Rationalization of the X-ray photoelectron spectroscopy of aluminium phosphates synthesized from different precursors

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The aim of this paper is to clarify the assignments of X-ray photoelectron spectra of aluminium phosphate materials prepared from the reaction of phosphoric acid with three different aluminium precursors [Al(OH)₃, Al(NO₃)₃, and AlCl₃] at different annealing temperatures. The materials prepared have been studied by X-ray photoelectron spectroscopy (XPS), powder X-ray diffraction (XRD), infrared spectroscopy and high-resolution solid-state ³¹P NMR spectroscopy. A progressive polymerization from orthophosphate to metaphosphates is observed by XRD, ATR-FTIR and solid state ³¹P NMR, and on this basis the oxygen states observed in the XP spectra at 532.3 eV and 533.7 eV are assigned to P–O–Al and P–O–P environments, respectively. The presence of cyclic polyphosphates at the surface of the samples is also evident.
a shift in the O(1s) peak from 531.4 eV for the orthophosphate to a broader peak at 531.8 eV for the metaphosphate. The latter peak clearly showed evidence for a second component at about 533.5 eV. However, the O(1s) data from the electrochemically treated surface were not so conclusive; the O(1s) spectrum of the “metaphosphate” was significantly broader than that of the orthophosphate and seemed to consist of two components. It is not clear whether the component due to bridging oxygens is at higher or lower binding energy than the component due to terminal oxygens.

In the present study, we have used solid-state $^{31}$P NMR, powder XRD and FTIR data to explore the structural changes that occur in aluminium phosphate materials synthesized from 3 : 1 mixtures of phosphoric acid and an aluminium precursor (either aluminium hydroxide, aluminium nitrate or aluminium chloride), which throw light on the information available from the surface specific XPS technique. Our results show that the expected polymerization occurs in materials prepared from all three precursors, and largely confirm the assignments of XP spectra based on existing literature. However, we also find evidence for the presence, at the surface, of polyphosphate species that do not contain aluminium, which distorts the Al : O : P ratios established from the XPS spectra.

## Experimental

### Synthesis of aluminium phosphate powders

The aluminium precursor $\text{AlX}_3$, where $X = \text{OH}$ (Sigma-Aldrich, Reagent Grade), $\text{NO}_3$ (Sigma-Aldrich, ACS Reagent, >98%) or $\text{Cl}$ (Sigma-Aldrich, 99.99%), was added to water [50 ml] at 298 K and mixed until dissolved. $\text{H}_3\text{PO}_4$ (8.75 ml, Sigma-Aldrich, 85% wt, 99.99% purity) was added to the $\text{AlX}_3$ solution to give a P : Al ratio of 3 : 1 and the solution was allowed to mix for 30 min. The reaction mixture was then heated on a hotplate to evaporate the water, resulting in the formation of a viscous gel. Separate samples of the gel were heated in air in a furnace for a period of one hour at three different temperatures (300, 500 and 800 °C), and then allowed to cool to room temperature before analysis. In the case of the materials prepared from $\text{Al} (\text{OH})_3$ and $\text{Al} (\text{NO}_3)_3$, a slightly wet white powder was obtained on heating to 300 °C which was observed (by eye) to dry completely on heating at the higher temperatures. However, the samples prepared from $\text{AlCl}_3$ remained gel-like at 300 °C. These samples were stable under vacuum and so could be analysed by XPS, but they had to be heated to 400 °C to obtain a solid that was suitable for analysis by solid-state NMR and ATR could be obtained. In the discussion below, the materials prepared from the $\text{Al} (\text{OH})_3$, $\text{Al} (\text{Cl})_3$, and $\text{Al} (\text{NO}_3)_3$ precursors are labelled $\text{AlPOH}$, $\text{AlPCl}$ and $\text{AlPNO}_3$, respectively.

### Materials characterization

XP spectra were recorded at room temperature on the powder samples using a Kratos Axis Ultra-DLD photoelectron spectrometer with a monochromatic Al K$\alpha$ X-ray source in the “hybrid spectroscopy” mode resulting in an analysis area of 700 × 300 μm$^2$ at a pass-energy of 40 eV for high-resolution scans and 160 eV for survey scans. The XPS data were analysed using CasaXPS with all binding energies referenced to the C(1s) peak at 284.7 eV with an uncertainty of ~0.2 eV. Since intensities for powder samples are dependent on the surface area analysed, which is poorly reproducible between different powder samples, the XP spectra shown in the figures are normalized to the point of maximum intensity. Curve fits were made using Gaussian-Lorentzian (GL [30]) line-shapes.

Powder X-ray diffraction (XRD) data were recorded at room temperature using a PANalytical X’Pert Pro diffractometer with a monochromatic Cu K$\alpha$ source ($\lambda = 0.154$ nm) operating at 40 kV and 40 mA. The data were recorded over the 2θ range 10–80° with a step size of 0.016°.

High-resolution solid-state $^{31}$P NMR spectra were acquired at room temperature on a Chemagnetics Infinity Plus spectrometer ($^{31}$P Larmor frequency, 121.50 MHz). The samples were contained in a 4 mm rotor with magic-angle spinning at 12 kHz. Methylidiphenylphosphine oxide (MDPPO) was used as a reference, with $^{31}$P chemical shift at 30.8 ppm.

FTIR spectra were recorded using a germanium crystal ATR on a Varian 3100 Excalibur system with Varian Resolutions Pro software.

## Results

### Bulk structural analysis using FTIR, XRD and solid-state $^{31}$P NMR

Throughout the discussion, we refer to each sample by the notation $\text{AlPOH}(T)$, $\text{AlPCl}(T)$ or $\text{AlPNO}_3(T)$, where the subscript identifies the precursor used in the synthesis and $T$ denotes the annealing temperature. All the analysis was performed after cooling to room temperature.

The crystalline phases present within each sample were investigated by powder XRD. Fig. 1, shows the XRD patterns for the $\text{AlPOH}$ samples. For $\text{AlPOH}(300)$ [i.e., the sample prepared from the $\text{Al} (\text{OH})_3$ precursor and annealed at 300 °C], only low intensity peaks are observed in the XRD data, and we have been unable to definitively match these peaks to a known structure. For the $\text{AlPOH}(500)$ sample, the XRD data show sufficient crystallinity to allow Le Bail fitting, although Rietveld refinement was not possible. The Le Bail fitting confirms that two distinct phases are present: a cubic aluminium metaphosphate [ICSD: 26759] and an aluminium hexacyclophosphate (ICSD: 260723). For the $\text{AlPOH}(800)$ sample, only the cubic aluminium metaphosphate is present. The results from Le Bail fitting are shown in Fig. S1 and S2 (in ESI†).

d’Yvoire14,15 also reported a pure cubic phase for a material prepared from $\text{Al} (\text{OH})_3$, and annealed at 800 °C, with a second phase present in the material annealed at 500 °C. However, d’Yvoire assigned the second phase to a monoclinic structure rather than the aluminium hexacyclophosphate observed here.

XRD indicates that the $\text{AlPNO}_3(500)$ and $\text{AlPNO}_3(800)$ samples (Fig. S3 and S4†) are a mixture of the cubic aluminium metaphosphate and aluminium hexacyclophosphate phases [similar to $\text{AlPOH}(500)$ but different from $\text{AlPOH}(800)$]. Among the samples prepared from the $\text{AlCl}_3$ precursor, the only crystalline
product was AlPCl(800), identified from XRD as pure cubic Al(PO3)3 (Fig. S5†).

Further structural insights are obtained from high-resolution solid-state 31P NMR spectra (Fig. 2). For several of the samples, a peak at 0 ppm is present and assigned as the phosphoric acid starting material. As the annealing temperature increases, there is a trend towards increasingly negative 31P chemical shifts, attributed to polymerization. For AlPOH(300), a broad set of overlapping peaks is observed between 5 ppm and −40 ppm, possibly suggesting an amorphous structure. The peaks at −21 ppm and −32 ppm (which represent ca. 20% of the total signal) are assigned† to aluminium tripolyphosphate (AlH2P3O10·H2O). These peaks are also observed for AlPNO3(300). For AlPOH(800), only one peak is observed (at 50.5 ppm) and is attributed unambiguously to cubic Al(PO3)3, consistent with the presence of a single phosphorus environment in this structure (Fig. 3). This peak is also present for the AlPOH(500) sample, together with peaks at −36.5 ppm and −43.0 ppm; the area ratio for these two peaks is 2 : 1, consistent with the presence of three crystallographically distinct phosphorus environments in the aluminium hexacyclophosphate phase (Fig. 3). Monoclinic Al(PO3)3, on the other hand, has 9 distinct phosphorus environments (Fig. 3).

For the AlPNO3(300) sample, three peaks are observed between 5 ppm and −35 ppm [in contrast to the overlapping set of peaks observed in this region for AlPOH(300)], including peaks at −21 ppm and −32 ppm assigned to aluminium tri-polyposphate (AlH2P3O10·H2O), which represents ~50% of the signal. The peak at −3.4 ppm is attributed to some remaining aluminium orthophosphate. For the AlPNO3(500) sample, the 31P NMR spectrum contains peaks characteristic of the hexacyclophosphate (43.3%) and cubic metaphosphate (52%). Annealing at a higher temperature does not complete the transformation to the cubic metaphosphate as the 31P NMR spectrum for the AlPNO3(800) sample clearly contains peaks due to the hexacyclophosphate phase (ca. 36% of the signal intensity).

The 31P NMR spectra for AlPCl(400) and AlPCl(500) are significantly different from those observed for the AlPOH and AlPNO3 materials. The major peaks are due to cubic metaphosphate (50.5 ppm), orthophosphate (1 ppm and 0 ppm), pyrophosphate (−12 ppm and −13.5 ppm) and polyphosphate (peaks in the range −20 ppm to −28 ppm), with only very weak peaks observed for hexacyclophosphate. In contrast, the AlPCl(800) sample is a pure phase of cubic Al(PO3)3. These observations suggest that the metaphosphate formed from the AlCl3 precursor may be produced via a slightly different pathway than from the Al(OH)3 and Al(NO3)3 precursors. We deduce that the stability of the Al–Cl bond hinders formation of aluminium phosphate from the Al(Cl)3 precursor at lower temperatures, leaving the phosphoric acid to react mostly with itself to form varying degrees of polyphosphates. However, at a higher annealing temperature of 800 °C, the phosphate transforms completely to cubic Al(PO3)3.

Surface analysis with XPS

The XP spectra in Fig. 4 show the O(1s) data for all samples. For samples prepared from each of the three precursors, there is a general shift in peak position towards lower binding energy as the annealing temperature is increased. Curve fitting confirms that two distinguishable peaks are present at binding energies of ca. 533.7 eV and 532.3 eV in all spectra, with a transfer of intensity from the higher binding energy peak to the lower binding energy peak as the annealing temperature is increased (peak area ratios are given in Table 1). In particular, we note the close similarity between the spectra for the AlPOH(800), AlPNO3(800) and AlPCl(800) samples, all of which show an approximately 2 : 1 intensity ratio (lower : higher binding energy peak) and resemble the O(1s) spectrum of aluminium metaphosphate published by Rotole and Sherwood.32

The Al(2p) XP spectra for all samples are shown in Fig. 5. For the AlPOH samples, a strong peak is present for all annealing temperatures at ~75.3 eV for AlPOH(300) and shifting slightly to
Fig. 2 High-resolution solid-state $^{31}$P NMR spectra of the aluminium phosphate materials prepared from the Al(OH)$_3$, Al(NO$_3$)$_3$ and AlCl$_3$ precursors at different annealing temperatures. All spectra were recorded at room temperature.

Fig. 3 Structures of (a) aluminium orthophosphate (AlPO$_4$), (b) monoclinic aluminium metaphosphate, (c) aluminium hexacyclophosphate, and (d) cubic aluminium metaphosphate.
Fig. 4 X-ray photoelectron spectra in the O(1s) region for the aluminium phosphate materials prepared from the Al(OH)₃, Al(NO₃)₃ and AlCl₃ precursors at different annealing temperatures. Solid lines show experimental data. The dotted curves show deconvolution using the methods described in the text. For all three precursors, the intensity of the peak at lower binding energy increases relative to the peak at higher binding energy as the annealing temperature is increased.

Table 1 Quantification of XP spectra

| Reactant | Annealing temperature/°C | 300/400 | 500 | 800 |
|----------|---------------------------|---------|-----|-----|
| Al(OH)₃  | O(1s) binding energy<a>  | 533.7   | 532.3| 533.5| 532.0| 533.8| 532.1|
|          | % of peak area            | 55.5    | 44.5| 44.6| 55.4| 39.0| 61.0|
|          | ~532 : 533 ratio         | 1.2     | 0.8 | 1.6 | 0.8 |
|          | O : P ratio               | 2.9     | 3.2 | 3.0 | 3.2 |
|          | P : Al ratio              | 3.8     | 5.3 | 5.3 | 5.3 |
| Al(NO₃)₃ | O(1s) binding energy<a>  | 533.6   | 532.2| 533.8| 532.3| 533.7| 532.1|
|          | % of peak area            | 64.4    | 30.9| 64.4| 35.6| 38.5| 61.5|
|          | ~532 : 533 ratio         | 1.6     | 0.43| 0.35| 1.6 |
|          | O : P ratio               | 2.9     | 2.8 | 2.9 | 3.4 |
|          | P : Al ratio              | 3.6     | n/a | 10.2| n/a |
| AlCl₃    | O(1s) binding energy<a>  | 534.0   | 532.5| 533.9| 532.4| 533.7| 532.1|
|          | ~532 : 533 ratio         | 40.2    | 62.4| 40.2| 40.2| 40.2| 40.2|
|          | Peak area ratio          | 1.5     | 0.6 | 0.7 | 0.6 |
|          | O : P ratio              | 2.8     | 3.3 | 3.1 | 3.3 |
|          | P : Al ratio             | 4.6     | n/a | n/a | n/a |

<a> Binding energies are referenced to the C (1s) peak at 284.7 eV; curves are fitted with Gaussian–Lorentzian (GL (30)) line shapes.
75.1 eV for AlPOH(800). However, for the AlPNO3(300) sample, no peaks are observed in the Al(2p) XP spectrum, while a peak at ca. 75 eV is present for both AlPNO3(500) and AlPNO3(800). For the AlPCl samples, again no peaks are observed in the Al(2p) XP spectra for the AlPCl(300) and AlPCl(500) samples, while a peak is observed at ca. 75 eV for AlPCl(800). This difference in behaviour parallels the physical appearance of the AlPCl samples, for which both AlPCl(300) and AlPCl(500) are tacky but AlPCl(800) is not.

The XP spectra in the P(2p) region (Fig. 6) have a single peak at ~134.8 eV for all samples with a small shift (~0.2 eV) to lower binding energy as the annealing temperature is increased to 800 °C. The observed peak is consistent with the average peak position for metaphosphates in the NIST database\textsuperscript{17} (134.8 eV; \(\sigma = 0.5\) eV) and with results of Rotole and Sherwood\textsuperscript{10,11} on aluminium phosphates. However, it is in marked contrast to sodium phosphates,\textsuperscript{4} for which the P(2p) binding energy shifts by 2 eV from the orthophosphate (132.5 eV) to the oxygen-bridged metaphosphate (134.5 eV).

The atomic ratios calculated from the XPS spectra (Table 1) are informative. The P : O ratio is very close to 1 : 3 for all samples but the P : Al ratio is always higher than 3 : 1. The fact that the XPS survey scans for materials prepared at lower annealing temperatures (Fig. S6\textsuperscript{†}) do not contain any signal for aluminium or for any other cation suggests that the surface is dominated by hydrogen phosphates, with the consistent O : P ratio of 3 : 1 indicating extensive polymerization at the surface. We also note that no XPS signal due to chlorine is observed for any of the samples (Fig. S7\textsuperscript{†}). Following annealing at 800 °C, the samples from all precursors are highly crystalline, and the presence of the Al(2p) peak in the XP spectra suggests that the surface is now dominated by aluminium metaphosphate. However, the P : Al ratio remains higher than the expected 3 : 1 ratio, particularly for AlCl3(800), suggesting that some hydrogen polyphosphates are present at the surface.

**Surface analysis with ATR-FTIR**

The sampling depth of ATR-FTIR spectroscopy, typically between 0.5–2 \(\mu\)m, is significantly larger than that for XPS, which detects only the top 2–4 nm of the surface for the elements studied here. Nevertheless, there is excellent agreement between the IR results and the results from the more surface sensitive XPS method, as illustrated in Fig. 7, which shows FTIR data for several samples.
For AlPOH(300), broad, weak bands are present in the range ca.
\(1100\)–\(1250\) cm\(^{-1}\), characteristic of an aluminium orthophosphate
with some indication of P–O–P bond formation from the weak
peak at \(1025\) cm\(^{-1}\). The spectra for AlPOH(500) and AlPOH(800)
are dominated by strong bands assigned to metaphosphates. In
particular, the peak at \(738\) cm\(^{-1}\) is assigned to Al–O–P, bands at
\(811, 1025, 1060\) and \(1070\) cm\(^{-1}\) are assigned to P–O–P modes, and
bands at \(1282\) and \(1305\) cm\(^{-1}\) are assigned to P–O bonds in the
aluminium metaphosphate.

**Discussion**

The structures of aluminium orthophosphate, aluminium hexacyclophosphate and two aluminium metaphosphates are shown
in Fig. 3, and can be used to rationalize the changes in the XP
spectra observed for different annealing temperatures. Small
shifts to lower binding energy are observed for the P(2p) and
Al(2p) peaks as the annealing temperature is increased, but the
most significant changes arise in the O(1s) spectra, which show
a transfer of intensity from a higher binding energy peak at
\(533.7\) eV to a lower binding energy peak at \(532.3\) eV. The best
starting point to understand these changes is the O(1s) spectra for
all samples annealed at \(800\) °C, as these spectra are all very
similar to those reported by several other authors, including
Gresch *et al.* for sodium metaphosphates,* Crobu *et al.* for zinc
phosphates,* and Rotole and Sherwood for model aluminium
phosphates.\(^{11,12}\) The FTIR, XRD and solid-state \(^{31}\)P NMR data
indicate the presence of mainly cubic aluminium metaphosphate
after annealing at \(800\) °C and we can therefore definitively assign
the peaks in the O(1s) region at \(533.7\) eV and \(532.3\) eV, respectively,
to oxygen atoms bridging between phosphorus atoms (P–O–P) and oxygen atoms bridging between phosphorus and
aluminium atoms (P–O–Al), in agreement with Gresch *et al.*

In the metaphosphate, these two bonding environments are
expected to be present in a \(2 : 1\) ratio of P–O–Al to P–O–P.
However, as shown in Table 1, quantification of the XPS data for
the samples annealed at \(800\) °C gives a peak area ratio (532.3
eV : 533.7 eV) of ca. \(1.6 : 1\), whereas the expected ratio is \(2 : 1\).
Thus, the P : Al ratio at the surface of these materials is higher
than the expected \(3 : 1\) ratio. To understand these differences,
we now consider the XP spectra recorded for samples annealed to lower temperatures.

A key observation is that the AlPNO3(300) and AlPCl(300) samples show no evidence, in the XP spectra, for the presence of aluminium. The AlPNO3(500) sample does show evidence for aluminium, but the AlPCl(500) sample does not. The absence of aluminium indicates a purely hydrogen terminated phosphate material at the surface. Unreacted phosphoric acid can be ruled out based on the O : P ratio of 3 : 1, but there is evidence from the solid-state $^{31}$P NMR results for hydrogen terminated or cyclic polyphosphates (giving peaks at $-28$ ppm and $-32$ ppm) which would have a 3 : 1 ratio. A cyclic polyphosphate such as $\text{P}_4\text{O}_{10}$ has a P–O–Pt oP $\text{O}$ bond ratio of 1.5 : 1, which could account for the XPS ratios if the oxygen in P–O overlaps with the XPS peak for the oxygen in P–O–Al. This assignment would be in agreement with Gresch et al. Finally, the “wet” physical appearance of samples annealed at lower temperatures is also consistent with the presence of hydrogen terminated polyphosphates which would be poorly crystalline.

From the data presently available, we cannot determine whether annealing ultimately leads to sublimation or decomposition of the polymeric phosphates, or whether further reaction with unreacted aluminium precursor occurs. However, for samples annealed at $800 \degree \text{C}$, the XP spectra are consistent with the presence of aluminium metaphosphates although the slightly higher P : Al ratio in the case of the material prepared from AlCl$_3$ suggests the surface contains some hydrogen terminated polyphosphates.

**Conclusions**

Aluminium metaphosphate is formed from the reaction of phosphoric acid with three different aluminium compounds (Al(OH)$_3$, Al(NO$_3$)$_3$ and AlCl$_3$) followed by annealing in air. XRD, XPS and FTIR measurements of the resulting materials show almost identical behaviour from all three precursors, but the solid-state $^{31}$P NMR spectra are significantly different at the lower annealing temperatures ($300 \degree \text{C}$ and $500 \degree \text{C}$). The unique solid-state $^{31}$P NMR spectra of the materials annealed at lower temperatures indicates the presence of amorphous materials which would not be identified by XRD, but explains the lack of an Al(2p) signal in XP spectra of the materials prepared from the Al(NO$_3$)$_3$ and AlCl$_3$ precursors at lower annealing temperatures. For all three precursors, a cubic metaphosphate is produced on annealing at $800 \degree \text{C}$, with a hexacyclophosphate present at lower annealing temperatures in the case of the Al(NO$_3$)$_3$ and Al(OH)$_3$ precursors. The XP spectra in the O(1s) region of the aluminium phosphate materials show two components at $532.3 \text{ eV}$ and $533.7 \text{ eV}$, which are definitively assigned to the P–O–Pt oP and P–O–P bonding environments, respectively. However, samples annealed at lower temperatures also exhibit surface species assigned as cyclic polyphosphates, with binding energies of $532.3 \text{ eV}$ and $533.7 \text{ eV}$ for the P–O and P–O–P bonding environments, respectively.

**Conflicts of interest**

There are no conflicts to declare.

**Acknowledgements**

VB, MT and LE are grateful for studentships funded by Cogent Power Ltd. XP spectra were recorded by the EPSRC National Facility for Photoelectron Spectroscopy (HarwellXPS), operated by Cardiff University and UCL under contract number PR16195. The UK $850 \text{ MHz}$ solid-state NMR Facility used in this research
was funded by EPSRC and BBSRC (contract reference PR140003), as well as the University of Warwick including via part funding through Birmingham Science City Advanced Materials Projects 1 and 2 supported by Advantage West Midlands (AWM) and the European Regional Development Fund (ERDF).

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