Rapid synthesis and properties of segmented block copolymers based on monodisperse aromatic poly(N-methyl benzamide) and poly(propylene oxide)†

T. Mori,a S. Masukawa,a T. Kikkawa,b A. Fujimori,b A. Satoh,c K. Matsumoto,c M. Jikei,c Y. Oishia and Y. Shibasaki a,∗a

Monodisperse aromatic N-methyl benzamide-based molecules (repeat number: x = 5–10, molar mass distribution Mω/Mn = 1.02–1.04) connected with terephthalic acid at both sides (MABCan) were rapidly prepared by a step-wise condensation reaction using the reported one-pot dendrimer synthetic method. These were copolymerized with NH2-terminated poly(propylene oxide) (PPOy) using a condensation reagent at ambient temperature, and the relationship between the structure and the properties was evaluated. The copolymers, poly(MABCan-b-PPOy), had Mv values up to 12 500 g mol−1 (Mω/Mn = 1.6), with only a single Tg at −70 °C by DSC, assignable to the pure PPO domain. Only for the copolymer consisting of a monodisperse MABCan segment, did the dynamic mechanical analysis (DMA) show two Tg values at −70 and above 50 °C, indicating the clear phase segregation of the copolymers. The mechanical properties of the copolymer films were mostly dependent on the fraction of the monodisperse hard segment, and the tensile modulus and elongation at break varied from 3.3 to 32.5 MPa and 150 to 540%, respectively.

Introduction

A material prepared of a block copolymer having hard and soft segments (HSs and SSs) in the same chain shows unique properties due to the immiscible polymer segments, and is thus applicable in various fields such as membranes,1–7 photonics,8–12 lithography,13–16 nanocarriers,17,18 and thermoplastic elastomers (TPEs).19 TPEs are currently receiving much attention since they are recyclable and do not require additional chemical cross-linking reactions. The hard domains are strongly governed by physical packing, rather than chemical crosslinking, and are melted or softened at high temperatures. TPEs exhibit elastomeric properties that are comparable to those of chemically cross-linked rubbers, and like conventional thermoplastics, show good processability. Poly(ether-b-amide) multiblock copolymers (PEBA) are a class of important TPE materials due to the tuneable properties with variable polyether SSs and the highly thermostable aliphatic amide segment that realizes the reliable mechanical strength even at high temperatures.20–25 One of the most important application of PEBA can be the gas separation membrane,26–31 in which the crystalline amide HS increases the separation selectivity while the soft poly(ether) segment functions as high gas permeability. The properties of TPE materials are, of course, dependent on the types, length, and the molar mass dispersity of the individual segments. However, in generally speaking, polymers like PEBA are prepared via condensation polymerization, and thus, the lengths of the HSs are not well aligned (broad molar-mass distribution (Mω/Mn) of around 2). On the condensation polymers, Harrell first prepared well-defined non-hydrogen-bonding segmented poly(urethane)s via step-wise reactions.32 The mechanical properties of these polymers were strongly affected by the length and size distribution of the HSs not by SSs. He also found that an HS with one repeat unit only cocrystallized with those containing two repeat units and not with those containing more repeat units. Samuels and Wilkes studied poly(urethane)s using WAXS, where an increasing HS length resulted in higher crystallinity with a thicker lamellar structure.33 Hydrogen-bondable segmented poly(urethane)s based on 4,4′-methylenebis(phenyl isocyanate) (MDI), poly(tetramethylene oxide) (PTMO), and butanediol were then reported by Miller and his co-workers, and found that the polymers having a low Mω/Mn in the HSs showed a smaller degree of phase mixing, and thus, higher strain in the polymers.34 Despite the
controllable toughness and the elongation ability by choosing an isocyanate molecule as a proper HS unit, the flow temperature of these polymers is typically from 50 to 150 °C, and they tend to degrade on heating over the melting temperature. Therefore, thermally more stable amide groups can be favored, especially for the high temperature TPE usage. As a novel type PEBA, Gaymans and Van Hutten reported a series of hydrogen-bondable segmented copolymers based on a monodisperse aliphatic- and aromatic-mixed bis(ester amide) as a HS unit. For example, a HS composed of two terephthaloyl with one 1,4-butylenediamino centre, called (T4T) was copolymerized with aliphatic diols. The T4T units (4 indicates the C number in aliphatic chain) crystallized very fast and almost completely, resulting in a low diamide concentration in the polyether phase.45 The uniform T4T segments formed perfect and stable crystalline lamellae, resulting in a temperature-independent rubbery plateau. They also prepared monodisperse aliphatic- and aromatic-mixed diamide HSs such as T2T,46 T6T,47 and oligoamide HSs of T6T6T and T6T6T6T.48 The tunable T_m from 111 to 225 °C with the monodisperse HS structure be useful as high-performance TPE materials, and also the materials were applied to gas separation membrane with the excellent CO_2 permeability.49 Despite such many useful properties, the synthesis of the hydrogen-bondable HSs with uniform length requires multi-step tedious procedures, which may constrict the industrial application.

Aromatic polyamides, such as poly(p-phenylene terephthalamide) (Kevlar) and poly(benzamide), have attracted extensive attention due to their inherently extended rigid-rod structure, which is caused by the para-linked benzene rings with the double bond character of the trans-amide linkage.40–42 A structural study based on X-ray crystallography, with conformational energy calculations of hydrogen-bondable terephthalamide model compounds, was performed by Flory and his coworker.43 On the other hands, nonhydrogen-bondable N-methyl benzamide was prepared, and the stereo chemistry of the amide bond structure was determined to be cis by X-ray analysis, whereas that of benzaniilde was trans.44-47 They also prepared a 1,2-bis(N-benzyol-N-methylamino)benzene molecule, from which optically active crystals were obtained. The amide linkage was again cis, but the molecule had C2 symmetry.48 Therefore, the circular dichroic spectra of the crystal solutions show positive and negative Cotton effects. Yokozawa and his co-workers prepared non hydrogen-bondable aromatic poly(benzamide) bearing a chiral side chain via condensative chain polymerization (thus, the M_w/M_n was around 1.1) and investigated the structure in solution and in the solid state, where the polymer had a unique one-handed helical structure depending on the chirality.49 We have prepared a series of non hydrogen-bondable poly(N-alkyl benzamide) [PABA_n (n indicates the C number of C_{x}H_{2x+1} at the side group)] by conventional direct condensation polymerization of N-(N-alkylamino)benzoic acids and elucidated the structures and packing of the PABA_n polymers using IR spectroscopy, elemental analysis, DSC, WAXD, and SAXS. The prepared PABA_2 and PABA_3 (where n = 2–5) samples were insoluble in all organic solvents, and the tight packing was assignable to a monoclinic and orthorhombic structure between the main-chain dipolar interaction of the cis amide function. In contrast, while PABA_7 also possessed crystalline structures, it crystallized in hexagonal lattices through weak van der Waals interactions of the alkyl side chains.50 Various higher-ordered layered structures of PABA_n were also manufactured by the Langmuir Blodgett method, and the color of the interference structure was observed from the differences in the layer thickness of the stepwise multiparticle layers.51–53 For the synthesis of aromatic polyamide dendrimers, we have demonstrated a simple and highly efficient convergent approach without protection and deprotection steps, using thionyl chloride.54 Therefore, utilizing this reliable step-wise technique, we expected that aromatic PABA_n molecules with a “perfectly” defined molecular weight can be rapidly prepared and applied to the synthesis of a novel type aromatic-based-PEBA with poly(propylene oxide) (PPO). The nonhydrogen-bondable monodisperse aromatic amide HS would show higher thermostability compared to the aliphatic-based-PEBA material. Moreover, the use of N-methylated aromatic amide HS with a cis configuration would show higher solubility, so that the design and the synthesis of the HS with the desired length be easily attainable. Here, we demonstrate the synthesis and properties of a block copolymer based on oligo(4-N-methyl benzamide) (crystalline molecule) having an almost uniform length with PPO (typical amorphous polymer), using the facile dendrimer synthetic approach.

Results and discussion

Synthesis of aramid molecules

We performed a step-by-step reaction to construct the rigid-rod aramid molecules; the two carboxylic acids at the terminals of the compound were activated in situ with thionyl chloride, followed by condensation with the building block, 4-(methylamino)benzoic acid (MAB), according to the procedure for our previously reported dendrimer synthetic method,54 as shown in Scheme 1.

The highly pure targeted aramid compounds with the desired number of repeating units (MAB_{x-y}, where x = 5, 6, 7, 8, 9, and 10) were readily obtained in good yields (77–83%). The aramid molecules are crystalline compounds, and thus, were recrystallized with methanol/water mixed solvent, giving pale brown fine powders. Fig. 1 depicts the 1H NMR spectra of the MAB_{x-y} molecules. Along with the protons of the main chain, a–d, the protons at the terminal units, c(t), were observed at 7.85 ppm to have reasonable integration values, which supported the successful formation of the highly pure targeted HS, MAB_{x-y} molecule. The 1H NMR spectra also showed that the prepared molecules had cis geometry at the amide function.44–47 The N-methylated aramid molecules were then subjected to the GPC analysis. Fig. 2 depicts typical MAB_{x-y} molecules prepared (a) by the step-wise and (b) by the conventional polycondensation methods using terephthalic acid in one-step reaction in NMP. The former MAB_{5} molecule shows the reasonable M_w of 1071 g mol⁻¹ (theoretically 1500 g mol⁻¹) along with the quite narrow M_w/M_n value of 1.03. In contrast, the conventional one-step polycondensation polymer (bMAB_{5}) shows the M_w of 1470 g mol⁻¹ with the M_w/M_n of 1.45. However,
the MALDI-TOF mass spectrum of the monodisperse MAB_{5–5} molecule shows that the molecule is not a single one, but still the mixture of MAB_{x–x} compounds as shown in Fig. 3. Although the MAB_{5–5} by the step-wise reaction has some molar-mass distribution, we thought that this still can be useful for the copolymerization with PPO to elucidate the phase segregation and the mechanical properties under the effect of the molar-mass distribution.

To investigate the thermal behaviour, we then performed a DSC experiment. Fig. 4 depicts the DSC profiles of MAB_{x–x} molecules prepared via step-wise process. In a previous paper, we reported that PABA_{n} (n = 1–5) were all crystalline polymers, whose crystal structures were well assigned by WAXD experiments. However, the melting temperature (T_{m}) of poly(N-methyl benzamide) (poly(MAB)) was too high to measure with DSC, and only T_{g} was observed at 200 °C. The MAB_{x–x} molecules prepared in this study showed a clear thermal transition at around 170–178 °C, and thus, can be assigned as the T_{g} of the molecules. As expected, these values increased with the length of the MAB_{x–x} molecules.

**Synthesis of block copolymers of MAB_{x–x} with PPO_{y}**

As described in Introduction, the mechanical properties of TPEs were primarily affected by the length and size distribution of the HSs, not of the SSs. Therefore, we then copolymerized

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**Scheme 1** Synthesis of the N-methyl poly(p-benzamide)s (MAB_{x–x}) and the block copolymerization with PPO_{y}.

**Fig. 1** ¹H NMR spectra (CDCl₃) of the N-methyl poly(p-benzamide)s (MAB_{x–x}) and the molecular-weight data by GPC.
monodisperse MAB₅₋₅ with commercially available PPO₃₀, whose \(M_n/M_w\) was not well controlled, as shown in Scheme 1. The resultant mixture was poured into water to precipitate the polymer. The viscous pale brown coloured polymer was washed with hot water several times to remove low-molecular-weight impurities.

Fig. 5 shows the GPC profiles of (a) PPO₃₀ \((M_n = 1600 \text{ g mol}^{-1}, M_w/M_n = 1.30)\) and (b) poly(MAB₅₋₅-b-PPO₃₀) \((M_n = 9400 \text{ g mol}^{-1}, M_w/M_n = 1.79)\).

![Fig. 2 GPC profiles of (a) MAB₅₋₅ \((M_n = 1070 \text{ g mol}^{-1}, M_w/M_n = 1.03)\) and (b) bMAB₅₋₅ \((M_n = 1470 \text{ g mol}^{-1}, M_w/M_n = 1.45)\).](image1)

![Fig. 3 MALDI-TOF MS spectrum of MAB₅₋₅ \((M_n = 1110 \text{ g mol}^{-1}, M_w/M_n = 1.03)\).](image2)

![Fig. 4 DSC profiles of MAB₅₋ₓ samples (heating rate 20 °C min⁻¹).](image3)

![Fig. 5 GPC profiles of (a) PPO₃₀ \((M_n = 1600 \text{ g mol}^{-1}, M_w/M_n = 1.30)\) and (b) poly(MAB₅₋₅-b-PPO₃₀) \((M_n = 9400 \text{ g mol}^{-1}, M_w/M_n = 1.79)\).](image4)

of the terminal \(e\)' protons should not change to be 6.00, and in this case, the integration at around 3.5 ppm became 124.6, in which about 91.6 integration values could be assigned to PPO₃₀ segment \(f\), \(g\), and \(h\) protons, so the 124.6–91.6 = 33.0 integration is assignable to the \(e\) protons of the MAB₅₋₅ segment.

Table 1 summarizes the results of the copolymerization. All polymers were obtained in good yield with \(M_n\) values of up to 12 500 g mol⁻¹. The morphology of the resulting polymers ranged from viscous sticky materials to hard solids, depending on the weight fraction \(f_{HS}\) of the aramid molecules. The resulting polymers showed excellent solubility to typical
common organic solvents including methanol, acetone, chloroform, tetrahydrofuran, and polar aprotic solvent, N,N-dimethylformamide. Fig. 7 depicts the DSC profiles of the copolymers. All polymers have a clear thermal transition at –70 °C, assignable to the PPO segments, but no other transitions were observed over the temperature range measured.

Properties of block copolymers [poly(MAB₅–ₓ-b-PPOₓₙ)] and the influence of Mₘ/Mₙ

Fig. 8 depicts the DMA and the tan δ curves of poly(MAB₅–₅–b-PPO₆₇) (Mₘ = 8800 g mol⁻¹, Mₙ/Mₘ = 2.2) and poly(bMAB₅–₅–b-PPO₆₇) (Mₘ = 9400 g mol⁻¹, Mₙ/Mₘ = 1.8) prepared with monodisperse (Mₙ/Mₘ = 1.03) and polydisperse (Mₙ/Mₘ = 1.45) MAB₅–₅ molecules as HSs, respectively. Even though these finally obtained polymers have similar Mₙ and Mₘ/Mₘ values, the initial storage modulus E' of 4.84 GPa of poly(MAB₅–₅–b-PPO₆₇) consisting of monodisperse HS was higher than that of poly(bMAB₅–₅–b-PPO₆₇) film (E' = 4.15 GPa). The peak intensity of tan δ at –50.7 °C of poly(MAB₅–₅–b-PPO₆₇) consisting of monodisperse HS was obviously lower than the other sample, and also it has two distinct peaks at –50.7 and 47.9 °C in tan δ, which indicated the existence of more clear phase segregation of the HS block in the copolymer. Indeed, the appearance of poly(MAB₅–₅–b-PPO₆₇) consisting of monodisperse HS film was
opaque and tough, while that of the other one [poly(bMAB5–7–b-PPO8)] film] was transparent and soft. These results indicated that in order to make the clear phase-segregated structure in the polymer film consisting of non-hydrogen-bondable N-methylated aromatic poly(p-benzamide), the use of the monodisperse HS block is essential.

Fig. 9 depicts the relationship between tan δ and the measured temperature of the copolymer films. In contrast to the weak intensity of tan δ peak observed in Fig. 8, more clear peaks were observed above 50 °C for the poly(MABx–y–b-PPOy) (x ≥ 6) samples. For example, poly(MAB6–6–b-PPO30) shows a small and large peak at -30.5 and 53.0 °C, respectively, suggesting the stronger effect of the HS molecules. However, as seen in the DMA curve (Fig. S3†), there is no plateau region. For the polymer having a longer HS, poly(MAB8–8–b-PPO30), the tan δ peak was shifted into higher region. For the polymers having a longer SS, such as poly(MAB6–6–b-PPO40), poly(MAB8–6–b-PPO40), poly(MAB9–7–b-PPO40), and poly(MAB9–9–b-PPO40), all showed a clearer drop in tan δ at a lower temperature of ~ -60 °C. In particular, for a poly(MAB9–9–b-PPO40) film, the tan δ peak intensity at ~ -57.7 and 90.3 °C became similar. In the DMA curve, E’ started to decrease at around ~71.2 °C and the second drop was observed at from 68.1 °C having clear plateau region between them.

From the data obtained above, we can conclude that the phase-separation conditions of the films are as follows. (A) For the polymers having a short SS, there can be a phase-separated region consisting of SSs significantly contaminated with HSs and HSs containing some quantities of SSs. (B) For the polymers having long SSs with long HSs, the two phases can be separated, and the phase separation should be better for longer HS samples.

To investigate the polymer microstructures in the films, we then performed a WAXD experiment (Fig. 10). In our previous study, we prepared poly(MAB) via conventional polycondensation of MAB with a small amount of terephthalic acid, and the crystal structure was assigned to be monoclinic with parameters a = 7.7 Å, b = 5.7 Å, c = 7.9 Å, α = γ = 90°, and β = 67.5°. The WAXD profile of the polymer powder of poly(MAB5–7–b-PPO30) is depicted in Fig. 10(b), from which the crystal structure was assigned as triclinic with parameters a = 4.6 Å, b = 6.5 Å, c = 7.9 Å, α = γ = 86°, β = 67.5°, and γ = 100°. Note that the crystallized MAB5–5 from a methanol/water mixed solvent did not show a clear diffraction pattern in the WAXD experiment. Thus, we assumed that the HSs cannot align well alone. However, they are placed in a regular manner in the solid-state block copolymer, where the PPO segment (Tg = -68 °C) acted as the solvent even at ambient temperature. In contrast, for the poly(MABx–y–b-PPOy), whose x values range from 6 to 9, all samples showed only amorphous hollows in the profiles. However, the packing was recovered for the poly(MAB10–10–b-PPO30) sample, returning to the crystal structure to be original monoclinic lattice. This can be explained by the polymers having relatively short HSs being unable to crystallize into the stable monoclinic lattice because of the inadequate length of the HSs.

In order to evaluate the phase segregated structure in more detail, we measured AFM of the poly(MAB5–5–b-PPO30), poly(MAB8–8–b-PPO30) films (Fig. S11†). Unfortunately, not very clear phase-separation was observed in the images. The former sample seems the rough surface with large phase separation in topography, and the other one shows more precise separation.

In the prepared films, the phase separation mainly occurred within the films, and was not appeared on the film surface. Since the interaction between the HS units in our system (non-hydrogen bondable) is rather weak compared with those of hydrogen-bondable PEBA, the existence of the longer HS must
be critical to form clear phase segregation structure. This resulted in the mechanical properties of the films. Namely, the longer HS unit in the copolymer is the better for the mechanical properties. In other words, the mechanical properties can be slowly changed, and thus we have more opportunity to tune the properties.

Fig. 11 depicts the stress–strain curves of the block copolymer films, and the mechanical properties are summarized in Table 2. Polymer films having long SSs showed much better ductility, and thus, the $E_b$ values were over 500%. The tensile modulus ($T_s$) of the films depended on the fraction and length of the HSs; $T_s$ values were found to be in the following order: poly(MAB$_{b-a}$-b-PPO$_{30}$) ($f_{HS} = 55.3\%$) $>$ poly(MAB$_{b-a}$-b-PPO$_{55}$) ($f_{HS} = 48.8\%$) $>$ poly(MAB$_{b-a}$-b-PPO$_{67}$) ($f_{HS} = 36.5\%$) $>$ poly(MAB$_{b-a}$-b-PPO$_{30}$) ($f_{HS} = 33.8\%$) $>$ poly(MAB$_{b-a}$-b-PPO$_{55}$) ($f_{HS} = 30.7\%$). The mechanical strength data in Table 2 are all superior to those of PEBA material could allow us a fine tuning of the film properties, the results in Fig. 12 supported this assumption. In addition, the HS of the poly(MAB$_{b-a}$-b-PPO$_{y}$) shows an excellent thermal stability of 5% weight loss temperature in N$_2$ $T_d5 = 408$ °C, while aliphatic amide, nylon 6 shows the value of at 320 °C. The ease of preparation and higher thermostability can be useful of the application of this monodisperse N-methylated aromatic poly(p-benzamide) compound.

### Experimental

#### Materials

4-(N-Methylamino)benzoic acid (MBA) was purchased from Wako Pure Chemical Industries, Ltd., and purified by sublimation. Poly(propylene glycol) bis(2-aminopropyl ether)s (PPO$_{y}$) were purchased from Sigma-Aldrich Co. LLC., and the number-average molecular weight ($M_n$) and degree of polymerization ($y$) were estimated by $^1$H NMR measurement. Two samples of PPO$_y$ ($M_n = 1875$ g mol$^{-1}$, $y = 30$ and $M_n = 4025$ g mol$^{-1}$, $y = 67$) were used. N-Methylpyrrolidone (NMP) was dried over calcium hydride (CaH$_2$) and distilled under reduced pressure. Thionyl chloride was distilled prior to use. Diphényl (2,3-dihydro-2-thioxo-3-benozaxo)phosphonate (DBOP) was prepared according to the literature. The other reagents and solvents were used without further purification.

#### Synthesis of low molar mass distribution 4-(N-methyl) benzamide molecules (MAB$_{b-a}$)

**Typical procedure (MAB$_{b-a}$).** Terephthalic acid chloride (0.406 g, 2.00 mmol) and NMP (2 mL) were added to a 30 mL two-necked flask equipped with a three-way stopcock and a nitrogen inlet, at 20 °C. The building block, MBA (0.604 g, 4.00 mmol), was added to the mixture, and the reaction was performed at 20 °C for 30 min to give the condensation product MAB$_{b-a}$. Thionyl chloride (0.604 g, 4.00 mmol) was then added and stirred for 30 min at that temperature (activation). Again, MBA (0.604 g, 4.00 mmol) was added to the mixture to perform
the condensation reaction, resulting in the condensation product MAB$_{5-7}$. This activation and condensation reaction process was repeated another three times, and the resulting solution was poured into distilled water to precipitate the product. The collected light brown powder was recrystallized with methanol/water mixed solvent and dried at 100 °C for 8 h to afford the title compound.

MAB$_{5-7}$: Yield: 1.24 g (82.8%). $^1$H NMR (CDCl$_3$, 400 MHz, ppm): δ 3.31-3.37 (m, 30H, -CH$_3$), 6.71-7.36 (m, 40H, -ArH), 7.85 (m, 4H, -ArH). $^{13}$C NMR (CDCl$_3$, 100 MHz, ppm): δ 35.5, 120.5, 126.2, 127.5, 129.1, 130.3, 136.9, 141.6, 143.3, 163.7, 172.0. IR (KBr, cm$^{-1}$): ν 3466 (COOH), 3048 (ArH), 2938 (CH$_3$), 1650 (C=O), 1602 (C=C), 1534 (amide C-N), 1102 (CH$_2$-N). Anal. calcd for (C$_{168}$H$_{146}$N$_2$O$_{24}$): C, 71.32; H, 5.20; N, 9.90%.

3.25 mmol of the title compound was formed on a Bruker AC-400P spectrometer at 400 MHz for $^1$H NMR measurements and a Rigaku R-axis Rapid diffractometer equipped with a graphite monochromator, with CuK$_\alpha$ radiation. The crystal data of the title compound are as follows.

Yield: 78.5%. $^1$H NMR (CDCl$_3$, 400 MHz, ppm): δ 3.28-3.69 (m, 48H, -CH$_3$), 6.70-7.40 (m, 64H, -ArH), 7.87 (m, 4H, -ArH). $^{13}$C NMR (CDCl$_3$, 100 MHz, ppm): δ 35.9, 120.5, 126.2, 126.7, 127.1, 129.2, 130.3, 136.8, 141.6, 142.8, 162.5, 171.8.

MAB$_{5-9}$: Yield: 76.7%. $^1$H NMR (CDCl$_3$, 400 MHz, ppm): δ 3.25-3.72 (m, 54H, -CH$_3$), 6.66-7.38 (m, 72H, -ArH), 7.86 (m, 4H, -ArH). $^{13}$C NMR (CDCl$_3$, 100 MHz, ppm): δ 35.7, 121.5, 126.5, 127.8, 132.4, 138.1, 143.6, 145.6, 163.8, 171.1, 175.4. IR (KBr, cm$^{-1}$): ν 3478 (COOH), 3051 (ArH), 2940 (CH$_3$), 1645 (C=O), 1602 (C=C), 1533 (amide C-N), 1102 (CH$_2$-N).

MAB$_{10-10}$: Yield: 77.5%. $^1$H NMR (CDCl$_3$, 400 MHz, ppm): δ 3.22-3.73 (m, 60H, -CH$_3$), 6.61-7.45 (m, 80H, -ArH), 7.75 (m, 4H, -ArH). $^{13}$C NMR (CDCl$_3$, 100 MHz, ppm): δ 34.7, 120.3, 127.2, 128.5, 130.1, 132.4, 136.0, 137.1, 141.6, 143.1, 163.8, 172.1, 175.5. IR (KBr, cm$^{-1}$): ν 3498 (COOH), 3052 (ArH), 2935 (CH$_3$), 1647 (C=O), 1602 (C=C), 1531 (amide C-N), 1102 (CH$_2$-N). Anal. calcd for (C$_{180}$H$_{152}$N$_2$O$_{32}$): C, 71.32; H, 5.20; N, 9.90%. Found: C, 69.51; H, 5.22; N, 9.36%.

Multiblock copolymerization of MAB$_{5-7}$ with PPO$_y$ (poly(MAB$_{5-7}$-b-PPO$_y$))

The typical procedure for poly(MAB$_{5-7}$-b-PPO$_y$) was as follows. MAB$_{5-7}$ (0.375 g, 0.250 mmol), PPO$_y$ (0.469 g, 0.250 mmol), triethylamine (0.060 g, 0.600 mmol), and NMP (2 mL) were added to a 30 mL two-necked flask equipped with a three-way stopcock and a nitrogen inlet at 20 °C. DBOP (0.230 g, 0.600 mmol) was added to initiate the polymerization. After stirring at 20 °C for 6 h, the solution was poured into 3 wt% NaHCO$_3$ (aq.). The precipitate was collected and dried under vacuum at 100 °C for 8 h to afford a viscous brown polymeric compound. Yield: 84.2%. $^1$H NMR (CDCl$_3$, 400 MHz, ppm): δ 1.12-1.26 (m, -CH$_3$), 1.27 (-CH$_3$), 3.35-3.72 (m, -CH$_3$-CH$_3$), 6.75-7.21 (m, -ArH). $^{13}$C NMR (CDCl$_3$, 100 MHz, ppm): δ 17.6, 73.0, 73.5, 75.4, 120.4, 123.6, 126.1, 129.3, 129.9. IR (KBr, cm$^{-1}$): ν 3051 (N-H), 2923 (aliphatic C-H), 1645 (C=O), 1368 (Ar-N), 1175 (CH$_3$-N).

Measurements

Fourier transform IR spectra were measured with a Jasco IR-5500 (Jasco Co., Ltd.) by transmission absorbance spectroscopy (KBr tablet method). The molecular mass of polymers [number- and weight-average (Mn and Mw)] was determined with a Tosoh HLC-8120 GPC using a consecutive polystyrene gel column (TSK-GEL GMHHR-M and GMHHR-N) at 40 °C, eluted with NMP containing 0.01 mol L$^{-1}$ LiBr at a flow rate of 1.0 mL min$^{-1}$. Nuclear magnetic resonance (NMR) was performed on a Bruker AC-400P spectrometer at 400 MHz for $^1$H and 100 MHz for $^{13}$C measurements. Deuterated chloroform (CDCl$_3$) was used as a solvent with tetramethylsilane as an internal reference. Thermal analyses were performed on a Seiko thermal analyzer (SAC 5200 system) at a heating rate of 10°C min$^{-1}$ for thermogravimetric analysis (TGA) by TG/DTA 320) under air or nitrogen. Differential scanning calorimetry (DSC) was performed on a Shimadzu DSC-60 at a heating rate of 20°C min$^{-1}$ under nitrogen. Dynamic mechanical analysis (DMA) was performed on a Shimadzu DMS-210 at a heating rate of 20°C min$^{-1}$ under nitrogen. Polymer films were cut into sizes of 20 × 9 × 0.12 mm$^3$, and a load was applied to the films with a strain amplitude of 5 μm and a frequency of 1 Hz. The modes of several polymers in the crystalline phase were examined by wide-angle X-ray powder diffraction (WAXD) measurements using a Rigaku R-axis Rapid diffractometer equipped with a graphite monochromator, with CuK$_\alpha$ radiation.
radiation at 45 kV and 200 mA. Tapping mode atomic force microscopy (AFM) measurements were performed with a NanoNavi S-image SPM system (Hitachi High-Tech Sci. Co., Tokyo, Japan) using a cantilever with a force constant of approximately 15 N m⁻¹ and a resonance frequency of 120 kHz. The scan size and a tip velocity of the AFM measurements are 10 × 10 μm² and 200 nm s⁻¹, respectively. Matrix-assisted laser desorption/ionization time-of-flight mass spectra (MALDI-TOF MS) were recorded on an AXIMA Confidian (SHIMADZU Corp., Kyoto, Japan). Dithranol was used as the matrix.

Conclusions

In this study, we prepared N-methyl benzamide-based monodisperse molecules (repeat number: 5–10 × 2) MABₓ₋ₓ for copolymerization with PPOₓ, characterized the structures, and evaluated the properties. The MABₓ₋ₓ molecules prepared by an authentic step-wise condensation reaction, using the reliable dendrimer synthetic method, exhibited quite narrow molar-mass distribution below 1.04 and the Tₓ was 170–178 °C, depending on the length of the repeating unit. These rigid-rod molecules were successfully copolymerized with PPO, giving the copolymer poly(MABₓ₋ₓ-b-PPOₓ) in good yield with Mₓ up to 12,500 g mol⁻¹ (Mₓ/Mₓ = 1.6). As a control, the polydisperse HS (bMAB₅₋₅, Mₓ 1470, Mₓ/Mₓ = 1.45) was prepared, and copolymerized with PPO as well to form poly(bMAB₅₋₅-b-PPOₓ) (Mₓ 9400 g mol⁻¹, Mₓ/Mₓ = 1.8). Even though these finally obtained block copolymers have similar Mₓ and Mₓ/Mₓ values, the initial storage modulus E', the peak intensity of tan δ at −50.7 °C, the tan δ peak at 47.9 °C, and the appearance showed large differences. Thus, for the formation of the phase-segregated structure with non-hydrogen-bondable block copolymer system, the use of the monodisperse HS block is essential. The copolymers using monodisperse HS blocks showed a single Tₓ at −60 °C, assignable to the pure PPO domain by DSC. DMA analysis of the copolymer films clearly showed two Tₓ values at −60 and above 60 °C, indicating the phase segregation of the copolymers. The mechanical properties of the copolymer films were mainly dependent on the fraction of the hard segment, and the tensile modulus and elongation at break varied from 3.3 to 28.0 MPa and 150 to 540%, respectively.

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