Supporting Information

for

Gas sorption porosimetry for the evaluation of hard carbons as anodes for Li- and Na-ion batteries

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Additional figures
**Figure S1:** Publication trend since the year 2000 on carbons for SIBs according to Web of Science.

**Figure S2:** Powder X-ray diffraction pattern of HT2. \( d_{002} = 4.02 \text{ Å} \) (\( 2\theta = 22.1^\circ \)), \( d_{101} = 2.06 \text{ Å} \) (\( 2\theta = 43.9^\circ \)).
**Figure S3:** Powder X-ray diffraction patterns of (a) RF-1000 and (b) RF-1600. (a) $d_{002} = 4.28$ Å ($2\theta = 20.8^\circ$), $d_{101} = 2.09$ Å ($2\theta = 43.2^\circ$), (b) $d_{002} = 3.94$ Å ($2\theta = 22.5^\circ$), $d_{101} = 2.06$ Å ($2\theta = 44.0^\circ$).
Figure S4: Raman spectrum of HT2. $I_D/I_G = 0.98$. 
Figure S5: Raman spectra of (a) RF-1000 and (b) RF-1600. (a) $I_D/I_G = 1.12$, (b) $I_D/I_G = 1.29$. c) Micropore size distributions of the carbons calcined at different temperatures obtained from the corresponding N\textsubscript{2} isotherms by the Horvath–Kawazoe (HK) method. The carbons treated above 1600 °C possessed negligible micropores.

Figure S6: SEM images of RF carbons. (a) RF-1000, (b) RF-1600.
**Figure S7:** Failed N\textsubscript{2} sorption isotherms of HT1.

**Figure S8:** Charge-discharge curves of HT2 vs (a) Li and (b) Na metal electrode.