Green synthesis of benzonitrile using ionic liquid with multiple roles as the recycling agent†

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Preparation of benzonitrile from benzaldehyde and hydroxylamine hydrochloride is one of the most advantageous approaches. Nevertheless, it suffers from various constraints such as longer reaction time, corrosion and recovery of hydrochloric acid, the use of metal salt catalysts and their separation. For these reasons, a novel green benzonitrile synthetic route was proposed with ionic liquid as the recycling agent in this study. The results indicated that hydroxylamine 1-sulfobutyl pyridine hydrotosolate salt ([NH2OH]2−[HSO3−-Py]-HSO4) was an expert alternative to hydroxylamine hydrochloride. Meanwhile, the ionic liquid [HSO3−-Py]-HSO4 exhibited the multiple roles of co-solvent, catalysis and phase separation, thus the use of metal salt catalyst was eliminated, and no additional catalyst was needed. Hence, the separation process was greatly simplified. When the molar ratio of benzaldehyde to (NH2OH)2−[HSO3−-Py]-HSO4 was 1 : 1.5, the volume ratio of paraxylene to [HSO3−-Py]-HSO4 was 2 : 1, the benzaldehyde conversion and benzonitrile yield were both 100% at 120 °C in 2 h. Even better, the ionic liquid could be recovered easily by phase separation, and recycled directly after reaction. Additionally, this novel route is applicable to the green synthesis of a variety of aromatic, heteroaromatic and aliphatic nitriles with excellent yields.

Introduction

Benzonitrile is a versatile chemical that plays key roles in the synthesis of benzoic acid, benzyamine, benzamide, pesticides and dyes. It is also a prominent intermediate for the production of benzoguanamine, which is a widely used advanced coating. Classical methods for benzonitrile synthesis are the cyanation of benzene halides, toluene halides and ammonia reaction, ammoniation of toluene, ammonia and air, benzoic acid and urea reaction, and the cyanation of benzaldehyde. Among them, preparation of benzonitrile from benzaldehyde and hydroxylamine hydrochloride (NH2OH·HCl) is one of the most interesting and advantageous approaches due to the constant carbon number in the reactant and final product, mild reaction conditions, low production cost and prospect for industrial-scale application.

The transformation of benzaldehyde and NH2OH·HCl to benzonitrile can be grouped into two-step and one-pot methods. For the two-step reaction, it is carried out to form benzaldoxime, and then the isolated benzaldoxime is dehydrated to benzonitrile. Both of the two steps can be catalyzed by acid. As for the one-pot route, it includes oxidation of benzaldehyde to benzaldoxime, followed by the dehydration of in situ generated benzaldoxime. Obviously, the one-pot reaction helps to reduce the energy consumption, solvent waste and reaction time, which in turn, substantially avoids the isolation of benzaldoxime. It is a valuable strategic way in terms of economic and environmental aspects. Owing to these advantageous features, several studies for the one-pot synthesis of benzonitrile have been conducted in recent years.

Enthaler et al. presented a direct conversion of benzaldehyde to benzonitrile catalyzed by Zn(CH3COO)2 (5 mol%). When benzaldehyde was 0.72 mmol, n(benzaldehyde) : n(NH2·OH·HCl) = 1 : 1.2, toluene was 2 mL, reaction temperature was 100 °C, reaction time was 24 h, the benzonitrile yield was 79.0%. Zhuang et al. studied SnCl2·2H2O (10 mol%) catalyzed transformation of benzaldehyde to benzonitrile. When benzaldehyde was 1 mmol, n(benzaldehyde) : n(NH2OH·HCl) = 1 : 1.1, CH2CN was 0.5 mol, NaHCO3 was 1 mmol, the benzonitrile yield was 80.0% at 80 °C within 24 h. Leggio et al. investigated the catalytic properties of TiCl4. When benzaldehyde was 1 mmol, n(benzaldehyde) : n(NH2OH·HCl) = 1 : 2, pyridine was 2 mL, TiCl4 was 2.5 mmol, the benzonitrile yield was 85.0% at 40 °C in 3 h. Wang et al. reviewed the synthesis of benzonitrile catalyzed by NaHSO4·H2O. Under the optimized conditions of molar ratio of n(benzaldehyde) : n(NH2OH·HCl) = 1 : 1.4, and...
reaction at 135 °C for 4.5 h in the presence of tetrabutylammonium bromide as phase transfer catalyst and DMF as solvent, benzonitrile was obtained in excellent yield of 92.8%. In addition to homogeneous catalysts, heterogeneous catalysts were also reported. Fe3O4-CTAB NPs (1.8 mol%) was applied as a competent catalyst for one-pot synthesis of benzonitrile.14 The combination of benzaldehyde (0.5 mmol), NH2OH·HCl (0.75 mmol), Fe3O4-CTAB NPs (1.8 mol%), and DMF (5 mL) at 80–90 °C for 1 h were found to be the optimized reaction conditions for the desired transformation, with the benzonitrile yield being 97.0%. A green chitosan supported magnetic ionic liquid (CSMIL) was prepared for direct conversion of benzaldehyde to benzonitrile. When benzaldehyde was 1 mmol, NH2OH·HCl was 1.2 mmol, dry-CSMIL was 15 mg, MeSO2Cl was 1.2 mmol, the benzonitrile yield was 93.0% (70 °C, 1.5 h). It can be seen that benzonitrile could be efficiently synthesized from benzaldehyde and NH2OH·HCl. However, it suffers from various constrains such as longer reaction time, use of metal salt catalysts and inconvenient post treatment.24,25 Furthermore, the using of NH2OH·HCl will inevitably lead to problems of equipment corrosion and environmental pollution because it releases HCl.

Recently, a series of ionic liquids (ILs) were used as alternatives to conventional inorganic acids in hydroxylamine stabilization, leading to the formation of several eco-friendly hydroxylamine liquid salts ([NH2OH]2·ILs) that exhibited better reactivity in one-step clean synthesis of caprolactam, phenol and aniline in comparison with NH2OH·HCl and [NH2OH]2·H2SO4.26 In continuation of our work for application of these [NH2OH]2·ILs, a green and high efficient reaction route was developed for direct synthesis of benzonitrile from benzaldehyde and (NH2OH)2·ILs. With ILs as the co-solvent, catalyst, phase separation and recycling agent, as shown in Scheme 1. Compared with the existing method with NH2OH·HCl as raw material, this novel route has the following advantages: (1) it is eco-friendly since no metal salt catalyst was added and no inorganic acid was released. (2) When the reaction was completed, the reaction mass was separated into organic phase and aqueous phase automatically. The organic phase could be analyzed by gas chromatography and the ionic liquid in aqueous phase could be easily recovered. (3) The recovered ionic liquid not only can be recycled as the co-solvent and catalyst, but also as the hydroxylamine stabilizer for preparation of (NH2OH)2·ILs. Thus, the recovery of the ionic liquid with multiple roles was fully realized. Moreover, the one-pot clean synthesis of benzonitrile was achieved.

**Experimental**

**Materials and methods**

All the chemicals (AR grade) were commercially available and used without further purification. FTIR spectra were recorded on a Bruker Vector 22 FTIR spectrometer in the 4000–400 cm−1 range using liquid film or KBr tablet. The 1H NMR and 13C NMR spectra were recorded on a Bruker AMX FT 400 MHz NMR spectrometer using D2O as solvent, and the chemical shifts were expressed in ppm.

**Preparation of the (NH2OH)2·ILs**

The (NH2OH)2·ILs were prepared according to the method described in the literature.26–27 The synthesis of hydroxylamine 1-sulfobutyl pyridine hydrosulfate salt is given here as an example. 1-Sulfobutyl pyridine hydrosulfate ([HSO3-·b-Py]·HSO4, 32.0 mmol) was added into a 250 mL three-necked flask equipped with a dropping funnel and a stirrer. The flask was then placed in a low-temperature reaction bath, and the temperature was continuously kept under 2 °C. The aqueous solution of hydroxylamine (80.0 mmol) was added dropwise to the three-necked flask that contained the [HSO3-·b-Py]HSO4 while stirring. When the neutralization process was completed, a clear solution was obtained. The solution was then evaporated to dryness under reduced pressure to obtain the white hydroxylamine 1-sulfobutyl pyridine hydrosulfate salt ([NH2OH]2·[HSO3-·b-Py]·HSO4). The preparation method of hydroxylamine N,N,N-trimethyl-N-sulfobutyl hydrosulfate salt ([NH2OH]2·[HSO3-·b(N(CH3)3)]·HSO4) and hydroxylamine 1-sulfobutyl-3-methyl imidazole hydrosulfate salt ([NH2OH]2·[HSO3-·b-mim]·HSO4) were similar to that of the [NH2OH]2·[HSO3-·b-Py]·HSO4, the only difference was to replace the [HSO3-·b-Py]·HSO4 with [HSO3-·b-N(CH3)3]·HSO4 or [HSO3-·b-mim]·HSO4.

**One-pot synthesis of benzonitrile from benzaldehyde and (NH2OH)2·ILs**

The reaction was performed in a 100 mL three-necked flask equipped with a stirrer and a reflux condenser. Typically, benzaldehyde (14.4 mmol), (NH2OH)2·[HSO3-·b-Py]·HSO4 (21.6 mmol), [HSO3-·b-Py]·HSO4 (16 mL) and the organic solvent (32 mL) were charged into the flask, and the reaction mixture was heated in an oil bath at 90–120 °C and kept for 0.5–2 h. Along with the reaction, the organic components (1 mL) were taken every five minutes and analyzed with an Agilent 7890B gas chromatograph connected with a DB-WAX capillary column and a flame ionisation detector (FID). The analysis conditions were as follows: N2 as carrier, FID temperature 320 °C, the injection port temperature 250 °C, and the program-controlled column temperature as follows: initial temperature of 90 °C for 1 min, increasing to 225 °C at a rate of 15 °C min−1 and holding for 1 min. The benzaldehyde conversion and benzonitrile yield were determined by the area normalization method.

![Scheme 1](image-url)
Recycling of the ILs

For recycling of the ILs, the one-step synthesis of benzonitrile from benzaldehyde and (NH2OH)2·[HSO3-b-Py]·H2SO4 was carried out under optimal reaction conditions. After the reaction, the resulting mixture was cooled and separated into organic phase and aqueous phase automatically. The organic phase was analyzed by gas chromatography, and the leaving ionic liquid in aqueous phase was easily recovered by distillation with a rotary evaporator under vacuum to remove the water. The recovered ILs could be recycled as the co-solvent and catalyst for one-pot synthesis of benzonitrile. Furthermore, it also could be used again as the hydroxylamine stabilizer to synthesize the (NH2OH)2·[HSO3-b-Py]·H2SO4.

Results and discussion

Co-solvent and catalysis role of the ionic liquid

The effect of solvent on direct synthesis of benzonitrile was presented in Fig. 1. It could be seen that solvent strongly affected the results. When bromobenzene and ortho-dichlorobenzene were used as solvents, the benzaldehyde conversion and benzonitrile yield were relatively low. While toluene, ethylbenzene and paraxylene were used, the benzaldehyde conversion and benzonitrile yield were very high, especially the latter two solvents. The conversion and yield were both improved to over 96.0%. The solvent effect of aryl halide was inferior to that of aromatic hydrocarbon. The probably reason is that, the present reaction includes oximation of benzaldehyde to benzaaldoxime, followed by the dehydration of the in situ generated benzaaldehyde to benzonitrile. Being a multistep process, the outcome of the reaction depends on the reactivity of the substrate towards both oximation and dehydration step. As for the first step, it is believed that benzaaldoxime is produced through the nucleophilic attack of nitrogen electron pairs in the NH2OH to the C==O carbon in the benzaldehyde. Free NH2OH resulted from the decomposition of hydroxylamine salt, is the nucleophile necessary for conversion of benzaldehyde to the corresponding benzaaldoxime. It was evident that, the aromatic hydrocarbon solvent, such as toluene, ethylbenzene and paraxylene, demonstrated an obvious promotion on the decomposition of hydroxylamine salt to form free NH2OH since high benzaldehyde conversions were obtained in these solvents. This solvent promoted decomposition role was probably due to the formation of hydrogen bond between the nitrogen electron pairs in the hydroxylamine salts and hydrogen in the solvent. Among them, paraxylene was a better choice for direct synthesis of benzonitrile.

Besides paraxylene, a certain amount of [HSO3-b-Py]·H2SO4 was also added. To clarify their roles, a comparative study on paraxylene alone, [HSO3-b-Py]·H2SO4 alone and different ratio of paraxylene to [HSO3-b-Py]·H2SO4 was made and the results were listed in Table 1. When paraxylene was added alone, a large number of white particles, which was the undissolved (NH2OH)2·[HSO3-b-Py]·H2SO4., was dispersed in the reaction mass. It indicated that pure paraxylene could not effectively dissolve the (NH2OH)2·[HSO3-b-Py]·H2SO4, which then inhibited the reaction between benzaldehyde and hydroxylamine, resulting in a low benzaldehyde conversion of 54.8%. Moreover, most of the product was benzaaldoxime, the benzonitrile yield was only 2.12%, indicating that paraxylene only acted as solvent and had no catalytic effect. When [HSO3-b-Py]·H2SO4 was added, the reaction turned to liquid–liquid two phase. It could be concluded that the addition of [HSO3-b-Py]·H2SO4 was favorable for the dissolution of (NH2OH)2·[HSO3-b-Py]·H2SO4. Moreover, [HSO3-b-Py]·H2SO4 being a Brønsted acid coordinated with the carbonyl oxygen of benzaldehyde thus increasing its electrophilicity and facilitating its reaction with hydroxylamine. Therefore, the [HSO3-b-Py]·H2SO4 prompted the benzoaldoxime oximation. The benzaldehyde conversion was greatly increased to over 95%. Furthermore, when the [HSO3-b-Py]·H2SO4 was

![Figure 1](image)

**Figure 1** Effect of solvent on direct synthesis of benzonitrile (BT). Reaction conditions: benzaldehyde (BD) 14.4 mmol, n(BD) : n(NH2OH)2·[HSO3-b-Py]·H2SO4 = 1 : 1.5, solvent 32 mL, [HSO3-b-Py]·H2SO4 16 mL, temperature 100 °C.

| No | Paraxylene (mL) | [HSO3-b-Py]·H2SO4 (mL) | XBD (%) | YBT (%) |
|----|-----------------|------------------------|--------|--------|
| 1  | 12              | 0                      | 54.8   | 2.12   |
| 2  | 10              | 2                      | 95.0   | 92.0   |
| 3  | 8               | 4                      | 98.0   | 97.0   |
| 4  | 6               | 6                      | 98.8   | 96.9   |
| 5  | 4               | 8                      | 100    | 96.7   |
| 6  | 0               | 12                     | 100    | 96.3   |
| γ  | 12              | 0                      | 3.71   | 2.43   |
| 8  | 12              | 0                      | 3.34   | 6.29   |

* Reaction conditions: benzaldehyde 3.6 mmol, n(BD) : n(NH2OH) = 1 : 3, temperature 100 °C, time 90 min. ** NH2OH·HCl. ³ [NH2OH]2·H2SO4.
added alone, the reaction was homogeneous. The benzaldehyde conversion and benzonitrile yield were 100% and 96.3% respectively, which were much higher than those in pure paraxylene. The results demonstrated that [HSO3-b-Py]-HSO4 not only can be used as solvent to promote the benzaldehyde oxidation, but also as catalyst to promote the benzaldoxime dehydration.53–57

Considering that when the [HSO3-b-Py]-HSO4 alone was applied as both the solvent and catalyst, the reaction liquid was too viscous to separate and recover the ionic liquid. For this reason, a mixture of paraxylene and [HSO3-b-Py]-HSO4 were employed as solvent for the synthesis of benzonitrile. It can be seen that the benzonitrile yield was related to their ratio. With the increase of the amount of [HSO3-b-Py]-HSO4, the benzonitrile yield improved first, then remained almost unchanged. The benzonitrile yield reached a maximum value of 97.0% when the volume ratio of paraxylene and [HSO3-b-Py]-HSO4 was 2:1. With NH2OH·HCl and [NH2OH]2·H2SO4 as raw materials and paraxylene as a solvent, the benzaldehyde conversion and benzonitrile yield were extremely low, which verified the cosolvent and catalytic role of the [HSO3-b-Py]-HSO4.

Influence of reaction parameters on direct synthesis of benzonitrile

(1) Temperature. The influence of reaction temperature on direct synthesis of benzonitrile was depicted in Fig. 2. It could be seen that the benzaldehyde conversion and benzonitrile yield were relatively low when the temperature was 90 °C. Meanwhile, the reaction rate was slower since it took longer time to achieve the same conversion and yield. However, the benzaldehyde conversion and benzonitrile yield increased significantly with rising temperature from 90 to 120 °C. The reaction rate also increased rapidly, indicating that higher reaction temperature could enhance the reaction rate, favoring the synthesis of benzonitrile. The highest conversion and yield obtained (120 °C) were 99.9% and 99.7% respectively.

(2) Time. The influence of reaction time was also investigated to see whether the benzaldehyde conversion and benzonitrile yield could reach 100% by prolonging the time. As shown in Fig. 3, the benzaldehyde conversion and benzonitrile yield depended greatly on time at the early stage of reaction. They increased rapidly with the extension of time, and reached a value of 98.6% and 98.4% respectively when the time was 1 h. Then the benzaldehyde conversion and benzonitrile yield increased slightly when further prolonging the time. When the reaction time was 2 h, they both reached a value of 100%.

Reactivity comparison of various hydroxylamine salts

A comparison between the (NH2OH)2·[HSO3-b-Py]-HSO4 and other [NH2OH]2·ILs, as well as the traditional hydroxylamine inorganic acid salts was summarized in Table 2. It can be seen that [NH2OH]2·[HSO3-b-mim]-HSO4 and [NH2OH]2·[HSO3-b-Py]-HSO4 demonstrated better reactivity than [NH2OH]2·[HSO3-b-N(CH3)3]-HSO4. The benzonitrile yield was around 99.9% and 100%, respectively. No other by-products were generated.
mechanism for the present reaction was proposed and depicted in Scheme 2. Firstly, free \( \text{NH}_2\text{OH} \) and \([\text{HSO}_3\text{-b-Py}]\cdot\text{HSO}_4\) were released by the decomposition of \([\text{NH}_2\text{OH}]_2\cdot[\text{HSO}_3\text{-b-Py}]\cdot\text{HSO}_4\). Then a nucleophilic attack of \( \text{NH}_2\text{OH} \) to benzaldehyde occurred, followed by a dehydration step to form the benzaldehyde oxime. At last, the hydroxyl group in the benzaldehyde oxime was protonated by the \( \text{H}^+ \) offered by the \([\text{HSO}_3\text{-b-Py}]\cdot\text{HSO}_4\) to form the intermediate, which then converted to benzonitrile through a dehydration step, along with the regeneration of \([\text{HSO}_3\text{-b-Py}]\cdot\text{HSO}_4\) to complete the catalytic cycle.

Plausible mechanism for the present reaction

Based on the reported reaction mechanism for aldehyde and hydroxylamine to produce aldehyde oxime,\(^{30,31}\) followed by benzaldehyde dehydration to benzonitrile,\(^{35,36}\) a plausible mechanism for the present reaction was proposed and depicted in Scheme 2. Firstly, free \( \text{NH}_2\text{OH} \) and \([\text{HSO}_3\text{-b-Py}]\cdot\text{HSO}_4\) were released by the decomposition of \([\text{NH}_2\text{OH}]_2\cdot[\text{HSO}_3\text{-b-Py}]\cdot\text{HSO}_4\). Then a nucleophilic attack of \( \text{NH}_2\text{OH} \) to benzaldehyde occurred, followed by a dehydration step to form the benzaldehyde oxime. At last, the hydroxyl group in the benzaldehyde oxime was protonated by the \( \text{H}^+ \) offered by the \([\text{HSO}_3\text{-b-Py}]\cdot\text{HSO}_4\) to form the intermediate, which then converted to benzonitrile through a dehydration step, along with the regeneration of \([\text{HSO}_3\text{-b-Py}]\cdot\text{HSO}_4\) to complete the catalytic cycle.

**Table 2** Effect of different hydroxylamine salts on direct synthesis of benzonitrile\(^a\)

| No | Hydroxylamine salt        | \( X_{\text{BD}} \) (%) | \( Y_{\text{BT}} \) (%) |
|----|---------------------------|-------------------------|-------------------------|
| 1  | \((\text{NH}_2\text{OH})_2\cdot[\text{HSO}_3\text{-b-N(C}_3\text{H}_3)_2]\)\cdot\text{HSO}_4\)\(^b\) | 84.3                     | 84.3                     |
| 2  | \((\text{NH}_2\text{OH})_2\cdot[\text{HSO}_3\text{-b-mim}]\)\cdot\text{HSO}_4\)\(^c\) | 99.9                     | 99.9                     |
| 3  | \((\text{NH}_2\text{OH})_2\cdot[\text{HSO}_3\text{-b-Py}]\)\cdot\text{HSO}_4\)\(^d\) | 100                      | 100                      |
| 4  | \(\text{NH}_2\text{OH} \cdot \text{HCl}\)\(^d\) | 99.7                     | 99.7                     |
| 5  | \((\text{NH}_2\text{OH})_2\cdot\text{H}_2\text{SO}_4\)\(^d\) | 94.3                     | 94.3                     |

\(^a\) Reaction conditions: benzaldehyde 3.6 mmol, \( n(\text{BD}) : n(\text{NH}_2\text{OH}) = 1 : 3 \), paraxylene 8 mL, IL 4 mL, temperature 120 °C, time 2 h. \(^b\) IL is \([\text{HSO}_3\text{-b-N(C}_3\text{H}_3)_2]\)\cdot\text{HSO}_4. \(^c\) IL is \([\text{HSO}_3\text{-b-mim}]\)\cdot\text{HSO}_4. \(^d\) IL is \([\text{HSO}_3\text{-b-Py}]\)\cdot\text{HSO}_4.

**Table 3** Calculation and comparison of \( H_0 \) values of different ionic liquids\(^a\)

| No | IL                        | \( A_{\text{max}} \) (%) | \( I \) (%) | \( H_0 \) (%) | \( H_0 \) (%) |
|----|---------------------------|-------------------------|-------------------------|
| 1  | Blank                     | 0.53                    | 100                      | 1.68                   |
| 2  | \([\text{HSO}_3\text{-b-N(C}_3\text{H}_3)_2]\)\cdot\text{HSO}_4\) | 0.44                    | 83.0                     | 17.0                   |
| 3  | \([\text{HSO}_3\text{-b-mim}]\)\cdot\text{HSO}_4\) | 0.38                    | 71.7                     | 28.3                   |
| 4  | \([\text{HSO}_3\text{-b-Py}]\)\cdot\text{HSO}_4\) | 0.37                    | 69.8                     | 30.2                   |

\(^a\) Indicator: 4-nitroaniline (5 mg L\(^-1\), \( pK_a = +0.99 \)), ILs (5 mmol L\(^-1\)).

**Scheme 2** Plausible mechanism for synthesis of benzonitrile from benzaldehyde and \((\text{NH}_2\text{OH})_2\cdot[\text{HSO}_3\text{-b-Py}]\)\cdot\text{HSO}_4.

**Fig. 4** FTIR spectra of the recovered and fresh \([\text{HSO}_3\text{-b-Py}]\)\cdot\text{HSO}_4.

**Fig. 5** FTIR spectra of the \((\text{NH}_2\text{OH})_2\cdot[\text{HSO}_3\text{-b-Py}]\)\cdot\text{HSO}_4 obtained from the recovered and fresh \([\text{HSO}_3\text{-b-Py}]\)\cdot\text{HSO}_4.
The recovery and recycling of the ionic liquid

(1) Recycling of the ionic liquid as co-solvent and catalyst. To investigate the feasibility of the recovery and recycling of the ionic liquid, direct synthesis of benzonitrile from benzaldehyde and \((\text{NH}_2\text{OH})_2\cdot\text{HSO}_3\cdot\text{b-Py}\cdot\text{HSO}_4\) was carried out under the above mentioned optimal reaction conditions. After completion of the reaction, the resulting mixture was cooled and separated into organic phase and aqueous phase automatically. The organic phase was analyzed by gas chromatography, and the

| No | Aldehyde | Nitrile | X (%) | Y (%) |
|----|----------|---------|-------|-------|
| 1  | ![1](image1) | ![1](image2) | 80.6  | 80.6  |
| 2  | ![2](image3) | ![2](image4) | 86.7  | 86.7  |
| 3  | ![3](image5) | ![3](image6) | 94.4  | 94.4  |
| 4  | ![4](image7) | ![4](image8) | 99.7  | 99.7  |
| 5  | ![5](image9) | ![5](image10) | 99.9  | 99.9  |
| 6  | ![6](image11) | ![6](image12) | 77.1  | 77.1  |
| 7  | ![7](image13) | ![7](image14) | 99.8  | 99.8  |
| 8  | ![8](image15) | ![8](image16) | 91.5  | 91.5  |
| 9  | ![9](image17) | ![9](image18) | 100   | 100   |
| 10 | ![10](image19) | ![10](image20) | 100   | 100   |
| 11 | ![11](image21) | ![11](image22) | 100   | 98.7  |
| 12 | ![12](image23) | ![12](image24) | 85.4  | 83.1  |
| 13 | ![13](image25) | ![13](image26) | 41.0  | 38.7  |
| 14 | ![14](image27) | ![14](image28) | 99.4  | 99.4  |
| 15 | ![15](image29) | ![15](image30) | 100   | 98.1  |
| 16 | ![16](image31) | ![16](image32) | 99.9  | 99.9  |

*Reaction conditions: aldehyde 3.6 mmol, \(n(\text{aldehyde}) : n((\text{NH}_2\text{OH})_2\cdot\text{HSO}_3\cdot\text{b-Py}\cdot\text{HSO}_4) = 1 : 1.5,\) paraxylene 8 mL, \([\text{HSO}_3\cdot\text{b-Py}\cdot\text{HSO}_4] 4\) mL, temperature 120 °C, time 2 h.*
leaving ionic liquid in aqueous phase was distilled with a rotary evaporator under vacuum to recover the [HSO₄⁻·b-Py]:HSO₄. The recovered [HSO₄⁻·b-Py]:HSO₄ was further used as co-solvent and catalyst for the next round. It can be seen that the benzaldehyde conversion and benzonitrile yield were still as high as 96.7% after five times of recycling (Table 4), indicating that [HSO₄⁻·b-Py]:HSO₄ was stable enough to be recycled.

After the [HSO₄⁻·b-Py]:HSO₄ was recycled for five times, it was recovered, collected and characterized. The FTIR spectra of the recovered and fresh [HSO₄⁻·b-Py]:HSO₄ were similar (Fig. 4). In addition, the ¹H NMR and ¹³C NMR spectra (Fig. S2–S5) of the recovered [HSO₄⁻·b-Py]:HSO₄ was consistent with that of the fresh [HSO₄⁻·b-Py]:HSO₄. These results demonstrated that the [HSO₄⁻·b-Py]:HSO₄ was structurally stable even after recycling five times.

(2) Recycling of the ionic liquid as hydroxylamine stabilizer. [HSO₄⁻·b-Py]:HSO₄ was recovered when it was recycled for five times, and then reused as the hydroxylamine stabilizer. The (NH₂OH)₂:[HSO₄⁻·b-Py]:HSO₄ was prepared again. The FTIR spectra of (NH₂OH)₂:[HSO₄⁻·b-Py]:HSO₄ obtained from the fresh and recovered [HSO₄⁻·b-Py]:HSO₄ was presented in Fig. 5. It can be seen that they were almost identical. The peaks at 3120 and 2725 cm⁻¹ were assigned to the NH₃⁺ stretching mode. The peaks at 2041 and 1540 cm⁻¹ were attributed to the characteristic peaks of primary ammonium salt and the NH₃⁺ deformation frequencies. Moreover, the ¹H NMR and ¹³C NMR spectra (Fig. S6–S9) of the (NH₂OH)₂:[HSO₄⁻·b-Py]:HSO₄ obtained from the recovered [HSO₄⁻·b-Py]:HSO₄ was consistent with the (NH₂OH)₂:[HSO₄⁻·b-Py]:HSO₄ obtained from the fresh one, indicating that the recovered [HSO₄⁻·b-Py]:HSO₄ could be used again as the hydroxylamine stabilizer even after it was recycled as co-solvent and catalyst for five times.

Next, the (NH₂OH)₂:[HSO₄⁻·b-Py]:HSO₄ obtained from the recovered [HSO₄⁻·b-Py]:HSO₄ was used in direct synthesis of benzonitrile under the optimal reaction conditions. The benzaldehyde conversion and benzonitrile yield were 100% and 98.2%, respectively. The results demonstrated that the reactivity of the (NH₂OH)₂:[HSO₄⁻·b-Py]:HSO₄ obtained from the recovered and fresh [HSO₄⁻·b-Py]:HSO₄ were much the same.

Reactivity of different aldehydes

In subsequent efforts, various aldehydes including structurally diverse aryl aldehydes, heterocyclic aldehydes, aliphatic aldehydes, even dialdehyde were chosen as potential substrates to produce their corresponding nitriles (Table 5). For substituted aromatic aldehydes, the substituents and their positions could affect their reactivity. For example, the conversion of the aldehyde with electron withdrawing group was increased along with the enhancement of the electron attract ability (line 1–3, Table 5). Moreover, the results of the para-substituted aldehydes were better than those of the meta-substituted aldehydes (line 4–5, 6–7, Table 5). For aliphatic aldehydes, the yield of aliphatic nitrile decreased with the increase of the chain length of the aliphatic aldehyde. The conversions of the heterocyclic aldehydes were nearly 100%. As evident from Table 5, stronger electron withdrawing groups (like nitro) have greater influence on the carbonyl group of aromatic aldehydes leading to an increase in its electrophilicity thus increasing the yield of nitrile. This effect is more pronounced for an electron withdrawing substituent in the para position compared to its meta position. Moreover, for aliphatic aldehydes, increasing chain length decreases the electrophilicity of the aldehydic group due to increase in positive inductive effect of the aliphatic chain thus decreasing the overall yield of aliphatic nitriles for higher aliphatic chain aldehydes.

Conclusions

In summary, a convenient and facile method for direct conversion of aldehydes into nitriles using ionic liquid as the recycling agent has been established. The optimum reaction condition was at 120 °C for 2 h. The suitable molar ratio of benzaldehyde to (NH₂OH)₂:[HSO₄⁻·b-Py]:HSO₄ was 1:1.5, and the suitable volume ratio of paraxyleno to [HSO₄⁻·b-Py]:HSO₄ was 2:1. Under the optimal reaction conditions, the benzaldehyde conversion and benzonitrile yield were both 100%. Ionic liquid exhibited the multiple roles of co-solvent, catalyst and phase separation. It could be recovered easily, and recycled directly after reaction. Additionally, this novel route is applicable to green synthesis of a variety of aromatic, heteroaromatic and aliphatic nitriles with excellent yields.

Conflicts of interest

The authors declare no competing financial interest.

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Notes and references

1. B. M. Xu, H. Y. Ruan, Y. Zhou, W. Xiong and K. Chen, Chem. World, 2012, 5, 316–320.
2. L. P. OuYang, X. Q. You, Y. Y. Huang, A. Q. Zhang and G. Y. Xie, Appl. Chem. Ind., 2017, 46, 1692–1694.
3. C. H. Yang and J. M. Williams, Org. Lett., 2004, 6, 2837–2840.
4. T. Schareina, A. Zape and M. Beller, Chem. Commun., 2004, 12, 1388–1389.
5. S. Iida, R. Ohmura and H. Togo, Tetrahedron, 2009, 65, 6257–6262.
6. H. F. Lu, Y. Zhou, H. F. Huang, H. Y. Liu and Y. F. Chen, Chin. J. Catal., 2005, 26, 101–105.
7. E. Rombi, I. Ferino, R. Monaci, C. Picciau, V. Solinas and R. Buzzoni, Appl. Catal., A, 2004, 266, 73–79.
8. K. V. R. Chary, K. R. Reddy, T. Bhaskar and V. S. Guggilla, Green Chem., 2002, 4, 206–209.
9. H. Chen, C. Y. Lu and X. H. Liu, J. Huaihai Inst. Technol., Nat. Sci. Ed., 1997, 6, 51–53.
10 K. Murugesan, T. Senthamarai, M. Sohail, M. Sharif, N. V. Kalevarua and R. V. Jagadeesh, Green Chem., 2018, 20, 266–273.
11 B. Mitra, G. C. Pariyar, R. Singha and P. Ghosh, Tetrahedron Lett., 2017, 58, 2298–2301.
12 C. J. Fang, M. C. Li, X. Q. Hu, W. M. Mo, B. X. Hu, N. Sun, L. Q. Jin and Z. L. Shen, RSC Adv., 2017, 7, 1484–1489.
13 V. P. Gozum and R. C. Mebane, Green Chem. Lett. Rev., 2013, 6, 149–150.
14 P. Ghosh, B. Saha, G. C. Pariyar, A. Tamang and R. Subba, Tetrahedron Lett., 2016, 57, 3618–3621.
15 S. L. Guan and W. B. Huang, Chem. World, 1993, 9, 418–423.
16 J. M. DeMott Jr and C. J. Kelley, J. Chem. Educ., 2001, 78, 780.
17 S. C. Xu, Organic Chemistry, Higher Education Press, Beijing, 2nd edn, 1993, p.284.
18 S.-I. Fukuzawa, Y. Yamaishi, H. Furuya, K. Terao and F. Iwasaki, Tetrahedron Lett., 1997, 38, 7203–7206.
19 Z. Wang and J. Luo, Fine Chem. Intermed., 2013, 43, 14–18.
20 S. Enthaler, M. Weidauer and F. Schröder, Tetrahedron Lett., 2012, 53, 882–885.
21 Y. J. Zhuang, J. Liu and Y. B. Kang, Tetrahedron Lett., 2016, 57, 5700–5702.
22 A. Leggio, E. L. Belsito, S. Gallo and A. Liguori, Tetrahedron Lett., 2017, 58, 1512–1514.
23 A. Khalafi-Nezhad and S. Mohammadi, RSC Adv., 2014, 4, 13782–13787.
24 S. H. Khezri, N. Azimi, M. Mohammed-Vali, B. Eftekhar-Sis, M. M. Hashemi, M. H. Baniasadi and F. Teimouri, ARKIVOC, 2007, 15, 162–170.
25 L. V.-B. Hoelz, B. T. Gonçalves, J. C. Barros and J. F. Mendes da Silva, Molecules, 2010, 15, 94–99.
26 Z. H. Li, Q. S. Yang, X. D. Qi, Y. Y. Xu, D. S. Zhang, Y. J. Wang and X. Q. Zhao, Chem. Commun., 2015, 51, 1930–1932.
27 Z. H. Li, Q. S. Yang, L. Y. Gao, Y. Y. Xu, D. S. Zhang, S. F. Wang, X. Q. Zhao and Y. J. Wang, RSC Adv., 2016, 6, 83619–83625.
28 Z. H. Li, X. D. Qi, L. Y. Gao, Y. Y. Xu, D. S. Zhang, S. F. Wang, X. Q. Zhao and Y. J. Wang, Chem. Lett., 2017, 46, 289–292.
29 H. J. Zang, M. L. Wang, B. W. Cheng and J. Song, Ultrasound Sonochim., 2009, 16, 301–303.
30 A. Elmakssoudi, K. Abdelouahdi, M. Zahouily, J. Clark and A. Solhy, Catal. Commun., 2012, 29, 53–57.
31 D. Lorenzo, A. Romero and A. Santos, Ind. Eng. Chem. Res., 2016, 55, 6586–6594.
32 M. Matsumoto, M. Tanimura, T. Akimoto, N. Watanabe and H. K. Ijuin, Tetrahedron Lett., 2008, 49, 4170–4173.
33 O. Phillips, J. M. Schwartz and P. A. Kohl, Polym. Degrad. Stab., 2016, 125, 129–139.
34 J. Wang, X. B. Zhang, F. M. Wang and W. F. Cai, Chem. Ind. Eng. Prog., 2016, 35, 1309–1313.
35 D. Saha, A. Saha and B. C. Ranu, Tetrahedron Lett., 2009, 50, 6088–6091.
36 D. Abolghasem, K. Amir, B. Mehdi and T.-H. Niloofar, Chin. J. Chem., 2011, 29, 978–982.
37 J. L. Zhang, X. R. Chen, Y. P. Yu, K. L. Xu and J. Huang, Chem. Res. Chin. Univ., 2011, 27, 712–714.
38 X. L. Zhang, H. L. An, H. Q. Zhang, X. Q. Zhao and Y. J. Wang, Ind. Eng. Chem. Res., 2014, 53, 16707–16714.