Ultra-Long Lived Luminescent Triplet Excited States in Cyclic (Alkyl)(amino)carbene Complexes of Zn(II) Halides

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Abstract: The high element abundance and d10 electron configuration make ZnII-based compounds attractive candidates for the development of novel photoactive molecules. Although a large library of purely fluorescent compounds exists, emission involving triplet excited states is a rare phenomenon for zinc complexes. We have investigated the photophysical and -chemical properties of a series of dimeric and monomeric ZnII halide complexes bearing a cyclic (alkyl)(amino)carbene (cAAC) as chromophore unit. Specifically, [(cAAC)Zn(n-(X2)ZnX(cAAC))] (X = Cl (1), Br (2), I (3) and [ZnX2(cAAC)(NCMe)] (X = Br (4), I (5)) were isolated and fully characterized, showing intense visible light photoluminescence under UV irradiation at 297 K and fast photo-induced transformation. At 77 K, the compounds exhibit improved stability allowing to record ultra-long lifetimes in the millisecond regime. DFT/MRCI calculations confirm that the emission stems from T1XCT/LEcAAC states and indicate the phototransformation to be related to asymmetric distortion of the complexes by cAAC ligand rotation. This study enhances our understanding of the excited state properties for future development and application of new classes of ZnII phosphorescent complexes.

Introduction

Photoactive transition metal (TM) complexes provide the opportunity to exploit the properties of triplet excited states in manifold applications, such as photocatalysis, photodynamic therapy, bioimaging and (organic) light emitting devices (OLEDs).[1–4] The benefit of the TM centre in comparison to purely organic systems is usually attributed to its potential to facilitate spin-forbidden processes, such as intersystem crossing (ISC) S1→Tn or phosphorescence T1→S0 by mediating spin-orbit coupling (SOC) between excited states of different multiplicity.[5] The SOC constant strongly depends on the effective nuclear charge and, consequently, research in the above mentioned application areas has mostly been focussed on complexes of 4d and 5d heavy metals, for example, IrI[6], PtII[7] or RuII. The success of this strategy has led to tremendous scientific and technological progress, involving utilization of such compounds in commercial second-generation (phosphorescence-based) OLEDs[6,7] or as industrial high-efficiency photocatalysts.[1,8]

Although 3d TM complexes in general suffer from much weaker operative SOC in comparison to their heavier congeners, which might hamper ISC processes, significant attention has been devoted to molecular systems based on metal cations with closed-shell (3d9 or 3d10) electron configurations.[8–12] Many factors triggered research in this direction, such as high element abundance, their variable coordination geometries, redox properties, absence of metal-centred d-d* states that might induce fast non-radiative decay, and the discovery of thermally activated delayed fluorescence (TADF). The latter process is particularly useful to harvest triplet excitons by reverse ISC Tn→S1 and subsequent fluorescence, bypassing the spin-forbidden phosphorescence and reducing the importance of SOC. This mechanism is often found in CuII complexes with a donor-metal-acceptor structure, and it has been shown that up to 100% quantum yield and (sub-)microsecond lifetimes can be achieved, e.g. in cyclic (alkyl)(amino)carbene (cAAC)[13,14] or amide compounds.[15,16] The resulting radiative rate constants k of up to 105–107 s−1 are competitive and, in some cases, even faster compared to the best heavy transition metals emitters.

However, judicious choice of ligands in combination with the low energy oxidation of CuII also allow to greatly increase the MLCT character and thus the operative SOC of the excited states, which can also enhance k. For example, the very strong σ-donating and π-accepting properties of the cAAC ligands in homoleptic [CuII(cAAC)4][PF6] (Scheme 1A) lead to destabilization of the HOMOs with pronounced d-orbital character, whereas the LUMO (π* character) orbital is stabilized.[17] As a result, the T1 state is strongly coupled to singlet excited states with high oscillator strength, leading to k of 9 x 104 s−1, which is the...
highest value reported so far for purely phosphorescent Cu compounds.

Interestingly, despite the d¹⁰ electron configuration of Zn²⁺ and its high abundance (~75 ppm in the Earth’s crust), only few triplet emitters have been developed so far, but a large number of purely fluorescent Zn²⁺ complexes is known. This may be due to the fact that MLCT states cannot be formed due to the much higher ionization energy of Zn²⁺ compared to Cu¹⁺, which precludes the metal d-orbitals to be involved in the excited states and mediate SOC. Thus, predominantly intra-ligand CT (ILCT) or LC states have been studied for phosphorescence or TADF.

For instance, Adachi et al. described a Zn²⁺ TADF emitter by employing ligands with a donor-acceptor motif. The significant spatial separation of the HOMO and LUMO within the ligand scaffold led to a small singlet-triplet energy gap, and external quantum efficiencies of up to 20% were achieved when applied in OLEDs. Roeshky et al. recently reported Zn²⁺ iminophosphonamide (NPN) complexes (Scheme 1B) displaying delayed fluorescence in the solid state. In this case, twisting of the NPN units in the T₁ state is the mechanism that supports the thermally driven spin flip to the singlet excited state as the parent compound with planar ligand arrangement showed prompt fluorescence. However, the quantum yields are relatively low (ϕ < 0.05) at 297 K, and thus this study nicely pinpoints the challenges and obstacles to overcome for the design of efficient Zn²⁺-based triplet emitters.

In contrast to TADF, phosphorescence is a surprisingly rare phenomenon for zinc-based compounds, except for porphyrins. Back in 1985, Crosby et al. reported an intriguing photophysical behaviour of [Zn(SR')₃(phen)] type complexes (phen = 1,10-phenanthroline, SR' = various thiophenolate derivatives) (Scheme 1C). The T₁ state of LLCT character is populated via internal conversion from an energetically higher lying ππ* state, that is associated with an energetic barrier. However, the main limitation is dominant non-radiative relaxation at room temperature, leading to efficient emission only at low temperatures. Yam and co-workers further modified the phen ligand by introducing dithienyl moieties in the 1,10-positions to design a series of photoswitchable compounds that emit from LLCT states at 298 K. A rigid octahedral (so-called hemi-cage) complex bearing three bipyridine units connected by a mesityl cap (Scheme 1D) has also been reported to phosphoresce, in contrast to the fluorescent compound [Zn(bpy)₂][PF₆] ( bpy = 2,2'-bipyridine). The switch of emission occurring from the S₁ or T₁ state was associated with the hexa-alkyl-substituted benzene fragment, of which the authors argue to introduce a low-energy triplet state of ligand-centred (LC) character. However, our own experimental and theoretical studies on these compounds that we will report in due course suggest the emission to be fluorescence in nature instead of phosphorescence for both compounds.

In order to broaden the horizon of photoactive 3d TM complexes, we were interested in the investigation of zinc compounds featuring LLCT states, which have received only very limited attention so far. Considering their strong excited state π-acceptor properties, cAACs seem to be suitable ligands for the development of donor-Zn²⁺-acceptor compounds. Herein, we report on the synthesis, characterization and photophysical study of Zn²⁺-cAAC complexes of ZnX₂ (X = Cl, Br, I) (Scheme 1). The experimental study revealed dimeric character of ZnX₂(cAAC) adducts when crystallized from THF. In addition, these complexes show fluxional behaviour in CH₃CN solution, allowing the isolation of the monomeric complexes (ZnX₁(cAAC)(NCCH₃)₃) for X = Br and I. Both monomeric and dimeric species undergo photo-induced transformations upon UV light irradiation, forming ultra-long-lived triplet excited states of LLCT character. These first experimental and DFT/ MRCI studies on Zn²⁺ complexes bearing cAAC ligands as π-chromophores indicate great potential of this compound class for the design of long-lived states for future applications.

**Results and Discussion**

**Synthesis and characterization**

Our investigation of Zn²⁺ carbene complexes was inspired by a reaction reported by Roeský et al., who suggested that mixing of ZnCl₂ and K₂cAAC in THF at ~50°C would afford the monomeric adduct K₂cAAC·ZnCl₂ according to NMR spectroscopic studies. We further optimized the reaction conditions for its isolation and found that mixing of pre-cooled (~40°C) THF solutions of ZnCl₂ and K₂cAAC, respectively, and slow warming to room temperature over 2 h is a reliable and reproducible method for preparation of single crystals suitable for X-ray diffraction on up to 200 mg scale (Scheme 2). Instead of a monomer, a dimeric arrangement ([K₂cAAC]·ZnCl₂) (1) was isolated, which is a more intuitive outcome considering the preference of zinc(II) for a tetrahedral coordination geometry (Figure 1). In the case of ZnBr₂ and ZnI₂, slow diffusion of n-pentane into the reaction mixture also provided single crystals that showed analogous dimeric configurations ([K₂cAAC]·ZnCl₂) (X = Br (2), X = I (3)). Importantly, once crystallized, complexes 2 and 3 are not

![Scheme 1. Selected examples of phosphorescent and TADF complexes of Cu(0) and Zn(0) reported in literature.](Image)
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carbene complexes have been reported in the literature to compare the structural parameters, which are [(IPr)BrZn(μ-Br)ZnBr(IPr)][36] (IPr = 1,3-dimesityl-imidazol-4,5-dihydro-2-ylidene), [(IPr)ClZn(μ-Cl)ZnCl(IPr)][39] (IPr = 1-mesityl-3-(methoxyethyl)-imidazol-4,5-dihydro-2-ylidene) and [(BlcAAC)ClZn(μ-Cl)ZnCl(BlcAAC)] (BlcAAC = bicyclic (alkyl)(amino)carbene).[38] The zinc–halide bond lengths progressively increase in a row Cl > Br > I (Figure 1), whereas the Zn–cAAC bonds of 1–3 show only minor variations and are very similar to those found for the NHc- and BlcAAC-based dimers mentioned above (∼2.0 Å).[34–36]

Dimers 1–3 are not soluble in aliphatic or aromatic hydrocarbons, or chlorinated solvents; however, they could be redissolved in acetonitrile for further characterization by multinuclear NMR spectroscopy (see Experimental Section). Coordination of the Me6cAAC ligand to the Zn6 ion leads to a significant upfield shift of the C1 resonance in the 13C{H} NMR spectra in CD3CN to ∼250 ppm for 1–3 compared to 304.2 ppm of free Me6cAAC.[42] The chemical shift of the 1N resonance is a useful tool to analyze the degree of σ-donation and π-backbonding in the metal–carbene bond.[17] For complexes with dominant carbene–M σ-donating contribution, the nitrogen resonates at around −160 ppm, whereas for species with increased π-backbonding, the signal is considerably upfield shifted up to −300 ppm.[17] In the case of 1–3, the 13N resonances experience a very high chemical shift of ca. −146 ppm, which suggests almost pure σ-donor character of the Me6cAAC–Zn bond as expected considering the dicationic charge of the metal ion. The MALDI mass spectra (MS) of 1–3 revealed extensive fragmentation (Figure S9–S15) despite a mild ionization, including cleavage of both Zn–X and Zn–CacAAC bonds, producing cationic fragments [(Me6cAAC)ZnX]+ (m/z = 430 (X = Br), m/z = 476 (X = I)) and organic ions related to the carbene moiety, Me6cAAC+ (m/z = 286). In the case of 1, the spectrum also shows an MS peak with m/z = 807. This fragment was assigned to dimeric compounds with one abstracted chloride, [(Me6cAAC)Zn3Cl2]−.

For compound 1, crystallization from CH3CN/Et2O mixture gave the same dimeric arrangement. However, under the same conditions, the lower Zn–X bond dissociation energies of 2 and 3 lead to a higher tendency for monomerization upon solvent coordination and [ZnX1(Me6cAAC)(NCH3)][X = Br (4), I (5)] (Scheme 1) were isolated, which are also sensitive to ambient environment as found for the dimers. The Zn–X interatomic distances and Zn–CacAAC bonds of 4 and 5 have almost the same values as determined for 2 and 3 (compare Figures 2 and 3), and the NMR spectra in CD3CN of the respective halide derivatives are identical. We note that the 1H NMR spectrum contains an additional resonance with an integral intensity of ∼3 at δ = 1.96 ppm corresponding to free CH3CN.[39] Although simple ligand exchange of CH3CN for CD3CN might be responsible for the liberation of coordinated acetonitrile, photophysical measurements (see below) suggest that backward dimerization to 2 and 3 might take place in solution.

All attempts to induce analogous coordination of CH3CN and subsequent monomerization of the chlorido derivative 1 by increasing the reaction temperature led to progressive con-
version to a mixture of compounds, including 4\textsubscript{H}cAAC(H)CH\textsubscript{3}CN (6). Although it is unclear what species acts as a proton source for 4\textsubscript{H}cAAC, these decomposition reactions of the zinc(II) complex 1 probably occur by initial dissociation of 4\textsubscript{H}cAAC. A similar process, including protonation of the carbene with residual water, has been reported for \((\text{CuX}_2(\text{IDipp})) (\text{IDipp}=1,3\text{-bis}(2,6\text{-diisopropylphenyl})\text{imidazol-2-yldiene})\).\textsuperscript{119} In addition, free 4\textsubscript{H}cAAC is capable of C–H activation of toluene at elevated temperatures,\textsuperscript{40} which makes the non-metal-mediated addition of more reactive acetonitrile to the free carbene quite likely. To prove this hypothesis, we reacted free 4\textsubscript{H}cAAC with dry CH\textsubscript{3}CN and indeed found quantitative conversion to 6 (see Supporting Information), for which we confirmed its identity by single-crystal X-ray diffraction (Figure 2).

Photophysical and DFT/MRCI studies

The absorption spectra of 1 and 2 in acetonitrile are very similar in their appearance, with intense bands in the UV between \(\lambda = 200-230\) nm and extinction coefficients of \(\varepsilon = \text{ca. 25–100} \times 10^{3} \text{ M}^{-1} \text{ cm}^{-1}\) that we tentatively assign to LC transitions of the 4\textsubscript{H}cAAC ligand (Figure 3A).

Much weaker low-energy absorptions between \(\lambda = 250-280\) nm with \(\varepsilon = 2,800-4,200 \text{ M}^{-1} \text{ cm}^{-1}\) originate from in-plane \(\pi \)-type to out-of-plane \(\pi^*\)-type transitions, with some minor contribution from the halides. The respective \(S_0\rightarrow S_{1u}\) excitations depicted in Figure 4 are similarly structured as recently reported for the homoleptic complex \([\text{Cu}(4\text{H}cAAC)_2]\)\textsubscript{PF\textsubscript{6}}\textsuperscript{117}. The iodo-bridged dimer 3 shows comparable spectral absorption features, but with much higher oscillator strength at \(\lambda = 255\) nm with \(\varepsilon = 30 \times 10^{3} \text{ M}^{-1} \text{ cm}^{-1}\) and a very broad, bathochromically shifted lowest energy absorption band between \(\lambda = 280-320\) nm. We note that the participation of the terminal halide ligands in particular in the first transition increases in the series Cl < Br < I (Figure 4). Together with the enhanced singlet-triplet admixture in the iodo complex 3 due to strong SOC, the larger XCT contribution explains the marked redshift of the \(S_0\rightarrow S_{1u}\) band observed in both, experiment, and theory.

The experimental absorption spectra upon dissolving single-crystalline samples of monomeric \([\text{ZnX}_2(4\text{H}cAAC)(\text{NCCCH}_3)]\) (X = Br (4), I (5)) in MeCN are identical to those of the respective dimers 2 and 3 (Figures S17 and S18), although our DFT/MRCI studies predict non-structured characteristic bands at \(\lambda = 278\) (4) and 291 nm (5) (Figure 3B). These low energy bands would originate from XCT/LE transitions, that involve mainly the \(p\) orbitals of the halides and in-plane \(\sigma\)-type orbitals of the carbene as donors, and the C-N antibonding \(\pi^*(\text{cAAC})\) orbital as acceptor (Figure 4A–D). In addition, there are significant differences between the calculated oscillator strengths for 4/5 and the observed respective experimental absorption intensities in the region \(\lambda = 240-260\) nm, that match those of 2 and 3, respectively. Overall, these findings are likely caused by backward dimerization of 4 and 5 to structures 2 and 3 in MeCN solution.
In the single-crystalline solid-state, the dimers 1–3 show dual photoluminescence when irradiated at $\lambda_{\text{ex}} = 300$ nm, with a high energy band at $\lambda_{\text{max}} = 360$ (1), 370 (2) and 355 (3) nm, and a second, very broad low energy band ($\lambda_{\text{max}} = 451$ (1), 515 (2) and 570 (3) nm) (Figure 5A). The ratio of the two bands strongly depends on the halide, with the chloride complex 1 featuring a relatively intense high energy emission, while it is very weak for the bromide (2) and iodide (3) compounds (for details see Figure S19). We observe photochemical transformations occurring on the minute timescale, which lead to an intensity decrease of the high energy bands and an increase of the low energy emission. The latter shows pronounced shoulders in particular for 1 and 2, suggesting the formation of multiple products. Interestingly, the relative rates of photoconversion increase along the series $1 < 2 < 3$ and appear to be related to some extent to the Zn–X bond strengths (Figure 5B). We were able to determine a luminescence efficiency of the slowest transforming compound 1 of $\phi = 0.03$.

In an attempt to minimize the photochemical transformation, we studied the luminescence properties of 1–3 at 77 K (Figure 6A) and found that the steady-state emission indeed does not change over time. Albeit very weak, high-energy bands resembling the ones at 297 K are observed besides broad low energy bands that are – presumably due to reduced non-radiative decay – much more intense and defined than at room temperature. The luminescence lifetimes recorded at the respective maxima of the low energy bands are bi- or tri-exponential in the micro- to millisecond regime and decrease with increasing size of the halides (Figure 6B–D and Tables S4–S6). We therefore attribute the emission to stem from triplet excited states with weak SOC. Importantly, the time-resolved measurements reveal rise-times of 7 ms and 106 $\mu$s for 1 and 2, respectively, which can be interpreted as the timescale at which the photo-induced process occurs to reach the final emitting...
state. The relatively short lifetime of 3 does not allow to determine the rise-time with our excitation lamp set-up.

In order to identify the resulting new emissive species, we irradiated solid state samples at RT and at 77 K, and also solution samples of 1–3 overnight and tried to characterize the photoproducts by NMR spectroscopy or recrystallization from MeCN/EtO solutions. However, in solution the original dimers were reformed, and thus the photochemical process is apparently reversible and the result of structural changes in the excited state. Irradiation of single crystals for direct characterization on the diffractometer did not produce meaningful data.

The fact that the aforementioned high-energy emission band of 1–3 is observed at the earliest stage of the luminescence measurements and its intensity continuously decreases under UV light irradiation, suggests that it belongs to the original, symmetrical dimeric arrangement. This hypothesis is further supported by quantum chemical studies. Our DFT/MRSCI calculations confirm that the high-energy band is phosphorescence in nature and originates from a $^3\text{A}_i$, $\text{XCT/LE}$($S_i$) state, which is the triplet counterpart of the $S_1$ state and retains the inversion symmetry of the ground state (Figure 7A/B and Table S8).

The increasing degree of halide contribution and resulting operative SOC for the heavier congeners are reflected in shorter radiative lifetimes ($\tau = 10$ ms (1), 640 $\mu$s (2), 27 ns (3)) and more pronounced bathochromic shift of the luminescence wavelengths ($\lambda_{\text{em}} = 369$ (1), 372 (2) and 393 (3) nm), which is in good agreement with the experimental observations.

For compounds 1 and 2, we were able to identify a photoinduced excited state transformation that might be responsible for the broad low energy band observed in the experimental luminescence spectra. The $^1\text{LE}(\pi\pi^{\ast}\text{dipp})$ state forms a symmetry-broken structure on the $T_1$ potential energy surface with nearly perpendicular carbene orientations. Upon torsion of one cAAC ligand around the Zn–C bond, the excitation localizes on the other cAAC (Figure 7D). Its vertical emission wavelength of $\lambda_{\text{em}} = 482$ and 469 nm for 1 and 2, respectively, coincides nicely with the experimental emission spectra at 77 K and the early spectra at room temperature, although for 2 this $^1\text{LE}$ state appears as a shoulder (Figure 5) at 297 K besides other products (see below). The large conformational change in conjunction with the small stabilization energy of 0.04 eV with regard to the $C_2$-symmetric initial $^1\text{LE}(\pi\pi^{\ast}\text{dipp})$ structure explains the retarded build-up of the lower-energy band in the solid-state and its enormous width. Due to its local $\pi\pi^{\ast}\text{dipp}$ character without significant chlorine or zinc contributions, the phosphorescence emission of this state is very long-lived. The trend of the calculated radiative lifetimes of $\tau = 3$ s (1) and 79 ms (2) agrees with our findings (see above), although the observed lifetimes are shorter due to significant non-radiative decay, as exemplified by $\varphi_{\text{NR}} = 0.13$ for 1. For a $^1\text{LC}$ state, one would not expect significant premature relaxation at such low temperatures. However, the relatively large distortion in comparison to the optimized geometry of the electronic ground state apparently provides low energy pathways via conical intersections.

We note that population of the localized asymmetric $^1\text{LE}(\pi\pi^{\ast}\text{dipp})$ state cannot explain the highly bathochromically shifted low energy band of 3, leading us to conclude that other additional photochemical distortions occur that yet need to be identified. This interpretation would also explain the much
stronger halide-dependent kinetics of 3 depicted in Figure 5, and the multiple shoulders of the broad emission upon prolonged irradiation at room temperature for 1 and 2.

The monomers \([\text{ZnX}^{\text{caac}}(\text{NCCH}_3)] (X = \text{Br} \ (4), \ I \ (5))\) also undergo structural changes in the solid state under UV-irradiation at room temperature (Figure 8A). In contrast to their dimeric congeners 2 and 3, both high and low energy bands increase in intensity within minutes. At 77 K, the steady-state luminescence does not undergo any changes and broad emission bands with \(\lambda_{\text{max}} = 470\ (4)\) and 480 (5) nm are observed, and the time-resolved measurements reveal decay times in the milli- and microsecond regime for 4 and 5, respectively. It is noteworthy that the emission lifetime of 4 is very similar as found for the dimer 2, and that again a rise-time of 687 \(\mu\)s indicates a significant excited state distortion. Bearing in mind that the calculated local \(^3\pi\pi^*\) of all dimers is halide-independent with regard to the emission wavelength of ca. 470 nm, but that the calculated decay times do decrease with the heavier analogues (Table S8), it is quite likely that for 4 and 5 the excited state transformation also yields a \(^3\text{LE}_{\text{caac}}\) state at 77 K. At room temperature, more diverse photochemical changes can occur, which may include halide or solvent dissociation and even changes of intermolecular interactions.

**Conclusions**

In this work, we have reported the synthesis and photophysical properties of a series of zinc(II) cyclic alkyl(amino)carbene complexes comprising both dimeric structures with bridging halide ligands, namely \([\text{Zn}^{\text{caac}}(\mu-X)\text{ZnX}^{\text{caac}}] (X = \text{Cl} \ (1), \ \text{Br} \ (2), \ I \ (3))\), and monomeric species \([\text{ZnX}^{\text{caac}}(\text{NCCH}_3)] (X = \text{Br} \ (4), \ I \ (5))\) containing coordinated acetonitrile. The bromido and iodo compounds appear to have a tendency to switch between the monomeric and dimeric forms depending on the reaction conditions, i.e., solvent and concentration.

All compounds show visible photoluminescence from triplet excited states with ultra-long lifetimes (up to ms region), where the caac acts as \(\pi\)-chromophore ligand. At room temperature, photochemical transformations in the excited state occur, leading to multiple emitting species, which are reversed into the original structures upon dissolution and recrystallization. However, at 77 K the transformation appears to be very specific, allowing to elucidate the reaction path by combined experimental and theoretical studies.

According to our DFT/MRCl calculations, the \(^3\text{XCT/LE}_{\text{caac}}\) excitations represent minima on the \(T_1\) potential energy surface of the dimers, giving high-energy phosphorescence with lifetimes in the milli- to microsecond range. However, the global minima on the \(T_1\) potential energy surfaces result from localized \(\pi^*\) excitation accompanied by a large conformational change of the nuclear framework, i.e. rotation of the caac, but maintains its dimeric structure. While the relaxation energy to reach that minimum is small in the triplet state, the large structural distortion is very unfavourable in the electronic ground state, thus leading to a strong redshift of the long-lived phosphorescence emission. The monomeric complexes 4 and 5 also show excitations of \(^3\text{LE}(\pi\pi^*_{\text{caac}})\) at 77 K, but undergo more diverse transformations at room temperature.

Apart from porphyrin-based systems, phosphorescent Zn\(^{II}\) complexes are extraordinarily rare. Bearing in mind that the ultra-long luminescence lifetimes of the monomeric and dimeric species are mainly the result of the \([\text{Zn}^{\text{caac}}(\text{X})\text{ZnX}^{\text{caac}}]\) acting as chromophore unit, modification of the coordination sphere should provide opportunities for further improvement of the photophysical properties. Also, the identification of the mechanism of photo-induced transformation suggests potential routes for enhanced photostability, for example, halide exchange or increased steric demand of the ligand sphere. Our study thus represents an important step towards the design of new photoactive 3d metal complexes, and we will report on the mentioned modifications and their effects on the photophysics and photoreactivity of Zn\(^{II}\) carbene compounds shortly.

**Experimental Section**

**Materials and Techniques**

All operations were performed under an argon atmosphere by using conventional Schlenk-line techniques or glovebox. The solvents were dried using Technology Inc. Pure-Solv system or standard methods and degassed prior to use. The \([\text{Zn}^{\text{caac}}] \) was prepared according to a procedure reported in the literature. Zinc halides were dried by heating (up to 100 °C) under high vacuum. All other starting materials were available commercially and were used without further purification.
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H5QC, 1H-1C HSMB, 1H-1H COSY and 1H-15N HMBC were measured at 300 K on Bruker 500 Avance or Bruker 600 Avance spectrometers. The (C\(\text{H}_3\))N\(\text{H}_5\) shifts are given in ppm relative to residual signals of the solvent [\(\text{H}_2\)]. 13C CD(CN) (1.94, 118.69 ppm). MALDI-TOF MS spectra were recorded on a Thermo Scientific Q Exactive in positive ion mode connected with a TransMIT AP-SPMALDI containing a laser with a wavelength of 343 nm. Best results were achieved with an attenuator of 35°.

**Synthesis and Characterization**

\([\text{ZnBr}_2\text{[acac]}]\) (4)

100 mg (98 µmol) of 2 was dissolved in 2 mL of CH\(_3\)CN and the mixture was stirred at room temperature for one hour. After that, the solution was concentrated by vacuum evaporation of solvent to a volume of 0.5 mL and crystallization was induced by slow diffusion of Et\(_2\)O as antisolvent. The product was collected, washed with 3 mL of Et\(_2\)O and vacuum-dried. Yield: 94 mg (170 µmol, 87%). Elemental analysis (% calc. for C\(_{24}\)H\(_{33}\)Br\(_2\)ZnBr\(_2\) 47.90; H: 6.21; N: 5.08. Found: C: 48.1; H: 6.3; N: 5.2.

\(\text{H}^1\) NMR (600 MHz, CD\(_2\)CN): \(\delta\) ppm = 7.84 (t, \(\text{J}_{\text{HH}}\) = 7.1 Hz, 1H, H, \(\text{C}(\text{H})_{\text{Br}}\)), 7.39 (d, \(\text{J}_{\text{HH}}\) = 7.8 Hz, 7H, 2H, \(\text{C}(\text{H})_{\text{Br}}\)), 2.83 (sept., \(\text{J}_{\text{HH}}\) = 6.8 Hz, 6H, \(\text{CH}_3\)), 2.45 (s, 2 CH, Dipp), 2.55 (74.2 ppm. Single crystals suitable for X-ray diffraction analysis were prepared by slow diffusion of Et\(_2\)O into a solution of 4 in CH\(_3\)CN.

\([\text{ZnBr}_2\text{[acac]}]\) (5)

Compound 5 was prepared in similar manner as described for 4 but with 100 mg (83 µmol) of 3. Yield: 90 mg (139 µmol, 84%). Elemental analysis (% calc. for C\(_{22}\)H\(_{30}\)N\(\text{Zn(II)}\) 40.92; H: 5.31; N: 4.34. Found: C: 40.9; H: 5.3; N: 4.6. 1H NMR (600 MHz, CD\(_2\)CN): \(\delta\) ppm = 7.51 (t, \(\text{J}_{\text{HH}}\) = 7.1 Hz, 1H, H, \(\text{C}(\text{H})_{\text{Br}}\)), 7.42 (d, \(\text{J}_{\text{HH}}\) = 7.7 Hz, 7H, 2H, \(\text{C}(\text{H})_{\text{Br}}\)), 2.51 (sept., \(\text{J}_{\text{HH}}\) = 6.7 Hz, 6H, \(\text{CH}_3\)), 2.10 (2CH, Dipp), 29.5 (m, 2CH, Dipp), 25.5 (2 CH, Dipp). 11C NMR (125 MHz, CD\(_2\)CN): \(\delta\) ppm = 246.8 (s, 1 C, \text{C\text{eternous}}), 146.8 (s, 2 C, CH\(_2\)), Dipp), 134.9 (s, 1 C, Dipp), 131.2 (s, 1 C, Dipp), 126.3 (s, 2 C, CH\(_3\)), Dipp).

\([\text{ZnBr}_2\text{[acac]}]\) (6)

36.5 mg of \(\text{ZnBr}_2\text{[acac]}\) (129 µmol) was dissolved in 2 mL of CH\(_3\)CN and the solution was stirred at room temperature for one hour. After that, the solvent was vacuum evaporated to give white powder of compound 6. Yield: 43.5 mg (127 mg, 98.5%). 1H NMR (600 MHz, CD\(_2\)CN): 7.28 (m, 1H, H, Dipp), 7.19 (m, 2H, Dipp), 3.82 (dd, \(\text{J}_{\text{HH}}\) = 3.9 Hz, \(\text{J}_{\text{HH}}\) = 10.5 Hz, 1H, C(CH\(_3\))CH\(_2\)), 3.72 (sept., \(\text{J}_{\text{HH}}\) = 6.8 Hz, 1H, CH\(_3\)), 3.27 (sept., \(\text{J}_{\text{HH}}\) = 6.8 Hz, 1H, CH\(_3\)), 2.22–2.10 (AB pattern, \(\Delta\nu_{\text{ab}}\) = 14.7 Hz, \(\nu_{\text{ab}}\) = 2.5 Hz, CH\(_2\)), Dipp, 2.01 (AB pattern, \(\Delta\nu_{\text{bb}}\) = 0.1 ppm. 1H (\(\text{J}_{\text{HH}}\) = 13.1 Hz, \(\text{J}_{\text{HH}}\) = 1.37 Hz, 3H, CH\(_3\)), 1.34 (3H, CH\(_3\)), CH\(_3\)), 1.32 (d, \(\text{J}_{\text{HH}}\) = 6.8 Hz, 3H, CH\(_3\)), CH\(_3\)), 1.25 (m, 6H, CH\(_3\)), Dipp, 0.86 (3H, CH\(_3\)), Dipp. 11C NMR (125 MHz, CD\(_2\)CN): \(\delta\) ppm = 152.2 (s, 1 C, CH\(_3\)), Dipp), 151.27 (s, 1H, CH\(_3\)), Dipp).
Singlet excited-state geometries of the isolated complexes were calculated with the strength of the spin-allowed transitions were calculated with the level of quasi-degenerate perturbation theory using Spock. The heavier halogenides, spectra were determined including SOC at the Zn ion, the relativistic small-core effective core potential (ECP) and C, N, H and Cl, were provided free of charge by the joint Cambridge Crystallographic Data Centre and Fachinformationszentrum Karlsruhe Access Structures service.

**Computational Studies**

The quantum chemical calculations were performed using the Turbomole program package for all geometry optimizations and for the generation of the molecular orbitals and two-electron integrals while the DFT/MRCI program and the spin–orbit coupling kit Spock were applied for computing the spectra and other photophysical properties of the complexes. Solvent effects were accounted for by the implicit solvent model COSMO using dielectric constants of $\varepsilon = 4.711$ for chloroform and $\varepsilon = 37.50$ for acetonitrile solutions. To model the adducts 4 and 5, one CH$_2$CN molecule was explicitly included in the quantum chemical treatment in addition to the COSMO surrounding. The equilibrium geometries of the electronic ground states were determined with Kohn–Sham density functional theory (DFT) employing the BHLYP functional including empirical dispersion corrections.

Singlet excited-states geometries of the isolated complexes were optimized at the level of time-dependent DFT (TDDFT). Triplet-state geometry optimizations additionally made use of the Tamm-Dancoff approximation. Solvent effects on the nuclear arrangement of the complexes in the excited states were neglected. Carbon, nitrogen, chlorine and hydrogen atoms were represented by def-SV(P) basis sets from the Turbomole basis set library. For the Zn ion, the relativistic small-core effective core potential (ECP) defpp-ECP and the associated contracted 6–3sp3d basis set were chosen.

The heavier halogenides employed relativistic small-core ECPs as well (ecp-10-mdf for bromine and ecp-28-mdf for iodine) in conjunction with the cc-pVTZ-PP basis for bromine and the SVP basis for iodine. Electronic excitation energies and oscillator strengths of the spin-allowed transitions were calculated with the DFT/MRCI method employing the redesigned R2016 Hamiltonian.

Due to the large size of the complexes, a tight configuration selection criterion (threshold 0.80 $E_r$) was used. To compute the absorption spectra, the DFT/MRCI secular equations were solved for 20 singlet and 20 triplet roots, except for the C$_2$-symmetric dimers 2 and 3 where more states were required to cover the spectral range up to 200 nm. In addition to 20 $A_v$ and 20 $A^*$ states, 20 (30) $A_v$ and $A^*$ states were computed for compound 2 (3), respectively. Because of the large spin-orbit coupling (SOC) constants of the heavier halogenides, spectra were determined including SOC at the level of quasi-degenerate perturbation theory using Spock. The Breit-Pauli SOC integrals, employed for C, N, H and Cl, were evaluated in the atomic mean-field approximation using AMFI, whereas spin-orbit ECPs were used for Zn, Br and I. To avoid convergence problems of the perturbation expansions, phosphorescence rate constants were determined using the multireference spin-orbit configuration interaction (MRSOCl) approach implemented in Spock.

**Photophysical Measurements**

To avoid decomposition of samples and quenching of photoluminescence, all photophysical measurements were performed under an inert atmosphere of argon. The acetonitrile used for electronic spectra measurement was stored over activated 4 Å molecular sieves and purified (to remove traces of molecular sieves and oxygen) by vacuum-transfer technique prior to use. Absorption spectra were recorded on a Cary 5000 UV-Vis-NIR spectrophotometer in the range 200–800 nm using quartz cuvettes with 1 cm path length. The concentration of samples was adjusted in the range of 1–7 $\times$ 10$^{-3}$ mol/l. Emission and excitation spectra with lifetimes were recorded on an Edinburgh Instrument FLS1000 spectrometer, equipped with a double monochromator for the excitation and emission pathways and a red-sensitive photomultiplier (PMT-980, 200–880 nm) as detector. For measurements of emission and excitation spectra, 450 W Xenon arc lamp was used, whereas lifetimes were recorded using a pulsed 60 W Xenon microsecond flashlamp with a repetition rate of 1–100 Hz. Low temperature measurements were performed using an Oxford Optistat DN cryostate.

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**Conflict of Interest**

The authors declare no conflict of interest.

**Data Availability Statement**

The data that support the findings of this study are available in the supplementary material of this article.

**Keywords:** carbenes · DFT/MRCI · luminescence · phosphorescence · zinc

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