Sulfation made simple: a strategy for synthesising sulfated molecules†

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Organosulfates play a variety of important roles in biology, from xenobiotic metabolism to the downstream signalling of steroidal sulfates in disease states.1 Sulfate groups on glycosaminoglycans (GAGs) such as heparin, heparan sulfate and chondroitin sulfate, facilitate molecular interactions and protein ligand binding at the cellular surface,2 an area of interest in drug discovery.3 Heparin (an anticoagulant),4 avibactam5 (a β-lactamase inhibitor), sotradecol (a treatment for varicose veins),6 the sulfate metabolite of paracetamol (an analgesic),7 and the bioactive sulfated compounds.9 sulfate motifs (Fig. 1). There are many other natural sources of active sulfated compounds.9

The incorporation of polar hydrophilic organosulfate groups onto drug-like molecules is timely to facilitate research investigating sulfated GAGs as potential new therapies.10 However, the insertion and isolation of sulfate groups into target molecules remains a challenging aspect of their synthesis,11 prompting recent advances into sulfate revealing pro-drugs.12 The presence of one or more sulfate group makes chemical synthesis and purification of (per)sulfated compounds challenging, primarily due to their poor solubility in organic solvents.13 Therefore, the insertion of organosulfate groups is typically the final step in a synthetic method, limiting further chemical modifications.14

† Electronic supplementary information (ESI) available: Preparative routes, compound characterisation, copies of 1H and 13C spectra and the cif file for the X-ray crystal structure of 1. CCDC 1894165 (1). For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c9cc01057b

Fig. 1 Examples of drug molecules containing organosulfates: heparin, sotradecol and avibactam; a chemical tool compound (C3) and a metabolite of paracetamol.

A variety of methods to prepare organosulfates are shown in Chart 1.

The preparation of organosulfate esters include a microwave assisted approach to the sulfation of alcohols, using Me3NSO3 and PySO3, Chart 1(i)).15,16 The addition of catalytic diaryl borinic acid (used to study atherosclerosis),8 contain organosulfate motifs (Fig. 1). There are many other natural sources of

Chart 1. Examples of methods to prepare organosulfates.
nucleoside phosphosulfates\textsuperscript{22} highlighted that the solubility of the organosulfate ester can be modulated by increasing the lipophilicity of the corresponding cation.

During a medicinal chemistry programme we encountered difficulties using the established sulfation methods to persulfate compounds, due to the poor solubility of the organosulfate intermediates and their resulting purification. We sought to develop an all in one reagent, to improve the solubility of the intermediate organosulfate ester. Combining tributylamine (pK\textsubscript{a}, 10.89) with SO\textsubscript{3}, we envisioned this would create a complex (Bu\textsubscript{3}NSO\textsubscript{3}, \textsuperscript{1}) for the persulfation of alcohols, with increased lipophilicity of the intermediate sulfate ester, improving the overall solubility of the organosulfates in organic solvents. We rationalised that \textsuperscript{1} would retain similar activity to SO\textsubscript{3} complexes with Et\textsubscript{3}N and Me\textsubscript{3}N, due to their similar Lewis base strengths, pK\textsubscript{a} = 11.01 and 10.63, respectively. Overall, \textsuperscript{1} may permit sequential chemical steps in organic solvents, with a simple purification method to the corresponding sodium salt, streamlining the synthesis of organosulfates.

The complexation of sulfur trioxide to nitrogen or oxygen containing molecules (such as pyridine, NMe\textsubscript{3}, NEt\textsubscript{3}, DMF, THF and dioxane) is well known,\textsuperscript{11,15} the use of an organic solubilising partner, tributylamine, is however not. To the best of our knowledge, the only literature report of the synthesis and physical study of \textsuperscript{1} was by Moede in 1949.\textsuperscript{23} It was not until 1976 that Parshikov and co-workers\textsuperscript{24} studied \textsuperscript{1} as a sulfating agent on simple aliphatic alcohols (without spectroscopic characterisation) and found that \textsuperscript{1} reacts via an S\textsubscript{N}2 mechanism driven by the hydrogen-bonding propensity of the alcohol under study. However, to the best of our knowledge no further use or development of this reagent has been reported.

We synthesised \textsuperscript{1} by reaction of tributylamine with chlorosulfonic acid, affording a 90% yield on a 60 g scale (Scheme 1(a)).\textsuperscript{25} For the first time both NMR spectral data and the crystal structure of \textsuperscript{1}, obtained from small molecule single crystal X-ray diffraction, was determined (Scheme 1(b)).\textsuperscript{26} Bu\textsubscript{3}NSO\textsubscript{3} (\textsuperscript{1}) adopts a Gauche conformation within an asymmetric unit cell caused by hydrogen bonding between the methylene hydrogen atoms to the nitrogen and the oxygens of SO\textsubscript{3}. The measured N-S bond length in \textsuperscript{1} is 1.886(3) Å, a comparable bond length to a single N-S bond (typically: 1.73–1.83 Å versus 2.06 Å for a donor–acceptor system),\textsuperscript{27} suggesting that \textsuperscript{1} exists as a betaine in the solid state which may have implications for the other unsolved amine–SO\textsubscript{3} complexes and their associated mechanisms.

Benzyl alcohol (\textsuperscript{2a}) was selected for the optimisation study (Chart 2) due to a distinct down-field shift (+0.35 ppm) of the benzylic signal after sulfation (by \textsuperscript{1}H-NMR spectroscopy).

We examined the sulfation of \textsuperscript{2a} with varying equivalents of \textsuperscript{1} (entries 1 to 4). It was found that 2.0 equivalents of \textsuperscript{1} was optimal for high conversions (>99%) and isolated yields (95%, entry 3).

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Less than 2.0 equivalents of 1 (entries 1 and 2) gave incomplete conversion to 3a. We next surveyed the effect of temperature (entries 5–8 vs. entry 3). 90 °C and 2.0 equivalents of 1 gave complete conversion to 3a within 2 h. Notably, reaction completion was achieved, even at 30 °C (with increased reaction time) demonstrating that 1 could be a suitable reagent for temperature sensitive substrates such as proteins.13

We compared the use of two commercially available amine–SO3 complexes, namely pyridine and trimethylamine (entries 9 and 12, respectively). Py–SO3 was more reactive than 1, giving complete conversion to 3a in 0.5 h but only a 17% isolated yield of the pyridinium species. As a control experiment, the sequential exchange of the pyridinium salt with different Bu4N cations afforded the isolation of 4a in 47% (Bu4NI) and 76% (Bu4NOAc) yield, respectively (entries 10 and 11). The use of Me3N–SO3 gave poor results (0% isolated yield of 4a from a 29% conversion to 3a). The addition of Bu4NI to the reaction mixture significantly improved the reaction, affording an 88% conversion to 3a (entry 13). Unfortunately, a more complex reaction mixture was detrimental to the isolation of 4a. These results demonstrate the isolation benefits associated with the protocol developed with 1.

We observed that 1 can achieve high conversions of 2a to 3a/4a without microwave irradiation (entries 1–7 vs. 14 and 15) unlike other reported sulfating agents. The addition of a hetero or homogenous base (entries 16–19) was investigated and in all cases this led to a decrease in conversion to 3a. Most notably, the addition of tributylamine (entry 19) initially increased the rate conversion to 3a, but reduced the overall conversion contradicting the original report.24 We rationalised that the addition of Et3N (with a higher pKₐ value than Bu4N) competes in the reaction, forming Et3NSO3 in situ. This amine exchange has also been observed in the reaction of Et3N with PySO3.21

We applied the optimal conditions to the synthesis of a range of sulfate esters (Chart 3). In all examples (Chart 3) we observed a near-quantitative conversion to the corresponding sulfate ester as the tributylammonium salt, independent of both electronic and steric factors. Differences occurred when converting the intermediate into a crystalline sodium salt; most likely due to the variability in precipitation of the sodium salt. Therefore differences of electron-withdrawing and electron-donating factors are difficult to draw with certainty. Steric bulk adjacent to the reacting centre was accommodated with ease (4h) and the reaction could also be applied to phenolic alcohols (4i).

Driving an organosulfate ester reaction to completion, by sulfating all possible hydroxyl sites in a structure remains a synthetic challenge. This is influenced by anionic crowding, making per-sulfation progressively more difficult.16 Therefore we applied the conditions to examples of compounds requiring two or more sulfation events, including natural products, and compounds containing different hydroxyl moieties within the same structure to probe the selectivity profile (Chart 4).

Cyclohexane-1,2-diol afforded the disulfate in high conversion and isolated yield, likewise with glycerol, the tri-sulfated analogue was prepared with ease (5 and 6, respectively). The reaction of 1 with menthol (7) delivered the NBu3 salt with ease.

To probe the alcohol selectivity profile, a primary alcohol was functionalised in preference to a phenol in 8. Therefore, β-estradiol afforded 17β-estradiol sulfate28 over the more common metabolite 3β-estradiol in 60% isolated yield (9a) using 1.5 eq. of 1. Using 5.0 eq. of 1 both the 17- and 3-positions were sulfated in 84% isolated yield (9b).

We then applied the methodology to the original medicinal chemistry challenge, the sodium salt of glycomimetic C3 (10).8 High conversion and an isolated yield of 76% on a 500 mg scale was achieved. The methodology was applied to the Fmoc-protected amino acid, serine, resulting in excellent yields (95% under normal conditions) and 90% at a non-denaturing temperature (11) both using 4.0 eq. of 1. Importantly, no loss in enantiomeric ratio was observed upon sulfation (>99 : 1).29

In summary, we have reported the first scalable preparation and reaction scoping study of 1 as a mild, bench-stable, and chromatography-free method to access organic sulfate esters as their ammonium or sodium salts. The reaction holds promise with the ability to install up to three sulfate groups on complex scaffolds including examples where sterics would limit other...
methods. Further work to elucidate the scope of the reaction on human sulfate metabolites is ongoing.

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Conflicts of interest

There are no conflicts to declare.

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