to the d–d transition of isolated Cu$^{2+}$ in weak disordered octahedral coordination of O-containing ligands.

The broad bands in the range 320–380 nm were attributed to the Mn$^{3+}$ → O$^{2-}$ charge transfer transition superimposed on the $^3$B$_{1g}$ → $^3$B$_{2g}$ crystal field d–d transition. The band at 322 nm was tentatively assigned to the Mn$^{3+}$ → O$^{2-}$ charge transfer in Mn$_3$O$_4$ in which manganese was octahedrally coordinated with oxygen. The band at 255–276 nm could be assigned to the Mn$^{2+}$ → O$^{2-}$ charge transfer transition in tetrahedral oxygen coordination. In the α-Mn$_3$O$_4$ structure, Mn$^{3+}$ ions occupied octahedral sites and, if highly symmetric, a single spin-allowed absorption band in the d–d transition region was expected similarly to [Mn(H$_2$O)$_6$]$^{3+}$ at 500 nm.

As manganese loading increased, the band at 320–380 nm increased gradually especially when the manganese loading changed from 2% to 4%, indicating that the percentage of Mn$^{3+}$ increased.

Meanwhile the bands at 255–276 nm were generated, indicating that the percentage of Mn$^{2+}$ decreased. Consistent with

![Fig. 7](image_url) DR UV-vis spectra recorded under ambient atmosphere of xMn–2Cu/SAPO.

![Fig. 8](image_url) XPS spectra of (a) Mn 2P and (b) Cu 2P.
the H₂-TPR result, the UV-vis results showed the partial oxidation of Mn²⁺ to Mn³⁺ or Mn⁴⁺, when manganese loading increased. The UV-vis spectra of xMn–2Cu/SAPO were similar to that of xMn/SAPO (x = 1, 2, 4, 8) in that the percentage of Mn³⁺ increased suddenly as manganese loading increased from 2% to 4%, indicating that the distribution of manganese species might not have been influenced by the copper species.

3.9 XPS

Table 3 Binding energies (eV) of Mn²⁺ 2P₃/₂ and the percentage of Mn³⁺

| Compound          | Mn²⁺ (eV) | %  | Mn³⁺ (eV) | %  | Mn⁴⁺ (eV) | %  |
|-------------------|-----------|----|-----------|----|-----------|----|
| 1Mn–2Cu/SAPO      | 640.97    | 26.3 | 642.19    | 38.8 | 643.51    | 34.9 |
| 2Mn–2Cu/SAPO      | 640.69    | 0.5  | 641.83    | 55.8 | 643.36    | 44.2 |
| 4Mn–2Cu/SAPO      | 639.99    | 0.1  | 641.58    | 56.2 | 643.18    | 43.8 |
| 8Mn–2Cu/SAPO      | 640.42    | 9.0  | 641.83    | 60.7 | 643.44    | 30.3 |

Table 4 Binding energies (eV) of Cu¹⁺ 2P₃/₂ and the percentage of Cu²⁺

| Compound          | Cu¹⁺ (eV) | %  | Cu²⁺ (eV) | %  |
|-------------------|-----------|----|-----------|----|
| 2Cu/SAPO          | 932.75    | 78.3 | 934.93    | 21.7 |
| 1Mn–2Cu/SAPO      | 932.75    | 79.9 | 934.81    | 20.1 |
| 2Mn–2Cu/SAPO      | 932.64    | 75.6 | 934.56    | 24.4 |
| 4Mn–2Cu/SAPO      | 932.84    | 65.7 | 934.93    | 34.3 |
| 8Mn–2Cu/SAPO      | 932.82    | 67.2 | 934.80    | 32.7 |

Table 5 Binding energies (eV) of Mn²⁺ 2P₃/₂ and the percentage of Mn³⁺

| Compound          | Mn²⁺ (eV) | %  | Mn³⁺ (eV) | %  | Mn⁴⁺ (eV) | %  |
|-------------------|-----------|----|-----------|----|-----------|----|
| 1Mn–2Cu/SAPO      | 640.97    | 26.3 | 642.19    | 38.8 | 643.51    | 34.9 |
| 2Mn–2Cu/SAPO      | 640.69    | 0.5  | 641.83    | 55.8 | 643.36    | 44.2 |
| 4Mn–2Cu/SAPO      | 639.99    | 0.1  | 641.58    | 56.2 | 643.18    | 43.8 |
| 8Mn–2Cu/SAPO      | 640.42    | 9.0  | 641.83    | 60.7 | 643.44    | 30.3 |

the H₂-TPR result, the UV-vis results showed the partial oxidation of Mn²⁺ to Mn³⁺ or Mn⁴⁺, when manganese loading increased.

The UV-vis spectra of xMn–2Cu/SAPO were similar to that of xMn/SAPO (x = 1, 2, 4, 8) in that the percentage of Mn³⁺ increased suddenly as manganese loading increased from 2% to 4%, indicating that the distribution of manganese species might not have been influenced by the copper species.

By performing peak-fitting deconvolution, the Mn 2P₃/₂ peak could be separated into three peaks, i.e. 640.8–640.9 eV, 642.1–642.3 eV and 643.8–644.1 eV, which correspond to the Mn²⁺ species, Mn³⁺ species and Mn⁴⁺ species, respectively. The atomic percentages of manganese species in different valence states are shown in Table 3.

When manganese loading increased from 1% to 2%, the percentage of Mn³⁺ and Mn⁴⁺ increased and the percentage of Mn²⁺ decreased a lot, while when manganese loading kept increasing, the percentage of manganese species did not change much. This was consistent with the NO oxidation, H₂-TPR and UV-vis results.

As manganese loading increased, the percentage of Mn³⁺ species increased, and the percentage of Mn⁴⁺ species also increased, except for 8Mn–2Cu/SAPO. This was consistent with the UV-vis and H₂-TPR results. Though H₂-TPR, UV-vis and XPS all showed that there were many Mn³⁺ and Mn⁴⁺ species on 8Mn–2Cu/SAPO, the XRD patterns showed the Mn₂O₃ phase, which indicated the aggregation of manganese species and decreased manganese dispersity, and the NH₃-TPD results showed that there were much less acid sites. The decreased manganese dispersity and the less acid sites could lead to the lower SCR activity.

The XPS spectra of xMn/SAPO (x = 1, 2, 4, 8) also showed that most of the manganese species were Mn³⁺ and Mn⁴⁺.

Fig. 8a shows the XPS spectra of Mn 2P of xMn–2Cu/SAPO. Two main peaks assigned respectively to Mn 2P₃/₂ at 642.5 eV and Mn 2P₃/₂ at 654 eV were observed. By performing peak-fitting deconvolution, the Mn 2P₃/₂ peak could be separated into two peaks, i.e. 640.8–640.9 eV, 642.1–642.3 eV and 643.8–644.1 eV, which correspond to the Mn²⁺ species, Mn³⁺ species and Mn⁴⁺ species, respectively. The atomic percentages of manganese species in different valence states are shown in Table 3.

When manganese loading increased from 1% to 2%, the percentage of Mn³⁺ and Mn⁴⁺ increased and the percentage of Mn²⁺ decreased a lot, while when manganese loading kept increasing, the percentage of manganese species did not change much. This was consistent with the NO oxidation, H₂-TPR and UV-vis results.

As manganese loading increased, the percentage of Mn³⁺ species increased, and the percentage of Mn⁴⁺ species also increased, except for 8Mn–2Cu/SAPO. This was consistent with the UV-vis and H₂-TPR results. Though H₂-TPR, UV-vis and XPS all showed that there were many Mn³⁺ and Mn⁴⁺ species on 8Mn–2Cu/SAPO, the XRD patterns showed the Mn₂O₃ phase, which indicated the aggregation of manganese species and decreased manganese dispersity, and the NH₃-TPD results showed that there were much less acid sites. The decreased manganese dispersity and the less acid sites could lead to the lower SCR activity.

The XPS spectra of xMn/SAPO (x = 1, 2, 4, 8) also showed that most of the manganese species were Mn³⁺ and Mn⁴⁺.

3.10 DRIFTS

In situ DRIFTS was performed to identify the adsorbed NO species. Fig. 10 shows the DRIFTS spectra of 4Mn–2Cu/SAPO at 150 °C exposed to different gas phases. In Fig. 10a, there was only one positive band at 1360 cm⁻¹ in the range of 1200-
1600 nm, which belongs to nitrite groups,\textsuperscript{32,33} suggesting that those NO species were nitrite groups. The negative bands at 3280 cm\textsuperscript{-1} and 2730 cm\textsuperscript{-1} might be attributed to the depletion of Bronsted base sites by the nitrite groups. Three bands at 3675 cm\textsuperscript{-1}, 3623 cm\textsuperscript{-1} and 3598 cm\textsuperscript{-1} could be attributed to the appearance of O–H bond of nitrite groups (–N–O–H) in different chemical environments. When NH\textsubscript{3} was introduced into the system, as shown in Fig. 10b, the bands associated with nitrite groups decreased rapidly, indicating reaction between NH\textsubscript{3} and NO related species.

In Fig. 10c, when NH\textsubscript{3} was introduced into the system, the bands at 1463 cm\textsuperscript{-1} increased greatly, indicating the generation of...
of NH₃ species on the Brønsted acid sites,²⁸ while the bands at 1210 cm⁻¹ increased slightly, which was consistent with the fact that most of the acid sites on SAPO-34 were Brønsted acid sites. The negative bands at 3681 cm⁻¹, 3625 cm⁻¹ and 3600 cm⁻¹ were attributed to the depletion of Brønsted acid sites, which was caused by the adsorption of NH₃. Specifically, the negative bands at 3681 cm⁻¹ were assigned to the occupation of P–OH sites by NH₃. The other two bands were related to the depletion of Si–OH–Al acid sites. The band at 3147–3396 cm⁻¹ was attributed to the N–H stretching vibrations of NH₄⁺.²⁹ When NO was introduced into the system, as shown in Fig. 10d, the bands related to the NH₃ species decreased gradually, which was caused by the reaction between NO and NH₃.

Fig. 10e and c are similar in that only bands related to NH₃ are visible, while no other bands appeared, especially the negative bands at 2700 cm⁻¹ and 3000–3400 cm⁻¹ were invisible.

This was strong evidence that only the Elay–Rideal mechanism was effective on the 4Mn–2Cu/SAPO at 150 °C. Since if Langmuir–Hinshelwood mechanism was effective, the adsorbed and well activated NO and NH₃ related species would react with each other, while the NO species not well activated would accumulate and be detectable for DRIFTS, leading to the appearance of negative bands at 2700 and 3000–3400 nm⁻¹, as Fig. 10a and b show. Due to the fact that those peaks were not observed in Fig. 10e, all these observations lead to the exclusion of the Langmuir–Hinshelwood mechanism and the assurance of the Elay–Rideal mechanism. In the Elay–Rideal mechanism, NO molecules from gas phase directly react with the well activated NH₃ species adsorbed on the surface and leave as products to gas phase with the active site left.

4 Conclusion

In this work, the SCR activity and physicochemical properties of xMn–2Cu/SAPO were investigated. The H₂–TPR, UV-vis and XPS results showed that Mn⁴⁺ and Mn³⁺ species began to appear when manganese loading reached 2%. The NO oxidation results showed that 2Mn–2Cu/SAPO had the best NO oxidation activity, indicating that good interaction existed between copper and manganese, which was consistent with the XPS results of Cu 2p binding energy.

The graph of the promotion effect of manganese showed that for 2Mn–2Cu/SAPO, manganese had a good promotion effect on 2Cu/SAPO. According to H₂–TPR, UV-vis, XPS and NO oxidation results, the promotion effect of manganese on 2Cu/SAPO comes from the generation of Mn³⁺ and Mn⁴⁺ species. The addition of manganese increased the ratio of Cu²⁺. The Mn³⁺ and Mn⁴⁺ species might take part in the catalytic cycle and accelerate the oxidation of Cu⁺ to Cu²⁺.

From the in situ DRIFTS results, the acid sites on the surface were mainly Brønsted acid sites, the main mechanism on 4Mn–2Cu/SAPO-34 at low temperature was the Elay–Rideal mechanism.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This research is sponsored by the China National Key Research and Development Program via Project No. 2016YFB0600603.

References

1. T. Bopingari and P. G. Smirniotis, Impact of nitrogen oxides on the environment and human health: Mn-based materials for the NOx abatement, Curr. Opin. Chem. Eng., 2016, 13, 133–141.
2. P. G. Smirniotis, D. A. Penä and B. S. Uphade, Low-Temperature Selective Catalytic Reduction (SCR) of NO with NH₃ by Using Mn, Cr, and Cu Oxides Supported on Hombikat TiO₂, Angew. Chem., Int. Ed., 2001, 40, 2479–2482.
3. P. R. Ettireddy, N. Ettireddy, T. Bopingari, R. Pardemann and P. G. Smirniotis, Investigation of the selective catalytic reduction of nitric oxide with ammonia over Mn/TiO₂ catalysts through transient isotopic labeling and in situ FT-IR studies, J. Catal., 2012, 292, 53–63.
4. B. Thirupathi and P. G. Smirniotis, Nickel-doped Mn/TiO₂ as an efficient catalyst for the low-temperature SCR of NO with NH₃: Catalytic evaluation and characterizations, J. Catal., 2012, 288, 74–83.
5. D. K. Pappas, T. Bopingari, P. Boolchand and P. G. Smirniotis, Novel manganese oxide confined interwoven titania nanotubes for the low-temperature Selective Catalytic Reduction (SCR) of NOₓ by NH₃, J. Catal., 2016, 334, 1–13.
6. P. R. Ettireddy and A. Kotra, Low Temperature SCR Catalysts Optimized for Cold-Start and Low-Load Engine Exhaust Conditions, SAE Technical Paper 2015-01-1026, 2015.
7. C. Wang, J. Wang, J. Wang, T. Yu, M. Shen, W. Wang and W. Li, The effect of sulfate species on the activity of NH₃-SCR over Cu/SAPO-34, Appl. Catal., B, 2017, 204, 239.
8. C. Paolucci, I. Khurana, A. A. Parekh, R. Gounder, et al., Dynamic multinuclear sites formed by mobilized copper ions in NOₓ selective catalytic reduction, Appl. Catal., B, 2018, 220, 161–170.
9. F. Gao, J. Kwak, J. Szanyi and C. F. Peden, Current Understanding of Cu-Exchanged Chabazite Molecular Sieves for Use as Commercial Diesel Engine DeNOₓ Catalysts, Top. Catal., 2013, 56, 1441–1459.
10. X. Dong, J. Wang, H. Zhao and Y. Li, The promotion effect of CeO₂ on Cu/SAPO-34 catalyst for selective catalytic reduction of NOₓ with ammonia, Catal. Today, 2015, 258, 28–34.
11. Y. Cao, S. Zou, L. Lan, Y. Yang, H. Xu, T. Lin, M. Gong and Y. Chen, Promotional effect of Ce on Cu/SAPO-34 monolith catalyst for selective catalytic reduction of NOₓ with ammonia, J. Mol. Catal. A: Chem., 2015, 398, 304–311.
12. K. Leistner, M. Mihai, K. Wijayanti, A. Kumar, K. Kamasanamudram, N. W. Currier, A. Yezersits and L. Olsson, Comparison of Cu/BEA, Cu/SSZ-13 and Cu/SAPO-34 for ammonia-SCR reactions, Catal. Today, 2015, 258, 49–55.
13. L. Ma, Y. Cheng, G. Cavataio, R. W. McCabe, L. Fu and J. Li, In situ DRIFTS and temperature-programmed technology

6118 | RSC Adv., 2018, 8, 6110–6119
study on NH$_3$-SCR of NO$_x$ over Cu-SSZ-13 and Cu-SAPO-34 catalysts, Appl. Catal., B, 2014, 156–157, 428–437.

14 C. Pang, Y. Zhuo and Q. Weng, Mn/SAPO-34 as an efficient catalyst for the low-temperature selective catalytic reduction of NO$_x$ with NH$_3$, RSC Adv., 2017, 7, 32146–32154.

15 G. Qi and R. T. Yang, Low-temperature selective catalytic reduction of NO with NH$_3$ over iron and manganese oxides supported on titania, Appl. Catal., B, 2003, 44, 217–225.

16 D. Wang, Y. Jangjou, Y. Liu, M. K. Sharma, J. Luo, J. Li, K. Kamasamudram and W. S. Epling, A comparison of hydrothermal aging effects on NH$_3$-SCR of NO$_x$ over Cu-SSZ-13 and Cu-SAPO-34 catalysts, Appl. Catal., B, 2015, 165, 438–445.

17 D. Wang, L. Zhang, J. Li, K. Kamasamudram and W. S. Epling, NH$_3$-SCR over Cu/SAPO-34 – Zeolite acidity and Cu structure changes as a function of Cu loading, Catal. Today, 2014, 231, 64–74.

18 D. K. Pappas, T. Boningari, P. Boolchand and P. G. Smirniotis, Novel manganese oxide confined interweaved titania nanotubes for the low-temperature Selective Catalytic Reduction (SCR) of NO$_x$ by NH$_3$, J. Catal., 2016, 334, 1–13.

19 D. A. Peña, B. S. Uphade and P. G. Smirniotis, TiO$_2$-supported metal oxide catalysts for low-temperature selective catalytic reduction of NO with NH$_3$: I. Evaluation and characterization of first row transition metals, J. Catal., 2004, 221, 421–431.

20 M. A. Zanjanchi and M. K. Rashidi, Structural charge transfer in the aluminophosphate molecular sieves by diffuse reflectance spectroscopy, Spectrochim. Acta, Part A, 1999, 55, 947–954.

21 S. Bennici, A. Gervasini, N. Ravasio and F. Zaccheria, Optimization of tailoring of CuO$_x$ Species of silica alumina supported catalysts for the selective catalytic reduction of NO$_x$, J. Phys. Chem. B, 2003, 107, 5168–5176.

22 J. Y. Luo, H. Oh, C. Henry and W. Epling, Effect of C$_3$H$_6$ on selective catalytic reduction of NO$_x$ by NH$_3$ over a Cu/zeolite catalyst: A mechanistic study, Appl. Catal., B, 2012, 123–124, 296–305.

23 S. A. Yashnik, Z. R. Ismagilov and V. F. Anufrienko, Catalytic properties and electronic structure of copper ions in Cu-ZSM-5, Catal. Today, 2005, 110, 310–322.

24 O. P. Taran, S. A. Yashnik, A. B. Ayusheev, A. S. Piskun, R. V. Prihod’ko, Z. R. Ismagilov, V. V. Goncharuk and V. N. Parmon, Cu-containing MFI zeolites as catalysts for wet peroxide oxidation of formic acid as model organic contaminant, Appl. Catal., B, 2013, 140–141, 506–515.

25 S. Bennici, A. Gervasini, N. Ravasio and F. Zaccheria, Optimization of tailoring of CuO$_x$ Species of silica alumina supported catalysts for the selective catalytic reduction of NO$_x$, J. Phys. Chem. B, 2003, 107, 5168–5176.

26 T. Kharlamova, G. Mamontov, M. Salaev, V. Zaikovskii, G. Popova, V. Sobolev, A. Knyazev and O. Vodyankina, Structural charge transfer in the aluminophosphate molecular sieves by diffuse reflectance spectroscopy, Appl. Catal., A, 2013, 467, 519–529.

27 Q. Tang, S. Hu, Y. Chen, Z. Guo, Y. Hu, Y. Chen and Y. Yang, Highly dispersed manganese oxide catalysts grafted on SBA-15: synthesis, characterization and catalytic application in trans-stilbene epoxidation, Microporous Mesoporous Mater., 2010, 132, 501–509.

28 I. Spassova, T. Tsontcheva, N. Velichkova, M. Khristova and D. Nihitanova, Catalytic reduction of NO with decomposed methanol on alumina-supported Mn–Ce catalysts, J. Colloid Interface Sci., 2012, 374, 267–277.

29 Z. Qu, Y. Bu, Y. Qin, Y. Wang and Q. Fu, The improved reactivity of manganese catalysts by Ag in catalytic oxidation of toluene, Appl. Catal., B, 2013, 132–133, 353–362.

30 T. Boningari, P. R. Ettireddy, A. Somogyvari, Y. Liu, A. Vorontsov, C. A. McDonald and P. G. Smirniotis, Influence of elevated surface texture hydrated titania on Ce-doped Mn/TiO$_2$ catalysts for the low-temperature SCR of NO$_x$ under oxygen-rich conditions, J. Catal., 2015, 325, 145–155.

31 T. Boningari, D. K. Pappas, P. R. Ettireddy, A. Kotrba and P. G. Smirniotis, Influence of SiO$_2$ on M/TiO$_2$ (M = Cu, Mn, and Ce) Formulations for Low-Temperature Selective Catalytic Reduction of NO$_x$ with NH$_3$: Surface Properties and Key Components in Relation to the Activity of NO$_x$ Reduction, Ind. Eng. Chem. Res., 2015, 54, 2261–2273.

32 C. Lamberti, S. Bordiga, M. Salvalaggio, G. Spoto and A. Zecchina, XAFS, IR, and UV-vis Study of the Cu$^+$ Environment in Cu$^+$-ZSM-5, J. Phys. Chem. B, 1997, 101, 344–360.

33 M. Iwasaki and H. Shinjoh, NO evolution reaction with NO$_2$, adsorption over Fe/ZSM-5: in situ FT-IR observation and relationships with Fe sites, J. Catal., 2010, 273, 29–38.