Biosynthesis of zinc oxide nanoparticle: a review on greener approach

Abstract

Past few decades application of nanoparticle in the different fields like biomedical, cosmetic, food and many more has gained much importance globally because of its size and unique properties, hence it becomes an important field among the researchers. Recently in nanotechnology research, synthesis of nanoparticle from green chemistry pathways has been preferred due to its natural biological reduction property which reduces the utilization and exposure of toxic chemical to the environment when compare to physical and chemical methods. Among the different inorganic nanoparticles, zinc oxide nanoparticle attracts more attention due to its, large bandwidth, high excitation binding energy, simplicity, easy fabrication, biosafe, non-toxic, biocompatible, eco and environmental friendly. Zinc oxide nanoparticle is readily soluble in biological fluids and tends to aggregate easily under different physiological condition. But physicochemical properties of the nanoparticle have an impact in the bioavailability. This review was to make an attempt to summarize and conclude green synthesis of zinc oxide nanoparticle and its pharmacokinetic properties.

Keywords: nanoparticle, zinc oxide, non-toxic, bioavailability, pharmacokinetic properties

Introduction

In the last few decades the revolution of nanotechnology exhibit significant and commercial applications in different fields like engineering, environmental science, agriculture, medicine, electrical and electronics, pharmaceutical, biological, biotechnology, mineral nutrition, food processing, reproduction, pharmacology applications and many more. The significant and increased commercial applications are owing to its nanoscale dimension, large surface area to volume ratio, unique and potential properties like electronic, magnetic, mechanical, non-linear optical performance, thermal conductivity, enhanced catalytic reactivity and chemical steadiness. Real research and development in nanotechnology is concerned with making modification in its atomic, molecular and supra-molecular levels to improve its physicochemical properties to produce desired effect.

Inorganic nanoparticles are made by combining organic materials with inorganic materials which forms hybrids and has unique properties. Different inorganic nanoparticles like quantum dots, magnetic, silver, gold, SiO₂, TiO₂, CuO and ZnO nanoparticles have been synthesized and utilized in various fields. Zinc oxide nanoparticle has large bandwidth and high excitation binding energy, made researchers to explore its potential application in different fields. Among the different inorganic nanoparticles, ZnO nanoparticle is the third highest nanoparticle produced globally and utilized in different fields.

Synthesis of nanoparticle

Processes for the preparation of inorganic nanoparticle by different physical and chemical method are:

- Vapor transport and condensation
- Amorphous crystallization
- Physical fragmentation
- Pulsed laser deposition

- Molecular beam epitaxy
- Thermal evaporation and decomposition
- Chemical micro emulsion
- Wet chemical
- Spray pyrolysis
- Electrode-position
- Direct precipitation
- Homogeneous precipitation
- Microwave assisted combustion
- Solvothermal method
- Sonochemical method
- Reverse micelles
- Sol gel method
- Microwave irradiation
- Reaction of zinc with alcohol
- Hydrothermal synthesis
- Micro emulsion synthesis
- Spray drying

However, the major drawback of the physical method is the involvement of high pressure and temperature whereas in the chemical method is the utilization of toxic chemicals, also capping and stabilizing agents are additionally needed.

Green synthesis

Hence, researchers focus towards the green synthesis of nanoparticle from microorganisms like bacteria, algae, fungi, yeast and different plant parts like leaves, fruits, seeds, roots and stems for various applications. Green synthesis of nanoparticle from plants are more stable, varied shape and size when compare to nanoparticle synthesized from other organisms. Green synthesis approach shows more catalytic activity and limits the exposure, utilization of toxic and
expensive chemicals; which can help to protect the environment from chemical toxicity and hazardous.

**Synthesis of zinc oxide nanoparticle**

Desired quantity zinc oxide or zinc nitrate or zinc acetate or zinc sulfate is dissolved in distilled water. To the zinc solution, defined volume of plant extract is added and mixed well in a magnetic stirrer. To adjust the pH of the above mixture to pH12, NaOH solution is added drop by drop with continuous stirring and kept for 1 hr as incubation time. The color of the reaction mixture is changed after 1h of incubation time. Formation of nanoparticle is confirmed in UV spectroscopy based on its surface plasmon resonance (SPR) effect. Formed white crystalline precipitate of Zinc oxide nanoparticle is separated from the reaction solution by centrifugation and dried in a hot air oven operating at 60-80°C for 2hr.\(^1\)

The chemical reaction process for the formation of zinc oxide nanoparticle from zinc acetate is shown in the following chemical equation\(^8\):

\[
\begin{align*}
Zn(CH_3COO)_2 + 2H_2O & \rightarrow Zn(OH)_2 + 2CH_3COONa + 2H_2O \quad [1] \\
Zn(OH)_2 + 2H_2O & \rightarrow Zn(OH)_2^2 + 2H + 2H_2O \quad [2] \\
Zn(OH)_2^2 & \rightarrow ZnO + 2OH + 2H_2O \quad [3]
\end{align*}
\]

Plants secrete phytoconstituents like amino acids, vitamins, terpenoids, alkaloids, polyphenol and polysaccharides which can acts as reducing agent as well as capping or stabilizing agent. Also these phytoconstituents reduce metal ions or metal oxides into zero valence metal nanoparticles. Hence, no capping or stabilizing agents are needed in this green approach of nanoparticle synthesis.\(^6,16\) This green synthesis approach is easy fabrication, cost effective, non-toxic, biocompatible, environment friendly and safe. Also US Food and Drug Administration (US FDA) enlisted ZnO with other four zinc compounds are generally recognized as safe (GRAS) material.\(^17\) In general for the ZnO nanoparticle synthesis zinc oxide or zinc nitrate or zinc acetate or zinc sulfate are used. Different factors like temperature, pH, time and concentration of extract have important role in obtaining desired size and shape of the nanoparticle.\(^18\) Formation ZnO Nanoparticle is confirmed in UV spectroscopy based on its surface plasmon resonance (SPR) effect.\(^4\) Different morphologies of ZnO Nanoparticles are nanoflower, nanoflake, nanorods, nanowire and nanobelt.\(^8\) Synthesis of zinc oxide nanoparticle from different natural plant source\(^9\) is shown in Figure 1 and the steps for the preparation of zinc oxide nanoparticle,\(^9\) are shown in Figure 2.

**Pharmacokinetic properties**

**Absorption**

Zinc oxide nanoparticles are less stable and readily soluble in biological fluids and subsequently release zinc ions.\(^21,22\) On oral administration first pass effect, gastrointestinal barrier, liver and gut wall function decreases the bioavailability of ZnO nanoparticles; whereas on intravenous administration shows 100% bioavailability.\(^13\) Different factors like pH, particle size, concentration and presence or absence of organic compounds influence the solubility of ZnO nanoparticles. However negative charge particles are absorbed more when compare to the positive charge particles.\(^29\) Dissolution, absorption and distribution of ZnO nanoparticles are depend upon the exposure dose amount and does not depend upon the particle size of the nanoparticle.\(^23\)

**Distribution**

On oral administration, zinc oxide nanoparticles are distributed in kidney, liver and spleen whereas on intraperitoneal administration zinc oxide nanoparticles are distributed in lungs, kidney, spleen, heart and liver.\(^23,26\) Kidneys and liver is the common target organ on both administrations. Distribution of zinc in the organs is depends upon the type of experimental animals, route of exposure and physicochemical properties of the nanoparticles.

**Excretion**

Metabolites and wastes from ZnO nanoparticles have less than 6 nm in hydrodynamic diameter are eliminated through urine by glomerular filtration,\(^27,28\) both absorbed and nonabsorbed entities are excreted through fecal, whereas metabolites are excreted via bile. Exposure route and physicochemical properties has an important role in the excretion of the ZnO nanoparticles; but biliary and fecal excretion are having major role in the elimination of the nanoparticle irrespective of its particle size, surface charge and exposure routes. Excretion kinetic of the nanoparticle is depends upon the particle size rather than the surface charge of the particle.\(^24,25\)
Protein interaction

When nanoparticles contact with biological fluid in systemic circulation, rapid interaction occurs leads to adsorption of protein on the nanoparticle surface. Adsorption of the protein on the nanoparticle surface is highly dependent on the surface charge and particle size of the nanoparticle. Plasma protein like fibrinogen, lipoprotein, albumin and immunoglobulins are utilized to assess the binding, adsorption and subsequent protein structure changes due to nanoparticle-protein interaction. Also plasma protein plays a crucial role in disposition, disposition and transportation.20–21 (Figure 3).

Conflict of interest

The author declares no conflict of interest.

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