A Two-Step SO$_3$H/ICG Catalyst Synthesis for Biodiesel Production: Optimization of Sulfonation Step via Microwave Irradiation

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Abstract

Conventional heating, a common method used for heterogeneous solid acid catalyst synthesis unknowingly consumes massive time and energy. In this study, acid catalyst was prepared through sulfonation process of incomplete carbonized glucose (ICG) via microwave-assisted technique to shorten the heating time and energy consumption. Optimization of the sulfonation process of ICG via microwave-assisted was carried out. Four-factor-three-level central composite design (CCD) was used to develop the design of experiments (DOE). Interaction between two factors was evaluated to determine the optimum process conditions. A quadratic model was proposed for prediction of biodiesel yield ($Y$) from palm fatty acid distillate (PFAD) and its conversion ($C$). The application of DOE successfully optimized the operating conditions for the two-step SO$_3$H/ICG catalyst synthesis to be used for the esterification process. The optimized conditions of the best performing SO$_3$H/ICG with maximum $Y$ and $C$ were at 7.5 minutes of reaction time, 159.5 mL of H$_2$SO$_4$ used, 671 rpm of stirring rate as well as 413.64 watt of power level. At these optimum conditions the predicted yield percentage and conversion percentage were 94.01% and 91.89%, respectively, which experimentally verified the accuracy of the model. The utilization of sulfonated glucose solid acid catalyst via microwave-assisted in biodiesel production has great potential towards sustainable and green method of synthesizing catalyst for biodiesel.

Keywords: Optimization; Palm Fatty Acid Distillate; Esterification; Sulfonated Glucose; Microwave-assisted

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1. Introduction

The process optimization in biodiesel production depends on independent variable process parameters, such as: percentage weight of the catalyst (wt%), molar ratio of alcohol to oil, reaction temperature (°C) as well as reaction time (min) [1]. However, there is lack of study in optimizing the independent variable during the preparation of the heterogeneous acid catalyst for biodiesel production itself. This catalyst preparation segment is imperative in biodiesel production in order to further reduce time and energy consumption for its synthesis since it possess high recyclability and are applied extensively in a conventional way. Zong [2] and Lokman [3] did great work in functionalizing the...
carbonized D-glucose using sulfuric acid in the esterification of high FFA feedstock for biodiesel production. However, both study reported that the time to sulfonate or functionalize the carbonized d-glucose required almost 15 h which consumed a lot of time and energy as well. Previous study by Shuit and Tan [4] modified the sulfonation method for multi walled carbon nano-tubes (MWCNTs) thermal decomposition and thermal treatment under nitrogen gas (\(N_2\)) flow for esterification of palm fatty acid distillate (PFAD). Optimization of these parameters was often done traditionally by using one-factor-at-a-time (OFAT) method where one factor is varied while the other factors are kept constant. This method is tedious if it involves large number of parameters.

Optimization using design of experiments (DOE) is more efficient statistically when dealing with more than 3 variables. Many researchers used DOE for optimizing their biodiesel production from high free fatty acid (FFA) feedstock such as waste frying oil [5], animal fats [6] as well as PFAD [7]. Most of the investigations optimized the variables of the biodiesel production for esterification including catalyst loading (wt%), reaction temperature (\(^{\circ}\)C), molar ratio of alcohol to oil, as well as reaction time (min). Although the optimizations for esterification for biodiesel production were well covered by researchers, the optimization for heterogeneous solid catalyst synthesis is not sufficiently dealt with. If the solid catalyst is well optimized for the sulfonation stage, massive benefit could be attained and this can lead biodiesel production to be more commercially viable. Besides systematically optimizing the synthesis of \(\text{SO}_3\text{H}/\text{ICG} \) catalyst, microwave irradiation was utilized to enhance the rate of heating which will rapidly speed up the duration of catalyst preparation. To the best of our knowledge, the application of microwave irradiation for biodiesel catalyst synthesis context via DOE has not been reported by others. Therefore, the main purpose of this study is to optimize the sulfonation variables including the volume of \(\text{H}_2\text{SO}_4 (\text{mL})\), sulfonation time (min), power level of microwave (W) and power stirring (rpm). Central Composite Design (CCD) of RSM in Design Expert 7\(^{th}\) version was employed for optimization for the sulfonation of the incomplete carbonized glucose (ICG). CCD was chosen because it is commonly applied in designing sequential experimental which consists of multivariate equations. The multivariate equations is simultaneously resolved by manipulations of the quantitative data in experimental design of the RSM [8]. This method has practically applied to determine the optimum conditions since RSM offers efficient and easier method compared to the conventional OFAT experiment design.

2. Materials and Methods

2.1 Materials

Main chemical used was D-(+) Glucose (100% purity) which had been bought from QRec (Asia) Sdn. Bhd. The PFAD used for catalytic performance was supplied by Mewaholeo Industries Sdn. Bhd., Pasir Gudang, Malaysia. Methanol (99.9% purity) and \(\text{H}_2\text{SO}_4\) (98.0% purity) was supplied from Merck and J.T. Baker, US, respectively. Gas chromatography (GC) Flame Ionization Detector (FID) Agilent brand and N7890 model was employed for the quantification. Methyl ester's standard for GC analysis, such as: methyl linoleate, methyl myristate, methyl oleate, methyl palmitate, and methyl stearate, were purchased from Sigma-Aldrich chemical company. All analytical grade products for all chemicals required no further purification. Equipment employed for catalyst preparation is Microwave SHARP R213CST.

2.2 Catalyst Preparation

20 g of D(+) Glucose powder was melted through heating process in a microwave for 20 minutes at medium power level (400 W). The ICG formed was then crushed into powder form and sieved at 150 nm. 4 g of the crushed ICG was added with 100 mL of concentrated \(\text{H}_2\text{SO}_4\) and stirred at 500 rpm for 5 minutes for sulfonate introduction to the glucose based. The mixture was then heated inside a microwave in medium high power level for sulfonation process for certain respective reaction times. The mixtures was then filtered and black precipitate was collected before washing process using hot distilled water at 85 °C to remove impurities contained in the mixture. Finally the catalyst was dried inside an oven for 5 hours to remove the moisture content in the catalyst.

2.3 Methyl Ester Production

FFA of PFAD can be reduced through esterification process. PFAD were weighed based on molar ratio and properly heated at 65 °C. Preheated PFAD was poured into a three neck round bottom flask following by the methanol and catalyst. Molar ratio of PFAD and methanol used was 10:1. 2.5 wt% of sulfonated glucose loading was used for the esterification.
2.4 Analysis of PFAD Methyl Ester

The percentage yield of biodiesel was calculated using Equation 1. Biodiesel in this study is defined as the top 5 components which are methyl linoleate, methyl myristate, methyl oleate, methyl palmitate and methyl stearate. From the esterification reaction, only methyl esters were obtained and no other compounds were present. The conversion of the biodiesel was determined through acid-base titration method. About 1 mL of biodiesel was added into 20 mL of the solvent, methanol. The mixture was then shaken until it dissolved. 2–3 drops of phenolphthalein was added into the flask and titrated with standardized potassium hydroxide (KOH) solution until the first permanent pink color showed up. The formula to calculate the percentage yield, acid value and percentage conversion were defined by equation (1), (2) and (3), respectively.

\[
\text{Yield}(\%) = \frac{\text{weight of experimental biodiesel}(g)}{\text{weight of theoretical biodiesel}(g)} \times 100
\]

\[
\text{Acid value} = \frac{N \times MW \text{ KOH} \times \text{Volume of KOH titrated (mL)}}{\text{mass of biodiesel}(g)}
\]

where, \(N\) refers to the normality of KOH used and \(MW\) indicating the molecular weight of KOH which is 56.11 g/mol.

\[
\text{Conversion}(\%) = \frac{(\text{acid value of PFAD} - \text{acid value of sample})}{\text{acid value of PFAD}} \times 100
\]

2.5 Optimizing Using CCD (Central Composite Design)

In this study, four factors Central Composite Design (CCD) require 30 experiments to be run, which includes six center points, six axial points, and eight factorial points. Yield and conversion response were analyzed to evaluate the performance of the process conducted. To produce experimental design for the synthesis of \(\text{SO}_3\text{H}\)/ICG catalyst, heating time of catalyst, volume of \(\text{H}_2\text{SO}_4\) used, power level and stirring rate were chosen as parameters that influence the efficiency of catalyst produced.

The fitness and adequacy of the models were validated using F-values, \(P\) values and affirmed further by the \(R^2\) value. Models achieving 95% confidence interval indicates acceptable accuracy. Some parts of the analysis were discussed and related to catalyst characterization results. The catalyst characterization are not discussed here but is referred to our previous report [10].

3. Results and Discussion

3.1 Single Response Optimization for Yield (Y) and Conversion (C)

In this chapter, five level of four factorial CCD were selected because it can fit a full quadratic model. CCD provides a factorial design with center points, augmented with a group of axial points which can estimate curvature. This level requires 30 series of experimental design including six center points, 6 axial points as well as 8 factorial points. The responses of this optimization were percentage yield of FAME and percentage of conversion of FFA in PFAD. Four independent variables used in sulfonation of the ICG were sulfonation time \((X_1)\), volume of \(\text{H}_2\text{SO}_4\) \((X_2)\), power level of microwave \((X_3)\), as well as stirring rate \((X_4)\). These 4 parameters were selected based on the most significantly variables influencing the catalytic activity as well as from previous study [9]. Table 1 tabulates experimental level coded and range for independent variables. The level coded ranging from the lowest \((-\alpha)\) to the median to the highest \((+\alpha)\). Time labelled as \(X_1\) was ranging from 3 to 11 min, volume of \(\text{H}_2\text{SO}_4\) \((X_2)\) was from 50 mL to 250 mL, power level of microwave \((X_3)\) from 80 W to 720 W and stirring rate \((X_4)\) from 200 rpm to 1000 rpm. Saimon and colleagues managed to obtain high percentage yield of FAME 91.41% using 100 mL of \(\text{H}_2\text{SO}_4\) at 7 min of sulfonation time and 560 W of microwave’s power level [10]. Thus,

| Parameter          | \(-\alpha\) | \(-1\) | 0   | +1  | \(+\alpha\) |
|--------------------|-------------|--------|-----|-----|-------------|
| Time, \(X_1\) (min) | 3           | 5      | 7   | 9   | 11          |
| Volume \(\text{H}_2\text{SO}_4, X_2\) (mL) | 50          | 100    | 150 | 200 | 250         |
| Power Level, \(X_3\) (watt) | 80          | 240    | 400 | 560 | 720         |
| Stirring rate, \(X_4\) (rpm) | 200         | 400    | 600 | 800 | 1000        |

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the ranges were extended to determine the optimum value for these independent variables. The respective catalyst characterizations for this investigation that include XRD, BET, SEM, FTIR and TPD–NH3 have been reported earlier [10].

As mentioned earlier, the responses of yield (Y) as well as conversion (C) were investigated using multiple regression analysis. Second order polynomial [11] was applied in this analysis as presented by Equation (4) as follow:

$$y = \beta_0 + \sum_{i=1}^{3} \beta_i x_i + \sum_{i=1}^{3} \beta_i x_i^2 + \sum_{i=1}^{3} \sum_{j=i+1}^{3} \beta_{ij} x_i x_j$$  \hspace{0.5cm} (4)

where y is the percentage yield of FAME or percentage of FFA conversion; $x_i$ and $x_j$ are the independent variables and $\beta_0$, $\beta_i$, $\beta_{ij}$ are intercept, linear, quadratic as well as interaction constant coefficients respectively.

The DOE complete with the independent variables as well as responses experimental and predicted are tabulated in Table 2. The interaction of the independent variables and the responses are derived as shown in Equation (5) and Equation (6).

$$Y = 93.20 + 1.346 X_1 + 1.53 X_2 + 0.55 X_3 + 1.48 X_4$$
$$+ 0.18 X_1 X_2 + 0.68 X_1 X_3 - 0.076 + 0.089 X_2 X_3$$
$$+ 0.081 X_2 X_4 - 0.4 X_3 X_4 - 1.91 X_1^2 - 2.83 X_2^2$$
$$- 3.52 X_3^2 - 1.21 X_4^2$$  \hspace{0.5cm} (5)

$$C = 91.40 + 1.36 X_1 + 1.31 X_2 + 0.64 X_3 + 1.45 X_4$$
$$+ 0.10 X_1 X_2 + 0.70 X_1 X_3 + 0.20 X_1 X_4$$
$$- 0.19 X_2 X_3 - 0.096 X_3 X_4 - 3.17 X_1^2$$
$$- 4.02 X_2^2 - 4.50 X_3^2 - 2.54 X_4^2$$  \hspace{0.5cm} (6)

| Run | $X_1$ | $X_2$ | $X_3$ | $X_4$ | Yexp (%) | Y RSM (%) | Cexp (%) | C RSM (%) |
|-----|-------|-------|-------|-------|----------|-----------|----------|-----------|
| 1   | 5     | 100   | 240   | 400   | 80.72    | 79.55     | 74.21    | 72.91     |
| 2   | 7     | 150   | 400   | 200   | 84.47    | 85.39     | 77.65    | 78.35     |
| 3   | 7     | 50    | 400   | 600   | 78.21    | 78.82     | 71.90    | 72.70     |
| 4   | 5     | 100   | 560   | 800   | 81.08    | 82.05     | 74.54    | 75.67     |
| 5   | 9     | 100   | 560   | 800   | 85.33    | 85.59     | 81.20    | 80.41     |
| 6   | 5     | 100   | 560   | 400   | 79.52    | 79.9      | 73.10    | 73.36     |
| 7   | 7     | 150   | 80    | 600   | 77.41    | 78.68     | 71.16    | 72.11     |
| 8   | 7     | 150   | 400   | 600   | 92.43    | 93.20     | 90.51    | 91.40     |
| 9   | 11    | 150   | 400   | 600   | 87.34    | 88.24     | 80.29    | 81.44     |
| 10  | 5     | 200   | 560   | 800   | 84.62    | 85.10     | 77.79    | 78.13     |
| 11  | 7     | 150   | 400   | 1000  | 92.17    | 91.29     | 84.73    | 84.16     |
| 12  | 7     | 150   | 400   | 600   | 93.51    | 93.20     | 91.76    | 91.40     |
| 13  | 7     | 150   | 400   | 600   | 92.85    | 93.20     | 90.31    | 91.40     |
| 14  | 5     | 200   | 240   | 400   | 81.45    | 81.92     | 74.88    | 76.13     |
| 15  | 9     | 200   | 560   | 400   | 86.54    | 87.18     | 79.01    | 79.34     |
| 16  | 7     | 150   | 400   | 600   | 93.54    | 93.20     | 92.43    | 91.40     |
| 17  | 9     | 100   | 560   | 400   | 84.72    | 83.74     | 77.88    | 77.3      |
| 18  | 9     | 200   | 240   | 400   | 85.46    | 83.74     | 78.56    | 76.84     |
| 19  | 9     | 200   | 240   | 800   | 87.16    | 87.50     | 80.13    | 80.34     |
| 20  | 7     | 150   | 400   | 600   | 94.01    | 93.20     | 92.76    | 91.40     |
| 21  | 3     | 150   | 400   | 600   | 83.75    | 82.89     | 76.99    | 75.98     |
| 22  | 9     | 100   | 240   | 800   | 85.11    | 84.09     | 78.24    | 77.54     |
| 23  | 9     | 100   | 240   | 400   | 80.41    | 80.65     | 73.92    | 74.04     |
| 24  | 5     | 200   | 560   | 400   | 82.36    | 82.63     | 75.71    | 75.82     |
| 25  | 5     | 200   | 240   | 800   | 85.76    | 85.98     | 78.84    | 78.83     |
| 26  | 7     | 250   | 400   | 600   | 85.52    | 84.95     | 78.62    | 77.96     |
| 27  | 5     | 100   | 240   | 800   | 83.21    | 83.29     | 75.48    | 75.61     |
| 28  | 7     | 150   | 720   | 600   | 82.11    | 80.88     | 75.48    | 74.67     |
| 29  | 7     | 150   | 400   | 600   | 92.84    | 93.20     | 90.63    | 91.40     |
| 30  | 9     | 200   | 560   | 800   | 88.93    | 89.35     | 81.75    | 82.45     |

| c: Centre value of all parameters. |

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Table 3. Analysis of variance (ANOVA) for yield response of biodiesel production.

| Source   | Sum of Squares | df  | Mean Square | F Value | p-value Prob > F | Comments |
|----------|----------------|-----|-------------|---------|------------------|----------|
| Model    | 672.62         | 14  | 48.04       | 41.16   | < 0.0001         | significant |
| X<sub>1</sub> | 42.99         | 1   | 42.99       | 36.83   | < 0.0001         |          |
| X<sub>2</sub> | 56.43         | 1   | 56.43       | 48.34   | < 0.0001         |          |
| X<sub>3</sub> | 7.28          | 1   | 7.28        | 6.24    | 0.0246           |          |
| X<sub>4</sub> | 52.27         | 1   | 52.27       | 44.78   | < 0.0001         |          |
| X<sub>1</sub>X<sub>2</sub> | 0.51          | 1   | 0.51        | 0.44    | 0.5181           |          |
| X<sub>1</sub>X<sub>3</sub> | 7.48          | 1   | 7.48        | 6.41    | 0.0230           |          |
| X<sub>1</sub>X<sub>4</sub> | 0.093         | 1   | 0.093       | 0.080   | 0.7816           |          |
| X<sub>2</sub>X<sub>3</sub> | 0.13          | 1   | 0.13        | 0.11    | 0.7470           |          |
| X<sub>2</sub>X<sub>4</sub> | 2.54          | 1   | 2.54        | 2.18    | 0.1605           |          |
| X<sub>3</sub>X<sub>4</sub> | 99.91         | 1   | 99.91       | 85.59   | < 0.0001         |          |
| X<sub>1</sub>X<sub>2</sub>X<sub>3</sub> | 56.43         | 1   | 56.43       | 48.34   | < 0.0001         |          |
| X<sub>1</sub>X<sub>2</sub>X<sub>4</sub> | 50.63         | 1   | 50.63       | 35.84   | < 0.0001         |          |
| X<sub>1</sub>X<sub>3</sub>X<sub>4</sub> | 0.18          | 1   | 0.18        | 0.12    | 0.7287           |          |
| X<sub>2</sub>X<sub>3</sub>X<sub>4</sub> | 7.92          | 1   | 7.92        | 5.61    | 0.0317           |          |
| Residual | 17.51         | 15  | 1.17        |         |                  |          |
| Lack of Fit | 15.80         | 10  | 1.58        | 4.61    | 0.0527           | Not significant |
| Pure Error | 1.71          | 5   | 0.34        |         |                  |          |
| Cor Total | 690.13        | 29  |             |         |                  |          |

R-square = 0.9746 ; Adjusted R-square = 0.9510

Table 4. Analysis of variance (ANOVA) for conversion response of biodiesel production.

| Source   | Sum of Squares | df  | Mean Square | F Value | p-value Prob > F | Comments |
|----------|----------------|-----|-------------|---------|------------------|----------|
| Model    | 1201.66        | 14  | 85.83       | 60.76   | < 0.0001         | significant |
| X<sub>1</sub> | 44.66         | 1   | 44.66       | 31.62   | < 0.0001         |          |
| X<sub>2</sub> | 41.45         | 1   | 41.45       | 29.34   | < 0.0001         |          |
| X<sub>3</sub> | 9.83          | 1   | 9.83        | 6.96    | 0.0186           |          |
| X<sub>4</sub> | 50.63         | 1   | 50.63       | 35.84   | < 0.0001         |          |
| X<sub>1</sub>X<sub>2</sub> | 0.18          | 1   | 0.18        | 0.12    | 0.7287           |          |
| X<sub>1</sub>X<sub>3</sub> | 7.92          | 1   | 7.92        | 5.61    | 0.0317           |          |
| X<sub>1</sub>X<sub>4</sub> | 0.64          | 1   | 0.64        | 0.45    | 0.5111           |          |
| X<sub>2</sub>X<sub>3</sub> | 0.57          | 1   | 0.57        | 0.40    | 0.5349           |          |
| X<sub>2</sub>X<sub>4</sub> | 0.000         | 1   | 0.000       | 0.000   | 1.0000           |          |
| X<sub>3</sub>X<sub>4</sub> | 0.15          | 1   | 0.15        | 0.10    | 0.7505           |          |
| X<sub>1</sub>X<sub>2</sub>X<sub>3</sub> | 276.21        | 1   | 276.21      | 195.52  | < 0.0001         |          |
| X<sub>1</sub>X<sub>2</sub>X<sub>4</sub> | 442.89        | 1   | 442.89      | 313.52  | < 0.0001         |          |
| X<sub>1</sub>X<sub>3</sub>X<sub>4</sub> | 556.25        | 1   | 556.25      | 393.76  | < 0.0001         |          |
| X<sub>1</sub>X<sub>2</sub>X<sub>3</sub>X<sub>4</sub> | 176.38        | 1   | 176.38      | 124.86  | < 0.0001         |          |
| Residual | 21.19         | 15  | 1.41        |         |                  |          |
| Lack of Fit | 15.58         | 10  | 1.56        | 1.39    | 0.3772           | Not significant |
| Pure Error | 5.61          | 5   | 1.12        |         |                  |          |
| Cor Total | 1222.58       | 29  |             |         |                  |          |

R-square = 0.9872 ; Adjusted R-square = 0.9665

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Based on the Equations (5) and (6), the effect of particular factor is respectively shown by the individual factor ($X_1$, $X_2$, $X_3$ and $X_4$) meanwhile the interaction between two factors is presented by the quadratic effects of the two factors ($X_1X_2$, $X_1X_3$, $X_1X_4$, $X_2X_3$, $X_2X_4$, $X_3X_4$, $X_1^2$, $X_2^2$, $X_3^2$, and $X_4^2$). The positives and negatives sign for each codes represent the parallel effects and adverse effects towards the responses respectively. Highest orders of polynomials was selected for this model indicates that the models are significant.

Regression model equation generated from the RSM is capable to analyze the adequacy of the model. The adequacy of the model indicates the interactions between the variables to the responses. The determination coefficient which also known as $R^2$ value is the most important factor to determine the appropriateness of the model adequacy through the revelation of the total variation of the mean observed of the system [12]. Table 3 and Table 4 show the sequential model sum of squares for the experimental design for both yield and conversion responses. The tables tabulate $F$-value, the significance of the model, the lack of fit as well as the $R^2$ value for responses yield as well as conversion.

The $R^2$ value for $Y$ and $C$ were 0.9746 and 0.9827, respectively indicated that high precision as well as the adequacy of the model developed. High value of $R^2$ describes a good validity of the models generated. Both of the $R^2$ values indicate 97.46% and 98.27% from the sample were attributed to the variables meanwhile the remaining 2.54% and 1.73% of $Y$ and $C$ respectively could not be explained by the regression model. The close agreement between the $R^2$ value and adjusted $R^2$ value for both $Y$ and $C$ describes the related variables of the model. The percentage of coefficient of variation (C.V.) were 1.26% and 1.49% for both $Y$ and $C$ response indicates a high reliability between fitted model and the experimental results [13].

Evaluation of the variance or the regression model’s adequacy can be obtained by using ANOVA. ANOVA is a statistical method to determine the differences between two or more means and evaluating the important effect between all variables and responses in order to fit with the second order polynomials model developed along with the experimental value. $F$ value is the most important value to be observed. $F$-value indicates the ratio of mean square based on the regression to the mean square to residual error. Significant $F$-value is obtained when the value is higher than the model’s tabulated value. From Table 3 and 4, $F$-value for both $Y$ and $C$ were 41.16 and 60.76, respectively. 5.27% and 37.72% of chance of error were reported for both $Y$ and $C$ respectively. The low value of the chance of error indicating that it is non-significant toward the model where it is a good sign since the model is required to be fit with a second order polynomial model.

The variables or parameters used for $Y$ and $C$ responses can be significantly observed via ANOVA. For $Y$ response, parameter of volume of $H_2SO_4$ is found to be the most significant or provide high influence on yield production. The $F$-value for this parameter is 48.34 with $p$-value less than 0.001. For $C$ on the other hand, it is found that stirring rate gives the highest $F$-value with 35.84 with $p$-value less than 0.001 to significantly affect the conversion percentage in biodiesel production. Due to the polar and nonpolar nature of sulfuric acid and $D(+)$-Glucose, the sulfonation suffers initial

**Figure 1.** Actual versus predicted value for biodiesel yield.

**Figure 2.** Actual versus predicted value for biodiesel conversion.
mass transfer limitations problem, similar to the situation reported by Namdeo [14]. This mass transfer limitation problem can be avoided by applying stirring on the system [15]. Optimum stirring for sulfonation enhances the attachment of $\text{−SO}_3\text{H}$. Noureddini and Zhu proposed an initial mass transfer controlled region followed by a kinetically controlled region for base catalytic transesterification of sunflower oil [16]. Hou et al. showed that the reaction is very slow initially due to mass transfer limitations between methanol and oil phase [17].

The distinction between the experimental data as well as the predicted data for both Y and C are illustrated in Figures 1 and 2, respectively. Based on Figure 1, a slight deviation was shown from $Y_{\text{act}} = Y_{\text{pred}}$, where $Y_{\text{act}}$ is referring to experimental data meanwhile $Y_{\text{pred}}$ is the data predicted for $Y$. The slight deviation proves the relevancy of the generated model which appropriately aligned with the model generated. The predicted values lie near the plotted line which contributing to higher value of $R^2$. This is almost the same with Figure 2 where $C_{\text{act}} = C_{\text{pred}}$ where $C_{\text{act}}$ referring to the conversion from experimental value meanwhile $C_{\text{pred}}$ is the predicted value from the modelling. Figure 2 shows that the regression model developed is appropriate.

Residuals on the other hand is the difference between $Y_{\text{act}}$ and $Y_{\text{pred}}$ as well as $C_{\text{act}}$ and $C_{\text{pred}}$. Residuals are estimation of experimental error obtained by subtracting the observed responses from the predicted responses. Normal probability plot of the residuals is a method to learn whether it is reasonable to assume the error terms are normally distributed. Normal distributions are achieved when the residuals values fall onto the linear line. Based on Figures 3 and 4, the normal distribution were achieved and satisfied both yield and conversion where the line adjacent to the straight line.

The effect of each variables on the $Y$ and $C$ can be evaluated from the Pareto chart. Pareto charts visually represents the absolute values of the effects of main factors and the effects of interaction of factors. The chart includes a ver-

![Figure 3](image1.png)  
**Figure 3.** Normal plot of residuals for biodiesel yield.

![Figure 4](image2.png)  
**Figure 4.** Normal plot of residuals for biodiesel conversion.

![Figure 5](image3.png)  
**Figure 5.** Pareto chart for biodiesel yield.

![Figure 6](image4.png)  
**Figure 6.** Pareto chart for biodiesel conversion.
tical reference line at the critical t-value for an alpha of 0.05 (magnitude for a 95% confidence level) to indicate that the factors which extend past this line are potentially important. From Figure 5 and Figure 6, all effects are statistically significant.

The significance of the interaction between effects could also be compared from the Pareto charts. Significance positive effects was identified in the interaction between the time and power level. These plots clearly indicated that the significance of the variables and their interaction between effects towards Y and C decreased in this order, $X_2 > X_4 > X_1 > X_5$, and $X_t > X_2 > X_3 > X_5$, respectively. Kefas and colleagues also did an optimization for sulfonation of glucose catalyst [18]. Four independent variables used including heating time (A), concentration of ($NH_3)_2SO_4$ (B), Volume of H$_2$SO$_4$ (C) and Temperature (D). However, based on the optimization result, parameter B and A were found to be the most significant with lowest p-value (< 0.05). The sulfonation of ICG using ($NH_3)_2SO_4$ heated conventionally within 1 to 9 hour was proven to be effective for attachment of $-SO_4^{2-}$ to the catalyst as the surface of catalyst may saturated negatively impact the catalytic activity of catalyst. The highest surface area from the catalyst prepared using microwave-heating method. The highest surface area from the best performing catalyst, from our previous report [10], was recorded at 16.94 m$^2$/g with acid density of 5.92 mmol/g, lower from the catalyst prepared using microwave-heating method. The response between two factors can be determined via contour plots, 3D surface or cube surface in design expert software by observing the effect of interaction of the response from the two variables. Figures 7 to 18 show the contour plots for every possible interaction between two independent parameters while the other two variables is kept at the center of their respective range.

Figures 7 and 8 portrays contour plots for interaction between volume of H$_2$SO$_4$ and time (min) where both power level (W) and stirring rate (rpm) were constant at 400 W and 600 rpm, respectively. The increment of Y and C were observed as the H$_2$SO$_4$ volume increased right until it slightly reduce after 163 ml and 160 mL of H$_2$SO$_4$ used. The presence of large quantity of H$_2$SO$_4$ increase the number of acid active sites available for reaction to occur, but excessive H$_2$SO$_4$ negatively impact Y value possibly due to great structural destruction that lead to less surface for $-SO_4H$ group attachment. Similar case is observed for heating time, Y increased with the increase of heating time. The 7 to 8 minutes is found to be the optimum value for heating time of catalyst in microwave to obtained maximum Y. Prolonged the catalyst heating time more than 8 minute negatively impact the catalytic activity of catalyst as the surface of catalyst may saturated with the $-SO_4H$ group that will hinder methanol react with the active sites [19]. The contour plot is portrayed in oval or elliptical shape which is proven by Erbay and colleagues as a perfect interaction between the independent variables [20].

3.2 Response Surface Contour Plot

The response between two factors can be determined via contour plots, 3D surface or cube surface in design expert software by observing the effect of interaction of the response from the two variables. Figures 7 to 18 show the contour plots for every possible interaction between two independent parameters while the other two variables is kept at the center of their respective range.

Figures 7 and 8 portrays contour plots for interaction between volume of H$_2$SO$_4$ and time (min) where both power level (W) and stirring rate (rpm) were constant at 400 W and 600 rpm, respectively. The increment of Y and C were observed as the H$_2$SO$_4$ volume increased right until it slightly reduce after 163 ml and 160 mL of H$_2$SO$_4$ used. The presence of large quantity of H$_2$SO$_4$ increase the number of acid active sites available for reaction to occur, but excessive H$_2$SO$_4$ negatively impact Y value possibly due to great structural destruction that lead to less surface for $-SO_4H$ group attachment. Similar case is observed for heating time, Y increased with the increase of heating time. The 7 to 8 minutes is found to be the optimum value for heating time of catalyst in microwave to obtained maximum Y. Prolonged the catalyst heating time more than 8 minute negatively impact the catalytic activity of catalyst as the surface of catalyst may saturated with the $-SO_4H$ group that will hinder methanol react with the active sites [19]. The contour plot is portrayed in oval or elliptical shape which is proven by Erbay and colleagues as a perfect interaction between the independent variables [20].
Figures 9 and 10 depict the contour plots of time (min) against power level (W) of microwave to obtain an optimum Y and C, respectively. The 150 mL volume of H₂SO₄ and 600 rpm stirring rate were fixed at this rate. Percentage Y and C were increased as heating time increased with 7.65 min and 7.49 min of optimum heating time (min) recorded based on the graph, respectively. As heating time increased from 5 min to 7 min, the yield and conversion percentage of the biodiesel increased indicating a pretty fast sulfonation process. The peak value appears around 7.65 min and 7 min for both Y and C, over which further prolonging the sulfonation makes no obvious sense. Theoretically, a sulfonation process requires certain duration to be connected with sulfonic acid groups from concentrated sulfuric acid to the carbon frames. Once the process is done, further prolonging the duration makes no sense. The same occurs with the power level (W) as Y and C amplified with the increment of the power level and started to reduce when the power level reached 400 W. A high power level hinder the −SO₃H group incorporated into the carbon network of catalyst [21]. The optimum range of the power level is located between 320 W to 480 W.

Figures 11 and 12 show the effect and interaction of time and stirring rate for both Y and C. From the contour plot of both graph, the value of Y and C continue to increase with the increment in time of catalyst heating. The interaction between Y or C and the stirring rate also
demonstrate the same pattern in which any increasing in stirring rate value results in inclination of $Y$ or $C$. It can be observed that to get maximum value of responses, the optimum value of heating time for catalyst as well as stirring rate is 718.92 rpm and 7.65 min for $Y$ and 659.46 rpm with 7.49 min for $C$. This is practically due to the mass transfer limitation problem as discussed previously. Optimum stirring rate for sulfonation induced the probability of attachment of $\text{SO}_3\text{H}$.

As for effect and interactions between volume of $\text{H}_2\text{SO}_4$ and power level, a plotted contour plot was presented in Figure 13 and Figure 14. The graph shows an increment in value of $Y$ and $C$ as the volume of $\text{H}_2\text{SO}_4$ used to sulfonate the catalyst increase. However, the $Y$ value started to decrease after the usage of $\text{H}_2\text{SO}_4$ volume reached 160 mL. The same pattern of response is also observed through the interactions between power levels with $Y$ and $C$. With the increasing in value of power level parameter, $Y$ and $C$ also increase, before started to decrease at power level value of 420 W. The highest temperature from the highest power level might lead to an excessive breakdown of the macromolecules and removal of organic portions [22].

While keeping the time for catalyst heating 7 minutes and power level 400 W, the effect and interactions between volume of $\text{H}_2\text{SO}_4$ used and stirring rate with $Y$ is demonstrated in Figure 15 and Figure 16. From the graph, it can be seen that the value of $Y$ increase and decrease as the volume of $\text{H}_2\text{SO}_4$ used to sulfonate the glucose is increased, with optimum value of $Y$ is obtained at 160 mL. As for the effects and interactions of stirring rate to $Y$ val-

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**Figure 13.** Yield contour plot volume $\text{H}_2\text{SO}_4$ (mL) against power level (W) for catalyst heated at 7 min and stirred at 600 rpm.

**Figure 14.** Conversion contour plot volume $\text{H}_2\text{SO}_4$ (mL) against power level (W) for catalyst heated for 7 min and stirred at 600 rpm.

**Figure 15.** Yield contour plot volume $\text{H}_2\text{SO}_4$ (mL) against stirring rate (rpm) for catalyst heated at 400 W for 7 min.

**Figure 16.** Conversion contour plot volume $\text{H}_2\text{SO}_4$ (mL) against stirring rate (rpm) for catalyst heated at 400 W for 7 min.
ue, the increment of stirring rate causes $Y$ value to increase up until 700 rpm before the pattern started to decrease. Efficient stirring rate with sufficient volume of $H_2SO_4$ effectively reduced the initial mass transfer limitation caused by the polar and non-polar of $H_2SO_4$ and D(+)glucose structure. Thus, optimum stirring rate enhanced the attachment of $-SO_3H$ group to the D(+)glucose surface yet increase the $Y$ and $C$ value. High stirring rate was not applicable due to the spillage of the solution thus reducing the volume of the $H_2SO_4$.

The effect and interaction of power level and stirring rate towards $Y$ were also studied through the contour plot, as illustrated in Figure 17 and Figure 18. It can be seen that $Y$ and $C$ shows an increasing pattern as the power level increase up to 400 W before it started declining. The microwave power level should not be too high, as it might cause a sharp rise in reaction temperature, resulting in more accumulated discontinuities and shortened molecular reorientation and vibration times. In contrast, if microwave output was too low, most of the microwave power would be used to warm up the reaction mixture, and the temperature would be lower than required in most cases, resulting in a lower yield. The appropriate microwave power density and its dissipation control were very important to maximize ester yield and minimize reaction time and microwave energy consumption [23].

Optimum value is the ideal value of the variables obtained when the highest $Y$ and $C$ is achieved. The optimum value of $Y$ and $C$ is tabulated in Table 5. The highest yield and conversion can be achieved is 94.01% and 91.89% respectively at optimum operating condition of

![Figure 17. Yield Contour Plot power level (W) against stirring rate (rpm) for catalyst heated for 7 min using 150 ml $H_2SO_4$.](image)

![Figure 18. Conversion Contour Plot power level (W) against stirring rate (rpm) for catalyst heated for 7 min and using 150 ml of volume $H_2SO_4$.](image)

### Table 5. Optimum operating condition to produce maximum $Y$.

| Parameters | Time (min) | Volume of $H_2SO_4$ used (mL) | Power level (W) | Stirring rate (rpm) | Max $Y$ (%) | Max $C$ (%) |
|------------|------------|-------------------------------|-----------------|---------------------|-------------|-------------|
|            | 7.53       | 159.51                        | 413.64          | 670.53              | 94.01       | 91.89       |

### Table 6. Validation test for yield using optimum value from optimization.

| No | Time (min) | Volume of $H_2SO_4$ (ml) | Power Level (W) | Stirring rate (rpm) | Yield (%) | Conv (%)  |
|----|------------|----------------------------|-----------------|---------------------|-----------|-----------|
| 1  | 7.5        | 160                        | 400             | 700                 | 92.06     | 89.09     |
| 2  | 7.5        | 160                        | 400             | 700                 | 92.89     | 88.74     |
| 3  | 7.5        | 160                        | 400             | 700                 | 92.06     | 89.09     |
|    |            |                             |                 |                     | **91.75** | **88.43** |

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7.5 min for time of heating catalyst, 159.5 mL of H₂SO₄, 413.6 W of power level, and stirring rate of 671 rpm.

The optimum result was then validated by conducting confirmatory experiment and the results are presented in Table 6. There are some limitation for validation test for optimum value. 3 replicate experiments had been done to validate the optimum value however, the experimental work was done using power level of 400 W and stirring rate at 700 rpm, since the automated function of the equipment cannot be set to the respective value. The average of the validation test was a bit off from the predicted optimum value since the variables such as microwave power level and stirring rate is according to analogue system and cannot be modified with the respective optimum value. The validation test obtained 91.75% and 88.43% for Y and C, respectively.

4. Conclusions

Response surface methodology was used optimized and study the effect of the sulfonation process of synthesizing SO₃H/ICG acid catalyst via microwave heating method. Quadratic model was proposed for prediction of biodiesel yield (Y) from palm fatty acid distillate (PFAD) and its conversion (C). The application of DOE successfully optimized the operating conditions for the two-step SO₃H/ICG catalyst synthesis to be used for the esterification process. The optimized conditions of the best performing SO₃H/ICG with maximum Y and C were at 7.53 minutes of reaction time, 159.5 mL of H₂SO₄ used, 671 rpm of stirring rate as well as 413.6 watt of power level. At these optimum conditions the predicted yield percentage and conversion percentage were 94.01% and 91.89%. The utilization of sulfonated glucose solid acid catalyst via microwave-assisted in biodiesel production has great potential towards sustainable and green method of synthesizing catalyst for biodiesel. The preparation of the catalyst is much convenient in time and energy consumption.

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