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Recent Progress in g-C₃N₄ based low cost photocatalytic system: activity enhancement and emerging applications

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Graphitic CₓNₓ (g-CₓNₓ) has continuously attracted attentions since it was reported as a metal free semiconductor for water splitting. However, its ability for evolving hydrogen from water is significantly dependent on the use of noble metal co-catalyst, mainly Pt. In recent years, a good progress has been achieved in developing co-catalysts containing earth abundant elements only for constructing low cost and efficient g-CₓNₓ based photocatalytic systems. Besides, exfoliation of bulk g-CₓNₓ into two dimensional g-CₓNₓ nanosheets offers large surface area and exposed active sites, which are beneficial for activity enhancement. Furthermore, oxygen evolution and CO₂ photoreduction over g-CₓNₓ have caught increasing interests due to the demand to achieve overall water splitting and conversion of CO₂ into chemicals and fuels. In this mini-review, we will briefly summarize the latest research works on g-CₓNₓ based photocatalytic systems during the last three years with emphasis on the progress achieved in enhancing hydrogen evolution activity of g-CₓNₓ by loading noble metal free co-catalysts, exfoliating bulk g-CₓNₓ into nanosheets, and the application of g-CₓNₓ system in photocatalytic O₂ evolution and CO₂ reduction.

1. Introduction

The energy and environment problems have become more and more severe over the recent years due to the overuse of fossil fuels and uncontrolled CO₂ emission from the combustion of fossil fuels. It is of utmost urgency to find green technology to address these concerns. One promising method is the utilization of solar energy with the help of semiconductor photocatalysts. A process known as photocatalytic water splitting converts water into H₂ which is considered to be a promising clean energy to replace fossil fuels; another process called artificial photosynthesis is focused on synthesizing hydrocarbon molecules from CO₂ which is a mimic of the natural photosynthesis in green plants.

Since Fujishima reported a photoelectrochemical (PEC) water splitting process by TiO₂ photoanode, researchers have spent efforts on improving the conversion efficiency of this process. Various semiconductors, e.g. oxide, (oxy)nitride, (oxy)sulfide have been shown capable of splitting water under light irradiation with suitable co-catalysts deposited on the surface. Unfortunately, the relatively low efficiency and noble metal containing materials of these systems make them unfavourable for the practical usage. The price of H₂ by this method is still much higher than expected. Further improving the solar to hydrogen (STH) efficiency and reducing the cost are of particular significance to realize the potential of this technology. Compared to water splitting, the study of CO₂ photoreduction is still in its infancy partially due to the fact that it is energetically more difficult to reduce CO₂ than proton. CO₂ reduction requires a much larger driving force than water reduction (Table 1). In addition, CO₂ reduction is a multiple electron process, which makes it kinetically harder to proceed. The selectivity of products is also an important aspect in CO₂ reduction. After all, the development of photocatalytic water splitting and CO₂ reduction depends on the efficient utilization of solar power and the enhancement of catalytic conversion of water and CO₂ into fuels and chemicals. These serve as the criteria for the selection of suitable photocatalysts.

Table 1 CO₂ reduction potentials (reported at pH 7). Adapted with permission from ref. 8. Copyright 2009 American Chemical Society.

| Reaction                        | E° (V) vs SCE |
|--------------------------------|--------------|
| CO₂ + 2H⁺ + 2e⁻ Y HCO₂H         | -0.85        |
| CO₂ + 2H⁺ + 2e⁻ Y CO + H₂O     | -0.77        |
| CO₂ + 4H⁺ + 4e⁻ Y C + 2H₂O    | -0.44        |
| CO₂ + 4H⁺ + 4e⁻ Y HCHO + H₂O  | -0.72        |
| CO₂ + 6H⁺ + 6e⁻ Y CH₃OH + H₂O | -0.62        |
| CO₂ + 8H⁺ + 8e⁻ Y CH₄ + 2H₂O  | -0.48        |

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Graphitic carbon nitride (g-C$_3$N$_4$), also known as melon, is the most stable allotrope among different carbon nitride materials. It can be synthesized via thermal condensation of low cost nitrogen rich precursors such as cyanamide, dicyandiamide, melamine, thiourea and urea. It has a graphene-like structure consisting of two dimensional frameworks of tri-s-triazine connected via tertiary amines. This unique structure and high degree of condensation make g-C$_3$N$_4$ stable at elevated temperature as high as 600 °C and in different chemical environment (acid, base or organic solvent). Furthermore, this tri-s-triazine ring structure makes it an indirect semiconductor with a band gap of ~2.7 eV, corresponding to an optical wavelength of 460 nm in the visible light range.\textsuperscript{9,10} In 2009, Wang et al reported that photocatalytic water splitting can be achieved using this metal free polymeric g-C$_3$N$_4$.\textsuperscript{11} Since then, enormous attention has been drawn on g-C$_3$N$_4$ because it is ideal for the construction of low cost photocatalysts. Nevertheless, the photoactivity of g-C$_3$N$_4$ suffers from several main drawbacks. Firstly, the relatively large band gap (2.7 eV) limits its effective utilization of visible light of longer wavelengths. It was estimated that a photocatalyst with a band gap as narrow as 1.8-2.0 eV and suitable band positions would be desirable from the viewpoint of both solar energy harvesting and surface kinetics of water splitting reactions.\textsuperscript{12} Secondly, the low charge carrier mobility inhibits the separation and transportation of electrons and holes. It was revealed by both first principle calculations and experimental results that the charge carrier mobility of the pristine g-C$_3$N$_4$ can be enhanced by doping non-metal elements which widen the valence band (VB) of g-C$_3$N$_4$.\textsuperscript{13-16} Furthermore, the surface inertness of g-C$_3$N$_4$ due to the nature of the covalent bonding leads to the low reaction rates of hydrogen and oxygen evolution half reactions.\textsuperscript{17} The surface area of bulk g-C$_3$N$_4$ is generally small since it is prepared by condensation of organic precursor compounds under high temperature. There also exist rich grain boundary defects on the surface of g-C$_3$N$_4$ resulted from incomplete condensation, which may cause the recombination of excited charges.\textsuperscript{17}

As a result, the efforts that have been made to address these issues include modification of the electronic structure of g-C$_3$N$_4$ by integrating certain organic molecules into its ring structure;\textsuperscript{18, 19} doping metal or non-metal ions in the bulk structure of g-C$_3$N$_4$ to form impurity level in its forbidden band to improve its visible light absorption property;\textsuperscript{14, 15, 20-23} and suppression of electron hole recombination by forming composites of g-C$_3$N$_4$ and other semiconductors with suitable band alignment or carbon materials to improve the charge transportation.\textsuperscript{24-28} Sensitization of the pristine g-C$_3$N$_4$ by organic dyes or quantum dots is also efficient to enhance its performance under visible light.\textsuperscript{29, 30} Besides, morphology control of g-C$_3$N$_4$ to increase the specific surface area as well as to expose active sites for photocatalytic reaction is another strategy to achieve photocatalytic activity improvement. Mesoporous g-C$_3$N$_4$ (mpg-C$_3$N$_4$) is commonly prepared using SiO$_2$ nanospheres as the template.\textsuperscript{31} In addition, g-C$_3$N$_4$ with other morphologies such as spherical nanoparticle,\textsuperscript{32} nanorod,\textsuperscript{33, 35} hollow sphere,\textsuperscript{36, 37} and ordered 3D structure\textsuperscript{29} have been prepared using template or non-template methods. There are many literature reports demonstrating the use of such modifications to improve the performance of g-C$_3$N$_4$ system, which has been shown as a promising material for solar fuel production.

In this mini review, we intend to summarize the recent progress on the development of efficient and low cost g-C$_3$N$_4$ systems on a few important aspects including deposition of noble metal free cocatalysts on g-C$_3$N$_4$, exfoliation of bulk g-C$_3$N$_4$ into g-C$_3$N$_4$ nanosheets and the application of g-C$_3$N$_4$ in water oxidation, overall water splitting and CO$_2$ photoreduction. These challenging research areas have drawn increasing interests from many research groups including our own. Some review articles have been recently published focusing on the fabrication, chemical modification, bandgap engineering and heterostructure formation of g-C$_3$N$_4$ based photocatalysts.\textsuperscript{5, 9, 10, 17, 24, 30} We believe this review would be a good complement to the literature on g-C$_3$N$_4$ based photocatalyst systems.

2. Noble metal free g-C$_3$N$_4$ system

Depositing suitable cocatalyst on the surface of semiconductor is of great importance to improve its efficiency. Cocatalysts significantly enhance the surface reaction rate by lowering the activation energy and facilitate the separation of electrons and holes by forming Schottky junction between semiconductor and cocatalyst. The role of cocatalyst on photocatalysis and photovoltaic photocatalysis has been well reviewed by Yang et al.\textsuperscript{40}

Pt is so far the most efficient and commonly used cocatalyst of g-C$_3$N$_4$ for hydrogen production. Despite that a series of non-noble metal containing hydrogen evolution reaction (HER) electrocatalysts have been explored, only a few of them are successfully deposited on g-C$_3$N$_4$ as cocatalysts.\textsuperscript{42} To further reduce the cost of the g-C$_3$N$_4$ based photocatalysts and make them available for industrial usage, noble metal free cocatalysts are highly needed.

As a pioneering work in this area, Dong et al reported in 2012 that a molecular cocatalyst Ni(TEOA)$_2$Cl$_2$ can be used in the system of g-C$_3$N$_4$ to replace Pt for stable hydrogen evolution.\textsuperscript{43} In this system, g-C$_3$N$_4$ plays the main role of light absorber. Ni(TEOA)$_2$Cl$_2$ is formed via self-assembly of Ni$^{2+}$ ions and triethanolamine (TEOA) which is the sacrificial reagent in solution. A similar nickel containing complex, nickel-thiourea-triethylamine (Ni-TU-TETN) formed in situ during photocatalytic reaction, was also reported to be an active cocatalyst on the surface of g-C$_3$N$_4$ for hydrogen production.\textsuperscript{43} But the quantum efficiency (QE) obtained using Ni-TU-TETN/g-C$_3$N$_4$ is only 0.2% at 400 nm irradiation, which is much smaller than 1.51% at 400 nm reported for Ni-TEOA/g-C$_3$N$_4$. Co(dmglH)$_2$pyCl has been studied as a electrocatalytic HER catalyst and was used for photocatalytic hydrogen production in a homogeneous system as well.\textsuperscript{44} When coupled with g-C$_3$N$_4$ in TEOA solution, it was able to produce hydrogen under visible light. The highest QE of this system was 0.62% at 365 nm irradiation.\textsuperscript{45} However, the activity decreased dramatically after 8 h of photocatalytic reaction due to the decomposition of Co(dmglH)$_2$pyCl. To overcome this, a more stable Ni(dmglH)$_2$ HER catalyst was used to construct a Ni(dmglH)$_2$/g-C$_3$N$_4$ photocatalyst. Ni(dmglH)$_2$ sub-microcrystals were grown on g-C$_3$N$_4$ via a simple precipitation method. Its activity remained almost the same after three recycle runs, indicating the enhanced stability of this system although the activity needs to be further improved.\textsuperscript{46}
Besides the transition metal containing complexes, some inorganic compounds have also been coupled with g-C₃N₄ for the construction of noble metal free g-C₃N₄ photocatalysts. For example, our group has successfully deposited NiS on g-C₃N₄ via a facile hydrothermal method. An aqueous solution of nickel acetate (NiAc) and thiourea was used for the synthesis of NiS in the presence of mpg-C₃N₄, allowing Ni²⁺ and sulfur precursor to enter into the pores of mpg-C₃N₄ for the formation of nanosized and well-dispersed NiS cocatalyst. The formation of NiS nanoparticles was confirmed by high resolution transmission electron microscopy (HRTEM) analysis, element mapping and X-ray photoelectron spectrum (XPS). The typical HRTEM images and elemental mapping results are shown in Fig. 1A. The loading of NiS has a significant effect on the activity and the optimum NiS loading was found to be 1.25 wt% (Fig. 1B). For comparison, the activity of a mixture of mpg-C₃N₄ and pre-synthesized NiS was much lower than that of NiS/mpg-C₃N₄ which is due to the poor interface between NiS and mpg-C₃N₄ and the relatively larger particle size of pre-synthesized NiS. A high QE of 1.9% at 440 nm was obtained over the optimized NiS/mpg-C₃N₄ photocatalyst. The same photocatalyst demonstrates a good stability with 84% of the initial activity retained after 4 runs of photoreaction. Subsequently, Chen et al reported the preparation of NiS/g-C₃N₄ via ion exchange that converts Ni⁴⁺/g-C₃N₄ to NiS/g-C₃N₄ using Na₂S. Another type of nickel sulfide supported on g-C₃N₄, NiS@g-C₃N₄, was prepared via a hydrothermal method using NiAc and thiourea in the presence of g-C₃N₄. The activity of the optimized NiS@g-C₃N₄ system is 4 times that of Pt@g-C₃N₄. However, the activity dropped to 60% of the initial value after five cycles. These results indicate that nickel sulfides represent promising cocatalysts for the construction of noble metal free g-C₃N₄ photocatalysts for hydrogen evolution. In addition, cobalt sulfide has been evaluated as the cocatalysts for mpg-C₃N₄ and exhibited lower activities than NiS under the same conditions.

In the family of metal sulfides, MoS₂ and WS₂ nanosheets have received a lot of interests as efficient noble free HER catalysts for both electrocatalytic and photocatalytic reactions. MoS₂ or WS₂ have also been loaded on g-C₃N₄ to act as cocatalysts. It is expected that the nanosheet morphology of these catalysts can facilitate charge transfer between g-C₃N₄ and the catalyst. Furthermore, the thin layer structure of the nanosheets can minimize the light blocking effect by cocatalysts in the aggregated form. Both MoS₂/mpg-C₃N₄ and WS₂/mpg-C₃N₄ were prepared by impregnation method. As schematically shown in Fig. 2A, (NH₄)₂MoS₄ or (NH₄)₂WS₄ was impregnated on mpg-C₃N₄ and the subsequent sulfidation was carried out in H₂S or H₂S/H₂ atmosphere at an elevated temperature. MoS₂ nanosheets were observed in HRTEM images (Fig. 2B). As can be seen in Fig. 2C, the activity of the optimized MoS₂/mpg-C₃N₄ was found higher than that of Pt/mpg-C₃N₄ and a QE of 2.1% was obtained at 400 nm. The stability of this system is also good with 80% of the initial activity maintained after recycling for four runs. Compared to MoS₂/mpg-C₃N₄, the activity of WS₂/mpg-C₃N₄ is relatively lower and decreased dramatically after recycled runs.

Hydroxides such as Ni(OH)₂ and Cu(OH)₂ were also reported to be capable of evolving hydrogen when coupled with g-C₃N₄. The preparation method of hydroxide@g-C₃N₄ was simply precipitating metal cations by NaOH with pre-added g-C₃N₄ followed by aging for a certain period of time. Both Ni(OH)₂@g-C₃N₄ and Cu(OH)₂@g-C₃N₄ were found to be stable and efficient for photocatalytic hydrogen production. A relatively high QE of 1.1% was obtained over Ni(OH)₂@g-C₃N₄ at 420 nm. XPS study for the hydroxide@g-C₃N₄ samples before and after photoreaction revealed that the in situ formed Ni⁰ or Cu²⁺ cluster via reduction of metal ions by photo electrons is the key to the hydrogen production process. Very recently, Zhang et al prepared a core shell structured Ni/NiO loaded on g-C₃N₄ by the impregnation method for hydrogen production. Ni⁴⁺ on g-C₃N₄ was first reduced to Ni nanoparticles in H₂ flow and then the surface of Ni nanoparticles were oxidized to NiO after being annealed in air leading to the formation of core shell nanostructure. Ni/NiO@g-C₃N₄ showed improved activity for photocatalytic hydrogen production compared with the pristine g-C₃N₄, Ni@g-C₃N₄ and NiO@g-C₃N₄ due to the enhanced charge transfer across the interface. The overpotential for HER was also
reduced for Ni/NiO-g-C₃N₄, confirming that the Ni/NiO core/shell structure is a suitable catalyst for HER.

In short, various noble metal free g-C₃N₄ based photocatalysts have been investigated. It is promising to find that the activities of some photocatalysts approach or surpass those of Pt/g-C₃N₄. However, it is still challenging to prepare stable and efficient noble metal free g-C₃N₄ photocatalysts with their activities comparable or better than other semiconductor based systems. The photoelectric current and the photocatalytic efficiency are still low due to the drawbacks mentioned before. One of the possible solutions could be the development of new preparation methods to enhance the interface between the cocatalysts and g-C₃N₄. Furthermore, similar to the case of graphene, the electronic and morphological properties of g-C₃N₄ itself could be tuned by exfoliation to enhance the photocatalytic performances, which is discussed in the next section.

3. Generation of g-C₃N₄ nanosheets by exfoliation toward enhanced photocatalytic activity

Among various methods for control and modification of the morphology of g-C₃N₄, exfoliation is a simple and facile process which has been shown effective to enhance the photocatalytic performance of g-C₃N₄. Bulk g-C₃N₄ has a graphite like structure which contains elementary layers built up from ring structure of carbon nitride and van der Waals interaction between the layers. This unique structure makes g-C₃N₄ possible to be exfoliated into a graphene like single layered nanosheets. By exfoliating bulk g-C₃N₄ into nanosheets, the specific surface area can be increased and the catalytic centers are exposed. Furthermore, due to the quantum confinement effect, the conduction band (CB) position of g-C₃N₄ nanosheets can be shifted to more negative values than that of its bulk compartment, which provides a larger driving force for photocatalytic reaction.

The exfoliation methods of g-C₃N₄ can be classified as thermal exfoliation and liquid exfoliation. Niu et al reported the preparation of graphene like g-C₃N₄ nanosheets by thermal etching of bulk g-C₃N₄. After the g-C₃N₄ sample was thermally oxidized in air at 500 °C for 2 h, g-C₃N₄ nanosheets with the thickness ranging from 1.62 nm to 2.62 nm were obtained corresponding to 4 to 7 carbon nitride layers. Compared to bulk g-C₃N₄, the optical absorption of g-C₃N₄ nanosheets exhibited 20 nm blue shift which is due to the quantum confinement. The surface area of g-C₃N₄ nanosheets increased from 50 m²/g to 306 m²/g. The synergistic effect of the higher surface area and the more negative CB position results in a significant increase in its hydrogen production activity. Xu et al further modified this thermal exfoliation method by firstly preparing a NH₄Cl/g-C₃N₄ composite with NH₃Cl intercalated in the interlayer space via hydrothermal reaction. Then the composite was annealed in N₂ to exfoliate g-C₃N₄ into nanosheets by the evolved gaseous NH₃. The preparation method is schematically illustrated in Fig. 3A. The surface area of g-C₃N₄ nanosheets is 10 times that of the bulk g-C₃N₄ prepared using dicyanamide as the precursor. As shown in Fig. 3B, the average thickness of the as-prepared g-C₃N₄ nanosheets is around 2.8 nm measured by atomic force microscopy (AFM). Although thermal exfoliation is considered to be a low cost, large scale and environmentally friendly method for preparing g-C₃N₄ nanosheets, the highest yield of the nanosheets obtained so far is 6 wt% of the starting bulk g-C₃N₄. 

![Fig. 3 (A) Schematics of thermal exfoliation of g-C₃N₄ into nanosheets and the optical images of g-C₃N₄ before and after exfoliation; (B) typical TEM images of bulk g-C₃N₄ and g-C₃N₄ nanosheets and AFM image and thickness profile of exfoliated nanosheets. Reproduced from ref. 58 with permission from The Royal Society of Chemistry.](image)

Liquid exfoliation is a more commonly used method and is done by ultrasonication of bulk g-C₃N₄ material in a suitable solvent. After this process, the exfoliated nanosheets can be easily separated from the remaining bulk g-C₃N₄ by centrifugation. It is considered as a facile method for the preparation of 2D materials. Zhang et al firstly reported the preparation of g-C₃N₄ nanosheets using water as the solvent during ultrasonication. As there is dangling hydrogen in the carbon nitride layer of g-C₃N₄, it was supposed that a polar solvent like H₂O is effective for the swelling and exfoliation of g-C₃N₄. The swelling and exfoliation process is shown in Fig. 4A. Besides H₂O, formamide, dimethylformamide, ethanol and methanol were used as exfoliation solvents. Only H₂O was found effective and the concentration of the resulting nanosheets suspension was measured to be 0.15 mg/mL. The change of UV-vis spectrum and photoluminescence (PL) spectrum before and after exfoliation can be attributed to the enlarged bandgap induced by quantum confinement (Fig. 4B). As shown in Fig. 4C, the thickness of as-prepared g-C₃N₄ nanosheets is ~2.5 nm corresponding to 7 atomic layers and the lateral length of the nanosheets ranges from 70 nm to 140 nm. The as-prepared nanosheets were used for bioimaging due to its high PL quantum yield and biocompatibility. Inspired by this result, researchers have adopted the liquid exfoliation method for the preparation of g-C₃N₄ nanosheets and expanded the application of g-C₃N₄ nanosheets to other areas. For example, Sun and co-workers applied the same method and used the resulting nanosheets as an efficient fluoro-sensor for detection of Cu²⁺. Cheng et al exfoliated g-C₃N₄ in water and loaded Au nanoparticles on the exfoliated g-C₃N₄ nanosheets via photodeposition. The composite was found to be able to efficiently degrade methyl orange under visible light. In addition, Zhang et al reported that the reaction pathway for photocatalytic phenol degradation is changed when exfoliated nanosheets were used as the photocatalyst. In this case, g-C₃N₄ nanosheets with a thickness of 2 nm was prepared by
sonication in water. The change of bandgap and CB position were measured by UV-vis spectrum and Mott-Schottky plot. It was found that oxygen was reduced to \( \text{H}_2\text{O} \) on g-C\(_3\)N\(_4\) nanosheets via two electron transfer process while \( \text{O}_2 \) was formed on the surface of bulk g-C\(_3\)N\(_4\) via one electron transfer process. This was attributed to the formation of 1,4-endoperoxide species on the surface of g-C\(_3\)N\(_4\) nanosheets.

The two electron transfer pathway promoted the efficient separation of excited electrons and holes and facilitated the formation of reactive species which resulted in the enhanced photocatalytic activity for phenol degradation.

Besides water, other polar solvents or mixed solvent have also been investigated for the exfoliation of g-C\(_3\)N\(_4\). For instance, isopropanol alcohol (IPA) was found another suitable solvent for exfoliation of g-C\(_3\)N\(_4\) into nanosheets due to the surface energy matching of IPA and g-C\(_3\)N\(_4\).\(^6\) An extended sonication duration of 10 h yielded g-C\(_3\)N\(_4\) nanosheets with an average thickness of 2 nm which is thinner than those prepared using \( \text{H}_2\text{O} \) as the solvent. Similarly, the exfoliated g-C\(_3\)N\(_4\) nanosheets exhibited a high surface area of 384 m\(^2\)/g. The electrochemical impedance spectrum (EIS) results indicated an enhanced charge separation and transfer ability. Compared to bulk g-C\(_3\)N\(_4\), the electron transfer resistance of g-C\(_3\)N\(_4\) nanosheets decreased by 75%. The photocatalytic hydrogen production activity of g-C\(_3\)N\(_4\) nanosheets is 10 times that of the bulk counterpart and is also higher than that of mpg-C\(_3\)N\(_4\). In another work, She et al prepared g-C\(_3\)N\(_4\) nanosheets via sonication in 1,3-butadiol (1,3-BUT).\(^6\) Due to the polarity and surface energy matching, 1,3-BUT can swell and exfoliate g-C\(_3\)N\(_4\) into nanosheets with a thickness of 0.9-2.1 nm. Mixed solvents have also been used for exfoliation of g-C\(_3\)N\(_4\) such as the mixture of ethanol and water.\(^6\) Interestingly, it was found that the ratio of the two solvents affect the yield and the highest concentration of nanosheets at 3 mg/mL was obtained with an ethanol/water volume ratio of 1:3. This concentration is much higher than that obtained in the single solvent. Furthermore, it is usual that g-C\(_3\)N\(_4\) monolayer was obtained using such mixed solvents. Monolayer g-C\(_3\)N\(_4\) nanosheets can also be produced via a combination of heat treatment and liquid exfoliation and the resultant nanosheets exhibited efficient photocatalytic disinfection activity for \text{Escherichia coli}.\(^6\) Fig. 5 shows the EIS curves and transient photoresponse of the bulk and single-layered g-C\(_3\)N\(_4\), indicating the enhanced charge transport property of g-C\(_3\)N\(_4\) nanosheets. Although facile and simple, the disadvantage of liquid exfoliation method is obvious. It usually takes long sonication duration and the yield is generally lower than 15%. In our recent work, the exfoliation of mpg-C\(_3\)N\(_4\) in ethanol was carried out using a probe sonicator.\(^6\) The probe sonicator with a higher power intensity is directly immersed into the suspension and hence is much more effective than a bath sonicator. It is widely used in the exfoliation of layered metal chalcogenide like MoS\(_2\).\(^6\) The result yield (25.8%) for exfoliated mpg-C\(_3\)N\(_4\) nanosheets was much higher than the previous reported values. Shown in Fig. 6A, the thickness of such mpg-C\(_3\)N\(_4\) nanosheets was 2-3 nm in average, corresponding to 5-8 carbon nitride layers. After depositing Pt or Co\(_3\)O\(_4\) on the surface of mpg-C\(_3\)N\(_4\) nanosheets as cocatalysts (Fig. 6B), the photocatalytic activities for hydrogen evolution or degradation of Rhodamine B, respectively, were greatly enhanced compared to those of the bulk g-C\(_3\)N\(_4\) (Fig. 6C and D). It is remarkable that the hydrogen evolution rate is 26 times that of the bulk g-C\(_3\)N\(_4\).

Inspired by the Hummers method for the exfoliation of graphite into graphene, Zhu et al developed a chemical etching method for the preparation of single atomic layered g-C\(_3\)N\(_4\) nanosheets.\(^6\) It was found that only nanoparticles were obtained in the presence of KMnO\(_4\) given g-C\(_3\)N\(_4\) is not stable enough against the oxidation by KMnO\(_4\).\(^5\) \( \text{H}_2\text{SO}_4 \) (98%) was firstly intercalated into the interlayer
space of g-C3N4. Then g-C3N4 nanosheets were obtained by the rapid heating effect of H2SO4 (98%) mixed with water. The yield of g-C3N4 nanosheets reached as high as 30%. AFM observation revealed that the thickness of 60% of the exfoliated nanosheets was around 0.4 nm which is close to the theoretical thickness of a single carbon nitride layer. Correspondingly, the surface area was significantly increased from 4.3 m2/g to 205.8 m2/g after exfoliation. After loading Pt on the surface of as-prepared g-C3N4 nanosheets, it exhibited an enhancement factor of 2.6 for hydrogen evolution compared to bulk g-C3N4. A subsequent research work demonstrated that the surface charge of the nanosheets prepared using this chemical etching method can be tuned by adding different charge guest. The electrostatic assembly of g-C3N4 nanosheets with CdS or BiOBr was prepared and investigated for the photocatalytic degradation of methyl orange (MO) and aminobenzoic acid. The CdS/g-C3N4 nanosheets composite exhibited a superior photocatalytic activity. However, before the nanosheets can be used for photocatalytic reaction, the suspension should be carefully and thoroughly washed to fully remove the acid.

Besides the abovementioned research efforts, the exfoliated g-C3N4 nanosheets have also been coupled with other materials such as carbon nanotube, reduced graphene oxide and WO3 arrays via electrostatic interaction for potential application in electrocatalysis or PEC reactions. In summary, g-C3N4 can be effectively exfoliated into monolayer or few layered nanosheets in polar solvents including water, some organic solvents and acid solution or by a simple thermal exfoliation. The exfoliated nanosheets present enhanced photocatalytic or PEC activities due to larger specific surface area, enhanced charge separation and transfer ability and more negative CB. The exfoliation process is simple but time consuming and the drawback of g-C3N4 is still need to be developed to mitigate this problem.

4. g-C3N4 based photocatalysts for oxygen evolution and overall water splitting

Water oxidation, the other half reaction in overall water splitting, is a four electron transfer process and considered to be the rate limiting step in overall water splitting. Most of the research efforts on g-C3N4 based photocatalysts have been focused on the water reduction half reaction, which is proceeded in the presence of a hole scavenger. There are only limited studies on water oxidation by g-C3N4 system, not to mention overall water splitting. Despite the energetic barrier and sluggish kinetics, water oxidation by g-C3N4 also suffers from the self-oxidation of the catalyst which leads to the evolution of N2 similar to the case of (oxy)nitrides e.g. Ta3N5, LaTiO2N. Efficient cocatalyst and protection layer need to be developed to mitigate this problem.

When g-C3N4 was first reported to be active for water splitting, its oxygen evolution activity was studied using AgNO3 as the electron acceptor. With RuO2 loaded as a water oxidation catalyst (WOC), g-C3N4 was shown to be able to oxidize water into O2 with a low reaction rate. Maeda et al showed that the photocorrosion can be significantly inhibited by loading RuO2 as an efficient WOC. Otherwise the evolution of N2 is significant. When the loading amount of RuO2 was optimized to 3 wt%, the highest oxygen evolution rate was obtained (12 µmol/h). McMillan et al studied the effect of the precursor and reaction parameter during g-C3N4 preparation on its oxygen evolution activity with RuO2 as the cocatalyst.

Besides the aforementioned drawbacks of g-C3N4, one of the main factors limiting its water oxidation activity is its VB edge position at about 1.4 V vs. NHE. Compared to the VB position of widely studied oxides such as WO3 (2.7 V vs. NHE), BiVO4 (2.8 V vs. NHE) and Fe2O3 (2.2 V vs. NHE), g-C3N4 can only provide moderate water oxidation ability. Wang and co-workers used trithiocyanuric acid as the precursor to prepare sulfur-mediated g-C3N4 (CNS). After condensation, the sulfur amount in the final product was less than 1 wt%. The release of sulfur species during synthesis altered the connectivity pattern, topology of g-C3N4 and lowered the VB position by ca. 0.2 V. It showed a 4-fold increase in O2 production compared to the pristine g-C3N4. However, the oxygen production rate of CNS was only 2.5 µmol/h under visible light irradiation since the photocatalytic experiment was conducted without loading a suitable WOC.

Then, the same group loaded cobalt species as WOC on CNS by an impregnation method and constructed an efficient noble-metal free photocatalyst system for water oxidation. The structure of the cocatalyst was confirmed to be mainly Co3O4 by XPS and HRTEM. Fluorescence quenching study indicated efficient charge transfer from CNS to Co3O4. The PEC study showed that Co3O4/CNS lowered the onset potential of anodic photocurrent by 120 mV and also enhanced the photocurrent response. Using AgNO3 as electron acceptor, the optimized Co3O4/CNS exhibited an oxygen evolution rate of 25.1 µmol/h under visible light. However, the oxygen evolution rate gradually decreased after 5 h which could be due to the deposition of Ag nanoparticles which blocked the catalytic centers as well as light absorption. Recently, Zhang et al reported the deposition of another cobalt based cocatalyst, Co(OH)3, on the surface of g-C3N4 via simple precipitation (Fig. 7A). The PL quenching study in Fig. 7B revealed the efficient charge transfer from g-C3N4 to Co(OH)3. During the first hour,
the oxygen evolution rate obtained was 27.4 μmol/h under UV-vis light irradiation (Fig. 7C). After calcination in air, Co(OH)₂/g-C₃N₄ can be converted to CoOOH/g-C₃N₄, which is less active. Furthermore, the electrocatalytic oxygen evolution reaction (OER) onset potential of Co(OH)₂/g-C₃N₄ was found lower than that of Co₃O₄/g-C₃N₄ by the linear sweep voltammetry shown in Fig. 7D.

In the area of overall water splitting by g-C₃N₄-based system, only limited literature has been found reporting the successfull construction of g-C₃N₄ photocatalysts that can evolve H₂ and O₂ simultaneously at a stoichiometric ratio. Lee et al. decorated cobalt hydroxide (CoPi) catalyst on the surface of mpg-C₃N₄ via direct photodeposition of Co²⁺ ions in a phosphate buffer solution or first deposition of metallic cobalt nanoparticles and then conversion of metallic nanoparticles to CoPi on mpg-C₃N₄. The as-prepared CoPi/mpg-C₃N₄ is active for both HER and OER half reactions in the presence of suitable hole and electron scavenger, respectively. CoPi, a widely studied OER catalyst, was found to be converted in situ to Co-oxo/hydroxo-phosphate which is an active photocatalyst for HER when a hole scavenger exists. In the absence of any charge scavenger, CoPi/mpg-C₃N₄ was found to be able to split water into H₂ and O₂ at the stoichiometric ratio in phosphate buffer solution with H₂ and O₂ produced at 13.6 μmol/g/h and 6.6 μmol/g/h, respectively. The activity for overall water splitting is stable after three runs. Though the reaction rate is still very low, this is the first example of overall water splitting by g-C₃N₄ based photocatalyst.

Besides the single particulate system described earlier, Z-scheme photocatalyst which contains two photon systems is another way to achieve overall water splitting. Inspired by natural photosynthesis, Z-scheme system is composed of a H₂-evolving photocatalyst, an O₂-evolving photocatalyst and an electron mediator. Photocatalysts that are only active for half reactions can be employed to construct Z-scheme system, which extends the choice of the photocatalysts for overall water splitting. However, the activity of a Z-scheme system depends highly on a suitable combination of its components. So far the most active Z-scheme system employs Ru/SrTiO₃/Rh and BiVO₄ as H₂-evolving photocatalyst and O₂-evolving photocatalyst, respectively. Other photocatalysts such as metal oxinitriles or metal sulfides are less active or inactive for Z-scheme systems. Tang and his co-workers for the first time constructed a Z-scheme system employing g-C₃N₄ as H₂-evolving photocatalyst. The mechanism is illustrated in Fig. 8A. Both BiVO₄ and WO₃ were used as the O₂-evolving photocatalyst and soluble Fe³⁺/Fe²⁺ and I³⁻/I⁻ pairs were used as electron mediator. With such a scheme, stoichiometric H₂ and O₂ were produced. As shown in Fig. 8, g-C₃N₄/WO₃ composite was the most efficient in this Z-scheme configuration with production rates of H₂ and O₂ at 74 μmol/g/h and 37 μmol/g/h, respectively, under full arc irradiation. The stability of this system is also confirmed to be good as stoichiometric H₂ and O₂ are evolved after three recycling runs and the gas production rate is almost the same as the initial run (Fig. 8D). Very recently, an effort was taken on the construction of mediator free Z-scheme overall water splitting system using g-C₃N₄ and WO₃ as the HER and OER photocatalysts, respectively. The g-C₃N₄/WO₃ composite was synthesized by in situ growth of WO₃ on the surface of g-C₃N₄ in a hydrothermal reaction. In addition, they used reduced graphene oxide (rGO) as electron mediator which is expected to enhance the performance. After loading Pt as the cocatalyst, the photocatalytic activity of g-C₃N₄/WO₃ and g-C₃N₄/rGO-WO₃ were measured in pure water without redox couples such as I⁻/I⁺. Stoichiometric H₂ and O₂ can be stably evolved under visible light irradiation. A QE of 0.9% was obtained with the optimized sample at 420 nm. These results indicate that the intimate contact between two photocatalysts and improved electron transfer ability may benefit the H₂/O₂ evolution in the two step photoexcitation Z-scheme system. Further improvement of the activity of these systems is still needed.

Recently, Kang et al. constructed a metal free g-C₃N₄ based composite photocatalyst by coupling g-C₃N₄ with carbon dots (CDs). The Solar to hydrogen (STH) efficiency reached 2% for the optimized CDs/g-C₃N₄ system, which is much higher than the thus reported values for photocatalytic systems except that of CoO (5.1%) which is self-corroded within 1 h. Such a STH efficiency reaches not far from the value of 5% set by the U.S. Department of Energy.
which corresponds to the H₂ production cost of $2.3/kg. Furthermore
CDs/g-C₃N₄ exhibited long term stability for 200 days with the
catalyst being separated, dried and reused after each day. The high
stability can be attributed to its chemical and structure stability.
A two-step OER mechanism including H₂O₂ production and
subsequent decomposition of H₂O₂ to O₂ was found to be key to the
high STH efficiency. This two-step pathway was verified by
electrochemical study and shown to be faster than conventional four-
electron process. Thus the STH efficiency was increased by
accelerating the rate-limiting step of OER.

In summary, non-noble metal containing g-C₃N₄ based
system is capable of efficiently evolving O₂ and is promising in
overall water splitting. The metal free feature of g-C₃N₄ makes
it a promising platform for the construction of low cost
photocatalytic system to further lower the cost of solar fuel
production.

5. CO₂ reduction using g-C₃N₄ based photocatalysts
CO₂ reduction is a rapidly developing research area as this
technology provides possible solutions to the environmental and
energy problems we are facing. Many studies and reviews have
been published showing the promising findings in reduction of CO₂ into
value added products.1, 3, 8, 85-93 Similar to the situation in water
oxidation and overall water splitting, there have not been extensive
studies on CO₂ photoreduction by g-C₃N₄ based system.
Nevertheless, during the last few years, CO₂ photoreduction using g-
C₃N₄ has received increasing attention due to the interesting
properties of g-C₃N₄ that may offer some opportunities in this area.

In 2012, Dong and co-workers prepared g-C₃N₄ and porous g-
C₃N₄ by heating melamine or melamine hydrochloride and
investigated their photocatalytic CO₂ reduction activity in the
presence of water vapor under visible light.92 Under these reaction
conditions, CO₂ was obtained as the reduction product. Subsequently,
Mao et al reported the use of g-C₃N₄ with different microstructures for CO₂
photoreduction in NaOH solution without depositing any
co-catalysts.93 In this work, urea and melamine were used for the
preparation of g-C₃N₄. It was found that the specific surface area of
u-g-C₃N₄ (39.5 m²/g) is much higher than that of m-g-C₃N₄ (3.7
m²/g), which leads to more efficient surface adsorption, better charge
separation and improved photoactivity. The CO₂ reduction products
over u-g-C₃N₄ were CH₂OH and CH₂H₂OH while only CH₂OH was
produced over m-g-C₃N₄. It is interesting that hydrocarbons were
produced by CO₂ reduction over g-C₃N₄ system instead of gaseous
CO₂ possibility because the reaction was conducted in aqueous
solution. However, CH₂OH or C₂H₄OH may be oxidized on the
surface of g-C₃N₄ by in situ evolved O₂ during long photocatalysis
indicated by the decreased reaction rate after irradiation beyond 9 h.
Tuning the electronic structure of g-C₃N₄ by doping S into the
pristine g-C₃N₄ can enhance its optical adsorption as well as the CO₂
reduction activity. S doped g-C₃N₄ was prepared by the condensation of thiourea and the CH₂OH yield over S doped g-C₃N₄ is 1.4 times
that over the pristine g-C₃N₄.94 To construct more efficient systems, one
strategy is to couple the highly active homogeneous catalyst
with g-C₃N₄. For instance, a Ru complex, cis,trans-[Ru{4,4'-
(CH₃PO)₂H₂-2,2'-bipyridine}(CO)₂Cl₂] (Ru) was adsorbed on the
surface of mpg-C₃N₄ with a high surface area of 180 m²/g.95 The
Ru/mpg-C₃N₄ was able to reduce CO₂ to formic acid under visible
light while a small amount of H₂ and CO were also detected in
acetone in the presence of TEOA as the sacrificial reagent.
Isotopic measurement result indicated that formic acid entirely came
from CO₂ reduction while 77% of the evolved CO originated from
the carbonyl ligand unit of Ru catalyst. The detachment of carbonyl
ligand from Ru was a slow process and has also been reported in
other papers including homogeneous and heterogeneous systems.96,97
However, the photocatalytic activity of Ru/mpg-C₃N₄ for formic
acid production was found to stay almost unchanged even after 5 h
of reaction. The Ru catalyst after carbonyl ligand detachment was
expected to be the active species for CO₂ reduction. Subsequently,
the same group studied the effect of pore-wall structure of mpg-C₃N₄
and the effect of Ru complex structure on CO₂ photo reduction.98,99
Based on these results, it is understood that the photocatalytic activity
is sensitive to the specific surface area of mpg-C₃N₄ and is not related
to the pore size and volume. Furthermore, introduction of too much
meso-porosity results in the shrinkage of mpg-C₃N₄ walls and leads
to activity drop. With I₃PO₄/H₂O used as the linker group, Ru/mpg-
C₃N₄ efficiently reduce CO₂ to HCOOH under visible light in N,N-
dimethylacetamide with TEOA as the sacrificial reagent. This hybrid
material gave a high turnover number (TON) of larger than 1000 and
QE of 5.7% at 400 nm. In another study, Lin et al prepared
Co(bpy)₃Cl₂/g-C₃N₄ hybrid material by self-assembly as the
photocatalyst for reduction of CO₂ in acetonitrile under visible light
in the presence of TEOA.100 CO and H₂ were the main products. A
TON of 4.33 with a relatively high selectivity of 88.4% for CO
production were obtained by the optimized hybrid system.
The surface of g-C₃N₄ or mpg-C₃N₄ was also modified with cobalt
species as oxidative promoters to enhance CO₂ photoreduction.
Co(bpy)₃Cl₂/CoO/mpg-C₃N₄ resulted in the highest TON of 13 and
the selectivity of CO to H₂ was 78.5%. Noble metal cocatalysts loaded
g-C₃N₄ have also been used for CO₂ photoreduction. Under UV-vis
light CO₂ can be reduced to hydrocarbons (mainly CH₄, CH₂OH and
HCHO) using Pt/g-C₃N₄ photocatalyst.101 Pt acts as an electron sink
to enrich the surface of g-C₃N₄ with electrons for efficient CO₂
reduction. The maximum yield can be obtained when the loading
amount of Pt was 0.75%. However, Pt can also act as a catalyst for
oxidation of HCHO over the time. Pt/g-C₃N₄ was also prepared via a
polyol process and used for photoreduction of CO₂ in the presence
of water vapor under daylight lamp irradiation.102 CH₄ was the main
product from CO₂ reduction and 5.1 fold enhancement of CH₄
production was obtained after 2% Pt was loaded on g-C₃N₄.

The composites of g-C₃N₄ and metal oxides have been
investigated by various research groups for CO₂
photoreduction. As in the case of hydrogen production,
coupling g-C₃N₄ with a suitable semiconductor enhances the
charge separation via band alignment, which leads to increased
activity. For example, the composite of NaNbO₃ nanowires and
g-C₃N₄ was prepared by annealing the mixture of NaNbO₃
nanowires and melamine at 520 °C in air.103 In this case, g-
C₃N₄ was expected to grow on the surface of NaNbO₃ and thus
to make a good interface, which is revealed by the SEM and
tem images in Fig. 9B and C. The band energy diagram of
NaNbO₃ and g-C₃N₄ is shown in the schematics in Fig. 9A. The
suitable band alignment
between NaNbO$_3$ and g-C$_3$N$_4$ facilitates the charge separation in the composite. After photodeposition of 0.5% Pt, the composite was capable to reduce CO$_2$ to CH$_4$ and the activity was much higher than those of the individual components loaded with Pt. Cao et al prepared In$_2$O$_3$/g-C$_3$N$_4$ photocatalyst by a solvothermal method in dimethyl sulfoxide. In$_2$O$_3$ nanocrystals were grown on the surface of sheet-like g-C$_3$N$_4$. In$_2$O$_3$/g-C$_3$N$_4$ exhibited a similar optical adsorption property as the pristine g-C$_3$N$_4$ but the transient photoresponse showed an increased photocurrent for In$_2$O$_3$/g-C$_3$N$_4$. After loading Pt as electron sink over 10% In$_2$O$_3$/g-C$_3$N$_4$ 159.2 ppm CH$_4$ can be evolved during 4 h. Transient PL decay also implied the inhibited charge recombination in In$_2$O$_3$/g-C$_3$N$_4$ due to enhanced charge separation at the interface. In another study, g-C$_3$N$_4$/TiO$_2$ heterojunction was prepared by an in situ growth method. The surface area of the composite increased with the percentage of TiO$_2$ in the composite. When the photoreduction of CO$_2$ was carried out with water vapor without cocatalyst under UV-vis irradiation, CO was found as the main product although a small amount of CH$_4$ was also produced. Bi$_2$WO$_6$ was previously prepared by a hydrothermal method and shown active for the reduction of CO$_2$ into CO under visible light. A solvothermal process was used to grow Bi$_2$WO$_6$ in situ to form g-C$_3$N$_4$/Bi$_2$WO$_6$ composite. The measured CB and VB positions of g-C$_3$N$_4$ and Bi$_2$WO$_6$ were used to explain the possible mechanism for the photoreduction of CO$_2$ to CO. Compared to the pure Bi$_2$WO$_6$ prepared using the hydrothermal or solvothermal method, the activity of the composite was largely enhanced. ZnO, a large band gap semiconductor, was also used to make a composite with g-C$_3$N$_4$ by the impregnation method for the CO$_2$ reduction. The charge separation and transportation were promoted by the suitable band alignment between g-C$_3$N$_4$ and ZnO, which leads to an enhanced activity. Besides oxides, carbon materials have also been coupled with g-C$_3$N$_4$ for CO$_2$ photoreduction. For example, a sandwich-like graphene/g-C$_3$N$_4$ hybrid nanostructure was fabricated using graphene oxide as a structure directing agent. The hybrid material shows enhanced activity for the conversion of CO$_2$ to CH$_4$ in the presence of water vapor under a daylight lamp. The enhanced photoactivity was attributed to the improved electron transfer induced by graphene.

Besides the electronic and catalytic aspects, modifying g-C$_3$N$_4$ with a component for CO$_2$ adsorption and enrichment has been found effective in enhancing the performance of the overall system. For example, Wang et al coupled g-C$_3$N$_4$ with a Co-containing zeolitic imidazole framework (Co-ZIF). Co-ZIF-9 has a high CO$_2$ adsorption capacity of 2.7 mmol/g and affords a high microporous surface area of 1607 m$^2$/g. As a result, Co-ZIF-9 can capture and concentrate CO$_2$ in its pores. After an electron mediator, bipyridine, was added into the reaction solution, the photoexcited electrons can be transferred from g-C$_3$N$_4$ to Co-ZIF-9 for CO$_2$ reduction as revealed by PL quenching study. CO was the main product in this system and a QE of 0.9% can be obtained, even without the loading of a cocatalyst. Our group reported the construction of a hybrid system using g-C$_3$N$_4$ and an anionic clay called layered double hydroxide (LDH) which has high affinity for CO$_2$. The structure of LDH/g-C$_3$N$_4$ composite is shown in Fig. 10A. Mg-Al-LDH with a positive surface charge was assembled with oppositely charged g-C$_3$N$_4$ by electrostatic interaction. In the TEM image shown in Fig. 10B, the hexagonal-shaped LDH nanosheets can be seen on the surface of g-C$_3$N$_4$. Owing to its unique layered structure, CO$_2$ molecules can be intercalated in the interlayer space of LDH in the form of CO$_3^2$ anions. After loading Pd as the cocatalyst, CH$_4$ was the main product from CO$_2$ reduction using this photocatalyst under UV-vis light. Isotopic measurement and blank control experiment indicated that most of the evolved CH$_4$ was from CO$_2$ instead of other carbon containing species. The photocatalytic activity of the assembly system is 4 times that of the control sample without LDH. Among the LDHs intercalated with dominantly NO$_3^-$, dominantly CO$_3^2$ and the mixture of the two, LDH-CO$_3^2$ shows superior performance (Fig. 10C), suggesting that the intercalated CO$_3^2$ anions from dissolved CO$_2$ in water can act as a carbon source and are easily
reduced since LDH is in close contact with both g-C₃N₄ and Pd. As shown in Fig. 10D, LDHs with different metal compositions were compared and Mg-Al-LDH exhibited the best performance for CO₂ reduction which is correlated with the highest CO₂ adsorption capacity of Mg-Al-LDH among all LDHs. The highest QE of 0.093% was obtained at 440 nm over the optimized Pd/LDH/g-C₃N₄ assembly. The QE is still low but this example shows the concept of coupling g-C₃N₄ with a CO₂ capturing material is promising in CO₂ reduction.

In summary, the study of g-C₃N₄ based photocatalysts for CO₂ reduction is still at its early stage, but has received increasing interests in recent years. Different g-C₃N₄ based photocatalysts were synthesized including organic-inorganic hybrids, metal deposited g-C₃N₄, nanocomposites of g-C₃N₄ with oxides or carbon materials, and composites of g-C₃N₄ with CO₂ adsorbing materials. The photocatalytic activity of g-C₃N₄ based photocatalysts depends on many factors and continuous efforts are needed for the development of more efficient and stable photocatalysts for CO₂ reduction. It is also worth noting that the reported experimental setups, reaction conditions and product analysis methods vary from one paper to another. This makes it difficult to directly compare the reported activity data from different research groups. While it is challenging to adopt the same experimental setup by different groups, it is important to use reliable methods to accurately quantify various products of CO₂ reduction. Several review articles including ours published in the past few years on this topic can provide useful practical guidelines to researchers in this field. 112-114

6. Conclusion and Outlook

This mini review summarizes the latest research efforts on the development of g-C₃N₄ based photocatalyst systems with the emphasis on the development of non-noble metal cocatalysts for g-C₃N₄, exfoliation of g-C₃N₄ to nanosheets with enhanced photoactivity, and application of g-C₃N₄ based photocatalysts for water oxidation, overall water splitting and CO₂ reduction. The study of non-noble metal cocatalyst loaded g-C₃N₄ has already achieved some breakthroughs. The activities of a few photocatalysts like MoS₂/g-C₃N₄ and Ni(OH)₂/g-C₃N₄ were reported to approach or surpass those of Pt/g-C₃N₄. Further improvement of efficiency and stability is still in need. Exfoliation is a promising method to improve the photocatalytic activity of g-C₃N₄. Facile and simple liquid exfoliation can be conducted in water, certain organic solvents or acid solutions. Single layered or few layered g-C₃N₄ nanosheets can be obtained after exfoliation. The larger surface area, improved charge transfer ability and enhanced charge separation contribute to the superior photocatalytic activities of g-C₃N₄ nanosheets. The performance can be further improved by preparing porous g-C₃N₄ nanosheets. However, the current exfoliation method suffers from long time sonication and low yield. It is necessary to develop more efficient exfoliation methods. Water oxidation can be achieved by g-C₃N₄ based photocatalysts such as cobalt oxides loaded g-C₃N₄, although there are only limited research works on water oxidation over g-C₃N₄. The performance and stability of current g-C₃N₄ water oxidation photocatalysts are quite poor. Meanwhile, the study of overall water splitting by g-C₃N₄ based photocatalysts is in its early stage although some systems including single particulate system and Z scheme have been shown capable of evolving stoichiometric H₂ and O₂ from water. Further development of efficient g-C₃N₄ based photocatalysts for water oxidation is thus needed for its application in overall water splitting. There have been increasing research interests on CO₂ photoreduction using g-C₃N₄ owing to its suitable electronic band structure. Drawbacks including moderate optical adsorption property, low charge carrier mobility, inert surface, low surface area and rich grain boundary defects need to be overcome to improve its photocatalytic activity. Some innovative ways and more systematic work are probably required to develop g-C₃N₄ to efficient and stable photocatalysts, which is a challenging task.

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