Role of Ethynyl-Derived Weak Hydrogen-Bond Interactions in the Supramolecular Structures of 1D, 2D, and 3D Coordination Polymers Containing 5-Ethynyl-1,3-benzenedicarboxylate

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Supporting Information

ABSTRACT: The influence of weak hydrogen bonds on the crystal packing of a series of heavy and transition metal coordination polymers synthesized using the ligand 5-ethyl-1,3-benzenedicarboxylic acid \((\text{H}_2\text{ebdc})\) has been evaluated. Five coordination polymers were prepared and crystallographically characterized. These comprise two 1D chains, \([\text{Pb}(\text{ebdc})(\text{DMSO})_2]\) (1) and \([\text{Pb}(\text{ebdc})(\text{DMF})]\) (2), two 2D nets, \([\text{Cu}_3(\text{ebdc})(\text{H}_2\text{O})_{1.5}(\text{MeOH})_{0.5}]\cdot\text{H}_2\text{O}\) (3) and \([\text{Pb}_2(\text{ebdc})(\text{DMF})_2]\cdot\text{H}_2\text{O}\) (4), and a single 3D framework, \([\text{HNEt}_3]\cdot\text{Zn}_3(\mu_1-\text{OH})(\mu_2-\text{H}_2\text{O})(\text{ebdc})_3(\text{MeOH})_{0.67}(\text{H}_2\text{O})_{0.33}\cdot\text{MeOH}\cdot\text{H}_2\text{O}\cdot\text{DMF}\) (5). The crystal structure of the free acid ligand form, \(\text{H}_2\text{ebdc}\cdot\text{H}_2\text{O}\), is also reported. Within the lead(II) coordination structures, ethynyl-derived \(\text{C}^\equiv\text{H}--\text{O}\) interactions are consistently found to provide the dominant influence over the crystal packing, as determined by solid-state structural analysis in combination with vibrational spectroscopy. The influence of weak hydrogen-bonding effects on the crystal packing of the transition metal coordination polymers that contain lattice water and methanol molecules was found to be far less prominent, which is interpreted in terms of the greater prevalence of strong hydrogen-bond donors and acceptors forming \(\text{O}--\text{H}--\text{O}\) interactions within these crystalline lattices.

INTRODUCTION

The design of crystalline architectures employing an understanding of the intermolecular interactions available to molecular subunits is the basis for crystal engineering, a key field straddling solid-state molecular assembly and structural analysis, whose importance lies in its capability for allowing control of crystal properties and functions as well as its structure. Much progress has been made using simple systems composed of covalent or strong hydrogen bonds; however, control over weak hydrogen-bonding interactions remains one of the most challenging areas of crystal engineering. A weak hydrogen bond can broadly be defined as an electrostatic interaction formed by a hydrogen atom between two structural moieties of moderate to low electronegativity, of which \(\text{C}^\equiv\text{H}--\text{X} (\text{X} = \text{O/N})\) interactions are a key example. The study of these interactions began with the discovery of increased polarization of haloforns with ketones, pyridines, and ethers. This observation was linked by discovery of increased polarization of haloforms with ketones, example. The study of these interactions began with the spectra of such species. While assignment of these interactions for large bathochromic shifts observed in the infrared (IR) spectra of such species. Recent advancements in the field of crystallography and in computational power have demonstrated and confirmed the fundamental importance of these interactions to supramolecular self-assembly. Such interactions can vary in strength from near equivalence to van der Waals interactions to being stronger than weak covalent bonds and rivalling traditional hydrogen bonds. In many crystalline networks, weak hydrogen bonds can be thought of as having a steering effect that may preferentially favor a single solid form, for example, a particular polymorph with solid-state packing governed largely by stronger intermolecular interactions but which is supplemented by contributions from weak directional \(\text{C}^\equiv\text{H}--\text{X}\) interactions. However, there are known instances wherein weak hydrogen bonds are formed preferentially over interactions involving available strong donors or acceptors. This preference for weak hydrogen bonds ensures that structural predetermination will fail if predictive methods rely solely on the traditional hierarchy of bonding strengths. Thus, the crystal engineer must view the sum of all interactions, both weak and strong, in order to predict the intermolecular architecture and crystal packing of a material, as opposed to focusing only on the strongest few interactions. This requirement necessitates an in-depth study of weak hydrogen-bonding interactions in order to understand fully, and integrate better, these stabilizing forces into the wider tenets of crystal engineering.

The ethynyl group is an attractive functionality for the study of weak hydrogen bonding owing to its highly activated nature
and acidity.\textsuperscript{15} It is also ideally suited to spectroscopic study in the solid state, owing to its prominent $C(sp)$−$H$ vibrational band in terms of intensity and unique location. Indeed examples of the $C(sp)$−$H$····$O$ hydrogen bond match moderately strong $O$−$H$····$O$ hydrogen bonds geometrically and, in terms of trends, spectroscopically.\textsuperscript{16} This finding has been supported by studies comparing the hydrogen-bonding potential energy curves of weak acidic hydrogen bonds to $O$−$H$····$O$ bonds that showed matching distributions.\textsuperscript{5,17} The ethynyl group can act cooperatively as a hydrogen-bond acceptor via the alkyne $\pi$ system, as well as containing a donor site from the $C(sp)$−$H$ hydrogen, enhancing hydrogen-bond strength by mutual polarization.\textsuperscript{5} Similar cooperatively is one factor that provides strength to traditional $O$−$H$····$O$−$H$····$O$−$H$ hydrogen-bonded chains.

Evaluation of the weak $C$−$H$····$O$ hydrogen bonding observed for coordination polymers in this study has been achieved by single-crystal diffraction studies in order to obtain accurate interatomic donor−acceptor distances ($D$) and angular properties ($\theta$) of the hydrogen donor approaching the acceptor atom, in conjunction with analysis of $C(sp)$−$H$ hydrogen bonds by vibrational spectroscopy and changes to thermal vibrations achieved by single-crystal di

### Table 1. Crystallographic Data for $H_2ebdc$ and Coordination Polymers 1−5$^a$

| Formula          | $H_2ebdc\cdot H_2O$ | 1 | 2 | 3 | 4 | 5 |
|------------------|---------------------|---|---|---|---|---|
| **Formula Weight** | $C_{14}H_{16}O_6PbS_2$ | $C_{20}H_{18}N_8O_8Pb_4$ | $C_{18}H_{20}C_6O_{31}$ | $C_{36}H_{32}N_2O_3Pb_3$ | $C_{75.34}H_{71.02}N_2O_{34.66}Zn_6$ |
| **Space Group**   | $P\bar{1}/n$ | $P\bar{1}/a$ | $P\bar{1}/a$ | $P\bar{1}/a$ | $P\bar{1}/c$ |
| **$a$, Å**        | 9.5290(6) | 17.8249(6) | 15.0418(5) | 18.2460(3) | 11.6687(1) |
| **$b$, Å**        | 13.7468(10) | 10.5770(4) | 15.0480(6) | 21.5618(4) | 14.9906(6) |
| **$c$, Å**        | 10.7970(6) | 21.5618(4) | 25.1110(4) | 24.3281(8) | 25.1110(4) |
| **$\alpha$, deg** | 90 | 90 | 90 | 90 | 90 |
| **$\beta$, deg**  | 90 | 71.172(3) | 90 | 90 | 90 |
| **$\gamma$, deg** | 90 | 90 | 90 | 90 | 90 |

$^a$Selected bond lengths and angles for the coordination polymers 1−5 are available as Supporting Information.

### EXPERIMENTAL METHODS

**Caution:** Metal perchlorates are potentially explosive! Only a small amount of material should be prepared and handled with great care. Starting materials and solvents were purchased from commercial sources and used without further purification, with the exception of $H_2ebdc$, for which single crystals were grown by slow diffusion of water into a concentrated methanol solution. X-ray diffraction data for $H_2ebdc$ and compounds 1−5 were collected on an Agilent Gemini A-Ultra diffractometer\textsuperscript{24} at the University of Bath using Mo Kα radiation, with the crystal being cooled to 150 K by an Agilent Cryojet.\textsuperscript{25} Powder X-ray diffraction patterns (PXRDs) were recorded on a Bruker AXS D8 Advance diffractometer with Cu Kα radiation of wavelength 1.5406 Å at 298 K. Samples were placed on a flat plate and measured with a 2θ range of 3−60°. Simulated X-ray powder patterns were generated from single-crystal data that were imported into PowderCell. Infrared spectra were recorded on a PerkinElmer Spectrum 100 spectrometer equipped with an ATR sampling accessory. Abbreviations for IR bands are s, strong; m, medium; w, weak. Elemental analyses (C, H, N) were performed on a CE-440 elemental analyzer (Exeter Analytical).

**Synthesis of $[Pb(ebdc)(DMSO)]_2$ (1).** Colorless prismatic crystals of 1 were obtained by solvothermal synthesis using $Pb(OAc)_2\cdot 3H_2O$ (0.08 g, 0.21 mmol) and 5-ethyl-1,3-benzenedicarboxylic acid ($H_2ebdc$) (0.04 g, 0.21 mmol) dissolved in DMSO (5 mL) in a sealed vial and placed in a 100 °C oven for 4 days. Bulk purity was determined by PXRD. Anal. Calcd (%) for $C_{14}H_{16}O_6PbS_2$: C, 30.48; H, 2.92. Found: C, 30.36; H, 2.83. FTIR: $\nu$ 3221 (m), 2931 (w), 1644 (s), 1593 (m), 1517 (s), 1424 (s), 1352 (s), 1101 (m), 953 (s), 790 (m), 679 (m), 621 (w), 531 (m), 408 (s), 331 (w), 293 (w), 243 (w), 167 (w), 156 (w), 141 (w), 135 (s), 1122 (w), 1099 (m), 1016 (s), 993 (s), 953 (s), 790 (m), 777 (s), 743 (w), 722 (s) cm$^{-1}$.

**Synthesis of $[Pb(ebdc)(DMF)]_2$ (2).** Colorless prismatic crystals of lead chain 2 were obtained by solvothermal synthesis of Pb(OAc)$_2\cdot 3H_2O$ (0.08 g, 0.21 mmol) and $H_2ebdc$ (0.04 g, 0.21 mmol) dissolved in DMF (5 mL) in a sealed vial and placed in a 140 °C oven for 2 days. Bulk purity was determined by PXRD. Anal. Calcd (%) for $C_{20}H_{18}N_8O_8Pb_4$: C, 30.48; H, 2.92. Found: C, 33.42; H, 2.83. FTIR: $\nu$ 3247 (m), 2931 (w), 1644 (s), 1593 (m), 1517 (s), 1424 (s), 1352 (m), 1233 (m), 1101 (m), 918 (m), 789 (m), 773 (s), 717 (s), 663 (s) cm$^{-1}$.

**Synthesis of $[Cu_2(ebdc)(H_2O)]_3(MeOH)_2\cdot 3H_2O$ (3).** Copper network 3 was crystallized by layering diffusion of $Cu(ClO_4)_2\cdot 3H_2O$ (0.04 g, 0.11 mmol) in $H_2ebdc$ (1 mL) and $H_2ebdc$ (0.02 g, 0.11 mmol)
in MeOH (1 mL) containing a drop of NEt₃ after several days. The layers were allowed to slowly diffuse over several days through a 1:1 (H₂O/MeOH) buffer layer (2 mL). Turquoise block crystals were obtained. Bulk purity was determined by PXRD. Anal. Calcd (%) for C₅₅H₅₃N₅O₂₄Pb₄: C, 35.58; H, 3.67; N, 1.44. Despite repeated attempts on multiple batches of Crystals of lead network 4 were obtained by solvothermal synthesis using Pb(OAc)₂·3H₂O (0.08 g, 0.21 mmol) and H₂ebdc (0.04 g, 0.21 mmol) dissolved in DMF (5 mL) in a sealed vial and placed in a 100 °C oven for 4 days. Bulk purity was determined by PXRD. Anal. Calcd (%) for C₅₅H₅₃N₅O₂₄Pb₄ (substitution of one DMF molecule for two water molecules, composed of three rings of molecular interactions is ultimately governed by covalent bonding interactions; hence, the nonbonded value for H₂ebdc (C(10)/C(9)) was determined to be 1.37, which is within the range found for known nonbonded terminal ethynyl groups, albeit near the lower limit.¹⁸

Five coordination polymers, comprising two 1D chains, [Pb₂(ebdc)₂(DMSO)₄] (1) and [Pb₂(ebdc)(DMF)] (2), two 2D nets, [Cu₅(ebdc)(H₂O)₃]₂(MeOH)ₐ·6H₂O (3) and [Pb₂(ebdc)₂(DMSO)₄]H₂O (4), and a single 3D framework, [HNEt₃][Zn₃(μ₂-OH)(μ₂-H₂O)(ebdc)₂](MeOH)₁·0.67·(H₂O)₀.₆₇·MeOH·1.₃₃H₂O (5), were synthesized from H₂ebdc with the crystal data obtained in this study summarized in Table 1. Tables of selected bond lengths and angles for each of the coordination polymers 1–5 are available as Supporting Information.

**Synthetic Overview.** The equimolar reaction of H₂ebdc with hydrated divalent metal salts gives rise to coordination polymers consistent with the known affinity of meta-substituted aromatic carboxylic acids for such species.²₇–²⁹. The hierarchy of molecular interactions is ultimately governed by covalent metal–carboxylate interactions involving the deprotonated ebdc ligand, which yield the primary structure. Inclusion of the ligand guarantees the presence of the ethynyl group, ensuring the possibility of weak hydrogen bonds influencing the secondary molecular packing. However, a range of polar protic and aprotic solvents were used during the synthesis of 1−5 (Scheme 1), and no effort was made to exclude water from the synthesis or crystallization processes; hence, competition with strong solvent-mediated hydrogen bonding has been allowed.
and may dominate crystal packing, as observed in the solid-state structure of H₂ebdc·H₂O. For Pb systems, the nature of the product obtained is dependent on the solvent used and the temperature of the reaction.

**Lead Chain 1.** The asymmetric unit of 1 was found to contain a single lead(II) center ligated by an ebdc ligand and two molecules of DMSO, one of which is disordered over two positions. The structure propagates such that the lead(II) is eight-coordinate, with each ebdc ligand coordinating to three lead(II) atoms and with each carboxylate chelating in a \(\mu-(\mu\text{O}:\kappa\text{O}')\) manner (Figure 3). This includes two secondary interactions within the limits of van der Waals interactions for lead and oxygen (2.75 < 3.30 Å), Pb(1)–O(4′) \((1 - x, -y, 2 - z), 3.248(7) \) Å; Pb(1)–O(2′) \((2 - x, -y, 2 - z), 3.079(7) \) Å. The long contacts pair together simple chains of [Pb(ebdc)·(DMSO)₂].

Packing of the 1D chains results from bifurcated donor C−H···O hydrogen bonding derived from the ethynyl group and links to O(2′) and O(4′) of two carboxylate groups (Figure 4). Distances for C(9)−O(2′) of 3.286(8) Å and C(9)−O(4′) of 3.213(9) Å were observed, and the angle of the ethynyl group is 170° relative to the plane of the carboxylate oxygen atoms, favoring the observed hydrogen-bonding interactions. A significant bathochromic shift of 80 cm⁻¹ was observed in the IR spectrum, and the ratio of thermal vibration for C(9) and C(8) decreased to 1.33, suggesting inhibition as a result of hydrogen bonding. These interactions occur on either side of the 1D chain and collectively yield a 2D hydrogen-bonded network that packs in a herringbone arrangement (Figure S1, Supporting Information).

**Lead 2D Net 2.** Lead network 2 was formed by solvothermal synthesis in DMF at 140 °C and yields a 1D tape motif that crystallizes in the triclinic space group P\(\overline{1}\). The asymmetric unit contains three unique lead(II) atoms that are ligated by three ebdc ligands and three coordinated molecules of DMF, two of which are disordered over two positions (Figure S2, Supporting Information). These propagate by symmetry to give repeating hexagonal rings of lead(II) atoms that are fused by bridging ebdc ligands (Figure 5). Pairs of ebdc ligands direct ethynyl groups above and below the tape, and a single ligand bridges the remaining orthogonal faces, ensuring that the periphery of the tape is shrouded by weak hydrogen donors (Figure 6). Bridging modes exhibited by the six carboxylate groups include \(\mu-(\mu\text{O}:\kappa\text{O}'), \mu_3-(\mu\text{O}3\mu\text{O}'), \) and \(\mu_4-(\mu\text{O}3\mu\text{O}'), \) which result in one nine-coordinate and two seven-coordinate lead(II) centers. There are three unique ethynyl groups that have the potential to form hydrogen-bond interactions (Figure 7). Bifurcated ethynyl C−H···O hydrogen...
bonding analogous to that of 1 was observed in two instances in the structure of 2, with paired distances of C(10)--O(5′), 3.279(10) Å; C(10)--O(8′), 3.308(11) Å; and C(20)--O(2′), 3.282(10) Å; C(20)--O(4′), 3.194(10) Å. The former hydrogen-bond interaction aligns with an angle of 164°, the latter, to 179°, with the plane of the carboxylate oxygen atoms. The final ethynyl group is directly oriented toward the oxygen atom of a carboxylate group (C(39)--O(5″)) with an angle of 170°; however, the distance is long at 3.562(17) Å. No other significant interactions form between the 1D chains, suggesting that weak hydrogen-bonding interactions are solely responsible for the crystal packing in 2. Despite the presence of three distinct ethynyl groups within the asymmetric unit of 2, only a single band with a prominent shoulder directed toward lower wavenumbers was observed in the IR spectrum. The dominant vibration exhibited a bathochromic shift of 56 cm⁻¹ relative to the nonbonded reference and has been assigned to the bifurcated pair of terminal alkynes, owing to its intensity relative to the shoulder band that is likely a result of its proximity to an oxygen acceptor. Furthermore, the shoulder exhibits greater red-shifting, which has been observed for directly aligned, as opposed to bifurcated, C–H···O interactions (vide infra). The ratios of C_{terminal}/C_{inner} thermal vibrations for the bifurcated hydrogen-bond interactions are 1.31 and 1.47, which suggests that the former, C(20)--O(4′), holds more hydrogen-bonding character, likely owing to its more favorable orientation angle of 179°. The final ethynyl group has a very high C_{terminal}/C_{inner} thermal parameter ratio of 1.73; however, this may be a result of its proximity to a neighboring disordered DMF site. In such instances, infrared analysis should be considered the most reliable indicator for weak hydrogen-bond identification.

**Copper 2D Net 3.** The 2D copper network 3 crystallizes in the monoclinic I2/a space group and exhibits formation of the characteristic copper paddlewheel motif. The asymmetric unit contains two copper(II) atoms bridged by a carboxylate group from three ebdc ligands, one of which coordinates the third copper(II) atom via the second carboxylate group. The apical positions of the paddlewheels are occupied by coordinated water molecules, one of which exhibits 50:50 substitutional site disorder with a methanol molecule. Five molecules of water are located in the lattice, four of which are disordered over more than one position. The extended molecular structure forms a t₃(6,3) 2D net, whereby the paddlewheels link to assemble a Kagome lattice (Figure 8). Each 2D sheet is offset relative to the sheets above and below. In this manner, the network is similar to that formed by copper(II) with other 1,3-benzenedicarboxylates (bdc), such as 5-nitro-1,3-bdc and 5-(methylsulfonylmethyl)-1,3-bdc.

Of the three unique ebdc ligands, two are oriented toward the carboxylate groups of the layer above, and one is oriented toward a lattice water molecule below (Figure 9). Competing interactions include π–π stacking that directly link the layers and a network of strong hydrogen bonding. The former is a parallel offset interaction with an intercentroid distance of 3.67 Å (closest π–π contact, 3.60 Å), linking this ligand with its symmetry equivalent in a second net. Interaction between the two nets is strengthened by a hydrogen-bonding array between the apical solvent molecules of the paddlewheel and lattice water molecules (Figure 10). While the quality of the crystallographic data was insufficient to locate the individual hydrogen atoms using the electron density map, interoxygen distances give strong evidence for the presence of a strong hydrogen-bonding network, exemplified by distances of 2.687(5) Å for O(15)--O(17) and 2.760(5) Å for O(17)--O(13′) (1/2 − x, 3/2 − y, 1/2 − z).

The closest internetwork ethynyl contacts are 3.217(4) Å for C(20)--O(2) and 3.198(4) Å for C(30)--O(8). The final ethynyl group is directed toward a lattice water molecule with a distance of 3.249(5) Å for C(10)--O(16). Despite the relatively

![Figure 7. Three unique ethynyl groups in 2 with potential for weak hydrogen bonding. Symmetry codes: O(4′) and O(8′) (1 − x, 1 − y, 1 − z), O(2′) and O(5″) (−x, 1 − y, 1 − z), and O(5″) (1 + x, y, z − 1).](image)

![Figure 8. 2D Kagomé lattice arrangement observed for 3.](image)

![Figure 9. Closest ethynyl contacts for the three unique ligands observed in 3. Thermal ellipsoids are shown with 50% probability. The methanol molecule coordinating Cu(2) exhibits substitutional disorder with a coordinated water molecule and O(16).](image)
close proximity of the ethynyl donor groups to the oxygen acceptors, the angles are not conducive of a significant hydrogen-bonding interaction, ranging from 143 to 152°.

The IR spectrum of 3 shows that the band intensity for the ethynyl C−H stretching vibrations centered at 3287 cm⁻¹ has decreased and broadened considerably, with each effect being characteristic of the absence of significant hydrogen bonding.31 Greater thermal vibration was observed between the terminal and inner sp carbons of the alkyne, with C_TERMINAL/C_INNER ratios of 1.66, 1.74, and 1.40, which provides further evidence for a lack of hydrogen-bonding interactions. In this instance, the crystal packing appears to be determined by interactions other than those of weak hydrogen bonding (i.e., strong hydrogen bonding and/or π−π stacking described earlier).

**Lead 2D Net 4.** Crystals of 2D lead network [Pb₂(ebdc)₂(DMF)₄]·H₂O (4) were obtained by heating a mixture of H₂ebdc and hydrated lead acetate in DMF at 100 °C. The network crystallizes in the tetragonal space group P4₂₁c, and the asymmetric unit contains a single lead(II) atom ligated via the carboxylate of a single ebdc ligand, two DMF molecules, of which one is disordered over a special position, and a single lattice water molecule with half occupancy that is also located on a special position (Figure S3, Supporting Information). The lead(II) atoms are eight-coordinate and aggregate to form a Pb₄O₄ cubane motif (Figure 11) in which each cubane node linked to four others by pairs of bridging ebdc ligands to give an extended (4,4) 2D net (Figure 12). The two carboxylate groups of ebdc have different coordination modes, whereby one carboxylate chelates a single lead(II) atom in a κO₂κO' manner, and the second carboxylate forms the corner of the cubane cluster, thereby coordinating three lead atoms in a μ₃-(μ₂O₂κO') fashion. Ethynyl-derived weak hydrogen bonds constitute the only interactions linking the 2D sheets, in which the ethynyl groups of bridging ebdc ligands coordinating to carboxylate groups located above or below the net in an alternating fashion. The C(10)−O(1') hydrogen bond to the carboxylate group of a symmetry generated cluster ($\frac{1}{2} - x, y - \frac{1}{2} - y, \frac{3}{2} - z$) has a distance of 3.152(5) Å, with a favorable angle for bonding of 171° (Figure 13).

Network 4 exhibits the largest bathochromic shift of compounds 1−5, wherein the terminal C−H stretch shifts 89 cm⁻¹ to 3212 cm⁻¹. Similarly, the lowest ratio of thermal vibration between the terminal and adjacent carbon atoms of

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**Figure 10.** Hydrogen-bonding interactions involving lattice water molecules that partition the 2D sheets in 3. Thermal ellipsoids are shown with 50% probability.

**Figure 11.** Pb₄O₄ cubane motif that acts as a secondary building unit in 4. Only the carboxylate groups of ebdc and the oxygen contacts of coordinated DMF molecules have been shown for clarity.

**Figure 12.** Extended (4,4) 2D network present in 4.
the alkyne, of 1.31, was observed for ebdc in 4. This observation can be explained by the alkyne group aligning to a single oxygen site (O(1'), D = 3.152(5) Å, θ = 171°), unlike in cases 1 and 2, where the interaction is spread over two oxygen receptors, allowing more vibration between the two sites.

**Zinc 3D Net 5.** The final coordination polymer, [HNEt3]·[Zn3(μ3-OH)(μ2-H2O)(ebdc)](MeOH)0.67(H2O)0.33·MeOH·1.33H2O (S) consists of three zinc(II) atoms arrayed in a triangular cluster about a central μ2-OH group, which are ligated by three ebdc ligands and two aqua ligands, of which one is substitutionally disordered with a molecule of methanol. The anionic structure is charge-balanced in the lattice by a triethylammonium cation that contains one ethyl arm ligated by three ebdc ligands and two aqua ligands, of which one is substitutionally disordered with a molecule of methanol.

As in the case of copper network 3, the band intensity for the ethynyl C–H stretching vibration is both decreased and broadened; however, three unique bands were observable at 3301, 3271, and 3252 cm⁻¹ that can be assigned to the three ethynyl groups. The non-hydrogen-bonding ethynyl ligand in S matches the frequency of the non-hydrogen-bonding reference of H2ebdc at 3301 cm⁻¹. The band at 3271 cm⁻¹ is red-shifted by 30 cm⁻¹, which is characteristic of C–H···π bonding, and also possesses the strongest intensity, which likely reflects the optimal T-shaped geometry of the interaction shown in Figure 17. The final ethynyl group is directed toward the carboxylate group of an adjacent ligand with a C(30)···O(8) distance of 3.453(7) Å and an offset angle of 153° (Figure 18).

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**Packing Influence of ebdc.** This study has identified several lead(II) networks, 1, 2, and 4, which exhibit ethynyl-based weak hydrogen-bonding interactions that directly influence crystal packing. Two factors appear to favor C(sp)–H···O interactions within these structures, the first being the flexible coordination polyhedra imparted by the lead(II) atoms in conjunction with O-donor ligands. This allows the ebdc ligands to adopt a wide range of coordination modes and thus possess more freedom to achieve optimal hydrogen-bonding geometry. The second factor, in part, may relate to the choice of 3.59 Å (closest π···π contact, 3.38 Å). The 3D network propagates by bridging ebdc ligands of fused [Zn3(μ3-OH)3(μ2-H2O)]10⁺ centers, forming 1D channels that contain solvent and counterions (Figure 16). Analogous hexanuclear zinc(II) secondary building units have been observed in a limited number of instances.

There are three unique ethynyl groups within the 3D network of S. Two of these groups align in a manner suggestive of C–H···π hydrogen-bond interaction, with one acting solely as a donor and the other solely as an acceptor (Figure 17). Such T-shaped dimers have been calculated to impart stabilization energies on the order of 1 to 2 kcal/mol. The average C(20)···H distance is 2.74 Å, which is equivalent to the mean distance (2.72 Å) observed in a study of such terminal alkyne interactions. The final ethynyl group is directed toward the carboxylate group of an adjacent ligand with a C(30)···O(8) distance of 3.453(7) Å and an offset angle of 153° (Figure 18).

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of polar aprotic solvents during synthesis of the lead(II) coordination polymers. This solvent choice may have resulted in the number of hydrogen-bond acceptors far exceeding the number of donors, which would maximize the likelihood of the ethynyl groups of ebdc achieving a high degree of utilization.

By contrast, the two transition metal structures, 3 and 5, are both found to contain extensive hydrogen-bonding arrays as well as π−π stacking interactions, both of which are lacking in the lead(II) structures. Only minor ebdc-based weak hydrogen-bonding interactions were observed by IR spectroscopy, suggesting that the crystal packing, and thus self-assembly process, for these structures was not governed by weak hydrogen-bonding interactions to any large extent. While methanol was found to participate in these hydrogen-bonding nets, in most cases, it was observed to be substitutionally disordered with water molecules; water was also available in the wet solvents used during synthesis of the lead(II) structures. This suggests that solvent choice is not the main reason for a lack of weak hydrogen-bonding interactions. Instead, it is likely that the more rigid coordination environment provided by the transition metals promotes the inclusion of more coordinated solvent, which in turn increases the availability of strong hydrogen donors. The apical positions of the copper paddlewheel motif, which are typically coordinated by aqua ligands, demonstrates this effect. Similarly, the high Lewis acidity of zinc(II) promotes formation of hydroxo species and, in turn, cluster formation, which also likely favors strong, as opposed to weak, hydrogen bonding through the inclusion of more strong hydrogen donors into the crystalline structure. This provides the rationale for the contrasting behavior of lead

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**Figure 15.** Pairs of ebdc π−π stacking interactions located either side of the \([\text{Zn}_6(\mu_3-\text{OH})_2(\mu_2-\text{H}_2\text{O})_2]^{10+}\) cluster core further stabilize this motif. Thermal ellipsoids are shown with 50% probability.

**Figure 16.** One-dimensional solvated channels propagating through structure 5. The lattice triethylammonium, methanol, and water molecules have been omitted to highlight the channels.
and transition metal complexes with respect to the influence of hydrogen bonding in their solid-state assembly.

While gaining control over weak hydrogen-bond interactions during the self-assembly of coordination polymers remains a distant goal at the present time, this study has identified lead(II) and its corresponding coordination polymers as interesting species for study using ligands containing weak hydrogen-bond donors, in which the resulting weak hydrogen bonds are influential in forming the solid state-structures. This work will be expanded to include a range of other heavy metals ligated by ebdc as well as varying the weak hydrogen-bond donor on similar diacids in combination with lead(II) to extend the scope of these findings.

CONCLUSIONS

Of the five coordination polymers studied, three examples have been identified where weak hydrogen bonding derived from the ethynyl groups of the dicarboxylate ligand ebdc dominates the crystal packing and thus molecular self-assembly, namely, [Pb(ebdc)(DMF)]·H2O (1), [Pb2(ebdc)2(DMF)4]·H2O (2), and [Pb2(ebdc)2(DMF)4]·H2O (4). These lead(II) networks have proven to be valuable systems for the study of weak C(sp2)−H···O hydrogen-bonding interactions, both by single-crystal X-ray crystallography and IR spectroscopy. Two transition metal networks were also investigated, [Cu3(ebdc)2(H2O)4(MeOH)6]·6H2O (3) and [HNEt3]2[Zn2(μ2-H2O)-(μ-OH)2(ebdc)3(MeOH)2(H2O)2]·MeOH·1.33H2O (5); however, in these instances, strong hydrogen-bonding interactions predominated. Future work will establish if other heavy metals ligated by ebdc exhibit similar behavior to that of lead(II) and the possibility of varying the ethynyl group for other weak hydrogen-bond donors that may control the architecture in lead(II) networks.

ASSOCIATED CONTENT

Supporting Information

Packing motif of 1 (Figure S1), asymmetric units of 2 (Figure S2) and 4 (Figure S3), hydrogen-bonding tables for H2ebdc and 5, selected bond lengths and angles, and PXRD spectra for networks 1–5. This material is available free of charge via the Internet at http://pubs.acs.org. CCDC 1025919–1025924 contain the supplementary crystallographic data for this article. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via http://www.ccdc.cam.ac.uk/Community/Requestastructure/Pages/DataRequest.aspx.

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The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

Notes

The authors declare no competing financial interest.

ACKNOWLEDGMENTS

We are grateful to the EPSRC for financial support of the project (EP/K004956/1) and the University of Bath for a studentship to J.V.K.

ABBREVIATIONS

H2ebdc, 5-ethynyl-1,3-benzenedicarboxylic acid; bdc, 1,3-benzenedicarboxylate; IR, infrared; CCDC, Cambridge Crystallographic Data Centre; 1D, one dimensional; 2D, two dimensional; 3D, three dimensional; PXRDs, powder X-ray diffraction patterns

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