VESIcal Part II: A critical approach to volatile solubility modelling using an open-source Python3 engine

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VESIcal Part II: A critical approach to volatile solubility modelling using an open-source Python3 engine

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Key Points:
- The Python3 tool VESIcal allows extensive comparisons to be drawn between different H$_2$O-CO$_2$ solubility models
- Solubility models are not interchangeable; for a single magma composition, different models can predict a wide range of solubility relationships
- The P-T-X calibration range of each solubility model must be critically evaluated before application to a specific volcanic system

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Abstract

Accurate models of H$_2$O and CO$_2$ solubility in silicate melts are vital for understanding volcanic plumbing systems. These models are used to estimate the depths of magma storage regions from melt inclusion volatile contents, investigate the role of volatile exsolution as a driver of volcanic eruptions, and track the degassing paths followed by magma ascending to the surface. However, despite the large increase in the number of experimental constraints over the last two decades, many recent studies still utilize the earlier generation of models, which were calibrated on experimental datasets with restricted compositional ranges. This may be because many of the available tools for more recent models require large numbers of input parameters to be hand-typed (e.g., temperature, concentrations of H$_2$O, CO$_2$, and 8–14 oxides), making them difficult to implement on large datasets. Here, we use a new open-source Python3 tool, VESIcal, to critically evaluate the behaviours and sensitivities of different solubility models for a range of melt compositions. Using literature datasets of andesitic-dacitic experimental products and melt inclusions as case studies, we illustrate the importance of evaluating the calibration dataset of each model. Finally, we highlight the limitations of particular data presentation methods such as isobar diagrams, and provide suggestions for alternatives, and best practices regarding the presentation and archiving of data. This review will aid the selection of the most applicable solubility model for different melt compositions, and identifies areas where additional experimental constraints are required. (242/250 words)

Plain Language Summary

Being able to accurately model the solubility of H$_2$O and CO$_2$ in magmas is very important for understanding a wide variety of volcanic processes, such as the depths at which magma is stored in the crust, the driving force behind volcanic eruptions, and the release of volatile elements into the atmosphere. However, there has been no easy way for volcanologists to perform calculations on large datasets, or to compare different models. This review uses a new, open-source tool called VESIcal written in the popular programming language Python3. This allows us to compare different models for a wide variety of melt compositions, temperatures and pressures, helping researchers to identify the most suitable model for their study. We also suggest areas where further experimental constraints are required. Finally, we highlight the limitations of particular data presentation methods such as isobar diagrams, and provide suggestions for alternatives, and best practices regarding the presentation and archiving of data.
1 Introduction

The most abundant volatile components found in terrestrial magmatic systems are H$_2$O and CO$_2$. It has been known for nearly a century (Bowen, 1928; Tuttle & Bowen, 1958) that these volatile species have profound effects on the chemical and material properties of magmas (e.g., phase equilibria, melting temperatures, magma viscosity and density; Burnham, 1979; Husen et al., 2016; Burnham & Davis, 1974; Hess & Dingwell, 1996; Ochs & Lange, 1999), so significantly affect their geochemical and dynamical behavior (e.g., eruption and degassing style, erupted volume; Papale et al., 1999; Huppert & Woods, 2002). Thus, it is vital to be able to predict how H$_2$O and CO$_2$ solubilities change as a function of intensive variables such as pressure, temperature, melt and fluid composition in order to understand plutonic and volcanic systems.

The solubility of a volatile species is defined at a given pressure and temperature as the maximum concentration that can be dissolved within a silicate melt of a specified composition. Ignoring disequilibrium effects, if the volatile content of the system exceeds this solubility limit, a separate fluid/vapour phase will exsolve from the magma. In this review, we favour the term fluid because of the supercritical nature of exsolved volatile phases at magmatic temperatures. In general terms, a magma is described as volatile undersaturated when there is no fluid phase, and volatile saturated once a fluid phase is present (also referred to as vapour undersaturated/saturated, or fluid undersaturated/saturated). In detail, different volatile species do not act as independent entities, but influence one another. For this reason, a magma may exsolve a mixed CO$_2$-H$_2$O fluid even if the dissolved concentrations of H$_2$O and CO$_2$ do not exceed the pure solubility limit of each species.

Despite the obvious importance of accurate volatile solubility modelling, very few studies of volcanic systems have evaluated results using several different solubility models to determine possible sources of systematic and random error, and assess the suitability of each model for the conditions of interest (e.g. temperature, pressure, and melt composition). This lack of intercomparison likely results from the fact that it is extremely time consuming to perform the large numbers of calculations necessary for thorough comparisons using available tools. For example, many solubility models were released as stand-alone Excel spreadsheets (e.g., G. Moore et al., 1998; Newman & Lowenstern, 2002; Allison et al., 2019) or web apps (e.g., Iacono-Marziano et al., 2012; Ghiorso & Gualda, 2015), where saturation pressures, dissolved volatile contents, degassing paths, and isobars can only be calculated for one sample and set of conditions at a time. The more recent models which include several terms accounting for the effect of melt composition on volatile solubility require users to hand-type a large number of input parameters. For example, to calculate a saturation pressure in MagmaSat (Ghiorso & Gualda, 2015), users must hand-type 9–14 oxide concentrations in addition to entering H$_2$O and
CO₂ concentrations, and a melt temperature. Similarly, the web app of Iacono-Marziano et al. (2012) requires users to input 8 major element oxide concentrations. Calculating isobars using these web apps is a particularly daunting task, as users must evaluate dissolved volatile contents at multiple fluid compositions, and then use curve fitting to produce a smooth isobar to display on plots. Other models were released with no calculator at all, requiring each user to correctly interpret and combine the relevant equations in the manuscript (Dixon, 1997; Shishkina et al., 2014).

Here, we take advantage of the recent release of VESIcal (Volatile Equilibria and Saturation Identification calculator; Iacovino et al., 2021), an open-source tool written in Python3. VESIcal contains functions to calculate saturation pressures, dissolved volatile contents, isobars, and degassing paths automatically for seven different models. Calculations can be performed based on melt compositions provided in an Excel spreadsheet, and users can take full advantage of Python’s extensive flexibility to perform large numbers of calculations automatically (e.g., creating for loops to perform calculations across a range of pressures, temperatures, and fluid compositions). To our knowledge, the only other model with similar functionality to VESIcal is the Linux program Solwcad supplied by Papale et al. (2006), which performs calculations automatically on a user-supplied .txt file containing melt compositions, pressures and temperatures (http://www.pi.ingv.it/progetti/eurovolc/). Solwcad was used alongside VESIcal in this review, through the Windows Subsystem for Linux (WSL2).

The overall aim of this review is to summarize the formulation, strengths and weaknesses of popular solubility models to inform users who wish to model volatile solubility in silicate melts, whether that be the calculation of melt inclusion saturation pressures, degassing paths, incorporating volatile exsolution in physical model of magma chambers (e.g. Huber et al., 2019), or calculating the dissolved volatile contents of experimental products where the pressure, temperature and fluid composition are known (e.g. Waters & Lange, 2015). We start by briefly summarizing the major results from volatile solubility experiments over the last century (section 2), before describing nine of the most popular solubility models (section 3). We then compare the solubility of pure H₂O, mixed H₂O-CO₂, and pure CO₂ predicted by different models for representative mafic and silicic compositions (section 4). We also explore the sensitivity of these models to parameters such as temperature and redox state, which are often poorly constrained in igneous systems (section 5). Finally, we evaluate the suitability of these models for intermediate melt compositions, where experimental constraints are sparse relative to basaltic and rhyolitic melts (section 6). We conclude by discussing best practices for presenting and archiving data related to volatile components in igneous systems (section 7). The aim of this review is to give readers an insight into the strengths and weaknesses of different models, as well as the large systematic errors associated with model choice. This will help users of VESIcal and other tools to
select an appropriate model for their specific application. All the Python code used to perform these comparisons is provided in the supporting material, which users can easily adapt to evaluate melt compositions for their specific application.

2 Major findings from experimental studies investigating volatile solubility in magmas

One of the earliest volatile studies was that of Goranson (1931), who investigated the effect of pressure on the solubility of water in granitic melts. The classic treatise of Tuttle and Bowen (1958) investigated the impact of H₂O on mineral phase equilibrium. This study led to a wider recognition of the importance of volatiles, and motivated the development of both the experimental and analytical approaches to determining volatile solubilities as a function of pressure, temperature, and melt composition. Hamilton et al. (1964) was one of the first to compare H₂O solubilities for differing melt compositions (basalt and andesite), while also investigating the effect of dissolved H₂O and oxidation state on the magmatic phase equilibria. These studies were followed by the fundamental experimental measurements of the Burnham group on the dissolution of H₂O in albite melts (e.g., Burnham & Davis, 1971, 1974).

Further investigation of volatile solubility over the next four decades in natural samples and experimental products was aided significantly by analytical developments, allowing volatile contents in quenched glasses to be measured by techniques such as Fourier Transform infra-red spectroscopy (FTIR; e.g., Stolper, 1982; Fine & Stolper, 1986; Silver et al., 1990) and secondary ion mass spectrometry (SIMS; Hervig & Williams, 1988; Hauri, 2002). In particular, the high spatial resolution of FTIR and SIMS (a few tens of micrometers) meant that volatile concentrations could be measured within quenched pockets of melt trapped within crystals (termed melt inclusions). Unlike subaerially-erupted lavas which have degassed almost all their H₂O and CO₂ following their ascent to shallow pressures, melt inclusions remain pressurized during ascent as they are trapped in relatively incompatible crystals, so retain high volatile contents.

Melt inclusion analyses have greatly advanced our understanding of the behavior of volatiles in volcanic systems (Lowenstern, 2003; Hauri et al., 2002; Roggensack, 2001; Wallace et al., 1995; Métrich & Wallace, 2008; Sides et al., 2014a). For example, melt inclusions provide insights into pre-eruptive volatile contents (e.g., Saal et al., 2002; Hervig et al., 1989), and links between melt volatile contents and eruption styles (Lucic et al., 2016). The strong pressure-dependence on volatile solubility means that H₂O and CO₂ contents within melt inclusions trapped from a volatile-saturated magma can be used to determine the pressure at which the inclusion was trapped (termed the saturation pressure or entrapment pressure). In turn, the distribution of saturation pressures in a suite of melt inclusions can
reveal the locations of the main regions of magma storage in a volcanic system. This explosion of new information from melt inclusions greatly increased the demand for flexible and accurate solubilities models that could be applied to a broad range of pressures, temperatures, and melt/fluid compositions (G. Moore, 2008; Ghiorso & Gualda, 2015).

It has become increasingly apparent from solubility experiments that while the solubility of H$_2$O is relatively insensitive to melt composition (e.g., G. Moore & Carmichael, 1998; Shishkina et al., 2010), CO$_2$ solubility is highly sensitive to melt composition, particularly in mafic melts where the carbonate ion is the dominant species (Dixon, 1997; Brooker et al., 2001a; Shishkina et al., 2010; Iacono-Marziano et al., 2012; Shishkina et al., 2014; Allison et al., 2019). This has led to a great diversity in the way that various models treat the dependence of CO$_2$ solubility on melt composition. In general, models have become more complex with time as the region of compositional space spanned by solubility experiments has increased to include more alkaline lavas. The individual role and relative importance of each cation species in the melt is still associated with a large degree of uncertainty (Allison et al., 2019), accounting for the larger discrepancies between different model predictions for CO$_2$ vs. H$_2$O.

Experimental work has also highlighted the complexities of mixing between H$_2$O and CO$_2$ in igneous systems. In the simplest case, the addition of one component in a melt-fluid system decreases the activity, and therefore the solubility of the other component in the melt (Lowenstern, 2001). This behavior is referred to as Henrian/ideal behavior. Henry’s Law states that the amount of a volatile dissolved in a liquid is proportional to its partial pressure in the gas phase in equilibrium with that liquid. Neglecting the possible entropic effects of speciation, the addition of H$_2$O to the fluid/gas phase acts to lower the partial pressure of CO$_2$, and therefore lowers the solubility of CO$_2$ in the liquid. Similarly, addition of CO$_2$ to the fluid/gas phase causes the solubility of H$_2$O in the melt to decrease.

Experimental studies have shown that Henry’s law is generally obeyed at low pressures (<1 kbar) in basaltic (Dixon et al., 1995) and rhyolitic melts (Blank et al., 1993). However, at higher pressures, some experimental observations have shown that the mixing behavior of CO$_2$ and H$_2$O becomes strongly non-Henrian (Papale, 1999). For example, Eggler (1973), Mysen et al. (1976) and Mysen (1976) show that the solubility of CO$_2$ in albite melt increases with the addition of H$_2$O at higher pressures. This has been attributed to the fact that the addition of small amounts of water as OH$^-\$ species decreases melt polymerization, and therefore enhances the solubility of CO$_2$ relative to an entirely anhydrous melt. After a certain amount of H$_2$O is added, solubility decreases once more because the addition of H$_2$O to the fluid phase causes the fugacity of CO$_2$ to decrease (Mysen, 1976; Dingwell,
1986; King & Holloway, 2002). More recently, this behavior has been demonstrated for dacitic and
rhyolitic melts by Behrens et al. (2004) and Liu et al. (2005).

The non-ideal behavior of H$_2$O and CO$_2$ in basaltic and andesitic melts at higher pressures is
less well constrained. King and Holloway (2002) show that at 1 kbar, andesitic melts (SiO$_2$=58.4 wt%)
exhibit a sharp increase in the solubility of CO$_2^-$ species with increasing melt H$_2$O contents between
0–3.39 wt%. In contrast, Jakobsson (1997) show that the solubility of CO$_2$ in an icelanditic melt (54.6
wt% SiO$_2$) at 10 kbar is essentially constant, despite variation in melt water contents between ~1–9
wt%. Similar behaviour to that observed by Jakobsson (1997) was noted for andesitic magmas (57.4
wt% SiO$_2$) at 5 kbar by Botcharnikov et al. (2006), who suggest that differences between H$_2$O-CO$_2$
mixing in their experiments and those of King and Holloway (2002) may result from differences in
oxygen fugacity of these experiments (Fe$^{3+}$/Fe$^T$=0.2–0.6 vs. Fe$^{3+}$/Fe$^T$=0.09–0.2). However, as we
discuss in section 4.1, the influence of oxygen fugacity on volatile solubility is still poorly constrained.
Recent basaltic H$_2$O-CO$_2$ experiments generally show a relatively flat plateau for CO$_2$ solubility with
increasing melt H$_2$O contents between ~0–4 wt% (Iacono-Marziano et al., 2012; Shishkina et al., 2010,
2014), although relatively large error bars on analyses of CO$_2$ in experimental products make it diffi-
cult to determine whether this plateau is truly flat, or shows a slight positive or negative gradient at
low H$_2$O contents (see Fig. 7a of Iacono-Marziano et al., 2012). The effect of H$_2$O-CO$_2$ mixing in the
9 solubility models evaluated here is discussed further in sections 4.2.2 and 4.3.2.

3 Models discussed in this review

In this review, we focus on the seven models implemented in VESIcal (Iacovino et al., 2021):

1. The simplified parameterization of the Dixon (1997) model for H$_2$O and CO$_2$, which was im-
plemented in the popular Excel tool VolatileCalc (Newman & Lowenstern, 2002), hereafter
VolatileCalc-Basalt.

2. The H$_2$O model of G. Moore et al. (1998), hereafter M–1998.

3. The H$_2$O and CO$_2$ models of Liu et al. (2005), hereafter L–2005.

4. The H$_2$O and CO$_2$ models of Iacono-Marziano et al. (2012), hereafter IM–2012.

5. The H$_2$O and CO$_2$ models of Shishkina et al. (2014), hereafter S–2014.

6. The combined H$_2$O and CO$_2$ model of Ghiorso and Gualda (2015), hereafter MagmaSat.

7. The CO$_2$ models of Allison et al. (2019), hereafter A–2019.

We also consider the two additional models, reflecting both their popularity and relative ease of calcu-
lation using previously published tools:
8. The combined H$_2$O and CO$_2$ model of Papale et al. (2006), hereafter P–2006, accessed using the
Linux program solwcad.

9. The Rhyolite functions in the VolatileCalc spreadsheet, hereafter VolatileCalc-Rhyolite

We do not consider the models of X. Duan (2014), Eguchi and Dasgupta (2018), or Burgisser et al.
(2015) because no tool exists to automate the necessary calculations. We also do not discuss models
with more limited pressure (P), temperature (T) or compositional (X) ranges.

A summary of the P, T and X range covered by the calibration dataset of each of the 9 models
evaluated, as well as available tools to perform calculations, is provided in Figure 1. The calibration
dataset of each model is shown on a total alkali-silica diagram in Figure 2. Detailed descriptions of
each model are provided below.
Summary of the calibration range of each model, as well as available tools to perform calculations.

| Publication            | Volatile Speciation | P (bar) | T (˚C) | Compositional range                                                                 | Formulation                                    | Redox sensitive? | Available Tools                                                                                                                                 | Notes |
|------------------------|---------------------|---------|--------|-------------------------------------------------------------------------------------|-----------------------------------------------|------------------|-------------------------------------------------------------------------------------------------------------------------------------------------|-------|
| Moore et al., 1998     | H2O                 | 0–3000  | 700–1200 | Broad compositional range: subalcaline basalt—rhyolites, alkali-trachybasalt—andesites, foidites, phonolites | Thermodynamic basis, empirical effect of melt composition | No (for FeOt) | Python code (VESIcal), web application (VESIcal.co.uk), Linux spreadsheet (SOLWCAD, Iacovino et al. 2021) | Author-suggested calibration range: The calibration dataset spans 190 to 6000 bar and 900–1300 ˚C for the carbon model, and 1–5000 bar and 700–1200 ˚C for the water model. |
| Newman and Lowenstern (2002) VolatileCalc–Basalt | H2O, CO2              | 0–5000  | 600–1500 | Alkali basalts: 40–49 w% SiO2                                                       | Thermodynamic basis, empirical effect of melt composition | No (input SiO2 only)  | Python code (VESIcal), web application (VESIcal.co.uk)                                                                                       | Warning implemented in VolatileCalc (Newman and Lowenstern, 2002). Calibration range suggested by Leone et al. (2011). Calibration range suggested by Iacono–Marziano et al. (2012) Calibration temperature of O’Hara (1997). If normalized (not recommended), different proportions of FeO and FeOt will slightly change the normalized SiO2 content. |
| Newman and Lowenstern (2002) VolatileCalc–Rhyolite | H2O, CO2              | 0–5000  | 600–1500 | Rhyolites                                                                           | Thermodynamic basis, non-compositional term   | No (input FeO and FeOt) | Python code (VESIcal), web application (VESIcal.co.uk)                                                                                       | Calibration range warning implemented in VolatileCalc (Newman and Lowenstern, 2002). |
| Liu et al. 2005        | H2O, CO2             | 0–5000  | 700–1200 | Haplogranites and rhyolites                                                        | Empirical basis, no compositional term        | No (input FeO and FeOt) | Python code (VESIcal), web application (VESIcal.co.uk)                                                                                       | Author-suggested calibration range for the mixed melt model. The calibration dataset covers 750–2000 bar and 900–1500 ˚C for the carbon model, and 1–5000 bar and 700–1200 ˚C for the water model. |
| Papale et al. (2006)   | H2O, CO2             | 0–10,000 | ~630–1630 | Broad compositional range (SiO2=37–85, Na2O+K2O=20–40) for H2O, more limited for CO2 and CO2+H2O liquids. Poor coverage of intermediate compositions (SiO2=55–75 w%). | Fully thermodynamic formulation + effect of melt composition | Yes (input FeO and FeOt) | Web application (VESIcalc), internet page: http://www.pi.ingv.it/volcan (last accessed 2021)                                                                 | Interception of H2O–CO2 fields on Fig. 15 of Papale (1989). |
| Iacono–Marziano et al., 2012 | H2O, CO2              | 100–10000 | 1100–14000 (preferably 1200–13000) | Predominantly mafic compositions: subalcaline and alkaline basalts–andesites          | Thermodynamic basis, empirical effect of melt composition | No (input FeO and FeOt) | Python code (VESIcal), web application (VESIcal.co.uk)                                                                                       | Range of calibration dataset, as authors do not state a range. We note that the vast majority of experiments were conducted at <500 bar. Authors state that most experiments were conducted between 1300–1300 ˚C (bar range 1500–1400 ˚C). |
| Shekhina et al. 2014   | H2O                  | 0–5000  | 1050–1400 (preferably 1150–1250) | Mafic and intermediate compositions: Subalcaline basalts–basaltic andesites, alkali basalts–phonolites. SiO2=65 w%. | Fully empirical (formulation + effect of melt composition) | Only for CO2 (as its equation uses FeOt). Calibrated with FeOt=2 | Python code (VESIcal), web application (VESIcal.co.uk)                                                                                       | Although their empirical expressions are for pure fluids, they were mostly calibrated on mixed CO2–H2O experiments. Authors suggested range Note: This model contains a temperature term. |
| Magmasat (Ghiorso and Guilda, 2015) | H2O as OH, CO2 as CaCO3 and CO2 as CaCO3                   | 0–20,000 | 550–1420 | Very broad compositional range of natural silicate melt compositions: subalcaline picobases–rhyolites, wide variety of mafic–silicic alkali compositions | Fully thermodynamic formulation + effect of melt composition | Yes (input FeO and FeOt) | Web application (VESIcalc), internet page: http://www.pi.ingv.it/volcan (last accessed 2021)                                                                 | Ranges extracted from Fig. 33 of Ghiorso and Guilda, 2015. |
| Allison et al., 2019   | CO2                 | 0–7000  | 1200 (1000–1400) | Alkaline mafic magmas from 8 volcanic fields (San Francisco Volcanic Field, Sunset Crater, Eburus, Vuelvite, Elma, Stromboli). Separate model coefficients for each composition. | Thermodynamic basis, separate parameters for each of 6 melt compositions | No (input FeO and FeOt) | Python code (VESIcal), web application (VESIcal.co.uk)                                                                                       | Although the model is for pure CO2, it was calibrated on mixed CO2–H2O experiments. Authors suggested range. The calibration dataset spans 50–4545 bar, Sunset Crater 470–4898 bar, Eburus 308–4545 bar, Vuelvite 386–6175 bar, Elma 458–668 bar, Stromboli 524–6000 bar. Notes: All calculations and experiments were performed at 1300 ˚C. Authors suggest applicable between 1000–1400 ˚C. |
Figure 2. Total alkalis (Na$_2$O+K$_2$O) vs. SiO$_2$ (TAS) diagram showing the composition of melts in the calibration dataset of each model for a) pure H$_2$O and b) pure CO$_2$ and mixed CO$_2$-H$_2$O. Pure CO$_2$ and CO$_2$-H$_2$O experiments are combined because pure CO$_2$ experiments are often contaminated by variable amounts of H$_2$O due to exposure with the earth’s atmosphere, and the high mobility of H$^+$ through experimental apparatus (Mangan et al., 2021). As dissolved H$_2$O contents in glasses from pure CO$_2$ experiments are rarely reported, it is nontrivial to distinguish these from a mixed H$_2$O-CO$_2$ experiment. For P-2006 and MagmaSat, points were extracted from the TAS diagrams shown in these papers using Web Plot Digitizer (Rohatgi, 2017). For the other models, the calibration dataset is provided in the supplementary information of Iacovino et al. (2021). TAS plot drawn using Python code from Stevenson (2015).
3.1 VolatileCalc-Basalt: a simplification of Dixon (1997)

The Dixon (1997) model calculates the solubility of H$_2$O and CO$_2$ in basaltic silicate melts, combining thermodynamic expressions as a function of pressure and temperature described in Dixon et al. (1995) with empirical parameters from Dixon (1997) accounting for the effect of melt composition in terms of melt SiO$_2$ content. The thermodynamic expressions are originally from Fine and Stolper (1986) for carbon, and Silver and Stolper (1989) for water. The Dixon (1997) model considers the solubility of the carbonate ion (CO$_3^{2-}$) for CO$_2$, and both molecular water (H$_2$O$_{mol}$) and hydroxyl groups (OH$^-$) for H$_2$O.

The solubility of molecular H$_2$O is calculated using an adapted version of equation 3 of Dixon et al. (1995). In the original equation the $X_{mol}^{H_2O} (P_0, T_0)$ term, representing the mole fraction of molecular H$_2$O in equilibrium with fluid with a fugacity of water specified by $f_{H_2O}(P_0, T_0)$ at 1473.15 K and 1 bar, was fixed at $3.28 \times 10^{-5}$. To account for the effect of melt composition on H$_2$O solubility, Dixon (1997) parameterize this term as a function of melt SiO$_2$ content:

$$X_{mol}^{H_2O} (P_0, T_0) = 3.04 \times 10^{-5} + 1.29 \times 10^{-6}[\text{SiO}_2]_{\text{wt} \%}$$

This relationship was derived from experimental observations of Cocheo and Holloway (1993), and predicts that there is a ~30% increase in the solubility of H$_2$O with increasing SiO$_2$ between nephelinite (~40 wt% SiO$_2$) and tholeiite (~49 wt% SiO$_2$) melt compositions. The concentration of OH$^-$ is then calculated as a function of the mole fraction of molecular H$_2$O in the melt using the solution model of Silver and Stolper (1989) (see equation 4 of Dixon, 1997). Interestingly, this is the only model discussed here which considers more than one species for dissolved H$_2$O in the melt.

For CO$_2$ solubility, Dixon (1997) adapted the model of Dixon et al. (1995) to account for the effect of melt composition, based on observations from experiments that CO$_2$ solubility increases from tholeiitic (49 wt% SiO$_2$) to basanitic (46 wt% SiO$_2$) to leucitic (44.1 wt% SiO$_2$) melts at 1200°C, 1 kbar. A linear regression with CO$_2$ solubility was achieved using a composition parameter (II) expressed in terms of the cation fractions, $X_i$ (Dixon, 1997):

$$II = -6.50(X_{Si^{4+}} + X_{Al^{3+}}) + 20.17(X_{Ca^{2+}} + 0.8X_{K^+} + 0.7X_{Na^+} + 0.4X_{Mg^{2+}} + 0.4X_{Fe^{2+}})$$

However, based on the strong correlation between II and SiO$_2$ in a suite of lavas from the North Arch Volcanic Field, Dixon (1997) simplify the $X_{mol}^{CO_2}$ (P$_0$, T$_0$) term from equation 6 of Dixon et al. (1995) solely as a function of melt SiO$_2$ content:

$$X_{mol}^{CO_2} (P_0, T_0) = 8.70 \times 10^{-6} - 1.70 \times 10^{-7}[\text{SiO}_2]_{\text{wt} \%}$$
Where $X_{\text{CO}_2^-}^{\infty}$ ($P_0$, $T_0$) is the mole fraction of carbonate in equilibrium with fluid with a fugacity of carbon dioxide specified by $f_{\text{CO}_2}(P_0, T_0)$ at 1473.15 K and 1 bar. Fugacities are calculated using the Redlich-Kwong equation of state (Holloway, 1977), with the correction of Flowers (1979).

This simplified expression was designed to aid the investigations of volatile solubility in the suite of lavas from the North Arch, where it effectively captures the observed 5$x$ decrease in CO$_2$ solubility from 40 to 49 wt% SiO$_2$. However, this simplified parameterization became very widely used in a wide variety of tectonic settings following its implementation in the Excel-based tool VolatileCalc (Newman and Lowenstern, 2002). Here, we refer to this model as VolatileCalc-Basalt, to differentiate it from the full $\Pi$ parameterization of Dixon (1997).

The advantage of the $\Pi$-SiO$_2$ simplification is that users only have to input the concentration of one oxide component (melt SiO$_2$) in addition to melt temperature and melt volatile contents to calculate saturation pressures or degassing paths. The limited number of inputs required by this model meant that users can calculate saturation pressures for large numbers of melt inclusions relatively quickly compared to more recent models such as Iacono-Marziano et al. (2012) and MagmaSat (which require users to input 8–14 oxide concentrations).

However, extreme care must be taken when using this simplified model to calculate CO$_2$ solubility. Firstly, the North Arch lavas span SiO$_2$ contents of only 40–49 wt%. Because of the rapid drop in $\Pi$ with increasing SiO$_2$, extrapolation beyond 51.2 wt% SiO$_2$ yields a negative value for $X_{\text{CO}_2^-}^{\infty}$ ($P_0$, $T_0$), corresponding to a negative amount of dissolved CO$_2$. To avoid this issue, VolatileCalc-Basalt returns an error, and will not perform the calculation if users enter a SiO$_2$ content $>$49 wt%. Most studies extrapolate beyond this by simply entering SiO$_2$=$49$ wt% into VolatileCalc-Basalt if their melts have higher silica contents (e.g. Sides et al., 2014a, 2014b; Tucker et al., 2019), and this approach is implemented in VESIcal for consistency. Newman and Lowenstern (2002) suggest that this extrapolation will be “generally applicable for other basaltic rocks with $<$52 wt% SiO$_2$”. However, if a large proportion of a sample suite has SiO$_2$ contents with $>$49 wt% SiO$_2$, the simplified $\Pi$-SiO$_2$ parameterization treats all melts as if they have the same composition, neglecting variations in solubility that may exist within that suite (see Wieser et al., 2021). Additionally, even if samples have SiO$_2$ contents between 40–49 wt%, this simplification can yield spurious results for melts which do not follow the same trend in $\Pi$-SiO$_2$ space to the North Arch lavas (see section 4.2.3). Thus, we suggest that any users wishing to apply VolatileCalc-Basalt to their system first check whether their melt compositions lie close to the trend defined by the North Arch lavas using the Jupyter notebook provided in the supplementary information of Iacovino et al. (2021).
It is also worth noting that, because VolatileCalc-Basalt parameterizes the effect of melt composition in terms of the absolute concentration of SiO\textsubscript{2} (rather than other empirical models which use cation fractions), it is extremely sensitive to normalization. For example, consider the MORB2 composition in Table 1 which has a measured SiO\textsubscript{2} content of 47.4 wt\%, and an anhydrous total of 97.375 wt\%. For 1000 ppm CO\textsubscript{2} and T=1200\degree C, using raw data (SiO\textsubscript{2}=47.4 wt\%) the calculated saturation pressure is 1206 bars. However, it is relatively common in the literature that major elements (excluding volatiles) are scaled to sum to 100\%, while volatile concentrations are left unchanged. This would cause the melt SiO\textsubscript{2} content to increase to 48.68 wt\% SiO\textsubscript{2}, corresponding to a saturation pressure of 1765 bar respectively (1.7\times higher!). We encourage users not to normalize their data, as we note that the II-SiO\textsubscript{2} plot of North Arch Glasses in Dixon (1997) is best recreated using unnormalized data (see Supporting Fig. S1). Unnormalized data is used throughout this review for all VolatileCalc-Basalt outputs, which gives results comparable to those produced in the VolatileCalc-Basalt spreadsheet when users enter the SiO\textsubscript{2} contents given in Table 1.

3.2 VolatileCalc-Rhyolite: Newman and Lowenstern, 2002

In addition to the functionality for basalts described above, the VolatileCalc spreadsheet also allows users to calculate saturation pressures, degassing paths, isobars and isopleths for rhyolitic magmas (hereafter, VolatileCalc-Rhyolite). For CO\textsubscript{2}, VolatileCalc-Rhyolite uses the simple thermodynamic model from Stolper et al. (1987) that was later applied to rhyolitic melts by Fogel and Rutherford (1990). The molar enthalpy change for CO\textsubscript{2} dissolution in the melt is from Fogel and Rutherford (1990), the single-O melt mass from Silver et al. (1990), and the CO\textsubscript{2} molar volume and solubility at standard state from Blank et al. (1993). The thermodynamic basis for the H\textsubscript{2}O model is the same as that used in VolatileCalc-Basalt. The fitted parameters for H\textsubscript{2}O solubility in the standard state is from Silver (1988), and the single-O melt mass and molar enthalpy change for H\textsubscript{2}O dissolution in the melt from Silver et al. (1990). The partial molar volume of H\textsubscript{2}O was adjusted to 5 cm\textsuperscript{3}/mol to provide a better fit to experimental data. There are two main differences of the rhyolite model relative to the basaltic model. First, while both the models use a temperature-dependent equation of state, VolatileCalc-Rhyolite model also contains a term for the heat of solution of volatile solubility, so is far more sensitive to temperature (see section 5). Second, unlike VolatileCalc-Basalt which require users to enter melt SiO\textsubscript{2} contents, VolatileCalc-Rhyolite is independent of melt composition. Thus, Newman and Lowenstern (2002) caution that this model may not be applicable for strongly peralkaline or peraluminous rhyolites.
3.3 M–1998 (Moore et al., 1998)

The Moore et al. (1998) model calculates the solubility of H$_2$O for a wide range of silicate melt compositions using an empirical expression valid between 700–1200°C and 0–3000 bars:

$$2 \ln(X_{\text{H}_2\text{O}}^{\text{melt}}) = \frac{a}{T} + \sum_i b_i X_i^P + c \ln(f_{\text{H}_2\text{O}}^{\text{fluid}}) + d$$  \hspace{2cm} (4)

Where $X_{\text{H}_2\text{O}}^{\text{melt}}$ is the mole fraction of H$_2$O dissolved in the melt, $T$ is the temperature in Kelvin, $P$ is the pressure in bars, and $X_i$ is the anhydrous molar fraction of each oxide component. $f_{\text{H}_2\text{O}}^{\text{fluid}}$ is the fugacity of H$_2$O in the fluid, calculated using the modified version of the Redlich-Kwong equation of state provided in the appendix of Holloway and Blank (1994). Equation 4 is associated with the following fit parameters (± standard error):

| Coefficient | Value (±1σ) |
|-------------|-------------|
| $a$ | 2565 (± 362) |
| $b_{\text{H}_2\text{O}}$ | -1.997 (±0.706) |
| $b_{\text{FeO}}$ | -0.9275 (±0.394) |
| $b_{\text{Na}_2\text{O}}$ | 2.736 (±0.871) |
| $c$ | 1.171 (±0.069) |
| $d$ | -14.21 (±0.54) |

As equation 4 includes a term for the fugacity of H$_2$O in the fluid, this model can be integrated with CO$_2$ models implemented in VESiCal (e.g., the II-SiO$_2$ simplification of Dixon, 1997, Liu et al., 2005, Iacono-Marziano et al., 2012, Shishkina et al., 2014, Allison et al., 2019) to investigate mixed H$_2$O-CO$_2$ fluids.

The model calibration dataset combines the authors’ pure H$_2$O experiments with literature data, spanning sub-alkaline basaltic to rhyolitic compositions, as well as some alkaline compositions (Fig. 2). As with other fully-empirical models implemented in VESiCal, or those including empirical expressions, extreme care must be taken when extrapolating this model outside of the calibration range. In particular, the authors warn against extrapolating this model to pressures exceeding 3 kbar, in part due to the complexities of the critical behavior of fluids at higher pressures.

3.4 L–2005 (Liu et al., 2005)

The Liu et al. (2005) model calculates the solubility of H$_2$O and CO$_2$ in metaluminous, high-silica rhyolitic melts using empirical expressions, valid between 700–1200°C, and 0–5000 bars. The following expression is used to calculate CO$_2$ solubility:

$$[\text{CO}_2]^{\text{ppm}} = \frac{P_{\text{CO}_2}(b_1 + b_2 P_{\text{W}})}{T} + P_{\text{CO}_2}(b_2 P_{\text{W}}^{0.5} + b_3 P_{\text{W}}^{1.5})$$  \hspace{2cm} (5)

$T$ is temperature in Kelvin, $b_1$–$b_4$ are fit parameters, and the $P_{\text{W}}$ and $P_{\text{CO}_2}$ terms account for the partial pressures of each volatile species in the co-existing fluid, with:

$$P_{\text{CO}_2} = X_{\text{CO}_2}^f P$$  \hspace{2cm} (6)

$$P_{\text{W}} = X_{\text{W}}^f P$$  \hspace{2cm} (7)
Where $P$ is pressure in MPa, $X_{f CO_2}$ is the mole fraction of CO$_2$ in the fluid, and $X_{f H_2O}$ is the mole fraction of H$_2$O in the fluid. These empirical terms mean that no equation of state is used (unlike M–1998 and VolatileCalc-Basalt). The fit parameters associated with equation 5 are shown below ($\pm$ error):

\[
\begin{array}{cccc}
\text{Coefficient} & b_1 & b_2 & b_3 & b_4 \\
\text{Value (Error)} & 5668 (\pm 127) & 0.4133 (\pm 0.0481) & 2.041 \times 10^{-3} (\pm 0.285 \times 10^{-3}) & -56.69 (\pm 8.38)
\end{array}
\]

Similarly, they provide the following expression for H$_2$O:

\[
[H_2O]_{wt\%} = \frac{a_1 P_{w}^{0.5} + a_2 P_{w} + a_3 P_{w}^{1.5} + P_{CO2} (a_5 P_{w}^{0.5} + a_6 P_{w})}{T}
\]

Using the following fit parameters:

\[
\begin{array}{cccc}
\text{Coefficient} & a_1 & a_2 & a_3 & a_4 \\
\text{Value (Error)} & 354.04 (\pm 4.65) & 0.623 (\pm 0.023) & -1.5223 (\pm 0.0722) & 0.0012439 (\pm 0.0000499)
\end{array}
\]

\[
\begin{array}{cccc}
& a_5 & a_6 \\
\text{-1.084} \times 10^{-4} (\pm 0.406 \times 10^{-4}) & -1.382 \times 10^{-5} (\pm 0.362 \times 10^{-5})
\end{array}
\]

The model calibration dataset combines pure H$_2$O solubility experiments by the authors between $\sim$1 and 250 bars with literature experiments investigating the solubility of pure H$_2$O, H$_2$O-H$_2$, CO$_2$ and CO$_2$-H$_2$O fluids, spanning significantly higher pressures (up to 5000 bars). Unlike the M–1998 model, their empirical expressions do not incorporate a term for melt composition, so care is needed when applying this model to melts with different major element compositions to the calibration dataset (Fig. 2).

### 3.5 P–2006 (Papale et al., 2006)

Papale et al. (2006) present a fully non-ideal thermodynamic model for H$_2$O and CO$_2$ solubility, which is a recalibration of the earlier models of Papale (1999) and Papale (1997). This updated model capitalizes on the large amount of volatile solubility experiments performed between 1997 and 2005, which nearly doubled the size of the calibration dataset, and allowed experimental data on CO$_2$ solubility collected prior to 1980 to be discarded (removing systematic errors associated with different analytical techniques, see Papale, 1999). Unlike the models discussed above which are calibrated on a specific subregion of compositional space and use empirical parametrizations to account for the effect of melt composition, the models of Papale et al. (2006) and Papale (1999) treat the composition of the silicate liquid using a thermodynamic approach based on Ghiorso et al. (1983). Papale et al. (2006) note that this thermodynamic approach means that for any specific region of composition space (e.g., comparing model results to a specific experiment), the fit may not be as good as an empirical model tuned to that composition. However, carefully calibrated thermodynamic models will be significantly more successful than empirical models when applied to melts which are not represented in the calibration dataset.
P–2006 considers a silicate liquid in mechanical, thermal and chemical equilibrium with a fluid phase containing \( \text{H}_2\text{O} \) and \( \text{CO}_2 \). The model uses the modified Redlich-Kwong equation of state of Kerrick and Jacobs (1981) to describe the fluid phase, and considers only the dissolution of \( \text{CO}_2 \) and \( \text{H}_2\text{O} \) in the melt (while natural silicate melts contain molecular \( \text{CO}_2 \) and \( \text{CO}_2^- \) species, and molecular \( \text{H}_2\text{O} \) and \( \text{OH}^- \) species). The model calculates the Gibbs free energy of mixing, considering 10 major oxide components in addition to \( \text{CO}_2 \) and \( \text{H}_2\text{O} \). Binary interaction coefficients, denoted by \( w_{ij} \), account for the attractive-repulsive behaviour between an oxide and volatile component. For example, \( w_{\text{CO}_2,\text{MgO}} \) describes the interaction of MgO with \( \text{CO}_2 \). Interaction coefficients for \( \text{CO}_2 \) are expressed as a function of pressure (relative to a reference pressure of 0.1 MPa) requiring two coefficients, while those for \( \text{H}_2\text{O} \) are invariant of pressure (requiring 1 coefficient):

\[
w_{\text{CO}_2,i} = w^0_{\text{CO}_2,i} + w^1_{\text{CO}_2,i} \ln \frac{P}{P_0}
\]

\[
w_{\text{H}_2\text{O},i} = w^0_{\text{H}_2\text{O},i}
\]

P–2006 uses interaction terms for \( \text{SiO}_2 \), \( \text{Al}_2\text{O}_3 \), MgO, CaO, Na\(_2\text{O} \), K\(_2\text{O} \), FeO and Fe\(_2\text{O}_3 \). The presence of two Fe terms means that the model is sensitive to melt redox. Papale et al. (2006) show that the inclusion of \( w \) terms for MnO and TiO\(_2 \) lead to overfitting, so the effect of these oxides on model outputs is only through the dilution of the concentration of components allocated \( w \) terms. The values of the 24 \( w \) terms, as well as 5 terms accounting for molar volumes and fugacities, are calculated from a calibration dataset comprising \( \sim 1100 \) solubility experiments with pure \( \text{CO}_2 \), pure \( \text{H}_2\text{O} \) and mixed \( \text{CO}_2\)-\( \text{H}_2\text{O} \) fluids. While the calibration dataset contains well populated clusters for basaltic and rhyolitic compositions, intermediate compositions and basaltic melts with high alkali contents are poorly represented, particularly for \( \text{CO}_2 \) (Fig. 2).

Papale et al. (2006) demonstrate that despite the addition of hundreds of new experimental datapoints for \( \text{H}_2\text{O} \), there are no significant changes in coefficients compared to those which were published with their 1999 model. The percent errors on the \( \text{H}_2\text{O} \) coefficients in the 2006 model are \(< 10\% \) for all species (defined as 100* \( \sigma \)/coefficient). In contrast, the addition of new \( \text{CO}_2 \) data to the calibration dataset resulted in significant changes in coefficients, and the percentage errors on these coefficients in the 2006 model remained large (\(~800\% \) for FeO, \(~150\% \) for Na\(_2\text{O} \), \(~190\% \) for MgO, Fig. 3). Based on these large error values, these coefficients would likely change again if this model was recalibrated to include all new \( \text{CO}_2 \) experiments published since 2006.

Finally, the pressure-dependence of the \( \text{CO}_2 \) melt interaction terms, combined with the fact that the \( w_0 \) and \( w_1 \) terms have different signs for all oxides except FeO, means that a given change in melt chemistry may cause an increase in \( \text{CO}_2 \) solubility at one pressure, but a decrease at another pressure (see Section 4.1). Ghiorso and Gualda (2015) note that the coefficient for the compressibility
of CO₂ in the P–2006 model is negative, which is physically impossible (implying the volume of the CO₂ fluid increases when pressure is increased), which they suggest may arise from the inclusion of pressure-dependent \( w \)-terms.

### P-2006

| Element | CO₂ | H₂O |
|---------|-----|-----|
|         | \( w_0 \) | % error (1σ) | \( w_1 \) | % error | \( w_2 \) | % error |
| SiO₂    | -59,962 | 18 | 6049 | 19 | -34,093 | 1.85 |
| Al₂O₃   | -590,957 | 32 | 41,395 | 47 | -189,117 | 2.54 |
| Fe₂O₃   | 4,469,623 | 9 | -529,301 | 9 | 135,935 | 9.32 |
| FeO     | 21,666 | 806 | 1214 | 1500 | -195,751 | 3.13 |
| MgO     | 52,866 | 189 | -13,446 | 78 | -86,418 | 7.06 |
| CaO     | -328,792 | 23 | 12,798 | 63 | -209,997 | 1.67 |
| Na₂O    | 140,034 | 146 | -35,213 | 60 | -322,253 | 1.42 |
| K₂O     | 309,070 | 48 | -58,010 | 27 | -349,798 | 1.79 |

### MagmaSat

| Element | CO₂-CO₃²⁻ | H₂O |
|---------|----------|-----|
|         | W (CO₂) | % error (1σ) | W (CaCO₃) | % error | W | % error |
| SiO₂    | 63.281 | 2.52 | 27.557 | 0.065 |
| TiO₂    | -19.266 | 24.3 | -79.203 | 0.46 | 88.199 | 2.87 |
| Al₂O₃   | 46.716 | 2.52 | 11.768 | 21.8 |
| Fe₂O₃   | 3.187 | 4.9 | 65.509 | 0.26 | 50.105 | 17.0 |
| Fe₂SiO₄ | -32.465 | 44.5 | -72.997 | 0.40 | 30.936 | 18.8 |
| Mg₂SiO₄ | -40.854 | 164.3 | -24.873 | 4.17 | 20.910 | 21.5 |
| CaSiO₃  | 30.012 | 80.0 | 37.534 | 2.70 | 9.715 | 27.5 |
| Na₂SiO₃ | -311.011 | 0.24 | -82.460 | 2.9 |
| KAlSiO₄ | -27.865 | 8.21 | 1.057 | 112 |
| Ca₃(PO₄)₂ | -3.473 | 172.6 | 37.534 | 0.13 | 44.133 | 0.76 |

**Figure 3.** Interaction coefficients for P–2006 and MagmaSat. Percentage errors calculated as 100*1σ/coefficient. Error are colored green if they are <10%, light pink if 10–25%, and red if >25%.

#### 3.6 IM–2012 (Iacono-Marziano et al. 2012)

The Iacono-Marziano et al. (2012) model expresses the solubility of H₂O and CO₂ in mafic melts by combining simplified thermodynamic expressions for melt-fluid thermodynamics with empirical formulations accounting for melt composition. For CO₂, they present the following expression:

\[
\ln[\text{CO}_2]_{\text{ppm}} = X_{\text{H}_2\text{O}d\text{H}_2\text{O}} + X_{\text{Al}d\text{Al}} + X_{\text{FeO}+\text{MgO}d\text{FeO}+\text{MgO}} + X_{\text{Na}_2\text{O}+\text{K}_2\text{O}d\text{Na}_2\text{O}+\text{K}_2\text{O}} + a_{\text{CO}_2}\ln[P_{\text{CO}_2}] + b_{\text{CO}_2}\left[\frac{\text{NBO}}{\text{OP}}\right] + B_{\text{CO}_2} + C_{\text{CO}_2} \frac{P}{T} \tag{11}
\]

...
Where \( P \) is the pressure in bars, \( T \) is the temperature in Kelvin, \( X_{\text{H}_2\text{O}} \) is the molar fraction of \( \text{H}_2\text{O} \) in the melt and \( P_{\text{CO}_2} \) is the partial pressure of \( \text{CO}_2 \) in bars. The partial pressure of \( \text{CO}_2 \) is calculated from the pressure multiplied by the mole fraction of \( \text{CO}_2 \) in the fluid. This means that this model does not rely on an equation of state (as with the L–2005 model). The other terms account for the effect of melt composition using molar fractions calculated on a hydrous basis; \( X_{\text{FeO}+\text{MgO}} \) is the sum of molar fractions of \( \text{FeO} \) and \( \text{MgO} \), \( X_{\text{Na}_2\text{O}+\text{K}_2\text{O}} \) is the sum of the molar fractions of \( \text{Na}_2\text{O} \) and \( \text{K}_2\text{O} \), and \( X_{\text{Al}} \) is the agpaitic index (AI):

\[
X_{\text{AI}} = \frac{X_{\text{Al}_2\text{O}_3}}{X_{\text{CaO}} + X_{\text{K}_2\text{O}} + X_{\text{Na}_2\text{O}}}
\]  

(12)

The NBO/O term represents the number of non-bridging oxygens divided by oxygen, expressing the availability of oxygen to form carbonate groups within the melt. NBO/O can be calculated from mol fraction of different oxides, \( X_i \), on an anhydrous or hydrous basis:

\[
\frac{\text{NBO}}{O}^{\text{Anhyd}} = \frac{2 (X_{\text{K}_2\text{O}} + X_{\text{Na}_2\text{O}} + X_{\text{CaO}} + X_{\text{MgO}} + X_{\text{FeO}} - X_{\text{Al}_2\text{O}_3})}{2X_{\text{SiO}_2} + 2X_{\text{TiO}_2} + 3X_{\text{Al}_2\text{O}_3} + X_{\text{MgO}} + X_{\text{FeO}} + X_{\text{CaO}} + X_{\text{Na}_2\text{O}} + X_{\text{K}_2\text{O}}}
\]

(13)

\[
\frac{\text{NBO}}{O}^{\text{Hyd}} = \frac{2 (X_{\text{H}_2\text{O}} + X_{\text{K}_2\text{O}} + X_{\text{Na}_2\text{O}} + X_{\text{CaO}} + X_{\text{MgO}} + X_{\text{FeO}} - X_{\text{Al}_2\text{O}_3})}{2X_{\text{SiO}_2} + 2X_{\text{TiO}_2} + 3X_{\text{Al}_2\text{O}_3} + X_{\text{MgO}} + X_{\text{FeO}} + X_{\text{CaO}} + X_{\text{Na}_2\text{O}} + X_{\text{K}_2\text{O}} + X_{\text{H}_2\text{O}}}
\]

(14)

In both cases, mole fractions are calculated on a hydrous basis (Iacono-Marziano, written comms).

Iacono-Marziano give coefficients for equation 11 for both cases (±2σ), leading to two forms of this model: IM–2012–A (anhydrous) and IM–2012–H (hydrous):

| Coefficient | \( d_{\text{H}_2\text{O}} \) | \( d_{\text{AI}} \) | \( d_{\text{H}_2\text{O},\text{MgO}} \) | \( d_{\text{Na}_2\text{O},\text{K}_2\text{O}} \) | \( a_{\text{CO}_2} \) | \( b_{\text{CO}_2} \) | \( c_{\text{CO}_2} \) | \( B_{\text{CO}_2} \) |
|-------------|----------------|----------------|----------------|----------------|----------------|----------------|----------------|----------------|
| **Hydrous (± error)** | -16.4 (±12) | 4.4 (±0.4) | -17.1 (±0.9) | 22.1 (±1.1) | 1 (±0.03) | 17.3 (±0.9) | 0.12 (±0.02) | -6 (±0.4) |
| **Anhydrous (± error)** | 2.3 (±0.5) | 3.8 (±0.4) | -16.3 (±0.9) | 20.1 (±1.1) | 1 (±0.03) | 15.8 (±0.9) | 0.14 (±0.02) | -5.3 (±0.4) |

We note for completeness that in the original publication, equation 11 was incorrectly expressed in terms of \( \ln[\text{CO}_2^2] \) (Iacono-Marziano, written comms).

For \( \text{H}_2\text{O} \), the authors state that it is statistically unjustified to include \( d_i \) terms similar to those in the \( \text{CO}_2 \) expression, due to the relatively small effect of melt composition on \( \text{H}_2\text{O} \) solubility. The effect of melt composition is incorporated only through the NBO/O term:

\[
\ln [\text{H}_2\text{O}]^{\text{wt\%}} = a_{\text{H}_2\text{O}} \ln [P_{\text{H}_2\text{O}}] + b_{\text{H}_2\text{O}} \left[ \frac{\text{NBO}}{O} \right] + B_{\text{H}_2\text{O}} + C_{\text{H}_2\text{O}} \frac{P}{T}
\]

(15)

Where \( P_{\text{H}_2\text{O}} \) is the partial pressure of \( \text{H}_2\text{O} \) in bars. As for \( \text{CO}_2 \), coefficients are provided for NBO/O calculated on a hydrous and anhydrous basis. The hydrous coefficients in the published paper differ from those used in the web app over the last decade (although a new web app using the published coefficients appeared briefly in 2021). VESiCal uses the web app hydrous coefficients by default, as recommended by Iacono-Marziano (written comms). The coefficients in the published paper were from an older version of the model, and predict extremely high \( \text{H}_2\text{O} \) solubility at \( \sim 10 \text{ kbar} \) (>100wt%).
The authors state that the differences between calculations performed with NBO/O calculated on a hydrous and anhydrous basis are relatively small, but that a slightly better fit to experimental data is obtained using the hydrous model (particularly for H$_2$O-rich, and CO$_2$-poor melts). For completeness, we perform calculations using both versions (referred to as IM–20120–H and IM–2012–A). Interestingly, we show that the anhydrous version is most similar to other models for MORB-like compositions.

The calibration dataset for CO$_2$ combines the authors experiments with those from a variety of literature studies for mixed H$_2$O-CO$_2$ fluids, spanning temperatures between 1100 and 1400°C, and pressures between 100 and 10,000 bars (but mostly <5000 bars). The calibration dataset for H$_2$O incorporates pure H$_2$O experiments from the literature (spanning 163–6067 bars, and 1000–1250°C), as well as the experiments on H$_2$O-CO$_2$ fluids used to calibrate the CO$_2$ expression. Melt compositions are predominantly mafic, spanning subalkaline-alkaline basalts to basaltic andesites (45–57 wt% SiO$_2$ for mixed H$_2$O-CO$_2$ experiments, and 46–63 wt% SiO$_2$ for pure H$_2$O; Fig. 2).

The empirical nature of the fitting terms incorporating melt composition, pressure and temperature means that users should be cautious when extrapolating this model to conditions lying outside the P-T-X range of the calibration dataset. In particular, Iacono-Marziano et al. (2012) highlight five weaknesses of their model:

1. The effect of melt MgO and FeO contents on CO$_2$ solubility is poorly constrained, because of the small variation in the concentrations of these oxides in the calibration database.

2. While their compositional terms for the effect of melt composition on CO$_2$ solubility gives equal weight to Na$_2$O and K$_2$O, the calibration dataset only includes K$_2$O-rich melts with a range of pressures, so the effect of substituting Na and K is poorly constrained.

3. The effect of temperature on the solubility of mixed H$_2$O-CO$_2$ is poorly constrained because the majority of experiments in the calibration dataset were performed at 1200–1300°C.

4. The relative role of molecular H$_2$O vs. OH$^-$ on melt structure, which in turn influences CO$_2$ solubility, needs to be evaluated further.

5. The model was calibrated assuming that all Fe was Fe$^{2+}$, so calculated solubilities are not sensitive to melt redox (unlike the model of P–2006).

These weaknesses are explored in more detail in sections 5 and 6.
### 3.7 S–2014 (Shishkina et al., 2014)

The Shishkina et al. (2014) model calculates the solubility of H$_2$O and CO$_2$ using fully-empirical expressions. Their expression for CO$_2$ solubility was calibrated on a dataset of mixed H$_2$O-CO$_2$ experiments on predominantly mafic compositions between 1200–1300°C, and 500–5000 bars:

$$\ln[\text{CO}_2]^{\text{ppm}} = 1.150 \ln(P) + 6.71 \Pi^* - 1.345$$  \hspace{1cm} (16)

Where $P$ is the pressure in MPa, and $\Pi^*$ is a compositional parameter expressed in terms of the cation fractions of 7 species:

$$\Pi^* = \frac{X_{\text{Ca}^{2+}} + 0.8 X_{\text{K}^+} + 0.7 X_{\text{Na}^+} + 0.4 X_{\text{Mg}^{2+}} + 0.4 X_{\text{Fe}^{2+}}}{X_{\text{Si}^{4+}} + X_{\text{Al}^{3+}}}$$  \hspace{1cm} (17)

We note for completeness that the expression provided in Shishkina et al. (2014) incorrectly states that CO$_2$ in equation 16 was in wt%, rather than ppm.

Their expression for H$_2$O solubility was calibrated on a dataset of pure H$_2$O experiments with mafic to intermediate compositions between 1200–1250°C, and 485–5009 bars. It incorporates a composition parameter expressed in terms of the anhydrous cation fractions of Na and K:

$$[\text{H}_2\text{O}]^{\text{wt}\%} = (3.36 \times 10^{-7} P^3 - 2.33 \times 10^{-4} P^2 - 0.0711 P - 1.1309)(X_{\text{Na}} + X_{\text{K}})$$

$$- 1.2 \times 10^{-5} P^2 + 0.0196 P + 1.1297$$  \hspace{1cm} (18)

In general, the compositional range of the Shishkina et al. (2014) dataset includes a larger variety of mafic compositions than that of Iacono-Marziano et al. (2012), particularly with respect to alkali-rich, or highly depolymerized melts (Fig. 2). However, as was true for IM–2012, the empirical nature of the compositional term means that extreme care is needed when extrapolating this model beyond the compositional range of the calibration dataset (see section 6).

One caveat of the implementation of this model in VESIcal is the treatment of mixing between CO$_2$ and H$_2$O. Shishkina et al. (2014) note that their experimental data shows evidence for significant non-ideality, with isobars remaining almost horizontal between 0–4 wt% H$_2$O (see their Fig. 6). However, the isobars shown on their plots are fitted to experimental data, rather than derived from their equations for CO$_2$ and H$_2$O solubility. These fits cannot be applied to melts with different compositions, and the authors give no guidance as to how to combine their equations for pure CO$_2$ and pure H$_2$O to reproduce this non-ideal behaviour for any given melt composition. Thus, due to an absence of other information, VESIcal treats mixing between H$_2$O and CO$_2$ as ideal in this model. To emphasize this assumption, the H$_2$O-CO$_2$ model called ShishkinaIdealMixing in VESIcal.

### 3.8 MagmaSat: Ghiorso and Gualda (2015)

Ghiorso and Gualda (2015) present a comprehensive thermodynamic model (MagmaSat) for mixed H$_2$O-CO$_2$ solubility, calibrated on the most chemically-diverse set of natural silicate melt com-
positions of all the models discussed thus far (Fig. 2). For this reason, it is the default model in VESI-
cal. MagmaSat uses the equation of state of Z. Duan and Zhang (2006) for the CO$_2$-H$_2$O fluid, and
is the volatile solubility model implemented in rhyolite-MELTS v.1.2 (Gualda et al., 2012). Thus,
it is currently the only model which can be directly integrated with phase equilibrium calculations
(e.g., to track microlite growth during degassing upon ascent, or post-entrapment modification to
melt inclusions). Like P–2006, the model considers the Gibbs free energy of solution using interac-
tion parameters (denoted with a capital W in this model), although, unlike in P–2006, these terms
are independent of pressure (as well as temperature). The exact choice of components differs from
that of P–2006, as MagmaSat adapts the formulation from Rhyolite-MELTS (e.g., Mg is considered as
Mg$_2$SiO$_4$, Ghiorso & Sack, 1995). Unlike P–2006, MagmaSat considers the dissolution of CO$_2$ as both
molecular CO$_2$ and carbonate species (as CaCO$_3$), with a set of W coefficients for each. It is appar-
ent from the percentage errors on these coefficients that the effect of melt composition on carbonate
solubility is much better constrained than the effect of melt composition on molecular CO$_2$ solubility
(percentage errors < 10% vs. errors up to 170%, Table 3). MagmaSat assumes that water dissolves
totally as a hydroxyl species, rather than considering both hydroxyl and molecular species. This helps
to reduce the number of interaction parameters for volatile-melt species, and seems to be a justified
simplification based on available experimental data (see Ghiorso and Gualda, 2015 for a more detailed
discussion).

The calibration dataset for H$_2$O spans 550–1420°C, and pressures of 0–20,000 bars, and for CO$_2$
spans 1140–1400 °C and 0–30,000 bars. Importantly, unlike P–2006, MagmaSat is not calibrated for
synthetic liquids (e.g., compositions only containing a small number of oxide species like albite), so
should only be applied to natural silicate liquid compositions.

3.9 A–2019: Allison et al. (2019)

Allison et al. (2019) present thermodynamic models to calculate CO$_2$ solubility for six different
basaltic compositions from Stromboli (alkali basalt), Etna (trachybasalt), Vesuvius (phonotephrite),
Erebus (phonotephrite), Sunset Crater (alkali basalt), and the San Francisco Volcanic Field (basaltic
andesite, Fig. 2). Specifically, they performed experiments at 1200°C, and ~4000–6000 bars to address
the paucity of experiments examining CO$_2$ solubility in alkali systems at mid crustal pressures. In
addition to these experiments, their models for Vesuvius, Etna and Stromboli incorporate experiments
from the literature, extending the calibration range of these three models to upper crustal pressures
(see Fig. 1). Unlike models which incorporate the effect of changing melt composition empirically
(e.g., Newman & Lowenstern, 2002; Iacono-Marziano et al., 2012; Shishkina et al., 2014), Allison et
al. (2019) determine the parameters $\Delta V_{\nu,m}$ (the molar volume change of the condensed components
of the reaction) and $K_0(P_0, T_0)$ (the equilibrium constant at the reference pressure and temperature) within their thermodynamic equation empirically for each of the six compositions they examine, and create 6 separate models (each of which contains no compositional dependence). The A–2019 models incorporate the modified Redlich-Kwong equation of state provided in the appendix of Holloway and Blank (1994).

Interestingly, Allison et al. (2019) show that CO$_2$ solubility does not simply scale with total alkali contents. Erebus melts have Na$_2$O+K$_2$O=8.8 wt%, but dissolve less CO$_2$ than Etna and Vesuvius melts (Na$_2$O+K$_2$O=5.2 and 7.8 wt% respectively). They suggest that CaO, MgO and Al$_2$O$_3$ may play a role in the lower solubility of Erebus compared to Etna, but the fact that 5 of the 7 major elements they examine show notable differences between these melt compositions make it difficult to conclusively determine the origin of solubility variations. Ideally, users would apply the A–2019 Etna model to lavas erupted at Etna, the Stromboli model to lavas erupted at Stromboli and so on. The absence of an empirical term for melt composition means that extreme care should be taken when applying these equations to alkaline lavas with different major element contents to those used in the experiments of Allison et al. (2019), even if the lavas originated from one of the 6 volcanoes they examine.

Allison et al. (2019) only present equations for CO$_2$ solubility, as their experiments weren’t designed to have a wide range of H$_2$O contents at different pressures, and their high $X_{CO_2}$ values mean that errors in their fluid fraction measurements propagate into large errors for H$_2$O fugacity (relative to the insignificant errors for CO$_2$ fugacity). In their supplementary spreadsheet, they integrate their CO$_2$ solubility models with an power law fit for water solubility developed specifically for Etna (Equation 2 of Lesne, Scaillet, Pichavant, Iacono-Marziano, & Beny, 2011). In VESIcal, users can combine any of the A–2019 carbon models with H$_2$O models from M–1998, IM–2012 and S–2014, or write their own.

4 Model Comparisons

To aid comparisons between models, a number of silicate melt compositions (Table 1) are used to examine the relationship between volatile solubility and pressure, the treatment of mixing between H$_2$O and CO$_2$ (manifested in the shapes of isobars), as well as sensitivity to parameters such as temperature and oxygen fugacity. For basalts, we compare four compositions; two mid-ocean ridge basalts (termed MORB1 and MORB2), one alkali basalt and one arc basalt (grey crosses on Fig. 2). MORB1 from Dixon et al. (1995) has 50.8 wt% SiO$_2$, MORB2 from table 3 of Ghiorso and Gualda (2015) has 47.4 wt% SiO$_2$ (originally from Allan et al., 1989), and the Etna alkali basalt from Allison et al. (2019) has 48.8 wt% SiO$_2$. Comparisons between these three compositions were performed at 1200°C.
To investigate model sensitivities with relevance to mafic melt inclusion studies in volcanic arcs, we use the composition of a Fuego melt inclusion from Lloyd et al. (2013) with 49.7 wt% SiO$_2$ at 1000°C.

For rhyolitic magmas, we perform calculations at 800°C for a rhyolite from Mono Craters (Eastern California) and Aluto (African Rift). The Mono Craters composition is a high-Si, metaluminous rhyolite with 77.19 wt% SiO$_2$ from Liu et al. (2005). The Aluto composition is a peralkaline rhyolitic melt inclusion with 73.5 wt% SiO$_2$ from Iddon and Edmonds (2020). The Aluto rhyolite has much lower Al$_2$O$_3$ and higher FeO contents than the Mono Craters Rhyolite (Table 1).

**Table 1.** Representative compositions used for comparisons. MORB1 is a Mid-Oceanic Ridge Basalt tholeiite from Dixon et al. (1995). MORB2 is the MORB composition given in Table 3 of Ghiorso and Gualda (2015), originally from Allan et al. (1989). Etna is sample ET-8 from the supplementary information of Allison et al. (2019). Fuego is the composition of a melt inclusion from Lloyd et al. (2013). Mono Craters is from Table 1 of Liu et al. (2005). Aluto rhyolite is the composition of a quartz-hosted melt inclusion from the East African Rift (MI70 from sample MER055A; Iddon & Edmonds, 2020).

| Name   | MORB1 | MORB2 | Etna | Fuego | Mono Craters | Aluto Rhyolite |
|--------|-------|-------|------|-------|--------------|----------------|
| SiO$_2$ | 50.8  | 47.4  | 48.77| 49.67 | 77.19        | 73.51          |
| TiO$_2$ | 1.84  | 1.01  | 1.79 | 1.17  | 0.06         | 0.23           |
| Al$_2$O$_3$ | 13.7 | 17.64 | 16.98| 16.50 | 12.8         | 9.18           |
| Fe$_2$O$_3$ | 2.1  | 0.89  | 2.51 | 1.65  | 0.26         | 1.41           |
| FeO     | 10.5  | 7.18  | 6.44 | 8.43  | 0.71         | 3.81           |
| MnO     | 0     | 0     | 0.18 | 0.19  | 0            | 0.25           |
| MgO     | 6.67  | 7.63  | 6.33 | 4.38  | 0.03         | 0              |
| CaO     | 11.5  | 12.44 | 11.26| 7.90  | 0.53         | 0.2            |
| Na$_2$O | 2.68  | 2.65  | 3.65 | 3.37  | 3.98         | 4.18           |
| K$_2$O  | 0.15  | 0.03  | 1.79 | 0.79  | 4.65         | 4.22           |
| P$_2$O$_5$ | 0.19 | 0.08  | 0.53 | 0.22  | 0            | 0              |

### 4.1 Redox sensitivity

Before proceeding with these comparisons, it is worth noting that the vast majority of studies report whole-rock, melt inclusion and matrix glass compositions in terms of FeO$_t$, because the proportions of FeO vs. Fe$_2$O$_3$ are difficult to determine precisely using common analytical techniques such as electron probe microanalysis (EPMA) and x-ray florescence (XRF). VolatileCalc-Rhyolite and L-2005 have no compositional terms, and VolatileCalc-Basalt is only parametrized in terms of the melt SiO$_2$ content, so these 3 models are not sensitive to the choice of FeO vs. Fe$_2$O$_3$ for the representative compositions in Table 1. Similarly, IM–2012 and M–1998 are parameterized using an FeO$_t$ term, so also show no sensitivity to melt redox. S–2014 is technically slightly redox-sensitive for CO$_2$, because the II* term is expressed in terms of Fe$^{2+}$ species (equation 17). However, given that the model was calibrated assuming Fe$^{2+}$=Fe$^T$, any sensitivity to redox is likely spurious, so VESIcal calculates II with
Fe\textsubscript{r} by default for consistency with their calibration. In Figure 4, we perform calculations for different Fe\textsuperscript{3+} ratios for completeness, but the rest of the figures in the manuscript for S–2014 are calculated using Fe\textsubscript{r}. H\textsubscript{2}O solubility in S–2014 is not redox sensitive, because the effect of melt composition is only parametrized in terms of cation fractions of Na and K (equation 18). Both P–2006 and MagmaSat have interaction parameters for Fe\textsuperscript{2+} and Fe\textsuperscript{3+}-bearing species, so are redox sensitive for both CO\textsubscript{2} and H\textsubscript{2}O solubility.

We examine the sensitivity of our calculations of volatile solubility to melt redox by performing calculations for 0, 10 and 20% Fe\textsuperscript{3+} for MORB2, and 0, 30 and 60% Fe\textsuperscript{3+} for Etna (the higher redox accounting for the highly oxidising conditions of experiments on Etna melts, e.g., Lesne, Scaillet, Pichavant, & Beny, 2011).

Briefly for Etna, pure H\textsubscript{2}O solubility in MagmaSat is relatively insensitive to redox, predicting variations in dissolved H\textsubscript{2}O which are well within model error (Fig. 4a). Pure CO\textsubscript{2} solubility in MagmaSat is more redox sensitive than H\textsubscript{2}O, predicting \( \sim 1.2-1.3 \times \) more CO\textsubscript{2} for 0% Fe\textsuperscript{3+} vs. 60% Fe\textsuperscript{3+} (Fig. 4b). Pure H\textsubscript{2}O solubility in P–2006 shows the same directionality as MagmaSat, but is more sensitive to redox (1.8\times more H\textsubscript{2}O dissolves at 0.1 kbar for 0% Fe\textsuperscript{3+} vs. 60% Fe\textsuperscript{3+}, dropping to 1.2\times at > 2 kbar). Pure CO\textsubscript{2} solubility in P–2006 is extremely redox-sensitive, with melts with 0% Fe\textsuperscript{3+} vs. 60% Fe\textsuperscript{3+} dissolving 25\times more CO\textsubscript{2} at 0.5 kbar, but 0.5\times less at 5 kbar. S–2014 is slightly less redox sensitive than MagmaSat for CO\textsubscript{2}.

Varying Fe\textsuperscript{3+} proportions between 0–20% for MORB2 produces similar patterns as Etna, with changes lying within model error for MagmaSat and S–2014, but showing significant differences for P–2006 (Fig. 4c-d). Isobars for different redox states for this composition can be found in Supporting Figure S2. For the smaller changes in Fe\textsuperscript{3+} proportions considered for MORB2, changes in dissolved H\textsubscript{2}O and CO\textsubscript{2} contents for MagmaSat and S–2014 are well within model uncertainty (generally stated as 10–20%). In contrast, P–2006 shows changes in dissolved CO\textsubscript{2} which are significantly larger than quoted errors on solubility models.
Figure 4. Relationship between volatile solubility and the proportion of Fe$^{3+}$ for: a-b) the Etna composition at 1200$^\circ$C; c-d) the MORB2 composition at 1200$^\circ$C. VESIcal uses FeO$t$ in S–2014 to calculate $\Pi^*$ for consistency with the calibration of this model. Here, dashed and dotted lines show the results from calculations where FeO$t$ is multiplied by Fe$^{2+}$/Fe$^T$, to show the results that would be obtained if VESIcal calculated $\Pi^*$ using only Fe$^{2+}$. Isobars for MORB2 and lines for intermediate Fe$^{3+}$ ratios for each composition are shown in Supporting Fig. S2.

The different sensitivities of MagmaSat and P–2006 are apparent from examining the interaction coefficients in Table 3. In MagmaSat, the $W_{Fe_2O_3, H_2O}$ coefficient is only 1.6× bigger than $W_{Fe_2SiO_4, H_2O}$ ($\sim$50 vs. 31), and these two coefficients overlap within ±1.5 $\sigma$ of the uncertainty of these coefficients. This accounts for the relatively weak effect of redox on calculated H$_2$O solubility. For the CaCaO$_3$ component representing the carbonate ion, the coefficients have similar magnitudes, but opposite signs ($W_{Fe_2O_3, CaCO_3}$ $\sim$66, $W_{Fe_2SiO_2, CaCO_3}$ $\sim$–73), and this difference is much larger than the error on the coefficients (accounting for the stronger effect of melt redox on CO$_2$ solubility compared with H$_2$O). The Fe$_2$O$_3$ and Fe$_2$SiO$_4$ coefficients for the CO$_2$ component (which becomes
more dominant in more evolved compositions) are also significantly different outside the quoted error but have the same sign \((W_{Fe_2O_3}, CO_2 \sim -32, W_{Fe_2SiO_2}, CO_2 \sim -3})

In P-2006, the \(w_{H_2O-FeO}^0\) coefficient is of similar magnitude, but opposite sign to \(w_{CO_2-FeO}^0\) (1.4 \times 10^5 vs. \(-2 \times 10^5\)), and clearly distinct outside the error on each coefficient. This accounts for the slightly stronger sensitivity of \(H_2O\) in P-2006 to redox compared with MagmaSat. In stark contrast to all the comparisons thus far, the \(w_{CO_2-Fe_2O_3}^0\) coefficient is \(>200\times\) larger than \(w_{CO_2-FeO}^0\). In fact, \(w_{CO_2-Fe_2O_3}^0\) is \(\sim 8\times\) higher than the next largest coefficient, suggesting that for a given mole fraction in the melt, it has the largest effect on carbonate ion solubility. The \(w_{CO_2-Fe_2O_3}^1\) coefficient in P-2006 model, which becomes more dominant at higher pressures, has the opposite sign to that for \(w_0\). This accounts for the fact that at low pressures (<3 kbar), increasing proportions of \(Fe^{3+}\) cause a decrease in \(CO_2\) solubility, while at higher pressures, increasing proportions of \(Fe^{3+}\) cause \(CO_2\) solubility to increase.

It is difficult to trust the extreme sensitivity of \(CO_2\) in P-2006 to redox given the large uncertainty associated with the proportions of \(Fe^{3+}\) in volatile solubility experiments. For example, S-2014 note that only 7 of the 48 experiments in their calibration dataset contain non-zero values of \(Fe_2O_3\), which is why they choose to calibrate the model using \(FeO_1\). In the P-2006 dataset, only 6 studies used in the calibration directly determined the proportion of \(Fe^{3+}\), and a further 9 reported the experimental oxygen fugacity. Thus, for the vast majority of their experimental calibration dataset, Papale et al. (2006) calculate the proportion of \(Fe^{3+}\) assuming the oxygen fugacity is controlled by the \(H_2O-H_2\) equilibrium at the stated experimental conditions. However, this method requires accurate measurements of fluid composition, is affected by \(Fe\) and \(H^+\)-loss during experiments, and it is unclear how applicable this method is for mixed \(H_2O-CO_2\) experiments (Botcharnikov et al., 2006). The P-2006 calibration dataset contains some very surprising values: in the calibration dataset for pure \(CO_2\) experiments, the experiments of Fogel and Rutherford (1990) have been allocated \(Fe^{3+}/Fe_T\) ratios of \(\sim 0.9\), despite the authors debating whether their experiments were actually reducing enough to stabilize a \(CO\) species. Similarly, the mixed \(CO_2-H_2O\) experiments of Paonita et al. (2000) have been allocated \(Fe^{3+}/Fe_T\) ratios of \(\sim 0.9\), despite the starting materials having \(Fe^{3+}/Fe_T\) ratios of 0.01 for rhyolites and 0.45 for basalts. Although it might seem that a few such studies will not influence the overall result, it is worth noting that the coefficients for \(CO_2\) are relatively underconstrained. While there are 10 coefficients for \(H_2O\) calibrated using 865 datapoints of pure-\(H_2O\) solubility, there are 20 coefficients for \(CO_2\), and only 173 datapoints for pure-\(CO_2\) solubility and and 84 for mixed fluids. In particular, the highly oxidising \(Fe^{3+}/Fe_T\) ratios calculated assuming \(H_2O-H_2\) equilibrium are more concentrated in experiments with low \(CO_2\) contents and pressures, making it difficult to deconvolve...
the differential effects of these parameters in a model with a large number of coefficients being cali-
ibrated on a relatively small calibration dataset.

Finally, we wish to draw attention to the very large error on the CO$_2$-FeO coefficients ($w_0 \sim$ 800% and $w_1 \sim 1500\%$), but the relatively small error on the CO$_2$-Fe$_2$O$_3$ coefficients in the P–2006
model ($w_0 \sim 9\%$ and $w_1 \sim 9\%$; Fig. 3). It is likely that the error on the FeO coefficient accurately
represents the large uncertainty on the effect of melt redox on CO$_2$ solubility, while the 9% error on
the CO$_2$-Fe$_2$O$_3$ coefficients is unrealistic, given the very large uncertainties associated with estimating
Fe$^{3+}$/Fe$_T$ proportions from experiments which did not report any information on redox conditions.

In sections 4.2 to 5, we show calculations using the Fe$^{3+}$/Fe$_T$ proportions in Table 1, as these
best-estimates for each center are representative of what a user would select when calculating melt
inclusion saturation pressures, dissolved volatile contents etc. For MORB1, MORB2, and Fuego, these
proportions are from the original publications. For Etna, Fe$_2$O$_3$ was calculated from FeO$_t$ assuming
Fe$^{3+}$/Fe$_T$=0.26 after Gaborieau et al. (2020). For Mono Craters and the Aluto pumice, Fe$^{3+}$/Fe$_T$ was
set at 0.25 based on available data on other rhyolites (e.g., Ghiorso and Gualda, 2015), and modelling
studies of the fractional crystallization path at Aluto (Gleeson et al., 2017).

4.2 Mafic Compositions

4.2.1 Pure H$_2$O

The 7 models applicable to H$_2$O in basaltic systems predict a sharp rise in pure H$_2$O solubility with
increasing pressure (Fig. 5). For all three melt compositions, S–2014, IM–2012–A, VolatileCalc-Basalt,
and M–1998 predict H$_2$O concentrations within ±10% of MagmaSat (grey envelope) between ∼1–5
kbar. For MORB1 and MORB2, IM–2012–H begins to deviate to higher H$_2$O contents than Mag-
maSat at >1 kbar. For Etna, IM–2012–H follows a similar trajectory between 0–3 kbar to the solu-
bility model of Lesne, Scaillet, Pichavant, Iacono-Marziano, and Beny (2011) developed specifically
for Etna melts (yellow line, Fig. 5d). In contrast, P–2006 plots to substantially higher H$_2$O solubili-
ies compared to all other models at >0.5 kbar (although P–2006 and IM–2012–H intercept at higher
pressures).

The fact that IM–2012–H predicts higher H$_2$O solubility relative to the cluster of other mod-
els lying within the error window of MagmaSat is an interesting observation. Iacono-Marziano et al.
(2012) favour their hydrous model, particularly for CO$_2$-poor, H$_2$O-rich melts, based on regressions
between predicted and measured H$_2$O contents, and by comparing the two models to experiments
conducted between 1 and 4 kbar on Etna melts (their Figure 7a). However, their Figure 8c, which
compares dissolved volatile contents calculated by the model for the entire calibration dataset, shows
that predicted H$_2$O contents using the hydrous version are overestimates for experimental products
Figure 5. Relationship between pure H$_2$O solubility and pressure for MORB1, MORB2 and Etna melts at 1200°C. The grey field shows a ±10% error window around MagmaSat. Model lines are dashed when extrapolated above the recommended pressure range. The low pressure region of a) is expanded in c), emphasizing the non-zero solubility of H$_2$O at P=0 bar in the S-2014 model.

with >6–7 wt% H$_2$O (although these predictions still lie within the ~17% error associated with their H$_2$O model). Our comparisons suggest that the anhydrous model is most similar to other models, so should not automatically be discounted in favour of the hydrous model.

Another notable oddity is the nearly linear trajectory of H$_2$O vs. P in S-2014 at <0.5 kbar, causing this model to predict a non-zero solubility of H$_2$O at 0 bar (Fig. 5c). This contrasts with the power-law shapes followed by the other models which intercept very close to the origin. This anomalous behaviour is because the S-2014 equation for H$_2$O solubility (equation 18) simplifies at P=0 to:

$$[H_2O]^{wt\%} = -1.1309(X_{Na} + X_{K}) + 1.1297$$

(19)
In the S–2014 calibration dataset, $X_{Na} + X_{K}$ varies from 0.05 to 0.25, which corresponds to solubilities of 0.85–1.07 wt% $H_2O$ at 0 bar. This demonstrates the issue with extrapolating empirical expressions beyond the calibration range (the lowest pressure experiment in the calibration dataset of S–2014 was conducted at 485 bar). When combined with the assumption of ideal mixing used in VESIcal, this non-zero solubility of $H_2O$ at 0 bar results in S–2014 predicting unusual degassing paths and isobar shapes relative to other models. For example, if a melt has <1 wt% $H_2O$, S–2014 predicts that the co-existing fluid contains no $H_2O$, despite abundant evidence that volcanic plumes in low $H_2O$ systems such as Hawai’i are dominated by $H_2O$ at low pressures (Gerlach, 1986). It also causes isobars to be entirely flat at low $H_2O$ contents (see section 4.2.2, Fig. 6).

Overall, excluding P–2006 and IM–2012–H based on their higher predictions of $H_2O$ solubility, and S–2014 based on anomalous behavior at low $H_2O$ contents, the remaining four solubility models predict dissolved $H_2O$ concentrations within error of one another at pressures lower than the upper calibration limit. This likely reflects the relatively small effect of melt composition of $H_2O$ solubility, meaning that more recent models calibrated on a wider compositional range display similar behavior to older models (G. Moore & Carmichael, 1998; Papale et al., 2006). The larger deviation between models at higher pressures reflect the fact that very few pure-$H_2O$ solubility experiments have been performed at > 5 kbar (Table 1). One reason for this shortage of higher pressure experimental data results from the fact that it is very difficult to quench silicate melts with >9 wt% $H_2O$ to a glass phase which can be analysed by FTIR or SIMS (Gavrilenko et al., 2019; Mitchell et al., 2017).

### 4.2.2 Mixed $H_2O-CO_2$

The majority of experiments used to calibrate expressions for pure-CO$_2$ solubility contained dissolved $H_2O$ and CO$_2$ (e.g., Iacono-Marziano et al., 2012; Shishkina et al., 2014; Allison et al., 2019), requiring authors to assess $H_2O-CO_2$ mixing behaviour to determine the solubility of pure-CO$_2$ fluids. Thus, it makes sense to consider the treatment of mixing between CO$_2$ and $H_2O$ species in each model before considering predictions of pure CO$_2$ solubility which are affected by these assumptions. The treatment of $H_2O-CO_2$ mixing is best demonstrated using isobar diagrams, which show the solubility of $H_2O$ and CO$_2$ in a given silicate melt composition at a given pressure for proportions of $X_{H_2O}$ in the co-existing fluid ranging from 0 (interception with the y axis) to 1 (interception with the x axis). The treatment of mixed fluids differs quite considerably in each model.

VolatileCalc-Basalt models mixed fluids under the assumption of Henrian (ideal) mixing in the fluid and melt phase. Thus, the addition of $H_2O$ always causes the solubility of CO$_2$ to decrease (and vice versa), and isobars possess a negative gradient, with a slightly sloping plateau at low $H_2O$ contents merging into a concave-down shape (Fig. 6).
S-2014 does not provide an equation for the treatment of non-ideal mixing, despite their experiments showing that increasing H$_2$O contents at high pressure cause almost no change in CO$_2$ solubility. Using the assumption of ideal mixing in VESIcal, S-2014 isobars exhibit a flat plateau at low H$_2$O contents, merging into a negative slope at higher H$_2$O contents. This flat plateau results from the fact that there are no partial pressures at which S-2014 yields H$_2$O < 1 wt%, so the y co-ordinate for lower H$_2$O contents is equal to the solubility of pure CO$_2$.

P-2006 is fully non-ideal, which causes isobars to have complex shapes, exhibiting both positive and negative gradients. In detail, the shape of isobars calculated using P-2006 differ as a function of both melt composition and temperature (see Fig. 12 of Papale, 1999). For the basaltic compositions considered here, isobars show a sharp increase to higher CO$_2$ contents at very low H$_2$O contents (Fig. 6a, c), because the model predicts a decrease in CO$_2$ solubility following the addition of small quantities of H$_2$O.

IM–2012–A and IM–2012–H incorporate empirical representations of non-ideality through the inclusion of a term for the molar fraction of H$_2$O in the melt in their expression for CO$_2$ solubility (equation 11). This means that these models predict that maximum CO$_2$ solubility occurs at non-zero H$_2$O concentrations, causing isobars to display prominent domed shapes (Fig. 6). Isobars calculated using IM–2012–A show a more extreme peak than IM–2012–H, because of the difference in the sign and magnitude of the dH$_2$O coefficient combined with the differences between NBO/O calculated on a hydrous and anhydrous basis (see Supporting Information for further detail).
Figure 6. 1 and 4 kbar isobars for MORB1 (a-b) and MORB2 (c-d) and Etna (e-f) at 1200°C. The 1 kbar isobar is expanded in b), d) and e). The distinctive shapes of isobars from different models reflects different treatment of H$_2$O-CO$_2$ mixing. This is most apparent at higher pressures. The prominent plateau at <$1$ wt% H$_2$O for S–2014 at all pressures results from the non-zero solubility of H$_2$O at low pressures.

Like P–2006, MagmaSat is fully non-ideal. However, unlike P–2006, the treatment of non-ideality in MagmaSat predicts that the addition of small amounts of H$_2$O always causes the solubility of CO$_2$
to increase (so isobars peak at non-zero H$_2$O concentrations; Fig. 6, c). This peak becomes more pronounced at higher pressures, but is generally smaller than that predicted by IM–2012.

These different mixing assumptions result in large discrepancies between the predicted volatile solubilities for melts in equilibrium with H$_2$O-CO$_2$ fluids, particularly at higher pressures where non-ideal behaviour is more pronounced. For example, while IM–2012–A predicts similar pure CO$_2$ and pure H$_2$O solubilities to VolatileCalc-Basalt and MagmaSat for MORB1 at 4 kbar (intersection with x and y axis on Fig. 6a), IM–2012–A predicts that melts with $\sim$ 4 wt% H$_2$O can dissolve more than twice as much CO$_2$ as that predicted by VolatileCalc-Basalt.

4.2.3 Pure CO$_2$

All basaltic compositions and models show a large increase in the solubility of pure CO$_2$ with increasing pressure (Fig. 7). The solubility of pure CO$_2$ is approximately an order of magnitude lower than for H$_2$O (compare Fig. 7 with Fig. 5). This solubility difference accounts for the fact that Mid Oceanic Ridge (MOR) magmas, which have similar concentrations of H$_2$O and CO$_2$ ($\sim$ 0.07 wt% H$_2$O, 0.1–0.2 wt% CO$_2$; Le Voyer et al., 2019), are almost always CO$_2$ saturated during crustal storage (Saal et al., 2002) but only exsolve measurable quantities of H$_2$O if erupted at very low pressures (Le Voyer et al., 2019).

For MORB1, IM–2012–A and H, and VolatileCalc-Basalt lie within, or close to the ± 10% error window on MagmaSat at <5 kbar, and S–2014 lie within ± 20%. The deviation at higher pressures is expected, because only P–2006 and MagmaSat are calibrated on large numbers of experiments performed at >5 kbar (Fig. 1). For example, the relationship between Π and CO$_2$ solubility of Dixon (1997) used in VolatileCalc-Basalt was based on experiments at 1 kbar, and Newman and Lowenstern (2002) suggest that it should not be extrapolated above 5 kbar. Similarly, only the experiments of Jakobsson (1997) in the IM–2012 database were conducted at >5 kbar, and there are no experiments in the calibration dataset of S–2014 performed at >5 kbar. Unlike for pure H$_2$O, IM–2012–A and H predict very similar pure CO$_2$ solubilities to one another. This reflects the fact the coefficients for CO$_2$ between these 2 model versions are very similar (apart from the $d_{H_2O}$ term, which is multiplied by a zero when calculating pure CO$_2$ solubility). In contrast, P–2006 plots to significantly higher pressures than the other models ($\sim$ 2× higher at ~8 kbar).

MORB2 shows a significantly larger discrepancy between different models at all pressures (Fig. 7b), although S–2014, IM–2012–H and –A follow very similar trajectories at <5 kbar. Most notably, P–2006 predicts that MORB2 dissolves $\sim$3370 ppm CO$_2$ at 2 kbar while MagmaSat predicts only $\sim$950 ppm (factor of 3.5×). VolatileCalc-Basalt also predicts higher CO$_2$ solubility relative to MagmaSat by a factor of 1.8× at 2 kbar.
Figure 7. Relationship between pure CO$_2$ solubility and pressure for three mafic melts at 1200°C: a) MORB1 from Dixon et al. (1995), b) MORB2 from Ghiorso and Gualda (2015), and c) Etna from Allison et al. (2019). Models extrapolated beyond their calibrated pressure range are shown as dashed lines, with the colored star marking the recommended upper calibration limit. d) The relationship between $\Pi$ and SiO$_2$ defined by the North Arch lavas is shown in blue (Dixon, 1997). Generally, VolatileCalc-Basalt is applied to melts with >49 wt% SiO$_2$ by setting SiO$_2$=49 wt%; the black dashed line represents this extrapolation. These simplified relationships incorporated into VolatileCalc-Basalt (blue and black lines) underestimate the true $\Pi$ value for Etna, and overestimate it for MORB2.

The high CO$_2$ solubility predictions by P–2006 relative to other models and experimental measurements has also been noted by Shishkina et al. (2010), Shishkina et al. (2014) and Mangan et al. (2021). This may result from the fact that the P–2006 has a negative compressibility for the CO$_2$ fluid (Ghiorso & Gualda, 2015). Alternatively, it may reflect the large errors on the CO$_2$ w coefficients, meaning that the effect of melt composition is uncertain, accounting for the larger discrepancy for MORB1 vs. MORB2 compared to other models. Finally, Shishkina et al. (2010) suggest that the
overprediction of CO₂ solubility by P-2006 may result from the inclusion of anomalously high CO₂ contents from the experiments of Freise (2004) in the calibration dataset of P-2006 (these values have now been revised to lower numbers, as the original FTIR thickness correction factor is thought to have been incorrect).

The fact that VolatileCalc-Basalt plots close to other models for MORB1 but not MORB2 is a good example of the main caveat of the II-SiO₂ simplification used to account for the effect of melt composition on CO₂ solubility. For melts with 40–49 wt% SiO₂, VolatileCalc-Basalt assumes that the relationship between CO₂ solubility and SiO₂ is identical to that defined by the North Arch lavas, shown in II vs. SiO₂ space as a blue line in Fig. 7d. This is a reasonable approximation for the MORB1 composition, which has a II value similar to North Arch Lavas with 49 wt% SiO₂. However, the MORB2 composition lies significantly below the line defined by North Arch lavas, so has a lower II value, and therefore a lower CO₂ solubility at a given SiO₂ content compared to the North Arch Lavas. Thus, VolatileCalc-Basalt likely overpredicts the solubility of CO₂ in this melt composition.

Furthermore, VolatileCalc-Basalt predicts that MORB2 dissolves ~1.7 times more CO₂ at a given pressure than MORB1. This is because MORB2 has 3.4 wt% less SiO₂ than MORB1, and VolatileCalc-Basalt predicts that CO₂ solubility increases drastically with decreasing SiO₂. However, if the full II expression of Dixon (1997) were used, MORB1 and MORB2 would have very similar CO₂ solubility, as they have similar II values despite different SiO₂ contents. P-2006 also predicts that MORB2 dissolves 5-6× more CO₂ at 0.4 kbar, and 1.9× more at 4 kbar than MORB1. In contrast, the models of S-2014, IM-2012, and MagmaSat predict that MORB1 and MORB2 dissolve similar amounts of CO₂ (MORB2/MORB1=~0.99×, ~1.125× and ~0.81–0.89× respectively). These three more recent models utilize significantly larger basaltic calibration datasets to parametrize the effect of multiple oxide species melt on CO₂ solubility (Fig. 2), so likely predict more realistic solubility relationships than VolatileCalc-Basalt and P-2006. CO₂ solubility in melt compositions that do not follow a similar trajectory in II-SiO₂ space as the North Arch Lavas (Fig. 7d) is unlikely to be accurately predicted by VolatileCalc-Basalt.

There is also significant deviation between different models for Etna melts (Fig. 7c), which is far greater than that observed for H₂O (Fig. 5). The A-2019 model, developed specifically for the composition of Etna magmas, predicts much higher CO₂ solubility at a given pressure than VolatileCalc-Basalt, S-2014, and MagmaSat, while P-2006 and IM-2012-H and -A follow similar trajectories to A-2019. The success of both IM-2012 models likely reflects the large number of alkaline compositions in their calibration dataset, including some from Etna. VolatileCalc-Basalt predicts the lowest CO₂ solubility (factor of 0.5-0.6× that of A-2019). The calculated II value for Etna lies significantly
above the line defined by North Arch lavas (so VolatileCalc-Basalt predicts lower CO$_2$ solubility; Fig. 7d). However, even the full $\Pi$ expression of Dixon (1997) is unlikely to be successful, because alkaline magmas show considerable variation in CO$_2$ solubility at a given $\Pi$ value (Allison et al., 2019). As S–2014’s expression for CO$_2$ solubility incorporates a $\Pi^*$ term very similar to the $\Pi$ term of Dixon (1997), the deviation of this model from that of A–2019 (0.6–0.7×) may also result from variations in CO$_2$ solubility that are not incorporated by this simplified melt composition parameter (Allison et al., 2019).

Interestingly, MagmaSat also underpredicts CO$_2$ concentrations at a given pressure relative to A–2019 by a factor of 0.6–0.7×, despite incorporating CO$_2$ experiments on Etna basalts from Lesne, Scailllet, Pichavant, and Beny (2011) and Iacono-Marziano et al. (2012) in its calibration dataset. This is a good example of the main pitfall of comprehensive models such as MagmaSat and P–2006 which can predict volatile solubility across the entire range of natural silicate melt compositions (Papale et al., 2006). For any specific melt composition, the model is highly unlikely to be as well tuned as models calibrated on melts from a specific volcanic center (e.g., Allison et al., 2019) or heavily weighted towards a specific region of compositional space (e.g., Iacono-Marziano et al., 2012, for alkaline basalts). Tuning MagmaSat to provide a better fit to Etna would almost certainly cause this model to show larger discrepancies for experiments conducted on different melt compositions.

4.3 Silicic Compositions

4.3.1 Pure H$_2$O

All five H$_2$O models calibrated for silicic magmas (MagmaSat, P–2006, L–2005, VolatileCalc-Rhyolite and M–1998) predict very similar H$_2$O concentrations at <1–1.5 kbar for the Mono Craters rhyolite composition (Fig. 8a, Table 1). At higher pressures, P–2006, and to a much lesser extent L–2015, show a smaller increase in H$_2$O solubility with pressure compared to MagmaSat, M–1998 and VolatileCalc-Rhyolite (the difference in H$_2$O solubility between models reaches ~4 wt% at 5 kbar). As discussed for basalt, the large discrepancies at higher pressures and H$_2$O contents likely results from an absence of experimental constraints because of challenges related to quenching melts with >6-9 wt% H$_2$O (Gavrilenko et al., 2019).

L–2005 and VolatileCalc-Rhyolite have no compositional dependence, so follow identical trajectories for the Mono Craters rhyolite and the peralkaline Aluto rhyolite (Fig. 8b). MagmaSat and M–1998 also show very similar trends for these two compositions. In contrast, the P–2006 model predicts higher H$_2$O concentrations at <2 kbar than the other models for Aluto (and higher H$_2$O concentrations than predicted for Mono Craters). At >3 kbar, the P–2006 Aluto model shows a rapid reduction
**Figure 8.** Relationship between pure H$_2$O (a-b) and pure CO$_2$ (c-d) solubility and pressure at 800°C for the Mono Lake rhyolite from Liu et al. (2005), and a peralkaline Rhyolite from Aluto in the East African Rift (Iddon and Edmonds, 2020). Models extrapolated beyond their calibrated pressure range are shown as dashed lines, with the colored star marking the recommended upper calibration limit.

in gradient, and predicts lower H$_2$O concentrations than L–2005 (although the discrepancy between these models at > 2 kbar is much smaller than for Mono Craters).

### 4.3.2 Mixed H$_2$O-CO$_2$

Differences in the treatment of H$_2$O-CO$_2$ mixing for rhyolitic melts are more subtle than for basaltic compositions (Fig. 9). Unlike for basalts, the differences in isobar positions mostly result from large differences between the pure CO$_2$ solubility predicted by different models rather than treatment of mixing. Only VolatileCalc-Rhyolite assumes ideal mixing of H$_2$O-CO$_2$, causing isobars to have a negative gradient at all pressures. L–2005 accounts for non-ideal mixing through the inclusion of a
term for the mole fraction of $H_2O$ in the fluid in their expression for $CO_2$ solubility (equation 5). This empirical representation of non-ideality causes isobars to exhibit a prominent peak at low $H_2O$ contents (Fig. 9). MagmaSat and P–2006, which include a fully non-ideal treatment of mixing, show a far less prominent peak than L–2005. The slight up-tick in the P–2006 isobars at very low $H_2O$ contents is much smaller than for basaltic compositions (e.g., Fig. 6).

![Figure 9. 1 and 4 kbar isobars for Mono Craters (a) and the Aluto rhyolite (b) at 800°C.](Image)

4.3.3 Pure $CO_2$

Unlike the relatively good correspondence between rhyolite $H_2O$ models (particularly at low pressures), there is substantial divergence between $CO_2$ models at all pressures (Fig. 8c, d). For the Mono Craters rhyolite, VolatileCalc-Rhyolite predicts $\sim 1.8 \times$ more dissolved $CO_2$ than MagmaSat at 2 kbar, while Liu, P–2006 and MagmaSat plot relatively close to each other at <5 kbar. As VolatileCalc-Rhyolite and L–2005 have no compositional dependence, the model lines are identical for Mono Craters and Aluto. MagmaSat predicts that the Aluto composition has slightly higher $CO_2$ solubility at a given pressure compared to the Mono Lake composition (factor of $\sim 1.2 \times$), so the discrepancy between MagmaSat, VolatileCalc-Rhyolite and L–2005 is smaller for Aluto than Mono Craters.

The P–2006 model shows a substantially different trajectory for $CO_2$ vs. pressure compared to the other three solubility models for both rhyolite compositions, showing a strongly concave-up shape compared to the near linear trajectory of L–2005 and VolatileCalc-Rhyolite, and the slightly concave-up shape of MagmaSat (Fig. 8c–d). For Mono Craters, P–2006 predicts similar $CO_2$ solubility to MagmaSat at <4 kbar, but rapidly rises to higher $CO_2$ contents at higher pressures, predicting al-
most as much dissolved CO$_2$ as VolatileCalc-Rhyolite at $\sim$12 kbar (Fig. 8c). For Aluto, the curvature of the P–2006 model is even more prominent, predicting drastically lower CO$_2$ contents than all other models at <6 kbar, and then rapidly rising, predicting higher CO$_2$ solubility than even VolatileCalc-Rhyolite at $>9$ kbar (Fig. 8d). These large deviations between models, as well as the large errors on the interaction terms for CO$_2$ solubility in MagmaSat (Table 3) demonstrate that while H$_2$O solubility in rhyolites is well constrained, more work is required to determine the effect of melt composition on CO$_2$ solubility at a range of pressures and temperatures.

### 4.4 Comparisons between Basalts and Rhyolites

In this section, we briefly discuss the differences in solubility between basalts (using the MORB1 composition) and rhyolites (using the Mono Craters composition). To differentiate the effect of melt composition from temperature (because basaltic melts tend to be hotter), we perform calculations at 800 and 1000°C for Mono Craters, and 1000°C and 1200°C for MORB1.

When all solubility models are compared (4 applicable to rhyolites, 6 to basalts), there is substantial overlap between curves calculated for MORB1 at 1200°C and Mono Craters at 800°C (compare Fig. 11a vs. Fig. 8a). To get around this problem of large differences between models, we compare the predictions from the three models which can be applied to both Rhyolites and Basalts: MagmaSat (Fig. 10a-b), P–2006 (Fig. 10c-d) and VolatileCalc-Basalt and -Rhyolite (Fig. 10e-f).

MagmaSat and VolatileCalc (Rhyolite vs. Basalt) predict that Mono Craters dissolves more H$_2$O than MORB1, even if these melts are at the same temperature (1000°C). In MagmaSat, the difference in solubility between Basalt and Rhyolite is enhanced by the fact that basalts tend to be hotter (the curves get closer when compositions are compared at 1000°C). In contrast, P–2006 predicts that Mono craters at 800°C dissolves less H$_2$O than MORB1 at 1200°C, although their solubilities are nearly identical when compared at 1000°C.

MagmaSat and P–2006 predict that MORB1 dissolves more CO$_2$ than Mono Craters, with the difference increasing rapidly as a function of pressure. In contrast, VolatileCalc predicts that, at 800°C, Mono Craters dissolves more CO$_2$ than MORB1 at 1000°C or 1200°C, while the model predicts very similar CO$_2$ solubility when Mono Craters and MORB1 are both at 1000°C.

Overall, these comparisons demonstrate that at <5 kbar, the difference in solubility between basalts and rhyolites is relatively subtle and easily overwhelmed by differences in predictions from different solubility models, particularly given some models predict that solubility increases with temperature, and others predict the opposite (see section 5.3).
Figure 10. Comparison of solubility of basalt (represented by MORB1) to rhyolite (represented by Mono Craters) using MagmaSat (a-b), P-2006 (c-d) and VolatileCalc-Basalt and Rhyolite (d-e). Solubility curves are calculated for 1200°C and 1000°C for basalt, and 1000°C and 800°C for rhyolite.
5 Model Sensitivities

In this section, we explore the sensitivity of the different models to parameters such as temperature and variable proportions of H$_2$O and CO$_2$. Specifically, we consider how these inputs affect calculations of the pressure at which a melt inclusion was trapped (termed the saturation pressure). To calculate saturation pressures, the initial concentration of major and volatile elements as well as the temperature must be estimated at the time of melt inclusion entrapment. However, a number of processes, such as crystallization of the host mineral on the wall of the inclusion (termed post-entrapment crystallization, or PEC), growth of a vapour bubble or daughter phases within the inclusion, and diffusive re-equilibration with a changing carrier liquid composition can make it difficult to reconstruct initial major element and volatile contents (Lowenstern, 1995). Similarly, diffusive re-equilibration of the major elements in the melt inclusion and host mineral, as well as the errors associated with mineral-melt and melt-only thermometers, can lead to uncertainties in the entrapment temperature, which propagates into the saturation pressure. By investigating the effect of varying these parameters within realistic limits, insight can be gained into the uncertainties associated with estimating magma storage depths using melt inclusions.

5.1 Relationship between saturation pressure and dissolved H$_2$O content

Melt inclusion H$_2$O contents are vulnerable to diffusional re-equilibration with the melt surrounding the crystal (here termed the carrier melt), because of the fast diffusion rate of H$^+$ through silicate minerals (Portnyagin et al., 2008). H$^+$ diffusion is particularly fast in olivine (Gaetani et al., 2012), with melt inclusions losing significant amounts of water in hours to days (Bucholz et al., 2013). Thus, this discussion focuses on mafic compositions, where olivine-hosted melt inclusions are frequently analysed.

In relatively H$_2$O-poor mafic systems such as MORs and ocean islands (e.g., Hawai‘i), diffusive re-equilibration can increase melt inclusion H$_2$O contents if crystals are mixed into more H$_2$O-rich carrier melts (Hartley et al., 2015), or, more commonly, cause melt inclusion H$_2$O contents to drop if the crystal is in contact with a carrier melt that has degassed its H$_2$O upon eruption (Gaetani et al., 2012; Bucholz et al., 2013). To assess how uncertainty in initial H$_2$O contents translates into errors on saturation pressures, we calculate saturation pressures for the MORB1 melt composition with 200, 1000, and 3000 ppm CO$_2$ (representing melt inclusions trapped at low, medium and high pressures) for H$_2$O contents between 0–1.5 wt% (Fig. 11a-c). CO$_2$ contents are held constant while H$_2$O contents are varied, simulating the changes undergone by melt inclusions during diffusive re-equilibration (which strongly affects H$_2$O contents in the inclusion, but does not change the total CO$_2$ budget of the inclusion).
The relationship between saturation pressure and dissolved H$_2$O predicted by each solubility model is strongly dependent on the amount of CO$_2$ in the melt, and therefore the pressure. To quantify model sensitivity to H$_2$O, saturation pressures calculated at H$_2$O=1.5 wt% are divided by the saturation pressure calculated at H$_2$O=0 wt%, representing the possible discrepancy between the calculated saturation pressure and the real saturation pressure for melt inclusions which have undergone complete H$^+$ re-equilibration with a fully degassed erupted melt at 0 bar. At low pressures (200 ppm CO$_2$), all models show a decrease in calculated saturation pressure with decreasing H$_2$O contents, with entrapment pressures being 1.2–1.8× higher before complete H$_2$O-loss (Fig. 11a). MagmaSat shows the strongest sensitivity to H$_2$O content, and both IM–2012 models the weakest.

At moderate pressures (1000 ppm CO$_2$), loss of H$_2$O causes a significantly smaller decrease in saturation pressure for VolatileCalc-Basalt, P–2006 and S–2014 compared to the 200 ppm CO$_2$ scenario (Fig. 11b, 1.1–1.2×). Saturation pressures for 1000 ppm CO$_2$ calculated using MagmaSat and IM–2012–H first decrease, then increase with H$_2$O loss. This is because these models predict that the maximum CO$_2$ solubility occurs at H$_2$O contents at ∼0.5–1.25 wt% (see Fig. 6).

At higher pressures (3000 ppm CO$_2$), saturation pressures from VolatileCalc-Basalt, P–2006 and S–2014 only drop by ∼5-10% with progressive H$_2$O-loss, while saturation pressures continually increase with progressive H$_2$O-loss for IM–2012–H and -A and MagmaSat (because these models predict that maximum CO$_2$ solubility is found at H$_2$O contents >1.5 wt% at these pressures; Fig. 6).

Within a given suite of MORB or OIB melt inclusions, the range of measured H$_2$O contents, and the uncertainty involved in reconstructing initial H$_2$O contents following diffusional re-equilibration, is likely significantly smaller than the 1.5 wt% H$_2$O considered here (Koleszar et al., 2009; Sides et al., 2014a; Wieser et al., 2021). Thus, except at low pressures (<1 kbar), uncertainties in saturation pressures due to diffusive re-equilibration of H$_2$O in relatively anhydrous systems are likely comparable to the analytical errors associated with the measurements of volatile species by FTIR or SIMS (±5–10%), errors on each solubility model (∼10–20%), and significantly smaller than the differences between solubility models.

The higher H$_2$O contents of melt inclusions from subduction zones (∼2–6 wt%; Plank et al., 2013) mean that substantially more H$_2$O can be lost following diffusive re-equilibration with a degassed carrier melt upon eruption. Additionally, arc melt inclusions are vulnerable to diffusive re-equilibration during crustal storage. This is because these relatively hydrous magmas saturate in a H$_2$O-rich fluid at high pressures in the crust. Thus, as a melt and its crystal cargo ascends from a deeper storage reservoir to a shallower storage reservoir, significant quantities of H$_2$O will be degassed and the H$_2$O contents of melt inclusions will rapidly diffusively re-equilibrate with the new carrier melt.
composition (Gaetani et al., 2012). Even if samples are rapidly quenched upon eruption (preventing syn-eruptive H$_2$O diffusion), only the volatile contents of melt inclusions trapped in the shallowest storage reservoir can be reliably converted into saturation pressures (Gaetani et al., 2012). This contrasts with more H$_2$O-poor systems such as mid-oceanic ridges and oceanic islands, where H$_2$O only degasses in the upper few hundred metres of the crust, so ascent to a shallower reservoir is not accompanied by a drop in melt H$_2$O contents (although diffusive re-equilibration could occur if the resident melts in the shallower reservoir have different H$_2$O contents; Hartley et al., 2015).

Figure 11. a-c) Relationship between saturation pressure and melt H$_2$O content for H$_2$O-poor melts (using the MORB1 composition at 1200°C). Three different melt CO$_2$ contents (200, 1000, 3000 ppm) are shown in part a, b and c respectively. The numbers on the graphs show the saturation pressure at 1.5 wt% H$_2$O (square symbol) divided by the saturation pressure at 0 wt% H$_2$O for each model. d-f) Sensitivity tests using the composition of a Fuego melt inclusion (Table 1) at 1000°C, and H$_2$O contents between 0–6 wt%. Note that the y scale for parts a-c is significantly smaller than parts d-f.

To investigate the effect of H$_2$O re-equilibration on melt inclusion saturation pressures in arcs, we repeat the sensitivity test described above, using the major element composition of a Fuego melt
inclusion with 49.7 wt% SiO$_2$ from Lloyd et al. (2013, Table 1) and H$_2$O contents between 0–6 wt% (Fig. 11d-f). For melt inclusions with 200 ppm CO$_2$, complete diffusive loss of H$_2$O may result in saturation pressures being underestimated by a factor of $\sim$5–10$\times$. Even for melts with 3000 ppm CO$_2$ (the highest pressure regime examined, and thus the best case scenario), diffusive loss can affect saturation pressures by factors of 0.98–1.6$\times$ (similar in magnitude to the sensitivity displayed by H$_2$O-poor melts at the lowest pressures; Fig. 11a vs. f). Only saturation pressures calculated in MagmaSat for the most CO$_2$-rich melts display variations with variable H$_2$O-loss similar in magnitude to analytical errors. Thus, it is extremely important to determine whether melt inclusions have undergone H$_2$O-loss during ascent to a shallower reservoir or syn-eruptive degassing before using saturation pressures to deduce magma storage depths in H$_2$O-rich volcanic systems.

5.2 Relationship between saturation pressure and dissolved CO$_2$ content

Estimating the initial CO$_2$ contents of melt inclusions is also challenging. While the total CO$_2$ content of the inclusion is not affected by diffusive re-equilibration, CO$_2$ may be partitioned from the melt phase into a vapour bubble. Cooling following melt inclusion entrapment is accompanied by the formation of a denser mineral phase from a less dense silicate melt, and differential thermal contraction of the melt and crystal. These processes cause the internal pressure of the inclusion to drop (Steele-Macinnis et al., 2011; Maclennan, 2017), driving the nucleation and growth of a vapour bubble. This may be enhanced by the diffusive loss of H$_2$O, which also causes a pressure drop in the melt inclusion because of the high molar volume but low molecular weight of H$_2$O (Gaetani et al., 2012; Aster et al., 2016). A drop in pressure, combined with a decrease in the solubility of CO$_2$ in the melt phase because of changes to the major element composition accompanying post-entrapment crystallization, causes CO$_2$ to partition strongly into the vapour bubble (L. R. Moore et al., 2015; Steele-Macinnis et al., 2011; Maclennan, 2017; Wieser et al., 2021). A number of recent studies have quantified the amount of CO$_2$ in vapour bubbles using Raman Spectroscopy, and demonstrated that between 15–99% of the total CO$_2$ budget of the inclusion may be held within the vapour bubble (Hartley et al., 2014; L. R. Moore et al., 2015; Wieser et al., 2021; Allison et al., 2021). This means that a large proportion of literature melt inclusion data, which only measured the CO$_2$ content of the glass phase, may have significantly underestimated initial CO$_2$ contents (and therefore saturation pressures).

In relatively H$_2$O-poor systems such as Hawai’i and Iceland, where melt inclusion CO$_2$ contents have the dominant control on saturation pressures (shown by the near horizontal slopes of most model isobars at low H$_2$O contents; Fig. 6), it is readily apparent that saturation pressures will be significantly underestimated if a CO$_2$-rich vapour bubble is not measured. In arcs, H$_2$O contents inferred from melt inclusions or mineral hygrometers are sometimes used to place first order constraints on
saturation pressures (e.g. Plank et al., 2013; Blundy & Cashman, 2005; Goltz et al., 2020). However, even in very H$_2$O-rich melts, the non-vertical orientation of isobars at high H$_2$O contents indicates that CO$_2$ contents still have an important role in determining the saturation pressure (Fig. 6). Additionally, only a very small number of studies have measured CO$_2$ in melt inclusion vapour bubbles from arc systems (L. R. Moore et al., 2015; Aster et al., 2016; Venugopal et al., 2020; Mironov et al., 2020). Thus, it is vital to determine the effect of CO$_2$ on saturation pressures in H$_2$O-rich systems.

Using a similar method to that for H$_2$O discussed above, we calculate saturation pressures for the composition of a Fuego melt inclusion from Lloyd et al. (2013) with varying CO$_2$ and H$_2$O contents. The mean melt inclusion glass CO$_2$ content from this melt inclusion suite was 340 ppm (range of 59–786 ppm). However, Raman analyses of vapour bubbles in the same sample set by L. R. Moore et al. (2015) reveals that 993–4776 ppm of CO$_2$ has migrated from the glass phase into the vapour bubble following melt inclusion entrapment. Thus, we calculate saturation pressures for CO$_2$ contents between 0–5000 ppm at 1000°C for 2, 4 and 6 wt% H$_2$O respectively (after Plank et al., 2013).

![Figure 12.](image_url)

**Figure 12.** Relationship between saturation pressures and melt CO$_2$ contents for H$_2$O-rich melts (using the composition of a Fuego melt inclusion at 1000°C; Table 1). Three different melt H$_2$O contents (2, 4 and 6 wt%) are shown in part a, b and c) respectively. The numbers on the graphs show the saturation pressure at 5000 ppm CO$_2$ (square symbol) divided by the saturation pressure at 0 ppm CO$_2$ (circle), and the saturation pressure at 1000 ppm (star symbol) divided by the saturation pressure at 0 ppm CO$_2$ (square) for each model.

S–2014 is most sensitive to CO$_2$ content, and IM–2012–H and -A the least sensitive. With increasing H$_2$O, the change in saturation pressure with increasing CO$_2$ becomes smaller, but is still significant (Fig. 12). For example, calculating a H$_2$O-only saturation pressure for a melt inclusion with H$_2$O=6 wt% in MagmaSat would underestimate magma storage depths by a factor of 1.5 if the melt inclusion had 1000 ppm CO$_2$, and a factor of 3 if the inclusion had 5000 ppm CO$_2$. For a melt inclu-
sion with 4 wt% H$_2$O, H$_2$O-only saturation pressures underestimate by a factor of $\sim$2.2 for 1000 ppm CO$_2$, and 5.4 for 5000 ppm CO$_2$. These variations in saturation pressure overwhelm the other errors associated with melt inclusion barometry (e.g., uncertainty in crustal density profiles, analytical errors associated with volatile measurements by FTIR or SIMS, differences between solubility models). Further investigation of the prevalence of CO$_2$-rich vapour bubbles in arc lavas is clearly required to have confidence in published barometric estimates in studies which did not measure the vapour bubbles, or used mineral hygrometers.

Saturation pressures in rhyolitic magmas are also very sensitive to melt CO$_2$ contents (Fig. 13). For example, saturation pressures calculated for 1000 ppm CO$_2$ vs. 0 ppm CO$_2$ differ by factors of 5.7–8.8$\times$ for 2 wt% H$_2$O, and 1.6–2$\times$ for 6 wt% H$_2$O. Even saturation pressures calculated for 300 ppm CO$_2$ (0.03 wt%) vs 0 ppm CO$_2$ are a factor of $\sim$2–3$\times$ higher for 2 wt% H$_2$O, and 1.2–1.3$\times$ higher for 6 wt% H$_2$O. The strong effect of CO$_2$ on saturation pressure is important to recognise when calculating saturation pressures using only melt H$_2$O contents, such as studies using mineral-melt hygrometers (Waters & Lange, 2013), or volatiles-by-difference methods to estimate H$_2$O contents of melt inclusions. It is also interesting to note that, to our knowledge, there are no published Raman measurements of CO$_2$ in vapour bubbles which grew after melt inclusion entrapment in dacitic-rhyolitic melt compositions. While the extremely low CO$_2$ contents of many rhyolitic melt inclusions are commonly interpreted to result from shallow crustal storage, it is becoming increasingly recognised that mafic melt inclusions with CO$_2$ below detection limit contain large quantities of CO$_2$ in the vapour bubble (Wieser et al., 2021). Thus, examination of vapour bubbles in melt inclusions from more silicic systems (e.g., Fig. 1 of Lowenstern, 2001) is likely warranted, to rule out the possibility that these melts crystallized at greater depths than calculated using measurements of CO$_2$ hosted in just the glass phase.
5.3 Sensitivity to Temperature

The temperature of the melt at the time of entrapment is another source of uncertainty when calculating saturation pressures, as melt or mineral-melt thermometers are relatively imprecise. For example, the liquid MgO thermometer presented in eq. 13 of Putirka (2008) has a standard error of $\pm 71^\circ C$, while the clinopyroxene–liquid thermometer presented in their equation 28b has a standard error of $\pm 48^\circ C$. Changes in the major element compositions of the melt inclusion during PEC and diffusive $H_2O$-loss can also introduce errors when estimating entrapment temperatures (as most thermometers are highly sensitive to the MgO and $H_2O$ content of the melt). Additionally, almost all solubility experiments are performed at supra-liquidus conditions, while melt inclusion formation must take place at sub-liquidus conditions, so extrapolation to lower temperatures is an unfortunate necessity.

To investigate the sensitivity of different models to temperature, we calculate the 0.5 and 2 kbar isobars for the MORB1 composition at 1000$^\circ C$ and 1400$^\circ C$. Only S–2014 shows no temperature dependency, because there is neither a temperature or fugacity term in their equations. Interestingly, there is considerable disagreement between the other models as to whether a hotter magma dissolves more or less volatiles. MagmaSat and P–2006 predict an increase in pure CO$_2$ solubility with increasing temperature, while VolatileCalc-Basalt and IM–2012–A and –H predict a much smaller decrease (Fig. 14a-c). In all models but IM–2012–A, isobars calculated for lower temperatures intercept the x axis (CO$_2$=0 wt%) at higher $H_2O$ contents, so the temperature dependency of $H_2O$ solubility is
opposite to that for CO\textsubscript{2} solubility. To visualize the effect of these trends on calculated saturation pressures for the MORBI\textsubscript{1} composition, the calculated saturation pressures for melts with volatile contents represented by the yellow and cyan stars on Figure 14b are plotted against temperature (between 1000–1400°C; Fig. 14c-d). MagmaSat and P–2006 show the strongest temperature sensitivity, with a slope opposite to that of the more subtle changes predicted by VolatileCalc-Basalt and IM–2012.

The lack of consensus as to whether increasing temperature increases or decreases the solubility of H\textsubscript{2}O and CO\textsubscript{2} indicates that this effect is relatively subtle, and overwhelmed by analytical errors associated with measuring experimental products (and other sources of experimental scatter; e.g., Fig. 16a-b). This makes it very difficult for empirical models to fully constrain the temperature sensitivity, particularly given that the experiments conducted by any given study are usually performed at a single temperature. For example, all the experiments used to calibrate the VolatileCalc-Basalt model were conducted at 1200°C, so the temperature dependency of this model results from the fugacity function, and 1/T terms from Dixon et al. (1995), rather than experimental observations.

IM–2012 is calibrated on experiments mostly performed at 1200–1300°C (with a few spanning 1100–1400°C). Their empirical expressions contain $C_{H_2O} \times \frac{P}{T}$ and $C_{CO_2} \times \frac{P}{T}$ terms in their expressions for H\textsubscript{2}O and CO\textsubscript{2} solubility respectively (equation 11 and 15), where $C_{H_2O}$ and $C_{CO_2}$ are empirically-derived constants, and P and T are pressure and temperature. In the hydrous model, $C_{H_2O}$ is negative (-0.02±0.02), so H\textsubscript{2}O solubility increases with increasing temperature, while in the anhydrous model $C_{H_2O}$ is positive (0.02±0.02), so H\textsubscript{2}O solubility decreases with increasing temperature (Fig. 14). As the magnitude of these coefficients is small, the temperature effect on H\textsubscript{2}O solubility is small, and only visible at higher pressures (because of the P part of these terms; Fig. 14a vs. c). It is also worth noting that these coefficients in both models are within error of zero, showing that the experimental dataset used to calibrate this model showed very little evidence for a change in H\textsubscript{2}O solubility with temperature. In both the hydrous and anhydrous models, $C_{CO_2}$ is positive (0.12±0.02 and 0.14±0.02 respectively) and larger in magnitude than $C_{H_2O}$, so CO\textsubscript{2} solubility decreases with increasing temperature (see Fig. 14c).

The approach taken by S–2014 and A–2019 is an interesting alternative when constructing solubility models. While S–2014 is calibrated on experiments conducted between 1200–1300°C for CO\textsubscript{2}, and 1200–1250°C for H\textsubscript{2}O, their solubility equations contain no temperature or fugacity term. Instead, these authors suggest that the H\textsubscript{2}O model should ideally be used between 1150–1250°C. They perform additional tests on experiments between 1050–1400°C not used in the calibration, and show that their model predicts H\textsubscript{2}O solubility within ±10% for 78% of experiments for this wider temperature range. The S–2014 testing dataset for CO\textsubscript{2} only has a slightly different temperature range than the calibra-
tion dataset (1170–1250°C vs. 1200–1250°C), so they do not suggest an expanded temperature range for CO₂. Similarly, the spreadsheet for A–2019 (and the implementation of this model in VESIcal) performs all calculations at 1200°C, regardless of the user-input temperature. Allison et al. (2019) suggest that this approach is likely valid between 1000–1400°C.

Unlike empirical models, the temperature sensitivity of P–2006 and MagmaSat arises from the entropy differences between melt and fluid components. Given the limited experimental evidence for changes in solubility with temperature, the directionality inferred by thermodynamical models is more likely to be correct, as the relative entropy differences between components are easier to constrain than deconvoluting subtle differences between the solubility of CO₂ and H₂O in experiments run at different temperatures, and because all empirical models are being extrapolated to lower temperatures than the supra-liquidus experiments used to calibrate them. However, the fact that P–2006 doesn’t account for volatile speciation for either H₂O or CO₂, and MagmaSat doesn’t account for the two H₂O species in the melt means that these thermodynamic temperature sensitivities may also need further interrogation.

Overall, although the differences between models is interesting, and important to recognise when extrapolating beyond the range of the calibration dataset, the uncertainty that temperature variations introduce to saturation pressure calculations are relatively small for mafic melts. For example, an uncertainty in the initial entrapment temperature of ±100°C introduces an uncertainty similar to that associated with in-situ measurements of melt inclusion volatile contents (±5–10%; Fig. 14).
Similarly, temperature sensitivity in rhyolitic melts was evaluated by calculating isobars at 0.5 and 2 kbar for 700 and 900°C using the Mono Craters rhyolite composition. As for the basaltic example, the directionality and magnitude of effect of temperature on saturation pressures for melts with volatile contents indicated by the colored stars is shown in Fig. 15c-d for temperatures between 700 and 1000°C. VolatileCalc-Rhyolite shows the strongest temperature sensitivity, predicting that the solubility of CO$_2$ decreases with increasing temperature. L–2005 also predicts decreasing CO$_2$ solubility with increasing temperature, although this effect is smaller than in VolatileCalc-Rhyolite. Decreasing solubility of molecular CO$_2$ with increasing temperature was demonstrated experimentally by Fogel
and Rutherford (1990). In contrast, the two thermodynamic models, P–2006 and MagmaSat, predict that pure CO$_2$ solubility increases with increasing temperature. VolatileCalc-Rhyolite also predicts that H$_2$O solubility decreases with increasing temperature, but the effect is smaller than for CO$_2$.

Figure 15. Sensitivity of saturation pressures for the Mono Craters rhyolite to temperature. a-b) Isobars calculated for different solubility models at 700 and 900°C and 0.5 and 2 kbar. c) Relationship between saturation pressure and temperature for a melt with 1000 ppm CO$_2$, 1 wt% H$_2$O (yellow star in b), d) 100 ppm CO$_2$, 6 wt% H$_2$O (cyan star in b). Ratios of saturation pressures at 900°C (square symbol) to 700°C (circle symbol) are shown on the figure.

It is noteworthy that the temperature sensitivity of CO$_2$ solubility predicted by L–2005 and VolatileCalc-Rhyolite is much greater than that shown by any of the basaltic models (Fig. 14d-e vs. Fig 15d-e), and significant considering other sources of error associated with saturation pressure calculations. MagmaSat and P–2006 also show a far greater sensitivity to H$_2$O solubility between
700–850°C in rhyolites than any of the basaltic models between 1000–1400°C (Fig. 14f vs. Fig 15f), although the sensitivity decreases between 800–900°C.

Given the contrasting behaviour of empirical and thermodynamic models, and the relatively strong effect of temperature in rhyolitic melts, we suggest that users proceed with caution when extrapolating empirical models to temperatures significantly lower or higher than the calibration temperature of each model. It may be best to use empirical models at the calibration temperature (e.g., 1200°C for VolatileCalc-Basalt, 850°C for VolatileCalc-Rhyolite), which is the approach used by in the models of S–2014 and A–2019, rather than introduce a temperature sensitivity with the wrong sign. This is discussed in further detail for VolatileCalc-Rhyolite in section 6.

6 Intermediate Compositions

In this section, we compare the predictions of different solubility models for intermediate melt compositions (andesites to dacites). Lavas with these compositions are dominant within subduction zones, and volcanoes erupting these compositions are extremely hazardous. Yet, there is a notable paucity of solubility experiments for andesitic and dacitic compositions relative to basalts and rhyolites (Fig. 2; King & Holloway, 2002; Botcharnikov et al., 2006). This section builds on the sensitivity tests performed in section 5 to evaluate possible discrepancies between model outputs and experimental constraints.

The calibration dataset of MagmaSat has the broadest coverage of andesitic-dacitic compositions of all the models described here (although it is far from extensive). While there are a number of pure H₂O experiments, MagmaSat only includes one pure CO₂ experiment on an andesitic melt (conducted at 1 GPa; King & Holloway, 2002), and no pure CO₂ experiments on dacitic melts. Similarly for mixed H₂O-CO₂, the calibration dataset for andesitic melts includes only four experiments from King and Holloway (2002), 21 from Botcharnikov et al. (2006) and three from Botcharnikov et al. (2007). Dacitic liquids are represented by the 12 experiments on mixed H₂O-CO₂ solubility by Behrens et al. (2004). As the P–2006 model had a decade fewer experimental constraints available for calibration, it only includes the one pure CO₂ and four mixed H₂O-CO₂ andesitic experiments of King and Holloway (2002). The IM–2012 model includes two pure H₂O experiments that lie within the andesite field on a TAS diagram, but no H₂O-CO₂ experiments, and no experiments in the dacitic field. None of the other models contain any andesitic or dacitic melts in their calibration datasets.

6.1 Comparing solubility models to experimental products

The suitability of different solubility models for andesitic-dacitic compositions can be evaluated by calculating isobars using the melt compositions, pressures and temperatures of different experi-
tal studies, and comparing these isobars to measurements of dissolved volatile contents in experimental products (similar to the method used in the supplementary material of Ghiorso & Gualda, 2015). The 2 and 5 kbar andesitic experiments of Botcharnikov et al. (2006) are shown in Fig. 16a-b, the 1, 2, and 5 kbar dacitic experiments of Behrens et al. (2004) are shown in Fig. 16c-e, and the 10 kbar andesite experiments of King and Holloway (2002) are shown in Fig. 16f. Additional isobar diagrams for the 3–12 kbar basaltic-andesite experiments of Mangan et al. (2021) are shown in the supporting information.

Isobar diagrams show that S–2014 significantly underpredicts CO$_2$ for all experiments except the most H$_2$O-poor composition of King and Holloway (2002). L–2005, VolatileCalc-Basalt and -Rhyolite also mostly plot to lower CO$_2$ contents than experimental products. IM–2012–H and –A do a reasonable job of recreating the most H$_2$O-poor experiments at <5 kbar, but curve rapidly down to intercept the x axis at lower H$_2$O contents than experimental products and other models. MagmaSat is a good match to experimental data in Fig. 16d, e and f, but plots to lower CO$_2$ contents than experiments in Fig. 16 a, b and c. Using Fe$^{3+}$ proportions best representing the experimental conditions, P–2006 only passes through experimental data on Figs. 16a-b, and e, and plots to significantly lower CO$_2$ contents than experiments (lower than MagmaSat) on Fig. 16c and d. P–2006 is a better match to most experiments if Fe$^{3+}$/Fe$_T$=0, but overpredicts CO$_2$ solubility at 10 kbar for all redox states. Interestingly, none of the available models recreate the near-flat trajectory of dissolved CO$_2$ contents with increasing H$_2$O from Botcharnikov et al. (2006).

The underprediction of CO$_2$ solubility by S–2014 is a good example of the dangers of extrapolating models accounting for the effect of melt composition using empirical expressions beyond the compositional range of the calibration dataset. The S–2014 model expresses CO$_2$ solubility as a function of the composition parameter, Π$^*$, with CO$_2$ solubility increasing as an exponential function of Π$^*$ at a given pressure (equation 16). The melt compositions for the three sets of experimental studies shown in Fig. 16 all plot to much lower Π$^*$ values than any of the melts in the calibration dataset (orange diamonds; Fig. 17a). These low Π$^*$ values mean that the S–2014 model predicts that these melts have very low CO$_2$ solubilities. However, CO$_2$ solubility for melts with Π$^*$ values outside the range of the calibration dataset may not follow the same exponential function of this parameter as melt compositions within the calibration range. Additionally, the exponential dependency of CO$_2$ solubility on Π$^*$ incorporated by S–2014 likely breaks down in more evolved melts, because Π$^*$ represents the ability of the melt to form carbonate-bearing species, while more evolved melts contain increasing proportions of molecular CO$_2$ (Shishkina et al., 2014). For example, the proportion of molecular CO$_2$ to carbonate
species varies between 0–4 wt% in the experiments of Botcharnikov et al. (2006), and 3–30 wt% in the dacitic experiments of Behrens et al. (2004).

VolatileCalc-Basalt and VolatileCalc-Rhyolite underpredict CO$_2$ solubility for all intermediate experiments, with VolatileCalc-Rhyolite predicting lower pure CO$_2$ solubility than VolatileCalc-Basalt. This is noteworthy, because many publications have calculated saturation pressures for andesitic and dacitic melts using VolatileCalc-Rhyolite (e.g., Blundy et al., 2006; Atlas et al., 2006; Cassidy et al., 2015; Koleszar et al., 2012). In the original publication, Newman and Lowenstern (2002) state: “because many andesites contain rhyolitic interstitial melt, VolatileCalc may also be applicable to these intermediate compositions.” However, this should not be taken to suggest that VolatileCalc-Rhyolite is safely extrapolated to andesitic-dacitic melts. Instead, this statement is referring to the fact that many lavas with andesitic bulk compositions have rhyolitic groundmass/melt inclusion compositions (e.g., Tamura & Tatsumi, 2002; Reubi & Blundy, 2009), where VolatileCalc-Rhyolite may be applicable. Interestingly, VolatileCalc-Rhyolite isobars calculated for the reference temperature of this model (850°C) rather than the experimental temperature are a good match to experimental data at < 5 kbar (see Supplementary Fig. S3). Thus, the main failure of this model in intermediate compositions appears to result from the fact that this model is extremely sensitive to temperature, and these melts have much higher temperatures than the experiments used to calibrate this model. This supports our suggestion in section 4.2.1 that it may be better run models at their reference temperature, rather than extrapolate beyond the calibration range.

It is also worth noting that all these experimental products have negative values of the Dixon Π parameter (equation 2). As highlighted by G. Moore (2008), the full Π expression of Dixon (1997) cannot be used to calculate CO$_2$ solubility in calc-alkaline lavas, because Π is negative (yielding a negative solubility of CO$_2$; Fig. 17).

Differences between experimental data and isobars from IM–2012–H and -A are more complicated, because the discrepancies between models and experiments are very dependent on the pressure. For example, at 1 kbar, IM–2012–H and -A underpredict H$_2$O and CO$_2$ solubility relative to the experiments of Behrens et al. (2004), while at 5 kbar, they overpredict CO$_2$ solubility, but underpredict H$_2$O solubility (Fig. 16c vs.e). These discrepancies likely reflect this model being extrapolated towards the limits of its calibration dataset in terms of both pressure (most experiments were conducted at <5 kbar) and melt composition (Fig. 17, see the next section for more discussion).
Figure 16. Assessing model fits to experimental data for andesitic and dacitic compositions a-b) Experiments from Botcharnikov et al. (2006). Isobars were calculated for 1200°C and Fe\(^{3+}\)/Fe\(^T\)=0.2 (\(\sim\)QFM+1.5). P–2006 isobars are also shown for Fe\(^{3+}\)/Fe\(^T\)=0.7 (the upper estimate of Fe\(^{3+}\)/Fe\(^T\) in experimental products; dash-dotted line), and Fe\(^{3+}\)/Fe\(^T\)=0 (dotted line) c-e) Experiments from Behrens et al. (2004). Isobars are shown for Fe\(^{3+}\)/Fe\(^T\)=0.59 (\(fO_2\) of QFM+4.7). For P–2006 and MagmaSat, isobars are also shown for Fe\(^{3+}\)/Fe\(^T\)=0. f) Experiments from King and Holloway (2002). Isobars are shown for Fe\(^{3+}\)/Fe\(^T\)=0.54. For P–2006 and MagmaSat, isobars are also shown for Fe\(^{3+}\)/Fe\(^T\)=0. VolatileCalc-Rhyolite isobars are not shown, as the spreadsheet doesn’t calculate isobars above 5 kbar.

Error bars on all plots shows the 2\(\sigma\) uncertainties from measurements of volatile contents in experimental products.

Fe\(^{3+}\)/Fe\(^T\) ratios were calculated from author-stated buffers using MELTS for Excel (Gualda & Ghiorso, 2015).
**Figure 17.** Comparing the calibration datasets of S–2014 and IM–2012 to the andesitic experiments of Botcharnikov et al. (2006), (B2006), Behrens et al. (2004), (B2004) and King and Holloway (2002), (K2002) shown in Fig. 16, and the melt inclusions from Popocatépetl (Atlas et al., 2006) and Soufrière Hills (Cassidy et al., 2015) shown in Fig. 18. Anhydrous molar fractions are used to calculate compositional parameters in parts c-f, because when accounting for discrepancies between isobars (e.g., on Fig. 18, the H₂O content and therefore hydrous cation fraction varies as a function of the pressure).
The fact that no model passes through all available experiments demonstrates that further investigation of solubility in andesitic to dacitic melts is warranted. Using representative experimental $\text{Fe}^{3+}/\text{Fe}_T$ ratios, MagmaSat is the most accurate model, predicting dissolved volatile contents within $\sim 20\%$ of experimental products (considering reported error bars on dissolved volatile contents).

The extreme sensitivity to the $\text{Fe}^{3+}/\text{Fe}_T$ ratio makes it very difficult to assess the accuracy of the P–2006 model (particularly given the relatively large uncertainties in the oxygen fugacity of experimental run products; King & Holloway, 2002; Botcharnikov et al., 2006). In all of the examples shown, P–2006 isobars calculated for $\text{Fe}^{3+}/\text{Fe}_T=0$ are a better fit to the experimental data than isobars calculated using estimates of the $\text{Fe}^{3+}/\text{Fe}_T$ ratio of experiments. This suggests that, in relatively oxidising intermediate melts, P–2006 is overestimating the effect of $\text{Fe}^{3+}$ species on volatile solubility. It is noteworthy that P–2006 is a particularly poor match to the high pressure andesitic experiments of King and Holloway (2002), despite the fact that these are the only intermediate experiments in the calibration dataset of this model.

As all the experiments shown were used to calibrate MagmaSat, similar analysis applied to new experimental data on andesitic compositions when it becomes available will provide further constraints on the accuracy of this model. Isobars for recent solubility experiments on a basaltic-andesite at 4–12 kbar by Mangan et al. (2021) are shown in the Supplementary Fig. S4 and S5. MagmaSat is a much better fit to this new data than P–2006, with experiments performed 400–815 MPa plotting within error of calculated MagmaSat isobars. In contrast, P–2006 overpredicts CO$_2$ solubility (regardless of $\text{Fe}^{3+}/\text{Fe}_T$) at <600 MPa.

6.2 Case Study: Intermediate melt inclusions

To assess the impact of these model differences on the depths of magma storage reservoirs inferred from melt inclusions in volcanic arcs, we calculate saturation pressures using a variety of models for two suites of melt inclusions with andesitic-dacitic liquid compositions: 1) 34 melt inclusions from Volcán Popocatépetl, Mexico with 55.7–73.4 wt\% SiO$_2$ (Figs. 2, 18a; Atlas et al., 2006), and 2) 8 melt inclusions from Soufrière Hills Volcano, Montserrat with 58.7–68.0 wt\% SiO$_2$ (Figs. 2, 18b; Cassidy et al., 2015). Both studies calculated saturation pressures (and therefore magma storage depths) using VolatileCalc-Rhyolite.

Cumulative frequency distributions for the Popocatépetl melt inclusions (Fig. 18a) shows that P–2006 predicts the lowest saturation pressures, and S–2014 the highest, with MagmaSat, IM–2012–A, VolatileCalc-Basalt, VolatileCalc-Rhyolite, and IM–2012–H lying in-between these model extremes. Based on our analysis in the previous section suggesting that MagmaSat is the best calibrated model for intermediate melt compositions, we ratio saturation pressures from each model to those determined
using MagmaSat (allowing model differences to be quantified). Additionally, because MagmaSat is a thermodynamic model that has been shown to work well for basaltic and rhyolitic compositions, it is effectively being interpolated to andesitic–dacitic compositions which are not represented in its calibration dataset (i.e., these melt inclusion compositions), rather than extrapolated (as for empirical models such as S–2014, VolatileCalc–Basalt, and IM–2012 which are primarily calibrated on more mafic melt compositions; Fig. 2).

The median saturation pressure for Popocatépetl melt inclusions calculated for VolatileCalc–Rhyolite is 1.26× higher than for MagmaSat. The median for MagmaSat is ∼1.2× higher than for P–2006. The deviation relative to MagmaSat increases with SiO$_2$ content for VolatileCalc–Rhyolite (∼1.1 to 1.5×), VolatileCalc–Basalt (∼1 to 1.35×), and L–2005 (∼1 to 1.4×). The deviation between S–2014 and MagmaSat is very large, increasing from a factor of 2 at ∼55 wt% SiO$_2$ to ∼5.5× at 72.5 wt% SiO$_2$. The deviation between IM–2012–H and -A and P–2006 vs. MagmaSat shows no clear correlation with SiO$_2$ content (Supporting Fig. S6). The presence of discrepancies which correlate with melt composition is particularly concerning, because it means that choice of an inappropriate solubility model may introduce systematic error into a dataset as a function of melt inclusion composition.
MagmaSat predicts the lowest saturation pressures for the Soufrière Hills melt inclusions, with L–2005, P–2006, and VolatileCalc-Rhyolite predicting reasonably similar pressures. VolatileCalc-
Basalt, S–2014 and IM–2012–A and –H are offset to higher pressures. Similar to the results for Popocatépetl, the ratio of saturation pressures for Soufrière Hills melt inclusions relative to MagmaSat for L–2005 (~1 to 1.1×), VolatileCalc-Basalt (~1.25 to 1.35×) and VolatileCalc-Rhyolite (~1.05 to 1.17×) increase with increasing SiO$_2$ (Supporting Fig. S7). Unlike Popocatépetl melt inclusions, the discrepancy between IM–2012 and MagmaSat increases dramatically with increasing SiO$_2$ (from ~1.5 to 2.5× for anhydrous, and 2 to 4.5× for hydrous), while that for S–2014 shows a weak negative correlation with SiO$_2$ (from ~1.7–1.35×).

The large discrepancies shown by IM–2012 (for Soufrière Hills) and S–2014 (for Popocatépetl) are best understood by examining an isobar calculated for the inclusion showing the largest deviation relative to MagmaSat in each dataset (SiO$_2$=72.4 wt% for Popocatépetl, and 67.6 wt% for Soufrière Hills; yellow pentagon on Fig. 18c-d) at the saturation pressure obtained from MagmaSat. For both sets of inclusions, S–2014 isobars intercept the x axis at similar H$_2$O contents to other models, but intercept the y axis at drastically lower CO$_2$ contents. This underestimation of CO$_2$ solubility likely results from the fact that the Π$^*$ values of these inclusions lie well below that of the calibration dataset (as discussed for andesitic experiments; Fig. 17b).

In contrast, IM–2012 predicts very high pure CO$_2$ solubility, and low pure H$_2$O solubility relative to the other models. Both IM–2012 models express the compositional dependence of H$_2$O solubility in terms of the parameter NBO/O, with melts with higher NBO/O values having higher H$_2$O solubility (equation 15). Both sets of melt inclusions possess much lower NBO/O values than the calibration dataset of IM–2012 (Fig. 17c-d). This is problematic, because the empirical relationship between NBO/O and pure H$_2$O solubility incorporated by IM–2012 has not been validated for these melt compositions. For example, Shishkina et al. (2014) show that IM–2012 drastically overestimates H$_2$O solubility in their basanite and nephelinite melt compositions. They point out that while IM–2012 conclude that there is only a small effect of melt composition on H$_2$O composition, the IM–2012 model ends up showing a strong sensitivity to melt composition when extrapolated to the high NBO/O ratio in their basanite and nephelinite melts (and we invoke a similar explanation for the lower NBO/O ratios in melt inclusions discussed here).

The IM–2012 expression for CO$_2$ solubility is more complicated, containing terms for the Agpaitic index (AI), NBO/O, $X_{Na_2O+K_2O}$, $X_{FeO+MgO}$ and $X_{H_2O}$ (equation 11). These two sets of melt inclusions have higher AI, lower NBO/O ratios, similar values of $X_{Na_2O+K_2O}$, and lower $X_{FeO+MgO}$ values than the calibration dataset (Fig. 17d-f). While the effect of NBO/O is more convoluted because it also affects the solubility of H$_2$O (which feeds back into the expression for CO$_2$), it is readily apparent that the positive coefficient attached to the AI term combined with the negative coefficient
attached to the MgO+FeO term causes this model to predict higher CO$_2$ solubilities than the calibration dataset for the andesitic-dacitic melt inclusions considered here.

The discrepancy between isobars for S–2014 and IM–2012 relative to MagmaSat are relatively similar for the Popocatépetl and Soufrière Hills melt compositions, while discrepancies for saturation pressures differ markedly (Fig. 18a-b vs. c-d). This is because the volatile contents of Popocatépetl melt inclusions are significantly more CO$_2$-rich (∼0.02–0.15 wt% and higher), and H$_2$O-poor (∼1–3 wt%) than Soufrière Hills melt inclusions (<0.04 wt% CO$_2$ and 5–6 wt% H$_2$O). For this reason, calculated saturation pressures for Popocatépetl melt inclusions are sensitive to the treatment of both CO$_2$ and H$_2$O in solubility models (Fig. 18b), while those for Soufrière Hills melt inclusions are mostly sensitive to pure H$_2$O solubility. Thus, S–2014 overestimates saturation pressures for Popocatépetl melt inclusions because this model drastically underestimates the solubility of pure CO$_2$. In contrast, S–2014 only slightly underestimates H$_2$O solubility relative to MagmaSat, so only slightly overpredicts saturation pressures for H$_2$O-rich Soufrière Hills melt inclusions. The discrepancy between IM–2012 and MagmaSat is much smaller for Popocatépetl relative to Soufrière Hills because, fortuitously, the IM–2012–H isobar intercepts the MagmaSat isobar at H$_2$O contents similar to these Popocatépetl melt inclusions. IM–2012–H and -A drastically underestimates the solubility of pure H$_2$O, so overestimates saturation pressures for the H$_2$O-rich Soufrière Hills melt inclusions relative to other models.

It is worth noting that Iacono-Marziano never intended their model to be applied to andesites, and when discussing the limitations of their model, they explicitly warn that their empirical expressions poorly incorporates the effect of MgO and FeO on CO$_2$ solubility because of the restricted range of these oxides in the calibration dataset. We have included this discussion as an extreme example of the danger of extrapolating empirical models beyond their calibration range. However, the sensitivity of this model to the FeO and MgO content of the melt also presents issues when applied to high MgO basaltic liquids. For example, Wieser et al. (2021) show that IM–2012 predicts higher saturation pressures for highly primitive (high MgO) melt inclusions from Kilauea Volcano relative to S–2014 and MagmaSat. This likely reflects the higher values of X$_{FeO+MgO}$ in these melt inclusions relative to the calibration dataset, which causes IM–2012 to predict lower CO$_2$ solubility (the opposite directionality to that seen for the intermediate melt inclusions discussed here).

The H$_2$O-rich nature of Soufrière Hills melt inclusions means that VolatileCalc-Rhyolite predicts much more similar saturation pressures to MagmaSat (1.07–1.15×; Supporting Figure S6) than for Popocatépetl (1.1–1.5× higher; Supporting Figure S5), because the main failure of VolatileCalc-Rhyolite for intermediate compositions at moderate to high temperatures (>850 C) is its prediction of pure CO$_2$ solubility (Figs. 16, 18c-d). The discrepancy for both VolatileCalc models and L–2005
relative to MagmaSat is significantly smaller than for S–2014 and IM–2012. This is because the solubility differences during evolution for basaltic to rhyolitic compositions are relatively small (30–40%) compared to the error associated with the extrapolating an empirical model far beyond its compositional range. Overall, this case study shows the importance of checking that the calibration dataset of a model contains melts similar to those in the sample set of interest, particularly if the effect of melt composition is parameterized empirically.

7 Best practices for data visualization and curation

7.1 Isobar Diagrams: limitations and alternatives

As demonstrated by the preceding section, isobar diagrams are a useful tool to visualize variations in volatile solubility for a specific melt composition. However, many suites of melt inclusions have considerable major element variability, which translates into differing solubilities of H$_2$O-CO$_2$ at a given pressure, and different isobar shapes at a given pressure (Wieser et al., 2021; Roggensack, 2001; Iacovino et al., 2021). We use two suites of melt inclusions to demonstrate this point.

First, we consider a suite of basaltic melt inclusions from Butajira volcanoes in the Main Ethiopian Rift from Iddon and Edmonds (2020). We calculate isobars for a representative inclusion composition (BJ08; Fig. 19a), and then we compare these to the isobars calculated for each individual melt inclusion composition at 1 and 3 kbar (Fig. 19a-b). 3 kbar isobars calculated from the composition of each individual melt inclusion from Butajira (Fig. 19b) cover the entire region of H$_2$O-CO$_2$ space that is bracketed by the 2 and 4 kbar isobars calculated in MagmaSat for the representative melt inclusion composition (Fig. 19b vs. a). In fact, saturation pressures calculated using any given melt inclusion composition vs. the major element composition of the melt inclusion in question can affect the saturation pressure by almost a factor of 2. By extension, visual inspection of melt inclusion volatile concentrations plotted on isobar diagrams may be associated with an error of this magnitude).

Second, we calculate isobars at 300 and 700 bars for a representative subset of melt inclusions from the 2018 eruption of Kilauea with PEC-corrected MgO contents between 6.4 to 13.5 wt% (Wieser et al., 2021). Although less extreme than the Butajira example, the spread of isobars calculated for a single sample at 700 bars for multiple melt inclusions is wider than the distance between isobars calculated 100 bar apart for a specific melt inclusion. Similar isobar spreads are seen for both Kilauea and Butajira using IM–2012 (Supporting Fig. S8 and S9).

VESIcal solves the problem of potentially misleading isobar diagrams by facilitating rapid calculations of saturation pressures for large suites of melt inclusions. For example, calculating saturation pressures for the 33 melt inclusions in the dataset of Iddon and Edmonds (2020) using VolatileCalc-Basalt, S–2014, and IM–2012 takes only 4.2 seconds if VESIcal is run on the ENKI server. MagmaSat
Figure 19. a-b) Isobars for olivine-hosted melt inclusions from Iddon and Edmonds (2020) calculated at 1170°C. a) MagmaSat isobars for melt inclusion BJ08_7 from Butajira with melt inclusion volatile data overlain. b) 1 and 3 kbar isobars calculated for the major element composition of each individual melt inclusion from Butajira (underlain as light lines in a). c-d) As for a-b, but using a representative subset of melt inclusions from Kīlauea (Wieser et al., 2021). To save computation time as this study analyses >100 melt inclusions, PEC-corrected compositions were sorted by MgO content, and every 5th inclusion composition was used to calculate an isobar. Isobars are calculated for a temperature calculated from the melt inclusion MgO content using the thermometer of Helz and Thornber (1987). MagmaSat calculations may run faster if ThermoEngine is installed locally (calculations take 26.3 s using a Dell Inspiron laptop with 16 GB.
Once users have calculated saturation pressures for each inclusion, a number of different x-y plots will provide more information than isobar diagrams. For example, G. Moore (2008) suggests that users could plot two graphs, one showing melt inclusion CO$_2$ content against saturation pressure and one showing melt inclusion H$_2$O contents against saturation pressure. This allows clustering of saturation pressures to be observed, and determination of the importance of each volatile species when calculating saturation pressure. Alternatively, saturation pressures could be plotted as histograms, cumulative density functions, or box/violin plots to assess clustering, and to compare distributions from different crystal populations, eruptions, or different volcanic centres. It may also be advantageous to compare saturation pressures to host-crystal chemistry. For example, Wieser et al. (2021) show that melt inclusions hosted in low forsterite olivine crystals from Kilauea Volcano crystallized at $\sim$1–2 km depth, while melt inclusions hosted within higher forsterite olivines crystallized at $\sim$3–5 km depth.

### 7.2 Assessing Errors

The ability to calculate saturation pressures using a number of different models in VESIcal is advantageous, because it can provide assessment of the systematic errors associated with model choice. If different solubility models produce saturation pressures which are statistically distinguishable using tests such as ANOVA or the Kolmogorov-Smirnov test, or differences between models exceed 10–20% (the approximate quoted error on most models), users need to evaluate their melt compositions, pressures and temperatures in the context of the calibration dataset of each solubility model. Several Jupyter notebooks aiding these comparisons are provided alongside VESIcal part I (Iacovino et al., 2021). Additionally, comparisons between any available experimental data for relevant melt compositions and different solubility models using a workflow similar to that used here for andesites (e.g., isobar diagrams as in Fig. 16, plots of melt composition vs. calibration datasets) will help users select a suitable model. As well as examining melt compositions, users should also evaluate whether they are extrapolating temperature-sensitive models beyond the calibration range (as discussed here for VolatileCalc-Rhyolite).

In general, if a natural silicate melt composition is poorly represented by experimental data, MagmaSat is probably the best model to use, as its thermodynamic nature is more suitable to extrapolation to melt compositions not represented in the calibration dataset than empirical models such as IM–2012, S–2014, or VolatileCalc-Basalt. However, the comprehensive nature of MagmaSat means that the fit to experimental data from any specific region of major element space is compromised by the fact the model is optimizing the overall fit to many different major element compositions. Thus,
where present, models developed for specific volcanic centres, or highly weighted towards specific melt compositions (e.g., A–2019 for the 6 centers they investigate, or IM–2012 for alkaline compositions) may return a better fit. While these composition-specific models may be well calibrated in terms of melt composition, users must also check that they are applying the model within the recommended pressure and temperature range.

VESIcal can also be used to explore the effects of uncertainty in input parameters (as in the sensitivity tests presented here). Error bars for calculated saturation pressures are asymmetrical, even for input parameters with symmetric error distributions. Thus, to calculate an error bar on each inclusion, three saturation pressure calculations must be performed. Using the older generation of tools, this would be exceptionally time consuming. Using VESIcal, calculations can be performed automatically, and users could simply input three spreadsheets, or a spreadsheet with three columns for the uncertain parameter (containing the preferred, maximum and minimum value) to calculate error bars (e.g., uncertainties in total CO\textsubscript{2} contents resulting from 2D measurements of bubble volumes and analytical uncertainties in diad splittings; Wieser et al., 2021).

VESIcal also opens up the capability to handle errors arising from volatile solubility modelling using the advanced functionality of Python3 packages such as Numpy, SciPy and PyMC to perform Bayesian statistical techniques (e.g., Markov Chain Monte Carlo methods). This would allow the uncertainty in all input parameters (e.g., temperature, analytical uncertainties in volatile and major element contents, uncertainty arising from post-entrapment crystallization corrections) into a resulting error distribution for each melt inclusion. These techniques are increasingly being utilized by igneous petrologists, with recent applications including calculating error distributions for diffusion timescales (Mutch et al., 2019), the contribution of melts from distinct mantle sources (Gleeson et al., 2020), and propagating uncertainties in vapour bubble growth models (Rasmussen et al., 2020).

7.3 Data Curation

Now that VESIcal makes it possible to calculate saturation pressures for large melt inclusion datasets in short amounts of time, it is vital that data is published in a way that allows such calculations to be performed retrospectively (e.g., using the same solubility model to calculate saturation pressures for a literature compilation). The concentration of major elements and volatile elements should be provided within a single spreadsheet or database, ideally alongside the composition of the mineral host. Additionally, if melt inclusions are corrected for post-entrapment crystallization (PEC), both raw and corrected major and volatile element concentrations should be published. Particularly for more chemically-complex host minerals like pyroxene and plagioclase, it is highly likely that community standards on the best PEC procedure will change with time. Even for olivine-hosted melt
inclusions, there a variety of correction techniques used in the literature, which can predict very variable amounts of PEC (Wieser et al., 2021, show that the PEC routine in Petrolog3 predicts 30% more PEC than corrections based on Mg# disequilibrium between the host and melt inclusion). Finally, given that numerous recent studies have shown that bubble CO\textsubscript{2} contents can change calculated saturation pressures by up to an order of magnitude (section 5), authors should also make note of whether each melt inclusion contained a vapour bubble (and ideally an estimate of the volume proportion of the vapour bubble). This information may allow the CO\textsubscript{2} contents in vapour bubbles to be reconstructed theoretically by future studies, given the recent proliferation of vapour bubble growth models (e.g. Aster et al., 2016; Rasmussen et al., 2020; Maclennan, 2017).

Proper data curation is particularly vital in the world of volatile solubility modelling because it is very likely that many more experiments will be published over the next decade, allowing the creation of new solubility models that are calibrated over an even wider region of P-T and compositional space. Many publications could not be used for comparisons in this study, because there was simply no way to combine volatile element concentrations and major element concentrations (which were often reported in different tables with non-unique or non-matching sample names), or not reported at all. It would be a great shame if published melt inclusion datasets could not be input into these new models to recalculate saturation pressures, and gain a greater understanding of magma storage in the Earth’s crust.

8 Future work

The comparisons drawn in this review highlight several research areas where further experimental work is required to be able to distinguish which model behaviors are accurate. First, significantly more experiments are needed on andesitic-dacitic melt compositions. Figure 16 shows that it is currently impossible to differentiate a potential failure in any given solubility model from anomalies in any given set of experiments (e.g., the differential effect of addition of H\textsubscript{2}O on CO\textsubscript{2} solubility in different experiments; Fig. 16b. vs f).

One of the challenges when assessing CO\textsubscript{2} solubility in andesitic-dacitic melts is the fact that CO\textsubscript{2} is present as both carbonate and molecular CO\textsubscript{2}. Carbon species do appear separately in FTIR spectra, but the accuracy of FTIR-derived volatile concentrations can be affected by peak overlap (e.g., Brooker et al., 2001b), as well as uncertainty in peak baseline and absorption coefficients (Botcharnikov et al., 2006; Mangan et al., 2021). SIMS cannot distinguish different carbon species (only yields total carbon), but may help to resolve issues with FTIR as a result of increased understanding of the optimal analysis conditions for volatiles in silicate glasses of the last few decades. However, in addition to its substantial cost relative to FTIR, accurate SIMS measurements are reliant
on having a suite of standards with similar major element compositions and a range of volatile contents (and these standards are often characterized by FTIR, so are subject to the caveats mentioned above).

Second, the effect of redox on volatile solubility across the range encountered in terrestrial magmas is still poorly constrained (section 5). This discrepancy largely reflects the fact that the redox conditions at which many experiments in the literature were conducted are uncertain and/or highly variable (e.g., Botcharnikov et al., 2006). Because of this uncertainty, many calibration datasets are built without being able to constrain the quantities of Fe$_2$O$_3$ and FeO for each experimental run.

While our investigation of intermediate melts indicates that the strong sensitivity of CO$_2$ solubility to melt redox shown by P–2006 is likely anomalous because of the presence of experiments with extremely high calculated Fe$^{3+}$/Fe$_T$ ratios in the calibration dataset, further experiments where Fe$_2$O$_3$ and FeO proportions are accurately measured are needed to be certain that this behaviour is not real. It is also noteworthy that almost all the andesitic experiments were performed at higher oxygen fugacities than lavas erupted at volcanic arcs. When the calibration dataset for a given set of compositions is so small, this makes it difficult to deconvolve changes in volatile solubility with melt composition compared to redox. Recent advances in measurements of Fe$^{3+}$/Fe$_T$ using Fe K-edge micro-X-ray absorption near-edge structure (XANES) spectroscopy in hydrous glasses (Cottrell et al., 2018) could provide an avenue to better constrain this parameter in future (and past) experimental products.

It is also worth noting that all the models discussed here only consider the effect of redox through terms for Fe$^{2+}$ and Fe$^{3+}$ in the melt, constraining their applicability to melts more oxidising than the IW buffer. In more reducing conditions, the co-existing CO$_2$-rich phase may be graphite or diamond rather than a CO$_2$-rich vapour phase (Eguchi & Dasgupta, 2018), and the dissolved volatile species may be CO, CH$_4$ and H$_2$ (Mysen et al., 2009). This means that extreme caution is required when applying these solubility models to highly reducing conditions such as those found on other planetary bodies (e.g., the Moon, Mars and Mercury; Li et al., 2017).

Third, there is still significant uncertainty regarding the exact nature of H$_2$O and CO$_2$ mixing at higher pressure. This reflects the difficulty in measuring mixed H$_2$O-CO$_2$ fluids that were in equilibrium with the melt during the experiment. If measured at all, methods in the literature span from puncture and weight loss of frozen capsules (i.e., when frozen the CO$_2$ is released, but not the H$_2$O; Shishkina et al., 2010) to more sophisticated and accurate vacuum line manometry (G. Moore et al., 2008; Iacovino et al., 2013; Allison et al., 2019). New infinite path laser spectrometry technology may offer potential improvement of this critical measurement, but challenges associated with small sample
sizes remain. More work determining the pure CO$_2$ solubility as a function of pressure and temperature would also be of great benefit in constraining the behavior of H$_2$O-poor fluids.

Fourth, we show that the sensitivity of dissolved volatile contents to temperature is highly model-specific. Given the difficulties with constraining temperature sensitivity experimentally (Iacono-Marziano et al., 2012), we suggest that it may be best to parametrize future empirical model at a single temperature (e.g., A-2019, S-2014), or incorporate the temperature sensitivity predicted by thermodynamical models, rather than introduce a spurious temperature dependency which is not founded in experimental data, nor consistent with the relative entropy of melt and fluid terms.

Finally, it is worth noting that all of the solubility models discussed only consider H$_2$O-CO$_2$ in the vapour phase, while in reality, natural fluids in volcanic systems may contain relatively large proportions of F, Cl, and S, as well as a separate brine phase (Botcharnikov et al., 2007). Additionally, at higher pressures and temperatures, significant quantities of major element species will dissolve into a H$_2$O-rich fluid (e.g., Si, Na, K), with silicate melt and hydrous fluids becoming completely miscible above a critical temperature (Bureau & Keppler, 1999). This causes a pure H$_2$O model to underestimate the true solubility of H$_2$O at these conditions. Combined with the fact that it is near-impossible to quench silicate melts with $\geq 9\pm 1$ wt% H$_2$O to a glass phase which can be analysed by SIMS or FTIR (Gavrilenko et al., 2019; Mitchell et al., 2017), quantifying the solubility of H$_2$O at condition relavant to lower crustal magma storage will require experimental innovations (e.g. Makhluf et al., 2020; Mitchell et al., 2017), in addition to developing models capable of calculating equilibria between a silicic melt phase and a complex aqueous fluid (Ghiorso & Sverjensky, 2016; Sverjensky et al., 2014; Huang & Sverjensky, 2019).

9 Conclusion

This review uses the new open-source Python3 tool VESIcal (Iacovino et al., 2021), in addition to VolatileCalc (Newman & Lowenstern, 2002) and Solwcad (Papale et al., 2006) to draw extensive comparisons between the behaviour of 9 different solubility models for a range of melt compositions. We show that these models predict surprisingly different volatile solubilities, particularly for pure CO$_2$ or mixed CO$_2$-H$_2$O fluids. Even for melt compositions that are well represented in the calibration datasets of multiple models (e.g., MORBs), calculated solubilities for pure CO$_2$ can deviate from one another by factors of $\sim 2$. Differential treatment of H$_2$O-CO$_2$ mixing enhances these differences when calculating volatile solubility for melts containing both volatile species. The solubility of CO$_2$ predicted by different rhyolitic models also differs substantially, overwhelming other sources of uncertainty such as analytical errors on measurements of volatile contents or uncertainties in crustal density.
profiles. Differences are most pronounced for peralkaline rhyolites where there are fewer experimental constraints.

Overall, these comparisons demonstrate that it is vital to pick a model which is calibrated for the pressure, temperature, and melt composition of interest. Choice of a poorly calibrated model could introduce a systematic error of a factor of 2 or more in estimates of saturation pressures. This has widespread implications for published estimates of magma storage depths within volcanic systems, and indicates that re-evaluation of published magma storage depths calculated using older models may be warranted.

We also investigate the sensitivity of different models to variation in parameters such as H_2O content (with relevance to diffusive re-equilibration), CO_2 content (with relevance to melt inclusion vapour bubble growth), temperature and oxygen fugacity. We suggest that by performing similar sensitivity tests in the future, the uncertainties affecting calculations of volatile solubility in magmatic systems (and therefore the limitations of each study) can be quantified. We also demonstrate that isobar diagrams are a poor visualization method for determining magma storage depths in systems where melt inclusions possess diverse melt inclusion chemistry, so encourage users to take advantage of the ease and speed of calculations in VESIcal to determine the saturation pressure for each melt inclusion of interest. Saturation pressures can then be visualized as various cumulative frequency distributions, histograms, or violin plots, and plotted against melt inclusion H_2O and CO_2 contents, or parameters relating to host crystal chemistry, to gain greater insight into the factors controlling magma storage depths within volcanic systems.

Finally, we identify that further experimental constraints are required to accurately estimate volatile solubility in andesitic-dacitic melts, and that further work is needed to understand the effect of temperature, redox, and non-ideal mixing between H_2O-CO_2 on volatile solubility.

Acknowledgments

PW acknowledges funding from a NERC DTP studentship (NE/L002507/1) and a National Science Foundation grant (1948862). KI and GMM were supported by the NASA Jacobs JETS Contract (NNJ13HA01C). We thank Jackie Dixon, Giada Iacono-Marziano, Paolo Papale, and Mark Ghiorso for help tracking down and using existing calculation tools, and for help constructing model calibration datasets. We also thank Mark Ghiorso and Paolo Papale for helpful reviews which improved the clarity and robustness of this manuscript, and the late Peter Fox for his editorial handling of the first draft of this manuscript.
**Data Availability** The Jupyter notebooks and associated Excel spreadsheets used to create the figures in this manuscript are currently hosted at https://github.com/PennyWieser/VESIcal-II, and will be added to the main VESIcal github upon acceptance of this article. Videos showing how to use VESIcal are hosted on YouTube https://www.youtube.com/channel/UCpvCCs5KMXz0xXWm0seF8Qw.

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Supporting Information for ”VESIcal Part II: A critical approach to volatile solubility modelling using an open-source Python3 engine”

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Introduction

This supplement presents a number of additional figures, and further explanation with respect to how the coefficients in the IM-2012 models account for some of the behaviors shown by this model.

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0.1. IM–2012 - Difference in isobar curvature

As noted in the main text, the anhydrous version of this model shows a more prominent peak in isobar space compared to the hydrous model. This reflects differences in the coefficient $d_{H_2O}$ attached to the term for the molar fraction of $H_2O$ in the melt. In the anhydrous version, $d_{H_2O}$ is positive (2.3±0.5), so the addition of small amounts of $H_2O$ to the melt causes $X_{H_2O}$, and therefore $CO_2$ solubility to increase. In the hydrous expression, this coefficient is negative (-16.4±1.2). Alone, this would cause the solubility of $CO_2$ to decrease with the addition of small amounts of water to the melt. However, NBO/O calculated on a hydrous basis increases with the addition of $H_2O$ (because $X_{H_2O}$ appears twice in the numerator, but only once in the denominator of the NBO/O term), and the NBO/O term is attached to a positive coefficient (17.3±0.9). This increase in NBO/O overwhelms the product of the negative coefficient multiplied by $X_{H_2O}$.

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Figure S1.  Π vs. SiO\textsubscript{2} graph shown in Fig. 2C of Dixon (1997). Red dots show the raw data presented in Table 1 of Dixon (1997), with analytical totals ranging between 97.38 to 99.995 wt%, while cyan dots show this data following the VESIcal Fixed Volatiles normalization routine. It is apparent that raw data is a significantly better fit; all normalization routines shift SiO\textsubscript{2} contents to higher values.
Figure S2. Parts c-d as in Fig. 4 of the main text, with the addition of isobars calculated for the MORB2 composition for different proportions of Fe$^{3+}$ at 1 and 4 kbar.
Figure S3. Recreation of the figure in the main text for andesitic isobars, but also showing VolatileCalc-Rhyolite isobars calculated at the reference temperature of this model (850°C). These are a significantly better fit to experimental products.
Figure S4. Isobars calculated for the major element composition of run number 1890 from Mangan et al. (2021) at 1125°C. As in the main text, calculations for P-2006 are performed for no Fe$^{3+}$, and a best estimate of the proportion of Fe$^{3+}$ in the experiments (from MELTS for Excel based on reported oxygen fugacity). Isobars in b) calculated for 1035 MPa, isobars in d) calculated for 875 MPa.
Figure S5. Isobars calculated for the major element composition of run number 1890 from Mangan et al. (2021) at 1125°C. As in the main text, calculations for P-2006 are performed for no Fe³⁺, and a best estimate of the proportion of Fe³⁺ in the experiments (from MELTS for Excel based on reported oxygen fugacity). Considering the error bars on measured volatile contents and on models (not shown, but likely 10-20%), MagmaSat recreates all experimental solubilities within error.
Figure S6. Discrepancies between various models and MagmaSat for the suite of Popocatépetl melt inclusions from Atlas et al. (2006). The strong correlation between SiO$_2$ and II$^*$ for these inclusions (Fig. 15a, main text) may account for the increased discrepancy between Shishkina and MagmaSat with increasing SiO$_2$. 
Figure S7. Discrepancies between various models and MagmaSat for the suite of Soufrière Hills melt inclusions from Cassidy et al. (2015). NBO/O decreases with increasing SiO$_2$ in the Soufrière Hills inclusions (Fig. 15c, main text), which may explain why the discrepancy between Iacono-Marziano and MagmaSat increases with increasing SiO$_2$. 
Figure S8. Comparison of isobars calculated for Butajira using MagmaSat (as shown in the main text) vs. IM-2012-H. This demonstrates that both models predict a spread of isobar positions.
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