Donor- and acceptor-functionalized dibenzo[a,e]pentalenes: modulation of the electronic band gap†

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Dibenzo[a,e]pentalene (DBP) is a low-bandgap organic semiconductor. A versatile synthetic route to DBPs substituted with aryl or alkyl groups in the 5,10-positions and electron-donating or -accepting functionalities in the 2,7-positions is described. Six donor- or acceptor-functionalized DBP derivatives were synthesized that show amphoteric redox behavior and band gaps around 2 eV. Through choice of the 2,7-substituents, the HOMO/LUMO energy levels and band gaps can be adjusted within a range of up to 0.6 eV. In the solid state, the DBP derivatives assume herringbone-type packing structures.

Introduction

Dibenzo[a,e]pentalene (DBP, 1, Chart 1) is a ladder-type conjugated hydrocarbon that has received a significant amount of attention in recent years. As organic semiconductors with a low band-gap, derivatives of 1 have been employed in organic field-effect transistors (OFETs), exhibiting p-type behavior. First synthesized in 1912 by Brand,7 a variety of methods for the synthesis of DBP derivatives have been developed in the past ten years.8 While Brand’s original synthesis relies on the addition of Grignard reagents to diphenyl succindandione followed by a two-fold elimination of water,7,9,10 derivatives of DBP (1) can now be accessed from (di)aryl-alkynes by reductive11,12 or Lewis-acid-induced cyclizations13 as well as transition-metal-catalysed or -mediated processes.14-18 These methodologies were also applied to the synthesis of π-extended, ring-fused derivatives of 1.19-22 Related hydrocarbons consisting of fused five- and six-membered rings are indeno[1,2-b]fluorenes,23-26 bispentalenes,27 and other systems that include heteroatoms in the ring structure.28-31 In spite of the recent synthetic developments towards DBP derivatives, systematic studies on the modulation of its electronic properties through substitution have been undertaken only twice. Kawase et al. reported on the synthesis and properties of 2,3,7(8)-aryl-substituted DBPs,16 while Orita and co-workers attached aryl-ethynlenes to the 5,10-positions of the DBP core.32 We herein present the synthesis and optoelectronic properties of six DBP derivatives of type 2 (Chart 1), substituted with aryl-alkynes in the 2,7-positions. Electron-donating or -withdrawing groups R1 were chosen that allow for a modulation of the HOMO and LUMO energy levels. The aryl-alkyne groups were attached through posterior modification of the DBP core starting from 2,7-dibromo-DBPs of type 3. Six derivatives of 3 as precursors to 2 were synthesized carrying different aryl or alkyl groups in 5,10-position that allow for adapting the solubility properties. The solid-state structure of six DBP derivatives of type 2 and 3 are presented and discussed.

In order to most strongly influence the HOMO energy of the DBP core, which is relevant for an application as p-type material in OFETs, the C2 and C7 positions on the six-mem-
bered rings were chosen for attachment of the electron-donating or -withdrawing groups, since these possess the largest MO coefficients in the HOMO. The aryl-R substituents were connected to the DBP core through alkyne bridges, since these allow for co-planarity of the latter with the electronically modulating aryl group and hence maximum orbital overlap.

At C5 and C10 of the five-membered rings aryl or alkyl substituents R were attached to modulate the solubility and crystallinity of the resulting compounds. A high crystallinity would be advantageous for the charge carrier mobility in OFETs, while a good solubility in common solvents would allow for facile processing of the materials.

Results and discussion

Syntheses

For the synthesis of derivatives 3, we chose a method based on Brand’s procedure, since the more recent methods for DBP synthesis do not give access to derivatives suitable for further functionalization, such as with bromo-substituents. Bis-brominated diphenyl succindandione, synthesized in an improved synthetic procedure in three steps and 52% yield from ethyl-4-bromophenylacetate (see ESI†), was reacted with different Grignard or lithiated alkyl or aryl compounds R2M to yield 3a–f (Scheme 1). DBP derivatives 3a (R2 = Me), 3b (R2 = Et), and 3d (R2 = Ph) were synthesized through initial addition of the corresponding Grignard reagent to the keto groups in Scheme 1. DBP derivatives 3a (R2 = Me), 3b (R2 = Et), and 3d (R2 = Ph) were synthesized through initial addition of the corresponding Grignard reagent to the keto groups in Scheme 1. DBP derivatives 3a (R2 = Me), 3b (R2 = Et), and 3d (R2 = Ph) were synthesized through initial addition of the corresponding Grignard reagent to the keto groups in Scheme 1. DBP derivatives 3a (R2 = Me), 3b (R2 = Et), and 3d (R2 = Ph) were synthesized through initial addition of the corresponding Grignard reagent to the keto groups in Scheme 1. DBP derivatives 3a (R2 = Me), 3b (R2 = Et), and 3d (R2 = Ph) were synthesized through initial addition of the corresponding Grignard reagent to the keto groups in Scheme 1. DBP derivatives 3a (R2 = Me), 3b (R2 = Et), and 3d (R2 = Ph) were synthesized through initial addition of the corresponding Grignard reagent to the keto groups in Scheme 1. DBP derivatives 3a (R2 = Me), 3b (R2 = Et), and 3d (R2 = Ph) were synthesized through initial addition of the corresponding Grignard reagent to the keto groups in Scheme 1. DBP derivatives 3a (R2 = Me), 3b (R2 = Et), and 3d (R2 = Ph) were synthesized through initial addition of the corresponding Grignard reagent to the keto groups in Scheme 1. DBP derivatives 3a (R2 = Me), 3b (R2 = Et), and 3d (R2 = Ph) were synthesized through initial addition of the corresponding Grignard reagent to the keto groups in Scheme 1. DBP derivatives 3a (R2 = Me), 3b (R2 = Et), and 3d (R2 = Ph) were synthesized through initial addition of the corresponding Grignard reagent to the keto groups in Scheme 1.

The bis-brominated DBPs 3a, 3b and 3d–f were further investigated by X-ray crystallography; their solid-state structures will be discussed below.

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Optoelectronic properties

(TD)DFT calculations were performed in order to gain information on the orbital energies and electronic transitions in the absorption spectra (see ESI† for details). For calculations all mesityl groups in 5a–f and the ethyl groups in 5f were replaced by methyl groups, denoted with an asterix (5a–f*). Four different density functionals in combination with D3 and the def2-TZVP basis set were tested with varying amounts of HF exchange from 0% to 50%, since this has been shown to have a pronounced effect: BP86 (0%), TPSSH (10%), B3LYP (20%) and BHLYP (50%). The closest match between TDDFT calculations and experimental absorption spectra was found with B3LYP, while TPSSH gave the best orbital energies (see ESI† for details), and hence these results will be used in the further discussion. Selected calculated orbital energies of 5a–f* as well as the molecular orbitals of 5c* are shown in Fig. 1. With increasing electron-withdrawing character of the substituent R the frontier molecular orbital energies of 5a–f* decrease. The HOMO energies experience a stronger shift (ΔE = 1.0 eV) compared to the LUMO energies (ΔE = 0.88 eV), which we had anticipated before based on the size of the MO coefficients in the HOMO and LUMO of DBP (I, see above). The order of decrease in energy of the orbitals from 5a* to 5f* corresponds well with the Hammett substitution constants σ of R1: the σpar constants of the substituents R1 amount to ~0.83 for NMe2 (5a*), ~0.27 for OMe (5b*), 0 for H (5c*), 0.45 for CO2Me (5d*), 0.66 for CN (5e*), and ~0.8 § for CHC(CN)(CO2Me) (5f*).

†(TD)DFT calculations were performed with the TURBOMOLE 6.6 program package (TURBOMOLE, University of Karlsruhe, Germany, 2012, http://www.turbomole.com).‡A value of 0.84 was reported for the electronically very similar substituent CHC(CN)(CN).§
The absorption spectra of 5a-f are displayed in Fig. 2. All derivatives show three characteristic absorption bands, which were assigned to electronic transitions using TDDFT calculations (B3LYP-D3/def2-TZVP). The lowest energy bands from 470–700 nm, well visible on the logarithmic scale in the inset in Fig. 2, can be ascribed to the HOMO → LUMO transitions (S1 to S2). These transitions are symmetry forbidden, since both orbitals are of a symmetric (Laporte’s rule, see also Fig. 1), which is reflected in the relatively low absorption coefficients of \( \log \varepsilon = 2.9–3.4 \). This low energy band appears at very similar wavelengths for all substituents with maxima from 530–534 nm except for 5a and 5f. The NMe2 substituent in 5a affects a bathochromic shift to 567 nm, while the CHC(CN)(CO2Et) substituent in 5f leads to a hypsochromic shift to 511 nm. The second absorption bands from 380–470 nm can be assigned to the HOMO–1 → LUMO, HOMO–2 → LUMO and other transitions, while the most intense (\( \log \varepsilon = 5.0–5.1 \)) and highest energy absorption bands from 270–380 nm can be attributed to the HOMO → LUMO–1 transitions. The CHC(CN)(CO2Et) substituent presents an exception, since the most intense absorption is the HOMO–1 → LUMO transition at 414 nm. Overall, the NMe2-substituted derivative 5a experiences the strongest bathochromic shift of all absorption bands. This is in accordance with the electrochemical results and calculations that show a strong increase in HOMO and LUMO–1 energy relative to the other derivatives, while the LUMO energy is less strongly affected (see also Fig. 1 and Table 1). This observation is in agreement with our initial assumption, that substitution at the 2,7-positions of the DBP core would most strongly affect the HOMO energy level (see above). The optical transitions in 5f are different compared to 5a–e, likely due to the extended conjugation into the CHC(CN)(CO2Et) substituent. The band gaps of 5a–f were obtained from the onset of the longest wavelength absorption and are listed in Table 1. None of the DBP derivatives 5a-f showed fluorescence. This has also been observed for the indeno[1,2-b]fluorenes, in which case it was rationalized through transient absorption spectroscopy measurements that showed that the excited state lifetimes were shorter than the timescale at which fluorescence is observed.

The electrochemical properties of 5a-f were investigated through cyclic voltammetry measurements in CH2Cl2 solution (Fig. 3 and Table 1), where the first oxidation and reduction can be used to determine the HOMO and LUMO energies, respectively. All DBP derivatives 5a-f show amphoteric redox behaviour. The redox potentials correspond well with the electron-donating or withdrawing character of the substituents \( \mathbf{R}^1 \): CO2Me (5d), CN (5e) and CHC(CN)(CO2Et) (5f) lead to a positive shift in reduction and oxidation potential relative to 5c with \( \mathbf{R}^1 = \text{H} \), while this trend is reversed for the OMe (5b) and NMe2 (5a) substituents. The NMe2 substituent significantly lowers the oxidation potential and hence the HOMO

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**Fig. 1** Molecular orbitals of 5c* (\( \mathbf{R}^1 = \text{H} \), left) and energy diagram (eV, right) of calculated MOs of DBP derivatives 5a–f* (TPSSH-D3/def2-TZVP, a CPM = CHC(CN)(CO2Me)).

**Fig. 2** Absorption spectra of 5a–f in CH2Cl2 solution (CPE = CHC(CN)(CO2Et)). Inset: Spectra plotted on a logarithmic scale.

**Table 1** Electrochemical and optical data for aryl–alkyne-substituted DBPs 5a–f

| \( \mathbf{R}^1 \) | \( E_{\text{red}} \) (V) | \( E_{\text{ox}} \) (V) | \( E_{\text{HOMO}} \) (eV) | \( E_{\text{LUMO}} \) (eV) | \( E_{\gamma} \) (V) | \( E_{\text{Eg}} \) (opt) (eV) |
|---|---|---|---|---|---|---|
| 5a | -1.79 | 0.39 | -5.17 | -3.18 | 1.99 | 1.82 |
| 5b | -1.77 | 0.78 | -5.57 | -3.19 | 2.39 | 1.99 |
| 5c | -1.74 | 0.83 | -5.68 | -3.22 | 2.46 | 2.08 |
| 5d | -1.71 | 0.93 | -5.73 | -3.27 | 2.46 | 2.07 |
| 5e | -1.67 | 0.98 | -5.77 | -3.28 | 2.49 | 2.07 |
| 5f | -1.68 | 0.89 | -5.69 | -3.34 | 2.35 | 1.95 |

* From the onset of the reduction/oxidation peak vs. Fe/Fc*, assuming an ionization energy of 4.8 eV for ferrocene.50 50 From the onset of the longest wavelength absorption band. 50 Cathodic peak potential. 50 Anodic peak potential.

Large peak separations might be due to a temperature effect (CVs were measured in the glove box at 25–30 °C) or the specific cell setup.
energy of 5a by $\Delta E_{\text{HOMO}} = 0.51$ eV relative to 5c ($R^1 = H, E_{\text{HOMO}} = -5.68$ V). For 5f ($R^1 = \text{CHC(CN)(CO}_2\text{Et})$) two reduction waves were observed, of which the first one was not reversible. This might be related to the close proximity in energy of the LUMO and LUMO+1, as found in DFT calculations (see Fig. 1). For all derivatives the electrochemically determined band gaps are larger than the optical band gaps, which has been observed for other non-alternant hydrocarbons before.

Our investigations show that by attaching electron-donating or -withdrawing substituents to the 2,7-positions of DBP, its electronic properties can be significantly modulated. The HOMO energies show the largest effect, as anticipated, with a difference of 0.60 eV between 5a and 5e, while the LUMO energies vary within 0.16 eV (5a vs. 5f).

### Structural properties

The molecular packing in the solid state is important for the performance of organic semiconductors in electronics devices. The solid-state structures of six of the synthesized DBP derivatives were investigated by X-ray crystallography. Single crystals suitable for X-ray diffraction were obtained by slow evaporation of solutions in CH$_2$Cl$_2$ (3a, d, e), CDCl$_3$ (3b), EtOAc/cyclohexane (3f) or CHCl$_3$ (5e). The molecular structures of 5c, 3a, 3b and 3f are shown in Fig. 4 (see ESI† for 3d and 3e). In all derivatives, the DBP core assumes an almost planar conformation. The phenyl-alkyne substituents in 5c are slightly rotated relative to the DBP core by $\sim25^\circ$. The dihedral angle between the mesityl groups and the DBP core amounts to $\sim70^\circ$ due to steric hindrance associated with the two ortho-methyl groups. Hence electronic communication through π-orbital overlap between the mesityl groups and the DBP core can be excluded. The bond lengths in the six-membered rings of the DBP core in 5c are relatively homogeneous (1.377–1.341 Å), while in the five-membered rings the C10–C10’ bond is short (1.362 Å) and the C5’–C10’, C4’–C5’ and C1’–C10 bonds are comparably long (1.464–1.477 Å). In the crystal lattice 5c assumes a herringbone pattern, in which the planes of the DBP cores of two non-parallel aligned molecules are oriented at an angle of $\sim70^\circ$ (Fig. 4b). 3a, b, and f also assume herringbone-type packing structures in the solid state (Fig. 4c–e). The distances between to parallel aligned molecules, viewed along the crystallographic a-axis, amount to 3.19 Å for 3a, 3.38 Å for 3b, and 4.26 Å for 3f.

### Conclusions

In summary, we have established a versatile synthetic route to 2,7-dibromo-functionalized dibenzo[α,ε]pentalenes 3 with aryl or alkyl substituents in 5,10-position that can be further derivatized through cross-coupling reactions. Six electron-donating or -accepting aryl alkynes were attached through Sonogashira reactions, and the resulting 2,7-aryl-alkyne-substituted DBPs 5a–f were investigated regarding their optoelectronic properties. The choice of the substituent allowed modulating the HOMO or LUMO energy level within a range of 0.60 eV or 0.16 eV, respectively, and the band gap within a range of 0.26 eV. In the solid state, the DBP derivatives 3a, 3b, 3f and 5c assume herringbone-type packing structures. The compounds show...
amphoteric redox behaviour and are attractive candidates for application in organic electronics devices.

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Notes and references

1 H. Hopf, Angew. Chem., Int. Ed., 2013, 52, 12224–12226.
2 C. Liu, S. Xu, W. Zhu, X. Zhu, W. Hu, Z. Li and Z. Wang, Chem. – Eur. J., 2015, 21, 17016–17022.
3 G. Dai, J. Chang, W. Zhang, S. Bai, K.-W. Huang, J. Xu and C. Chi, Chem. Commun., 2015, 51, 503–506.
4 T. Kawase, T. Fujiwara, C. Kitamura, A. Konishi, Y. Hirao, K. Matsumoto, H. Kurata, T. Kubo, S. Shimamura, H. Mori, E. Miyazaki and K. Takimia, Angew. Chem., Int. Ed., 2010, 49, 7728–7732.
5 C. Li, C. Liu, Y. Li, X. Zhu and Z. Wang, Chem. Commun., 2015, 51, 693–696.
6 M. Nakano, I. Osaka, K. Takimia and T. Koganewaza, J. Mater. Chem. C, 2014, 2, 64–70.
7 K. Brand, Chem. Ber., 1912, 45, 3071–3077.
8 M. Saito, Symposium, 2010, 2, 950–969.
9 K. Brand and F. W. Hoffmann, Chem. Ber., 1920, 53, 815–821.
10 K. Brand and F. Schläger, Chem. Ber., 1923, 56, 2541–2545.
11 M. Saito, M. Nakamuk, T. Tajima and M. Yoshioka, Angew. Chem., Int. Ed., 2007, 46, 1504–1507.
12 H. Zhang, T. Karasawa, H. Yamada, A. Wakamiya and S. Yamaguchi, Org. Lett., 2009, 11, 3076–3079.
13 C. Chen, M. Harhausen, R. Liedtke, K. Bussmann, A. Fuakazawa, S. Yamaguchi, J. L. Petersen, C. G. Danilic, R. Fröhlich, G. Kehr and G. Erker, Angew. Chem., Int. Ed., 2013, 52, 5992–5996.
14 A. S. K. Hashmi, M. Wieteck, I. Braun, P. Nösel, L. Jongbloed, M. Rudolph and F. Rominger, Adv. Synth. Catal., 2012, 354, 555–562.
15 Z. U. Levi and T. D. Tilley, J. Am. Chem. Soc., 2009, 131, 2796–2797.
16 T. Kawase, A. Konishi, Y. Hirao, K. Matsumoto, H. Kurata and T. Kubo, Chem. – Eur. J., 2009, 15, 2653–2661.
17 T. Maekawa, Y. Segawa and K. Itami, Chem. Sci., 2013, 4, 2369–2373.
18 J. Zhao, K. Oniwa, N. Asao, Y. Yamamoto and T. Jin, J. Am. Chem. Soc., 2013, 135, 10222–10225.
19 Z. U. Levi and T. D. Tilley, J. Am. Chem. Soc., 2010, 132, 11012–11014.
20 H. Li, X.-Y. Wang, B. Wei, L. Xu, W.-X. Zhang, J. Pei and Z. Xi, Nat. Commun., 2014, 5, 4508.
21 F. Xu, L. Peng, K. Shinohara, T. Nishida, K. Wakamatsu, M. Uejima, T. Sato, K. Tanaka, N. Machida, H. Akashi, A. Orita and J. Otera, Org. Lett., 2015, 17, 3014–3017.
22 J. Shen, D. Yuan, Y. Qiao, X. Shen, Z. Zhang, Y. Zhong, Y. Yi and X. Zhu, Org. Lett., 2014, 16, 4924–4927.
23 D. T. Chase, B. D. Rose, S. P. McClintock, L. N. Zakharov and M. M. Haley, Angew. Chem., Int. Ed., 2011, 50, 1127–1130.
24 D. T. Chase, A. G. Fix, S. J. Kang, B. D. Rose, C. D. Weber, Y. Zhong, L. N. Zakharov, M. C. Lonergan, C. Nuckolls and M. M. Haley, J. Am. Chem. Soc., 2012, 134, 10349–10352.
25 T. Maekawa, H. Ueno, Y. Segawa, M. M. Haley and K. Itami, Chem. Sci., 2016, 7, 650–654.
26 J. L. Marshall, N. J. O’Neal, L. N. Zakharov and M. M. Haley, J. Org. Chem., 2016, 81, 3674–3680.
27 J. Cao, G. London, O. Dumele, M. van Wantoch Rekowski, N. Trapp, L. Ruhlmann, C. Boudon, A. Stanger and F. Diederich, J. Am. Chem. Soc., 2015, 137, 7178–7188.
28 J. D. Wood, J. L. Jellison, A. D. Finke, L. Wang and K. N. Plunkett, J. Am. Chem. Soc., 2012, 134, 15783–15789.
29 X.-Y. Wang, A. Nairita, X. Feng and K. Mullen, J. Am. Chem. Soc., 2015, 137, 7668–7671.
30 A. Konishi, T. Fujiwara, N. Ogawa, Y. Hiroa, K. Matsumoto, H. Kurata, T. Kubo, C. Kitamura and T. Kawase, Chem. Lett., 2010, 39, 300–301.
31 G. Dai, J. Chang, J. Luo, S. Dong, N. Aratani, B. Zheng, K. W. Huang, H. Yamada and C. Chi, Angew. Chem., Int. Ed., 2016, 55, 2693–2696.
32 F. Xu, L. Peng, A. Orita and J. Otera, Org. Lett., 2012, 14, 3970–3973.
33 B. Esser, Phys. Chem. Chem. Phys., 2015, 17, 7366–7372.
34 J.-J. Shen, J.-Y. Shao, X. Zhu and Y.-W. Zhong, Org. Lett., 2015, 18, 256–259.
35 T. Imamoto, N. Takiyama, K. Nakamura, T. Hatajima and Y. Kamiya, J. Am. Chem. Soc., 1989, 111, 4392–4398.
36 C. Hansch, A. Leo and R. W. Taft, Chem. Rev., 1991, 91, 165–195.
37 S. Grimm, J. Antony, S. Ehrlich and H. Krieg, J. Chem. Phys., 2010, 132, 154104.
38 S. Grimm, S. Ehrlich and L. Goerigk, J. Comput. Chem., 2011, 32, 1456–1465.
39 A. Schäfer, C. Huber and R. Ahlrichs, J. Chem. Phys., 1994, 100, 5829–5835.
40 G. Zhang and C. B. Musgrave, J. Phys. Chem. A, 2007, 111, 1554–1561.
41 A. D. Becke, Phys. Rev. A, 1988, 38, 3098–3100.
42 J. P. Perdew, Phys. Rev. B: Condens. Matter, 1986, 33, 8822–8824.
43 J. P. Perdew, Phys. Rev. B: Condens. Matter, 1986, 34, 7406.
44 J. Tao, J. P. Perdew, V. N. Staroverov and G. E. Scuseria, Phys. Rev. Lett., 2003, 91, 146401.
45 V. N. Staroverov, G. E. Scuseria, J. Tao and J. P. Perdew, J. Chem. Phys., 2003, 119, 12129.
46 P. J. Stephens, F. J. Devlin, C. F. Chabalowski and M. J. Frisch, J. Phys. Chem., 1994, 98, 11623–11627.
47 A. D. Becke, J. Chem. Phys., 1993, 98, 5648–5652.
48 A. D. Becke, J. Chem. Phys., 1993, 98, 1372–1377.
49 B. D. Rose, L. E. Shoer, M. R. Wasielewski and M. M. Haley, Chem. Phys. Lett., 2014, 616–617, 137–141.
50 B. W. D’Andrade, S. Datta, S. R. Forrest, P. Djurovich, E. Polikarpov and M. E. Thompson, Org. Electron., 2005, 6, 11–20.
51 R. Gleiter and G. Haberhauer, Aromaticity and Other Conjugation Effects, Wiley-VCH, Weinheim, Germany, 2012.
52 X. Yang, X. Shi, N. Aratani, T. P. Goncalves, K.-W. Huang, H. Yamada, C. Chi and Q. Miao, Chem. Sci., 2016, 7, 6176–6181.
53 R. Ahlrichs, M. Bär, H. Marco, H. Horn and C. Kölmel, Chem. Phys. Lett., 1989, 162, 165–169.