Karsili, T. N. V., Marchetti, B., & Ashfold, M. N. R. (2016). Mechanistic Insights into Excited State Intramolecular Proton Transfer in Isolated and Metal Chelated Supramolecular Chemosensors. *Dalton Transactions, 45*, 18921-18930. https://doi.org/10.1039/C6DT03906E
Mechanistic insights into excited state intramolecular proton transfer in isolated and metal chelated supramolecular chemosensors†

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We report a computational study of excited state intramolecular proton transfer in a series of related and progressively more complex organic chromophores ranging from 2-[[2'-hydroxyphenyl]-benzoxazole (HBO) through to the 5-[[benzo[d]oxazol-2'-yl]-2-(4-[[bis(pyridin-2'-ylmethyl)amino)methyl]benzo[d] oxazol-2'-yl]-4-hydroxyphenolate (HDBO') anion. The latter chelates group 12 metal cations (X = Zn²⁺, Cd²⁺ and Hg²⁺), and can serve as a fluorescence-based sensor for such metals. Initial π* ← π excitation of the ground (S₀) state enol-tautomer induces charge separation in the first excited singlet (S₁) state and drives the subsequent proton transfer (i.e., enol → keto tautomerism). The keto-tautomer constitutes a local minimum on the S₁ PES, and is responsible for highly Stokes shifted fluorescent emission; S₁(enol) → S₀ fluorescence is proposed to account for the shorter wavelength emission from the X−HDBO' complexes. Derivatives of HDBO' that should retain the favourable visible absorption and heavily Stokes shifted emission properties but, additionally, offer higher fluorescence quantum yields (i.e., enhanced metal sensing capability) are proposed.

1. Introduction

Luminescent sensors are typically functionalised supramolecules that enhance or quench fluorescence (or phosphorescence) upon an external stimulus, such as a change in the chemical environment.¹,² Such species are generally optimised synthetically in order to maximise their selectivities and specificities whilst steering their photophysical properties towards a particular intended function. Many studies have focussed on the synthesis of luminescent supramolecules that are specific for the detection of malignant tumours.³ Upon detection and subsequent complexation to specific markers released by cancerous tissue, changes to the luminescence signal and the Stokes shift reveal their specific positions and activities, thus enhancing the efficiency and efficacy of the diagnosis. Hence the major interest in photochemically active supramolecules as bio-diagnostics.⁴,⁵

†Electronic supplementary information (ESI) available: (i) Minimum energy geometries and dominant orbital promotions associated with excitation to the lowest three singlet excited states of Cd²⁺−HDBO' and Hg²⁺−HDBO'; (ii) permanent dipole moment vectors in the S₀ and S₁ (enol and keto) states of HBO. See DOI: 10.1039/c6dt03906e

Trace metal detection in, for example, water supply systems is another crucial analytical problem and one that is required for the regulation of safe drinking water.⁶ Supramolecular systems that are benign towards human health are ideal for this purpose, allowing their addition to the water supply to test and ultimately regulate the concentrations of trace metals. Binding of such metals to the supramolecule causes a measurable (photo)chemical change. One such example is the detection of alkali metal cations using N-bridgehead cryptand-naphthalene supramolecules which, when cation-bound, fluoresce following electronic excitation.⁷ In the absence of the metal, however, photoexcitation of the free supramolecule causes photo-induced electron transfer (PET) from the naphthalene donor to the cryptand-hole acceptor; the fluorescence quantum yield is negligible. The mechanisms responsible for this switch in behaviour are, as yet, poorly understood.

Other studies have demonstrated the tunability of such supramolecular sensors.⁸–¹⁰ Aza-crown ethers, for example, with chromophores (such as coumarin) attached to both ends, have been shown to exhibit similar, but opposite, PET and fluorescence propensities, i.e., the supramolecule fluoresces when metal-free but undergoes PET when metal-bound.¹¹ These systems are also referred to as molecular LOGIC switches, since they ‘switch’ the fluorescence or PET signal ‘on’ or ‘off’ depending on the presence or absence of the
intended ‘to be detected’ solute.\textsuperscript{12} In these specific examples, sensing relies on a change in the electronic properties of the sensor upon chelation with the alkali metal. Such donor-acceptor supramolecules can also be tuned, synthetically, in order to function as molecular nanowires.\textsuperscript{13-16}

As well as changes in electronic structure, molecules used as luminescent probes can also undergo geometry changes upon chelation and/or (subsequent) photoexcitation. One notable example is \textit{E}-diazobenzene, which can function as a molecular photo-switch upon irradiation with light and has found application in pH sensing, photochromism and in non-linear optics.\textsuperscript{17} These applications all rely on the same photo-physics: upon near-UV irradiation, \textit{E}-diazobenzene undergoes $E \to Z$ photoisomerism on the excited state potential energy surface \textit{(PES)}, followed by rapid internal conversion \textit{(IC)} and subsequent branching into either \textit{Z-} (30\%) or \textit{E}-diazobenzene (70\%) on the ground state \textit{PES}. The timescales of such $E \to Z$ photoisomerisms and subsequent IC processes are typically ultrafast – sub-ps in the specific case of diazobenzene, as revealed by transient absorption studies in deuterated dimethyl sulfoxide\textsuperscript{18} and n-hexane solutions.\textsuperscript{19} Such ultrafast IC is generally associated with the presence of low energy crossings between adiabatic \textit{PESs} which, when orthogonal motions are considered, develop into conical intersections \textit{(CIs)}.\textsuperscript{20} CIs are ubiquitous in photochemically and excited state driven phenomena and are recognised as regions of the \textit{PES} that promote efficient transfer of excited state population back to the ground state (or between excited states). Such processes are termed non-adiabatic, as the electronic and nuclear motions are usually strongly coupled.

Non-adiabatic processes can be driven by a variety of nuclear motions, but the present study focusses on proton/electron coupled motions – commonly referred to as proton-coupled electron transfer \textit{(PCET)} processes – which drive much of the redox chemistry associated with transition metal complexes.\textsuperscript{21} In a hydrogen-bonded complex, photoinduced-PCET occurs when light absorption initiates electron promotion from an acidic donor (D) orbital to a hitherto unoccupied, basic acceptor (A) orbital. This creates a D’H–A’ charge-separated \textit{(CS)} excited state that is neutralised by transfer of a proton from D to A. The potential energy of a CS excited state almost always decreases upon proton transfer \textit{(PT)}, \textit{i.e.} there is a driving force for PT in a CS excited state. The ground state, in contrast, is stable in the closed-shell DH–A configuration, so its \textit{PES} shows a large barrier with respect to PT. Thus the CS and ground state potentials inevitably cross along the PT coordinate and the resulting CI facilitates non-adiabatic decay of CS state population to the ground state – \textit{i.e.} facilitates ultrafast PCET. In cases where the PT occurs at short range, the charge separation in the excited CS state may not fully neutralise and further nuclear rearrangements may be necessary to cancel the remaining charge-separation ultimately leading to a curve crossing.\textsuperscript{22-25} Such molecular distortions may be restricted in strained molecular systems, however. In these cases, the excited state population may be trapped in a metastable CS state, which then decays to the ground state by an alternative route, by interacting with surrounding solvent molecules or, in extreme cases, by fluorescence. The fluorescence from such constrained systems can be exploited when designing effective supramolecular fluorescent probes\textsuperscript{22,23,26} – as we illustrate here. The class of (supra-)molecular sensors known to function via PCET is still rather small but, as shown in a recent review,\textsuperscript{27} is growing fast.

The present study explores the mechanism of excited state intramolecular proton-transfer \textit{(ESIPT)} (a specific class of PCET) in a series of organic chromophores that can be tuned synthetically so as to function as supramolecular chemosensors for trace metal detection. We start with 2-(2′-hydroxypyridin-phenyl)-benzoxazole (HBO, depicted in Fig. 1), which is known to undergo ESIPT upon electronic excitation.\textsuperscript{27} Transient absorption measurements following 340 nm excitation of HBO in cyclohexane solution show enol $\to$ keto tautomerism occurring as a result of ESIPT on a timescale of $\sim$100 fs or less\textsuperscript{28-31} – consistent with the results of mixed quantum/classical molecular dynamics calculations by Daengngern \textit{et al.}\textsuperscript{32} The thio-analogue of HBO has also been studied recently by Barbatti and co-workers\textsuperscript{33} who, proposed, on the basis of the deduced ultrafast excited state intra- and inter-molecular PT, that this molecule could be a good candidate for use as a protein fluorescent marker.

The potential utility of the fast and efficient ESIPT in HBO is likely to be limited by its absorption maximum ($\lambda_{\text{abs(max)}} \sim$330 nm), however, which is much shorter than the (visible) wavelengths preferred for supramolecular sensing applications.\textsuperscript{34} Adding another benzoxazole to HBO yields 2-(2′-hydroxyphenyl)-dibenzoazole (HDBO, termed bis(HBO) in ref. 34, and also shown in Fig. 1), shifts the absorption maximum closer to the visible ($\lambda_{\text{abs(max)}} = 412$ nm) and still maintains the ESIPT upon photoexcitation.\textsuperscript{34} Functionalising HBO with \textit{N},\textit{N}-dipyriddimethylamylamine results in formation of 5-\textit{[benzo[d]} oxazol-2-yl]-2-(4-\textit{[bis(pyridin-2-ylmethyl]amino)methyl]benzo [d’]oxazol-2-yl]-4 hydroxyphenolate \textit{(henceforth termed HDBO’}, but abbreviated as Zinhbo-1 in ref. 34). Protonated HDBO’ shows a similar absorption maximum $\lambda_{\text{abs(max)}} \sim$410 nm in non-polar solvents, attributable to the neutral free ligand, and a weak fluorescent emission centred at $\lambda_{\text{em}} = 543$ nm (\textit{i.e.} a Stokes shift of $\sim$130 nm). HDBO’ is of particular interest as a ligand sensor for trace metal detection in more polar solvents. The group 12 X-HDBO’ complexes (X = \textit{Zn}$^{2+}$, \textit{Cd}$^{2+}$ and \textit{Hg}$^{2+}$, with the phenol proximal to the cation deprotonated, also illustrated in Fig. 1) show much larger Stokes shifts than bare HDBO’. \textit{Zn}$^{2+}$-HDBO’, for example, shows $\lambda_{\text{abs(max)}} \sim$480 nm both in DMSO and in aqueous solution, and $\lambda_{\text{em}} = 712$ nm – a spectacularly large Stokes shift of $\sim$230 nm.\textsuperscript{34}

The present work reports gas phase (isolated molecule) electronic structure calculations designed to offer mechanistic insights into these photoinduced turn-on processes. Specifically, we explore the topographies of the excited state \textit{PESs} that support these efficient ESIPT processes, the origin of the large differences in Stokes shift upon progressing from HBO to HDBO to X-HDBO’, and the ways in which chelation with metal ions affects the potential energy (PE) profile along
2. Computational methodology

Using the Gaussian 09 computational package, the ground state minimum energy geometries of all species of current interest were optimized using the long-range correlated Coulomb Attenuated Model Becke-3rd parameter-Lee–Yang–Parr (CAM-B3LYP) functional in Density Functional Theory (DFT).

Relaxed PE profiles for the ground and first few excited electronic states of all systems investigated were scanned along the ESIPT coordinate using time dependent (TD)-DFT/CAM-B3LYP using the O1–H bond length \( R_{\text{O1-H}} \) as the driving coordinate. A 6-31G(d)\(^{37} \) Pople basis set was used for all first and second row atoms (i.e. H, C, N and O), together with 6-31G(d) (for Zn) and LANL2DZ\(^{38} \) (for Cd and Hg) basis functions for calculations of the metal complexes.

\( S_1-S_0 \) absorption profiles of HBO and HDBO were simulated using Newton-X\(^{39,40} \) by first generating an ensemble of 100 initial geometries \( (N) \) – sampled using a Wigner distribution based on the ground state normal mode harmonic wave-numbers (calculated using the system specific level of theory described above). Following ref. 41, the vertical excitation energy and oscillator strength were then calculated for each of the \( N \) initial geometries using TD-DFT/CAM-B3LYP/6-31G(d) and the excitation energy dependent photoabsorption cross section \( P(E) \) obtained using eq. (1)

\[
P(E) = \frac{\pi e^2}{2m \varepsilon_0} \sum_{i=0}^{j-1,3} \sum_{k=0}^{\text{100}} f_{ij}(R_k)g(E - \Delta E(R_k), \delta). \tag{1}
\]

The internal sum in eqn (1) is over the set of Wigner points \( (N_{\text{tot}} = 100) \) while the external sum includes transitions from the initial state \( (i, \text{the } S_0 \text{ state}) \) to final states \( (j = S_1, S_2 \text{ and } S_3 \text{ in the present modelling}) \) with respective oscillator strengths \( f_{ij} \). \( R_k \) represents the initial coordinates of the \( N^\text{th} \) point and \( g(E - \Delta E(R_k), \delta) \) is a Lorentzian line shape function given by eqn (2)

\[
g(E - \Delta E(R_k), \delta) = \frac{\hbar \delta}{2\pi} \left( \frac{E - \Delta E_{ij}}{2} \right)^{-1}. \tag{2}
\]

\( \delta \) in eqn (2) is a broadening factor, which was arbitrarily set to 0.05 eV for each of the calculated absorption profiles reported here. \( S_1-S_0 \) emission profiles were calculated in a similar manner, by sampling an ensemble of 100 initial geometries using a Wigner distribution based on the normal mode wave-numbers associated with the optimized geometry of the first singlet excited electronic state (calculated using the CAM-B3LYP/6-31G(d) level of theory as above).

3. Results and discussion

3.1. Ground state minimum energy geometry and vertical excitation energies

Fig. 1 showed the ground state minimum energy geometries of HBO, HDBO, protonated HDBO’ and Zn\(^{2+}\)-HDBO’. The Cd\(^{2+}\)-HDBO’ and Hg\(^{2+}\)-HDBO’ complexes display the same struc-
The minimum energy structure in each case involves planar HBO and HDBO moieties. The acidic O(1)–H group shows an intramolecular hydrogen bond with the basic N(10) atom in all species, as does the O(8)–H group (with the N(18) atom) in HDBO and protonated HDBO′; both of these N atoms possess an in-plane lone pair that can act as the hydrogen bond acceptor.

Table 1 lists vertical excitation energies (VEEs) and oscillator strengths for the $S_j \rightarrow S_0$ ($j = 1$–3) transitions returned by the TD-DFT calculations for all systems featured in the present study, while the dominant associated orbital promotions are shown in Fig. 2. For completeness, the calculated VEEs to the corresponding triplet excited states are also included in Table 1; the PE profiles of some of the latter states will be considered later, when considering the metal complexes (vide infra).

The first three singlet excited states of HBO and HDBO are each calculated to be of $\pi\pi^*$ character. In both cases, the excitation from the $\pi$ highest occupied molecular orbital (HOMO) to the lowest unoccupied $\pi^*$ (LUMO) as shown in Fig. 2 is calculated to contribute >80% of the total weight of the $S_1 \rightarrow S_0$ transition. These transitions have appreciable oscillator strengths, and their respective VEEs are in reasonable accord with the long wavelength absorption maxima determined by experiment. The calculated $S_1 \rightarrow S_0$ VEE in HBO (Table 1) corresponds to a wavelength $\lambda_{\text{abs, calc}} \sim 285$ nm (cf. the experimental $\lambda_{\text{abs(max)}} \sim 330$ nm), while the equivalent comparison in the case of HDBO is $\lambda_{\text{abs, calc}} \sim 365$ nm, cf. $\lambda_{\text{abs(max)}} \sim 412$ nm. We should expect this (already acceptable) level of agreement...
between theory and experiment to improve further once zero-point energy differences are included. The $S_2\rightarrow S_0$ and $S_3\rightarrow S_0$ transitions are poorly described in terms of single orbital promotions; in both molecules, the orbital promotions shown in Fig. 2 are calculated to be dominant, but only to contribute ~40% of the total transition strength.

The electronic promotions responsible for the $S_1$ and $S_2$ states of protonated HDBO' are analogous to those for HBO, whereas the largest contributor to the $S_3$ state is predicted to involve electron promotion from an orbital centred on the $N,N$-dipyridylmethylamine group to the analogous $\pi^*$ LUMO as in HBO. Theory (Table 1) and experiment both suggest that functionalization with $N,N$-dipyridylmethylamine has little impact on the long wavelength $S_1\rightarrow S_0$ absorption of HDBO.

Upon deprotonating O(8), this and the N(18) atom are now both effective donor sites enabling complexation to metal atoms with low oxidation states. The same orbital promotions dominate when forming the $S_1$ states of protonated and deprotonated HDBO' but, as Table 1 shows, the calculated VEE of the $S_1\rightarrow S_0$ transition in the latter is significantly reduced. This can be rationalised in light of the loss of conjugation upon deprotonating O(8). As Fig. 2 shows, deprotonating HDBO' destabilises the HOMO by increasing the $\pi$-density on O(8) and thus reducing the conjugation of the C(5)–O(18) bond. The $\pi^*$ LUMO, in contrast, is largely unaffected. Thus the net effect of deprotonating O(8) is a reduction in the HOMO–LUMO energy gap — consistent with the bathochromic shift of the $S_1\rightarrow S_0$ absorption maximum observed for the singly deprotonated, unchelated HDBO' moiety in aqueous solution.34

The visible absorption maximum shifts further upon chelation with Zn$^{2+}$, Cd$^{2+}$ or Hg$^{2+}$. The metal atoms in these complexes all have filled d shells so, as Fig. 2 shows, the HOMO in each case is similar to that in HDBO and the first excited states are formed by ligand centred $\pi^*\rightarrow\pi$ transitions. As Table 1 shows, the respective $S_1\rightarrow S_0$ transitions are each predicted to have substantial oscillator strengths and to show a bathochromic shift relative to the corresponding transitions in HBO, HDBO and protonated HDBO' — again in accord with experimental observation34 — but to be blue-shifted relative to that in deprotonated HDBO'. The latter trend is understandable in terms of a partial recovery of conjugation within the ligand $\pi$-system induced by dative bonding to the metal centre. We recognise that heavy metal species like Hg$^{2+}$, in particular, are likely to induce strong spin–orbit coupling, but the inclusion of such effects is beyond the scope of the present work. The $S_2\rightarrow S_0$ and $S_3\rightarrow S_0$ transitions in these X–HDBO' complexes also involve $\pi^*\rightarrow\pi$ excitations. Reference to Fig. 2 and $S_1^\dagger$ suggests that the energetic ordering of these closely spaced $\pi^*$ orbitals is X dependent, but in all cases, one of these $\pi^*$ orbitals is largely localised on the $N,N$-dipyridylmethylamine moiety while the other is delocalised over the entire HDBO' ligand. Neither overlap well with the $\pi$ HOMO, thus accounting for the small oscillator strengths calculated for these transitions (Table 1).

### 3.2. Excited state intramolecular proton transfer, emission and Stokes shift

**HBO and HDBO.** Fig. 3(a) and (b) show PE profiles for the ground and first few excited states of, respectively, HBO and HDBO along the ESIPT coordinate. In both cases, the filled black circles indicate the relaxed profile of the ground state along $R_{O(1)-H}$ and the open coloured circles show the energies of various singlet and triplet excited states calculated at the corresponding $S_0$ relaxed geometries along $R_{O(1)-H}$. The profile shown by the filled red circles is for the relaxed geometries of the $S_1$ state along $R_{O(1)-H}$, while the open black circles show the $S_0$ energies at the corresponding relaxed $S_1$ geometries.

Fig. 3 confirms that the *enol* tautomer is the minimum energy structure in the electronic ground states of both HBO and HDBO and that the *keto* tautomer corresponds to a local

![Fig. 3 PE profiles along the R_{O(1)-H} ESIPT coordinate in (a) HBO and (b) HDBO. The black and red filled circles show the relaxed profiles for, respectively, the $S_0$ and $S_1$ ($t^\pi\pi^*(CT)$) states, while the open black circles show the energy of the $S_0$ state computed at the relaxed $S_1$ state geometries. Other PE profiles shown by colored open circles are for the excited states indicated, calculated at the relaxed $S_0$ state geometries.](image.png)
maximum along $R_{O(1)-H}$. The respective $S_1$ potentials, in contrast, show minima for both enol and keto tautomers, with the latter calculated to be more stable by $\sim 0.3$ eV in HBO, and by $\sim 0.1$ eV in the case of HDDBO. The calculated barrier to enol $\rightarrow$ keto tautomerism on the $S_1$ PES of HBO is only $\sim 0.05$ eV; the corresponding barrier for HDDBO is calculated to be larger ($\sim 0.13$ eV). Thus the present calculations reveal a clear driving force for enol $\rightarrow$ keto tautomerism following vertical excitation to the $S_1$ state of HBO. This accords with the predicted change in both the magnitude and direction of the permanent dipole moment upon electronic excitation (see Fig. S2 of ESI†), and indicates that the $S_1$ state develops substantial CT character upon extending $R_{O(1)-H}$. Based solely on the present calculations, the tautomerism probability following vertical excitation in HDDBO is less clear cut and further complicated by the additional O(8)H⋯N(18) hydrogen bond, which opens the possibility of double ESIPT (along the $R_{O(1)-H}$ and $R_{O(8)-H}$ coordinates).

Fig. 4 shows calculated $S_0$ and $S_1$ PESs of HDDBO plotted as functions of $R_{O(1)-H}$ and $R_{O(8)-H}$. For each $R_{O(1)-H}$ and $R_{O(8)-H}$ combination, the $S_0$ energy was calculated after relaxing all other nuclear degrees of freedom and the $S_1$ energies were then computed at these relaxed $S_0$ geometries. The $S_1$ PES calculated in this way displays minima corresponding to the keto tautomers formed by single proton transfer along either $R_{O(1)-H}$ and $R_{O(8)-H}$. These are equivalent, as required by the parent symmetry. The $S_1$ PES also shows a local minimum at the geometry corresponding to double ESIPT, but the calculated energy of this double-keto structure is $\sim 0.27$ eV higher than the minima accessed by single ESIPT and $\sim 0.2$ eV above the starting double-enol minimum. The calculated energy barrier associated with the concerted double-enol $\rightarrow$ double-keto tautomerism on the $S_1$ PES is $\sim 0.3$ eV (cf. $\sim 1.75$ eV on the $S_0$ PES). Based on this topography, we predict a dominant role for single ESIPT (cf. double ESIPT) in the photochemistry of HDDBO.

As Fig. 3 showed, the unrelaxed PE profiles for the $S_1$ and $S_2$ states of both HBO and HDDBO show local minima for both the enol and keto tautomers and can thus also be expected to exhibit CT character upon increasing $R_{O(1)-H}$. These unrelaxed potentials all show barriers upon extending $R_{O(1)-H}$, beyond which the $S_2$ states (in particular) stabilise and, in both molecules, are predicted to show an energetic preference for the keto-structure. The $T_1$ states of HBO and HDDBO share the same electronic configurations as the respective $S_1$ states. The calculated PE profiles for these triplet states also show minima for both the enol and keto tautomers, whereas the profiles for the $T_2$ and $T_3$ states (the electronic configurations of which are analogous to those of the respective $S_2$ and $S_3$ states) simply increase with increasing $R_{O(1)-H}$ implying no driving force for ESIPT in these states. We note that the $T_2$ and $T_3$ potentials are predicted to cross other singlet and/or triplet potentials, which could provide non-radiative decay pathways for any triplet state population formed by photoexcitation. However, given the absence of heavy atoms in HBO or HDDBO, we suggest that the photochemistry of these molecules is likely to be dominated by nuclear motions on the singlet potentials.

We now focus on the $S_1$ state and the likely photochemistry induced by $S_1$$\rightarrow$$S_0$ absorption in these two molecules. Following photoinduced tautomerism, we consider two possible fates for a keto tautomer in its $S_1$ state: fluorescence and internal conversion. Analogy with molecules like 2-(2′-hydroxyphenyl)benzotriazole$^{23}$ and 2-(2′-hydroxyphenyl)benzothiazole$^{23}$ suggests that $E \rightarrow Z$ isomerism around the C(3)–C(9) bond (in the case of HBO, C(6)–C(17) in the case of HDDBO)) might offer a route from the keto minimum to a minimum energy (ME)CI between the $S_1$ and $S_0$ PESs. Fig. 5, which shows PE profiles for $E \rightarrow Z$
isomerism about C(3)=C(9) in HBO calculated as a function of the C(2)–C(3)–C(9)–N(10) dihedral angle, \( \phi_{C(3)-C(9)} \), confirms this expectation. The displayed energy at \( \phi_{C(3)-C(9)} = 0^\circ \) is the relaxed S1 energy at \( R_{O(1)-H} = 1.8 \) Å (see Fig. 3(a)), while the final energy point (at \( \phi_{C(3)-C(9)} = 110^\circ \)) is for the geometry at which the S1/S0 energy gap in the initial S1 relaxed scan along \( \phi_{C(3)-C(9)} \) was smallest. The S1 electronic configuration varies across the intermediate geometries spanned by the S1 relaxed scan, so the intermediate geometries and associated energies were obtained by constructing a linear interpolation in internal coordinate (LIIC) between \( \phi_{C(3)-C(9)} = 0^\circ \) and 110°. The ∼0.4 eV barrier in the PE profile returned in this way thus represents an upper limit. Thus the present calculations serve to reinforce previous conclusions regarding the twisted minimum energy geometry of the keto-tautomer of HBO in its S1 state.42 They also suggest that \( E \rightarrow Z \) isomerisation about the C(3)=C(9) bond could be a viable non-radiative decay route by which keto HBO(S1) population (and, by extrapolation, similar populations in HBO and the X–HDBO’ complexes) could undergo internal conversion to the S0 state. We will revisit this path later.

Fluorescence is another possible S1 population loss mechanism, evidence for which is provided by the reported (albeit weak) emission spectra of the present systems.34 Fig. 6 shows S1–S0 absorption and emission profiles of both HBO and HDBO calculated using Wigner sampling methods and Newton-X. The respective absorption spectra (shown in blue) are dominated by the parent enol tautomer, while the displayed emission spectra are from both the keto tautomer formed upon single ESIPT (shown in red) and, for completeness, from the respective S1 enol minima (shown in purple). These calculations reinforce previous suggestions that single ESIPT is responsible for the large Stokes shifts observed experimentally,14 and succeed in reproducing (semi-quantitatively, at least) the magnitudes of these shifts which, as Fig. 6 shows, are predicted to be ∼110 nm in HBO and ∼150 nm in HDBO.

The present results agree well with previous explorations of the excited state proton transfer of HBO and HDDBO (and substituted analogues) and serve to re-emphasise the necessity for long-range correlation in describing the charge-separated states involved in ESIPT.27,42 The extents to which these earlier conclusions, and the available experimental data, are reproduced by the present long-range corrected CAM-B3LYP functional encourages confidence in extrapolating the mechanism identified in HBO and HDBO to the metal-containing systems.

X–HDBO’ (where X = Zn2+, Cd2+ and Hg2+). The calculated PE profiles along the ESIPT coordinate of Zn2+-HDBO’ displayed in Fig. 7 show many similarities with those obtained for HBO and HDBO. Again, the minimum energy configuration in the S0 state is the enol tautomer, which is predicted to undergo enol → keto tautomerisation following excitation to the S1 state. The S1 keto tautomer is predicted to be slightly more stable than the enol form (by ~0.04 eV). The S1 keto minimum is calculated to lie ~2.7 eV above the S0 minimum but, because of the very different topographies of the S1 and S0 PESs, the S1–S0 energetic gap at the geometry of the S1 keto tautomer is only ~1.7 eV (corresponding to a wavelength of ~720 nm). As Fig. 7 shows, the S1 keto tautomer is predicted to be slightly more stable than the enol form (by ~0.04 eV). The S1 keto minimum is calculated to lie ~2.7 eV above the S0 minimum but, because of the very different topographies of the S1 and S0 PESs, the S1–S0 energetic gap at the geometry of the S1 keto tautomer is only ~1.7 eV (corresponding to a wavelength of ~720 nm). As Fig. 7 shows, the T1-PES shows a similar topography along \( R_{O(1)-H-N(10)} \) coordinate (consistent with the similar electronic configurations of the S1 and T1 states), hinting at a similar propensity for ESIPT.

Given these topographies, we can now suggest interpretations for the reported X–HDBO’ emission spectra and their

![Fig. 6](image1.png) Calculated gas phase absorption and emission profiles of (a) HBO and (b) HDBO. The blue trace reflects vertical excitation from the S0 enol minimum, while the calculated emissions displayed in red and purple originate from, respectively, the keto and enol minima on the S1 PES. For display purposes, the three spectra in each panel are shown with the same peak areas.

![Fig. 7](image2.png) PE profiles for ground and first few singlet and triplet excited states of Zn2+-HDBO’ along the \( R_{O(1)-H-N(10)} \) ESIPT coordinate. The black and red filled circles show the relaxed profiles for, respectively, the S0 and S1 (11st*(CT)) states. The open black circles show the energy of the S2 state computed at the relaxed S1 state geometries. Other PE profiles shown by colored open circles are for the excited states indicated, calculated at the relaxed S0 state geometries.
The present analysis suggests that, as with HBO and HDBO, the peak wavelengths for S1 emission from both the keto and enol forms of Zn2+-HDBO', Cd2+-HDBO' and Hg2+-HDBO' and the experimentally measured emission spectra of these same complexes in aqueous solution (from ref. 31 and shown by the solid black curves). The calculated wavelengths for vertical S1 \rightarrow S0 absorption from the respective optimised S0 geometries are shown by blue sticks, while the peak wavelengths for S1 \rightarrow S0 emission from the optimised S1 enol and keto geometries are shown by, respectively, the purple and red sticks. The different stick heights reflect the respective calculated oscillator strengths (shown on the right hand vertical axis).

Fig. 8 Comparison between calculated gas phase absorption and emission maxima for (a) Zn2+-HDBO', (b) Cd2+-HDBO' and (c) Hg2+-HDBO' and the experimentally measured emission spectra of these same complexes in aqueous solution (from ref. 31 and shown by the solid black curves). The calculated wavelengths for vertical S1 \rightarrow S0 absorption from the respective optimised S0 geometries are shown by blue sticks, while the peak wavelengths for S1 \rightarrow S0 emission from the optimised S1 enol and keto geometries are shown by, respectively, the purple and red sticks. The different stick heights reflect the respective calculated oscillator strengths (shown on the right hand vertical axis).

low fluorescence quantum yields. The experimental emission spectra are reproduced in Fig. 8, along with sticks showing the calculated wavelengths of peak S1 \rightarrow S0 absorption and of peak emission from both the enol and keto minima on the S1 PES. The present analysis suggests that, as with HBO and HDBO, the longest wavelength emission arises from the keto S1 minimum. This emission shows a substantial bathochromic shift, larger than in the cases of HBO or HDBO, and consistent with experiment. The low fluorescence quantum yield from the X-HDBO' complexes is plausibly attributed to competition from internal conversion via E \rightarrow Z isomerisation of the keto tautomer, as demonstrated here for the simpler HBO system. Nonetheless, HDBO' satisfies many of the requirements of a sensor for trace metal detection, i.e. the X-HDBO' complex absorbs at (short) visible wavelengths and fluoresces with a sufficient quantum yield and a large Stokes shift.

As Fig. 8 shows, the experimental spectra display an additional feature at \sim 540 nm that has no analogue in the cases of HBO and HDBO. Xu and Pang noted the very similar excitation profiles for the 540 and 712 nm emissions from Zn2+-HDBO' and concluded that both emissions were from the same complex. Given the PE profiles shown in Fig. 7, we can propose the following explanation for the emissions observed in the solution phase experiments. Vertical excitation projects the X-HDBO' complex to energies above the S1 minimum. The excited state molecules so formed will evolve under the influence of the S1 PES, and start dissipating internal energy by interacting with the surrounding solvent. Some fraction of the S1 state population will undergo ESPT. This fraction can then make a radiationless transition to the ground state (e.g. by E \rightarrow Z isomerism) or relax into the S1 keto minimum and decay by fluorescing (at \lambda_{em} \sim 712 nm). Any S1 state population that relaxes into the S1 enol minimum (R_{X(1)-H} = 1.0 Å), however, would likely fluoresce at \lambda_{em} \sim 540 nm as suggested by Xu and Pang. Recalling the experimental emission profiles (Fig. 8), it is tempting to suggest that the latter process is (relatively) most important in Cd2+-HDBO'. This could be explained if the energy barrier to enol \rightarrow keto tautomerisation on the S1 PES is larger in Cd2+-HDBO' than in Zn2+-HDBO'. As an alternative, however, the observation could also indicate that, relative to Zn2+-HDBO', the S1 keto form of Cd2+-HDBO' has a smaller barrier to E \rightarrow Z isomerism (and thus a lower fluorescence quantum yield for this tautomer).

We note that the present calculations for the isolated (i.e. gas phase) systems cannot allow for the additional complexities that would manifest from considering the bulk aqueous environment. Electronic absorptions to charge-separated states typically show a bathochromic shift in polar solutions (reflecting the stabilisation of the charge-separated state). Notwithstanding, the absorption and emission maxima returned by the present (isolated molecule) calculations are seen to mimic the experimental (aqueous phase) data sufficiently well to enable interpretation of the observed photophysics. Inspecting Fig. 8, it is clear that the peak emission wavelength of the enol-form of Zn2+-HDBO' is produced least well; relative to the other X-HDBO' species, we predict a notably smaller Stokes shift for Zn2+-HDBO', implying a smaller change in equilibrium geometry upon S1 \rightarrow S0 excitation in this case. It is unclear at this stage to what extent this is a real effect, or rather a reflection of the different basis sets used for Zn2+-HDBO' (cf. Cd2+-HDBO' and Hg2+-HDBO').

Having offered a rationale for the large and, from the sensing perspective, highly desirable Stokes shifts exhibited by these X-HDBO' complexes, we now turn attention to potential...
ESIPT to the diketo \( \text{HDBO} \) analogue, which is predicted to absorb at similar wavelengths to \( \text{HDBO} \), to undergo ESIPT by a C atom (henceforth labelled C(18)) and a two-carbon bridge linking C(18) to C(4). As Fig. 9(a) shows, BC-HBO is an analogue of HDBO in which O(13) is replaced by a C atom (henceforth labelled C(18)) and a two-carbon bridge linking C(18) to C(4). As Fig. 9(a) shows, BC-HBO is predicted to absorb at similar wavelengths to HDBO, to undergo ESIPT to the keto-tautomer upon \( S_1 \leftrightarrow S_0 \) excitation, and to show substantially red-shifted \( S_1 \rightarrow S_0 \) emission. Crucially, however, the two-carbon bridge linking C(18) to C(4) not only holds all atoms in a common plane, but should mitigate against any non-radiative population loss via rotation about \( \phi_{C(3)-C(9)} \) in the keto-conformer. Fig. 9(b) depicts another analogue of HDBO, quinolino[8,7-h]quinolone-5-11-diol (henceforth QQ), which is calculated to display similar tendencies for ESIPT, red-shifted \( S_1 \rightarrow S_0 \) emission and, again, the additional structural rigidity to preclude excited state \( E \rightarrow Z \) isomerism as a rival to fluorescence decay. Analogy with HDBO and, particularly, the deprotonated variant of \( \text{HDBO}^-' \), suggests that suitably derivitized variants of either of these ligands could constitute improved ESIPT-based sensors for trace metal detection.

4. Conclusions

The present computational study reveals mechanistic details of a prototypical system that can function as a fluorescent sensor for trace metal cations. We have scanned PE profiles along selected coordinates for the ground and first few excited states for the target metal-chelated complexes (X–HDBO, with X = Zn\( ^{2+} \), Cd\( ^{2+} \) and Hg\( ^{2+} \)) and for simpler molecular building blocks like HBO and HDBO from which these complexes derive, and shown that all are likely to undergo ESIPT following \( S_1 \leftrightarrow S_0 \) photoexcitation. The initial excitation in all of these cases involves a \( \pi^* \leftrightarrow \pi \) transition, that induces charge separation in the \( S_1 \) state and drives a subsequent proton transfer which, in these molecules, is synonymous with \( \text{enol} \rightarrow \text{keto} \) tautomerism. In each case, the resulting keto-tautomer constitutes a local minimum on the \( S_1 \) PES, and is identified as the carrier of Stokes shifted fluorescent emission – consistent with recent experimental studies\(^{34}\) that reveal measurable fluorescence quantum yield from this tautomer. The present study succeeds in reproducing the relative energies of the observed \( S_1 \leftrightarrow S_0 \) \( \text{enol} \) absorptions across the series HBO, HDBO, protonated and deprotonated HDBO' and the X–HDBO' complexes themselves, and of their respective \( S_1(\text{keto}) \rightarrow S_0 \) emissions. We also suggest \( S_1(\text{enol}) \rightarrow S_0 \) fluorescence as an explanation for hitherto unassigned shorter wavelength emission from the X–HDBO' complexes. Derivatives of HDBO' are also proposed that are predicted to retain attractive features like (short wavelength) visible absorption and heavily Stokes shifted emission but, additionally, offer higher fluorescence quantum yields (\( i.e. \) enhanced metal sensing capability) by precluding possible non-radiative decay via \( E \rightarrow Z \) isomerism from the \( S_1 \) keto minimum to the \( S_0 \) PES.

Acknowledgements

The authors are grateful to the Engineering and Physical Sciences Research Council via Programme grant EP/L005913/1. BM and TNVK are grateful to TUM for the award of postdoctoral research fellowships. The underlying data upon which this manuscript is based has been stored in the University of Bristol research data repository and can be access via the following link: DOI: 10.5523/bris.836bu3nhnbzg1pszfc2xowr44.

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