Research Article

Synthesis, Characterization, and In Vitro Cytotoxic Activities of Benzaldehyde Thiosemicarbazone Derivatives and Their Palladium (II) and Platinum (II) Complexes against Various Human Tumor Cell Lines

Wilfredo Hernández,1 Juan Paz,1 Abraham Vaisberg,2 Evgenia Spodine,3 Rainer Richter,4 and Lothar Beyer4

1 Facultad de Ingeniería Industrial, Universidad de Lima, Avenue Javier Prado Este Cda 46, Monterrico-Santiago de Surco, Lima 33, Peru
2 Facultad de Ciencias y Filosofía, Universidad Peruana Cayetano Heredia, Avenue Honorio Delgado 430, Urb. Ingeniería-San Martin de Porras, Lima 31, Peru
3 Facultad de Ciencias Químicas y Farmacéuticas and CIMA T, Universidad de Chile, Santiago 8380000, Chile
4 Fakultät für Chemie und Mineralogie, Universität Leipzig, Johannisallee 29, 04103 Leipzig, Germany

Correspondence should be addressed to Wilfredo Hernández, whernandez79@yahoo.es

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The palladium (II) bis-chelate Pd (L1−3)2 and platinum (II) tetranuclear Pt4(L4)4 complexes of benzaldehyde thiosemicarbazone derivatives have been synthesized, and characterized by elemental analysis and IR, FAB(+)-mass and NMR (1H,13C) spectroscopy. The complex Pd(L2)2 [HL2 = m-CN-benzaldehyde thiosemicarbazone] shows a square-planar geometry with two deprotonated ligands (L) coordinated to PdII through the nitrogen and sulphur atoms in a trans arrangement, while the complex Pt4(L4)4 [HL4 = 4-phenyl-1-benzaldehyde thiosemicarbazone] has a tetranuclear geometry with four tridentate ligands coordinated to four PtII ions through the carbon (aromatic ring), nitrogen, and sulphur atoms where the ligands are deprotonated at the NH group. The in vitro antitumor activity of the ligands and their complexes was determined against different human tumor cell lines, which revealed that the palladium (II) and platinum (II) complexes are more cytotoxic than their ligands with IC50 values at the range of 0.07–3.67 μM. The tetranuclear complex Pt4(L4)4, with the phenyl group in the terminal amine of the ligand, showed higher antiproliferative activity (CI50 = 0.07–0.12 μM) than the other tested palladium (II) complexes.

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1. INTRODUCTION

The synthesis of transition metal complexes with thiosemicarbazone ligands has been receiving considerable attention due to the pharmacological properties of both ligands and complexes [1–3]. Thiosemicarbazone derivatives exhibit a great variety of biological activities, such as antitumor [4], antifungal [5, 6], antibacterial [6, 7], and antiviral [8] properties.

The deprotonated thiosemicarbazone ligands usually coordinate to platinum, palladium, copper, ruthenium, and osmium through oxygen, nitrogen, and sulphur donor atoms in their (N, S) bidentate form or (N, N, S or O, N, S) tridentate form, to form metallic complexes of different molecular geometry [9–11].

The square planar platinum (II) and palladium (II) complexes of M(HL)Cl2 and M(L)Cl type with thiosemicarbazone ligands derived from phenylacetaldehyde and 2-formylpyridine showed high cytotoxicity in vitro against HL60 leukemia and P388 mouse leukemia cell lines [12], while platinum(II) and palladium(II) binuclear complexes with p-isopropylbenzaldehyde thiosemicarbazone ligands exhibit strong cytotoxic activities on mouse tumor cell growth inhibition [4, 13].

On the other hand, C,N,S thiosemicarbazone ligands can also be coordinated to palladium (II) to form complexes with
two fused five-membered chelate rings containing a carbon-metal σ bond [14]. Although there is little information about the antitumor activity of these tetranuclear complexes, their pharmacological applications could be relevant.

As part of our continuing investigations about metal complexes with ligands derived from thiourea [15–19] such as thiosemicarbazones, we report here the synthesis, characterization, and antitumor activity of palladium(II) bis-chelates Pd(L1–3)2 and platinum (II) tetranuclear complex Pt2(L4)4 with benzaldehyde thiosemicarbazone and 4-phenyl-1-benzaldehyde thiosemicarbazone ligands, R=PhCH=N-NH-C(=S)-NHR1, HL1(R, R1=H), HL2(R=m-CN, R1=H), HL3(R=o-NO2, R1=H), and HL4(R=H, R1=Ph).

2. EXPERIMENTAL

2.1. Materials and measurements

Chemicals were reagent grade. Palladium (II) bis(acetylacetonate), ammonium tetrachloroplatinate, thiosemicarbazide, 4-phenyl-thiosemicarbazide, benzaldehyde, m-CN-benzaldehyde, and o-NO2-benzaldehyde were purchased from Aldrich. Melting points were determined on a Büchi melting point B-545 apparatus. Elemental analyses were determined on a Fisons-Carlo Erba Elemental Microanalyzer. The infrared (IR) spectra were recorded in solid state (KBr pellets) on a Bruker FT-IR IFS 55 Equinox spectrophotometer in the 4000–400 cm⁻¹ range. The FAB(+) mass spectra were recorded on a ZAB-HSQ (V.G. Analytical Ltd. Floats Roads, Wythenshawe, Manchester, England) spectrometer, using 3-nitrobenzyl alcohol as the matrix. The 1H (300 MHz) and 13C (75.5 MHz) NMR spectra were recorded on a Bruker Advance DRX 300 spectrometer at 300 K, using DMSO-d6 as solvent. The chemical shifts (δ) in ppm were measured relative to tetramethylsilane (TMS).

2.2. Synthesis of the ligands

The thiosemicarbazone derivatives (HL) were prepared according to the literature [20] as shown in Scheme 1.

**General method**

To a hot solution of thiosemicarbazide (1.82 g, 20 mmol) or 4-phenylthiosemicarbazide (3.3 g, 20 mmol) in 160 mL, methanol was added dropwise a solution of the corresponding benzaldehyde (20 mmol) in 70 mL methanol during 30 minutes. The mixture was stirred and refluxed for 4 hours, it was filtered and the filtrate was concentrated to half the volume under reduced pressure. After a slow evaporation of the concentrate at room temperature, crystals were collected by filtration, washed with cold ethanol, and dried in vacuo. For ligand HL4, the filtrate was kept in the refrigerator and after several hours small rectangular crystals were obtained. These crystals were suitable for structure analysis by X-ray diffraction.

**2.2.1. Benzaldehyde thiosemicarbazone (HL1)**

Yellow crystals. Yield 80%, m.p. 167–169°C. Anal. Calc. For C9H9N3S (224.3 g/mol): C, 53.6%; H, 3.1%; N, 24.9%; S, 17.9%. Found: C, 53.5%; H, 3.5%; N, 23.5%; S, 17.7%. FAB(+)-MS: m/z 199 (M+, 100%); IR (KBr, cm⁻¹): ν(NH) 3250; ν(C=N) 1600; ν(=S) 885. 1H-NMR (DMSO-d6): δ 7.85 (d, 2H ortho, Ph, J=6.8 Hz), 7.65 (t, 2H meta, Ph, J= 7.2 Hz), 7.40 (t, 1H para, Ph, J= 7.2 Hz); 8.05 (s, 1H, HC=N); 8.19, 7.98 (d, 2H, NH2); 11.42 (s, 1H, =N–NH). 13C-NMR (DMSO-d6): δ 128.5, 127.2, 129.8, 118.6; 134.1 (Ph); 142.2 (HC= N); 178.3 (C=O).

**2.2.2. m-cyanobenzaldehyde thiosemicarbazone (HL2)**

Colorless crystals. Yield 76%, m.p. 203–204°C. Anal. Calc. For C9H9N3S (204.3 g/mol): C, 52.9%; H, 3.9%; N, 27.4%; S, 15.6%. Found: C, 52.8%; H, 3.7%; N, 27.6%; S, 15.5%; FAB(+)-MS: m/z 205 (M+, 100%); IR (KBr, cm⁻¹): ν(NH2) 3410, 3397; ν(NH) 3250; ν(C=N) 1600; ν(=S) 885. 1H-NMR (DMSO-d6): δ 7.81, 7.79 (s, 2H ortho, Ph, J=5.9 Hz); 7.58 (t, 1H meta, Ph, J= 7.7 Hz); 7.83 (t, 1H para, Ph, J= 5.7 Hz); 8.03 (s, 1H, HC=N); 8.31, 8.26 (d, 2H, NH2); 11.60 (s, 1H, =N–NH). 13C-NMR (DMSO-d6): δ 129.87, 118.6, 135.69, 132.32, 132.68 (Ph); 111.97 (CN); 139.66 (HC=N); 178.35 (C=O).

**2.2.3. o-nitrobenzaldehyde thiosemicarbazone (HL3)**

Colorless crystals. Yield 90%, m.p. 214–215°C. Anal. Calc. For C9H8N4S (204.3 g/mol): C, 52.9%; H, 3.9%; N, 27.4%; S, 14.3%. Found: C, 42.6%; H, 3.5%; N, 24.6%; S, 14.1%. FAB(+)-MS: m/z 226 (M+, 100%); IR (KBr, cm⁻¹): ν(NH2)
Scheme 2: Synthesis of the palladium (II) bis-chelate complexes and the platinum (II) tetracuclear complex.

2.2.4. 4-phenyl-1-benzaldehyde thiosemicarbazone (HL₄)

Yellow crystals. Yield 75%, m.p. 192–194°C. Anal. Calc. For C₁₄H₁₃N₃S (255.3 g/mol): C, 65.9%; H, 5.1%; N, 16.5%; S, 12.5%. Found: C, 65.4%; H, 5.3%; N, 16.7%; S, 12.6%. FAB(+)–MS: m/z 255 (M⁺, 48%); IR (KBr, cm⁻¹): ν(NH) 3245; ν(C=N) 1625; ν(C=S) 915. ¹H-NMR (DMSO-d₆): δ 7.42 (t, 1H para, Ph, J = 6.8 Hz); 7.43 (t, 2H meta, Ph, J = 5.5 Hz); 7.34 (d, 2H ortho, Ph, J = 6.3 Hz); 7.21 (t, 2H para, NHPh, J = 6.5 Hz); 7.07 (d, 2H ortho, NHPh, J = 5.7 Hz); 8.17 (s, 1H, HC=N); 10.11 (s, 1H, =N–NH). ¹³C-NMR (DMSO-d₆): δ 127.6, 128.6, 130.0, 134.0 (Ph); 125.3, 125.9, 128.0, 139.1 (NH-Ph); 142.9 (HC=N); 176.0 (C=S).

2.3. Synthesis of the palladium (II) and platinum (II) complexes (see Scheme 2)

A solution of Pd(acac)₂ (0.30 g, 1.0 mmol) in CH₂Cl₂/CH₃OH (30 mL, 2:1 v/v) or a solution of (NH₄)₂PtCl₄ (0.1865 g, 0.5 mmol) in water/ethanol (2:1, 15 mL) was added dropwise to a stirred solution of the corresponding thiosemicarbazone (2.0 mmol) in 60 mL of methanol. Sodium acetate (0.16 g, 2 mmol) in 3 mL of water was then added. The solution was refluxed for 2 hours and stirred for 24 hours at room temperature. The precipitate was collected by filtration and dried in vacuo.

2.3.1. Palladium (II) complex of benzaldehyde thiosemicarbazone, Pd(L₁)₂

Yellow solid. Yield 70%, m.p. 204-205°C. Anal. Calc. For C₁₆H₁₆N₆S₂Pd (462.9 g/mol): C, 41.5%; H, 3.5%; N, 18.2%; S, 13.9%. Found: C, 40.9%; H, 3.6%; N, 18.6%; S, 13.5%. FAB(+)–MS: m/z 463 (M⁺, 60%); IR (KBr, cm⁻¹): ν(NH₂) 3390, 3367; ν(C=N) 1582; ν(C=S) 805. ¹H-NMR (DMSO-d₆): δ 7.49, 7.45, 7.41, 7.39 (m, Ph); 8.13 (s, 2H, HC=N); 8.29, 8.21 (d, 4H, NH₂).
2.3.2. Palladium (II) complex of m-cyanobenzaldehyde thiosemicarbazone, Pd(L\textsuperscript{1})\textsubscript{2}

Crystals suitable for X-ray structure determination were obtained by slowly evaporating a methanol/dichloromethane (2:1) solution at room temperature.

*Orange crystals.* Yield 63%, m.p. > 240°C (decomp.). Anal. Calc. for C\textsubscript{18}H\textsubscript{14}N\textsubscript{2}S\textsubscript{2}Pd·H\textsubscript{2}O (530.9 g/mol): C, 40.7%; H, 3.0%; N, 21.1%; Found: C, 42.0%; H, 2.7%; N, 21.6%; S, 12.3%. FAB(+)−MS: m/z 513 (M+-H\textsubscript{2}O, 100%); IR (KBr, cm\textsuperscript{-1}): ν(NH\textsubscript{2}) 3405, 3377; ν(CN) 2230; ν(C=N) 1570; ν(C=S) 815. \textsuperscript{1}H-NMR (DMSO-d\textsubscript{6}): δ 7.75, 7.65, 7.55 (m, Ph); 8.06 (s, 2H, HC=N); 8.68 (s, 2H, NH\textsubscript{2}).

2.3.3. Palladium (II) complex of o-nitrobenzaldehyde thiosemicarbazone, Pd(L\textsubscript{2})\textsubscript{2}

*Yellow solid.* Yield 61%, m.p. > 260°C (decomp.). Anal. Calc. for C\textsubscript{28}H\textsubscript{26}N\textsubscript{2}S\textsubscript{2}Pd·2C\textsubscript{2}H\textsubscript{5}OH (1885.8 g/mol): C, 38.2%; H, 3.0%; N, 8.9%; S, 6.8%. Found: C, 38.0%; H, 3.0%; N, 8.6%; S, 6.5%; IR (KBr, cm\textsuperscript{-1}): ν(NH\textsubscript{2}) 3200; ν(C=N) 1590; ν(C=S) 840. \textsuperscript{1}H-NMR (DMSO-d\textsubscript{6}): δ 7.50–7.80 (m, Ph); 7.0–7.4 (m, NH\textsubscript{2}); 8.0 (d, 4H, HC=N); 9.15, 9.68 (d, 4H, NH\textsubscript{2}).

2.3.4. Platinum (II) tetranuclear complex, Pt\textsubscript{4}(L\textsubscript{4})\textsubscript{4}

Crystals suitable for structure determination by X-ray methods, which revealed the positions of all nonhydrogen atoms and refined in riding mode. The hydrogen atoms for HL\textsubscript{4} and the palladium (II) complex Pd(L\textsubscript{2})\textsubscript{2}, and the platinum (II) tetranuclear complex Pt\textsubscript{4}(L\textsubscript{4})\textsubscript{4} are summarized in Table 1.

![Cartoon of a molecule](image)

2.5. Biological activity

2.5.1. Cell culture

The antitumor assays were performed employing the following cell lines: H460 (human lung large cell carcinoma), ME180 (human cervix epidermoid carcinoma), M-14 (human amelanotic melanoma), DU145 (human prostate carcinoma), MCF-7 (human breast adenocarcinoma), HT-29 (human colon adenocarcinoma), PC3 (human prostate carcinoma), and K562 (human chronic myelogenous leukemia). Cells were maintained in Dulbecco's Modified Eagle's Medium (DMEM) supplemented with 10% fetal calf serum and 50 μg/mL gentamycin, and grown at 37°C in a 5% CO\textsubscript{2} humidified environment.

2.5.2. Assessment of cytotoxicity

Cells were inoculated into 96-well tissue culture plates at a density of 3000–5000 cells per well and incubated at 37°C with their corresponding growth medium for 24 hours to allow cells to attach. A plate containing each of these cells was fixed in situ with trichloroacetic acid (TCA) in order to obtain the cell values at zero time before adding the test compounds. The rest of the plates containing the different cell lines received serial dilutions of the ligands and palladium (II) complexes in DMSO to be incubated at 37°C for 48 hours. The assay was terminated by the addition of cold TCA. The cell numbers in each well was determined using the sulforhodamine B (SRB) assay [23]. TCA-treated plates were incubated at 4°C for 1 hour and then the cells were washed five times with tap water and dried completely at room temperature. The cells were stained for 20 minutes with a solution of 0.4% sulforhodamine B in 1% acetic acid. At the end of the staining period, unbound dye was removed by washing four times with 1% acetic acid until the washing solution became colorless. After complete drying, bound dye was solubilized with 10 mM Tris buffer (pH 10.5) and the absorbance reads on an automated plate reader at a wavelength of 550 nm. The IC\textsubscript{50} value was defined as the concentration of test sample resulting in a 50% reduction of absorbance as compared with untreated controls that received a serial dilution of the solvent in which the test samples were dissolved, and was determined by linear regression analysis. For K562 cells, which grow in suspension, instead of fixing and staining with SRB, cells were counted using a Coulter counter.

3. RESULTS AND DISCUSSION

3.1. IR spectra of the ligands and their complexes

The infrared absorption bands become very useful for determining the mode of coordination of the ligands to metal. In the IR spectra, the broad bands of the −NH group observed at 3236–3250 cm\textsuperscript{-1} for the ligands disappear in
Table 1: Crystal data and refinement summary.

|                         | HL₄ | Pd(L²)₂·H₂O | Pt₃(L²)₄·2C₂H₅OH |
|-------------------------|-----|-------------|------------------|
| Empirical formula       | C₁₄H₁₃N₃S | C₁₉H₁₉N₆OPdS₂ | C₆₀H₅₆N₁₂O₂Pt₄S₄ |
| Formula weight (g/mol)  | 255.33 | 530.91      | 1885.77          |
| Crystal habit, color    | Yellow plates | Orange prisms | Red prisms       |
| Crystal system          | Triclinic | Monoclinic  | Monoclinic       |
| Space group             | P-1  | C2/c        | P2₁              |
| a (Å)                   | 5.988(1) | 20.016(4)   | 12.708(1)        |
| b (Å)                   | 10.285(2) | 6.421(1)    | 13.639(1)        |
| c (Å)                   | 11.410(2) | 18.511(4)   | 17.031(1)        |
| α (°)                   | 68.040(2) |             |                  |
| β (°)                   | 82.514(3) |             |                  |
| γ (°)                   | 86.886(2) |             |                  |
| Volume (Å³)             | 0.44 × 0.20 × 0.04 | 0.43 × 0.39 × 0.07 | 0.16 × 0.16 × 0.05 |
| Z; F(000)               | 2; 268 | 4; 1064     | 2; 1784          |
| Density, g/cm³          | 1.313 | 1.725       | 2.127            |
| Crystal size (mm)       | 0.44 × 0.20 × 0.04 | 0.43 × 0.39 × 0.07 | 0.16 × 0.16 × 0.05 |
| μ(MoKα)(mm⁻¹)           | 0.235 | 1.141       | 9.669            |
| 2θ range (°)            | 3.8–50.0 | 6.8–56.2   | 3.8–59.0         |
| Temperature (K)         | 220   | 213         | 213              |
| Measured reflections    | 3388  | 9478        | 19123            |
| Rint                   | 0.0290 | 0.0635      | 0.0224           |
| Unique reflections      | 2252  | 2462        | 13693            |
| Observed reflections (I > 2σ(I)) | 1939 | 2012        | 12253            |
| R₁ (observed reflections) | 0.0618 | 0.0331      | 0.0338           |
| Largest difference peak and hole (e/Å³) | 0.38/−0.31 | 1.26/−0.88 | 1.71/−1.14 |

the complexes spectra, which indicates the deprotonation of the NH-CS group. The strong bands observed at 1596–1625 cm⁻¹ range in the free ligands have been assigned to ν(C=N) stretching vibrations [24]. On complexation, these bands were observed to be shifted to lower frequencies (1570–1590 cm⁻¹), which are in agreement with the wave numbers for other bischelate complexes [6, 25, 26]. These results indicate that the imine nitrogen is coordinated to the metal ion. All ligands showed medium bands in the 880–915 cm⁻¹ range ascribed to ν(C=S) vibrations. These absorption bands shift 65–80 cm⁻¹ to lower frequencies on the coordination of the thiocarbonyl sulfur to palladium (II) or platinum (II) ion. These results are in agreement with other thiosemicarbazone complexes [24, 27]. In addition, the vibrational frequencies of the –NH₂ groups remain unchanged for both the ligands and the complexes. This evidence indicates the noncoordination of the –NH₂ group to the Pd(II) center.

3.2. NMR spectra of the ligands and their complexes

In the ¹H-NMR spectra of the ligands, the signals of the =N–NH protons were observed as singlets at δ 8.03–8.17 in the complexes spectra, which indicates the deprotonation of the =N–NH group. The signals of the HC=N protons which appear as singlets at δ 8.03–8.17 in the ligands show a shift to downfield in δ 0.03–0.80 after complexation. This shift indicates the coordination of the imine nitrogen to the metal center [28]. The signals of the aromatic protons of the ligands appeared at δ 7.21–7.91, and the resonance lines found correspond to the calculated multiplicity. These signals do not suffer relevant changes in the chemical shifts for the palladium (II) and platinum (II) complexes. The NH₂ signal in the ligands HL₁, HL₂, and HL₃ appears as doublets at δ 7.98–8.45 due to the nonequivalence of the amine protons. This evidence is attributed to the restricted rotation around C–N bond (thiocarbonyl carbon and terminal amine nitrogen) due to its partial double bond character [14, 29]. The presence of the phenyl group on the terminal amine (NHPh) of the ligand HL₄ produces a downfield chemical shift at δ 2.1 with respect to the NH₂ group of the ligand HL₁. This reveals that HL₄ is slightly less basic than HL¹. The resonance signals of the –NH₂ or NHPh groups in the palladium (II) and platinum (II) complexes do not change, and this evidence indicates that the amine groups are not coordinated to the metal ion [14].

In the ¹³C-NMR spectra, the carbon resonance signals of the HC=N group appeared at δ 139.66–148.3. The
results are similar to the chemical shifts found for the ligands benzophenone thiosemicarbazide and phenylpropanal thiosemicarbazone (both found at $\delta$ 141) \[30, 31\]. The C=S signals observed at $\delta$ 178.5–176.0 are characteristic for this group, while the aromatic carbons were observed at $\delta$ 139.1–124.5.

### 3.3. Structural data

The molecular structures of HL$^4$, Pd(L$^2$)$_2$, and Pt$_4$(L$^4$)$_4$ are shown in Figures 1, 2, and 3, respectively, whereas their selected bond lengths and bond angles are presented in Tables 2 and 3.

#### 3.3.1. 4-phenyl-1-benzaldehyde thiosemicarbazone HL$^4$

The reaction product of benzaldehyde and 4-phenyl thiosemicarbazide shows the expected bond lengths, especially the N1–C8 double bond with a length of 1.284(4) Å (see Figure 1). The molecular fragment N3–C1(S1)–N2–N1–C8 is nearly planar. The C9–C14 phenyl ring deviates only slightly from this mean plane and forms an angle of 56.4(1)$^\circ$ with the C2–C7 phenyl ring.
There are two hydrogen bonds: an intramolecular N3–H···N1 [N3–H 0.85 Å, H···N1 2.17 Å, N3···N1 2.615 Å, N3–H···N1 112°] hydrogen bond and an intermolecular N2–H···S1 [N2–H 0.88 Å, H···S1 2.62 Å, N2···S1 3.466 Å, N2–H···S1 162°] hydrogen bond. The latter leads to the formation of pairs of molecules in the crystal structure.

### 3.3.2. Bis(3-cyanophenyl-1-benzaldehyde thiosemicarbazonato)palladium (II) Pd(L²)₂

3-cyanophenyl-1-benzaldehyde thiosemicarbazone reacts with palladium (II) acetylacetonate to form a bis-chelate with C₅ molecular symmetry (see Figure 2). The deprotonated ligand coordinates bidentately through S and N. The coordination of the Pd atom is square planar with a ligand coordinates bidentately through S and N. The coordination of the Pd atom is square planar with a ligand coordinates bidentately through S and N.

In the crystal structure, one molecule of water per formula unit of the chelate is included.

There are three hydrogen bonds comprising the atom O1 of the water molecule and the atom N4 of the cyano group: O1–H···N4 2.07 Å, N3···N4 2.898 Å, N3–H···N4 162° and N3–H’···N4 2.999 Å, N3–H’···N4 173°. The formation of tetranuclear compounds was previously observed for Pd complexes with similar thiosemicarbazone ligands [14].

### 3.3.3. Tetrakis (4-phenyl-1-benzaldehyde thiosemicarbazonato)tetrplatium (II) Pt₄(L⁴)₄

4-phenyl-1-benzaldehyde thiosemicarbazone reacts with ammonium tetrachloroplatinate (II) to form a tetrannuclear complex with slightly distorted square planar geometry (see Figure 3). The tridentate ligands are deprotonated at the NH group and coordinated through S, N, and C (aromatic ring). The fourth coordination site at each Pt atom is occupied by a sulfur atom of a neighboring ligand. In this way, a puckered eight-membered ring of alternating Pt and S atoms is formed as the core of the molecule. Each of the four Pt atoms belongs to two fused five-membered chelate rings: the C, N metallocycle and the N, S chelate moiety.

The Pt–Pt distances range from 3.43 Å to 3.84 Å. The Pt–S bonds form two distinct groups with significantly differing lengths: Pt–S_chelating 2.351(2) Å and Pt–S_bridging 2.298(2) Å (mean values). The coordination of the ligand to the Pt atoms leads to a lengthening of the S–C₁ bond (increased single-bond character) and a shortening of the neighboring N2–C₁ bond (increased double-bond character) compared with the free ligand.

In the crystal structure, two molecules of ethanol per formula unit of the tetrannuclear complex are included, which stabilize the crystal structure by hydrogen bonds. The O1 and O2 atoms of the ethanol molecules are bonded by hydrogen bonds to N3 atoms: N3–H···O1 [N3–H 0.86 Å, H···O1 2.07 Å, N3···O1 2.898 Å, N3–H···O1 162°] and N3–H’···O2 [N3–H’ 0.86 Å, H’···O2 2.14 Å, N3···O2 2.999 Å, N3–H’···O2 173°].

The Pt₄(L⁴)₄ complex with a cyano group in the meta/situation of the aromatic ring and the Pt₄(L⁴)₄ tetrannuclear complex with the phenyl group in the terminal amine of the ligand showed to be more cytotoxic (IC₅₀ = 0.45–3.67 and 0.07–12.46 μM, resp.) than the other Pt(L¹)₂ and Pt(L²)₂ complexes against all tested human tumor cell lines. These results indicate that the cytotoxic activity is enhanced when ligands are coordinated to the metal ion [18, 19, 32].

### 3.4. Antitumor evaluation

All ligands had a 50% inhibitory concentration (IC₅₀) > 40 μM against the used human tumor cell lines. As shown in Table 4, the palladium (II) and platinum (II) complexes were more cytotoxic (IC₅₀ = 0.08–12.46 μM) than their respective ligands. These results reveal that the cytotoxic activity increases dramatically when ligands are coordinated to the metal ion [18, 19, 32].

The Pt(L²)₂ complex with a cyano group in the meta/situation of the aromatic ring and the Pt₄(L⁴)₄ tetrannuclear complex with the phenyl group in the terminal amine of the ligand showed to be more cytotoxic (IC₅₀ = 0.45–3.67 and 0.07–12.46 μM, resp.) than the other Pt(L¹)₂ and Pt(L²)₂ complexes against all tested human tumor cell lines. These results indicate that the cytotoxic activity is enhanced when ligands are coordinated to the metal ion [18, 19, 32].
the cytotoxicity shown by the square planar copper (II) complexes with carboxamidrazide ligands (IC$_{50}$ = 3.0 $\mu$M), assayed in vitro against to the (MCF-7) human breast adenocarcinoma cell line [35], the Pt$_4$(L$^4$)$_4$ complex resulted to be more cytotoxic, while the Pd(L$^2$)$_2$ complex showed a similar inhibition concentration (IC$_{50}$ = 2.09 $\mu$M). In relation to other palladium (II) and platinum (II) complexes of the M(HL)Cl$_2$ type with phenyl acetaldehyde thiosemicarbazone ligands (IC$_{50}$ = 38 and 9 $\mu$M, resp.) assayed in the K562 human chronic myelogenous leukemia cell line [12], the Pd(L$^2$)$_2$ and Pt$_4$(L$^4$)$_4$ complexes showed higher cytotoxic activity. In addition, the Pd(L$^2$)$_2$ complex resulted to be more cytotoxic (IC$_{50}$ = 2.09 $\mu$M) than the palladium (II) and copper (II) complexes of the M(L)Cl type (IC$_{50}$ = 12.94 and 3.98, resp.) and the NiL$_2$ complex (IC$_{50}$ = 2.25 $\mu$M) with 1,2-naphthoquinone-1-thiosemicarbazide ligands, tested in vitro against to the MCF-7 human breast adenocarcinoma cell line [34].

In summary, we have prepared the palladium (II) bis-chelate complexes and the platinum (II) tetranuclear complex, Pd(L$^{1-3}$)$_2$, and Pt$_4$(L$^4$)$_4$, which were more cytotoxic on all the human tumor cell lines at low micromolar concentrations with respect to the free ligands. The crystal structure of Pd(L$^2$)$_2$ shows that the palladium atom has a square-planar geometry with two bidentate ligands having sulfur and nitrogen as donor atoms in transpositions. The complex Pt$_4$(L$^4$)$_4$ has a tetranuclear structure with four tridentate(C,N,S) ligands coordinated to four platinum atoms.

**Supplementary material**

Further details of the crystal structure determination are available on request from the Cambridge Crystallographic Data Center (CCDC, 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033; email: deposit@ccdc.cam.ac.uk), on quoting the depositing numbers CCDC 694972 for HL$^2$, 694973 for Pd(L$^2$)$_2$, and 694974 for Pt$_4$(L$^4$)$_4$, the names of the authors, and the journal citation.

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